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Environmental applications of magnetic nanohybrid materials

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The distinctive characteristics and versatile functionalities of magnetic nanoparticles (MNPs) position them as promising candidates for numerous environmental applications, attributed to their structural and physiochemical properties. Magnetic nanoparticles can integrate organic polymers, carbon-based materials, and metal oxides to form multifunctional composites that effectively adsorb and degrade pollutants. This review article discusses the magnetic nanomaterials and the techniques employed to functionalize and tailor these materials and their application for wastewater treatment, cleanup of oil spills, and photocatalytic degradation of pollutants of emerging concern. Focusing on their regeneration capabilities, catalytic performance, and adsorption efficiency, the article analyzes how magnetic nanohybrids engage with pollutants remediation. Additionally, it evaluates these materials' stability, implications, and recyclability to assess their practical applicability in real-world environmental scenarios. By presenting a detailed overview of the recent advancements, challenges, and prospects, this work aims to assist researchers in developing and enhancing magnetic nanohybrids as a remediation technology to advance sustainable environmental remediation strategies. The magnetic nanoparticles discussed in this document exhibit remarkable efficacy in contaminant removal due to their magnetic retrievability, facilitating easier recovery and reuse. These advantages not only enhance their practical applications but also align with principles of sustainability and resource efficiency. The innovative integration of advanced magnetic materials with green chemistry principles outlined in this review significantly enhances environmental remediation initiatives. This review adopts a holistic approach by emphasizing the benefits while addressing challenges, including instability, elevated costs, quality and functionality problems, concerns regarding secondary pollution, and issues related to reusability.

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1. Introduction

The global population has been increasing consistently with urban areas experiencing more significant population growth, attributed to enhanced healthcare, educational opportunities, and job accessibility, with industrialization playing a key role in this urbanization process.¹ As the population increases, the need for goods and services rises, leading to industrial growth and additional factories.² The need to support larger populations and urban centers drives technological innovations in manufacturing, energy production and transportation. Many

industrial and non-industrial sources contribute to water pollution, which harms human health,³ aquatic ecosystems, and biodiversity.^{4,5} The chemical manufacturing, mining, oil, and textile sectors discharge toxic materials into water bodies through untreated wastewater, unintentional spills, and poor waste disposal.⁶ These substances include organic pollutants⁷ like solvents, dyes, hydrocarbons, and heavy metals⁸ such as lead, mercury, and arsenic. Moreover, agricultural runoff comprising fertilizers, pesticides, and animal manure causes eutrophication, toxic algal blooms, and pathogens⁹ contamination, in addition to deforestation and soil erosion.¹⁰ Furthermore, mining results in metal leaching and acid mine drainage,¹¹ and hydraulic fracturing in oil and gas operations puts groundwater at risk of contamination.¹² Marine environments face threats from plastic waste, oil spills, and the atmospheric deposition of pollutants such as nitrogen oxides.¹³ To protect water quality and ecosystem health, integrated legislation, sustainable practices, and efficient wastewater treatment technologies must be implemented.¹⁴ Managing the impact of urbanization and industrialization is crucial for environmental sustainability.

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In the realm of water and wastewater treatment, nanotechnology stands out as a highly effective and promising avenue for advancement.¹⁵ The utilization of magnetic nanoparticles (MNPs) has shown significant benefits in wastewater remediation, owing to their exceptional optical, electrical, and magnetic properties, as well as their high adsorption capacities, mobility, reactivity, and catalytic abilities. Furthermore, the features and applications of MNPs can be tailored through the application of magnetic fields, enhancing their efficacy in treatment processes. MNPs find extensive potential applications in catalysis, biomedicine, tissue targeting, colloidal photonic crystals, MRI, MPI, microfluidics, data storage, environmental remediation, and more.^{16,17} MNPs of many kinds, including iron oxides (Fe_3O_4), spinel ferrite (CoFe_2O_4), bio-metallic, carbon-based, and others are employed for wastewater treatment. With the potential for nearly 100% *in situ* treatment, MNPs in waste treatment improve efficiency, encourage reuse, and save money, time, and effort.¹⁸

Physical and chemical methods are essential for wastewater treatment. Physical methods involve mechanical techniques such as screening, filtration, and sedimentation to eliminate suspended solids and debris. Additionally, flotation and centrifugation are employed to separate materials based on density. Chemical methods utilize various substances to precipitate, neutralize, or oxidize contaminants. Coagulation and flocculation processes introduce chemicals that promote the aggregation of smaller particles into larger ones, improving their removal through sedimentation or filtration.¹⁹ Disinfection techniques such as chlorination and ozone treatment eliminate harmful pathogens. Chemical oxidation can break down complex organic pollutants into simpler compounds, often necessitating pH adjustments to optimize these processes. While these methods are effective, they can be costly, require careful handling of hazardous materials, and generate chemical sludge that must be disposed appropriately. Nevertheless, the implementation of both physical and chemical treatments is crucial for the effective removal of diverse pollutants, particularly in industrial wastewater.²⁰ Wastewater treatment utilizing physical and chemical techniques typically involves considerable expenses and substantial energy consumption, especially when dealing with large volumes. Applying these methods often requires the use of chemicals, which can pose environmental and safety risks, in addition to generating chemical sludge that requires proper disposal. Moreover, these methods may not effectively remove certain dissolved pollutants or biodegradable organic materials, leading to the need for further treatment processes.²¹

Membrane technologies are advanced methods employed in wastewater treatment, leveraging selective barriers to efficiently separate contaminants from water. Commonly utilized membrane processes include nanofiltration, microfiltration, ultrafiltration, and reverse osmosis, each characterized by specific pore sizes and varying effectiveness in pollutant removal. Microfiltration and ultrafiltration are adept at removing suspended solids, bacteria, and larger particles, while nanofiltration and reverse osmosis target smaller dissolved contaminants, such as salts, heavy metals, and organic

molecules.²² Nonetheless, membrane technologies can incur significant expenses owing to the energy requirements for sustaining pressure across the membrane and the regular necessity for cleaning or replacing membranes due to fouling. Membrane fouling diminishes efficiency and requires regular cleaning or replacement, leading to increased downtime and operational costs. Furthermore, this technology might only be financially viable for extensive applications with considerable investment.²³

Among the advanced oxidation processes (AOPs), photocatalysis stands out due to its effectiveness and versatility in wastewater treatment, surpassing methods like ozone oxidation, Fenton's reagent, and $\text{UV}/\text{H}_2\text{O}_2$. A notable advantage of photocatalysis is its ability to operate under ambient conditions, which negates the necessity for high temperatures or pressures, thus streamlining operations and reducing energy costs.²⁴ Unlike chemical methods like Fenton's reagent that rely on iron salts and hydrogen peroxide, photocatalysis does not require potentially harmful chemical additives, resulting in fewer secondary pollutants and a reduced need for subsequent neutralization processes.²⁵ The issue of electron-hole recombination in photocatalysis, which can restrict the efficiency of the process, is currently under investigation. Innovations like heterojunction photocatalysts and plasmonic materials contribute to the reduction of recombination and enhance overall performance.²⁶

Magnetic nanoparticles constitute a distinctive category of nanomaterials utilized in photocatalytic wastewater treatment. They provide specific advantages and functionalities that differentiate them from traditional photocatalysts such as TiO_2 , ZnO , CdS , and graphene-based materials.²⁷ Magnetic nanoparticles, primarily consisting of iron oxides such as Fe_3O_4 or $\gamma\text{-Fe}_2\text{O}_3$, demonstrate significant magnetic properties that facilitate their separation from treated water by applying an external magnetic field. This feature enhances the recovery and reuse of photocatalysts, tackling a considerable challenge of conventional photocatalysts, which often face time-consuming, costly, or inefficient separation from treated effluent. Materials such as TiO_2 and ZnO demonstrate high effectiveness; however, their recovery usually necessitates filtration or sedimentation, potentially elevating operational costs. Magnetic nanoparticles can be rapidly and effectively extracted using magnetic fields, enhancing the overall process's economy and scalability, particularly in large-scale wastewater treatment applications.²⁸ MNPs frequently act as a support or core material for other semiconductors, improving their photocatalytic activity. A photocatalyst such as TiO_2 or ZnO can be applied to the surface of MNPs, resulting in a composite that utilizes the photocatalytic characteristics of the semiconductor while also providing the benefit of magnetic separation. This hybrid method integrates the advantageous characteristics of both materials: the elevated photocatalytic efficiency of semiconductors and the straightforward recovery of magnetic particles.²⁹ MNPs contribute to the reduction of electron-hole recombination in these systems, a process where electrons and holes recombine, reducing the efficiency of the photocatalyst, as the iron oxide core facilitates electron transfer, enhancing



overall photocatalytic performance. This contrasts with conventional semiconductor materials such as TiO_2 , ZnO , or CdS , where recombination remains a significant challenge restricting reactive species' production. TiO_2 is susceptible to electron-hole recombination; however, incorporating magnetic nanoparticles as a support material can enhance charge separation and increase the efficiency of the photocatalyst.³⁰ MNPs display distinctive surface chemistry that can be utilized for additional functionalization. Modification or doping with metals, oxides, or organic molecules can enhance catalytic properties, improve light absorption, or introduce specific functionalities for pollutant binding. Combining Fe_3O_4 nanoparticles with silver or gold nanoparticles enhances light absorption. It broadens their activity into the visible spectrum, thereby overcoming a significant limitation of conventional photocatalysts such as TiO_2 , which predominantly function under UV light.³¹

Magnetic nanoparticles enhance process efficiency by facilitating improved charge separation and easier recovery post-treatment rather than directly influencing photocatalytic reactions. This renders them more complementary, in contrast to traditional photocatalysts designed to degrade pollutants independently.³² The interaction between MNPs and traditional photocatalysts presents a promising approach for enhancing wastewater treatment solutions' efficiency, sustainability, and cost-effectiveness.¹⁸ This review explores MNPs, focusing on their structural characteristics and photocatalytic capabilities, while highlighting their unique benefits in addressing different contaminants in wastewater. It examines surface modification and composite strategies employed to enhance the photocatalytic efficiency of magnetic nanoparticles, focusing on prevalent issues such as electron-hole recombination and material stability. This document also evaluates the versatility of these nanoparticles in addressing various pollutants, organic dyes, heavy metals, synthetic drugs, and oil spills across diverse environmental conditions. The review offers insights into recent advancements and applications, aiming to facilitate future innovations using magnetic nanoparticles for scalable and environmentally friendly wastewater treatment technologies. This is the first report that presents the regeneration capabilities, the photocatalytic performance, and adsorption efficiency of magnetic nanoparticles as a remediation technology to advance the renewable and sustainable water treatment strategies. It offers insights into recent advancements and applications, aiming to facilitate future innovations using magnetic nanoparticles for scalable and environmentally friendly wastewater treatment technologies.

2. Synthesis techniques for MNPs

MNPs are synthesized by various methods. In general, the synthesis of MNPs can be broadly classified into three categories (a) physical, (b) chemical, and (c) biological or microbial methods.³³ Physical methods are based on the top-down strategy, *i.e.*, synthesis starts from bulk material and depletes to generate NPs. Chemical and biological methods follow a bottom-up or constructive approach, where the atoms/

molecules assemble into different sizes of NPs.³⁴ Fig. 1 shows different synthesis methods of MNPs.

2.1 Top-down synthesis

This approach employs a destructive method to reduce bulk materials into tiny, nanoscale particles through various physical and chemical procedures. It can be demonstrated by destructive techniques such as physical vapor deposition, grinding, milling, or several others.³⁵

2.1.1 Mechanical method/ball milling method. This method of producing nanoparticles from bulk is cost-effective. It entails breaking down bulk materials directly into nanostructures. During mechanical milling, the bulk material receives the rollers which reduces grain size. Alloys of various metals are produced using this method.³⁶ In a study by Bououdina *et al.*, nanocrystalline magnesium ferrite was synthesized through ball milling of $\alpha\text{-Fe}_2\text{O}_3$ and MgO in powdered form and strengthened at 700–900 °C. The synthesized magnetic nanopowders were suitable for removing heavy metals from wastewater.³⁷ In another study, carbonaceous magnetic adsorbents were synthesized through ball milling, a combination of activated carbon and Fe_3O_4 nanoparticles. The authors found that the ball-milled magnetic adsorbents are environmentally friendly to synthesize, have high efficiency, low cost, and offer convenience in operation.³⁸

2.1.2 Thermal decomposition method. Thermal decomposition is an endothermic process where heat causes chemical decomposition and breaks the chemical link in a compound. The nanoparticles result from the decomposition of metal at a particular decomposition temperature at which an element undergoes chemical breakdown.³⁹ Hyeon and co-workers have also used a similar thermal decomposition approach to prepare monodisperse iron oxide nanoparticles. They used nontoxic and inexpensive iron(III) chloride and sodium oleate to generate an iron oleate complex *in situ*, which was then decomposed at temperatures between 240 and 320 °C in different solvents, such as 1-hexadecene, octyl ether, 1-octadecene, 1-eicosene, or trioctylamine.⁴⁰ The thermal decomposition of zero-valent metal precursor $\text{Fe}(\text{CO})_5$ leads to the formation of metal NPs, but if oxidation occurs, it may form high-quality iron oxide MNPs. On the other hand, if the decomposition of precursors occurs with cationic metal centers, it can directly form metal oxide NPs.⁴¹ This method was successfully generalized for synthesizing other magnetic nanocrystals, such as MnO , Co_3O_4 , and NiO .⁴²

2.1.3 Laser ablation. This method involves laser irradiation to reduce particle size to nanoscale. A solid surface is exposed to a laser beam, creating a low-flux plasma plume that is eventually sublimated or evaporated to create NPs. The materials undergo plasma conversion at higher flux. Due to its high yield, improved long-term stability, and improved particle size and shape control, this technique is a better substitute for traditional chemical procedures.⁴³ Laser ablation has also synthesized magnetic nanoparticles from transition metals and metal alloys.⁴⁴ Using the laser ablation method, Sakiyama *et al.* fabricated Ni particles of a selected size coated with a NiO shell.



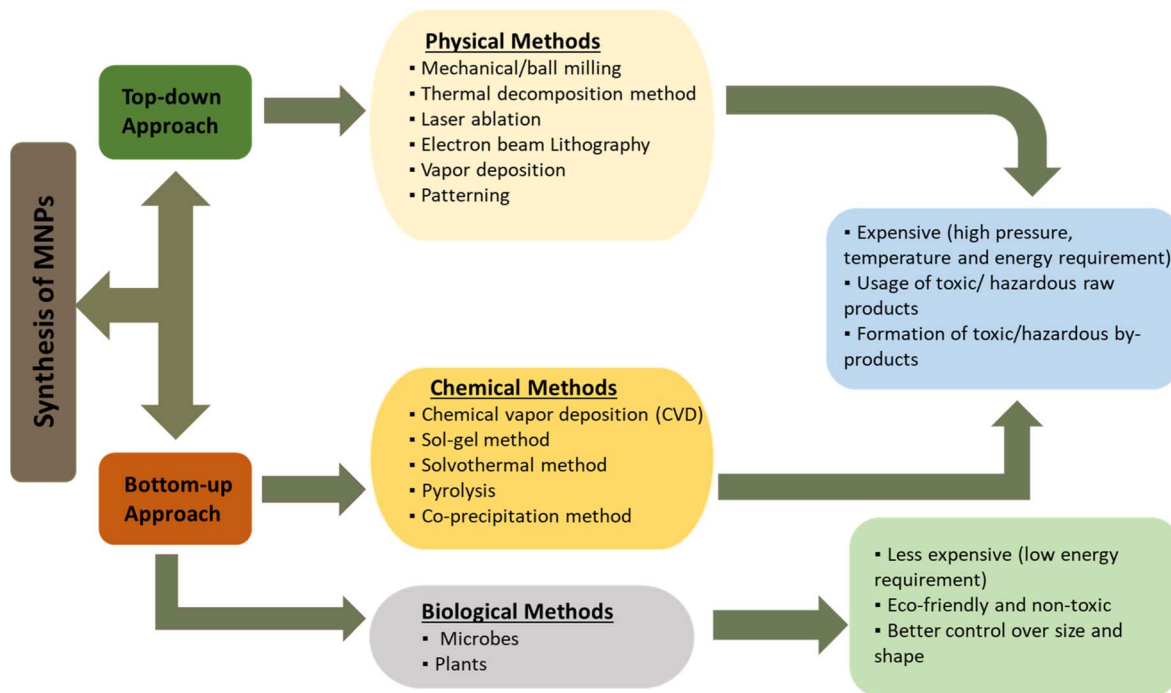


Fig. 1 Classification of synthesis methods for MNPs and their key features.

The oxidation of Ni nanoparticles successfully generated a Ni/NiO core-shell structure.⁴⁵ Shinde *et al.* also synthesized iron oxides and strontium ferrite with laser ablation and evaluated the magnetic characteristics.⁴⁶

2.1.4 Electron beam lithography. In the electron beam lithography process, iron particles are transformed into iron oxides (Fe_3O_4) using an electron beam, or e-beam. This process produces nanoscale iron oxide nanoparticles (NPs) by using specified e-beam emission across an iron particle-film surface. This approach is acknowledged as being inexpensive, versatile, and capable of producing stable NPs.^{18,47}

2.1.5 Vapour deposition and patterning. This technique can produce magnetic particles by filling holes in a template and creating a continuous film. Vapour deposition methods (laser ablation, sputtering, and evaporation) and electrodeposition methods are used to deposit the material that makes up the particles. This process has demonstrated the ability to generate high-quality NPs, lower alloying temperatures, and potentially be used on a wide scale.^{18,48}

2.2 Bottom-up synthesis

Bottom-up synthesis, the constructive method, uses chemical and biological processes to construct MNPs from relatively more straightforward substances such as atoms, clusters, and molecules. This process initially creates the nanostructured building blocks of nanoparticles, which are subsequently put together to produce the finished product (MNPs). The final product offers a more uniform shape (physical attributes), size, and chemical composition. Chemical vapor deposition, sol-gel, and pyrolysis are examples of bottom-up synthesis.⁴⁹

2.2.1 Chemical vapor deposition (CVD). In this method, the thin film of the gaseous reactant and a mixture of additional gas molecules are deposited onto a substrate that promotes substrate superheating. The ions are reduced when the combined gases and substrate come into contact. The outcome of this reaction is typically a film from which the MNPs must be scrapped. The production of intricate, robust, homogeneous, and highly pure NPs is a prominent feature of the CVD method. This technique is now highly significant in the semiconductor and electronics industries. However, the production of extremely hazardous gases as byproducts and the demand for specialized equipment are the main drawbacks of this method.⁵⁰ Niu *et al.* synthesized well-defined core/shell type single crystalline $\text{Fe}_7\text{S}_8/\text{Fe}_3\text{O}_4$ coated α -Fe hybrids ($\text{Fe}_7\text{S}_8/\text{Fe}_3\text{O}_4@\text{Fe}$) with vacuum CVD technique. The CVD process triggers the conversion of naturally formed Fe_3O_4 layer on the surface of commercial Fe nanoparticles from amorphous into a single crystalline phase.⁵¹ Another study used a room-temperature method to synthesize Ni, Ni/Zn, and Zn ferrite nanoparticles. Various shapes for the produced carbon structures have been observed using ferrite nanoparticles as a growth catalyst of the CVD process to prepare crystalline MNPs.⁵²

2.2.2 Sol-gel method. The sol-gel method is one of the most favored straightforward techniques for synthesizing magnetic nanoparticles. This method combines two words: sol (a colloid made up of solid suspended particles in continuous liquid) and gel (a solid macromolecule dissolved in solvent). An appropriate chemical solution, such as metal oxides and chlorides, is a precursor. The precursor is distributed throughout the host liquid *via* various techniques, including shaking, stirring, and sonication. The resulting solution comprises a liquid



and solid phase, separated using multiple methods like centrifugation, sedimentation, and filtration to extract the nanoparticles. It is a procedure that turns sol into gel using hydrolysis and condensation.⁵³ This technique works very well for creating silica-coated and iron oxide MNPs. It makes it possible to produce MNPs in large quantities with controlled sizes and distinct shapes.⁵⁴ Mohammad *et al.* reported a novel visible-light-driven photocatalyst of $\text{TiO}_2/\text{Au}/\text{SWCNT}$ nanohybrid prepared by a sol-gel approach and successfully assessed for the photocatalytic degradation of the organic methylene blue. $\text{TiO}_2/\text{Au}/\text{SWCNT}$ nanohybrid was obtained by doping TiO_2 lattice with Au and supporting with modified single-walled carbon nanotubes (SWCNTs).⁵⁵ In another study, barium ferrite spinel (BaFe_2O_4) was synthesized through the sol-gel method, and its nanohybrids with copolymers of carbazole and pyrrole were successfully synthesized *via* mechano-chemical mixing. The spectral, thermal, morphological, and magnetic properties of the synthesized nano-magnetic materials were analyzed *via* thermal gravimetric analysis (TGA), transmission electron microscopy (TEM), and vibrating sample magnetometer (VSM) techniques.⁵⁶

2.2.3 Solvothermal method. MNPs are created using the solvothermal process, also known as hydrothermal, in aqueous fluids at temperature above 200 °C and extremely high pressures (>13 790 kPa) using autoclaves. It primarily entails the rapid nucleation and growth of newly created MNPs, resulting in pure particles with controlled morphologies. During this process, MNPs are prepared by hydrolysis and oxidation reactions. The geometry of NPs can be better controlled using this technique.⁵⁷

The synthesis of Fe_3O_4 nanoparticles (NPs) with a size of 15 nm and a spherical shape is one of the impressive achievements obtained using this technique, which has been effectively applied in tumor MRI applications.⁵⁴ Modern nanomaterials that are challenging to synthesize through traditional methods are typically prepared thermodynamically in stable and metastable states using the solvothermal approach. Moreover, producing superior crystallized monodispersed nanocrystals is one of the features of solvothermal synthesis.⁵⁸

2.2.4 Pyrolysis. Thermal decomposition is a crucial additional method for the synthesis of magnetic nanoparticles. This process of endothermic decomposition utilizes heat to break the chemical bonds within the compound, causing the precursor to decompose. After that, force it into a chemical reaction that synthesizes MNPs and other impurities in ash. Then, NPs are separated by additional processing of the obtained ash. The precursor could be in a liquid or vapor state. A laser or plasma can also be employed in place of a flame to generate a high temperature that facilitates evaporation. It is an inexpensive and efficient continuous process with a high yield.⁵⁹ Nugroho *et al.* aimed to implement the one-pot pyrolysis for hybrid nanomaterial synthesis consisting of graphene quantum dots (GQDs) from solid citric acid and carbon dots (CDs), which were hydrothermally derived from the herbal plant extract of "*Houttuynia cordata*," denoted as CDs mixed GQDs. The nanohybrid was structurally characterized by XPS, FT-IR, fluorescence, and UV-visible absorption spectroscopy.⁶⁰ In another

study, $\text{Fe}/\text{Fe}_3\text{C}$ magnetic nanohybrids were synthesized with solid-phase pyrolysis of iron(III)-phthalocyanine. The pyrolysis temperature and synthesis duration were altered to produce a series of nanohybrids with distinct magnetic and structural properties. Nine samples were obtained with synthesis temperatures ranging from 700 °C to 1100 °C and pyrolysis duration of 5 or 15 minutes.⁶¹

2.2.5 Coprecipitation. This technique is most accurate and efficient approach to create superparamagnetic iron oxide nanoparticles (SPIONs), which have a mean diameter of less than 50 nm. In an aqueous monophasic liquid medium, it involves chemical interactions where the development and nucleation of coherent iron hydroxide nuclei must be controlled. A significant parameter in the magnetization of MNPs is the annealing temperature, with the most promising results occurring at temperature between 900 and 1000 °C.⁶² The typical MNPs synthesized through this method are Fe_3O_4 or $\gamma\text{-Fe}_2\text{O}_3$ NPs. Manganese ferrite (MnFe_2O_4) NPs were produced using sodium hydroxide (NaOH) as a precipitant and ferric chloride (FeCl_3), manganese(II) chloride (MnCl_2) as suppliers of metal ions.⁵⁴

2.3 Biological methods

Biological synthesis is a well-established method for synthesizing MNPs using living organisms, including plants and microorganisms such as fungi, viruses, bacteria, and actinomycetes.⁶³ The benefits of this method are its efficiency, eco-friendliness, and clean process. The disadvantage is its poor dispersion of the NPs.⁶⁴ The synthesis of NPs using plant tissue, extracts, exudates, and other plant parts has become an area of great interest for researchers.⁶⁵ For instance, particles with an average size of 60 nm ferromagnetic magnetite were reported to be biologically synthesized.⁶⁶ A biologically synthesized Fe_3O_4 magnetic material was used as a catalyst in the Suzuki-Miyaura reaction and photocatalysis.⁶⁷

2.4 Functionalization of MNPs

The functionalization of magnetic nanoparticles has emerged as a highly effective strategy for environmental remediation, especially in the area of photocatalytic degradation. By modifying the surfaces of these nanoparticles with photocatalytic materials, their ability to decompose organic pollutants in both water and air is significantly improved, leading to more efficient environmental cleanup processes.⁶⁸ The use of a magnetic core facilitates the efficient recovery and reuse of nanoparticles post-treatment, thereby reducing secondary pollution. When subjected to light, these functionalized nanoparticles produce reactive oxygen species capable of breaking down harmful contaminants, including dyes, pharmaceuticals, and pesticides. This mechanism not only purifies the environment but also presents a sustainable and cost-effective approach to managing industrial and agricultural waste. Additionally, the customization of the nanoparticles' surface chemistry allows for targeted interactions with specific pollutants, enhancing the overall efficiency of the degradation process. Consequently, the functionalization of magnetic nanoparticles represents



a significant advancement in utilizing photocatalytic degradation for effective environmental remediation.⁶⁹

The key element in the development of MNP is iron oxide materials, particularly magnetite and maghemite. Various synthesis methods, such as co-precipitation, thermal decomposition, and hydrothermal synthesis, are employed to produce these nanoparticles. Each technique allows for precise control over the size, shape, and magnetic properties of the particles, which in turn affects their performance characteristics. Among these methods, the chemical precipitation of iron salts is the most widely used for creating iron oxide nanoparticles. This process involves the precipitation of iron salts under basic conditions, resulting in the formation of Fe^{2+} and Fe^{3+} salts that aggregate into nanoparticles, typically at ambient or slightly elevated temperatures.⁷⁰ This approach is favored for its simplicity, cost-effectiveness, and scalability, making it suitable for large-scale industrial production. In contrast, thermal decomposition involves the breakdown of organometallic precursors in high-boiling organic solvents, often in the presence of surfactants, allowing for precise control over the nanoparticles' structure and size.⁷¹ High-performance materials that are advantageous for various applications require exceptional magnetic properties, making the intricate and expensive single-step decomposition of organometallic precursors in high-boiling organic solvents with surfactants a critical process. Hydrothermal synthesis, which involves chemical reactions at elevated temperatures and pressures within a sealed vessel containing an aqueous solution, offers remarkable control over the formation of crystals and their respective particle sizes. This production technique allows researchers to fabricate nanoparticles with outstanding magnetic saturation and stability, suitable for complex applications.¹⁶

