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Nanoporous [Ho₂(CO₂)₇(H₂O)₂]-organic frameworks for excellent catalytic performance on the cycloaddition of CO₂ into epoxides and the deacetalization-Knoevenagel condensation†

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Designing nanoporous lanthanide-based metal–organic frameworks (MOFs) as robust heterogeneous catalysts has received a lot of interest in recent years. Herein, we successfully constructed a novel isomorphous nanoporous MOF {[Ho₂(TDP) (H₂O)₂]·5H₂O·4DMF}_n (named as **NUC-55**, NUC = North University of China) by combining [Ho₂(CO₂)₇(H₂O)₂] (abbreviated as {Ho₂}) clusters with 2,4,6-tri(2,4-dicarboxyphenyl)pyridine (H₆TDP) as structure-oriented multifunctional ligands under acidic solvothermal conditions. **NUC-55** is a holmium(III)-based 3D MOF with a hierarchical porous architecture containing tetragonal microchannels (0.56 nm in diameter) and octagonal nanochannels (1.79 nm in diameter). In **NUC-55**, plenty of Lewis acidic and basic sites, including open Ho³⁺ sites and N_{pyridine} atoms, coexist. Moreover, it is worth mentioning that the void volume (~65%) is significantly higher in **NUC-55** than in most documented 3D lanthanide-based MOFs (Ln-MOFs). Catalytic experiments show that activated **NUC-55** exhibits high catalytic activity in the CO₂–styrene oxide cycloaddition reactions under mild conditions, with a high turnover number of 2475 and a high turnover frequency of 619 h⁻¹. In addition, activated **NUC-55** can remarkably accelerate the deacetalization-Knoevenagel condensation reactions of benzaldehyde dimethyl acetal and malononitrile. Taken together, this work can not only establish an effective self-assembly strategy for fabricating highly porous Ln-MOFs, but also provide new insights into their catalytic mechanism.

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1 Introduction

Metal–organic frameworks (MOFs), which are formed by combining inorganic secondary building units (SBUs) and organic ligands, have a wide range of applications in various fields, such as gas separation-storage,^{1,2} heterogeneous catalysis,^{3,4} fluorescent sensors and probes,^{5,6} and sustained drug-releasing systems.⁷ Nowadays, with the development of MOF chemistry, porous materials with pores of different sizes and shapes have been successfully fabricated by changing their structural units, leading to improved physicochemical

properties and wider applications.^{8–10} Although the specific structure of ligands connecting inorganic components into modularly assembled complex superstructure is most exciting, it remains challenging to establish a predictable synthetic route to assemble with a specific chemical composition and function of the topological structure of target materials. Therefore, a systematic exploration of predictable synthetic pathways and characterization of MOFs is still needed^{11,12}.

In recent years, excess carbon dioxide (CO₂) in the atmosphere has caused great damage to the ecosystem, such as rainstorms, floods, mudslides, and landslides.^{13–16} Therefore, researchers are developing various technologies to capture and convert CO₂ into energy products. Among them, CO₂-epoxide cycloaddition (CEC) has attracted extensive attention, because the generated cyclic carbonate can be widely used in green solvents, lithium batteries, polymers, chemical intermediates, and so on.^{17–20} However, some limitations, such as slow reaction rates and lack of efficient catalysts, restrict the application of CEC in CO₂ capture and fixation. Thus, it is very important to develop recyclable high-performance heterogeneous catalysts under mild conditions.^{21,22} Recently, MOFs have been considered one of the most promising heterogeneous catalysts for the

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CEC owing to the presence of easily accessible Lewis acidic and/or basic sites and highly tunable porosity. Many transition metal-based MOFs have been developed, such as ZIF-67,²³ gea-MOF-1,²⁴ BIT-103,²⁵ USTC-253,²⁶ MIL-101,²⁷ MOF-505,²⁸ and UIO-66.²⁹ However, due to the lack of effective synthetic strategies, there are few reports on lanthanide-based MOFs (Ln-MOFs). In fact, Ln-MOFs are expected to exhibit more excellent catalytic activity because of the strong Lewis acidity of Ln³⁺ ions, as evidenced by Friedel-Crafts,³⁰ Diels-Alder,³¹ aldol-allylation,³² and Michael addition reactions.³³ Thus, in order to construct highly catalytic Ln-MOFs, there is an urgent need to develop effective synthetic strategies for building large-pore and high-surface-area structures to facilitate the interactions between Ln³⁺ ions and organic ligands.^{34–38}

