



Cite this: *RSC Adv.*, 2025, **15**, 22376

## Supersolubility and solubility of lithium phosphate in sodium carbonate solution†

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The tail liquid generated from lithium carbonate production in salt lake brine is termed lithium-bearing mother liquor. This mother liquor exhibits a complex composition, with the  $\text{Li}^+$  concentration typically around  $1.5 \text{ g L}^{-1}$ , representing a significant lithium resource. Preparing lithium phosphate ( $\text{Li}_3\text{PO}_4$ ) from this mother liquor is critical for efficient lithium recovery. However, the lack of data on the thermodynamic behavior and  $\text{Li}_3\text{PO}_4$  crystallization in such complex solutions has hindered the high-efficiency recovery of lithium resources. In this study, the solubility of  $\text{Li}_3\text{O}_4$  in sodium carbonate solutions was determined using the dynamic dissolution equilibrium method. The effects of temperature and sodium carbonate concentration on solubility were analyzed, and experimental data were correlated using an exponential equation. Results indicated that the solubility of  $\text{Li}_3\text{PO}_4$  in pure water and sodium carbonate solutions increases with temperature and sodium carbonate concentration. The supersolubility of  $\text{Li}_3\text{PO}_4$  in  $\text{LiCl}-\text{Na}_2\text{CO}_3$  electrolyte solutions was measured via turbidimetric analysis, and the metastable zone width (MSZW) was determined. The supersolubility of  $\text{Li}_3\text{PO}_4$  significantly decreased with rising temperature. In contrast, supersolubility initially increased and then decreased with higher  $\text{Na}_2\text{CO}_3$  concentrations, with reactant concentration being the decisive factor driving the crystallization reaction. Furthermore, the MSZW narrowed at elevated temperatures. Thermodynamic functions ( $\Delta S_d$ ,  $\Delta H_d$ , and  $\Delta G_d$ ) for the dissolution process were calculated via the van't Hoff equation, confirming that  $\text{Li}_3\text{PO}_4$  dissolution is a spontaneous and endothermic process. Based on solubility and supersolubility data, a novel process was developed to prepare battery-grade  $\text{Li}_3\text{PO}_4$  (purity: 99.80%) from salt lake mother liquor. The results of Raman, FTIR, TG and SEM suggested that the prepared lithium phosphate was pure phase. This study provides fundamental physicochemical data and theoretical insights for the efficient separation and extraction of lithium resources from lithium precipitation mother liquor.

Received 18th April 2025  
 Accepted 17th June 2025

DOI: 10.1039/d5ra02716k  
[rsc.li/rsc-advances](http://rsc.li/rsc-advances)

### 1. Introduction

The rapid growth of new global energy industries has underscored the importance of developing and utilizing lithium resources, which play a pivotal role in renewable energy applications.<sup>1,2</sup> As the world transitions towards carbon neutrality, lithium has emerged as a crucial raw material for lithium-ion batteries, making it indispensable for electric vehicles, energy storage systems, and various electronic devices, resulting in a significant increase in market demand.<sup>3–5</sup> Consequently, ensuring the efficient development and utilization of lithium resources and maintaining a stable supply chain have become

a shared objective among researchers. Current methods for lithium extraction primarily include the recovery of lithium from salt lake brines and mineral ores.<sup>6–8</sup> Salt lake brine extraction techniques typically involve processes such as solvent extraction,<sup>9,10</sup> ion exchange,<sup>11,12</sup> precipitation,<sup>13,14</sup> and membrane separation<sup>15–17</sup> to concentrate and separate lithium ions, ultimately yielding lithium compounds. China possesses abundant lithium resources in its salt lakes, with their reserves accounting for 83% of the nation's total lithium resources.<sup>18</sup> However, challenges such as harsh environmental conditions, varying brine grades, and high  $\text{Mg}/\text{Li}$  ratios complicate the extraction process.<sup>7,19</sup> Addressing these challenges and advancing extraction technology are crucial for the efficient utilization of salt lake resources and for supporting the development of the new energy sector. The mainstream processes for lithium extraction from salt lakes with high magnesium-to-lithium ratios primarily include the following steps: magnesium–lithium separation, impurity removal from lithium-bearing solutions, and lithium carbonate precipitation/

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† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d5ra02716k>



conversion. Among these, the lithium carbonate precipitation step involves adding sodium carbonate solution to a lithium-enriched solution to produce lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) *via* precipitation. Solid–liquid separation subsequently generates a lithium carbonate mother liquor, also referred to as brine lithium extraction mother liquor. This alkaline mother liquor retains a lithium concentration of 1.0 to 2.5 g L<sup>-1</sup> and contains major components such as  $\text{Na}^+$ ,  $\text{Li}^+$ ,  $\text{Cl}^-$ , and  $\text{CO}_3^{2-}$ . As a critical resource, lithium urgently requires efficient recovery and utilization.<sup>20–23</sup>

$\text{Li}_3\text{PO}_4$ , a high value-added lithium compound, has gained attention due to its applications in lithium-ion batteries,<sup>24,25</sup> catalysts,<sup>26,27</sup> and other advanced materials.<sup>28,29</sup> With its low solubility ( $K_{\text{sp}} = 2.37 \times 10^{-11}$  at 25 °C),  $\text{Li}_3\text{PO}_4$  can precipitate efficiently from solutions even at low lithium concentrations, making it a promising candidate for enhancing lithium recovery rates.<sup>30,31</sup> Liu Dongfu *et al.*<sup>32</sup> prepared an anolyte solution by electrochemically deintercalating lithium from salt lake brine with a high Mg/Li ratio, using phosphate precipitation to remove impurities. By optimizing factors such as the initial lithium concentration, precipitation time, precipitant dosage, and reaction temperature, combined with anhydrous ethanol and seed crystal induction, a lithium precipitation rate of 82.5% was achieved, yielding a relatively pure  $\text{Li}_3\text{PO}_4$  product. Sun Jianzhi *et al.*<sup>33</sup> utilized tailings from salt lake lithium extraction to precipitate  $\text{Li}^+$  as  $\text{Li}_2\text{CO}_3$ , which was then mixed with  $\text{Na}_3\text{PO}_4$  solution, adjusted to pH 11, and reacted in a high-pressure reactor at 120–150 °C for 3 to 10 hours. The resulting  $\text{Li}_3\text{PO}_4$  particles, with sizes of 5–8  $\mu\text{m}$ , enabled a lithium yield of 90% to be achieved. Furthermore, Zhu Jun *et al.*<sup>34</sup> produced high-purity, uniform lithium phosphate powder by removing impurities from crude lithium phosphate using oxalate precipitation and a two-stage ion exchange process. These studies provide a solid foundation for enhancing lithium resource utilization efficiency to meet diverse market needs.

Generally, in the crystallization process of  $\text{Li}_3\text{PO}_4$ , solubility, supersolubility, and metastable zone width (MZW) are core parameters requiring critical consideration. Industrial crystallization is typically controlled within the metastable zone to obtain products with high purity, high yield, ideal morphology, and uniform particle size distribution. Supersolubility and metastable zone width are often influenced by factors such as temperature, impurities, feeding rate, and solution kinetics. The mother liquor after  $\text{Li}_2\text{CO}_3$  precipitation contains abundant carbonate ions, which may exert complex effects on  $\text{Li}_3\text{PO}_4$  crystallization behavior. To our knowledge, no data on  $\text{Li}_3\text{PO}_4$  solubility or supersolubility in the carbonate impurity system have been documented. Additionally, reported data on  $\text{Li}_3\text{PO}_4$  solubility in aqueous solutions at the same temperature ranges have discrepancies, possibly due to the hydrolysis of phosphate ions. Specifically, the solubilities determined at 298.15 K were reported as 0.0270%,<sup>35</sup> 0.0297% and 0.0239%.<sup>36</sup> In this study, the influence of experimental parameters (*e.g.*, temperature and sodium carbonate concentration) on  $\text{Li}_3\text{PO}_4$  solubility, supersolubility, and MZW was investigated. Based on the observed patterns of  $\text{Li}_3\text{PO}_4$  solubility, supersolubility, and MZW

variations, a battery-grade  $\text{Li}_3\text{PO}_4$  precipitate from the lithium-bearing mother liquor was obtained.