MNPs can be effectively utilized in biomedical applications by incorporating organic polymers, which enhance their biocompatibility and stability while improving dispersibility. This integration typically involves coating the nanoparticles with biocompatible polymers such as polyethylene glycol (PEG), polydopamine, or chitosan. The hydrophilic nature of PEG significantly contributes to the nanoparticles' compatibility with biological environments, resulting in prolonged circulation times in biological fluids. For instance, Xu *et al.*, successfully immobilized β -Glucosidase (β -Glu) onto MNPs, creating MNP- β -Glu, which was further modified with PEG to produce MNP- β -Glu-PEG. This modification not only enhances the stability of the nanoparticles in blood serum but also boosts the enzyme's activity.⁷² PEGylated nanoparticles are designed to avoid detection by the reticuloendothelial system, making them ideal for applications in drug delivery and imaging. These PEG-coated nanoparticles serve as effective drug carriers, enabling the targeted delivery of therapeutic agents to specific tissues or tumors while employing controlled release systems to minimize side effects during treatment. The robust and stable coating of magnetic nanoparticles is derived from polydopamine, a versatile biopolymer inspired by mussel adhesive proteins. This material contains reactive groups that allow healthcare professionals to attach targeting ligands or therapeutic agents. Consequently, polydopamine-coated nanoparticles can fulfill

multiple roles, including targeted drug delivery, photothermal therapy, and functioning as biosensors.⁷³ Targeted cancer cell therapy is significantly improved through the precise delivery of cancer cells using polydopamine-coated Fe_3O_4 nanoparticles in conjunction with antibodies or peptides. These magnetic nanoparticles are enhanced by a biocompatible and biodegradable chitosan coating, a natural polysaccharide that facilitates chemical interactions with drugs, proteins, and nucleic acids. This characteristic makes chitosan an ideal candidate for applications in genetic delivery and tissue engineering. By utilizing chitosan-based coatings on nanoparticles, researchers have established a promising approach for delivering genetic material, which holds potential for advancements in gene therapy and regenerative medicine.⁷⁴

MNPs demonstrate improved characteristics when combined with carbon-based materials like graphene oxide (GO) and carbon nanotubes (CNTs). This synergy leads to a notable increase in surface area, enhanced electrical conductivity, and greater mechanical strength. Consequently, these hybrid composites prove to be particularly effective for environmental applications, including the removal of pollutants and the treatment of wastewater.⁷⁵ The oxygen functional groups present on GO derivatives serve as effective binding sites for magnetic nanoparticles, which, along with their adsorption capabilities and stability, define the performance of GO-coated nanoparticles. These advanced materials play a crucial role in extracting heavy metals, dyes, and organic pollutants from water sources.⁷⁶ Research has demonstrated that $\text{Fe}_3\text{O}_4/\text{GO}$ composites are effective in removing toxic heavy metals such as Pb^{2+} and Cd^{2+} from contaminated water, leveraging their magnetic recovery properties. Additionally, cylindrical CNTs are distinguished by their remarkable electrical, thermal, and mechanical characteristics, making them notable nanostructures in various applications.⁷⁷ The integration of magnetic nanoparticles with CNTs enhances the performance of the composite system. These hybrid materials are utilized in catalytic pollutant degradation and as electrode components in energy storage applications. The $\text{Fe}_3\text{O}_4/\text{CNTs}$ composites serve as efficient catalysts for degrading wastewater dyes, attributed to the high surface area and catalytic efficiency of CNTs.⁷⁸

By functionalizing MNPs with TiO_2 or ZnO , their photocatalytic properties are significantly improved, making them suitable for environmental remediation. These composites utilize a magnetic core that facilitates easy separation and reuse while effectively degrading organic pollutants under light irradiation. TiO_2 is recognized as a reliable photocatalyst due to its stability when exposed to UV light, and TiO_2/MNP composites exhibit strong degradation capabilities for various organic pollutants, including water-soluble dyes, pesticides, and pharmaceutical residues.⁷⁹ Specifically, $\text{Fe}_3\text{O}_4/\text{TiO}_2$ composites have proven to be effective in the photodegradation of methylene blue dye under UV light, while also allowing for convenient magnetic recovery for recycling purposes.⁸⁰ The photocatalytic performance of ZnO is significantly enhanced under UV light, and when ZnO is applied as a coating on MNPs, their photocatalytic efficiency improves, making them suitable for environmental applications. The use of $\text{Fe}_3\text{O}_4/\text{ZnO}$ composites



facilitates water purification through sustainable methods, promoting the degradation of organic pollutants like phenol and bisphenol A. Furthermore, the functionalization of MNPs represents a significant advancement in nanotechnology, providing effective solutions to various application challenges. By integrating MNPs with organic polymers, carbon-based materials, and metal oxides, researchers are addressing specific needs in both medical and environmental fields.⁸¹

2.5 Stabilization of nanoparticles

Nanoparticles possess a high surface area-to-volume ratio, which contributes to their inherent instability and susceptibility to aggregation, oxidation, or dissolution over time. This instability is further influenced by factors such as surface charge, ionic strength, pH, temperature, and surface chemistry. Consequently, stabilizing nanoparticles is crucial to maintain their unique properties, including enhanced reactivity, capacity, and optical characteristics. If these properties are compromised, the nanoparticles may become ineffective or behave unpredictably in various applications. Effective stabilization ensures that nanoparticles can be reliably stored, transported, and utilized in fields such as medicine, catalysis, and environmental science.⁸² When nanoparticles aggregate, they form tightly bound collections that are difficult to separate into individual particles, causing them to behave more like bulk materials and resulting in a loss of their unique qualities, such as high reactivity or catalytic efficiency. By implementing stabilization techniques, aggregation is prevented, allowing the nanoparticles to retain their distinct and active characteristics.⁸³

Chemically synthesized nanoparticles are typically stabilized by surface-capping ligands that control their final dimensions. In the absence of these stabilizing agents, van der Waals forces would cause the colloidal particles to aggregate and precipitate. Stabilizers or capping agents counteract this tendency by creating a repulsive force between the particles. For instance, charged capping agents like sodium citrate facilitate electrostatic stabilization, where negatively charged citrate-capped nanoparticles attract positively charged counter-cations from the surrounding medium, resulting in the formation of an electrical double layer that induces coulombic repulsion. Alternatively, steric stabilization is achieved through the use of long-chain thiols or larger molecules such as polymers. Various stabilization methods exist, including the chemical binding of stabilizing molecules to the nanoparticle surface, the introduction of particle charges, or the creation of physical barriers to prevent aggregation.⁸⁴

2.5.1 Electrostatic stabilization. Electrostatic stabilization originates from the repulsive forces experienced by nanoparticles when they are surrounded by a double layer of electric charges, which effectively prevents their agglomeration. This stabilization process involves charging the surfaces of nanoparticles through the introduction of ions, acids, or bases, or by ionizing surface functional groups at specific pH levels, thereby minimizing the risk of aggregation.⁸⁵ The pH of the surrounding medium plays a crucial role in determining the

ionization state of these functional groups, thereby influencing the surface charge of the nanoparticles. Optimal stability is observed at a specific pH level, where increased deprotonation of surface groups leads to a higher negative charge on the nanoparticles. This enhanced negative charge promotes greater repulsion between particles, thereby improving their stability. The zeta potential is a key measure of the effective surface charge and potential stability of nanoparticles.⁸⁶ A low zeta potential indicates inadequate stability and a likelihood of particle aggregation, while a high zeta potential, typically exceeding ± 30 mV, suggests strong repulsive forces among particles. The pH level can influence the dispersion characteristics by altering the zeta potential or surface charge. In aqueous and colloidal suspensions, electrostatic stabilization plays a crucial role in maintaining stability.

2.5.2 Steric stabilization. Steric stabilization refers to the mechanism by which large molecules, including polymers and surfactants, adhere to the surfaces of nanoparticles to prevent their aggregation.⁸⁵ As the number of these adsorbed molecules rises, the osmotic repulsive forces also increase, thereby improving the stability of the nanostructured particles. The repulsive forces that facilitate steric stabilization can be understood by examining the extended interaction potential. Factors such as temperature, average chain length, polymer concentration, and solubility influence steric stability, which remains unaffected by particle size due to its nature as a short-range interaction.⁸⁷

2.5.3 Electrosteric stabilization. Electro-steric stabilization refers to the synergistic effect of electrostatic and steric stability,⁸⁸ which is achieved by adsorbing ionic polymers onto the surfaces of charged particles. This process generates both electrostatic and steric barriers that effectively inhibit the agglomeration of dispersed nanoparticles. The electrostatic double layer formed through this technique significantly improves the dispersion of nanofluids compared to methods relying solely on steric repulsion. Typically, this approach employs polyelectrolyte polymers, which consist of repeating units linked by at least one ionized functional group, such as carboxylic or sulfonic acid groups. Commonly used anionic polyelectrolytes include polyacrylic acid, poly(styrene sulfonic acid), and poly(1-vinylpyrrolidone-co-acrylic acid), while cationic options like polyethyleneimine and poly(vinyl sulfonic acid) are also utilized for the electro-steric stabilization of nanoparticles in suspension.⁸⁹ The ionized functional groups attach to the polymers and are distributed within the base fluid, forming an electrical double layer around incoming particles. For effective stabilization, it is essential that the ionic polyelectrolytes and nanoparticles possess opposite charges. Additionally, polyampholytes, which feature both anionic and cationic functional groups, can be utilized for electro-steric stabilization. Ionic liquids contribute to the stabilization of nanoparticles by offering both electrostatic repulsion and steric hindrance, primarily due to their high charge density and strong chemisorption on the NP surface. Table 1 outlines the pros and cons of various synthesis techniques for magnetic nanohybrid materials.



Table 1 Comparison of the synthesis methods

| Synthesis method | Cost | Size control | Yield | Scalability | Advantages | Disadvantages | Ref. |
|----------------------------|------|--------------|----------|-------------|--|--|-----------|
| Mechanical ball milling | Low | Good | High | High | Wide particle size distribution, ease of operation, high efficiency | Irregular shape, high energy consumption and less production efficiency | 90 |
| Thermal decomposition | High | Excellent | Low | High | Very narrow size distribution, highly scalable and good shape control | Complicated synthesis, inert atmosphere and high temperature requirement | 54 |
| Laser ablation | High | Good | Low | Low | Narrow size distribution, monodispersity, simple & efficient, high purity & crystallinity | High input energy & uneconomical | 91 |
| Electron beam lithography | High | Fine | Low | Low | Well controlled inter-particle spacing | Expensive & extremely costly machinery needed | 18 |
| Chemical vapour deposition | High | Good | Limited | High | Production of intricate, robust, homogeneous, and highly pure NPs | Production of toxic gases as byproduct, demand for specialized equipment | 18 and 92 |
| Sol gel method | Low | Good | Variable | High | Good control of particle size and microstructure, monodispersity, desirable shape and length of the products, good crystallinity and high purity | Contamination of product with matrix component, weak bonding, low wear resistance, high permeability | 93 and 94 |
| Solvothermal | High | Excellent | Marginal | High | Very narrow size distribution & high crystallinity, precise size & shape control | High reaction temperature & pressure | 95 |
| Spray pyrolysis | Low | Good | Moderate | High | Tunable magnetic properties, homogeneous coatings | High energy consumption, risk of aggregation, low production efficiency | 96 and 97 |
| Coprecipitation | Low | Poor | High | High | Simple, narrow size distribution and high production rate | Poor morphology and non-stoichiometric magnetite | 98 and 99 |

3. Environmental applications of MNPs

3.1 Water purification using MNPs as adsorbents

Freshwater consumption is likely to increase due to the growing population. Over the past century, the demand for water has increased almost sevenfold, while the global population has risen fourfold.^{100–102} Water availability per capita is expected to decrease by 2050, increasing the likelihood of water stress.¹⁰³ Despite 2.6 billion people gaining access to quality drinking water since 1990, projections indicate that access to water-stressed areas will be available to 3.9 billion people by 2030.¹⁰⁴ Contaminants in water resources, resulting from both natural and human activities, are threatening various sectors, and the quality of these resources is declining significantly. Continued poor results in managing water quality, variations in rainfall, over-exploitation of groundwater sources, and water pollution are equally a problem, as are imbalances in the allocation of resources and the occurrence of droughts. Sewages from industries, urban/peri-urban settlements, poor disposal of wastes, traffic irrigation water, and animal wastes, especially in rural areas, play an excessive role in water pollution.¹⁰⁵

3.1.1 Adsorption of dyes. Carboxymethyl cellulose (CMC), a cellulose derivative, has several attributes: biodegradable, biocompatibility, hydrophilic, nontoxicity, and the least expensive. Nasiri *et al.* employed CMC to develop a nano-magnetic adsorbent, $\text{CoFe}_2\text{O}_4\text{@CMC}/\text{HZSM-5}$, to remove metronidazole (MNZ) from water. The structures of the adsorbents were characterized by using several methods. The maximum MNZ removal was obtained at optimum conditions of 6 pH, 20 °C temperature, 50 mg L⁻¹ concentration of MNZ, and a reaction time of 60 minutes. The average removal efficiency was 94% for the synthetic wastewater samples and 85% for the actual wastewater samples. The adsorbent lost only 14% of its capacity in six consecutive cycles.¹⁰⁶

The iron oxide-flax seed-based hybrid nanocomposites were synthesized by Choudhry *et al.* and were utilized and examined for the adsorption of malachite green. Experimental outcomes proved that the composite achieved high removal efficiency, as 90% of malachite green was obtained from a 10.0 mg L⁻¹ solution, with 1.0 g L⁻¹ of the composite, within 15 minutes at 30 °C and pH 7. The most time-consuming process in the adsorptive removal of malachite green on the catalyst was intraparticle diffusion, which might occur due to weak electrostatic and hydrogen bonding. In conclusion, this work shows



that this magnetic flax seed-based composite is effective for water treatment using adsorption technology.¹⁰⁷ Saber *et al.* synthesized two categories of nanohybrids. First, cobalt iron oxide nanoparticles were combined with Al/Zn nanolayers. The second one used long-chain fatty acids, including stearic acid, to increase the distance of Al/Zn nanolayers and create incorporation pathways for the NPs of cobalt iron oxide. Additionally, the zinc oxide nanohybrids demonstrated a lowering in band gap energy value from 3.20 to 2.75 eV, thus increasing the performance of the cells in the sunlight. This high efficiency of the nanohybrid was illustrated by the total decolorization of Acid Green 1 dye within 10 minutes of exposure to sunlight, unlike the pure and doped ZnO, which took 360–840 minutes. The pseudo-first-order kinetic profile revealed that the ZnO nanohybrid prepared from an inorganic-magnetic-organic composite had higher activity than that prepared from an inorganic-magnetic composite. This approach offers a viable plan for addressing energy and water purification issues using clean, non-harmful power sources.¹⁰⁸

Composites of natural polymers have great potential as sorbents for magnetic solid-phase extraction (MSPE). Chitosan, a natural polymer known for its amino and carboxylic groups, offers biodegradability, biocompatibility, and ease of modification.^{109–111} Recently, magnetic chitosan nanohybrid materials were created by dispersing ferrites throughout the chitosan matrix, forming ternary ferrites chitosan microspheres (TFCM).¹¹² The nanoparticles were prepared using a co-precipitation method, and the chitosan matrix was incorporated to produce an active adsorbent for the degradation of methylene blue dye. Zhou *et al.* prepared magnetic $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-NH}_2/\text{F}_{13}$ hybrid material using a one-step reaction to modify the surface of SiO_2 by an amino group and an octyl-perfluorinated chain through a sol-gel procedure. The resulting material ($\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-NH}_2/\text{F}_{13}$) exhibited tremendous adsorption ability of perfluorinated compounds (PFC) from the aqueous sample, as confirmed by HPLC-MS/MS analysis. 50 mg of the $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-NH}_2/\text{F}_{13}$ sorbent in a 500 mL water sample reached the adsorption equilibrium within 30 minutes. Fluorine-fluorine interactions, electrostatic solid attraction, and

size exclusion effects are attributed to the high adsorption efficiency. The recovery rates ranged from 90.65% to 106.67%, with low limits of detection (LODs) between 0.029 and 0.099 ng L⁻¹. Therefore, $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-NH}_2/\text{F}_{13}$ is an effective adsorbent for measuring the concentration of PFC in large volumes of water samples (Fig. 2).¹¹³

Qin *et al.* significantly synthesized MNPs, which were coated with mesoporous silica, which was modified using methyl dimethoxy and *p*-toluenesulfonic acid (PTSA) to prepare $\text{Fe}_3\text{O}_4/\text{mSiO}_2\text{-Me-PTSA}$ material. This material was tested for removing polychlorinated biphenyls (PCBs) from wastewater. The material facilitated rapid magnetic separation, with an adsorption time of just 10 minutes for PCBs. The $\text{Fe}_3\text{O}_4/\text{mSiO}_2\text{-Me-PTSA}$ showed high enrichment factors for PCBs (119 to 147) and an adsorption efficiency of 46.3 mg g⁻¹. The process demonstrated promising outcomes with recoveries between 85.25% and 118.60% and low LODs ranging from 0.16 to 0.91 ng L⁻¹.¹¹⁴ Wu *et al.* synthesized a polymorphic graphene-like porous carbon with N-doping/Ni magnetic nanohybrids. The NC@Ni-1 sample demonstrated exceptional long-term cycle stability, exceptional rate performance, and high specific capacity. The NC@Ni-0.75 sample showed a high adsorption capacity for MG dye (889 mg g⁻¹). It also demonstrated a distinctive magnetic separation characteristic, making it suitable for the adsorptive removal of cationic MG dye.¹¹⁵

Choudhry *et al.* discuss synthesizing magnetic flax seeds (MFS), a unique hybrid nanocomposite based on iron oxide and flax seeds for water treatment. A straightforward co-precipitation technique was used to create the composite. With an equilibrium adsorption capacity of roughly 10 mg g⁻¹ and the highest adsorption capacity (Q_{max}) of 52.91 mg g⁻¹, the experimental results revealed that MFS successfully absorbed over 90% of malachite green dye (MG) from a ten mg L⁻¹ solution at 30 °C at neutral pH in 15 minutes. Weak electrostatic and hydrogen bonding interactions drove the adsorption process, which followed the Freundlich isotherm, indicating multilayer physisorption. The rate-determining phase was intraparticle diffusion. According to kinetic analyses, the pseudo-second-order model described MG adsorption. While

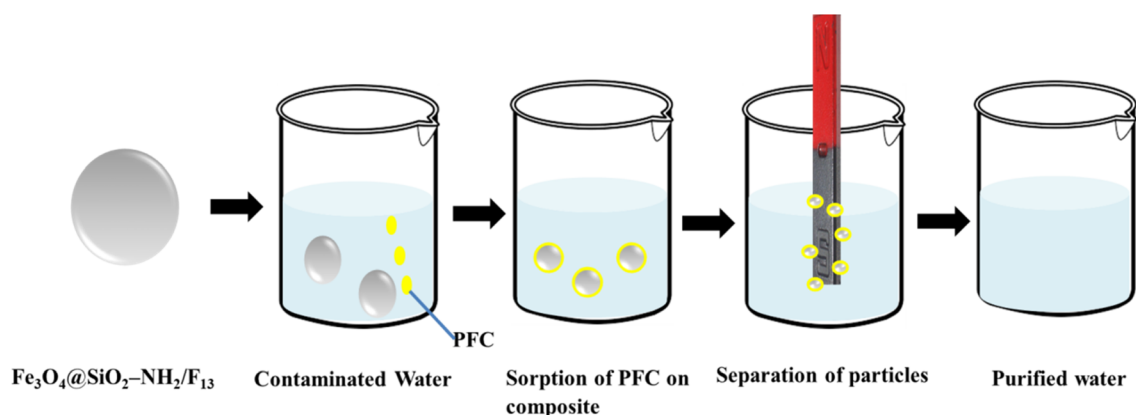


Fig. 2 Removal of PFC particles from surface water using $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-NH}_2/\text{F}_{13}$ magnetic hybrid composite. This figure has been reproduced from ref. 113 with permission from Elsevier, copyright 2016.

MFS proved effective for water treatment, the study suggests further investigation of desorption and regeneration for future applications.¹⁰⁷

Elanchezhian *et al.* effectively synthesized zirconium-imprinted manganese ferrite ($\text{ZrMnFe}_2\text{O}_4@\text{rGO}$) nanohybrids based on rGO and assessed their capacity to adsorb PFOA and PFOS. As validated by Raman spectroscopy, appropriate I_D/I_G ratios demonstrated the intercalation of $\text{Zr-MnFe}_2\text{O}_4$ nanoparticles within rGO. Interactions with the rGO sheets decreased the $\text{Zr-MnFe}_2\text{O}_4$ nanohybrids' saturation magnetization. According to kinetic studies, the highest adsorption efficiency for PFOA and PFOS on $\text{Zr-MnFe}_2\text{O}_4@\text{rGO}$ was achieved in less than 12 hours. The adsorption course was pH-dependent for both pollutants and adhered to a pseudo-second-order kinetic model. The Langmuir model, with maximum adsorption capacities of 10.1 mg g^{-1} for PFOA and 12.6 mg g^{-1} for PFOS, mainly attributable to electrostatic and hydrophobic interactions, was found to have the best fit to the experimental data, according to adsorption isotherm analysis. Notably, the adsorption density of $\text{Zr-MnFe}_2\text{O}_4@\text{rGO}$ remained unchanged after four consecutive recycling tests, with no observable mass loss of the material.¹¹⁶

3.1.2 Emerging contaminants. Wastewater comprises diverse and complex chemical contaminants; although around 700 have been identified, mainly harmful heavy metals, and numerous others, especially organic pollutants, remain unknown due to their diversity and low levels.¹¹⁷ Mostly ions of Pb, Ni, Cr, Cu, Co, As, Cd, Zn, and Hg, originate from mining, tanneries, fertilizers, painting, metal plating, and batteries.^{118–121} Heavy metals are not biodegradable and toxic; they persist in the aquatic system for extended periods in various oxidation states. Urban pollution also brings heavy metals from contaminated air, dust, and sediment, transportable commodities. These metals are most toxic in their biological forms, particularly to ecosystems and humans. Consequently, it is necessary to filter out these metals from the wastewater before they are released into water sources. Due to their large surface area and adequate adsorption capacities, magnetic nanohybrid materials such as magnetic activated carbon, magnetic metal–organic frameworks (MOFs), and magnetic mesoporous carbon are frequently employed to eliminate heavy metals from wastewater. Magnetic nanoparticles doped or deposited on biomass materials, such as chitosan-coated magnetic nanoparticles or magnetic carbon generated from coconut shells, can create biomass-based magnetic nanohybrid adsorbents that additionally showed efficacy in applications involving the removal of heavy metals.

3.1.3 Removal of metal ions. Amin *et al.* aim to design affordable magnetite nanoparticles for the removal of heavy metals from water. MNPs were made using the thermal decomposition of iron oleate precursors, and the surface of these nanoparticles was covered with layers of mesoporous silica using cetyltrimethylammonium bromide. Chitosan coating was then applied to the coated particles. The adsorption study revealed that the adsorption was monolayer chemical adsorption, and the adsorption capacity was 150.33 mg g^{-1} for Pb^{2+} and 126.26 mg g^{-1} for Cd^{2+} , which was entirely

satisfactory.¹²² Creating a magnetic nanocomposite material, magnetic graphene oxide-covalently functionalized tryptophan, required functionalizing graphene oxide with tryptophan and adding magnetic characteristics to enhance its retrievability after use. The significant adsorption capacity of the material of 766.1 mg g^{-1} demonstrated its effectiveness in eliminating lead(II) ions, indicating its potential for use in lead remediation. GO was used to eliminate cadmium(II) ions, indicating a noteworthy adsorption capacity of 530 mg g^{-1} , confirming its efficacy in eliminating cadmium. The high adsorption capabilities attained imply that the materials worked well in the experimental setup of the study.¹²³

Muntean *et al.* synthesized carbon magnetic nanocomposite and used it as an adsorbent to extract heavy metal ions from aqueous solutions, including copper, lead, and zinc. Working at the natural pH of 5.8 for 240 minutes and utilizing a dose of 1 g L^{-1} of the nanocomposite were the ideal circumstances for the process, which produced high removal efficiencies of 81.36% for copper, 84.50% for lead and 72.68% for zinc ions. The investigation results showed that these parameters were appropriate for effective heavy metal ion removal operations employing the carbon magnetic nanocomposite material since they struck a good balance within removal rate, practicability, and adherence to ecological and socioeconomic issues.¹²⁴ The chitosan matrix can promote the growth of MNPs and, at the same time, hinder their aggregation. Tolessa *et al.* synthesized two μm magnetic chitosan hybrid materials through suspension cross-linking to eliminate Ag NPs. They suspended MNPs in 1% chitosan solution and added toluene. The mixture was stirred at 500 rpm for half an hour; then, the solutions of NaOH and glutaraldehyde were added. The magnetic chitosan composite material was separated from the solution with the help of an external magnet. Alongside inductively coupled plasma-mass spectrometry (ICP-MS), the material provided low LODs for Ag NPs, from 0.016 to $0.023 \mu\text{g L}^{-1}$. The high removal efficiency of Ag nanoparticles, between 84.9% and 98.8%, is due to the positive charges on the chitosan's surface, making it an excellent adsorbent. The magnetic chitosan hybrid material remains effective after three recycling uses, with an efficiency of approximately $77.2 \pm 2.2\%$.¹²⁵

In another study, a core-shell $\text{Fe}_3\text{O}_4@\text{RF}@m\text{TiO}_2$ material was developed and tested as an MSPE sorbent for removing arsenic from highly acidic samples by Zhao *et al.* The material featured a 130 nm Fe_3O_4 core, a 50 nm resorcinol–formaldehyde (RF), and a mesoporous TiO_2 shell. This design provided quick adsorption, which is $1.16 \text{ g mg}^{-1} \text{ h}^{-1}$ and a 139 mg g^{-1} adsorption capacity. The RF layer offered hydrophobic protection to the Fe_3O_4 core against acid etching, while the mesoporous TiO_2 enabled surface complexation and electrostatic interactions with arsenate. The magnetic properties allowed facile separation and recycling using an external magnetic field. This multi-layer material shows promise for long-term wastewater treatment.¹²⁶ The adsorption capabilities of magnetic iron oxide nanoparticles for heavy metals, including Pb, Cd, Cu, Zn, and Co, in wastewater treatment procedures are significantly increased by modification with chemicals such as thio-salicylhydrazide. Hybrid magnetic nanohybrid materials have



been synthesized to overcome the limits of original magnetic nanoparticles in water remediation and to advance heavy metal removal efficiency by combining magnetic capabilities with other compositions. Magnetic nanohybrid materials with enhanced adsorption capabilities and practicality for environmental remediation applications have been developed due to hybridization operations' new compositions and chemical structure alterations.¹²⁷

A flower-like magnetic $\text{MoS}_2/\text{Fe}_2\text{O}_4$ nanohybrid has been successfully synthesized by Wang *et al.* and used for the removal of $\text{Pb}(\text{II})$ and $\text{Hg}(\text{II})$ from contaminated wastewater and soil. The nanohybrid was highly efficient in removing heavy metal ions, with rates of 99.56% for $\text{Pb}(\text{II})$ and 99.08% for $\text{Hg}(\text{II})$. pH, contact time, initial metal ion concentration, and temperature significantly influence the removal efficiency. The nanohybrid also demonstrated excellent reusability after multiple adsorption and desorption cycles, indicating its economic and environmental sustainability. The study concludes that the flower-like magnetic MoS_2 nanohybrid is a promising material for efficient remediation of these contaminants, with its high removal efficiency, reusability, and sensitivity to various parameters making it a valuable tool for environmental cleanup processes.¹²⁸ Using a facile approach, Dai *et al.* synthesize a ternary adsorbent, PEHA-Phos-GO/ MnFe_2O_4 . This material showed improved $\text{Pb}(\text{II})$ adsorption properties compared to GO because of the large number of adsorption sites offered by the amino and phosphate groups. Experimental data showed that the adsorption process depended on the pH and involved pseudo-second-order, which pointed to chemisorption as the rate-controlling step. The data obtained from the Langmuir isotherm agreed with the experimental data. The maximum adsorption capacity for $\text{Pb}(\text{II})$ was 366.4 mg g^{-1} at pH 5.5, which was higher than that of pristine GO, 212.1 mg g^{-1} . Recycling tests further supported the higher capacity of regeneration of PEHA-Phos-GO/ MnFe_2O_4 , and the work demonstrated the possibility of applying the synthesized material in water treatment processes, as shown in Fig. 3.¹²⁹

Pirhaji *et al.* successfully synthesized nano $\text{NiFe}_2\text{O}_4/\text{HNTs}/\text{GQDs}$ as a novel adsorbent for $\text{Pb}(\text{II})$ removal (Fig. 4). Batch experiments revealed that its strong adsorption performance is mainly due to complexation interactions between Pb^{2+} ions and the functional groups on the adsorbent. The Box-Behnken

design and response surface methodology determined optimal experimental parameters, achieving a maximum Pb removal efficacy of 97.14%. The Langmuir isotherm model best explained the adsorption behavior, with a Q_{max} of 42.02 mg g^{-1} at 25°C . The Elovich model suggested chemical adsorption and kinetic studies indicated that the process followed a pseudo-second-order model. The thermodynamic analysis confirmed that the adsorption is endothermic and spontaneous, making $\text{NiFe}_2\text{O}_4/\text{HNTs}/\text{GQDs}$ a cost-effective and efficient option for $\text{Pb}(\text{II})$ removal.¹³⁰

A new magnetic solid-phase extraction (MSPE) approach using Ni-Al layered double hydroxide/magnetite (Ni-Al LDH/ Fe_3O_4) nano sorbent was synthesized by Abdolmohammad-Zadeh *et al.* for the enrichment and separation of $\text{Mn}(\text{VII})/\text{Mn}(\text{II})$ in water samples by flame atomic absorption spectrometry. The nanosorbent selectively adsorbs $\text{Mn}(\text{VII})$ because of the electrostatic attraction and intercalation of MnO_4^- ions into the LDH layers. $\text{Mn}(\text{II})$ is quantified indirectly by calculating the difference between the total Mn and $\text{Mn}(\text{VII})$ concentrations. The nano-hybrid is paramagnetic, allowing simple separation using an external magnet without filtration or centrifugation. The nanosorbent is non-toxic, recyclable more than 100 times, and offers high sensitivity, accuracy, and selectivity for Mn speciation in water samples.¹³¹

In another study, Katubi *et al.* focused on synthesizing the $\text{MnFe}_2\text{O}_4/\text{GO}$ nanocomposite using an *in situ* method for the adsorption of Pb^{2+} ions in water. The adsorption parameters, including initial concentration, adsorbent dose, pH, and reaction time, were determined to identify the best condition for adsorption. The study revealed that with the increase in the initial concentration of Pb^{2+} ions, the adsorption capacity ranged between 63 and 625 mg g^{-1} . Also, the extent of removal increased with the increase in the adsorbent dose from 39% to 98.8% as the number of active sites increased. The results also revealed that the adsorption process was most effective at pH 6. The adsorption capacity increased sharply during the first 15 minutes and then progressively at a slower rate as the surface became fully occupied, the maximum contact time being 120 minutes. The study also exhibited that Pb^{2+} adsorption fitted the Langmuir isotherm model, which supports monolayer adsorption with a maximum 636.94 mg g^{-1} adsorption capacity. Moreover, reusability tests proved the stability and efficiency of

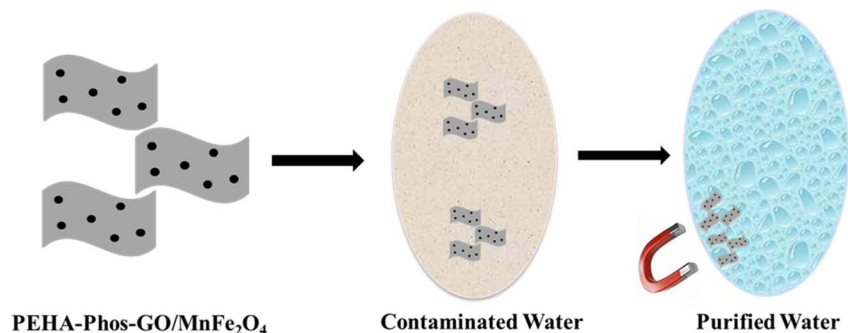


Fig. 3 Removal of lead ions using PEHA-Phos-GO/ MnFe_2O_4 nanohybrid.

