



Cite this: *RSC Adv.*, 2025, **15**, 18512

Received 7th March 2025

Accepted 9th May 2025

DOI: 10.1039/d5ra01656h

rsc.li/rsc-advances

Chemospecific cobalt-catalyzed hydroboration of CO₂[†]

Andrey Fedulin,^{ab} Lea Luxenberger^a and Axel Jacobi von Wangelin^{ID}*^a

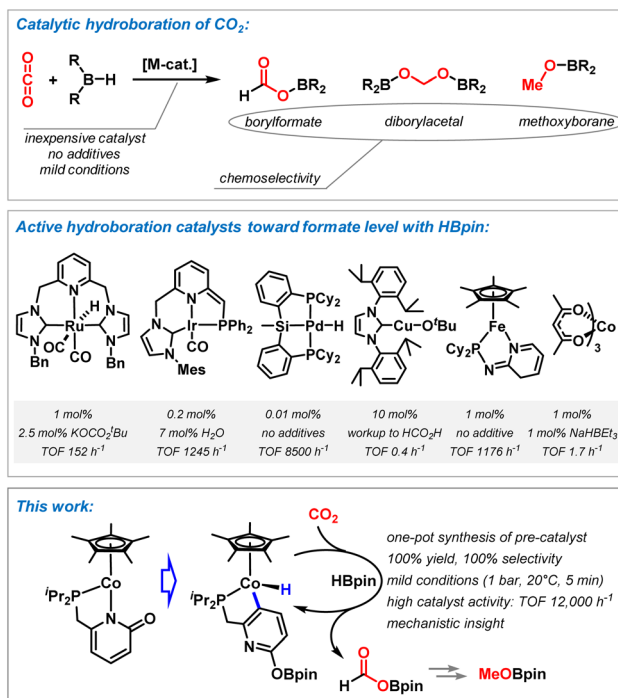
A bifunctional cobalt pyridonate complex effectively catalyzes the hydroboration of CO₂ to the boryl formate at very mild conditions (0.1–1 mol% cat., 1 bar CO₂, r.t., 5 min, 100% yield, TON 1000, TOF 12 000 h^{−1}). At higher temperature, clean conversion to the methoxyborane was achieved (98% yield). Mechanistic studies indicate formation of a ligand-derived cobalt hydride species.

Introduction

Reductive transformations of carbon dioxide are key to all future scenarios in sustainable energy and chemical production technologies.^{1,2} The use of CO₂ as chemical building block is strongly limited by its thermodynamic and kinetic stability, so that efficient catalytic mechanisms at mild conditions constitute a prime area of research. Various synthetic strategies of CO₂ reduction have been developed to formate, formaldehyde, and methanol derivatives and methane. Technical processes mostly utilize hydrogenation reactions under high pressures of H₂ and elevated temperatures, whereas lab-scale reactions often operate with more convenient liquid hydrogen surrogates such as boranes and silanes. Metal-catalyzed hydroborations of CO₂ have been demonstrated to enable facile reductions to borylformates, diborylacetals, and methoxyboranes (Scheme 1, top).³ Major challenges reside in the highly selective formation of a single reduced C1 building block, the use of inexpensive yet highly reactive catalysts and reducing reagents, and the operation under mild conditions with no excess reagents and without waste formation. For example, the two-electron reduction of CO₂ with boranes in the presence of metal catalysts provides O-boryl formates that constitute valuable formyl and formate building blocks. So far, pincer-ligand supported noble metal catalysts (Pd,^{4,5} Ir,⁶ Ru⁷) were among the most active (Scheme 2, middle).⁸ An NHC-copper alkoxide catalyst gave 85% formic acid after hydrolysis;⁹ ligand-coordinated zinc hydride catalysts afforded moderate activity.^{10,11} Recently, an iron catalyst with an anionic PN ligand showed very good activity.¹² Further examples of Mn,^{13,14} Fe,¹⁵ and Ni^{16,17} catalysts are known for CO₂ hydroborations to diborylacetal and borylmethanol. A single example of a cobalt-catalyzed reduction of CO₂ selectively to boryl

formate was reported, with low to moderate yields and little mechanistic insight.¹⁸

Pyridonate ligands entertain a rich coordination chemistry with most transition metals due to their flexible binding modes, hemilability, and potential metal–ligand cooperativity.¹⁹ We reasoned that such multi-functional behavior of metal pyridonates may be effectively exploited for CO₂ hydroborations by sequential BH bond splitting, CO₂ coordination, and hydride transfer onto CO₂. Herein, we report a highly efficient cobalt pyridonate catalyst (TOF >12 000 h^{−1}) that operates under very