In this study, 2,4,6-tri(2,4-dicarboxyphenyl)pyridine (H₆TDP) was used as a difunctional N-containing heterocyclic ligand to construct Ln-MOFs with the ability to catalyze the CEC under mild conditions (101.32 kPa, low temperatures). Here, we successfully fabricated a dual-channel Ho-based MOFs ($\{[\text{Ho}_2(\text{TDP})(\text{H}_2\text{O})_2] \cdot 5\text{H}_2\text{O} \cdot 4\text{DMF}\}_n$, NUC-55), which possessed excellent physico-chemical properties, such as high porosity, large specific surface areas, solvent-free dual channels, thermal stability, and solvent tolerance. Because of the coexistence of Lewis acidic and basic sites, such as Ho³⁺ ions, N=O and C=O groups, and N_{pyridine} atoms, NUC-55 exhibited very high catalytic activity in the cycloaddition of CO₂ with styrene oxide under solvent-free and mild conditions. Moreover, NUC-55 could remarkably accelerate the deacetalization-Knoevenagel condensation of various substrates.

2 Results and discussion

2.1 Crystal structure

Based on single crystal X-ray diffraction analysis, NUC-55 crystallizes in the tetragonal space group *P*4₃22 and processes a dinuclear hierarchical porous framework formed by combining [Ho₂(CO₂)₇(H₂O)₂] SBUs and 4-connected TDP⁶⁻ moieties. PLATON calculation shows that there are two kinds of edge-sharing channels, including tetragonal micropores (*ca.* 0.56 nm) and octagonal nanopores (*ca.* 1.79 nm), leading to a total potential solvent-accessible void volume of *ca.* 4489.1 Å³ and a porosity of 64.9%, which are larger than those of other 3D Ln-MOFs, such as {Ln(BTB)}_n (49.6%),³⁹ {[Pr₃(ATPT)₂(HATPT)₄(NO₃)(H₂O)₈]}_n (47.9%),⁴⁰ Er₂(TBDC)₃(phen)₂ (38.9%),⁴¹ {Ln(TPO)}_n (48.4–49.0%),⁴² {Ln₃(PTTBA)₂}_n (51.2%),⁴³ and NUC-38 (56.3–56.7%).⁴⁴ It's worth noting that NUC-55 can act as a Lewis acid catalyst since the unoccupied coordination sites associated with each Ho³⁺ ion in octagonal-nanoporous channels can be converted to defective hexa-coordination configuration after releasing water molecules. Furthermore, Lewis basic sites (*i.e.*, uncoordinated N_{pyridine} atoms) are found on the internal surface of microporous channels. Thus, NUC-55 is an ideal candidate for gas separation/reservoir, catalysis, and fluorescence recognition.

In NUC-55, two Ho³⁺ ions are spanned together by three μ₂-η¹:η¹ carboxyl groups from three separate TDP⁶⁻ moieties to form a [Ho₂(CO₂)₇(H₂O)₂] cluster (abbreviated as {Ho₂}, Fig. 1a),

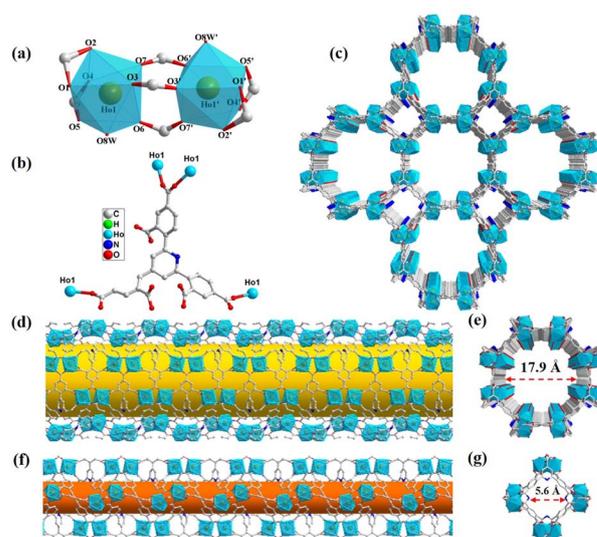


Fig. 1 (a) [Ho₂(CO₂)₇(H₂O)₂] SBUs and its simplified design. (b) The asymmetrical unit of NUC-55. (c) The neutral dual-channel framework of NUC-55. (d) Schematic diagram of the three-dimensional structure of channel I. (e) Side view of channel I. (f) Schematic diagram of the three-dimensional structure of channel II. (g) Side view of channel II.