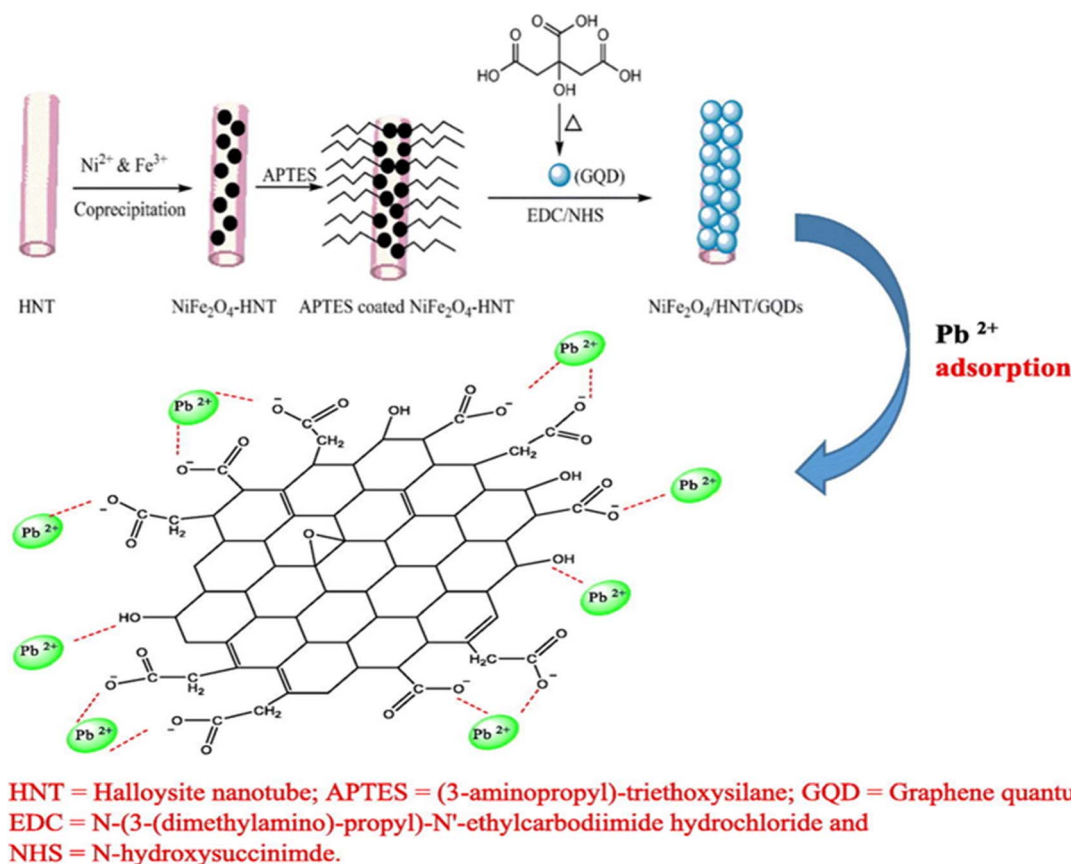


Fig. 4 Synthesis of NiFe₂O₄/HNTs/GQDs and plausible adsorption mechanism of Pb(II) ions onto NiFe₂O₄/HNTs/GQDs. This figure has been adapted from ref. 130 with permission from Elsevier, copyright 2020.

the nanocomposite for at least five cycles, thus making it an efficient and cost-effective adsorbent for Pb²⁺ ions in water treatment.¹³² rGO was synthesized utilizing the modified Hummers' method, and ZnFe₂O₄ nanoparticles were incorporated onto the surface of rGO to fabricate the adsorbent material to remove lead ions. Experimental parameters, including the amount of the adsorbent, initial concentration of the pollutant, and pH, were thus adjusted to maximize the pollutant removal efficiency. The developed rGO nanocomposite had an adsorption capacity of 89.8 mg g⁻¹ for lead ions. Vacancy defects and oxygen-containing functional groups on the rGO surface were identified as the primary mode for pollutant elimination. Adsorption was consistent with the Langmuir isotherm and fitted a pseudo-second-order kinetic model, suggesting that the adsorption process was chemisorptive and through pollutant monolayers. The synthesized rGO nanocomposite was stable and reusable for up to five cycles, which is suitable for application in heavy metals and dye removal in the water treatment process.¹³³

In another work, coprecipitation and Hummer's modified method were combined to synthesize the magnetic Fe₃O₄-GO nanohybrid composite material. The morphology of the hybrid sample demonstrated that the diffusion of Fe₃O₄ nanoparticles into the mesoporous GO layers of porous channels limited GO nanosheet restacking and stopped magnetic nanoparticle

leaching and agglomeration. The pseudo-second-order kinetic and Langmuir isotherm models suit the results of the kinetic and isotherm analyses conducted to assess the adsorption mechanism quite well. After 60 minutes, the As(v) adsorption efficiency, H, peaked at 99.37%. The GO/Fe₃O₄ nanohybrid material in an acidic media showed a Q_{max} of 14.1 mg g⁻¹. Further research was also done on the function and contribution of GO and Fe₃O₄ nanoparticles in adsorptive capacity and the enhancement of As(v) adsorption efficiency.¹³⁴

3.2 MNPs for the cleanup of oil spills

NH₂-functionalized magnetic materials can be used for extracting oil droplets from water.¹³⁵ High magnetization and a large surface area are required for nanocomposites to have a strong oil absorption capability.¹³⁶ Atta *et al.*¹³⁷ synthesized a novel class of magnetic sorbents by employing the coprecipitation approach to coat magnetite nanoparticles with rosin amidoximes, which produced a hydrophobic coating on the surface of Fe₂O₃ nanoparticles. This hydrophobic coating can increase the oil uptake selectivity, stop Fe₂O₃ nanoparticles from oxidizing, and prevent magnetic nanoparticles from aggregating. By comparing iron oxide coated with RK-AN amidoxime to that coated with R-AN and bare Fe₃O₄ nanoparticles, oil recovery measurements indicate that the latter is a worse oil sorbent. Fe₃O₄ nanoparticles coated with RK-AN amidoxime



exhibit a higher level of oil removal than those coated with R-AN amidoxime, as shown by their respective magnetization saturation values of 60.9 and 70.8 emu g⁻¹.

Using the unique physicochemical characteristics of ZnO-T, PDMS, and Fe₂O₃-NR, a magnetic nanosorbent was effectively synthesized by Sharma *et al.* in 2019. With a WCA of 157°, the as-synthesized PZF nanosorbent exhibited superhydrophobic super oleophilic properties. With an oil removal (%) of 96% and an adsorption capacity of 1135 mg g⁻¹, respectively, this magnetic nanosorbent efficiently separates diesel oil from water. The kinetics of oil adsorption showed that diesel oil adsorptive capacity on PZF nanosorbent is high, with equilibrium potentially reached in as little as 50 minutes. The pseudo-first-order kinetics and Langmuir isotherm models matched the oil adsorption process well. The simple synthesis process and the superhydrophobic-superoleophobic and magnetic responsive characteristics of the as-prepared nano sorbent offer an alternate method for treating oily wastewater.¹³⁸ Shao and Yu successfully made a sepiolite/graphene oxide (Sep/GO) membrane for wastewater treatment. The excellent oil/water selectivity and low oil adherence of the synthetic Sep/GO membranes made them highly successful in separating oil-in-water emulsions. Moreover, the membrane preserved its underwater high stability and superoleophobicity in severe settings. As a result, the Sep/GO membrane works well in processes involving wastewater treatment and oil/water separation.¹³⁹

Qian and Chen described a TiO₂/sulfonated GO/Ag NPs membrane with beneficial photocatalytic and wettability capabilities.¹⁴⁰ When solubilized methylene blue (MB) was exposed to UV light, the synthesized membrane demonstrated good photodegradation and an effective oil/water separation performance. As a result, the membrane has interesting uses in wastewater treatment and oil/water separation.¹⁴⁰ Zhang *et al.* synthesized a superoleophilic and oil-superhydrophobic composite nanofibrous membrane (GPNM) for gravity-driven surfactant-stabilized water-in-oil emulsions using a straightforward one-step electrospinning technique. The synthetic composite membrane demonstrated a remarkable separation efficiency of 99.8% for water-in-oil emulsions. These findings revealed that the composite nanofibrous membrane could potentially treat oil-contaminated wastewater.¹⁴¹

In another study, Cheng and Barras effectively created a superoleophilic/superhydrophobic rGO-polydopamine membrane functionalized with 1H,1H,2H, and 2H-perfluorodecanethiol (rGO-PDA-PFDT) using a straightforward two-step procedure. The rGO-PDA-PFDT membrane was utilized to clean water by separating oil. The membrane's superhydrophobicity and superoleophilicity allowed it to extract chloroform from water. Therefore, in oil/water separation operations, the synthesized rGO-PDA-PFDT membrane is exceptionally favorable.¹⁴² Shokry *et al.* synthesized a novel magnetic sorbent material from water hyacinth to remove oil from polluted water. The chemical and thermal treatment established that the root segment of the water hyacinth had a higher oil adsorption capacity than the shoot segment. The comparison of the two segments with nano-activated carbon

and NMAC (nano-magnetic activated carbon) showed that NMAC had a more crystalline structure, better pore structure, and thermal stability. It also registered a saturation magnetization of 17.2587 emu g⁻¹. Oil removal efficiency increased with the solution temperature and NMAC dosage increase. The results indicated that the Freundlich isotherm model best described the oil adsorption process rather than the Langmuir isotherm and Dubinin-Radushkevich models, implying monolayer adsorption and surface heterogeneity. The pseudo-second-order and intra-particle diffusion models were utilized to determine adsorption kinetics. The primary process through which oil was adsorbed on NMAC was physical absorption through van der Waals forces.¹⁴³

Sharma *et al.* investigated the physicochemical characteristics of ZnO-T, PDMS, and Fe₂O₃-NR to give large surface area, low surface energy, and super-paramagnetism, respectively, in synthesizing a magnetic nanosorbent. The PZF nanosorbent exhibited superhydrophobic and superoleophilic properties, exhibiting a WCA 157°. With an oil removal percentage of 96% and an adsorption capacity of 1135 mg g⁻¹, respectively, this magnetic nanosorbent efficiently separates diesel oil from water. According to the oil adsorption kinetics, diesel oil can quickly adsorb on PZF nanosorbent, reaching equilibrium in as little as 50 minutes. The Langmuir isotherm and pseudo-first-order kinetics models fit the oil adsorption process well. The easily synthesized nanosorbent's superhydrophobic and superoleophilic properties, combined with its magnetic responsiveness, offer a novel approach to treating oily wastewater.¹³⁸

3.3 MNPs for photocatalytic degradation of organic dyes

Among the various pollutants, waterway dyes are recognized as a global ecological concern that can impact humans, plants, and animals. Materials that are widely studied for the photocatalytic degradation of organic dyes found in water effluents are magnetic nanohybrids. Organic dye degradation *via* photocatalysis using magnetic nanohybrid materials follows the mechanism shown in Fig. 5.

Yuan *et al.* prepared Fe₃O₄@SiO₂@meso-TiO₂ nanocomposites with a core-shell architecture by a simple and efficient method with increased photocatalytic performance. The magnetic core, SiO₂ interlayer, and mesoporous TiO₂ outer shell constitute the as-made core-shell architecture. It was produced with TiOSO₄ as the titanium source and cetyltrimethylammonium bromide (CTAB) as a pore-forming agent. Fe₃O₄@SiO₂ was made using a sol-gel method. In comparison to Fe₃O₄@SiO₂@solid-TiO₂, the Fe₃O₄@SiO₂@meso-TiO₂ composite showed better photocatalytic efficiency for the degradation of Rh B in aqueous suspension due to the large BET surface area produced by the mesoporous TiO₂ shell.¹⁴⁴ Aliyan *et al.* have explained the immobilization of Fe₃O₄ to modify the surface of meso-structured silica (SBA-15) to produce a nanocatalyst. Aliyan *et al.* used a UV lamp as a light source to investigate the catalytic activity of Fe₃O₄@SBA-15 toward the photodegradation of MG in a photo-catalytic reactor. The impact of several experimental factors on the process's



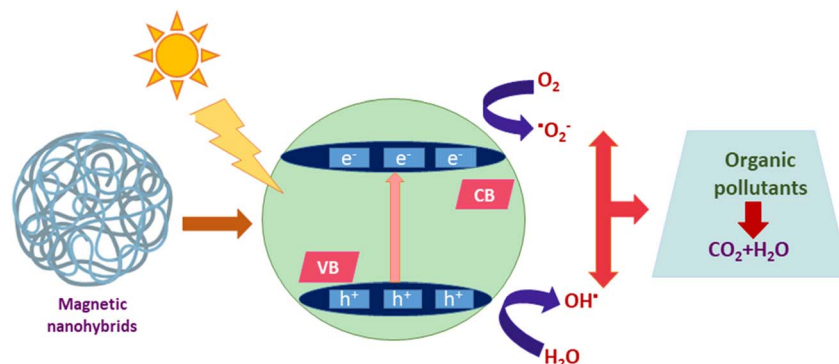
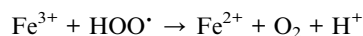
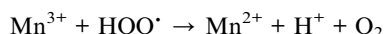
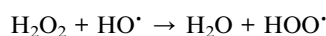
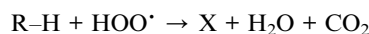


Fig. 5 General mechanism for organic dye degradation by photocatalysis.

degrading performance was assessed by analyzing catalyst dose, initial dye concentration, and dye solution pH in the presence of $\text{Fe}_3\text{O}_4@\text{SBA-15}$ as a photo-catalyst. Even after five MG degradation cycles, the described photocatalyst showed notably high catalytic stability and low activity loss.¹⁴⁵ Sahoo *et al.* investigated the adsorption of methyl orange (MO) on monodispersed magnetic mesoporous manganese ferrite composites with a size of around 200 nm. Sonication and photolysis in the presence of sunshine improved the MO's degrading efficiency. The authors concluded that their catalyst exhibited adsorption, degradation, recovery, and reusability in a single system. The electron transfer activities inside the MnFe_2O_4 particles cause the potential dye degradation reaction. The following describes the potential method by which Mn^{2+} ions decompose H_2O_2 .¹⁴⁶ Yonghong Ni *et al.* found that adding Mn^{2+} to TiO_2 nanoparticles increased their photocatalytic activity and that the metal ions, such as Fe^{2+} , actively contributed to the degradation of H_2O_2 to produce OH radical species.¹⁴⁷



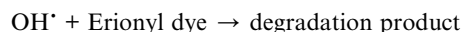
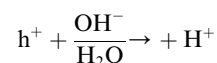
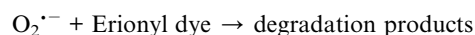
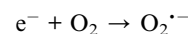
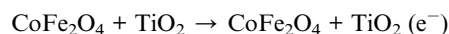
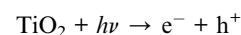
The organic dye (RH) reacts with the O_2 molecule and OH produced above.



X = degraded organic product.

Santos *et al.* prepared the magnetic $\text{CoFe}_2\text{O}_4/\text{TiO}_2$ composite through a facile solvothermal route. Compared with pristine TiO_2 , the $\text{CoFe}_2\text{O}_4/\text{TiO}_2$ catalyst had a reduced band gap, allowing it to exhibit higher photocatalytic efficiency. The nanohybrid catalyst showed exceptional activity in degrading Erionyl Red A-3G dye, with 69.4% TOC eradication after 120 minutes and 95.39% color removal after 15 minutes of reaction. One possible explanation for the degradation of Erionyl Red A-3G is that when the $\text{CoFe}_2\text{O}_4/\text{TiO}_2$ composite is exposed to UV

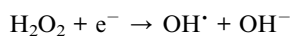
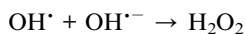
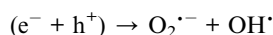
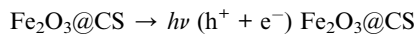
light, the TiO_2 in the composite absorbs light and produces more charges. As a result, holes (h^+) are formed on the valence band (VB) of TiO_2 due to an electron transfer (e^-) from VB to CB. A successful e^-/h^+ pair separation is subsequently promoted by the photogenerated electrons (e^-) migrating from the CB of TiO_2 to that of CoFe_2O_4 , which helps to limit photogenerated charge recombination and enhance photodegradation activity. The e^- on the CB of CoFe_2O_4 reacts with O_2 to produce $\text{O}_2^{\bullet-}$ radical, which destroys pollutants. While the h^+ on the VB of TiO_2 directly oxidizes organic molecules and combines with OH^- or H_2O to produce OH^\bullet radical, which also functions as an oxidizing species.¹⁴⁸



Rehman *et al.* synthesized $\text{Fe}_3\text{O}_4@\text{GO}$ by modifying 3-aminopropyl triethoxysilane (APTES) and tetraethyl orthosilicate (TEOS) using a mechanical stirring technique. Under UV light, the synthesized nanocomposite was evaluated as a possible heterogeneous catalyst for methylene blue (MB) degradation. Over time, the catalysts' photocatalytic activity increased progressively. The $\text{Fe}_3\text{O}_4@\text{GO}$ composite catalyst had the highest MB removal efficiency of 70.06%, which was more excellent than pure Fe_3O_4 (57.56%).¹⁴⁹ Vigneshwaran *et al.* reported that co-precipitation created the Fe_2O_3 -reinforced chitosan ($\text{Fe}_2\text{O}_3@\text{CS}$) nanocomposite. The surface of the chitosan was made available to develop a beneficial nanocomposite. Because of the increased partition efficiency and quicker



transfer of the photo-generated charge carriers, these Fe₂O₃ wrappings on chitosan offer synergistically enhanced characteristics. MO and OG dyes were degraded using this combination. According to the data, the most significant degradations for MO and OG were 89.2% and 94.6%, respectively. The trapping study highlighted how OH radicals contributed to dye degradation over Fe₂O₃@CS composites. The suggested method used Fe₂O₃@CS to oxidize anionic MO and OG dyes, as shown in Fig. 6.



Furthermore, the production of reactive species increases quickly due to the e^- - h^+ pair's delayed recombination. It has a significant impact on the materials' photocatalytic efficiency.¹⁵⁰

Sobhani-Nasab *et al.* successfully created a novel MnWO₄/TmVO₄ ternary nanohybrid using a straightforward sonochemical technique. Under visible light, the degradation of phenol red (Ph R), eosin Y (EY), 2-naphthol, and rhodamine B (Rh B) by MnWO₄/TmVO₄ was examined. The corresponding photocatalytic efficiencies for Rh B, Na, Ph R, and EY are 99.2%, 93.3%, 88.3%, and 83.1%.¹⁵¹ When MnWO₄/TmVO₄ is excited by visible light, the photogenerated holes on the VB of TmVO₄ will move to MnWO₄, while the photoinduced electrons on the CB of MnWO₄ will migrate to TmVO₄ (depicted in Fig. 7). The heterojunction forms relatively tiny amounts of O₂^{•−} and OH[•]. The degradation results demonstrate that MnWO₄/TmVO₄ has a greater photodegradation capacity than pure TmVO₄ or MnWO₄.¹⁵²

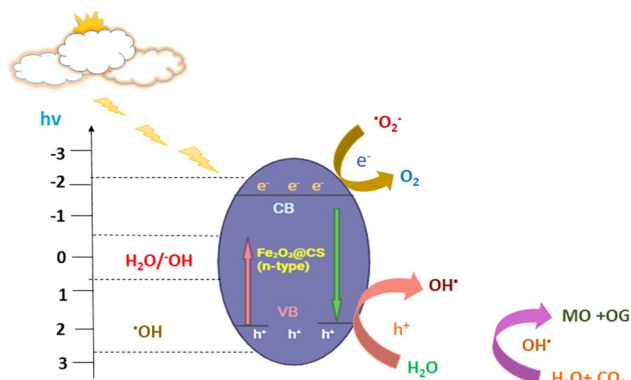
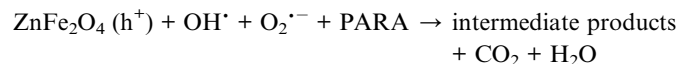
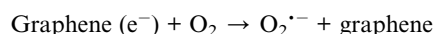
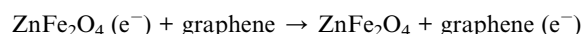
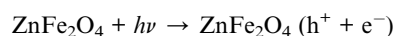


Fig. 6 Illustration of MO and OG dye molecule oxidation after e^- - h^+ production on a Fe₂O₃@CS photocatalyst.

Roostaei *et al.* investigated the reduction process by sodium citrate, Hummer's method, and hydrothermal approach was used to synthesize Fe₃O₄@Au core-shell nanoparticles decorated on rGO nanocomposite, respectively. These nanostructures were used to break down crystal violet (CV) by photocatalysis. Under ideal circumstances and visible light irradiation, it was found that the Fe₃O₄@Au-rGO performed better than Fe₃O₄@Au and rGO alone. In the presence of 0.008 g of the produced nano-photocatalyst, 100% degradation was obtained after 1 minute, and in the presence of 0.005 g, 100% degradation was obtained after 5 minutes.¹⁵³ Because rGO can function as an effective electron transporter for separating photo-generated electron-hole pairs through interfacial charge transfer, Fe₃O₄@Au-rGO demonstrated remarkable photocatalytic activity to destroy crystal violet.¹⁵⁴

3.4 MNPs for photocatalytic degradation of emerging contaminants

Emerging contaminants (EC) are new types of pollutants recently introduced into aquatic systems that harm the environment and the health of living things. Personal care products (PPCPs) and medications are prominent examples of emerging pollutants.¹⁵⁵ Since PPCPs are more difficult to degrade than other organic contaminants, magnetic nanohybrids are an excellent option for eradication.¹⁵⁶ Using the sol-gel technique, Rostami *et al.* synthesized ZnFe₂O₄-4 weight percent GR nanohybrids. It is evident that when exposed to visible light, ZnFe₂O₄ itself is nearly photo-catalytically inert; however, the addition of graphene significantly improves photocatalytic activity. Under exposure to visible light, the ZnFe₂O₄-x weight percent GR nanohybrids demonstrated enhanced charge separation and nano-photocatalytic activity towards PARA decomposition. Fig. 8 is a potential mechanism for increasing nano-photocatalytic activity.¹⁵⁷



Tang *et al.*, using a hydrothermal process, prepared rGO-CdS/ZnS heterostructure composites that have been effectively synthesized by building the CdS/ZnS heterostructure nanoparticles on rGO sheets. In contrast, GO is reduced at the same time. With 15% RGO content, the rGO-CdS/ZnS composites, as synthesized, show very active TC photodegradation (about 85%). When RGO is added, the rGO-CdS/ZnS heterostructure photocatalysts show less electron-hole pair recombination than CdS/ZnS, facilitating better charge-carrier migration and photocatalytic activity.¹⁵⁸ The charge transfer and photocatalytic activity by RGO-CdS/ZnS composites are schematically

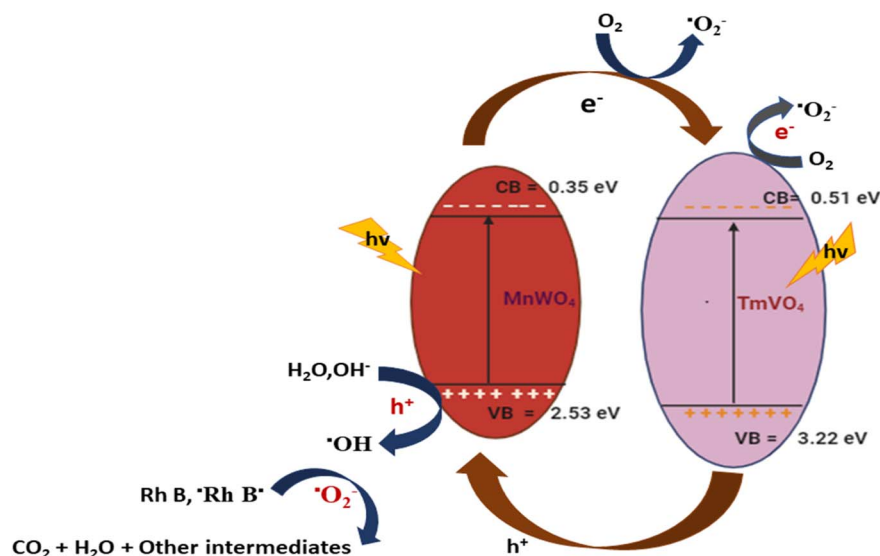


Fig. 7 Rh B degradation scheme in the presence of ternary nanohybrids.