## 2. Experimental

### 2.1 Chemicals

The reagents used in this study included lithium phosphate ( $\text{Li}_3\text{PO}_4$ , 99.9%, McLean), sodium carbonate ( $\text{Na}_2\text{CO}_3$ , 99.9%, McLean), lithium chloride ( $\text{LiCl}$ ) (99.9%, Macklin), sodium hydroxide ( $\text{NaOH}$ ) (99.0%, Macklin), and concentrated hydrochloric acid (36–38%, Macklin). All chemicals were used as received without further purification. Deionized water (resistivity 18.25 MΩ cm) was prepared with an ultra-pure water preparation system (UPT-II-20T, Chengdu Ultra-Pure Technology Co., Ltd).

### 2.2 Solubility determination

The solubility of  $\text{Li}_3\text{PO}_4$  in  $\text{Na}_2\text{CO}_3$  solutions was determined using a static isothermal dissolution method. A series of  $\text{Na}_2\text{CO}_3$  solutions with different mass fractions were prepared and transferred into polytetrafluoroethylene (PTFE) bottles. These bottles were placed in a thermostatic water bath with a temperature control accuracy of  $\pm 0.01$  °C, and heated to the desired temperature. Subsequently, 1.5 g of solid  $\text{Li}_3\text{PO}_4$  was added to each bottle and the solutions were stirred at a constant temperature until equilibrium was reached. Once equilibrium was achieved, the solid and liquid phases were separated and the lithium ion concentration in the liquid phase was measured using inductively coupled plasma optical emission spectrometry (ICP-OES). The measured lithium ion concentrations were then used to calculate the solubility of  $\text{Li}_3\text{PO}_4$  in each solution. The solid phase was analyzed by XRD.

### 2.3 Supersolubility and solubility determination

Fig. 1(a) displays the CrystalSCANPolyBlock system (E1061, HEL LIMITED) used for measuring supersolubility. The crystallizer consisted of a 100 mL glass reactor with an internal overhead stirrer and temperature and turbidity sensors controlled by the PolyBlock. The turbidity sensor detected crystal nuclei with an IR laser reflected by an optical lens. The temperature was controlled with the PolyBlock through a thermostatic bath (FP-50, JULABO Labortechnik GmbH).

To measure supersolubility, 70.0 g (2 g L<sup>-1</sup>)  $\text{LiCl}$  solution was settled in the crystallizer, and the overhead stirrer, temperature probe and turbidity probe were immersed in the solutions. The CrystalSCANPolyBlock system and thermostatic bath were then initiated with a stirring rate of 250 rpm and the solution temperature was kept constant. A  $\text{Na}_3\text{PO}_4$  solution (wt 8%) was pumped into the  $\text{LiCl}$  solution, except when assessing the impact of  $\text{Na}_2\text{CO}_3$  concentration on  $\text{Li}_3\text{PO}_4$  supersolubility, *via* a liquid dosing system at a feeding rate of 1 mL min<sup>-1</sup>. Upon nucleation, as indicated by a rapid rise in the turbidity curve, pumping of the solution was stopped and the time interval between the initial pumping and nucleation was recorded as the pumping time. Supersolubility was determined based on the added amount of  $\text{Na}_3\text{PO}_4$ , which was calculated based on



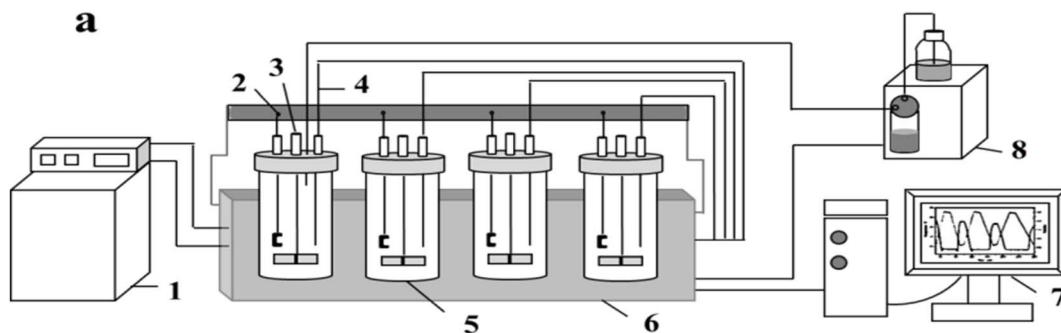


Fig. 1 1-Low-temperature thermostat tank; 2-turbidity probe; 3-suspension stirring paddle; 4-temperature probe; 5-crystallization reactor; 6-four-channel reaction platform; 7-control system; 8-liquid dosing system.

concentration, feeding rate, and pumping time. Each measurement was repeated twice to verify the experimental reproducibility. The supersolubility and solubility of  $\text{Li}_3\text{PO}_4$  in electrolyte solutions are represented by  $c_{\text{super}}$  and  $c_{\text{sol}}$ , respectively. The difference between the supersolubility and solubility is defined as the MSZW ( $\Delta c = c_{\text{super}} - c_{\text{sol}}$ ).

#### 2.4 Preparation of battery-grade lithium phosphate

The process flow for the preparation of battery-grade  $\text{Li}_3\text{PO}_4$  is shown in Fig. 1. The preparation process uses a fully automatic anti-solvent crystallization screening instrument as the synthesis equipment. The equipment allows for precise control of reaction conditions such as temperature, stirring speed, and feeding speed through program settings. The raw material was lithium-bearing mother liquor from a salt lake in the Qaidam Basin (composition shown in Table 1). First, concentrated hydrochloric acid was used to adjust the pH of the lithium precipitation mother liquor to 7.0 to remove some carbonate ions. Then, a 30% sodium hydroxide solution was added to adjust the pH to 13.0, followed by the dropwise addition of 8% sodium phosphate for the precipitation reaction. The crude product obtained was further processed using a hydrochloric acid recrystallization method to produce battery-grade  $\text{Li}_3\text{PO}_4$ .

FTIR spectroscopy analysis was conducted with a Nicolet Nexus 670 FTIR spectrophotometer (Thermo Nicolet Corporation, Madison, WI, USA) in solid films using KBr salt tablets in a range of 500–4000  $\text{cm}^{-1}$ . Raman spectra were recorded at 25 °C with a Raman spectrometer (DXR, Thermo Fisher Scientific, USA).

The thermal stability of  $\text{Li}_3\text{PO}_4$  was measured by TG-DSC (Mettler Toledo, TGA/DSC3+). The measurements were carried at a temperature range of 30–900 °C with a heating rate of 10 °C  $\text{min}^{-1}$ . The blowing gas was nitrogen with a blowing flow rate of 50  $\text{mL min}^{-1}$ . A 100  $\mu\text{L}$  platinum crucible with a perforated cover was used for measurement.