which is subsequently caught by five TDP⁶⁻ moieties (Fig. 1b) to propagate into a rare 3D Ho-organic framework (Fig. 1c), with a potential solvent-accessible volume of 4489.1 Å³ calculated by PLATON program.^{45,46} In NUC-55, each of eight rows of {Ho₂} clusters with the aid of ambient carbon framework was built into a hydrophilic 1D octagonal channel I (Fig. 1d), with a diameter of 1.79 nm (Fig. 1e), indicating the presence of nearly mesoporous channels in NUC-55. For channel II (Fig. 1f and g), four {Ho₂} clusters are connected through two completely deprotonated ligands to generate a 1D square-shaped channel that is hydrophobic and has an inner diameter of 0.56 nm. Meantime, a 3D ultramicroporous architecture is formed by assembling two types of channels along the *c* axis in an edge-sharing way (Fig. 1c). In addition, the hexapodal TDP⁶⁻ ligand exhibits a self-adaptive behavior in the non-structural deformation due to three dihedral angles (46.4°, 40.0°, and 47.5°) among three branched benzene plane and the central pyridine.

Based on topological analysis,^{47,48} the neutral host framework of NUC-55 can be simplified into a 3D nut-type (4,5)-c network with the Schläfli symbol of (3²·4²·5²)(3⁴·4³·5⁴·6⁹·7) by taking TDP⁶⁻ moieties and {Ho₂} clusters as a linear 5-connected triangle and a 4-connected octahedron, respectively (Fig. S3†).

2.2 Gas adsorption performance

The PXRD patterns of NUC-55 show that the phase purity is very high (Fig. S4†). To test the gas adsorption performance, NUC-55 was activated by exchanging it with methanol three times, with methanol renewed every 24 h, and dried in a vacuum oven at 80 °C for 24 h. The characteristic diffraction peaks in the PXRD pattern of activated NUC-55 were highly consistent with those of the as-synthesized samples (Fig. S5†), indicating that NUC-55



has excellent stability after desolvation. The N₂ adsorption isotherm of NUC-55 at 77 K (Fig. S6†) shows that the permanent porosity increases sharply at low pressures of 0–0.2 bar (Fig. S6†), indicating that it belongs to a typical type I adsorption–desorption isotherm. The overall N₂ uptake of NUC-55 at 1 atm was 305 cm³ g⁻¹. The Brunauer–Emmett–Teller (BET) surface area and Langmuir surface area are 1214 and 1547 m² g⁻¹, respectively. We used non-local density functional theory to fit the N₂ adsorption isotherm at 77 K, and the results showed that the pore sizes of NUC-55 were concentrated at 0.58 nm and 1.72 nm (Fig. S7†), which were consistent with those obtained from the crystal structure.

Furthermore, the CO₂ adsorption curve of NUC-55 was a typical type I adsorption isotherm at 273 K and 298 K, and the maximum volumetric uptakes of CO₂ were 120.3 and 59.3 cm³ g⁻¹, respectively (Fig. S8†). The accurate prediction of CO₂ adsorption at the saturation point was obtained by fitting isotherms measured at 273 and 298 K using the same method. At low coverage, the estimated isosteric heat of adsorption (Q_{st}) for NUC-55 was 23.5 kJ mol⁻¹ (Fig. S9†), indicating that the adsorption of CO₂ by NUC-55 is mainly physical adsorption. Because of the low heat of adsorption, it was easy for NUC-55 to be regenerated through desorption at low temperatures.