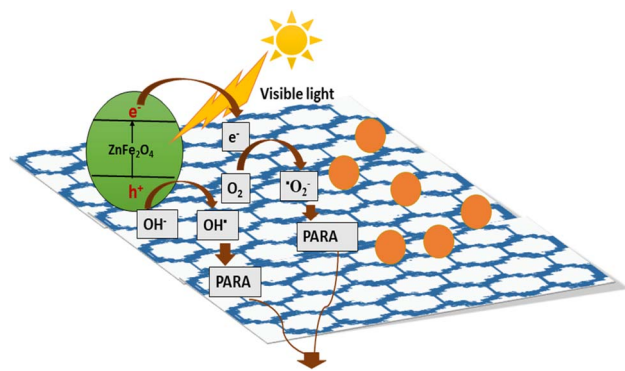


Fig. 8 A graphical representation of PARA's photocatalytic degradation over the ZnFe_2O_4 -x wt% GR nano-hybrid.

illustrated in figure. The production of holes in the VB is caused by the very initial excitation of the electron-hole pairs in CdS, which stimulates the electrons to the conduction band (CB). Because of the band potential, the electrons in the CB of CdS may be immediately transported into the CB of ZnS. The potential of ZnS is around 0.91 eV, which is less negative than the CB level of CdS (about 0.95 V). Additionally, RGO may trap the free photo-generated electrons in the ZnS CB, which are linked to the 2D carbon network. Thus, the visible-light activity was significantly produced by the photo-generated electrons of ZnS and RGO being so active that they participated in the surface reaction to make radicals.¹⁵⁹

To degrade amoxicillin (AMX), a new MIL-68(In)- NH_2 /GO composite was prepared as a visible-light-driven photocatalyst. Compared to pure MIL-68(In)- NH_2 , the MIL-68(In)- NH_2 /GO composite demonstrated much greater photocatalytic activity. After 120 and 210 minutes of irradiation, respectively, 93% degradation and 80% TOC removal toward AMX were possible using 0.6 g per L MIL-68(In)- NH_2 /GO and pH = 5. The modification by employing GO, which worked as a sensitizer to

increase visible light absorption in addition to acting as an electron transporter to reduce photogenerated carrier recombination, is responsible for the increased activity for MIL-68(In)- NH_2 /GO.¹⁶⁰ The photocatalytic mechanism for the degradation of AMX is shown in Fig. 9.

MIL-68(In)- NH_2 /GO's ECB and EVB have been estimated to be -0.40 and 2.03 eV vs. NHE, respectively. The photogenerated electrons on the CB of MIL-68(In)- NH_2 are beneficial for reducing the O_2 because the ECB of MIL-68(In)- NH_2 /GO (-0.40 eV) is more damaging than the O_2/O_2^- potential (-0.046 eV vs. NHE).¹⁶¹ Similarly, these electrons can also reduce O_2 to H_2O_2 ($\text{O}_2/\text{H}_2\text{O}_2$ is 0.915 eV vs. NHE), and by trapping one electron, the resulting H_2O_2 can be further converted into $\cdot\text{OH}$. Due to the lower positive potential of MIL-68(In)- NH_2 /GO's VB (2.03 eV) compared to the typical $\cdot\text{OH}/\text{OH}^-$ (2.38 eV vs. NHE), the

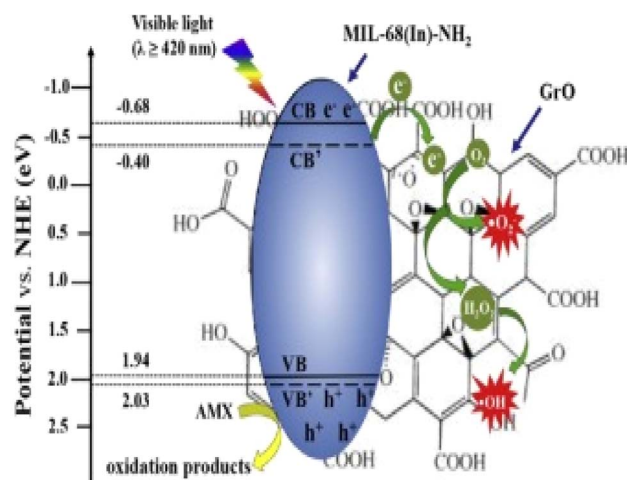


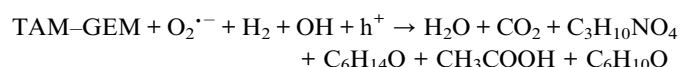
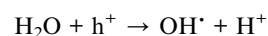
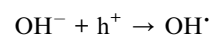
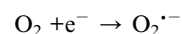
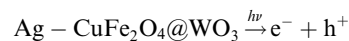
Fig. 9 Photocatalytic mechanism of MIL-68(In)- NH_2 /GO composite. This figure has been adapted from ref. 160 with permission from Elsevier, copyright 2017.



photogenerated holes on the VB of MIL-68(In)-NH₂/GO can simultaneously oxidize the AMX directly, but not the OH⁻ to form 'OH.¹⁶² Bastami *et al.* reported that magnetic Fe₃O₄/Bi₂WO₆ nanohybrids were created in two steps. A solvothermal technique was used to develop Fe₃O₄ nanospheres in a polyol medium, and a subsequent hydrothermal process produced Bi₂WO₆ nanocrystals. Tungstophosphoric acid hydrate (H₃PW₁₂O₄₀) was used as an acidic agent in the Bi₂WO₆ nanoparticle manufacturing. Ibuprofen (IBP) was photo-catalytically degraded from an aqueous solution under sunlight to assess the photocatalyst's activity. Particularly at pH 4.7, the as-prepared nanohybrid demonstrated significant efficiency in the photocatalytic destruction of IBP and was readily recovered by a magnet.¹⁶³ The OH[•] radical may be produced due to a chain reaction when hydrogen peroxide comes into contact with Fe(III) sites on the surface of the Fe₃O₄/Bi₂WO₆ photocatalyst.¹⁶⁴ First, the photocatalyst surface produces the Fe(III)H₂O₂ complex. Then, the generated species is transformed into Fe(II) and HO₂[•]. Additionally, Fe(II) species are produced on the photocatalyst surface. Every type of Fe(II) species produced on the photocatalyst surface has the potential to react with H₂O₂ and create reactive radicals (HO₂[•] and HO[•]) (Fig. 10).

The Fe₃O₄/Bi₂WO₆ photocatalyst's polyoxometalate-based materials (POM) are considered to increase photo-efficiency by inhibiting the rapid recombination of a charge pair (h⁺-e⁻) on nanohybrids and generating a potent oxidant O₂^{•-}.¹⁶⁵ When sunlight irradiation excites Fe₃O₄/Bi₂WO₆, the electron shifts from its conducting band to the POM's LUMO; following the attack of radicals on organic molecules, the adsorbed O₂ and H₂O₂ can readily trap an electron in the LUMO of the POM anion to produce the oxidizing species OH[•]. Consequently, the enhanced photocatalytic activity can slow down the rapid recombination of photo-induced electrons and holes. Sayadi *et al.* produced a novel Ag-CuFe₂O₄@WO₃ magnetic photocatalytic nanosystem and examined how well it degraded two drug pollutants, gemfibrozil (GEM) and tamoxifen (TAM), when

exposed to UV light. At pH 5, a photocatalyst dosage of 0.2 g L⁻¹, a drug concentration of 5 mg L⁻¹, and a contact time of 150 minutes, the highest levels of photodegradation for GEM (81%) and TAM (83%) were attained. Excellent catalytic efficiencies were produced by the slow electron-hole recombination brought about by the photocatalyst's heterostructure. The following reactions can be used to characterize the primary mechanism of the drug models' photodegradation process, and the mechanism can be seen in Fig. 11.¹⁶⁶



Xie *et al.* synthesized Fe₃O₄/BiOI magnetic photocatalyst to break down sulfadiazine sodium, which is used to treat infectious disorders. As mentioned above, the photocatalyst activity was about 80%. It was discovered that adding iron oxide to the nanocomposite enhanced the separation of electrons and holes and gave it magnetic properties that made it easier to separate after drug degradation.¹⁶⁷ Kumar *et al.* reported N-TiO₂@-SiO₂@Fe₃O₄ magnetic nanocomposite for prescription photodegradation ibuprofen, carbamazepine, and benzophenone-3 (UV filter). The combination exhibited superior degradation against ibuprofen (94%), benzophenone-3 (93%), and carbamazepine (71%). The remarkable photodegradation efficiency has been ascribed to the modification of the electronic band structure caused by nitrogen doping. Following the degradation phenomena, the nanocomposite produced easy separation (95% in 20–25 minutes) under an external magnetic field.¹⁶⁸ Khan *et al.* described the synthesis of a magnetic photocatalyst BiOBr/Fe₃O₄@SiO₂ by solvothermal method for large-scale ibuprofen decontamination. Ibuprofen was nearly wholly (99%) degraded, and the nanocomposite was recycled using an electromagnet in under five minutes, demonstrating the photocatalyst's high stability.¹⁶⁹ Additionally, Fe₃O₄/ZnO has been synthesized to eliminate four distinct antibiotics. All four antibiotics showed photocatalytic efficiencies of greater than 90%. Additionally, it was noted that the composite has comparable photocatalytic capability in subsequent reuse cycles, indicating its stability and reusability.¹⁷⁰

Nasseh *et al.* reported a magnetic FeNi₃@SiO₂@ZnO nanocomposite to remove tamoxifen from wastewater successfully. ZnO was the primary component of this composite because of its appropriate bandgap. FeNi₃ was a magnetic core, and the silica layer shielded the magnetic core from oxidation and light.¹⁷¹ There have been reports of the photocatalytic

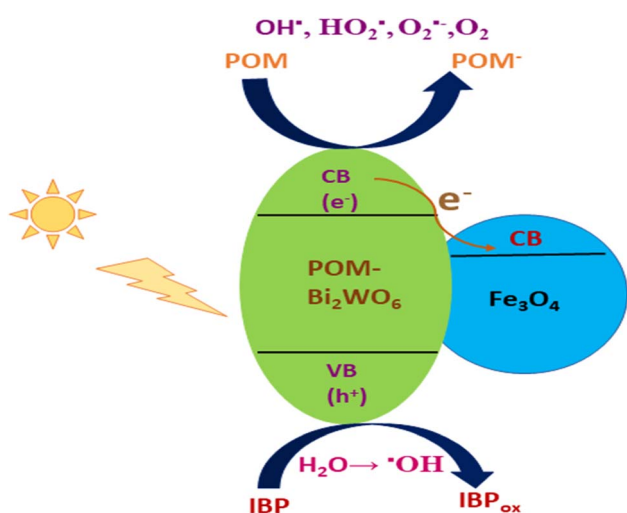


Fig. 10 A schematic illustration of the IBP photocatalytic reaction on the Fe₃O₄/Bi₂WO₆ nanohybrid.



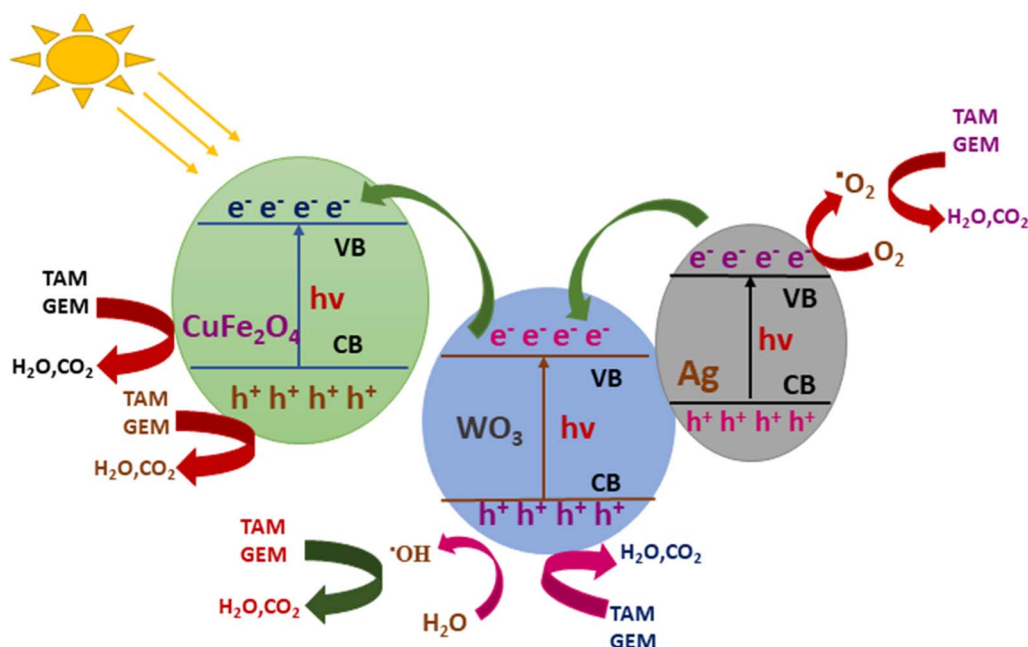


Fig. 11 Photodegradation mechanism of the Ag–CuFe₂O₄@WO₃ nanoparticles in the drug models.

breakdown of anticancer medications using the magnetically separable nanocomposite Fe₃O₄/Au. Perhaps because of its magnetic iron core, the nanocomposite, as mentioned above, was easily separated by a simple magnet. However, gold nanoparticles stopped iron oxide from oxidizing and clumping together. After five cycles, the photocatalyst demonstrated extraordinary reusability and outstanding photocatalytic effectiveness for imipenem (84%) and imatinib (82%), both exposed to visible light.¹⁷² Ahmadvour *et al.* synthesized ZnFe₂O₄@TiO₂/Cu using a solvothermal process and applied it against naproxen photodegradation. Comparing the ZnFe₂O₄@TiO₂/Cu (80.73%) nanocomposite to ZnFe₂O₄@TiO₂ (73.81%) and ZnFe₂O₄ (50.58%), the former demonstrated greater photoactivity. Cu particles have been associated with increasing ZnFe₂O₄@TiO₂/Cu's photocatalytic effectiveness by decreasing electron/hole pair recombination and producing more free radicals.¹⁷³

Wang *et al.* described the SnFe₂O₄/ZnFe₂O₄ photocatalyst that is active in visible light for the degradation of TC from aqueous media. Because of its ideal magnetic saturation value of 24.8 emu g⁻¹, this magnetic heterojunction photocatalyst demonstrated 93.2% degrading ability and simple separation from aqueous solution. According to the theorized process, the holes in the valence band of SnFe₂O₄ were neutralized by the photoinduced electrons in ZnFe₂O₄, and these electron–hole pairs produced potent radicals that degraded tetracycline.¹⁷⁴ Xu *et al.* revealed a Fe₃O₄/BiVO₄/CdS nanocomposite that is visible light active and can photocatalytically break down tetracycline. They found that combining BiVO₄ and CdS reduced the rate at which photoinduced charge carriers reacted, improving the nanocomposite's capacity to absorb light and increasing its surface area. On the other hand, Fe₃O₄ loading enabled its high separation *via* the magnetic field. These characteristics in one

material have been linked with exceptional photocatalytic performance (87%) against tetracycline. Even after five runs, the stability and recyclability were determined to be above 81% in catalytic efficiency.¹⁷⁵

Gabelica *et al.* showed that magnetite-based MNPs were effectively created utilizing a quick microwave-assisted synthesis method. The Fe₃O₄/SiO₂/TiO₂ nanocomposite was a potential option for ciprofloxacin degradation. After three consecutive replicates, Fe₃O₄/SiO₂/TiO₂ showed a remarkable 94% photodegradation efficiency and remained intact. The loading of Fe₃O₄ also confirmed the straightforward catalyst recovery using a magnet.¹⁷⁶ ZnFe₂O₄/carbon derivatives (ZnFe₂O₄@CNTs, ZnFe₂O₄@GO, and ZnFe₂O₄@fullerene) have been reported for the photocatalytic degradation of norfloxacin. Carbon moieties enhanced light absorption in the visible spectrum because of the carbon derivatives' conjugated π – π structure and efficient electron trapping. However, ZnFe₂O₄@carbon nanotubes exhibited the highest degradation efficiency (91.36%) against norfloxacin because of the outstanding charge carrier separation efficiency brought about by the good interfacial connection between ZnFe₂O₄ and CNTs.¹⁷⁷ Another visible light-active nanocomposite that had been synthesized for the efficient breakdown of 94% tetracycline is Fe₃O₄/g-C₃N₄/MoO₃. The heterojunction creation between g-C₃N₄ and MoO₃ has been attributed to this exceptional efficiency, as it has improved the segregation of charge carriers and increased the production of active species.¹⁷⁸ Using a hydrothermal process, a different magnetic composite (g-C₃N₄/NiO/ZnO/Fe₃O₄) was synthesized for the photocatalytic degradation of esomeprazole at 95.05%. The synergistic contribution of the individual composite components that provided the structural and optical alterations is the only reason for this efficiency. The creation of light-induced charge carriers, their separation, and decreased





Table 2 MNPs for the removal of different contaminants

| MNPs | Contaminant removed | Removal mechanism | Contact time | Removal efficiency | Saturation magnetization (M_s) | Size (nm) | Reusability | Ref. |
|---|-------------------------------------|----------------------------|--------------|--------------------|--|---------------------|--|------|
| Humic acid functionalized Fe_3O_4 | Carcinogenic malachite green dye | Adsorption | 35 min | 97% | 70.05 emu g^{-1} | 5–15 nm | Up to five cycles for 85% removal | 181 |
| Combination of Fe_3O_4 and Fe_2O_3 | Bromophenol blue (BPPB) | Photocatalytic degradation | 60 min | 98% | — | 10–80 nm | Up to three cycles for 95% removal | 182 |
| Superparamagnetic iron-oxide nanoparticles (SPION) and SPION/ β -CD | Bisphenol A and malachite green dye | Photocatalytic degradation | 10 min | 82.5% | 68.34 emu g^{-1} | 10–15 nm | 70% removal for first five cycles | 183 |
| | | | | | 39.21 emu g^{-1} | | Up to five cycles for 90.13% removal | |
| Epoxy-triazinetrione-functionalized MNPs (Fe_3O_4 -ETT) | Malachite green (MG) | Adsorption | 15 min | 95% | 30.7 emu g^{-1} | 25–52 nm | Up to six cycles for 61% removal | 184 |
| JC- Fe_3O_4 and CT- Fe_3O_4 NPs | Organic dyes | Adsorption | 120 min | 99% | 38.46 emu g^{-1} | 20–42 nm | — | 185 |
| | | | | | 34.35 emu g^{-1} | 26–35 nm | — | |
| Fe_3O_4 @NiO core shell MNPs | Alizarin red S dye | Adsorption | 60 min | 96.58% | 24.7 emu g^{-1} | 65 \pm 10 nm | Up to six cycles for 43.93% removal | 186 |
| MAPE-mediated iron oxide nanoparticles (MION) | Sunset yellow dye | Adsorption | 180 min | 92% | 44.14 emu g^{-1} | 20–60 nm | Up to five cycles for 75.4% recovery | 187 |
| Yeast cells immobilized Fe_3O_4 MNPs | Methyl orange dye | Biosorption | 140 min | 96.52% | — | 87.46–231.4 nm | — | 188 |
| Copper-doped ZrO_2 MNPs | Methyl orange dye | Photocatalytic degradation | 100 min | 98% | — | 20–45 nm | Up to four cycles for 90% removal | 189 |
| MCPEI (polyethyleneimine and magnetic nanoparticle) | Black 5 dye | Adsorption | 180 min | 100% | 3.3 emu g^{-1} (before adsorption) | 1.92 nm (pore size) | Up to five cycles for 48% removal | 190 |
| | | | | | 3.2 emu g^{-1} (after adsorption) | | | |
| Zeolite/ Fe_3O_4 | Basic violet 16 (BV16) | Adsorption | 45 min | 99% | 18.11 emu g^{-1} | 1–4 nm | Up to five cycles for 64.17% removal | 191 |
| Fe_3O_4 MNPs | Methylene blue (MB) | Adsorption | 60 min | 88.8% | 49.48 emu g^{-1} | 21–32 nm | Up to four cycles for 77% removal | 192 |
| Magnesium oxide modified magnetic nanoparticles (MgO@MNPs) | Eriochrome black T (EBT) dye | Adsorption | 5 min | 98% | 11.70 emu g^{-1} (before adsorption) | 33.33–50.80 nm | Up to seven cycles for 96% removal | 193 |
| | | | | | 10.62 emu g^{-1} (after adsorption) | | | |
| Fe_3O_4 MNPs | Anionic azo dye | Adsorption | 120 min | 99.99% | — | 50 nm | — | 194 |
| Magnetic core-shell Fe_3O_4 @ SiO_2 nanoparticles | Methyl orange | Adsorption | 30 min | 83% | 72.08 emu g^{-1} | 120 nm | Reusable up to 5 cycles | 195 |
| Magnesium doped $CoFe_2O_4$ NPs | Methylene blue (MB) dye | Adsorption | 30 min | 98.6% | 15.2971 emu g^{-1} | 7.9 nm | Up to five cycles 82.5% removal for MB and 79.6% removal for RhB | 196 |
| | Rhodamine B (RhB) dye | | | 95.3% | | | | |



Table 2 (Contd.)

| MNPs | Contaminant removed | Removal mechanism | Contact time | Removal efficiency | Saturation magnetization (M_s) | Size (nm) | Reusability | Ref. |
|---|---|-------------------|----------------------|--|--|----------------------|---|------|
| Nanoparticles of titanomagnetite (NTM) | Methylene blue (MB) | Adsorption | 20 min | — | 62 emu g^{-1} | 35 nm | Reusable for up to three successive cycles | 197 |
| $\text{Fe}_3\text{O}_4\text{@Ag@MESNa}$ | Hg(II) | Adsorption | 30 s | 100% | — | — | Up to two cycles for 100% removal | 198 |
| Magnetic eggshell membrane (MESM) | Lead | Adsorption | 48 h | 95% | 66 emu g^{-1} | ~50 nm | Reusable | 199 |
| $\text{Fe}_3\text{O}_4\text{@EPS}$ | PO_4^{2-} | Adsorption | 13 h | 91% | 5.0 emu g^{-1} | 10–20 nm | — | 200 |
| Magnetic tubular carbon nanofibers (MTCFs) | Cu(II) | Adsorption | 10 min | $99.9 \pm 0.1\%$ | 10.65 emu g^{-1} | 25–110 nm | Up to six cycles for 85% removal | 201 |
| Magnetic nanocomposite activated hydrochar ($\text{Fe}_3\text{O}_4\text{-ACH}$) | Cr(VI) | Adsorption | 120 min | 94.1% | 0.110334 emu g^{-1} | 0.7–1.5 nm | Up to five cycles for 82.7% removal | 202 |
| Cabuya fibers impregnated magnetite nanoparticles (FC-MNPs) and MNPs | Hg(II) | Adsorption | 4 h | 93% | — | 19 nm | — | 203 |
| JC- Fe_3O_4 and CT- Fe_3O_4 NPs | Co^{2+} and Cu^{2+} | Adsorption | 120 min | 90% | 38.46 emu g^{-1} 34.35 emu g^{-1} | 20–42 nm 26–35 nm | — | 185 |
| Hyperbranched polyglycerol HPG-MNPs | Ni, Cu and Al | Adsorption | 130 min | 94% | –77 to +77 emu g^{-1} | 21–30 nm | Up to nine cycles for 90% removal | 204 |
| Chitosan-coated magnetic nanoparticles (cMNPs) | Bio-refinery wastewater containing heavy metals (Cu, Cr, and As) | Adsorption | 90 min | Cu (42.2%), Cr (18.7%), and As (2.44%) | 26.96 emu g^{-1} | — | Up to five cycles for 20% of initial removal efficiency | 205 |
| Superparamagnetic $\text{Fe}_3\text{O}_4\text{@SiO}_2\text{@GLYMO(S)-en}$ | Pb^{2+} and Cd^{2+} | Adsorption | 55 min | — | 40 emu g^{-1} | 16 nm | Up to five cycles for 90% removal | 206 |
| Epoxy-triazinetriene-functionalized MNPs ($\text{Fe}_3\text{O}_4\text{-ETT}$) | Pb(II) | Adsorption | 15 min | 95% | 30.7 emu g^{-1} | 25–52 nm | Up to six cycles for 61% removal | 184 |
| Eucalyptus leaf extracts (EL-MNP@zeolite) | PO_4^{2-} NH_4^+ | Adsorption | 30 min | 99.8% 43.3% | — | 60 nm | — | 207 |
| MNP@ $\text{SiO}_2\text{-PEI-DTPA}$ | Pb^{2+} Cd^{2+} | Adsorption | 72 h | >90% | 85 emu g^{-1} | 25–150 nm | Up to five cycles for 80% removal | 208 |
| Nano-chitosan coated MNPs | Pb(II) Cu(II) Cd(II) Pb^{2+} | Adsorption | 10–30 min | 57.6–95.5% | — | 9.32–20.57 nm | Up to three cycles for 96% removal | 209 |
| C- $\text{CoFe}_2\text{O}_4/\text{N}$, S-BC | Pb^{2+} | Adsorption | 30 min | 99.1% | — | — | Up to eight cycles with >80% removal efficiency | 210 |
| $\text{Fe}_3\text{O}_4/\text{CNTs}$ nanocomposites | Pb^{2+} | Adsorption | 60 min | 91.64% | 8.9 and 16.4 $\text{A m}^2 \text{kg}^{-1}$ | 2–30 nm | Reusable up to six cycles | 211 |
| Ionic-liquid modified MNPs ($\text{Fe}_3\text{O}_4\text{@SiO}_2\text{-[C}_3\text{C}_1\text{Im][Cl NPs]}$) | Platinum (Pt) Palladium (Pd) | Adsorption | 15 min 1 h 6 h | 22–24% 90–93% | 51 emu g^{-1} | 65 nm | — | 212 |
| DES/ $\text{GO-Fe}_3\text{O}_4$ nanohybrid | Pb(II) | Adsorption | 20 min | 80% | — | 6–30 nm | — | 213 |
| Magnetic nanoparticles coated mixed fungal biomass (MNP-FB) | Cr(VI) | Adsorption | 60 min | 99.4% | — | — | — | 214 |



Table 2 (Contd.)