Table 1 Lithium-bearing mother liquor composition ( $\text{g L}^{-1}$ )

Sample	$\rho$ ( $\text{g cm}^{-3}$ )	$\text{Li}^+$	$\text{Mg}^{2+}$	$\text{Ca}^{2+}$	$\text{Na}^+$	$\text{Cl}^-$	$\text{CO}_3^{2-}$	$\text{OH}^-$	$\text{B}_2\text{O}_3$	$\text{SO}_4^{2-}$
Mother liquor	1.1685	1.36	0.002	0.003	49.39	60.46	17.33	2.35	0.44	0.17

The morphology of  $\text{Li}_3\text{PO}_4$  was examined by SEM (JSM-5610LV, JEOL, Japan) in combination with energy dispersive X-ray spectroscopy mapping (X-MAXN).

## 3 Results and discussion

### 3.1 Solubility of lithium phosphate in $\text{Li}_2\text{CO}_3$ solution

Based on the data listed in Table S1,† when the dissolution equilibrium time was more than 48 h, the solubility of  $\text{Li}_3\text{PO}_4$  remained stable. Thus, a 48-hour equilibrium time was adopted to determine the  $\text{Li}_3\text{PO}_4$  solubility in this study.

Table 2 shows the solubility of  $\text{Li}_3\text{PO}_4$  in aqueous solution ranging from 298.15 K to 353.15 K. It is noted that the solubility increased with increasing temperature. A solubility of 0.0244% at 298.15 K was consistent with the reported value,<sup>36</sup> but marginally higher than that reported<sup>37</sup> at higher temperature. This is because the reported values showed a fluctuating correlation with the rising temperature.

Using the solubility exponential model equation (eqn (1)), the data across various temperatures were fitted as depicted in Fig. 2 with a correlation coefficient of  $R^2 = 0.9837$ . The relative errors between the calculated solubility values and the experimental data are presented in Table 3, with all errors within

Table 2 Solubility of  $\text{Li}_3\text{PO}_4$  in aqueous solution at 293.15 K–353.15 K

T/K	Solubility of $\text{Li}_3\text{PO}_4$ (%)	Solid phase
298.15	0.0244	$\text{Li}_3\text{PO}_4$
303.15	0.0283	$\text{Li}_3\text{PO}_4$
313.15	0.0327	$\text{Li}_3\text{PO}_4$
323.15	0.0358	$\text{Li}_3\text{PO}_4$
333.15	0.0376	$\text{Li}_3\text{PO}_4$
343.15	0.0405	$\text{Li}_3\text{PO}_4$
353.15	0.0437	$\text{Li}_3\text{PO}_4$

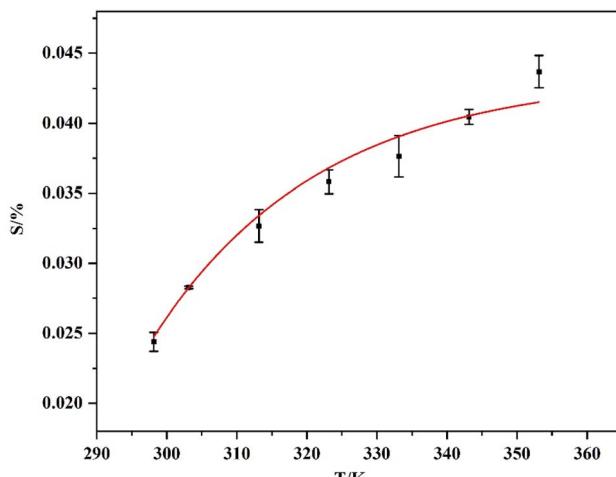


Fig. 2 Exponential form of the solubility of lithium phosphate in water as a function of temperature.

Table 3 Experimental values for the solubility of lithium phosphate in water, calculated values in exponential form and relative errors

Temperature/K	Experimental value (%)	Calculated value (%)	Relative error (%)
298.15	0.0244	0.0251	2.73
303.15	0.0283	0.0276	-2.27
313.15	0.0327	0.0320	-1.95
323.15	0.0358	0.0356	-0.52
333.15	0.0377	0.0386	2.48
343.15	0.0405	0.0410	1.24
353.15	0.0437	0.0430	-1.73

±1.7%. The linear relationship between the experimental and calculated values is shown in Fig. 3, demonstrating that the model equation effectively describes the solubility behavior of lithium phosphate in water.

$$S (\%) = 0.0517 - 11.12 \exp(-0.0202T) \quad (1)$$

$$\varepsilon = \left( m_{\text{Li}_3\text{PO}_4}^{\text{cal}} - m_{\text{Li}_3\text{PO}_4}^{\text{exp}} \right) / m_{\text{Li}_3\text{PO}_4}^{\text{exp}} \quad (2)$$

The solubility of  $\text{Li}_3\text{PO}_4$  in  $\text{Na}_2\text{CO}_3$  solutions was measured from 303.15 K to 353.15 K, with results displayed in Fig. 4 and Table 4. The relative deviations between experimental and calculated values are shown in Tables S2–S8,<sup>†</sup> with all errors within ±2.7%. The linear relationship between the experimental and calculated values is shown in Fig. S1,<sup>†</sup> indicating a good fitting model for the solubility of  $\text{Li}_3\text{PO}_4$  in  $\text{Na}_2\text{CO}_3$  solutions.

It is noticeable that the addition of  $\text{Na}_2\text{CO}_3$  significantly increased the solubility of  $\text{Li}_3\text{PO}_4$ , especially at higher dosages. However, when the mass fraction of  $\text{Na}_2\text{CO}_3$  was more 15%, the solubility appeared to be near to the solubility limit values. This is because higher amounts of  $\text{Na}_2\text{CO}_3$  would cause precipitation of  $\text{Li}_2\text{CO}_3$ , thus leading to a limit of solubility of the  $\text{Li}_3\text{PO}_4$  in

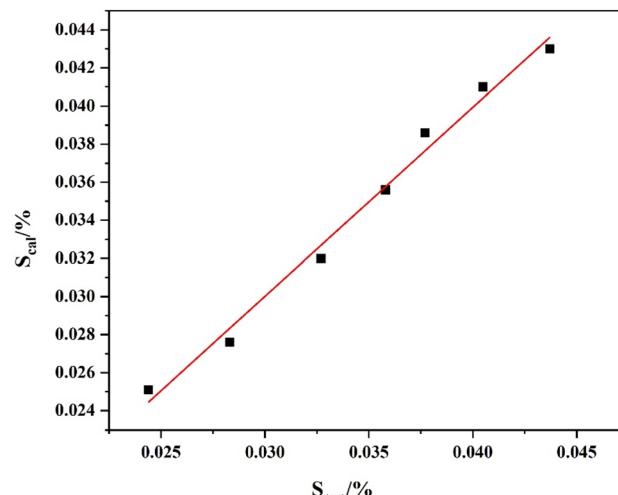


Fig. 3 Correlation between the experimental and exponential calculated values of lithium phosphate solubility in water at different temperatures.

solution. Thus, based on the salting-in effect of the  $\text{Na}_2\text{CO}_3$  on the  $\text{Li}_3\text{PO}_4$  solubility, it is necessary to remove carbonate ions from the lithium-bearing mother liquor to enable high yields of  $\text{Li}_3\text{PO}_4$  during the crystallization process.