2.3 Catalytic cycloaddition of CO₂ to styrene oxide under mild conditions

Activated NUC-55 had many merits as follows. (i) After losing coordinated solvent molecules, Ho³⁺ ions can be used as Lewis acidic active sites to activate small molecules with unbalanced charge distributions. (ii) Uncoordinated pyridine groups can be used as complementary catalytic sites for CO₂ cycloaddition. (iii) Nanoscale pores can facilitate the mass transfer of substrates. The catalytic effect of NUC-55 on the CEC was studied using styrene oxide as the standard substrate. The yield of the products was quantitatively analyzed using GC-MS. The structure of the products was determined using NMR (Fig. S10†). Particularly, the synergistic effect of *n*-Bu₄NBr as a cocatalyst for NUC-55 was also tested. As shown in Table 1, the conversion of styrene oxide was very low in the presence of only a single catalyst in 48 h (entries 1 and 2). In contrast, in the presence of both 0.10 mol% NUC-55 and 5 mol% *n*-Bu₄NBr, the yield increased to 27% at room temperature only in 8 h (entry 3). With the prolongation of reaction time from 16 to 48 h, the yield increased significantly from 45% to 96% (entries 4–8), indicating that *n*-Bu₄NBr could synergistically improve the conversion efficiency of CO₂ since free Br⁻ ions could polarize the β-carbon on ethylene oxide through nucleophilic attack to initiate the ring-opening reaction of epoxides. Next, we observed the effects of reaction time, reaction temperature, and catalyst dosage on the catalytic performance. When the reaction time was extended to 48 h (entries 3–8), the conversion rate reached 96%. Interestingly, as the reaction temperature increased from 35 to 80 °C (entries 9–14), the reaction could be almost completed at a very short reaction time (4 h). At 0.04–0.10 mol% of NUC-55, the dosage had little effect on the catalytic performance (entries 14–17). However, at doses of

NUC-55 below 0.02 mol%, the yield decreased significantly (entry 18–19). Therefore, the optimal reaction conditions for the cycloaddition of CO₂ with epoxides were selected as follows: 80 °C, $P_{CO_2} = 1$ atm, 0.04 mol% NUC-55, and 5 mol% *n*-Bu₄NBr. It is noteworthy that the catalytic efficiency of NUC-55 for styrene oxide is much higher than those of most reported MOF catalysts, such as Zn-2PDC,⁴⁹ Rh-PMOF-1,⁵⁰ and JLU-MOF58 (Zr) (Table S4†).⁵¹

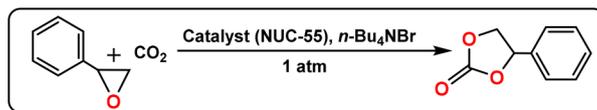
2.4 Cycloaddition of CO₂ with various epoxides

To test the catalytic ability of NUC-55, 10 typical epoxides with different steric hindrance groups and electron-withdrawing (or donating) groups were selected for the cycloaddition reaction with CO₂ under the optimal conditions (Table 2). NUC-55 had high catalytic activity for epoxides with small steric hindrance. The structures of all cycloaddition products were characterized using ¹H NMR and ¹³C NMR (Fig. S11–S19†). However, styrene oxide (entry 9) and 3-phenoxy-1,2-epoxypropane (entry 10) were not easily dispersed in the nanochannels of NUC-55 due to their large steric hindrance, so they could not be effectively activated at the exposed catalytic sites, leading to a slight decrease in the yield. In addition, the molecular sizes of epoxides varied with different substituted groups (Table S3†). In comparison with several MOF-based catalysts reported in the literature (Table S5†), we found that NUC-55 exhibited higher catalytic activity against an epoxide with a bulky substituted group, indicating that NUC-55 might have a wider application.

2.5 Recyclability and stability of NUC-55 in CO₂-epoxide cycloaddition

The recyclability and stability of NUC-55 in cycloaddition reactions were also investigated under optimal conditions. After each reaction, organic matter adsorbed to NUC-55 was removed by immersing NUC-55 in ethyl acetate, followed by vacuum drying. During five catalytic cycles of NUC-55, the yields ranged between 96% and 99% (Fig. S20†), indicating that NUC-55 had excellent catalytic performance, high recyclability and stability. Furthermore, after 5 catalytic cycles, recycled NUC-55 exhibited similar characteristic PXRD peaks to those of original NUC-55, indicating that the structure of NUC-55 remained unchanged after 5 catalytic cycles (Fig. S21†). This can be observed from the SEM images before and after CO₂ cyclization reaction catalyzed by NUC-55 (Fig. S22†). The N₂ adsorption capacity of recycled NUC-55 was 286 cm³ g⁻¹ (Fig. 2), which was slightly less than 302 cm³ g⁻¹ of original NUC-55, possibly due to the blockage of pores and channels by residual styrene carbonate. In addition, the result obtained from ICP-MS showed that only a small amount of Ho³⁺ ions (0.016%) were leached from NUC-55, further confirming the excellent stability of NUC-55. Finally, the heterogeneity of NUC-55 was verified by the thermal filtration experiment (Fig. S23†). Once NUC-55 was removed from the reaction system after 2 h of reaction, the yield did not increase, indicating that the catalytic reaction was terminated, which confirms that NUC-55 possessed pure heterogeneous properties.