| MNPs | Contaminant removed | Removal mechanism | Contact time | Removal efficiency | Saturation magnetization (M_s) | Size (nm) | Reusability | Ref. |
|--|--|----------------------------|--------------|----------------------------|-------------------------------------|----------------------------|---|------|
| AG-g-PAN/Cu Fe ₂ O ₄ | Pd(II) | Adsorption | 30–120 min | 97.01% | 12 emu g ⁻¹ | 37 nm | Up to three cycles for 93.98% removal | 215 |
| Fe ₃ O ₄ and Fe ₃ O ₄ -chitosan NPs | Pb(II) | Adsorption | 31.2 min | 69.02 and 89.54% | 51.22 and 27.86 emu g ⁻¹ | 30 nm | Up to five cycles for 59.43% removal | 216 |
| Fe ₃ O ₄ @DA-DMSA magnetic nanoparticles (FDDMs) | Pb ²⁺ Cu ²⁺ Cd ²⁺ | Adsorption | 45 min | 90.11% 40.61% 13.87% | 21.50 emu g ⁻¹ | 446 nm | — | 217 |
| Core shell magnetic nanoparticles (cMNPs) | As(V) | Adsorption | 10 min | 97.5% | 36.4 emu g ⁻¹ | 4.9 nm | Up to five cycles for 80% removal | 218 |
| Fe ₃ O ₄ @ZIF-8 and Fe ₃ O ₄ @SiO ₂ @ZIF-8 | Pb ²⁺ | Adsorption | 15 min | 93.61% 99.2% | — | 110 ± 10 nm and 80 ± 10 nm | — | 219 |
| Magnetic hybrid alumina nanoparticles (MHAI-NPSP) | Cu(II) | Adsorption | 60 min | 85% | — | — | — | 220 |
| Gum arabic magnetic nanoparticles (GA-MNPs) | Ciprofloxacin (CIP) | Adsorption | 120 min | 96.30% | — | — | Up to seven cycles for initial removal efficiency of 96.34% | 221 |
| Fe ₃ O ₄ -ACLM | Ciprofloxacin (CIP) | Adsorption | 75 min | 100% | 37.6 emu g ⁻¹ | 9.8 ± 1.5 nm | Up to six cycles for 90% removal | 222 |
| C-CoFe ₃ O ₄ /N, S-BC | Ciprofloxacin (CIP) | Adsorption | 30 min | 94.7% | — | — | Up to eight cycles with >80% removal efficiency | 210 |
| Fe ₃ O ₄ @MIL-125-NH ₂ | (Pesticides) carbaryl Methomyl | Photocatalytic degradation | 150 min | 91% | — | 24.57 nm | — | 223 |
| Activated carbon magnetic nanoparticles (AC-MNPs) | Metronidazole (MTZ) | Adsorption | 90 min | 98% 100% | 24.93 emu g ⁻¹ | 20–35 nm | Up to five cycles for 90.5% removal | 224 |
| PAC@Fe ₃ O ₄ -MN | Ciprofloxacin (CIP) | Adsorption | 60 min | 100% | 51.5 emu g ⁻¹ | <80 nm | Up to eight cycles for 96% removal | 225 |
| MOM-Fe ₃ O ₄ | Pharmaceutical substance (metformin) | Adsorption | 720 min | 93.9% | — | — | — | 226 |
| Fe ₃ O ₄ /clinoptilolite | Tetracycline | Adsorption | 20 s | 98.6% | 30.0 emu g ⁻¹ | 60–80 nm | Up to four cycles for 80.6% removal | 227 |
| Fe ₃ O ₄ @gly@indole@CuNO ₃ magnetic nanoparticle (FGICu) | Tetracycline | Adsorption | 60 min | 95% | 38.4 emu g ⁻¹ | 17.8 nm | Up to five cycles for 70% removal | 228 |
| Magnetic sporopollenin supported magnesium nanoparticles (MSP@MgO) | Tetracycline | Adsorption | 180 min | 98.05% | ~32 emu | — | Up to 15 cycles for 84.72% removal | 229 |
| Magnetic Ho ₂ MoO ₆ /Fe ₂ O ₃ | Tetracycline | Adsorption | 90 min | 99.96% | 14.82 emu g ⁻¹ | 25–50 nm | Up to five consecutive cycles for 79.1% removal | 230 |
| FeNi ₃ @SiO ₂ @TiO ₂ | Humic acid (HA) | Adsorption | 30 min | 100% | 18.84 emu g ⁻¹ | 20–30 nm | — | 231 |
| Nano porous Co ₂ O ₃ /Cu ₂ O ₃ ·Al ₂ O ₃ ·SiO ₂ | <i>E. faecalis</i> | Photocatalytic degradation | 05 min | 100% | 0.157–0.24 emu g ⁻¹ | 23–31 nm | — | 232 |
| Chitosan-coated magnetic nanoparticles (cMNPs) | Lignocellulosic biorefinery containing phenol | Adsorption | 90 min | 46.2% | 26.96 emu g ⁻¹ | — | Up to five cycles for 20% of initial removal efficiency | 205 |



Table 2 (Contd.)

| MNPs | Contaminant removed | Removal mechanism | Contact time | Removal efficiency | Saturation magnetization (M_s) | Size (nm) | Reusability | Ref. |
|---|-----------------------------|--|--------------|---|--|---------------------|--------------------------------------|------|
| Novel MNP-alum conjugate | Natural organic matter | Adsorption cum enhanced coagulation-flocculation | 30 min | 98.7% | — | 20–30 nm | Up to five cycles with ~91% removal | 233 |
| NFO@SiO ₂ @APTS | Acetaminophen | Adsorption | 30 min | 94% | 22 emu g ⁻¹ | 130–150 nm | Up to four cycles for 89% removal | 234 |
| Magnetic mesoporous silica microspheres (MSMS) | Acetaminophen | Adsorption | 30 min | 97.4% | 8.3 emu g ⁻¹ | 26 nm | Reusable for up to 14 cycles | 235 |
| Fe ₃ O ₄ /Douglas fir biochar | Acetylsalicylic acid | Adsorption | 5 min | 70% | — | 12.3 ± 7.1 nm | Reusable for up to five cycles | 236 |
| NFO@SiO ₂ @APTS | Ibuprofen | Adsorption | 30 min | 97% | 22 emu g ⁻¹ | 130–150 nm | Up to three cycles for 92% removal | 234 |
| Magnetic activated carbon-Fe ₃ O ₄ (AC-Fe ₃ O ₄) | Promazine | Adsorption | 06 min | 99.97% | — | 20 nm | Up to five cycles for 99% removal | 237 |
| Magnetic silica-based nanoadsorbents (MMST and MMST-Ph) | Pharmaceutical substance | Adsorption | 200 min | 80% | 16.0 emu g ⁻¹ (MMST) 18.6 emu g ⁻¹ (MMST-Ph) | 3.75 nm (pore size) | Up to eight cycles for 100% removal | 238 |
| MNPs coated with rhamnolipids (Rh-cMNP) | Acetaminophen | Adsorption | 60 min | 96.7% | 16.5 emu g ⁻¹ (before adsorption) 8.2 emu g ⁻¹ (after adsorption) | — | Up to eight cycles for | 239 |
| Fe ₃ O ₄ @GNP | Ketoprofen (KP) | Adsorption | 30 min | For wastewater (65.0–83.6%) For tap water (67.1–90.0%) | — | 50 nm | Reusable for up to five cycles | 240 |
| Ibuprofen (IP) | Ibuprofen (IP) | Adsorption | 5–150 min | 94.42% | — | 150–250 nm | Up to five cycles for 73.56% | 241 |
| Rape straw biomass fiber-β-CD-Fe ₃ O ₄ (RSBCDF) | Naproxen (NX) | Adsorption | 30 min | For wastewater (65.0–83.6%) For tap water (67.1–90.0%) | — | 50 nm | Reusable for up to five cycles | 240 |
| Fe ₃ O ₄ @GNP | Diclofenac sodium salt (DF) | Adsorption | 30 min | Up to 80% | 66.815 emu g ⁻¹ | — | Up to four cycles for 74.25% | 242 |
| Biochar/MgFe ₂ O ₄ | Levofloxacin (LFX) | Adsorption | 5 min | 70% | — | 12.3 ± 7.1 nm | Reusable for up to five cycles | 236 |
| Fe ₃ O ₄ /Douglas fir biochar | Caffeine | Adsorption | 1 h | 85% | 55.4 emu g ⁻¹ | 80–90 nm | Up to five cycles for 85% removal | 243 |
| 3D-GC/MGO-SO ₃ H | Tetracycline | Adsorption | 40 min | 96.38% | — | 290.2 nm | Up to three cycles for 94% removal | 244 |
| Magnetic polydopamine-chitosan modified adsorbent (MPDA-CS) | Levofloxacin (LEV) | Adsorption | 24 h | 86.15% | — | 14.34 nm | Up to four cycles for 80.47% removal | 245 |



Table 2 (Contd.)

| MNPs | Contaminant removed | Removal mechanism | Contact time | Removal efficiency | Saturation magnetization (M_s) | Size (nm) | Reusability | Ref. |
|--|---------------------------|----------------------------|---------------------|--------------------|------------------------------------|----------------|--|------|
| Magnetic molecularly imprinted polymers (MMIP) | Norfloxacin | Adsorption | — | — | — | — | Reusable | 246 |
| Fe-TiO ₂ @Fe ₃ O ₄ | 2,4-Dichlorophenol | Sono-catalytic degradation | 90 min | 94% | 25 emu g ⁻¹ | 15.05–18.12 nm | Up to five cycles for 91% removal | 247 |
| Fe ₃ O ₄ @Phe NPs | Ciprofloxacin (CIP) | Adsorption | 30 min | 62.5% | 60 emu g ⁻¹ | 150–270 nm | Up to four cycles for 80% removal | 248 |
| Zn _{0.5} Ni _{0.5} FeCrO ₄ MNPs | 4-Nitrophenol | Photo-degradation | 180 min and 150 min | 95% and 80% | 7.06 emu g ⁻¹ | 25–35 nm | Up to four cycles reusable | 249 |
| Polyvinylidene fluoride (PVDF)/Fe ₃ O ₄ @carboxymethyl cellulose | Aniline | Sorption | 30 min | — | 62–55 emu g ⁻¹ | 20 nm | Up to five cycles reusable | 250 |
| Magnetic Janus nanoparticles (M-Janus NPs) | Oil sludge (oleic acid) | Phase-separation | 15 min | 96% | — | 2 mm | Up to five cycles for 90% removal | 251 |
| Surface-modified MNPs, AU-MNPs, and AP-MNPs | Cooking oil and crude oil | Adsorption | 3–20 min | 98% | 40.23 emu g ⁻¹ | 141.8 nm | Up to four cycles for 89% removal | 252 |
| VCL/Fe ₃ O ₄ and VAA/Fe ₃ O ₄ | Oil spills (crude oil) | Adsorption | 5–20 min | 98% | 43.85 emu g ⁻¹ | 133.4 nm | Up to four cycles for 80% removal | 253 |
| Fe ₃ O ₄ MNP | Crude oil | Adsorption | 5–30 min | 99.3% | 54.6 emu g ⁻¹ | 117.5 nm | Up to nine cycles for 64% removal | 254 |
| Fe ₃ O ₄ MNP | Crude oil | Adsorption | — | 94% | 49.4 emu g ⁻¹ | 3 nm | Up to 7 cycles for 90% removal | 255 |
| Nanoparticles of titanomagnetite (NTM) | Crude oil spills | Adsorption | 20 min | 83% | 51 ± 0.4 emu g ⁻¹ | 15–20 nm | Reusable for up to three successive cycles | 197 |
| HAN-Fe ₃ O ₄ | Crude oil | Adsorption | 8 min | 100% | 62 emu g ⁻¹ | 35 nm | Up to five cycles for reusable | 256 |
| HAT-Fe ₃ O ₄ | Crude oil (API = 23.3°) | Adsorption | 30 min | 89% | — | 10 nm | Up to 10th cycle with 41.98 g per g oil sorption capacity | 257 |
| Magnetic corn cob powder (CoFe ₂ O ₄) | Crude oil (API = 29.7°) | Adsorption | — | — | 32 emu g ⁻¹ | 30 nm | — | — |
| DOTAC-coated MNPs (M3) | Fresh motor oil | Adsorption | 3 min | 94% | ~26.0 emu g ⁻¹ | — | Up to 7 cycles in W/O emulsion with >90% efficiency | 258 |
| MNP-S@CTAB | Used motor oil | Adsorption | 20 min | 100% | — | 15–20 nm | Up to 10th cycles with 19 mg g ⁻¹ adsorption efficiency | 259 |
| Silica coated ferroferric oxide (Fe ₃ O ₄ @SiO ₂) | Kerosene | Transformation | 05 min | 98% | 69.1 emu g ⁻¹ | 50–100 nm | Up to five cycles for 95% removal | 260 |

recombination—all of which ultimately resulted in exceptional photocatalytic activity are appropriate for these modifications.¹⁷⁹ According to the structural analysis, the synthesized $\text{AgCuFe}_2\text{O}_4/\text{MC}/\text{AC}$ catalyst was created on a nanoscale with a crystalline structure, large surface area, appropriate magnetic characteristics, and optical activity. High efficiency in photo-degrading TC, which is 90.91% from synthetic solutions and 87.17% from actual wastewater, was demonstrated by the synthesized catalyst under ideal circumstances (neutral pH, starting TC content of 5 mg L^{-1} , nano-photocatalyst dosage of 0.5 g L^{-1} , one and half hour). The results showed that the heterogeneous magnetic nano-photocatalyst $\text{AgCuFe}_2\text{O}_4/\text{MC}/\text{AC}$ may be utilized successfully to clean wastewater from hospitals and pharmaceutical enterprises because of its strong capacity to degrade antibiotics, notably TC.¹⁸⁰ Table 2 represents the MNPs used in different environmental applications.

3.5 Scalability challenges

Scalability challenges encompass several factors, including the recovery rates and reusability of nanoparticles, treatment costs, and the economic feasibility of the processes involved, as well as the influence of complex environmental conditions such as high salinity and low temperatures. Among these, the reusability of MNPs stands out as a critical aspect. The economic viability of MNPs is significantly tied to their ability to be reused; however, issues like surface fouling and structural degradation can diminish their adsorption efficiency after multiple cycles.²⁶¹ For MNPs to be commercially feasible in large-scale applications, it is essential to develop more sustainable and efficient regeneration methods. Enhancing the reusability of MNPs can be achieved by exploring strategies that minimize nanoparticle degradation, such as enzymatic treatments or low-temperature regeneration techniques.¹⁸ Metal-organic frameworks and metal-based adsorbents exhibit high adsorption capacities due to the nature of adsorption through metal complexation. However, challenges remain regarding the recovery, stability, recycling, and potential metal leaching associated with these materials.²⁶² Furthermore, the economic feasibility of large-scale applications may be hindered by the energy costs associated with magnetic recovery, necessitating further optimization.²⁶³ Recent literature reveals a shift from traditional carbon-based materials to various promising alternatives that are not only more cost-effective for industrial applications but also demonstrate comparable performance to carbon-based adsorbents.²⁶⁴

Despite the exploration of advanced adsorbent materials such as ceramics, MOFs, graphene oxide, and composite adsorbents in small-scale studies using simulated hospital wastewater or simplified water sources, their long-term effectiveness in real hospital wastewater (HWW) remains largely unexamined. Currently, only granular and powdered activated carbons have seen widespread application. Challenges such as high operational costs, non-specific adsorption, degradation of material and adsorption capacity due to frequent regeneration, and the management of spent adsorbents pose significant economic and environmental hurdles. To address these issues,

the development of innovative adsorbent materials that offer high adsorption capacity, durability, and ease of regeneration is essential for effective full-scale wastewater treatment.²⁶⁵ While MOF-derived magnetic nanocomposites offer numerous advantages, such as significant structural and chemical flexibility, the costs involved in their synthesis still require optimization. This includes expenses related to raw materials, like metallic salts and organic binders, as well as processing methods that often utilize non-recyclable organic solvents and activation techniques. One potential solution is to adopt aqueous synthesis in place of solvothermal methods, which could lead to a cost reduction of approximately US\$ 35 for every US\$ 13 spent on synthesizing MOFs at a large scale.²⁶⁶ For industrial applications to be economically viable, the cost of these materials should ideally be below \$10 per kilogram. Consequently, it is essential to focus on less expensive organic ligands, advocating for the use of shorter and smaller binders rather than more expensive longer or aromatic alternatives.²⁶⁷ Nevertheless, there is a notable gap in the literature regarding cost evaluations, which are essential for assessing economic viability at scale. Research is increasingly focused on analyzing the economic feasibility of pollution control methods that incorporate these nanocomposites, taking into account production, operational costs, waste management, and material regeneration.²⁶⁸

Despite the advantages offered by nanostructured adsorbents, several challenges hinder their broader application, primarily their high production costs and difficulties in scaling.²⁶⁹ The expense associated with manufacturing these materials renders them impractical for large-scale water treatment initiatives. Additionally, while nanostructured adsorbents have demonstrated effectiveness in laboratory settings, translating this success to larger systems remains problematic. Key considerations, such as the efficient distribution and maintenance of these adsorbents in real-world applications, must be thoroughly addressed before they can be adopted on a wider scale.²⁷⁰

In light of these challenges, further research and development are essential to address the existing issues and enhance the potential benefits of nanostructured adsorbents in water treatment applications. These materials have proven effective in various wastewater treatment scenarios, including the removal of organic pollutants, dyes, and heavy metals. For instance, surface-modified iron oxide nanoparticles have successfully eliminated toxic elements such as lead and arsenic from industrial wastewater. Additionally, carbon-based nanostructured materials, including carbon nanotubes and modified graphene, have demonstrated significant efficacy in extracting organic contaminants and dyes due to their high surface area and adsorption capacity. These practical applications underscore the versatility and effectiveness of nanostructured adsorbents in tackling diverse challenges in wastewater treatment.

A thorough research into the use of nanostructured adsorbents for the continuous treatment of sewage, industrial effluents, and polluted water bodies is essential. While surface-functionalized nanostructured adsorbents have demonstrated effectiveness in laboratory settings, their application in



industrial environments is limited due to high costs and complex modification processes. To promote advancements in wastewater treatment technologies, it is crucial to focus on developing cost-effective preparation methods, selectively modifying surface functional groups, and further investigating pollutant removal mechanisms, recovery of valuable elements from wastewater, and the recycling of adsorbents.²⁷¹ Scalability challenges such as high salinity in wastewater and low temperatures can significantly hinder the effectiveness of MNPs in large-scale environmental remediation efforts. These conditions can diminish treatment efficiency, compromise material stability, and reduce adsorption capabilities. The presence of elevated salt concentrations, commonly found in oilfield brines, desalination processes, and various industrial effluents, poses a particular obstacle to the scalability of magnetic nanohybrid materials. Salts like NaCl, MgCl₂, and CaSO₄ interfere with the electrostatic interactions between contaminants and nanomaterials, disrupting surface chemistry and leading to decreased adsorption effectiveness due to competition for active sites.²⁷² Furthermore, chloride ions can promote the leaching of iron ions from magnetite, thereby destabilizing metal oxide nanoparticles and affecting their durability and magnetic separability. Additionally, elevated ionic strength can lead to the aggregation of nanoparticles, further reducing their reactivity and active surface area.²⁷³

The effectiveness of MNPs in environmental remediation can be significantly influenced by low ambient temperatures. Key processes such as diffusion rates, redox reactions, and the kinetics of pollutant adsorption, which are vital for nanohybrid-based remediation, are all temperature-dependent. For instance, semiconductor-metal nanohybrid-mediated photocatalytic degradation is notably less efficient at lower temperatures due to reduced activation energy and slower electron transfer rates. Additionally, colder temperatures increase the viscosity of wastewater, hindering the mobility of nanoparticles and their capacity to engage with contaminants. These thermal constraints pose challenges to the sustainability and cost-effectiveness of large-scale applications, making the deployment of MNPs in temperate or cold regions less feasible without additional energy support.²⁷⁴

The large-scale application of MNPs in real-world settings faces significant challenges due to the interplay of low temperatures and high salinity. These conditions can slow down reaction kinetics and promote the aggregation of nanoparticles, thereby reducing their effectiveness in pollutant removal. To address these issues, it is essential to develop more robust nanohybrids that maintain their dispersibility and catalytic performance across diverse environmental conditions, such as surface-functionalized or polymer-stabilized MNPs. Furthermore, this situation underscores the importance of conducting pilot studies and implementing site-specific adaptive designs before embarking on large-scale remediation efforts.²⁷⁵ Scalability is essential for enhancing productivity and transitioning laboratory-scale synthesis to large-scale operations, particularly in both batch and continuous processes. A key engineering challenge in designing modular reactors for large-scale water treatment lies in achieving uniform flow

dynamics and treatment efficacy across all modules. Typically arranged in parallel or series within decentralized systems, these modular reactors face theoretical difficulties in ensuring consistent flow distribution and hydraulic retention time. Such variations can lead to overloaded or underutilized reactor zones, ultimately diminishing pollutant removal efficiency. Additionally, non-uniform utilization of reactive media, such as MNPs, can hinder effective magnetic separation and result in unequal pollutant exposure, especially in high-flow industrial or municipal environments.²⁷⁶ While the application of nanoparticles in real-world scenarios is increasingly viable, challenges remain regarding their production and broader implementation. Most research is still conducted at the bench level, necessitating pilot-scale studies to validate findings. The scaling up of green nanoparticle production is promising, as larger-scale operations can provide more accurate data on cost-effectiveness, repeatability, and efficiency.

3.6 Environmental impact assessment

The increasing interest in magnetic nanohybrid materials highlights their unique characteristics and wide-ranging applications, especially in environmental remediation. While extensive research has been conducted on the efficacy and uses of MNPs, it is crucial to also assess their potential environmental impacts. An effective environmental impact assessment for MNPs should concentrate on three primary aspects: the release of harmful substances during their application, their influence on ecosystems, and the strategies for ensuring safe ecological remediation. A unified scientific framework is essential for researchers to evaluate sustainable practices associated with MNP applications. It is vital to investigate the potentially harmful substances that may be released during MNP usage, as their production involves various chemical compounds, some of which could be toxic. Furthermore, the degradation of MNPs can result in the release of nanoparticles, metal ions, or other byproducts that may endanger human health and the environment. For instance, a study by Osman *et al.* focused on synthesizing a highly active magnetic sorbent composite from mixed plastic and biomass waste, employing Life Cycle Assessment (LCA) to evaluate the environmental impacts of producing the magnetic char composite adsorbent, adhering to the methodologies and standards set forth in ISO 14040:2006 and ISO 14044:2006.^{277,278}

Ferric iron is primarily transported in the body by transferrin, which interacts with transferrin receptors on cell surfaces, while its storage occurs mainly in ferritin proteins within cells. Initially, MNPs were considered non-toxic due to their ability to release ferric iron into normal iron metabolism. However, their diminutive size enables them to accumulate in high concentrations within cells, complicating their elimination from the body.²⁷⁹ This accumulation raises significant concerns, particularly because excess free iron can produce harmful free radicals that threaten sensitive tissues, such as the brain. Research by Bucak *et al.* has shown that toxicity can emerge when serum proteins adhere to the surfaces of MNPs, thereby modifying the cellular environment. To mitigate this



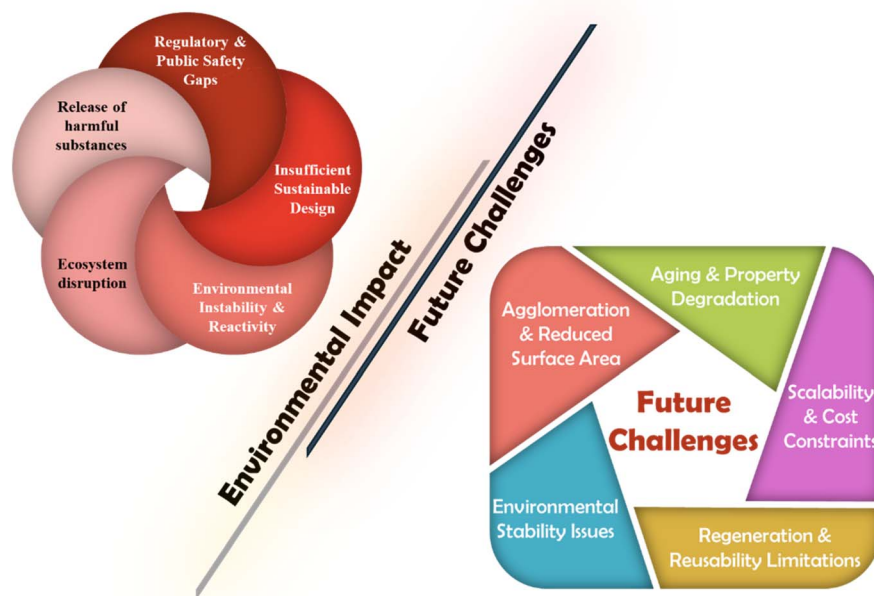


Fig. 12 Pictorial illustration of environmental impact and future challenges.

risk, coating MNPs with biocompatible materials can effectively reduce protein binding and surface interactions, ultimately lowering their toxicity.²⁸⁰

A comprehensive examination of the effects of MNPs on ecosystems is crucial due to their distinctive magnetic properties and large surface area, which allow them to interact with various environmental components such as water, soil, air, and living organisms. Although MNPs can effectively break down water pollutants, their degradation products may pose risks to aquatic ecosystems and disrupt food webs. Furthermore, the application of MNPs in soil remediation can modify microbial communities, potentially hindering plant growth. To assess these impacts, thorough ecotoxicological evaluations are essential. Establishing best practices and guidelines for the safe use of MNPs in environmental remediation is vital, emphasizing the principles of green chemistry and sustainable engineering in their synthesis and design. This approach should focus on utilizing sustainable chemical materials that foster non-toxic interactions and incorporate recyclable, eco-friendly processes to reduce harmful byproducts. Techniques such as coating and functionalization can enhance the stability and biocompatibility of MNPs, thereby lowering their toxicity. The integration of biocompatible polymers and organic ligands can improve the distribution of MNPs in environmental media and prevent aggregation, which in turn minimizes their chemical reactivity and potential negative effects.

The effective and safe integration of MNPs necessitates collaboration among researchers, industry stakeholders, and regulatory bodies to ensure sustainable practices. It is crucial to disseminate research findings on the environmental impacts of MNPs through scientific journals, conferences, and user-friendly online platforms. Industry stakeholders are

encouraged to adopt best practices for the synthesis, management, and disposal of MNPs, adhering to established standards and guidelines. Regulatory agencies must create clear regulations that outline the proper control of MNPs during environmental impact assessments at the company level. Additionally, public education and community engagement are vital for raising awareness about both the beneficial and harmful aspects of MNPs, empowering individuals to make informed decisions.

Environmental assessments of magnetic nanohybrid materials are essential for their sustainable application in environmental remediation. By carefully evaluating the potential release of toxic substances and their ecological effects, we can reduce the environmental risks linked to magnetic nanoparticles through the implementation of appropriate safety measures. This thorough evaluation framework not only bolsters the safety and security of these materials but also establishes a scientific foundation for their responsible and sustainable utilization. Continued collaboration in research and regulatory initiatives is vital to enhance our understanding of the environmental consequences of magnetic nanoparticles, facilitating the creation of safe and sustainable practices for their ecological use. Fig. 12 illustrates the environmental impact and future challenges associated with these materials.

4. Conclusions and future directions

Magnetic nanohybrid materials (MNMs) have gained recognition as effective agents for environmental remediation due to their distinctive properties, including high surface reactivity, magnetic separability, and customizable functionalities for pollutant extraction. Their versatility allows for applications in water



purification, soil decontamination, and air filtration, as they can efficiently adsorb, catalyze, or degrade various contaminants. Nevertheless, challenges such as agglomeration, long-term stability in environmental settings, scalability of production, and cost-effectiveness pose significant barriers to their broader implementation. To overcome these issues, future research should focus on innovative material designs and the integration of advanced technologies. A particularly promising avenue is the creation of stimuli-responsive MNMs, which can adjust to changes in environmental conditions, thereby enhancing pollutant capture and regeneration. Furthermore, the application of artificial intelligence and machine learning holds the potential to transform the optimization of nanomaterials by accurately predicting synthesis parameters, surface modifications, and interactions with pollutants, ultimately facilitating closed-loop systems for real-time performance improvements.