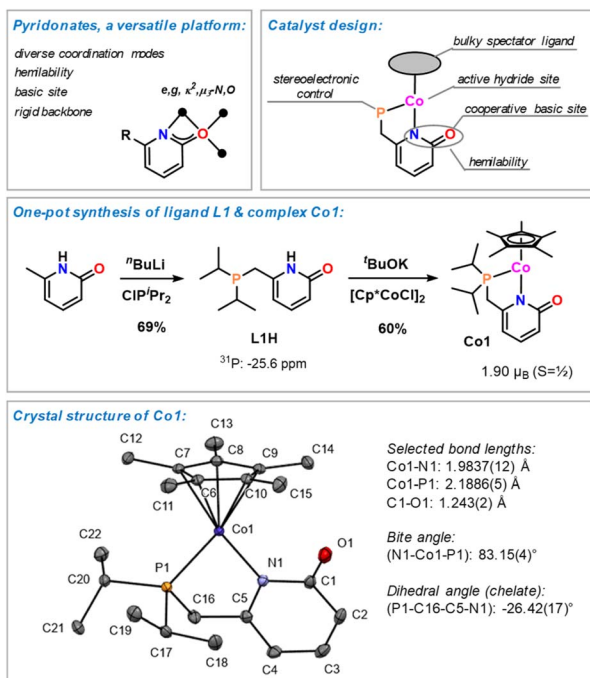
Scheme 1 Top: metal catalyzed hydroboration of carbon dioxide. Middle: selected examples of active metal catalysts and reaction conditions. Bottom: Cobalt pyridonate catalyst for chemospecific CO₂ hydroboration (= this work).

^aDept. of Chemistry, University of Hamburg, Martin Luther King Platz 6, 20146 Hamburg, Germany. E-mail: axel.jacobi@uni-hamburg.de

^bDept. of Chemistry, University of Regensburg, 93040 Regensburg, Germany

[†] Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d5ra01656h>





Scheme 2 Design concept, synthesis, and structure of cobalt pyridonate **Co1**.²³

mild conditions with perfect chemo-specificity toward the formate reduction level. Adjustment of reaction conditions enables further reduction to the methanol level (Scheme 1, bottom).

Results and discussion

Pyridonates confer great versatility of structure and reactivity patterns onto a metal complex by virtue of their multiple binding modes, wide stereoelectronic variation by substituents, hemilability, and redox-activity.¹⁹ Upon employment of polar substrates, ligand–metal cooperativity may be effectively exploited for dual activation modes at the metal and the ligand sites. From the presence of a pyridonate ligand, a strongly coordinating phosphine, and a bulky spectator ligand in the coordination sphere of a cobalt complex, we anticipated the suppression of unwanted aggregation, coordinative saturation with substrate molecules and at the same time favour a cobalt-pyridonate-centered hydroboration.¹⁹ Consequently, we prepared the easily accessible complex $\text{Cp}^*\text{Co}^{\text{II}}(\text{P}\text{ON})$ (**Co1**) with the 6-phosphinomethyl-2-pyridonate ligand (**L1**).²⁰ 6-Methyl-pyridone was converted to the phosphinopyridone **L1H** by double lithiation and substitution with ClP^iPr_2 . Base-mediated reaction of **L1H** with $[\text{Cp}^*\text{CoCl}]_2$ gave the desired half-sandwich complex **Co1** in 60% yield (recrystallized from MeCN). Single crystal structure analysis of the air-sensitive complex documented the chelating $\kappa^2\text{-P,N}$ -coordination to cobalt and a pendant $\text{C}=\text{O}$ double bond moiety with a C1–O1 bond distance of 1.243(2) Å.^{21,22} The magnetic moments (Evans method: 1.89 μ_{B} ; SQUID: T-dependent χ_{MT} at dc field of 0.5 T, 1.90 μ_{B}) are in full agreement with a 17-electron complex and

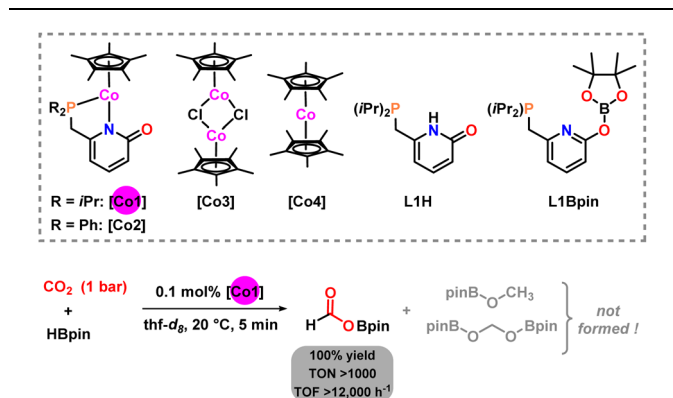
a low-spin $\text{Co}(\text{II})$ center (d^7 , $S = \frac{1}{2}$, see Fig. S41†). **Co1** is highly soluble in THF and toluene, moderately soluble in acetonitrile, sparingly soluble in ether and insoluble in hexane.²³