Table 1 CO₂-styrene oxide cycloaddition under different conditions^a

Entry	NUC-55 (mol%)	<i>n</i> -Bu ₄ NBr (mol%)	Temp. (°C)	Time (h)	Yield ^b (%)	TON ^c	TOF ^d
1	0	5	RT	48	13	—	—
2	0.10	0	RT	48	9	90	2
3	0.10	5	RT	8	27	271	34
4	0.10	5	RT	16	45	450	28
5	0.10	5	RT	24	64	640	26
6	0.10	5	RT	32	77	770	24
7	0.10	5	RT	40	90	900	221
8	0.10	5	RT	48	96	960	20
9	0.10	5	35	40	98	980	25
10	0.10	5	40	32	99	990	31
11	0.10	5	50	24	99	990	41
12	0.10	5	60	12	99	990	83
13	0.10	5	70	8	99	990	124
14	0.10	5	80	4	100	1000	249
15	0.08	5	80	4	99	1238	311
16	0.06	5	80	4	99	1650	413
17	0.04	5	80	4	99	2475	619
18	0.02	5	80	4	92	4600	1150
19	0	5	80	4	53	—	—

^a Reaction conditions: solvent free, styrene oxide (20 mmol), and CO₂ (1 atm). ^b Product yield was determined using GC-MS with *n*-dodecane as an internal standard. ^c TON (turnover number): mole of product/mole of catalyst. ^d TOF (turnover frequency): TON/time. RT: room temperature.

2.6 Possible mechanism of the catalytic activity of NUC-55 in the cycloaddition of CO₂ into epoxides

By combining previous studies^{41,52} and the structural features of NUC-55, we proposed a possible catalytic mechanism for NUC-55 to promote the cycloaddition of CO₂ into epoxides (Fig. 3). The catalytic process is composed of four steps, including polarization, loop opening, insertion, and loop closing. (i) Ethylene oxide is activated by Ho³⁺ ions after adsorption to its oxygen atoms. (ii) The loop was opened when nucleophilic Br⁻ anions attacked carbon atoms in epoxides with small steric hindrance and generated alkylcarbonate anions.⁵³⁻⁵⁷ (iii) CO₂ molecules adsorbed in the channels of NUC-55 were polarized and inserted into alkylcarbonate anions because of their electronic equilibrium and stability at the ground state.⁵⁸⁻⁶² (iv) Finally, the loop was closed, and one cycle of the reaction ended and the next one began.

2.7 Catalytic performance of NUC-55 for deacetalization-Knoevenagel condensation

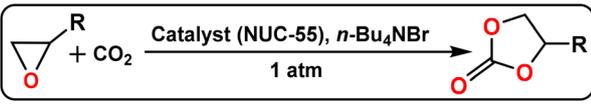
The catalytic performance of NUC-55 for deacetalization-Knoevenagel condensation was tested given the coexistence of Lewis acidic and basic sites, including Ho³⁺ ions, carbonyl oxygen atoms, and pyridinyl groups.^{25,63-66} Using benzaldehyde dimethyl acetal (BDA) and malononitrile as model reactants and DMSO as a solvent, the effects of reaction temperature, reaction time and catalyst dosage on the catalytic performance were investigated. When the reaction was carried out at 40 °C

for 12 h, the yields increased gradually from 3% to 76% as the doses of NUC-55 increased from 0 to 0.30 mol% (entries 2–6, Table 3), indicating that NUC-55 exerted catalytic effects on the deacetalization-Knoevenagel condensation reaction. At 55 °C, the yield reached 96.22% in 12 h, however, the yield could reach 99.33% in 6 h at 70 °C. These results indicated that NUC-55 had high catalytic effects on the tandem deacetalization-Knoevenagel condensation under the optimal condition (0.30 mol% NUC-55, temperature: 70 °C, and reaction time: 6 h). Moreover, compared with other MOFs reported in the literature (Table S7†) NUC-55 had higher catalytic efficiency for the Knoevenagel condensation of BDA.