A vital area of focus is the enhancement of MNMs' durability and reusability through advanced stabilization methods, such as atomic-layer deposition and bio-inspired coatings like poly-dopamine and cellulose derivatives, which help reduce aging and aggregation. Additionally, employing core-shell structures with protective layers can further bolster their chemical and magnetic stability in challenging environments. For broader application, it is essential to investigate green synthesis techniques that utilize plant extracts or microbial processes, thereby minimizing the use of hazardous chemicals and decreasing production costs. Moreover, the integration of MNMs into hybrid photocatalytic systems presents an opportunity for solar-driven degradation of persistent pollutants, facilitating easy magnetic recovery and effectively linking photocatalysis with practical remediation efforts.

In conclusion, while MNMs hold immense promise for addressing environmental pollution, their transition from laboratory innovations to practical applications, a multidisciplinary approach is essential. Emphasizing smart material design, AI-driven optimization, sustainable manufacturing, and hybrid catalytic systems will enable researchers to address current challenges and fully harness the capabilities of these nanomaterials. Collaborative efforts among chemists, engineers, and data scientists are crucial for refining synthesis methods, enhancing stability, and ensuring economic feasibility. With focused advancements in these domains, MNMs have the potential to significantly contribute to sustainable environmental remediation, fostering cleaner ecosystems and supporting a circular economy.

Data availability

Data are available within the manuscript in the form of figures and tables. No primary research results, software or code have been included and no new data were generated or analyzed as part of this review.

Conflicts of interest

The authors declare no conflict of interest regarding the publication of this manuscript.

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References

- 1 D. Lagakos, A. M. Mobarak and M. E. Waugh, The welfare effects of encouraging rural-urban migration, *Econometrica*, 2023, **91**(3), 803–837.
- 2 M. H. Brenner, Industrialization and economic growth: Estimates of their effects on the health of populations, in *Assessing Contributions/h*, Routledge, 2019, pp. 65–115.
- 3 P. Chowdhary, R. N. Bharagava, S. Mishra and N. Khan, Role of industries in water scarcity and its adverse effects on environment and human health, *Environmental Concerns and Sustainable Development: Volume 1: Air, Water and Energy Resources*, 2020, pp. 235–256.
- 4 S. Garg, Z. Z. Chowdhury, A. N. M. Faisal, N. P. Rumjit and P. Thomas, Impact of industrial wastewater on environment and human health, *Advanced Industrial Wastewater Treatment and Reclamation of Water: Comparative Study of Water Pollution Index during Pre-industrial, Industrial Period and Prospect of Wastewater Treatment for Water Resource Conservation* 2022, pp. 197–209.
- 5 X. Zhang, M. Yan, P. Chen, J. Li, Y. Li, H. Li, X. Liu, Z. Chen, H. Yang, S. Wang, J. Wang, Z. Tang, Q. Huang, J. Lei, T. Hayat, Z. Liu, L. Mao, T. Duan and X. Wang, Emerging MOFs, COFs, and their derivatives for energy and environmental applications, *Innovation*, 2025, **6**(2), 100778–100798.
- 6 Y. Wang, Q. Geng, J. Yang, Y. Liu and C. Liu, Hybrid system of flocculation-photocatalysis for the decolorization of crystal violet, reactive red X-3B, and acid orange II dye, *ACS Omega*, 2020, **5**(48), 31137–31145.
- 7 R. Sharma, A. Rana and D. Panthari, Wastewater pollution induced detrimental impacts on aquatic biodiversity: A review, *Advances in Environmental Pollution Management: Wastewater Impacts and Treatment Technologies*, Agro Environ Media, Haridwar, Uttarakhand, India, 2020, pp. 113–127.
- 8 S. Dutta, S. Adhikary, S. Bhattacharya, D. Roy, S. Chatterjee, A. Chakraborty, D. Banerjee, A. Ganguly, S. Nanda and P. Rajak, Contamination of textile dyes in aquatic environment: Adverse impacts on aquatic ecosystem and human health, and its management using bioremediation, *J. Environ. Manage.*, 2024, **353**, 120103.
- 9 V. M. Pathak, V. K. Verma, B. S. Rawat, B. Kaur, N. Babu, A. Sharma, S. Dewali, M. Yadav, R. Kumari and S. Singh, Current status of pesticide effects on environment, human health and its eco-friendly management as bioremediation: A comprehensive review, *Front. Microbiol.*, 2022, **13**, 962619.
- 10 J. A. Silva, Wastewater treatment and reuse for sustainable water resources management: a systematic literature review, *Sustainability*, 2023, **15**(14), 10940.



- 11 A. Singh, A. Sharma, R. K. Verma, R. L. Chopade, P. P. Pandit, V. Nagar, V. Aseri, S. K. Choudhary, G. Awasthi and K. K. Awasthi, Heavy metal contamination of water and their toxic effect on living organisms, in *The Toxicity of Environmental Pollutants*, IntechOpen, 2022.
- 12 W. U. Rehman, S. Y. Barykin, A. V. Sharkova, K. M. J. Iqbal, M. I. Khan, K. S. Alpysbayev and L. N. Danilova, A Novel Approach to Waste Load Management and Environmental Compliance for Marine Pollution Based on Econometric Modeling, *J. Leg. Ethical Regul. Issues*, 2021, **24**(Pt. 2), 1.
- 13 E. Lusweti, E. K. Kanda, J. Obando and M. Makokha, Effects of oil exploration on surface water quality—a review, *Water Pract. Technol.*, 2022, **17**(10), 2171–2185.
- 14 A. K. Sharma, M. Gupta, S. A. Khan, A. Neem, H. Kumar, A. Bhati and U. Patel, in *Effects of Air Pollution on Surface Water Contamination*, IOP Conference Series: Earth and Environmental Science, IOP Publishing, 2024, p. 012036.
- 15 X. Pan, J. Ji, N. Zhang and M. Xing, Research progress of graphene-based nanomaterials for the environmental remediation, *Chin. Chem. Lett.*, 2020, **31**(6), 1462–1473.
- 16 A. Ali, T. Shah, R. Ullah, P. Zhou, M. Guo, M. Ovais, Z. Tan and Y. Rui, Review on recent progress in magnetic nanoparticles: Synthesis, characterization, and diverse applications, *Front. Chem.*, 2021, **9**, 629054.
- 17 C. Verma, D. K. Verma, E. Berdimurodov, I. Barsoum, A. Alfantazi and C. M. Hussain, Green magnetic nanoparticles: a comprehensive review of recent progress in biomedical and environmental applications, *J. Mater. Sci.*, 2024, **59**(2), 325–358.
- 18 S. Shukla, R. Khan and A. Daverey, Synthesis and characterization of magnetic nanoparticles, and their applications in wastewater treatment: A review, *Environ. Technol. Innovation*, 2021, **24**, 101924.
- 19 E. Mutegea, Efficient techniques and practices for wastewater treatment: an update, *Discover Water*, 2024, **4**(1), 69.
- 20 P. Kumari and A. Kumar, Advanced oxidation process: a remediation technique for organic and non-biodegradable pollutant, *Results Surf. Interfaces*, 2023, **11**, 100122.
- 21 S. Kato and Y. Kansha, Comprehensive review of industrial wastewater treatment techniques, *Environ. Sci. Pollut. Res.*, 2024, **31**(39), 51064–51097.
- 22 A. H. Birniwa, S. Habibu, S. S. a. Abdullahi, R. E. A. Mohammad, A. Hussaini, H. Magaji, B. N. S. Al-dhawi, A. Noor and A. H. Jagaba, Membrane technologies for heavy metals removal from water and wastewater: A mini review, *Case Stud. Chem. Environ. Eng.*, 2024, **9**, 100538.
- 23 F. Saffarimiandoab, B. Y. Gul, R. S. Tasdemir, S. E. Ilter, S. Unal, B. Tunaboyle, Y. Z. Menceloglu and I. Koyuncu, A review on membrane fouling: Membrane modification, *Desalin. Water Treat.*, 2021, **216**, 47–70.
- 24 E. M. Cuerda-Correa, M. F. Alexandre-Franco and C. Fernández-González, Advanced oxidation processes for the removal of antibiotics from water. An overview, *Water*, 2019, **12**(1), 102.
- 25 X. Yang and D. Wang, Photocatalysis: from fundamental principles to materials and applications, *ACS Appl. Energy Mater.*, 2018, **1**(12), 6657–6693.
- 26 S. K. Lakhera, K. P. Kangeyan, C. Yazhini S, S. Golda A and N. Bernaudshaw, Advances in hybrid strategies for enhanced photocatalytic water splitting: Bridging conventional and emerging methods, *Appl. Phys. Rev.*, 2024, **11**(4), 041305.
- 27 C. Zuo, X. Tai, Z. Jiang, M. Liu, J. Jiang, Q. Su and X. Yan, S-Scheme 2D/2D Heterojunction of ZnTiO₃ Nanosheets/Bi₂WO₆ Nanosheets with Enhanced Photoelectrocatalytic Activity for Phenol Wastewater under Visible Light, *Molecules*, 2023, **28**(8), 3495.
- 28 F. Tanos, A. Razzouk, G. Lesage, M. Cretin and M. Bechelany, A Comprehensive Review on Modification of Titanium Dioxide-Based Catalysts in Advanced Oxidation Processes for Water Treatment, *ChemSusChem*, 2024, **17**(6), e202301139.
- 29 P. Tiwari, M. Verma, Ambika, H. Chutani, P. P. Singh, S. Kanodia and T. Verma, Titanium dioxide-based nanoparticles and their applications in water remediation, *J. Environ. Eng. Sci.*, 2023, **19**(1), 46–54.
- 30 J. Radić, G. Žerjav, L. Jurko, P. Bošković, L. Fras Zemljič, A. Vesel, A. Mavrič, M. Gudelj and O. Plohl, First Utilization of Magnetically-Assisted Photocatalytic Iron Oxide-TiO₂ Nanocomposites for the Degradation of the Problematic Antibiotic Ciprofloxacin in an Aqueous Environment, *Magnetochemistry*, 2024, **10**(9), 66.
- 31 L. F. O. Maia, M. S. Santos, T. G. Andrade, R. de Carvalho Hott, M. C. da Silva Faria, L. C. A. Oliveira, M. C. Pereira and J. L. Rodrigues, Removal of mercury (II) from contaminated water by gold-functionalised Fe₃O₄ magnetic nanoparticles, *Environ. Technol.*, 2020, **41**, 959–970.
- 32 S. K. Panigrahy, A. Nandha, M. Chaturvedi and P. K. Mishra, Novel nanocomposites with advanced materials and their role in waste water treatment, *Next Sustainability*, 2024, **4**, 100042.
- 33 J. Kudr, Y. Haddad, L. Richtera, Z. Heger, M. Cernak, V. Adam and O. Zitka, Magnetic nanoparticles: From design and synthesis to real world applications, *Nanomaterials*, 2017, **7**(9), 243.
- 34 N. Baig, I. Kammakakam and W. Falath, Nanomaterials: A review of synthesis methods, properties, recent progress, and challenges, *Mater. Adv.*, 2021, **2**(6), 1821–1871.
- 35 S. Kumari, S. Raturi, S. Kulshrestha, K. Chauhan, S. Dhingra, K. András, K. Thu, R. Khargotra and T. Singh, A comprehensive review on various techniques used for synthesizing nanoparticles, *J. Mater. Res. Technol.*, 2023, **27**, 1739–1763.
- 36 M. S. El-Eskandarany, A. Al-Hazza, L. A. Al-Hajji, N. Ali, A. A. Al-Duweesh, M. Banyan and F. Al-Ajmi, Mechanical milling: a superior nanotechnological tool for fabrication of nanocrystalline and nanocomposite materials, *Nanomaterials*, 2021, **11**(10), 2484.
- 37 M. Bououdina, T. Alwqyan, L. Khezami, B. Al-Najar, M. Shaikh, R. Gill, A. Modwi, K. K. Taha and O. Lemine,



- Fabrication and characterization of nanostructured MgO-Fe₂O₃ composite by mechanical milling as efficient adsorbent of heavy metals, *J. Alloys Compd.*, 2019, **772**, 1030–1039.
- 38 Y. Li, A. R. Zimmerman, F. He, J. Chen, L. Han, H. Chen, X. Hu and B. Gao, Solvent-free synthesis of magnetic biochar and activated carbon through ball-mill extrusion with Fe₃O₄ nanoparticles for enhancing adsorption of methylene blue, *Sci. Total Environ.*, 2020, **722**, 137972.
 - 39 A. A. Demessie, Y. Park, P. Singh, A. S. Moses, T. Korzun, F. Y. Sabei, H. A. Albarqi, L. Campos, C. R. Wyatt and K. Farsad, An advanced thermal decomposition method to produce magnetic nanoparticles with ultrahigh heating efficiency for systemic magnetic hyperthermia, *Small Methods*, 2022, **6**(12), 2200916.
 - 40 J. Park, K. An, Y. Hwang, J.-G. Park, H.-J. Noh, J.-Y. Kim, J.-H. Park, N.-M. Hwang and T. Hyeon, Ultra-large-scale syntheses of monodisperse nanocrystals, *Nat. Mater.*, 2004, **3**(12), 891–895.
 - 41 F. B. Effenberger, R. A. Couto, P. K. Kiyohara, G. Machado, S. H. Masunaga, R. F. Jardim and L. M. Rossi, Economically attractive route for the preparation of high quality magnetic nanoparticles by the thermal decomposition of iron (III) acetylacetonate, *Nanotechnology*, 2017, **28**(11), 115603.
 - 42 K. Khan, S. Rehman, H. U. Rahman and Q. Khan, Synthesis and application of magnetic nanoparticles, *Nanomagnetism*, J. M. Gonzalez Estevez, One Central Press (OCP), UK, 2014.
 - 43 A. Balachandran, S. P. Sreenilayam, K. Madanan, S. Thomas and D. Brabazon, Nanoparticle production via laser ablation synthesis in solution method and printed electronic application-A brief review, *Results Eng.*, 2022, **16**, 100646.
 - 44 M. Kim, S. Osone, T. Kim, H. Higashi and T. Seto, Synthesis of nanoparticles by laser ablation: A review, *KONA Powder Part. J.*, 2017, **34**, 80–90.
 - 45 K. Sakiyama, K. Koga, T. Seto, M. Hirasawa and T. Orii, Formation of size-selected Ni/NiO core-shell particles by pulsed laser ablation, *J. Phys. Chem. B*, 2004, **108**(2), 523–529.
 - 46 S. Shinde, S. Kulkarni, A. Banpurkar, R. Nawathey-Dixit, S. Date and S. Ogale, Magnetic properties of nanosized powders of magnetic oxides synthesized by pulsed laser ablation, *J. Appl. Phys.*, 2000, **88**(3), 1566–1575.
 - 47 Y. Wang, J. A. Pan, H. Wu and D. V. Talapin, Direct Wavelength-Selective Optical and Electron-Beam Lithography of Functional Inorganic Nanomaterials, *ACS Nano*, 2019, **13**(12), 13917–13931.
 - 48 T. A. Shifa, F. Wang, Z. Cheng, P. He, Y. Liu, C. Jiang, Z. Wang and J. He, High Crystal Quality 2D Manganese Phosphorus Trichalcogenide Nanosheets and their Photocatalytic Activity, *Adv. Funct. Mater.*, 2018, **28**, 1800548.
 - 49 S. Kumar, P. Bhushan and S. Bhattacharya, Fabrication of nanostructures with bottom-up approach and their utility in diagnostics, therapeutics, and others, *Environmental, Chemical and Medical Sensors*, 2018, pp. 167–198.
 - 50 A. A. Dinata, A. M. Rosyadi, S. Hamid and R. Zainul, *A Review Chemical Vapor Deposition: Process And Application*, 2018.
 - 51 H. Niu, Y. Yang, W. Zhao, H. Lv, H. Zhang and Y. Cai, Single-crystalline Fe₇S₈/Fe₃O₄ coated zero-valent iron synthesized with vacuum chemical vapor deposition technique: Enhanced reductive, oxidative and photocatalytic activity for water purification, *J. Hazard. Mater.*, 2021, **401**, 123442.
 - 52 H. Tavakol, S. Salimpour and W. Salvenmoser, The synthesis of sulfur-doped carbon nanofibers using chemical vapor deposition on the nickel-ferrite catalyst and the gold decoration of the product for morphine sensing, *SN Appl. Sci.*, 2020, **2**(12), 2018.
 - 53 D. Bokov, A. Turki Jalil, S. Chupradit, W. Suksatan, M. Javed Ansari, I. H. Shewael, G. H. Valiev and E. Kianfar, Nanomaterial by sol-gel method: synthesis and application, *Adv. Mater. Sci. Eng.*, 2021, **2021**(1), 5102014.
 - 54 G. F. Stiufiuc and R. I. Stiufiuc, Magnetic nanoparticles: synthesis, characterization, and their use in biomedical field, *Appl. Sci.*, 2024, **14**(4), 1623.
 - 55 M. K. Mohammed, Sol-gel synthesis of Au-doped TiO₂ supported SWCNT nanohybrid with visible-light-driven photocatalytic for high degradation performance toward methylene blue dye, *Optik*, 2020, **223**, 165607.
 - 56 J. Zia and U. Riaz, Studies on the spectral, morphological and magnetic properties of PCz-PPy copolymer encapsulated BaFe₂O₄ nanohybrids, *J. Mater. Sci.: Mater. Electron.*, 2020, **31**(24), 22856–22865.
 - 57 E. M. Materón, C. M. Miyazaki, O. Carr, N. Joshi, P. H. Picciani, C. J. Dalmaschio, F. Davis and F. M. Shimizu, Magnetic nanoparticles in biomedical applications: A review, *Appl. Surf. Sci. Adv.*, 2021, **6**, 100163.
 - 58 J. Li, Q. Wu and J. Wu, Synthesis of Nanoparticles via Solvothermal and Hydrothermal Methods, in *Handbook of Nanoparticles*, ed. M. Aliofkhaezrai, Springer, Cham, 2015, DOI: [10.1007/978-3-319-13188-7_17-1](https://doi.org/10.1007/978-3-319-13188-7_17-1).
 - 59 M. U. Zahid, E. Pervaiz, A. Hussain, M. I. Shahzad and M. B. K. Niazi, Synthesis of carbon nanomaterials from different pyrolysis techniques: a review, *Mater. Res. Express*, 2018, **5**(5), 052002.
 - 60 D. Nugroho, R. Benchawattananon, J. Janshongsawang, N. Pimsin, P. Porrawatkul, R. Pimsen, P. Nuengmatcha, P. Nueangmatcha and S. Chanthai, Ultra-trace analysis of chromium ions (Cr³⁺/Cr⁶⁺) in water sample using selective fluorescence turn-off sensor with natural carbon dots mixed graphene quantum dots nanohybrid composite synthesis by pyrolysis, *Arabian J. Chem.*, 2024, **17**(1), 105443.
 - 61 M. Tsompanoglou, Fe₃C/Fe Magnetic nanohybrids: Structural, magnetic features & biomedical applicability, Doctoral dissertation, Aristotle University Of Thessaloniki, 2023.
 - 62 M. A. Albalah, Y. A. Alsabah and D. E. Mustafa, Characteristics of co-precipitation synthesized cobalt nanoferrites and their potential in industrial wastewater treatment, *SN Appl. Sci.*, 2020, **2**(5), 804.



- 63 R. Verma, S. Pathak, A. K. Srivastava, S. Prawer and S. Tomljenovic-Hanic, ZnO nanomaterials: Green synthesis, toxicity evaluation and new insights in biomedical applications, *J. Alloys Compd.*, 2021, **876**, 160175.
- 64 A. Komeili, Molecular mechanisms of compartmentalization and biomineralization in magnetotactic bacteria, *FEMS Microbiol. Rev.*, 2012, **36**(1), 232–255.
- 65 S. Gul, S. B. Khan, I. U. Rehman, M. A. Khan and M. Khan, A comprehensive review of magnetic nanomaterials modern day theranostics, *Frontiers in Materials*, 2019, **6**, 179.
- 66 J. J. Lenders, C. L. Altan, P. H. Bomans, A. Arakaki, S. Bucak, G. de With and N. A. Sommerdijk, A bioinspired coprecipitation method for the controlled synthesis of magnetite nanoparticles, *Cryst. Growth Des.*, 2014, **14**(11), 5561–5568.
- 67 M. Duan, J. G. Shapter, W. Qi, S. Yang and G. Gao, Recent progress in magnetic nanoparticles: synthesis, properties, and applications, *Nanotechnology*, 2018, **29**(45), 452001.
- 68 U. Kumar, J. Z. Hassan, R. A. Bhatti, A. Raza, G. Nazir, W. Nabgan and M. Ikram, Photocatalysis vs adsorption by metal oxide nanoparticles, *J. Mater. Sci. Technol.*, 2022, **131**, 122–166.
- 69 N. S. Omar, L. I. Abd Ali, A. F. Qader and R. A. Omer, Use of magnetic nanoparticles for the removal of organic and inorganic pollution in wastewater treatment-a review, *Rev. Inorg. Chem.*, 2025, DOI: [10.1515/revic-2024-0117](https://doi.org/10.1515/revic-2024-0117).
- 70 N. Ajinkya, X. Yu, P. Kaithal, H. Luo, P. Somani and S. Ramakrishna, Magnetic iron oxide nanoparticle (IONP) synthesis to applications: present and future, *Materials*, 2020, **13**(20), 4644.
- 71 K. Hachem, M. J. Ansari, R. O. Saleh, H. H. Kzar, M. E. Al-Gazally, U. S. Altamari, S. A. Hussein, H. T. Mohammed, A. T. Hammid and E. Kianfar, Methods of chemical synthesis in the synthesis of nanomaterial and nanoparticles by the chemical deposition method: a review, *BioNanoScience*, 2022, **12**(3), 1032–1057.
- 72 Q. Xu, J. Hou, J. Rao, G.-H. Li, Y.-L. Liu and J. Zhou, PEG modification enhances the in vivo stability of bioactive proteins immobilized on magnetic nanoparticles, *Biotechnol. Lett.*, 2020, **42**(8), 1407–1418.
- 73 V. F. Cardoso, A. Francesko, C. Ribeiro, M. Bañobre-López, P. Martins and S. Lanceros-Mendez, Advances in magnetic nanoparticles for biomedical applications, *Adv. Healthcare Mater.*, 2018, **7**(5), 1700845.
- 74 M. I. Anik, M. K. Hossain, I. Hossain, A. Mahfuz, M. T. Rahman and I. Ahmed, Recent progress of magnetic nanoparticles in biomedical applications: A review, *Nano Sel.*, 2021, **2**(6), 1146–1186.
- 75 J. Guo, H. Jiang, Y. Teng, Y. Xiong, Z. Chen, L. You and D. Xiao, Recent advances in magnetic carbon nanotubes: synthesis, challenges and highlighted applications, *J. Mater. Chem. B*, 2021, **9**(44), 9076–9099.
- 76 M. Selvaraj, A. Hai, F. Banat and M. A. Haija, Application and prospects of carbon nanostructured materials in water treatment: A review, *J. Water Proc. Eng.*, 2020, **33**, 100996.
- 77 M. Lu, X. J. Wu, C. X. Wan, Q. P. Gong, J. X. Li, S. S. Liao, Y. A. Wang and S. H. Yuan, Evaluation of Fe₃O₄-MnO₂@RGO magnetic nanocomposite as an effective persulfate activator and metal adsorbent in aqueous solution, *Environ. Sci. Pollut. Res.*, 2023, **30**(17), 51125–51142.
- 78 A. Karnwal and T. Malik, Nano-revolution in heavy metal removal: engineered nanomaterials for cleaner water, *Front. Environ. Sci.*, 2024, **12**, 1393694.
- 79 C. Donga, S. B. Mishra, A. S. Abd-El-Aziz and A. K. Mishra, Advances in graphene-based magnetic and graphene-based/TiO₂ nanoparticles in the removal of heavy metals and organic pollutants from industrial wastewater, *J. Inorg. Organomet. Polym. Mater.*, 2021, **31**(2), 463–480.
- 80 S. Dhadda, N. Jangir, P. Sihag, S. K. Bagaria and D. K. Jangid, The designing of magnetic nanomaterials synthesis and applications in dye degradation: a review, *Environ. Sci. Pollut. Res.*, 2024, 1–23.
- 81 S. Yueyu, The synergistic degradation of pollutants in water by photocatalysis and PMS activation, *Water Environ. Res.*, 2023, **95**(10), e10927.
- 82 G. M. Nair, T. Sajini and B. Mathew, Advanced green approaches for metal and metal oxide nanoparticles synthesis and their environmental applications, *Talanta Open*, 2022, **5**, 100080.
- 83 E. M. Hotze, T. Phenrat and G. V. Lowry, Nanoparticle aggregation: challenges to understanding transport and reactivity in the environment, *J. Environ. Qual.*, 2010, **39**(6), 1909–1924.
- 84 A. Heuer-Jungemann, N. Feliu, I. Bakaimi, M. Hamaly, A. Alkilany, I. Chakraborty, A. Masood, M. F. Casula, A. Kostopoulou and E. Oh, The role of ligands in the chemical synthesis and applications of inorganic nanoparticles, *Chem. Rev.*, 2019, **119**(8), 4819–4880.
- 85 J. Hang, L. Shi, X. Feng and L. Xiao, Electrostatic and electrosteric stabilization of aqueous suspensions of barite nanoparticles, *Powder Technol.*, 2009, **192**(2), 166–170.
- 86 T. L. Doane, C.-H. Chuang, R. J. Hill and C. Burda, Nanoparticle ζ -potentials, *Acc. Chem. Res.*, 2012, **45**(3), 317–326.
- 87 P. May, U. Khan, J. M. Hughes and J. N. Coleman, Role of solubility parameters in understanding the steric stabilization of exfoliated two-dimensional nanosheets by adsorbed polymers, *J. Phys. Chem. C*, 2012, **116**(20), 11393–11400.
- 88 C. Vasilescu, M. Latikka, K. D. Knudsen, V. M. Garamus, V. Socoliuc, R. Turcu, E. Tombácz, D. Susan-Resiga, R. Ras and L. Vékás, High concentration aqueous magnetic fluids: structure, colloidal stability, magnetic and flow properties, *Soft Matter*, 2018, **14**(32), 6648–6666.
- 89 J. Hierrezuelo, A. Sadeghpour, I. Szilagyi, A. Vaccaro and M. Borkovec, Electrostatic stabilization of charged colloidal particles with adsorbed polyelectrolytes of opposite charge, *Langmuir*, 2010, **26**(19), 15109–15111.