### 3.2 Supersolubility and MSZW of $\text{Li}_3\text{PO}_4$

In the experiments, the concentrations of  $\text{LiCl}$  solution, feeding rate, and stirring speed were set to  $c_{\text{Li}^+} = 2 \text{ g L}^{-1}$ ,  $1 \text{ mL min}^{-1}$ , and 250 rpm, respectively. As shown in Fig. 5, the supersolubility of  $\text{Li}_3\text{PO}_4$  first increases and then decreases with the increase in  $\text{Na}_2\text{CO}_3$  solution concentration, reaching a maximum value when the  $\text{Na}_2\text{CO}_3$  concentration is 3%. This is mainly because the reactant concentration is the decisive factor driving the crystallization reaction. When the  $\text{Na}_2\text{CO}_3$  concentration is low, the nuclei of  $\text{Li}_3\text{PO}_4$  dissolve into the solution

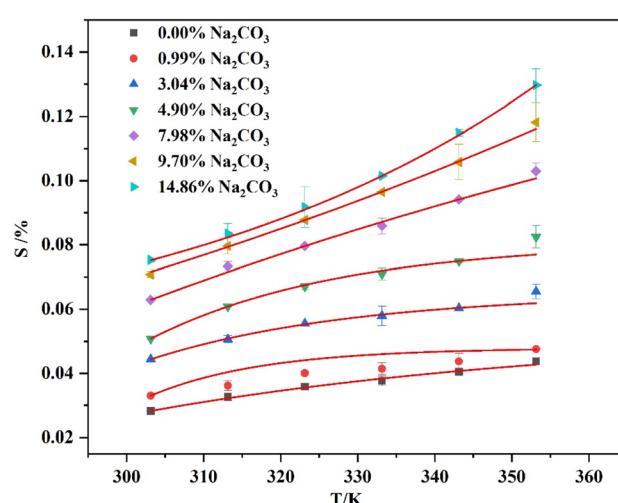
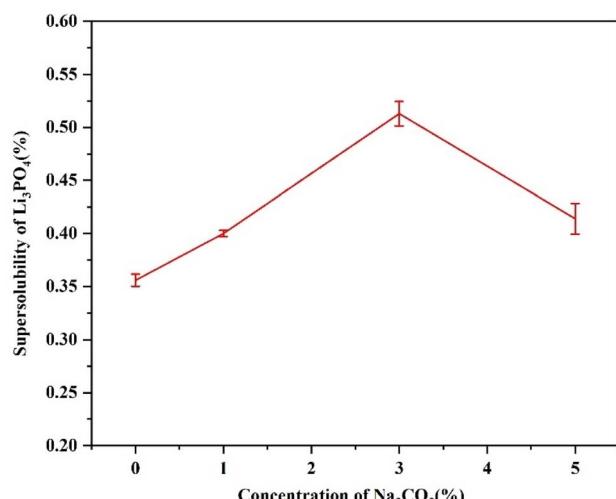


Fig. 4 Solubility of  $\text{Li}_3\text{PO}_4$  in sodium carbonate solutions as a function of temperature.



**Table 4** Exponential solubility equations of  $\text{Li}_3\text{PO}_4$  in sodium carbonates solutions at 303.15 K–353.15 K

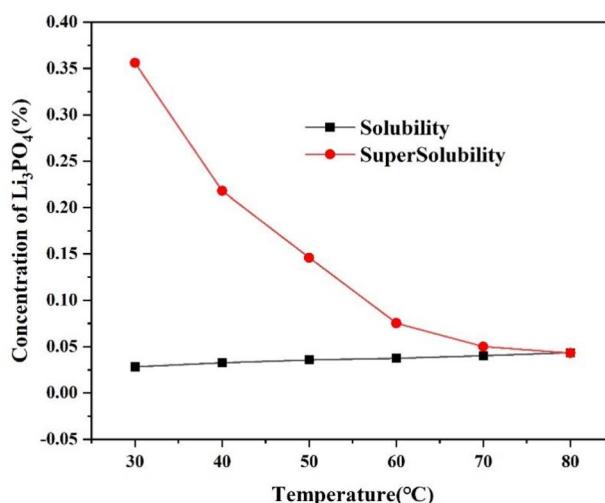
Sodium carbonate concentration (%)	Exponential equation	$R^2$
0.99	$S = 0.0700 - 0.6271\exp(-0.0093T)$	0.9739
3.04	$S = 0.0798 - 5.3270\exp(-0.0166T)$	0.9736
4.90	$S = 0.1057 - 6.1802\exp(-0.0156T)$	0.9720
7.98	$S = 0.3899 - 0.6968\exp(-0.0025T)$	0.9899
9.70	$S = -0.0093 + 0.0051\exp(0.0091T)$	0.9806
14.86	$S = 0.0393 + 0.0001\exp(0.0182T)$	0.9993



**Fig. 5** The effect of sodium phosphate concentration on the supersolubility of lithium phosphate at 30 °C.

before being detected by the probe, delaying the crystallization reaction and leading to an increase in supersolubility. However, as the  $\text{Na}_2\text{CO}_3$  concentration continues to increase, the higher concentration accelerates the crystallization reaction and nucleation rate, causing the nuclei to grow to a detectable size in a short time, which results in a decrease in supersolubility. Therefore, to ensure a wider metastable zone during the crystallization process, a  $\text{Na}_2\text{CO}_3$  solution concentration of 3% was selected for the experiments.

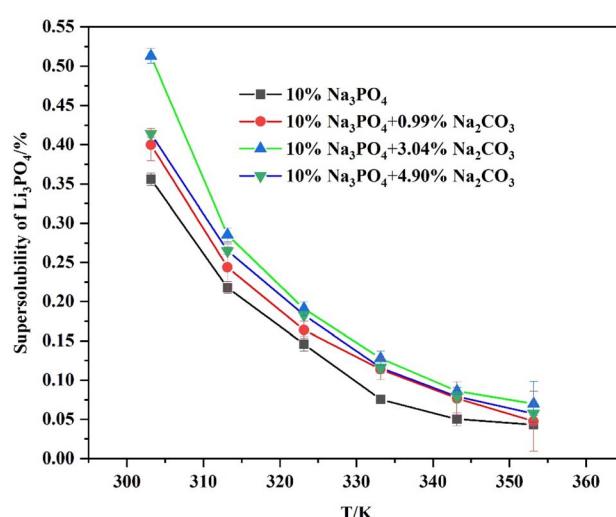
Fig. 6 shows the solubility and supersolubility curves of  $\text{Li}_3\text{PO}_4$  in a  $\text{LiCl}$  solution ( $c_{\text{Li}^+} = 2 \text{ g L}^{-1}$ ). As can be seen from Fig. 6, compared to solubility, the supersolubility of  $\text{Li}_3\text{PO}_4$  is more significantly affected by temperature, decreasing as the temperature increases. When the reaction temperature is 30 °C, the supersolubility is almost 15 times the solubility, whereas it became almost equal to the solubility at 80 °C, indicating that higher reaction temperatures can be used for lithium phosphate in practical production to obtain high yields. Supersolubility is the main driving force for crystallization and is influenced by various factors such as temperature, solute concentration, stirring conditions, feeding rate, and cooling rate. An increase in temperature promotes the movement of molecules or ions, thereby increasing the collision frequency



**Fig. 6** The supersolubility and solubility of  $\text{Li}_3\text{PO}_4$  in  $\text{LiCl}$  solution ( $c_{\text{Li}^+} = 2 \text{ g L}^{-1}$ ).

and mass transfer rate between particles, accelerating nucleation, and leading to a reduction in supersolubility.