2.8 Substituted groups and size selectivity in the deacetalization-Knoevenagel condensation reaction

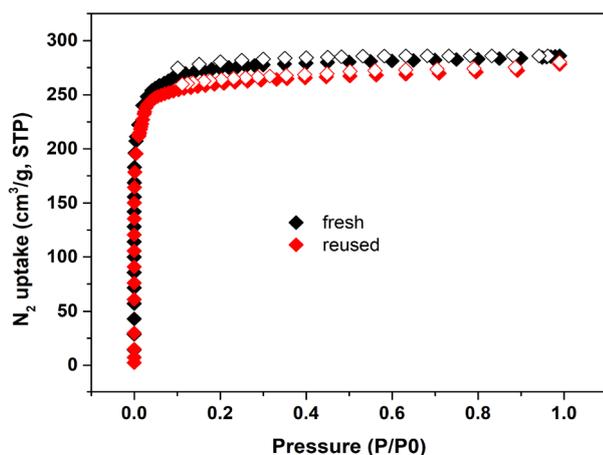
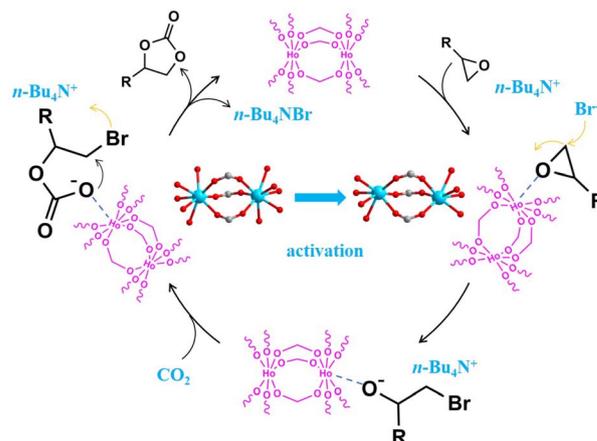
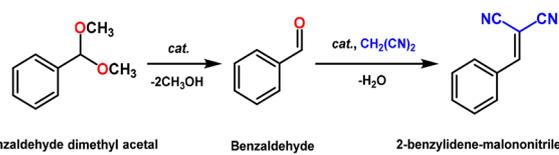
The deacetalization-Knoevenagel condensation reactions were carried out using various BDA derivatives containing different electron acceptors (–NO₂, –F, and –Br) and electron donors (–CH₃, –OCH₃, and –OPh) under the optimal condition. As shown in Table 4, NUC-55 could catalyze the condensation reaction of all BDA derivatives tested. However, it is noted that different substituted groups could affect the reaction to a certain extent, and the yields were slightly higher for BDA derivatives with electron acceptors than those with electron donors. In addition, adverse effects on the electron-donating groups could be verified by several substituted groups, such as –OCH₃ and –OPh (entries 7–10). Moreover, molecular size is also an important



Table 2 Cycloaddition of CO₂ with various epoxides with NUC-55^a


Entry	Epoxides	Selectivity (%)	Yield ^b (%)	TON ^c
1		99	99	2475
2		99	99	2475
3		99	99	2475
4		99	99	2475
5		99	99	2475
6		99	99	2475
7		98	99	2475
8		99	99	2475
9		99	98	2450
10		97	95	2375

^a Reaction conditions: substrates (20 mmol), *n*-Bu₄NBr (5 mol%), NUC-55 (0.04 mol%, based on the {HO₂} cluster), CO₂ (1 atm), 80 °C, and 4 h.
^b *n*-Dodecane was used as the internal standard and determined by GC-MS.
^c TON (turnover number) = mole of product/mole of catalyst.

Fig. 2 N₂ adsorption isotherms of NUC-55 before and after five catalytic cycles.Fig. 3 Proposed mechanism of NUC-55 to catalyze the cycloaddition of CO₂ into epoxides.Table 3 Catalytic Effects of NUC-55 on the deacetalization-Knoevenagel condensation reaction between benzaldehyde dimethyl acetal and malononitrile^a