We initiated our investigations into the catalytic reduction of CO_2 with the hydroboration reaction in the presence of the inexpensive reductant pinacolborane (HBpin) and the pyridonate complex **Co1**.²³ In contrast to the use of strongly hydridic boranes (e.g. $\text{L} \rightarrow \text{BH}_3$, MBH_4 , 9-BBN) that can operate in the presence of simple Lewis basic catalysts²⁴ or under catalyst-free conditions,^{25,26} the use of the less reactive HBpin may enable higher selectivities and controlled access to the individual reduction intermediates (formate, acetal, methoxy levels). Consequently, the solution of catalyst and HBpin in THF- d_8 was degassed and an ambient pressure of CO_2 (1 bar) was applied. The reaction progress was monitored by ^1H NMR spectroscopy (vs. internal mesitylene).

Full conversion was observed with 1 mol% **Co1** after only 5 min exposure to 1 bar CO_2 at room temperature (i.e. in the first recorded ^1H NMR spectrum). Furthermore, the NMR spectrum exhibited perfect chemoselectivity toward the borylformate HCO_2Bpin which had formed as the only product in 100% yield (Table 1, entry 1). The same productivity was afforded with 0.1 mol% catalyst loading after 5 min at 1 bar CO_2 (entry 2). A slightly altered reaction (0.85 mol% **Co1**, 2 bar CO_2) gave full conversion after 2 min (entry 4). Change of solvents to benzene and acetonitrile and a neat reaction gave lower yields (83/21/32%, entry 5), respectively. A wide set of control reactions were performed that documented the crucial role of each component of the modular catalyst **Co1**: The diphenylphosphino derivative of the catalyst (**Co2**) afforded similarly perfect chemoselectivity toward the borylformate, but with only 34% yield (entry 6). The hydroboration did not proceed in the absence of catalyst (entry 7). The pyridone-free complexes $[\text{Cp}^*\text{CoCl}]_2$ and Cp^*_2Co (**Co3** and **Co4**) were no competent catalysts, respectively (entry 8). The use of pyphos (which can be viewed as a truncated deoxo-derivative of **L1**) with **Co3** gave very low conversion and low yield of the borylformate; whereas the *in situ* formed catalyst (from **L1H**, $t\text{-BuOK}$, and **Co3**) afforded good conversion (entries 9 and 10). The ligand alone showed no activity in its neutral form (**L1H**) or by deprotonation with potassium hexamethyldisilazide (**L1K**), respectively (entry 11). Likewise, conversion of **L1H** to the borylated derivative **L1Bpin** – which may be operative under hydroboration conditions – did not afford an active catalyst (entry 12). It is important to note that the catalytic hydroboration of CO_2 operated with low amounts of the pre-catalyst **Co1** (0.1 mol%) under very mild conditions (room temp., 1 bar CO_2) in 5 min reaction time to completion and perfect chemoselectivity toward the borylformate (100% yield).

The active pre-catalyst (**Co1**) operated with a turnover number (TON) of greater than 1000 and a turnover frequency (TOF) of greater than 12 000 h^{-1} (determined after 5 min at full conversion; limited by time of sampling and ^1H NMR analysis, entry 2). To the best of our knowledge, these values document a higher catalytic activity than all literature methods based on main group element and 3d transition metal catalysts. There is a single metal-catalyzed hydroboration of CO_2 with HBpin that



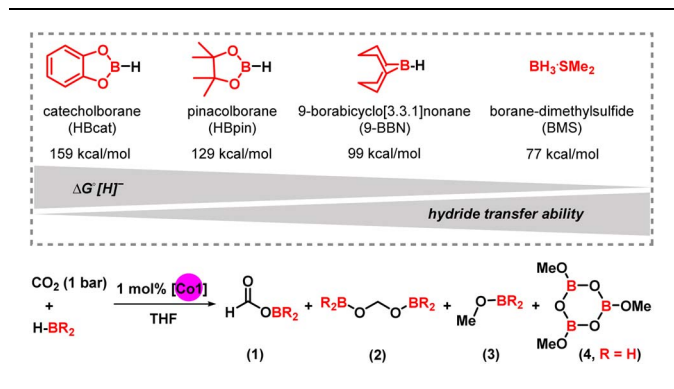
Table 1 Selected optimizations of the cobalt-catalyzed hydroboration of CO₂^a