- 90 J. Alonso, J. M. Barandiarán, L. Fernández Barquín and A. García-Arribas, Chapter 1 - Magnetic Nanoparticles, Synthesis, Properties, and Applications, in *Magnetic Nanostructured Materials*, ed. A. A. El-Gendy, J. M. Barandiarán and R. L. Hadimani, Elsevier, 2018, pp. 1–40.
- 91 S. Wang, D. Chen, Q. Hong, Y. Gui, Y. Cao, G. Ren and Z. Liang, Surface functionalization of metal and metal oxide nanoparticles for dispersion and tribological applications – A review, *J. Mol. Liq.*, 2023, **389**, 122821.
- 92 C. Pechyen, B. Tangnorawich, S. Toommee, R. Marks and Y. Parcharoen, Green synthesis of metal nanoparticles, characterization, and biosensing applications, *Sens. Int.*, 2024, **5**, 100287.
- 93 J. I. Orege, J. Wei, Q. Ge and J. Sun, Spinel-structured nanocatalysts: New opportunities for CO₂ hydrogenation to value-added chemicals, *Nano Today*, 2023, **51**, 101914.
- 94 N. Dudchenko, S. Pawar, I. Perelshtein and D. Fixler, Magnetite nanoparticles: Synthesis and applications in optics and nanophotonics, *Materials*, 2022, **15**(7), 2601.
- 95 P. Kumar, S. Pathak, K. Jain, A. Singh, G. A. Basheed and R. P. Pant, Low-temperature large-scale hydrothermal synthesis of optically active PEG-200 capped single domain MnFe₂O₄ nanoparticles, *J. Alloys Compd.*, 2022, **904**, 163992.
- 96 M. A. Dahamni, M. Ghamnia, S. E. Naceri, C. Fauquet, D. Tonneau, J.-J. Pireaux and A. Bouadi, Spray Pyrolysis Synthesis of Pure and Mg-Doped Manganese Oxide Thin Films, *Coatings*, 2021, **11**(5), 598.
- 97 M. Marć, W. Wolak, A. Drzewiński, S. Mudry, I. Shtablavyi and M. R. Dudek, The Concept of Using 2D Self-Assembly of Magnetic Nanoparticles for Bioassays, *Appl. Sci.*, 2024, **14**(5), 1906.
- 98 H. Al-Madhagi, V. Yazbik, W. Abdelwahed and L. Alchab, Magnetite nanoparticle Co-precipitation synthesis, characterization, and applications: mini review, *Bionanoscience*, 2023, **13**(2), 853–859.
- 99 N. Dudchenko, S. Pawar, I. Perelshtein and D. Fixler, Magnetite Nanoparticles: Synthesis and Applications in Optics and Nanophotonics, *Materials*, 2022, **15**, 2601.
- 100 R. R. Karri, J. Sahu and B. Meikap, Improving efficacy of Cr (VI) adsorption process on sustainable adsorbent derived from waste biomass (sugarcane bagasse) with help of ant colony optimization, *Ind. Crops Prod.*, 2020, **143**, 111927.
- 101 H. Rasoulzadeh, M. H. Dehghani, A. S. Mohammadi, R. R. Karri, R. Nabizadeh, S. Nazmara, K.-H. Kim and J. Sahu, Parametric modelling of Pb (II) adsorption onto chitosan-coated Fe₃O₄ particles through RSM and DE hybrid evolutionary optimization framework, *J. Mol. Liq.*, 2020, **297**, 111893.
- 102 M. Ruthiraan, N. Mubarak, E. Abdullah, M. Khalid, S. Nizamuddin, R. Walvekar and R. R. Karri, An overview of magnetic material: preparation and adsorption removal of heavy metals from wastewater, *Magnetic Nanostructures: Environmental and Agricultural Applications*, 2019, pp. 131–159.
- 103 M. Bora and D. C. Goswami, Water quality assessment in terms of water quality index (WQI): case study of the Kolong River, Assam, India, *Appl. Water Sci.*, 2017, **7**, 3125–3135.
- 104 R. L. Droste and R. L. Gehr, *Theory and Practice of Water and Wastewater Treatment*, John Wiley & Sons, 2018.
- 105 J. Machell, K. Prior, R. Allan and J. M. Andresen, The water energy food nexus—challenges and emerging solutions, *Environ. Sci.:Water Res. Technol.*, 2015, **1**(1), 15–16.
- 106 A. Nasiri, M. R. Heidari, N. Javid and G. Yazdanpanah, New efficient and recyclable magnetic nanohybrid adsorbent for the metronidazole removal from simulated wastewater, *J. Mater. Sci.: Mater. Electron.*, 2022, **33**(33), 25103–25126.
- 107 A. Choudhry, A. Sharma, T. A. Khan and S. A. Chaudhry, Flax seeds based magnetic hybrid nanocomposite: An advance and sustainable material for water cleansing, *J. Water Proc. Eng.*, 2021, **42**, 102150.
- 108 O. Saber, M. Osama, N. M. Shaalan, A. Osama, A. Alshoaibi and D. Osama, Designing novel strategy to produce active nanohybrids in sunlight for purification of water based on inorganic nanolayers, magnetic nanocomposites and organic species, *Molecules*, 2022, **27**(12), 3673.
- 109 A. Khan, N. Ali, S. Malik, M. Bilal, H. Munir, L. F. R. Ferreira and H. M. Iqbal, Chitosan-based green sorbents for toxic cations removal, in *Sorbents Materials for Controlling Environmental Pollution*, Elsevier, 2021, pp. 323–352.
- 110 A. Aziz, N. Ali, A. Khan, M. Bilal, S. Malik, N. Ali and H. Khan, Chitosan-zinc sulfide nanoparticles, characterization and their photocatalytic degradation efficiency for azo dyes, *Int. J. Biol. Macromol.*, 2020, **153**, 502–512.
- 111 N. Ali, M. Bilal, A. Khan, F. Ali, H. Khan, H. A. Khan, K. Rasool and H. M. Iqbal, Understanding the hierarchical assemblies and oil/water separation applications of metal-organic frameworks, *J. Mol. Liq.*, 2020, **318**, 114273.
- 112 A. Nawaz, A. Khan, N. Ali, N. Ali and M. Bilal, Fabrication and characterization of new ternary ferrites-chitosan nanocomposite for solar-light driven photocatalytic degradation of a model textile dye, *Environ. Technol. Innovation*, 2020, **20**, 101079.
- 113 Y. Zhou, Z. He, Y. Tao, Y. Xiao, T. Zhou, T. Jing, Y. Zhou and S. Mei, Preparation of a functional silica membrane coated on Fe₃O₄ nanoparticle for rapid and selective removal of perfluorinated compounds from surface water sample, *Chem. Eng. J.*, 2016, **303**, 156–166.
- 114 S. B. Qin, Y. H. Fan, X. X. Mou, X. S. Li and S. H. Qi, Preparation of phenyl-modified magnetic silica as a selective magnetic solid-phase extraction adsorbent for polycyclic aromatic hydrocarbons in soils, *J. Chromatogr. A*, 2018, **1568**, 29–37.
- 115 P. Wu, S. Yu, M. Feng, H. Liu, S. Liu and J. Fu, Controllable synthesis of the polymorphic porous carbon with N-doping/ Ni magnetic nanohybrids for high performance supercapacitor and environment applications, *Appl. Surf. Sci.*, 2021, **567**, 150875.
- 116 S. S. Elanchezhian, S. M. Prabhu, J. Han, Y. M. Kim, Y. Yoon and C. M. Park, Synthesis and characterization of novel magnetic Zr-MnFe₂O₄@ rGO nanohybrid for



- efficient removal of PFOA and PFOS from aqueous solutions, *Appl. Surf. Sci.*, 2020, **528**, 146579.
- 117 I. Ali, The quest for active carbon adsorbent substitutes: inexpensive adsorbents for toxic metal ions removal from wastewater, *Sep. Purif. Rev.*, 2010, **39**(3–4), 95–171.
 - 118 R. R. Karri and J. Sahu, Process optimization and adsorption modeling using activated carbon derived from palm oil kernel shell for Zn (II) disposal from the aqueous environment using differential evolution embedded neural network, *J. Mol. Liq.*, 2018, **265**, 592–602.
 - 119 B. Samiey, C. H. Cheng and J. Wu, Organic-inorganic hybrid polymers as adsorbents for removal of heavy metal ions from solutions: a review, *Materials*, 2014, **7**(2), 673–726.
 - 120 L. Järup, Hazards of heavy metal contamination, *Br. Med. Bull.*, 2003, **68**(1), 167–182.
 - 121 L. P. Lingamdinne, J. R. Koduru, Y.-Y. Chang and R. R. Karri, Process optimization and adsorption modeling of Pb (II) on nickel ferrite-reduced graphene oxide nano-composite, *J. Mol. Liq.*, 2018, **250**, 202–211.
 - 122 K. F. Amin, F. Gulshan, F. Asrafuzzaman, H. Das, R. Rashid and S. M. Hoque, Synthesis of mesoporous silica and chitosan-coated magnetite nanoparticles for heavy metal adsorption from wastewater, *Environ. Nanotechnol., Monit. Manage.*, 2023, **20**, 100801.
 - 123 S. Bosu, N. Rajamohan and M. Rajasimman, Enhanced remediation of lead (II) and cadmium (II) ions from aqueous media using porous magnetic nanocomposites-A comprehensive review on applications and mechanism, *Environ. Res.*, 2022, **213**, 113720.
 - 124 S. G. Muntean, L. Halip, M. A. Nistor and C. Păcurariu, Removal of metal ions via adsorption using carbon magnetic nanocomposites: optimization through response surface methodology, kinetic and thermodynamic studies, *Magnetochemistry*, 2023, **9**(7), 163.
 - 125 T. Tolessa, X.-X. Zhou, M. Amde and J.-F. Liu, Development of reusable magnetic chitosan microspheres adsorbent for selective extraction of trace level silver nanoparticles in environmental waters prior to ICP-MS analysis, *Talanta*, 2017, **169**, 91–97.
 - 126 Y. Zhao, C. Wang, S. Wang, C. Wang, Y. Liu, A. A. Al-Khalaf, W. N. Hozzein, L. Duan, W. Li and D. Zhao, Magnetic mesoporous TiO₂ microspheres for sustainable arsenate removal from acidic environments, *Inorg. Chem. Front.*, 2018, **5**(9), 2132–2139.
 - 127 Y. Pang, J. Yu, L. Tang, G. Zeng, C. Zhu and X. Wei, Magnetic nanohybrid materials for water-pollutant removal, in *Nanohybrid and Nanoporous Materials for Aquatic Pollution Control*, Elsevier, 2019, pp. 1–30.
 - 128 Z. Wang, J. Zhang, T. Wen, X. Liu, Y. Wang, H. Yang, J. Sun, J. Feng, S. Dong and J. Sun, Highly effective remediation of Pb (II) and Hg (II) contaminated wastewater and soil by flower-like magnetic MoS₂ nanohybrid, *Sci. Total Environ.*, 2020, **699**, 134341.
 - 129 K. Dai, G. Liu, W. Xu, Z. Deng, Y. Wu, C. Zhao and Z. Zhang, Judicious fabrication of bifunctionalized graphene oxide/MnFe₂O₄ magnetic nanohybrids for enhanced removal of Pb (II) from water, *J. Colloid Interface Sci.*, 2020, **579**, 815–822.
 - 130 J. Z. Pirhaji, F. Moeinpour, A. M. Dehabadi and S. A. Y. Ardakani, Synthesis and characterization of halloysite/graphene quantum dots magnetic nanocomposite as a new adsorbent for Pb (II) removal from water, *J. Mol. Liq.*, 2020, **300**, 112345.
 - 131 H. Abdolmohammad-Zadeh, Z. Ayazi and K. Nezami, Magnetic solid-phase extraction based on Ni–Al layered double hydroxide/magnetite nano-hybrid for speciation of Mn (vii)/Mn (ii) in water samples by FAAS, *Anal. Methods*, 2019, **11**(4), 462–471.
 - 132 K. M. M. Katubi, N. S. Alsaiani, F. M. Alzahrani, S. M. Siddeeq and M. A. Tahooun, Synthesis of manganese ferrite/graphene oxide magnetic nanocomposite for pollutants removal from water, *Processes*, 2021, **9**(4), 589.
 - 133 M. A. Ahmed, M. A. Ahmed and A. A. Mohamed, Facile adsorptive removal of dyes and heavy metals from wastewaters using magnetic nanocomposite of zinc ferrite@ reduced graphene oxide, *Inorg. Chem. Commun.*, 2022, **144**, 109912.
 - 134 M. D. Dung, T. T. V. Nga, N. T. Lan and N. K. Thanh, Adsorption behavior and mechanism of As (V) on magnetic Fe₃O₄–graphene oxide (GO) nanohybrid composite material, *Anal. Sci.*, 2022, **38**(2), 427–436.
 - 135 S. Ko, E. S. Kim, S. Park, H. Daigle, T. E. Milner, C. Huh, M. V. Bennetzen and G. A. Geremia, RETRACTED ARTICLE: Amine functionalized magnetic nanoparticles for removal of oil droplets from produced water and accelerated magnetic separation, *J. Nanopart. Res.*, 2017, **19**, 1–14.
 - 136 O. Saber, N. Mohamed and S. Arafat, Conversion of iron oxide nanosheets to advanced magnetic nanocomposites for oil spill removal, *RSC Adv.*, 2015, **5**(89), 72863–72871.
 - 137 A. M. Atta, H. A. Al-Lohedan and S. A. Al-Hussain, Functionalization of magnetite nanoparticles as oil spill collector, *Int. J. Mol. Sci.*, 2015, **16**(4), 6911–6931.
 - 138 M. Sharma, M. Joshi, S. Nigam, D. K. Avasthi, R. Adelung, S. K. Srivastava and Y. K. Mishra, Efficient oil removal from wastewater based on polymer coated superhydrophobic tetrapodal magnetic nanocomposite adsorbent, *Appl. Mater. Today*, 2019, **17**, 130–141.
 - 139 L. Shao, Z. Yu, X. Li, H. Zeng and Y. Liu, One-step preparation of sepiolite/graphene oxide membrane for multifunctional oil-in-water emulsions separation, *Appl. Clay Sci.*, 2019, **181**, 105208.
 - 140 D. Qian, D. Chen, N. Li, Q. Xu, H. Li, J. He and J. Lu, TiO₂/sulfonated graphene oxide/Ag nanoparticle membrane: In situ separation and photodegradation of oil/water emulsions, *J. Membr. Sci.*, 2018, **554**, 16–25.
 - 141 T. Zhang, C. Xiao, J. Zhao, X. Liu, D. Ji and H. Zhang, One-step facile fabrication of PVDF/graphene composite nanofibrous membrane with enhanced oil affinity for highly efficient gravity-driven emulsified oil/water separation and selective oil absorption, *Sep. Purif. Technol.*, 2021, **254**, 117576.



- 142 Y. Cheng, A. Barras, S. Lu, W. Xu, S. Szunerits and R. Boukherroub, Fabrication of superhydrophobic/superoleophilic functionalized reduced graphene oxide/polydopamine/PFDT membrane for efficient oil/water separation, *Sep. Purif. Technol.*, 2020, **236**, 116240.
- 143 H. Shokry, M. Elkady and E. Salama, Eco-friendly magnetic activated carbon nano-hybrid for facile oil spills separation, *Sci. Rep.*, 2020, **10**(1), 10265.
- 144 Q. Yuan, N. Li, W. Geng, Y. Chi and X. Li, Preparation of magnetically recoverable Fe₃O₄@ SiO₂@ meso-TiO₂ nanocomposites with enhanced photocatalytic ability, *Mater. Res. Bull.*, 2012, **47**(9), 2396–2402.
- 145 H. Aliyan, R. Fazaeli and R. Jalilian, Fe₃O₄@ mesoporous SBA-15: A magnetically recoverable catalyst for photodegradation of malachite green, *Appl. Surf. Sci.*, 2013, **276**, 147–153.
- 146 B. Sahoo, S. K. Sahu, S. Nayak, D. Dhara and P. Pramanik, Fabrication of magnetic mesoporous manganese ferrite nanocomposites as efficient catalyst for degradation of dye pollutants, *Catal. Sci. Technol.*, 2012, **2**(7), 1367–1374.
- 147 Y. Ni, Y. Zhu and X. Ma, A simple solution combustion route for the preparation of metal-doped TiO₂ nanoparticles and their photocatalytic degradation properties, *Dalton Trans.*, 2011, **40**(14), 3689–3694.
- 148 S. B. F. Santos, L. R. Hollanda, Y. Vieira, G. L. Dotto, E. L. Foletto and O. Chivavone-Filho, Enhanced UV-light driven photocatalytic performance of magnetic CoFe₂O₄/TiO₂ nanohybrid for environmental applications, *Environ. Sci. Pollut. Res.*, 2023, **30**(30), 75078–75088.
- 149 G. U. Rehman, A. Ismail, P. Goh, M. R.-D. Arzhandi and N. Ismail, APTES and TEOS modified binary recyclable hybrid Fe₃O₄@ GO nanocomposite for photocatalytic dye removal, *Jurnal Teknologi*, 2018, **80**(4), 157–164.
- 150 S. Vigneshwaran, P. Sirajudheen, C. P. Nabeena, V. P. Sajna and S. Meenakshi, Photocatalytic performance of chitosan tethered magnetic Fe₂O₃-like (3D/2D) hybrid for the dynamic removal of anionic dyes: degradation and mechanistic pathways, *Int. J. Biol. Macromol.*, 2021, **183**, 2088–2099.
- 151 M. Ghiyasiyan-Arani, M. Salavati-Niasari and S. Naseh, Enhanced photodegradation of dye in waste water using iron vanadate nanocomposite; ultrasound-assisted preparation and characterization, *Ultrason. Sonochem.*, 2017, **39**, 494–503.
- 152 A. Sobhani-Nasab, S. Pourmasoud, F. Ahmadi, M. Wysokowski, T. Jesionowski, H. Ehrlich and M. Rahimi-Nasrabadi, Synthesis and characterization of MnWO₄/TiVO₄ ternary nano-hybrids by an ultrasonic method for enhanced photocatalytic activity in the degradation of organic dyes, *Mater. Lett.*, 2019, **238**, 159–162.
- 153 M. Roostaei and I. Sheikhshoaei, Low-temperature synthesis of hetero-structures of magnetically separable iron oxide@ Au-rGO nanocomposite for efficient degradation of organic dye under visible light irradiation, *Environ. Res.*, 2022, **205**, 112510.
- 154 L. Yu, L. Wang, X. Sun and D. Ye, Enhanced photocatalytic activity of rGO/TiO₂ for the decomposition of formaldehyde under visible light irradiation, *J. Environ. Sci.*, 2018, **73**, 138–146.
- 155 J. Wilkinson, P. S. Hooda, J. Barker, S. Barton and J. Swinden, Occurrence, fate and transformation of emerging contaminants in water: An overarching review of the field, *Environ. Pollut.*, 2017, **231**, 954–970.
- 156 I. A. Al-Baldawi, A. A. Mohammed, Z. H. Mutar, S. R. S. Abdullah, S. S. Jasim and A. F. Almansoori, Application of phytotechnology in alleviating pharmaceuticals and personal care products (PPCPs) in wastewater: Source, impacts, treatment, mechanisms, fate, and SWOT analysis, *J. Cleaner Prod.*, 2021, **319**, 128584.
- 157 M. Rostami, R. M. Zamani, K. M. Aghajanzadeh and H. Danafar, Sol-gel synthesis and characterization of zinc ferrite-graphene nano-hybrids for photo-catalytic degradation of the paracetamol, *J. Pharm. Invest.*, 2018, **48**, 657–664.
- 158 Y. Tang, X. Liu, C. Ma, M. Zhou, P. Huo, L. Yu, J. Pan, W. Shi and Y. Yan, Enhanced photocatalytic degradation of tetracycline antibiotics by reduced graphene oxide-CdS/ZnS heterostructure photocatalysts, *New J. Chem.*, 2015, **39**(7), 5150–5160.
- 159 C. Cui, Y. Wang, D. Liang, W. Cui, H. Hu, B. Lu, S. Xu, X. Li, C. Wang and Y. Yang, Photo-assisted synthesis of Ag₃PO₄/reduced graphene oxide/Ag heterostructure photocatalyst with enhanced photocatalytic activity and stability under visible light, *Appl. Catal., B*, 2014, **158**, 150–160.
- 160 C. Yang, X. You, J. Cheng, H. Zheng and Y. Chen, A novel visible-light-driven In-based MOF/graphene oxide composite photocatalyst with enhanced photocatalytic activity toward the degradation of amoxicillin, *Appl. Catal., B*, 2017, **200**, 673–680.
- 161 L. Ye, J. Liu, Z. Jiang, T. Peng and L. Zan, Facets coupling of BiOBr-g-C₃N₄ composite photocatalyst for enhanced visible-light-driven photocatalytic activity, *Appl. Catal., B*, 2013, **142**, 1–7.
- 162 H. Cheng, B. Huang, Y. Dai, X. Qin and X. Zhang, One-step synthesis of the nanostructured AgI/BiOI composites with highly enhanced visible-light photocatalytic performances, *Langmuir*, 2010, **26**(9), 6618–6624.
- 163 T. R. Bastami, A. Ahmadpour and F. A. Hekmatikar, Synthesis of Fe₃O₄/Bi₂WO₆ nanohybrid for the photocatalytic degradation of pharmaceutical ibuprofen under solar light, *J. Ind. Eng. Chem.*, 2017, **51**, 244–254.
- 164 T. R. Bastami and M. H. Entezari, Activated carbon from carrot dross combined with magnetite nanoparticles for the efficient removal of p-nitrophenol from aqueous solution, *Chem. Eng. J.*, 2012, **210**, 510–519.
- 165 A. Dolbecq, P. Mialane, B. Keita and L. Nadjo, Polyoxometalate-based materials for efficient solar and visible light harvesting: application to the photocatalytic degradation of azo dyes, *J. Mater. Chem.*, 2012, **22**(47), 24509–24521.
- 166 M. H. Sayadi, N. Ahmadpour and S. Homaeigohar, Photocatalytic and antibacterial properties of Ag-



- CuFe₂O₄@ WO₃ magnetic nanocomposite, *Nanomaterials*, 2021, **11**(2), 298.
- 167 X. Xie, Y. Liu, X. Dong, C. Lin, X. Wen and Q. Yan, Synthesis and characterization of Fe₃O₄/BiOI n-p heterojunction magnetic photocatalysts, *Appl. Surf. Sci.*, 2018, **455**, 742–747.
 - 168 A. Kumar, M. Khan, L. Fang and I. M. Lo, Visible-light-driven N-TiO₂@ SiO₂@ Fe₃O₄ magnetic nanophotocatalysts: synthesis, characterization, and photocatalytic degradation of PPCPs, *J. Hazard. Mater.*, 2019, **370**, 108–116.
 - 169 M. Khan, C. S. Fung, A. Kumar and I. M. Lo, Magnetically separable BiOBr/Fe₃O₄@ SiO₂ for visible-light-driven photocatalytic degradation of ibuprofen: mechanistic investigation and prototype development, *J. Hazard. Mater.*, 2019, **365**, 733–743.
 - 170 L. Fernández, M. Gamallo, M. González-Gómez, C. Vázquez-Vázquez, J. Rivas, M. Pintado and M. Moreira, Insight into antibiotics removal: Exploring the photocatalytic performance of a Fe₃O₄/ZnO nanocomposite in a novel magnetic sequential batch reactor, *J. Environ. Manage.*, 2019, **237**, 595–608.
 - 171 N. Nasseh, T. J. Al-Musawi, M. R. Miri, S. Rodriguez-Couto and A. H. Panahi, A comprehensive study on the application of FeNi₃@ SiO₂@ ZnO magnetic nanocomposites as a novel photo-catalyst for degradation of tamoxifen in the presence of simulated sunlight, *Environ. Pollut.*, 2020, **261**, 114127.
 - 172 S. Mirsadeghi, H. Zandavar, M. Yousefi, H. R. Rajabi and S. M. Pourmortazavi, Green-photodegradation of model pharmaceutical contaminations over biogenic Fe₃O₄/Au nanocomposite and antimicrobial activity, *J. Environ. Manage.*, 2020, **270**, 110831.
 - 173 N. Ahmadpour, M. H. Sayadi, S. Sobhani and M. Hajiani, A potential natural solar light active photocatalyst using magnetic ZnFe₂O₄@ TiO₂/Cu nanocomposite as a high performance and recyclable platform for degradation of naproxen from aqueous solution, *J. Cleaner Prod.*, 2020, **268**, 122023.
 - 174 J. Wang, Q. Zhang, F. Deng, X. Luo and D. D. Dionysiou, Rapid toxicity elimination of organic pollutants by the photocatalysis of environment-friendly and magnetically recoverable step-scheme SnFe₂O₄/ZnFe₂O₄ nano-heterojunctions, *Chem. Eng. J.*, 2020, **379**, 122264.
 - 175 G. Xu, M. Du, T. Li, Y. Guan and C. Guo, Facile synthesis of magnetically retrievable Fe₃O₄/BiVO₄/CdS heterojunction composite for enhanced photocatalytic degradation of tetracycline under visible light, *Sep. Purif. Technol.*, 2021, **275**, 119157.
 - 176 I. Gabelica, L. Ćurković, V. Mandić, I. Panžić, D. Ljubas and K. Zadro, Rapid microwave-assisted synthesis of Fe₃O₄/SiO₂/TiO₂ core-2-layer-shell nanocomposite for photocatalytic degradation of ciprofloxacin, *Catalysts*, 2021, **11**(10), 1136.
 - 177 A. Behera, S. Mansingh, K. K. Das and K. Parida, Synergistic ZnFe₂O₄-carbon allotropes nanocomposite photocatalyst for norfloxacin degradation and Cr (VI) reduction, *J. Colloid Interface Sci.*, 2019, **544**, 96–111.
 - 178 T. He, Y. Wu, C. Jiang, Z. Chen, Y. Wang, G. Liu, Z. Xu, G. Ning, X. Chen and Y. Zhao, Novel magnetic Fe₃O₄/g-C₃N₄/MoO₃ nanocomposites with highly enhanced photocatalytic activities: visible-light-driven degradation of tetracycline from aqueous environment, *PLoS One*, 2020, **15**(8), e0237389.
 - 179 S. Raha and M. Ahmaruzzaman, Enhanced performance of a novel superparamagnetic g-C₃N₄/NiO/ZnO/Fe₃O₄ nanohybrid photocatalyst for removal of esomeprazole: Effects of reaction parameters, co-existing substances and water matrices, *Chem. Eng. J.*, 2020, **395**, 124969.
 - 180 M. Pourshaban-Mazandarani and A. Nasiri, Photocatalytic degradation of tetracycline in wastewater with bio-based matrix magnetic heterogeneous nanocatalyst: performance and mechanism study, *J. Polym. Environ.*, 2024, 1–25.
 - 181 R. K. Gautam and I. Tiwari, Humic acid functionalized magnetic nanomaterials for remediation of dye wastewater under ultrasonication: Application in real water samples, recycling and reuse of nanosorbents, *Chemosphere*, 2020, **245**, 125553.
 - 182 I. Fatimah, E. Z. Pratiwi and W. P. Wicaksono, Synthesis of magnetic nanoparticles using *Parkia speciosa* Hassk pod extract and photocatalytic activity for Bromophenol blue degradation, *Egypt. J. Aquat. Res.*, 2020, **46**(1), 35–40.
 - 183 A. Kumar, G. Sharma, M. Naushad and S. Thakur, SPION/ β -cyclodextrin core-shell nanostructures for oil spill remediation and organic pollutant removal from waste water, *Chem. Eng. J.*, 2015, **280**, 175–187.
 - 184 S. Bakhshi Nejad and A. Mohammadi, Epoxy-triazinetriene-functionalized magnetic nanoparticles as an efficient magnetic nanoadsorbent for the removal of malachite green and Pb (II) from aqueous solutions, *J. Chem. Eng. Data*, 2020, **65**(5), 2731–2742.
 - 185 C. Das, S. Sen, T. Singh, T. Ghosh, S. S. Paul, T. W. Kim, S. Jeon, D. K. Maiti, J. Im and G. Biswas, Green synthesis, characterization and application of natural product coated magnetite nanoparticles for wastewater treatment, *Nanomaterials*, 2020, **10**(8), 1615.
 - 186 R. Nodehi, H. Shayesteh and A. Rahbar-Kelishami, Fe₃O₄@NiO core-shell magnetic nanoparticle for highly efficient removal of Alizarin red S anionic dye, *Int. J. Environ. Sci. Technol.*, 2022, **19**(4), 2899–2912.
 - 187 M. N. Begam, K. Muthukumaran and P. Thamarai, Application of *Musa acuminata* mediated iron oxide nanoparticles in the removal of sunset yellow dye, *Desalin. Water Treat.*, 2024, **317**, 100022.
 - 188 R. A. Azeez and F. K. I. Al-Zuhairi, Biosorption of dye by immobilized yeast cells on the surface of magnetic nanoparticles, *Alexandria Eng. J.*, 2022, **61**(7), 5213–5222.
 - 189 C. V. Reddy, I. N. Reddy, K. Ravindranadh, K. R. Reddy, N. P. Shetti, D. Kim, J. Shim and T. M. Aminabhavi, Copper-doped ZrO₂ nanoparticles as high-performance catalysts for efficient removal of toxic organic pollutants