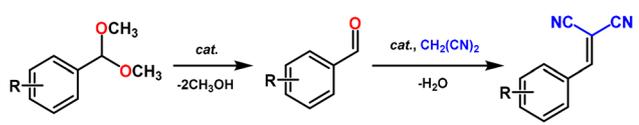
Entry	NUC-55		Temp. (°C)	Yield ^b (%)	TON ^c	TOF ^d
	(mol%)	Time (h)				
1	0	12	40	3	—	—
2	0.10	12	40	36	360	30
3	0.15	12	40	47	313	26
4	0.20	12	40	61	305	25
5	0.25	12	40	71	284	24
6	0.30	12	40	76	253	21
7	0.30	12	45	81	270	23
8	0.30	12	50	89	297	25
9	0.30	12	55	96	320	27
10	0.30	10	60	98	327	33
11	0.30	8	65	99	330	41
12	0.30	6	70	99	330	55

^a Reaction conditions: benzaldehyde dimethyl acetal (10 mmol), malononitrile (20 mmol), DMSO 3 mL. ^b Product yield was determined using GC-MS with *n*-dodecane as an internal standard.
^c TON (turnover number) = mole of product/mole of catalyst. ^d TOF (turnover frequency) = TON/time.

factor influencing the effects of different substituted groups on reaction efficiency. Nevertheless, due to its large channel size, NUC-55 could still promote cycloaddition reactions of BDA derivatives with larger reactive substrates (entries 9 and 10 in Table S6†), with yields reaching 95% and 92%, respectively (Table 4). The results indicate that NUC-55 could effectively catalyze the deacetalization-Knoevenagel condensation reaction.^{67,68} In addition, the structures of deacetalization-Knoevenagel condensation products were confirmed using ¹H NMR and ¹³C NMR (Fig. S24–S33†).



Table 4 Effects of NUC-55 on the deacetalization-Knoevenagel condensation reaction of benzaldehyde dimethyl acetal derivatives containing different substituted groups^a



Entry	Substrates	Selectivity (%)	Yield ^b (%)	TON ^c
1		99	99	330
2		99	99	330
3		98	99	330
4		99	99	330
5		99	99	330
6		99	99	330
7		99	98	327
8		98	97	323
9		98	95	317
10		97	92	307

^a Reaction condition: catalyst NUC-55 (0.30 mol%, based on the active {Ho₂} cluster), malononitrile (20 mmol), benzaldehyde dimethyl acetal derivatives (10 mmol), malononitrile (20 mmol), DMSO 3 mL, 70 °C, 6 h. ^b Product yield was determined using GC-MS with *n*-dodecane as an internal standard. ^c TON (turnover number) = mole of product/mole of catalyst.

2.9 Recyclability and heterogeneity of NUC-55 in the deacetalization-Knoevenagel condensation reaction

Finally, the recyclability and heterogeneity of NUC-55 in the tandem decondensation-Knoevenagel condensation reaction

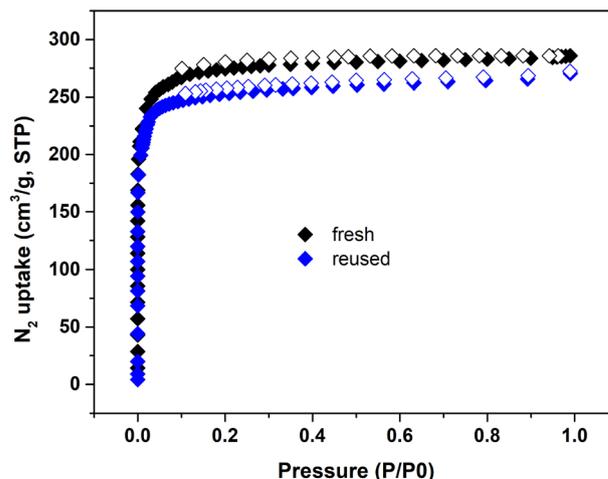


Fig. 4 N₂ adsorption isotherms of NUC-55 before and after five catalytic cycles of deacetalization-Knoevenagel condensation reactions.