Entry	Change from conditions above ^a	Yield [%]
1	1 mol% [Co1]	100
2	None	100
3	0.05 mol% [Co1]	33
4	0.85 mol% [Co1], 2 bar CO ₂ , 2 min	100
5	C ₆ D ₆ , CD ₃ CN, neat; each with 1 mol% [Co1]	83/21/32
6	1 mol% [Co2]	34
7	without catalyst	0
8	1 mol% [Co3] or [Co4]	0/0
9	2 mol% [pyphos] + [Co3]	8
10 ^b	2 mol% [L1H] + ^t BuOK + [Co3]	56
11	2 mol% [L1H] or [L1H] + Khmds	0/0
12 ^c	2 mol% L1-Bpin	0

^a Reaction conditions: an NMR tube was charged in argon-filled glovebox with HBpin (0.27 mmol), mesitylene (as internal NMR reference), the catalyst (solid or stock solution), and 0.6 mL solvent. The mixture was degassed by two cycles of freeze-pump-thaw and backfilled with 1 bar CO₂. The NMR tube was sealed, shaken vigorously, and after 5 min the gas was released and the reaction subjected to ¹H and ¹¹B NMR analysis. ¹H NMR yields were determined by integration vs. internal mesitylene as an average of two runs. ^b Equimolar amounts of L1H, ^tBuOK and Co3 were pre-mixed in a vial. ^c A stock solution of L1Bpin in THF-d₈ was prepared by heating equimolar amounts of L1H and HBpin at 60 °C for 16 h.

exhibited higher activity utilizing a Pd-silyl pincer complex (TON 37 200 after 1 day; 63 500 after 5 day; TOF 8500 h⁻¹).⁵

The choice of using pinacolborane (HBpin) as reductant for highly selective hydroboration of CO₂ was evident from a brief screening of alternative boranes (Table 2). The observed trend of borane reactivities can be partially interpreted with their thermodynamic hydricities ($\Delta G^\circ(\text{H}^-)$).²⁷ The least hydridic borane HBcat ($\Delta G^\circ(\text{H}^-) = 159\text{ kcal mol}^{-1}$) afforded very low conversions and only minor amount of reduction product (8% methoxycatecholborane). The absence of the borylformate may be a consequence of lower steric bulkiness of the catechol vs. the pinacol substituent. 9-BBN (more hydridic with a $\Delta G^\circ(\text{H}^-) = 99\text{ kcal mol}^{-1}$) gave full conversion in <1 h to a mixture of the diborylacetate (46%) and the methoxy borane (26%). Again, no formate intermediate was observed. The most hydridic borane in this series, BH₃·SMe₂ ($\Delta G^\circ(\text{H}^-) = 77\text{ kcal mol}^{-1}$), led to rapid

Table 2 Selection of boranes and chemoselectivities of CO₂ reduction.^a

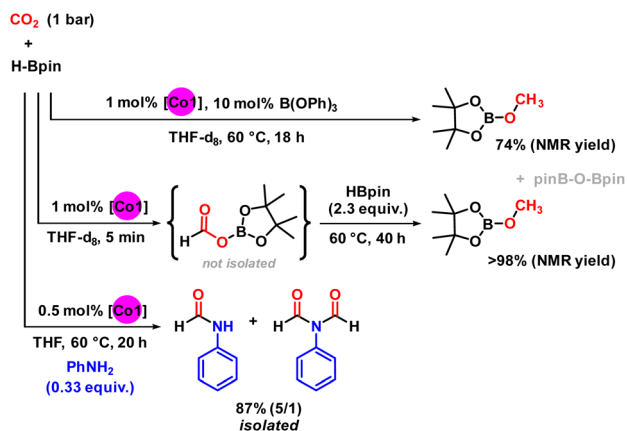
Entry	Borane	T [°C]	t [h]	1/2/3/4 [%] ^b
1	HBcat	60	18	0/0/9/-
2	HBpin	20	0.1	100/0/0/-
3	9-BBN	20	24	0/46/26/-
4	BH ₃ ·SMe ₂	60	18	-/-/-/0

^a Reaction conditions: an NMR tube was charged with HBpin (0.27 mmol), mesitylene (as internal NMR reference), Co1 (2.7 μmol, 1 mol%), and 0.6 mL THF-d₈. The mixture was degassed by two cycles of freeze-pump-thaw and backfilled with 1 bar CO₂ for 5 min. The NMR tube was sealed, shaken vigorously, (heated in an oil bath if required). Reaction progress was monitored by ¹H and ¹¹B NMR. ¹H NMR yields were determined by integration vs. internal mesitylene as an average of two runs. ^b 1, 2 and 3 equiv. of HBR₂ are required for the formations of products 1, 2 and 3, respectively. 1 equiv. BH₃·SMe₂ and 3 equiv. CO₂ are required to produce 4. Equimolar amounts of diboryloxide and methoxyborane 3 are formed.

catalyst decomposition to a dark precipitate without any detectable formation of CO₂ reduction products.