- and stable solar water oxidation, *J. Environ. Manage.*, 2020, **260**, 110088.
- 190 A. H. Nordin, S. Wong, N. Ngadi, M. M. Zainol, N. A. F. Abd Latif and W. Nabgan, Surface functionalization of cellulose with polyethyleneimine and magnetic nanoparticles for efficient removal of anionic dye in wastewater, *J. Environ. Chem. Eng.*, 2021, **9**(1), 104639.
 - 191 A. Peyghami, A. Moharrami, Y. Rashtbari, S. Afshin, M. Vosuoghi and A. Dargahi, Evaluation of the efficiency of magnetized clinoptilolite zeolite with Fe₃O₄ nanoparticles on the removal of basic violet 16 (BV16) dye from aqueous solutions, *J. Dispersion Sci. Technol.*, 2023, **44**(2), 278–287.
 - 192 E. Ghoohestani, F. Samari, A. Homaei and S. Yosuefinejad, A facile strategy for preparation of Fe₃O₄ magnetic nanoparticles using *Cordia myxa* leaf extract and investigating its adsorption activity in dye removal, *Sci. Rep.*, 2024, **14**(1), 84.
 - 193 A. N. Tiwari, K. Tapadia and C. Thakur, A Facile Synthesis of MgO Decorated Spinel Magnetic Nanoparticles Using Terminalia catappa Leaf Extract: Its Application in EBT Dye Removal, *Chem. Afr.*, 2024, 1–22.
 - 194 N. S. Ali, E. H. Khader, M. A. Abdulrahman, I. K. Salih and T. M. Albayati, Removal of anionic azo dye from wastewater using Fe₃O₄ magnetic nanoparticles adsorbents in a batch system, *Desalin. Water Treat.*, 2024, **317**, 100033.
 - 195 H. Li, H. Jin, R. Li, J. Hua, Z. Zhang and R. Li, Magnetic Fe₃O₄@ SiO₂ study on adsorption of methyl orange on nanoparticles, *Sci. Rep.*, 2024, **14**(1), 1217.
 - 196 S. J. Salih, L. I. Abd Ali and W. M. Hamad, Novel synthesis and characterization of magnesium-doped CoFe₂O₄ nanoparticles–SiO₂–3-aminopropylethoxysilane–gallic acid magnetic nanocomposite for effective removal of cationic dyes, *Arabian J. Chem.*, 2024, **17**(3), 105647.
 - 197 A. Hethnawi, O. Kashif, R. Jeong, F. Sagala, K. Hashlamoun, A. D. Manasrah and N. N. Nassar, Green synthesis of novel titanomagnetite nanoparticles for oil spill cleanup, *Colloids Surf., A*, 2023, **664**, 131191.
 - 198 Y. Vicente-Martínez, M. Caravaca and A. Soto-Meca, Total removal of Hg (II) from wastewater using magnetic nanoparticles coated with nanometric Ag and functionalized with sodium 2-mercaptoethane sulfonate, *Environ. Chem. Lett.*, 2020, **18**(3), 975–981.
 - 199 A. Peigneux, J. D. Puentes-Pardo, A. B. Rodríguez-Navarro, M. T. Hincke and C. Jimenez-Lopez, Development and characterization of magnetic eggshell membranes for lead removal from wastewater, *Ecotoxicol. Environ. Saf.*, 2020, **192**, 110307.
 - 200 M. Govarthanan, C. H. Jeon, Y. H. Jeon, J. H. Kwon, H. Bae and W. Kim, Non-toxic nano approach for wastewater treatment using *Chlorella vulgaris* exopolysaccharides immobilized in iron-magnetic nanoparticles, *Int. J. Biol. Macromol.*, 2020, **162**, 1241–1249.
 - 201 M. Ahmad, J. Wang, J. Xu, Q. Zhang and B. Zhang, Magnetic tubular carbon nanofibers as efficient Cu(II) ion adsorbent from wastewater, *J. Cleaner Prod.*, 2020, **252**, 119825.
 - 202 I. K. Rind, N. Memon, A. Sari, M. Y. Khuhawar, M. Tuzen, S. N. ul Hasan, A. A. Memon, W. A. Soomro, R. O. Z. Brohi and T. A. Saleh, Magnetic nanoparticles loaded hydrochar for effective Cr (VI) removal from water: Batch and column studies, *Mater. Chem. Phys.*, 2024, **318**, 129077.
 - 203 H. Sánchez-Moreno, L. García-Rodríguez and C. Recalde-Moreno, Natural cellulose fibers (*Agave Americana* L. ASPARAGACEAE) impregnated with magnetite nanoparticles as a novel adsorbent of mercury (Hg) in aqueous solutions, *Adsorption*, 2025, **31**(1), 1–22.
 - 204 F. Almomani, R. Bhosale, M. Khraisheh, A. kumar and T. Almomani, Heavy metal ions removal from industrial wastewater using magnetic nanoparticles (MNP), *Appl. Surf. Sci.*, 2020, **506**, 144924.
 - 205 A. K. R. Kumar, K. Saikia, G. Neeraj, H. Cabana and V. V. Kumar, Remediation of bio-refinery wastewater containing organic and inorganic toxic pollutants by adsorption onto chitosan-based magnetic nanosorbent, *Water Qual. Res. J.*, 2020, **55**(1), 36–51.
 - 206 A. Masjedi, E. Askarizadeh and S. Baniyaghoob, Magnetic nanoparticles surface-modified with tridentate ligands for removal of heavy metal ions from water, *Mater. Chem. Phys.*, 2020, **249**, 122917.
 - 207 Q. Xu, W. Li, L. Ma, D. Cao, G. Owens and Z. Chen, Simultaneous removal of ammonia and phosphate using green synthesized iron oxide nanoparticles dispersed onto zeolite, *Sci. Total Environ.*, 2020, **703**, 135002.
 - 208 N. Zhu, B. Zhang and Q. Yu, Genetic engineering-facilitated coassembly of synthetic bacterial cells and magnetic nanoparticles for efficient heavy metal removal, *ACS Appl. Mater. Interfaces*, 2020, **12**(20), 22948–22957.
 - 209 A. N. Hosain, A. El Nemr, A. El Sikaily, M. E. Mahmoud and M. F. Amira, Surface modifications of nanochitosan coated magnetic nanoparticles and their applications in Pb (II), Cu (II) and Cd (II) removal, *J. Environ. Chem. Eng.*, 2020, **8**(5), 104316.
 - 210 Y. Wang, S. Yu, H. Yuan and L. Zhang, Constructing N, S co-doped network biochar confined CoFe₂O₄ magnetic nanoparticles adsorbent: Insights into the synergistic and competitive adsorption of Pb²⁺ and ciprofloxacin, *Environ. Pollut.*, 2024, **343**, 123178.
 - 211 F. H. Alkallas, S. M. Alghamdi, O. Albeydani, E. A. Mwafy, A. B. G. Trabelsi, W. B. Elsharkawy, E. Alsubhe and A. M. Mostafa, Effect of ablation time for loading amounts of magnetic nanoparticles on CNTs for removal of Pb (II) ions from aqueous solution, *Appl. Phys. A: Mater. Sci. Process.*, 2024, **130**(2), 128.
 - 212 J. C. Almeida, C. E. Cardoso, M. C. Neves, T. Trindade, M. G. Freire and E. Pereira, Extraction of platinum and palladium from aqueous solutions using ionic-liquid-modified magnetic nanoparticles, *J. Water Proc. Eng.*, 2024, **63**, 105510.
 - 213 N. Mehrabi, U. F. A. Haq, M. T. Reza and N. Aich, Application of deep eutectic solvent for conjugation of magnetic nanoparticles onto graphene oxide for lead (II)



- and methylene blue removal, *J. Environ. Chem. Eng.*, 2020, **8**(5), 104222.
- 214 A. Saravanan, P. S. Kumar, M. Govarthan, C. S. George, S. Vaishnavi, B. Mouliswaran, S. P. Kumar, S. Jeevanantham and P. R. Yaashikaa, Adsorption characteristics of magnetic nanoparticles coated mixed fungal biomass for toxic Cr(VI) ions in aquatic environment, *Chemosphere*, 2021, **267**, 129226.
 - 215 F. Hassanzadeh-Afruzi, G. Heidari and A. Maleki, Magnetic nanocomposite hydrogel based on Arabic gum for remediation of lead (II) from contaminated water, *Mater. Chem. Horiz.*, 2022, **1**(2), 107–122.
 - 216 N. Karami, A. Mohammadpour, M. R. Samaei, A. M. Amani, M. Dehghani, R. S. Varma and J. N. Sahu, Green synthesis of sustainable magnetic nanoparticles Fe₃O₄ and Fe₃O₄-chitosan derived from *Prosopis farcta* biomass extract and their performance in the sorption of lead(II), *Int. J. Biol. Macromol.*, 2024, **254**, 127663.
 - 217 T. Lei, X. Jiang, Y. Zhou, H. Chen, H. Bai, S. Wang and X. Yang, A multifunctional adsorbent based on 2,3-dimercaptosuccinic acid/dopamine-modified magnetic iron oxide nanoparticles for the removal of heavy-metal ions, *J. Colloid Interface Sci.*, 2023, **636**, 153–166.
 - 218 H. Zeng, L. Zhai, J. Zhang and D. Li, As(V) adsorption by a novel core-shell magnetic nanoparticles prepared with Iron-containing water treatment residuals, *Sci. Total Environ.*, 2021, **753**, 142002.
 - 219 N. Rafie, M. Khodadadi, M. Zamani, A. Zarepour and A. Zarrabi, Magnetic silica nanoparticles adorned with a metal-organic framework; a novel nanosorbent for elimination of aqueous Pb ions contaminant, *Environ. Res.*, 2023, **226**, 115694.
 - 220 A. H. Ragab, M. F. Mubarak, H. A. El-Sabban, J. Kang, A. El Shahawy, H. A. Alshwyeh and M. Hemdan, Exploring the sustainable synthesis pathway and comprehensive characterization of magnetic hybrid alumina nanoparticles phase (MHAI-NPSP) as highly efficient adsorbents and selective copper ions removal, *Environ. Technol. Innovation*, 2024, **34**, 103628.
 - 221 A. H. Birniwa, R. E. A. Mohammad, M. Ali, M. F. Rehman, S. S. a. Abdullahi, S. M. Eldin, S. Mamman, A. C. Sadiq and A. H. Jagaba, Synthesis of gum Arabic magnetic nanoparticles for adsorptive removal of ciprofloxacin: equilibrium, kinetic, thermodynamics studies, and optimization by response surface methodology, *Separations*, 2022, **9**(10), 322.
 - 222 M. Yilmaz, T. J. Al-Musawi, M. k. Saloot, A. D. Khatibi, M. Baniyadi and D. Balarak, Synthesis of activated carbon from *Lemna minor* plant and magnetized with iron (III) oxide magnetic nanoparticles and its application in removal of Ciprofloxacin, *Biomass Convers. Biorefin.*, 2022, 1–14.
 - 223 R. M. Abdelhameed, O. M. Darwesh and M. El-Shahat, Titanium-based metal-organic framework encapsulated with magnetic nanoparticles: Antimicrobial and photocatalytic degradation of pesticides, *Microporous Mesoporous Mater.*, 2023, **354**, 112543.
 - 224 A. Sheikhmohammadi, E. Asgari and J. Yeganeh, Application of Fe₃O₄@ activated carbon magnetic nanoparticles for the adsorption of metronidazole from wastewater: optimization, kinetics, thermodynamics and equilibrium studies, *Desalin. Water Treat.*, 2021, **222**, 354–365.
 - 225 T. J. Al-Musawi, A. H. Mahvi, A. D. Khatibi and D. Balarak, Effective adsorption of ciprofloxacin antibiotic using powdered activated carbon magnetized by iron(III) oxide magnetic nanoparticles, *J. Porous Mater.*, 2021, **28**(3), 835–852.
 - 226 L. F. Cusioli, H. B. Quesada, A. L. d. B. P. Castro, R. G. Gomes and R. Bergamasco, Development of a new low-cost adsorbent functionalized with iron nanoparticles for removal of metformin from contaminated water, *Chemosphere*, 2020, **247**, 125852.
 - 227 M. Rouhani, S. D. Ashrafi, K. Taghavi, M. N. Joubani and J. Jaafari, Evaluation of tetracycline removal by adsorption method using magnetic iron oxide nanoparticles (Fe₃O₄) and clinoptilolite from aqueous solutions, *J. Mol. Liq.*, 2022, **356**, 119040.
 - 228 M. Moghaddam-Manesh, R. Darvishi, A. Barati, A. Moshkriz and M. Seyedsharifi, Synthesis of novel Fe₃O₄@ gly@ Indole@ CuNO₃ magnetic nanoparticle as high-performance antibiotic absorbent, antimicrobial agent, and reusable magnetic Nano catalyst, *J. Environ. Chem. Eng.*, 2025, **13**(1), 114970.
 - 229 J. S. Algethami, K. K. Yadav, A. Gacem, I. H. Ali, S. Rezaia, M. S. O. Alhar, A. Mezni, B.-H. Jeon and S. Chairapat, Magnetic sporopollenin supported magnesium nanoparticles for removal of tetracycline as an emerging contaminant from water, *Environ. Sci. Pollut. Res.*, 2024, **31**(28), 40257–40268.
 - 230 F. Ghadami, M. Valian, F. Atoof, E. A. Dawi, M. B. Miranzadeh, M. A. Mahdi and M. Salavati-Niasari, Response surface methodology for optimization of operational parameters to remove tetracycline from contaminated water by new magnetic Ho₂MoO₆/Fe₂O₃ nano adsorbent, *Results Eng.*, 2024, **21**, 101746.
 - 231 M. Khodadadi, T. J. Al-Musawi, H. Kamani, M. F. Silva and A. H. Panahi, The practical utility of the synthesis FeNi₃@ SiO₂@ TiO₂ magnetic nanoparticles as an efficient photocatalyst for the humic acid degradation, *Chemosphere*, 2020, **239**, 124723.
 - 232 A. B. Abou Hammad, A. M. El Nahwary, B. A. Hemdan and A. L. K. Abia, Nanoceramics and novel functionalized silicate-based magnetic nanocomposites as substitutional disinfectants for water and wastewater purification, *Environ. Sci. Pollut. Res.*, 2020, **27**, 26668–26680.
 - 233 M. Kumari and S. K. Gupta, A novel process of adsorption cum enhanced coagulation-flocculation spiked with magnetic nanoadsorbents for the removal of aromatic and hydrophobic fraction of natural organic matter along with turbidity from drinking water, *J. Cleaner Prod.*, 2020, **244**, 118899.
 - 234 S. C. Kollarahithlu and R. M. Balakrishnan, Adsorption of pharmaceuticals pollutants, Ibuprofen, Acetaminophen,



- and Streptomycin from the aqueous phase using amine functionalized superparamagnetic silica nanocomposite, *J. Cleaner Prod.*, 2021, **294**, 126155.
- 235 R. Natarajan, S. Venkataraman, D. S. Rajendran, B. Tamilselvam, H. Zaveri, N. Jeyachandran, H. Prashar and V. K. Vaidyanathan, Adsorption performance of magnetic mesoporous silica microsphere support toward the remediation of acetaminophen from aqueous solution, *J. Water Proc. Eng.*, 2022, **48**, 102835.
- 236 A. S. Liyanage, S. Canaday, C. U. Pittman Jr and T. Mlsna, Rapid remediation of pharmaceuticals from wastewater using magnetic Fe₃O₄/Douglas fir biochar adsorbents, *Chemosphere*, 2020, **258**, 127336.
- 237 B. D'Cruz, M. Madkour, M. O. Amin and E. Al-Hetlani, Efficient and recoverable magnetic AC-Fe₃O₄ nanocomposite for rapid removal of promazine from wastewater, *Mater. Chem. Phys.*, 2020, **240**, 122109.
- 238 M. E. Peralta, D. O. Mártire, M. S. Moreno, M. E. Parolo and L. Carlos, Versatile nanoadsorbents based on magnetic mesostructured silica nanoparticles with tailored surface properties for organic pollutants removal, *J. Environ. Chem. Eng.*, 2021, **9**(1), 104841.
- 239 R. Natarajan, M. Anil Kumar and V. K. Vaidyanathan, Synthesis and characterization of rhamnolipid based chitosan magnetic nanosorbents for the removal of acetaminophen from aqueous solution, *Chemosphere*, 2022, **288**, 132532.
- 240 L. A. Al-Khateeb, W. Hakami, M. A. Salam, J. A. Sanari, R. El-Shaheny and M. El-Maghrabey, Solid phase-fabrication of magnetically separable Fe₃O₄@ graphene nanoplatelets nanocomposite for efficient removal of NSAIDs from wastewater. Perception of adsorption kinetics, thermodynamics, and extra-thermodynamics, *Anal. Chim. Acta*, 2022, **1223**, 340158.
- 241 G. Wu, Q. Liu, J. Wang, S. Xia, H. Wu, J. Zong, J. Han and W. Xing, Facile fabrication of rape straw biomass fiber/ β -CD/Fe₃O₄ as adsorbent for effective removal of ibuprofen, *Ind. Crops Prod.*, 2021, **173**, 114150.
- 242 B. Yao, Z. Luo, S. Du, J. Yang, D. Zhi and Y. Zhou, Sustainable biochar/MgFe₂O₄ adsorbent for levofloxacin removal: Adsorption performances and mechanisms, *Bioresour. Technol.*, 2021, **340**, 125698.
- 243 Y. Liu, R. Liu, M. Li, F. Yu and C. He, Removal of pharmaceuticals by novel magnetic genipin-crosslinked chitosan/graphene oxide-SO₃H composite, *Carbohydr. Polym.*, 2019, **220**, 141–148.
- 244 M. Sun, Q. Sun, C. Zhao, Y. Huang, J. Jiang, W. Ding and H. Zheng, Degradation of diclofenac sodium with low concentration from aqueous milieu through polydopamine-chitosan modified magnetic adsorbent-assisted photo-Fenton process, *Sep. Purif. Technol.*, 2022, **289**, 120771.
- 245 S. Altaf, R. Zafar, W. Q. Zaman, S. Ahmad, K. Yaqoob, A. Syed, A. J. Khan, M. Bilal and M. Arshad, Removal of levofloxacin from aqueous solution by green synthesized magnetite (Fe₃O₄) nanoparticles using *Moringa olifera*: Kinetics and reaction mechanism analysis, *Ecotoxicol. Environ. Saf.*, 2021, **226**, 112826.
- 246 M. M. Gaho, G. Z. Memon, J. u. R. Memon, J. B. Arain, A. J. Arain, A. Shah and M. Q. Samejo, Synthesis of novel magnetic molecularly imprinted polymers by solid-phase extraction method for removal of norfloxacin, *Chin. J. Anal. Chem.*, 2022, **50**(6), 100079.
- 247 E. Norabadi, A. Jahantig and H. Kamani, Synthesis of Fe-TiO₂@ Fe₃O₄ magnetic nanoparticles as a recyclable sonocatalyst for the degradation of 2, 4-dichlorophenol, *Environ. Sci. Pollut. Res.*, 2023, **30**(11), 31446–31460.
- 248 X. Fu, S. Sarker, W. Ma, W. Zhao, Y. Rong and Q. Liu, Novel phenylalanine-modified magnetic ferroferric oxide nanoparticles for ciprofloxacin removal from aqueous solution, *J. Colloid Interface Sci.*, 2023, **632**, 345–356.
- 249 S. Taghavi Fardood, F. Moradnia and T. M. Aminabhavi, Green synthesis of novel Zn_{0.5}Ni_{0.5}FeCrO₄ spinel magnetic nanoparticles: Photodegradation of 4-nitrophenol and aniline under visible light irradiation, *Environ. Pollut.*, 2024, **358**, 124534.
- 250 C. Anushree, F. A. Rahim, S. C. Vanithakumari, C. Thinaharan and J. Philip, Electrospun superparamagnetic fibrous composite nanofiber films for enhanced oil spill recovery: Effect of capping and magnetic nanoparticle loading on oil sorption efficiency, *Composites, Part A*, 2023, **171**, 107591.
- 251 X. He, Q. Liu and Z. Xu, Treatment of oily wastewaters using magnetic Janus nanoparticles of asymmetric surface wettability, *J. Colloid Interface Sci.*, 2020, **568**, 207–220.
- 252 M. M. Abdullah, M. S. Ali and H. A. Al-Lohedan, Oil spill cleanup employing surface modified magnetite nanoparticles using two new polyamines, *J. Chem.*, 2023, **2023**(1), 7515345.
- 253 M. M. Abdullah, H. A. Al-Lohedan, A. Al-Khwilani and B. M. Al-Maswari, Efficient oil spill uptake using surface-modified magnetite nanoparticles with pet waste derivatives, *ACS Omega*, 2023, **8**(46), 43955–43963.
- 254 R. Malhas, J. H. El Achkar, B. Misbah and S. Al Radhwan, Optimizing Oil Removal from Oil-Water Emulsions Using Novel Iron Oxide Magnetic Nanoparticles, *Water, Air, Soil Pollut.*, 2023, **234**(9), 564.
- 255 W. F. Elmobarak and F. Almomani, Application of magnetic nanoparticles for the removal of oil from oil-in-water emulsion: Regeneration/reuse of spent particles, *J. Pet. Sci. Eng.*, 2021, **203**, 108591.
- 256 M. Abdullah and H. A. Al-Lohedan, Facile fabrication of magnetite nanoparticles with new hydrophobic amides and their application in oil spill remediation, *Environ. Sci. Pollut. Res.*, 2024, **31**(25), 36986–36994.
- 257 F. Rezaei and S. Hassanajili, Facile fabrication of superhydrophobic magnetic bio-waste for oil spill cleanup, *Ind. Crops Prod.*, 2023, **201**, 116848.
- 258 T. Lü, X. Zhang, R. Ma, D. Qi, Y. Sun, D. Zhang, J. Huang and H. Zhao, Quaternary ammonium siloxane-decorated magnetic nanoparticles for emulsified oil-water separation, *Sep. Purif. Technol.*, 2023, **309**, 123097.



- 259 H. Hamed, S. Zendehboudi, N. Rezaei, A. Azizi and F. Shahhoseini, Application of functionalized Fe₃O₄ magnetic nanoparticles using CTAB and SDS for oil separation from oil-in-water nanoemulsions, *Langmuir*, 2023, **39**(23), 7995–8007.
- 260 T. Lü, D. Qi, D. Zhang, K. Fu, Y. Li and H. Zhao, Fabrication of recyclable multi-responsive magnetic nanoparticles for emulsified oil-water separation, *J. Cleaner Prod.*, 2020, **255**, 120293.
- 261 M. Zhang, K. Lian, L. Ai, W. Kang and T. Zhao, Simultaneous determination of 11 antiseptic ingredients in surface water based on polypyrrole decorated magnetic nanoparticles, *RSC Adv.*, 2020, **10**(61), 37473–37481.
- 262 C. Du, Z. Zhang, G. Yu, H. Wu, H. Chen, L. Zhou, Y. Zhang, Y. Su, S. Tan and L. Yang, A review of metal organic framework (MOFs)-based materials for antibiotics removal via adsorption and photocatalysis, *Chemosphere*, 2021, **272**, 129501.
- 263 C. K. Gadupudi, L. Rice, L. Xiao and K. Kantamaneni, Endocrine disrupting compounds removal methods from wastewater in the United Kingdom: a review, *Sci*, 2021, **3**(1), 11.
- 264 G. G. Haciosmanoğlu, C. Mejías, J. Martín, J. L. Santos, I. Aparicio and E. Alonso, Antibiotic adsorption by natural and modified clay minerals as designer adsorbents for wastewater treatment: A comprehensive review, *J. Environ. Manage.*, 2022, **317**, 115397.
- 265 M. W. Nugraha, S. Kim, F. Roddick, Z. Xie and L. Fan, A review of the recent advancements in adsorption technology for removing antibiotics from hospital wastewater, *J. Water Proc. Eng.*, 2025, **70**, 106960.
- 266 S. Ramu, I. Kainthla, L. Chandrappa, J. M. Shivanna, B. Kumaran and R. G. Balakrishna, Recent advances in metal organic frameworks-based magnetic nanomaterials for waste water treatment, *Environ. Sci. Pollut. Res.*, 2024, **31**(1), 167–190.
- 267 N. Shehata, D. Egirani, A. Olabi, A. Inayat, M. A. Abdelkareem, K.-J. Chae and E. T. Sayed, Membrane-based water and wastewater treatment technologies: Issues, current trends, challenges, and role in achieving sustainable development goals, and circular economy, *Chemosphere*, 2023, **320**, 137993.
- 268 Y. Miyah, N. El Messaoudi, M. Benjelloun, J. Georgin, D. S. P. Franco, Y. Acikbas, H. S. Kusuma and M. Sillanpää, MOF-derived magnetic nanocomposites as potential formulations for the efficient removal of organic pollutants from water via adsorption and advanced oxidation processes: A review, *Mater. Today Sustain.*, 2024, **100985**.
- 269 Y. Gadelhak, M. El-Azazy, M. F. Shibl and R. K. Mahmoud, Cost estimation of synthesis and utilization of nano-adsorbents on the laboratory and industrial scales: A detailed review, *Sci. Total Environ.*, 2023, **875**, 162629.
- 270 Z. Deng, S. Cheng, N. Xu, X. Zhang and B. Pan, Pilot-scale field demonstration of environmental nanotechnology for groundwater defluoridation, *ACS ES&T Eng.*, 2022, **3**(2), 226–235.
- 271 X. Zhang, K. Zhang, Y. Shi, H. Xiang, W. Yang and F. Zhao, Surface engineering of multifunctional nanostructured adsorbents for enhanced wastewater treatment: A review, *Sci. Total Environ.*, 2024, **920**, 170951.
- 272 M. Arienzo and L. Ferrara, Environmental fate of metal nanoparticles in estuarine environments, *Water*, 2022, **14**(8), 1297.
- 273 S. Ethaib, S. Al-Qutaifia, N. Al-Ansari and S. L. Zubaidi, Function of nanomaterials in removing heavy metals for water and wastewater remediation: A review, *Environments*, 2022, **9**(10), 123.
- 274 M. A. Mudassir, M. Farhan, S. Kousar and S. Akram, Role of magnetic hybrid/composite nanomaterials for removal of metal ions from aqueous matrices, in *Functionalized Magnetic Nanohybrids*, Elsevier, 2025, pp. 361–388.
- 275 J. Govan, Recent advances in magnetic nanoparticles and nanocomposites for the remediation of water resources, *Magnetochemistry*, 2020, **6**(4), 49.
- 276 B. García-Merino, E. Bringas and I. Ortiz, Synthesis and applications of surface-modified magnetic nanoparticles: Progress and future prospects, *Rev. Chem. Eng.*, 2022, **38**(7), 821–842.
- 277 Standard I., *ISO Environmental Management—Life Cycle Assessment: Principles and Framework*, ISO Standard 2006, p. 14040.
- 278 Standard I., *Environmental Management-Life Cycle Assessment-Requirements and Guidelines*, ISO, London, 2006.
- 279 I. Lynch, T. Cedervall, M. Lundqvist, C. Cabaleiro-Lago, S. Linse and K. A. Dawson, The nanoparticle–protein complex as a biological entity; a complex fluids and surface science challenge for the 21st century, *Adv. Colloid Interface Sci.*, 2007, **134**, 167–174.
- 280 S. Bucak, B. Yavuztürk and A. D. Sezer, Magnetic nanoparticles: synthesis, surface modifications and application in drug delivery, *Recent Adv. Novel Drug Carrier Syst.*, 2012, **2**, 165–200.