Fig. 7 and 8 showed the variation curves of the supersolubility and MSZW of  $\text{Li}_3\text{PO}_4$  with the concentration of sodium carbonate. The supersolubility and MSZW of  $\text{Li}_3\text{PO}_4$  first increased and then decreased with the addition of  $\text{Na}_2\text{CO}_3$ , which was not consistent with the effect on the solubility. According to Table 1, the equilibrium content of  $\text{Li}^+$  and  $\text{CO}_3^{2-}$  ions in the mother liquor was  $1.36 \text{ g L}^{-1}$  and  $17.33 \text{ g L}^{-1}$  (about 2.62% in the form of  $\text{Na}_2\text{CO}_3$ ), respectively. In our system, the lithium content used was  $2 \text{ g L}^{-1}$ , and the  $\text{Na}_2\text{CO}_3$  concentration varied from 0.99%, 3.04% to 4.90%. The higher concentration of 4.90% may facilitate the  $\text{Li}_2\text{CO}_3$  precipitation during the  $\text{Li}_3\text{PO}_4$  crystallization, thus leading to a decrease in the supersolubility of  $\text{Li}_3\text{PO}_4$ , especially at lower temperatures. This



**Fig. 7** The supersolubility curves of  $\text{Li}_3\text{PO}_4$  in  $\text{Na}_2\text{CO}_3$  systems with different concentrations.



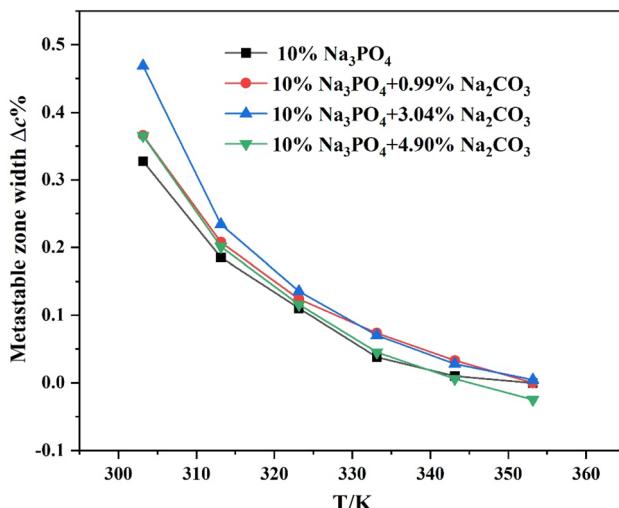


Fig. 8 Metastable zone width of Li<sub>3</sub>PO<sub>4</sub> in Na<sub>2</sub>CO<sub>3</sub> electrolyte solution.

indicates that the role of CO<sub>3</sub><sup>2-</sup> ion impurities in the crystallization process of Li<sub>3</sub>PO<sub>4</sub> cannot be ignored. Therefore, to prepare the Li<sub>3</sub>PO<sub>4</sub> product with high yield and purity, the removal of CO<sub>3</sub><sup>2-</sup> ion impurities from the mother liquor is required.

### 3.3 The thermodynamic properties of Li<sub>3</sub>PO<sub>4</sub> dissolution in Na<sub>2</sub>CO<sub>3</sub> solutions

The Van't Hoff equation (eqn (3)) reveals a linear relationship between the logarithm of the molar fraction of the solute and the reciprocal of absolute temperature. By assuming that the enthalpy (ΔH<sub>d</sub>) and entropy (ΔS<sub>d</sub>) of dissolution remain constant over the temperature range studied, the thermodynamic properties of Li<sub>3</sub>PO<sub>4</sub> during dissolution in Na<sub>2</sub>CO<sub>3</sub> solutions were calculated and shown in Table 5. The Gibbs free energy change (ΔG<sub>d</sub>) during dissolution was also determined from eqn (4), as shown in Fig. 9.

$$\ln x = -\frac{\Delta H_d}{RT} + \frac{\Delta S_d}{R} \quad (3)$$

$$\Delta G_d = \Delta H_d - \Delta S_d T \quad (4)$$

The positive ΔG<sub>d</sub> values indicate that the dissolution process is endothermic, and an increase in temperature favors the

Table 5 The calculated ΔS<sub>d</sub> and ΔH<sub>d</sub> of Li<sub>3</sub>PO<sub>4</sub> in Na<sub>2</sub>CO<sub>3</sub> solutions

ω <sub>Na<sub>2</sub>CO<sub>3</sub></sub> (%)	ΔS <sub>d</sub> (J mol <sup>-1</sup> )	ΔH <sub>d</sub> kJ (mol <sup>-1</sup> K <sup>-1</sup> )
0	-55.44	-8.34
0.99	-61.49	-6.20
3.04	-57.88	-6.44
4.90	-51.50	-7.96
7.98	-48.22	-7.34
9.70	-45.47	-8.93
14.86	-42.46	-9.59

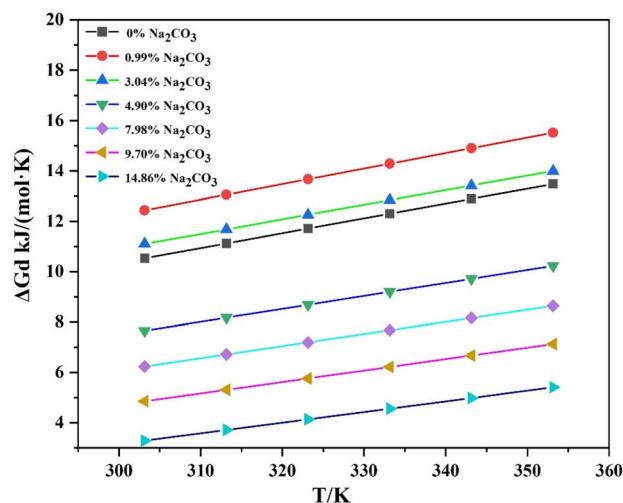


Fig. 9 Plot of the Gibbs free energy (ΔG<sub>d</sub>) change versus temperature.

dissolution of Li<sub>3</sub>PO<sub>4</sub>. As the Na<sub>2</sub>CO<sub>3</sub> concentration increased, the ΔG<sub>d</sub> values decreased, suggesting that higher Na<sub>2</sub>CO<sub>3</sub> concentrations reduced the energy barrier for Li<sub>3</sub>PO<sub>4</sub> dissolution, thereby enhancing its solubility by the salting-in effect.

### 3.4 Preparation of battery-grade Li<sub>3</sub>PO<sub>4</sub>

Based on the solubility and supersolubility of Li<sub>3</sub>PO<sub>4</sub> in the Na<sub>2</sub>CO<sub>3</sub> solutions, it is known that the effect of CO<sub>3</sub><sup>2-</sup> ions on the crystallization process of Li<sub>3</sub>PO<sub>4</sub> cannot be ignored. It is necessary to remove the CO<sub>3</sub><sup>2-</sup> ion impurities from the mother liquor. Therefore, in this experiment, we first removed the CO<sub>3</sub><sup>2-</sup> ions by adding hydrochloric acid until the mother liquor reached approximately pH 7.0; this was followed by the drop-wise addition of an 8% sodium phosphate solution to prepare lithium phosphate. The yield and purity of lithium were investigated at 30 °C (sample LTS1) and 80 °C (sample LTS2). The lithium yield at 30 °C was only 15.23% after 24 hours, while it reached to 82.94% at 80 °C after 30 minutes. This is because Li<sub>3</sub>PO<sub>4</sub> has a high supersolubility at low temperatures, which is not favorable for the crystal nucleation and growth of Li<sub>3</sub>PO<sub>4</sub>. The chemical analyses of the lithium phosphate product are shown in Table 6. Based on the content of PO<sub>4</sub><sup>3-</sup>, the contents of Li<sub>3</sub>PO<sub>4</sub> were calculated to be 84.05% and 80.898% for LTS1 and LTS2, respectively, with sodium ions as the primary impurity. Therefore, we purified the crude product obtained above using a hydrochloric acid recrystallization method (the process flow is shown in Fig. 10). The phase of the resulting