were studied using BDA and malononitrile as substrates under the optimal reaction condition. After 6 h of reaction, NUC-55 was collected through filtration, washed with methanol, and reused directly in the next catalytic cycle. As shown in Fig. S34,† the yields remained nearly unchanged during five catalytic cycles, indicating that NUC-55 had excellent recyclability. Moreover, the PXRD patterns of recycled and original NUC-55 showed similar characteristic diffraction peaks (Fig. S35†), indicating that the skeleton of NUC-55 was not damaged and NUC-55 had excellent stability and recyclability. Moreover, this can be observed from the SEM images before and after the catalytic Knoevenagel of NUC-55 (Fig. S36†). The low-temperature N₂ adsorption experiment showed that the amount of N₂ uptake by recycled NUC-55 was almost the same as that of original NUC-55 (Fig. 4), indicating that NUC-55 had high stability. In addition, the heterogeneous properties of NUC-55 were investigated by removing it in the deacetalization-Knoevenagel condensation reaction. As shown in Fig. S37,† when NUC-55 was removed after 2 h of reaction, the reaction was almost terminated. Meanwhile, the result of ICP-MS showed that there was a few amount of Ho³⁺ ions (0.012%) in the filtrate, which explained the slight increase in the yield at 6 h.

2.10 Possible mechanism of the catalytic activity of NUC-55 in the deacetalization-Knoevenagel condensation reaction

Based on the structural features of NUC-55 and relevant literature,⁶⁹⁻⁷¹ we proposed a possible catalytic mechanism of NUC-55 (Fig. 5). First, BDA was hydrolyzed to benzaldehyde with the help of Lewis acidic sites of Ho³⁺ ions. Second, the oxygen atom on the aldehyde group was coordinately activated by the Lewis acidic sites of Ho³⁺ ions. Meanwhile, the methylene group of malononitrile was activated by Lewis basic sites, including the oxygen atom of the carbonyl group and the adjacent N-atom of pyridine. Third, the intermediate 1-phenyl-2,2'-dicyanoethanol was formed. Finally, benzonitrile products were formed



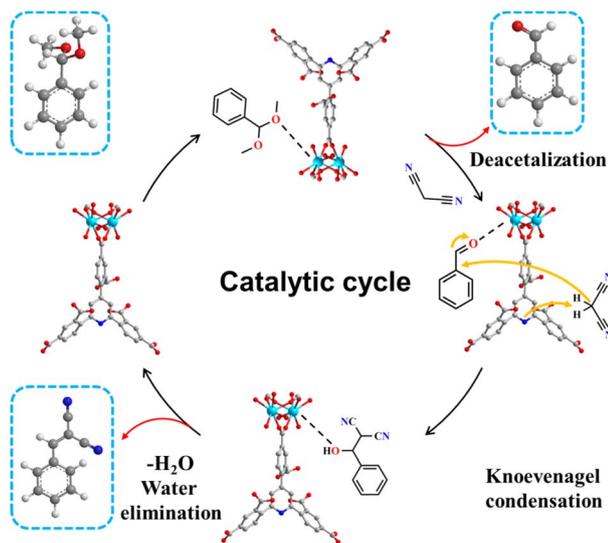


Fig. 5 Proposed catalytic mechanisms of NUC-55 in the deacetalization-Knoevenagel condensation reaction.

after eliminating water molecules, then NUC-55 was released and could be reused in the next cycle.⁷²

3 Conclusions

In conclusion, the solvothermal self-assembly of Ho^{3+} ions and structure-oriented 2,4,6-tri(2,4-dicarboxyphenyl)pyridine (H_6TDP) ligand was used to fabricate NUC-55, a highly robust $\{\text{Ho}_2\}$ -organic frameworks with embedded hierarchical tetragonal-microporous and octagonal-nanoporous channels. NUC-55 possessed a plentiful of coexisted Lewis acid-base sites in its inner wall, including defective Ho^{3+} sites, carbonyl oxygen atoms, and $\text{N}_{\text{pyridine}}$ atoms. NUC-55 showed high catalytic activity for THE cycloaddition of CO_2 to epoxides under mild conditions, with high turnover numbers and turnover frequencies. Moreover, as a heterogeneous catalyst, activated NUC-55 could remarkably accelerate the deacetalization-Knoevenagel condensation of benzaldehyde, dimethylacetal and malononitrile. Thus, this work provides an effective synthetic route for constructing hierarchically porous Ln-organic frameworks, which not only could be used as a robust catalyst, but also could serve as a platform in various fields, such as energy storage, gas adsorption/separation, composite container, etc.

Data availability

All experimental datasets, including crystallographic data, are available in the Cambridge Crystallographic Data Centre (CCDC) under deposition numbers 2125269. These data can be accessed free of charge via [<https://www.ccdc.cam.ac.uk/structures/>]. Additional supporting data are provided in the ESI.†

Conflicts of interest

The authors have no relevant financial or non-financial interests to declare.

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