Table 6 Composition of Li<sub>3</sub>PO<sub>4</sub>

Sample	Li %	PO <sub>4</sub> %	Na %	K %	Mg %	Ca %
LTS1	16.04	68.93	2.48	0.053	0.052	0.059
LTS2	15.63	66.34	2.55	0.054	0.053	0.061
LTS3	18.15	79.96	0.078	0.035	0.012	0.066



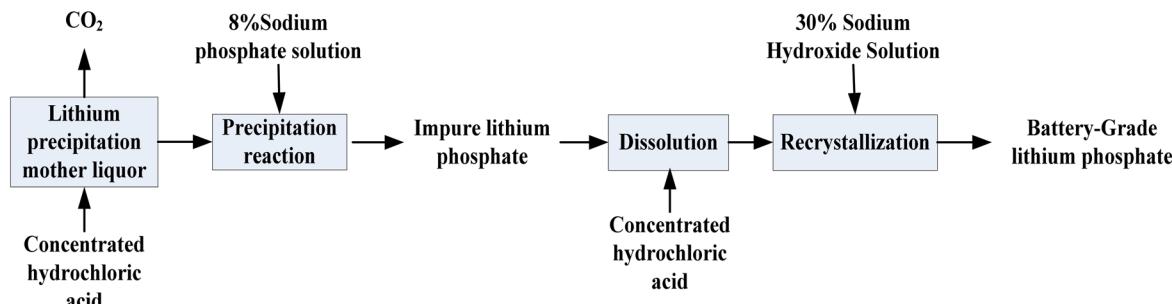


Fig. 10 Process flow diagram for lithium phosphate preparation.

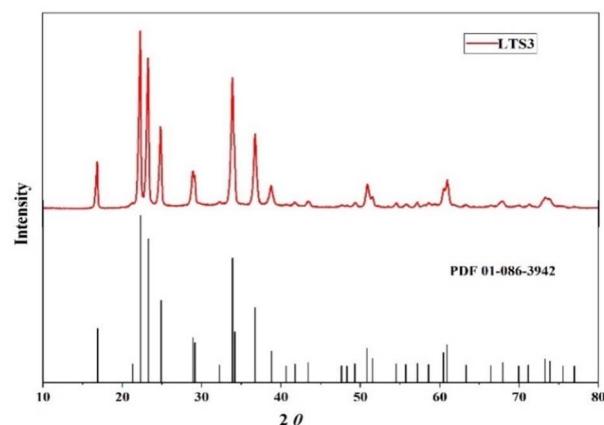
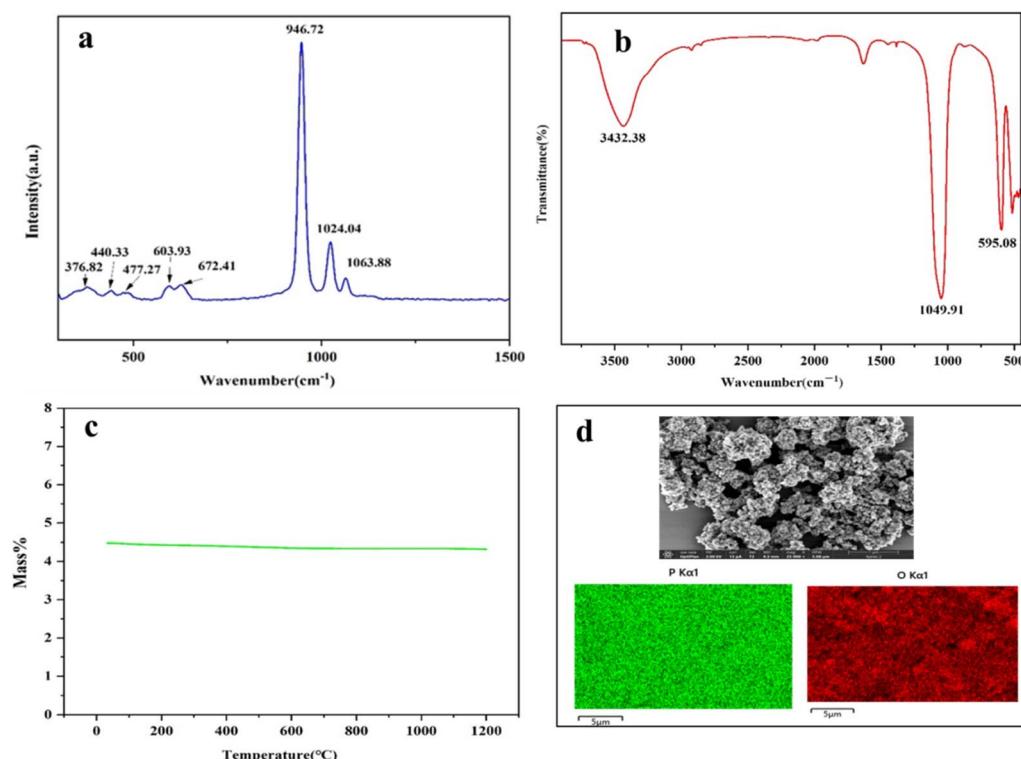
Fig. 11 XRD pattern of  $\text{Li}_3\text{PO}_4$ .

Fig. 12 Raman (a), FTIR (b), TG (c) and SEM images (d) of lithium phosphate at 25 °C.



about 477.27 and 440.33  $\text{cm}^{-1}$  was assigned to Li–O stretching vibrations.<sup>38,39</sup> The FTIR spectrum of lithium phosphate is shown in Fig. 12(b), which reveals characteristic absorption peaks at 1049.91  $\text{cm}^{-1}$  and 541.8  $\text{cm}^{-1}$  corresponding to stretching and asymmetrical stretching vibrations of  $\text{PO}_4^{3-}$ , respectively. The results of TG experiments with lithium phosphate, shown in Fig. 12(c), reveal that lithium phosphate remained undecomposed in the range of 30 °C to 1200 °C. From Raman, FTIR, TG and SEM analyses, it was concluded that the prepared lithium phosphate was pure phase. Fig. 12(d) shows an SEM image and EDS mapping spectrum of lithium phosphate. It can be clearly observed that the morphology of the lithium phosphate is non-uniform and blocky. The position and distribution of P and O elements can be clearly observed from Fig. 12(d). The content of P in lithium phosphate was lower than that of O, which was consistent with the composition of  $\text{PO}_4^{3-}$ . From Raman, FTIR, TG and SEM analyses, it was concluded that the prepared lithium phosphate was pure phase.

## 4. Conclusions

In this study, the solubility data of  $\text{Li}_3\text{PO}_4$  in 0% to 14.86%  $\text{Na}_2\text{CO}_3$  solutions was determined within the temperature range of 303.15 K to 353.15 K by using the dynamic dissolution equilibrium method. The effects of temperature and sodium carbonate concentration on the solubility of lithium phosphate were investigated, and the experimental data were correlated using an index equation. The results indicate that the solubility of lithium phosphate in pure water and sodium carbonate increases with rising temperature, but the change in solubility in pure water is relatively small. The solubility of lithium phosphate increases with the concentration of sodium carbonate due to two factors: the solubility of lithium carbonate being greater than that of  $\text{Li}_3\text{PO}_4$  at the same temperature, and the salting-out effect of  $\text{Na}_2\text{CO}_3$ , which enhances the solubility of lithium phosphate. This suggests that high concentrations of  $\text{Na}_2\text{CO}_3$  solution reduce the recovery rate of  $\text{Li}_3\text{PO}_4$ . The index equation can effectively describe the solubility properties of lithium phosphate. The supersolubility of lithium phosphate in different concentrations of sodium carbonate electrolyte solutions was measured using the turbidity method, and the MZW of  $\text{Li}_3\text{PO}_4$  was calculated. The supersolubility of  $\text{Li}_3\text{PO}_4$  is significantly influenced by temperature, decreasing as temperature increases, indicating that nucleation is slower at lower temperatures. The supersolubility of  $\text{Li}_3\text{PO}_4$  first increases and then decreases with the concentration of  $\text{Na}_2\text{CO}_3$  solution, reaching a maximum when the concentration of  $\text{Na}_2\text{CO}_3$  is 3%, due to the reactant concentration being the decisive factor driving the crystallization reaction. The MZW of  $\text{Li}_3\text{PO}_4$  decreases with increasing temperature. The thermodynamic functions  $\Delta S_d$ ,  $\Delta H_d$ , and  $\Delta G_d$  for the dissolution process of lithium phosphate were calculated based on the Van't Hoff equation. The dissolution is a spontaneous endothermic process. As the concentration of sodium carbonate increases,  $\Delta G_d$  gradually decreases, indicating that higher  $\text{Na}_2\text{CO}_3$  concentrations reduce the energy barrier for  $\text{Li}_3\text{PO}_4$  dissolution, thereby enhancing its solubility by the salting-in effect. Based

on the solubility and supersolubility data of  $\text{Li}_3\text{PO}_4$  in pure water and sodium carbonate solutions, a new process for preparing battery-grade lithium phosphate from salt lake mother liquor by first acidification and then lithium precipitation was developed. The purities of the prepared lithium phosphate reached 99.80%. Raman, FTIR, TG and SEM analyses of  $\text{Li}_3\text{PO}_4$  suggest that the prepared lithium phosphate was pure phase. This study provides fundamental physical chemistry data and theoretical support for the efficient separation and extraction of salt lake mother liquor.

## Data availability

Data are present within the article.

## Author contributions

Huaiyou Wang: methodology, investigation, formal analysis, and writing – original draft; Jia Zhang: investigation, formal analysis, and writing – review & editing; Xu Liu: software and measurement; Haiwen Ge: formal analysis and conceptualization; Zhibo Luo: measurement and supervision; Min Wang: funding acquisition, resources, and supervision. All the authors have read and agreed to the published version of the manuscript.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

This project was supported by the Key R&D and Transformation of Qinghai Province, China (No. 2022-GX-102).

## References

- Y. E. Milian, N. Jamett, C. Cruz, S. Herrera-León and J. Chacana-Olivares, A comprehensive review of emerging technologies for recycling spent lithium-ion batteries, *Sci. Total Environ.*, 2024, **910**, 168543, DOI: [10.1016/j.scitotenv.2023.168543](https://doi.org/10.1016/j.scitotenv.2023.168543).
- B. Swain, Recovery and recycling of lithium: A review, *Sep. Purif. Technol.*, 2017, **172**, 388–403, DOI: [10.1016/j.seppur.2016.08.031](https://doi.org/10.1016/j.seppur.2016.08.031).
- A. Razmjou, M. Asadnia, E. Hosseini, A. Habibnejad Korayem and V. Chen, Design principles of ion selective nanostructured membranes for the extraction of lithium ions, *Nat. Commun.*, 2019, **10**, 5793, DOI: [10.1038/s41467-019-13648-7](https://doi.org/10.1038/s41467-019-13648-7).
- J. Shin, J. M. Jeong, J. B. Lee, H. J. Cho, Y. H. Kim and T. Ryu, Preparation of lithium carbonate from waste lithium solution through precipitation and wet conversion methods, *Hydrometallurgy*, 2022, **210**, 105863, DOI: [10.1016/j.hydromet.2022.105863](https://doi.org/10.1016/j.hydromet.2022.105863).
- J. Xiao, B. Niu and Z. Xu, Highly efficient selective recovery of lithium from spent lithium-ion batteries by thermal



reduction with cheap ammonia reagent, *J. Hazard. Mater.*, 2021, **418**, 126319, DOI: [10.1016/j.hazmat.2021.126319](https://doi.org/10.1016/j.hazmat.2021.126319).

6 S. A. Jose, J. L. Stoll, T. Smith, C. Jackson, T. Dieleman, E. Leath, N. Eastwood and P. L. Menezes, Critical Review of Lithium Recovery Methods: Advancements, Challenges, and Future Directions, *Processes*, 2024, **12**, 2203, DOI: [10.3390/pr12102203](https://doi.org/10.3390/pr12102203).

7 J. H. Han, Z. Nie, Z. H. Fang, Q. Wu, Q. Cao, Y. S. Wang, L. Z. Bu and J. J. Yu, Analysis of existing circumstance of supply and demand on China's lithium resources, *Inorg. Chem.: Indian J.*, 2021, **53**, 61–66, DOI: [10.19964/j.issn.1006-4990.2021-0002](https://doi.org/10.19964/j.issn.1006-4990.2021-0002).

8 T. Ding, M. P. Zheng, X. F. Zhang, Q. Wu and X. Y. Zhang, Development of lithium extraction technology and industrialization in brines of salt lake, *Sci. Technol. Rev.*, 2020, **38**, 16–23.

9 C. Shi, H. Li, B. Liu, Y. Qin and G. Song, Solvent extraction of lithium from aqueous solution using an ammonium ionic liquid, *J. Mol. Liq.*, 2020, **304**, 112756, DOI: [10.1016/j.molliq.2020.112756](https://doi.org/10.1016/j.molliq.2020.112756).

10 T. Kanagasundaram, O. Murphy, M. N. Haji and J. J. Wilson, The recovery and separation of lithium by using solvent extraction methods, *Coord. Chem. Rev.*, 2024, **509**, 215727, DOI: [10.1016/j.ccr.2024.215727](https://doi.org/10.1016/j.ccr.2024.215727).

11 N. Kim, X. Su and C. Kim, Electrochemical lithium recovery system through the simultaneous lithium enrichment via sustainable redox reaction, *Chem. Eng. J.*, 2021, **420**, 127715, DOI: [10.1016/j.cej.2020.127715](https://doi.org/10.1016/j.cej.2020.127715).

12 K. Zhao, J. Li, J. Yuan, X. Yu, Y. Guo, Z. Jiang, M. Li, J. Duo and T. Deng, A novel Co-doped  $H_2TiO_3$  spinning composite for efficient lithium recovery from alkaline lithium precipitation mother liquor, *Chem. Eng. J.*, 2024, **482**, 148989, DOI: [10.1016/j.cej.2024.148989](https://doi.org/10.1016/j.cej.2024.148989).

13 L. Zhang, H. P. Yang, L. Liu and G. F. Ding, Global Technology Trends of Lithium Extraction, *Conservation and Utilization of Mineral Resources*, 2020, **40**, 24–31, DOI: [10.13779/j.cnki.issn1001-0076.2020.05.004](https://doi.org/10.13779/j.cnki.issn1001-0076.2020.05.004).

14 United States Pat.. US9988280, 2018.

15 X. Li, Y. Mo, W. Qing, S. Shao, C. Y. Tang and J. Li, Membrane-based technologies for lithium recovery from water lithium resources: A review, *J. Membr. Sci.*, 2019, **591**, 117317, DOI: [10.1016/j.memsci.2019.117317](https://doi.org/10.1016/j.memsci.2019.117317).

16 Y. Li, M. Wang, Y. J. Zhao, H. Y. Wang, H. J. Yang and Z. H. Zhu, Study on separation of magnesium and lithium from salt lake brine with high magnesium-to-lithium mass ratio by nanofiltration membrane, *CIESC journal*, 2021, **72**, 3130–3139, DOI: [10.11494/0438-1157.20201594](https://doi.org/10.11494/0438-1157.20201594).

17 H. Wen, Z. Liu, J. Xu and J. P. Chen, Nanofiltration membrane for enhancement in lithium recovery from salt-lake brine: A review, *Desalination*, 2024, **591**, 117967, DOI: [10.1016/j.desal.2024.117967](https://doi.org/10.1016/j.desal.2024.117967).

18 Z. J. Cao, M. Gao, Z. Y. Ning and W. Wang, Lithium Resources and Lithium Extraction Technology in Qinghai Salt Lake, *Chem. Eng. Des. Commun.*, 2019, **45**, 190+207.

19 M. X. Li, Master, *Study on Lithium and Boron Extraction Process from Brine of Longmucuo Salt Lake in Tibet*, Jiangxi University Of Science And Technology, 2012.

20 J. Z. Qin, C. X. Li, X. L. Zhang and M. X. Xu, Difficulties and Prospects of the Production of Battery Grade Lithium Carbonate from Lithium Resources in Salt Lake, *J. Salt Sci. Chem. Ind.*, 2022, **51**, 15–18, DOI: [10.16570/j.cnki.issn1673-6850.2022.11.007](https://doi.org/10.16570/j.cnki.issn1673-6850.2022.11.007).

21 N. Linneen, R. Bhave and D. Woerner, Purification of industrial grade lithium chloride for the recovery of high purity battery grade lithium carbonate, *Sep. Purif. Technol.*, 2019, **214**, 168–173, DOI: [10.1016/j.seppur.2018.05.020](https://doi.org/10.1016/j.seppur.2018.05.020).

22 S. Xu, J. Song, Q. Bi, Q. Chen, W.-M. Zhang, Z. Qian, L. Zhang, S. Xu, N. Tang and T. He, Extraction of lithium from Chinese salt-lake brines by membranes: Design and practice, *J. Membr. Sci.*, 2021, **635**, 119441, DOI: [10.1016/j.memsci.2021.119441](https://doi.org/10.1016/j.memsci.2021.119441).

23 L. J. Li, C. L. Liu, X. F. Song and J. G. Yu, Optimization of Crystallization Process for Preparation of Lithium Phosphate in Alkaline Solution System, *J. East China Univ. Sci. Technol.*, 2020, **46**, 598–607, DOI: [10.14135/j.cnki.1006-3080.20190430003](https://doi.org/10.14135/j.cnki.1006-3080.20190430003).

24 C. Zhang, L. Wang, X. Ji and G. Li, Characterization of All Solid State Batteries with LiPON Thin Films Obtained with Different Substrates and RF Sputtering Times, *Mater. Trans.*, 2018, **59**, 1156–1160, DOI: [10.2320/matertrans.M2017419](https://doi.org/10.2320/matertrans.M2017419).

25 L. Li, Doctor, *Study of Solid Electrolyte LiPON and All Solid-State Thin-Film Lithium-ion Batteries' Fabrications and Characteristics*, Lanzhou University, 2018.

26 D. Y. Wang, Master, *Preparation of  $Li_3PO_4$ ,  $Au/Li_3PO_4$  and Isomerization Mechanism of Propylene Oxide Catalyzed by  $Li_3PO_4$* , Nanjing University of Science and Technology, 2018.

27 Y. N. Wang and W. H. Ma, *Presented in Part at the Steam Treatment on Hollow Lithium Phosphate: enhancing Carbon deposition Resistance and Improving Their Catalytic Performances Of propylene Oxide Rearrangement*, Dalian, Liaoning, China, 2016.

28 W. Zhang, L. Chen, Z. Yang and J. Peng, An optical humidity sensor based on  $Li_3PO_4$  hollow nanospheres, *Sens. Actuators, B*, 2011, **155**, 226–231, DOI: [10.1016/j.snb.2010.11.052](https://doi.org/10.1016/j.snb.2010.11.052).

29 A. Gassmann, C. Melzer and H. von Seggern, The  $Li_3PO_4/Al$  electrode: An alternative, efficient cathode for organic light-emitting diodes, *Synth. Met.*, 2012, **161**, 2575–2579, DOI: [10.1016/j.synthmet.2011.08.009](https://doi.org/10.1016/j.synthmet.2011.08.009).

30 C. Xiao and L. Zeng, Thermodynamic study on recovery of lithium using phosphate precipitation method, *Hydrometallurgy*, 2018, **178**, 283–286, DOI: [10.1016/j.hydromet.2018.05.001](https://doi.org/10.1016/j.hydromet.2018.05.001).

31 C. Zhao, Y. Zhang, H. Cao, X. Zheng, T. Van Gerven, Y. Hu and Z. Sun, Dataset of lithium phosphate recovery from a low concentrated lithium-containing solution, *Data Brief*, 2019, **25**, 104044, DOI: [10.1016/j.dib.2019.104044](https://doi.org/10.1016/j.dib.2019.104044).

32 D. F. Liu, J. C. Xiong, W. H. Xu, L. H. He, X. H. Liu and Z. W. Zhao, Lithium selective extraction from lithium-enriched solution by phosphate precipitation, *Chin. J. of Nonferrous Met.*, 2021, **31**, 2541–2550.

33 J. Z. Sun, Preparation of monodisperse micron lithium phosphate, *Ind. Miner. Process.*, 2008, **4–5**, DOI: [10.16283/j.issn1000-0938.2008.04.001](https://doi.org/10.16283/j.issn1000-0938.2008.04.001).



j.cnki.hgkwyjg.2008.08.008, DOI: [10.16283/j.cnki.hgkwyjg.2008.08.008](https://doi.org/10.16283/j.cnki.hgkwyjg.2008.08.008).

34 J. Zhu, X. P. Li, X. H. Liu, Y. Sun and J. Yu, Experimental Study on Purification of Crude Lithium Phosphate from Salt Lake, *Min. Metall. Eng.*, 2023, **43**, 101–105, DOI: [10.3969/j.issn.0253-6099.2023.03.023](https://doi.org/10.3969/j.issn.0253-6099.2023.03.023).

35 RLD, *Handbook of Chemistry and Physics [M]*, 84th edn, CRC Pres, Boca Raton, 2004.

36 S. Wu, M. L. Zhang, C. Zan and H. Zhou, Process of Deep Recovery of Lithium from the Mother Liquor of Lithium Carbonate by Phosphate Precipitation, *J. Tianjin Univ. Sci. Technol.*, 2023, **38**(2), 35–41, DOI: [10.13364/j.issn.1672-6510.20220181](https://doi.org/10.13364/j.issn.1672-6510.20220181).

37 BWDDEHPV, *The NBS Tables of Chemical Thermodynamic Properties :selected Values for Inorganic and C1 and C2 Organic Substances in SI Units [M]*, American Institute of Physics for the National Bureau of Standards, New York, 1982.

38 L. Popović, B. Manoun, D. de Waal, M. K. Nieuwoudt and J. D. Comins, Raman spectroscopic study of phase transitions in  $\text{Li}_3\text{PO}_4$ , *J. Raman Spectrosc.*, 2003, **34**, 77–83, DOI: [10.1002/jrs.954](https://doi.org/10.1002/jrs.954).

39 R. W. Berg, A. V. Nikiforov and N. J. Bjerrum, Vapor pressure and specific electrical conductivity in the  $\text{H}_2\text{O}-\text{LiH}_2\text{PO}_4-\text{LiPO}_3$  system—a novel electrolyte for water electrolysis at elevated temperature, *Ionics*, 2021, **27**, 703–719, DOI: [10.1007/s11581-020-03867-0](https://doi.org/10.1007/s11581-020-03867-0).