While the hydroboration of CO₂ with HBpin and catalytic Co1 under standard conditions (1 bar CO₂, THF, 20 °C, 5 min) cleanly afforded the boryl formate in perfect yield and selectivity, change of the reaction conditions enabled onward reduction to the methanol level. Addition of the Lewis acid B(OPh)₃ as co-catalyst (10 mol%) fully inhibited the hydroboration at 20 °C (no conversion of HBpin after 1 h), but afforded the corresponding methoxyborane as single reduction product in 74% yield after 16 h at 60 °C (Scheme 3, top).^{5,14} More conveniently, full conversion to the methoxyborane could be easily achieved when adding excess amounts of HBpin to the crude boryl formate and further reaction at 60 °C (Scheme 3, middle, and Fig. S1 and S2†). The synthetic utility of the borylformate product was explored by addition of aniline to the hydroboration reaction (Scheme 3, bottom). Reaction of aniline and 3 equiv. HBpin in the presence of Co1 (1 mol%) in THF under 1 bar CO₂ at room temperature resulted in the formation of a mixture of HCO₂Bpin and the undesired dehydro-coupling product PhNHBpin (2/1). The same reaction with 0.5 mol% Co1 at 60 °C gave clean formylation of the aniline to afford *N*-formanilide and *N,N*-diformyl aniline in overall 87% isolated yield (5/1; see Fig. S8† for details).¹²

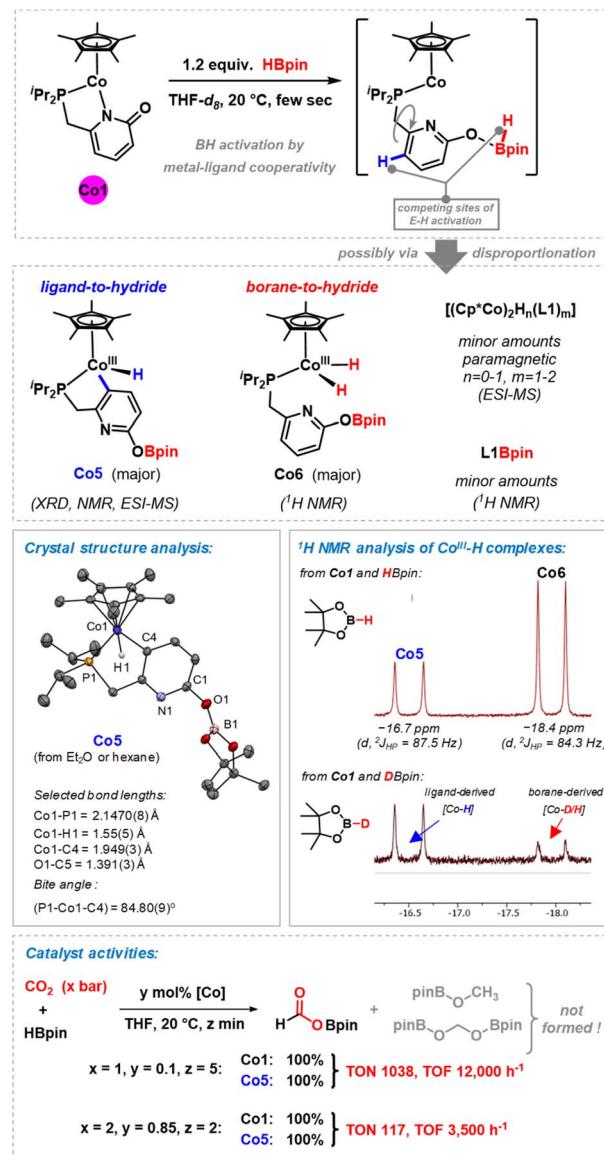




Scheme 3 Variations of the general protocol. Top: selective hydroboration to the boryl-formate and methoxyborane. Bottom: *in situ* formylation to give *N*-phenyl formamides.

Mechanistic studies

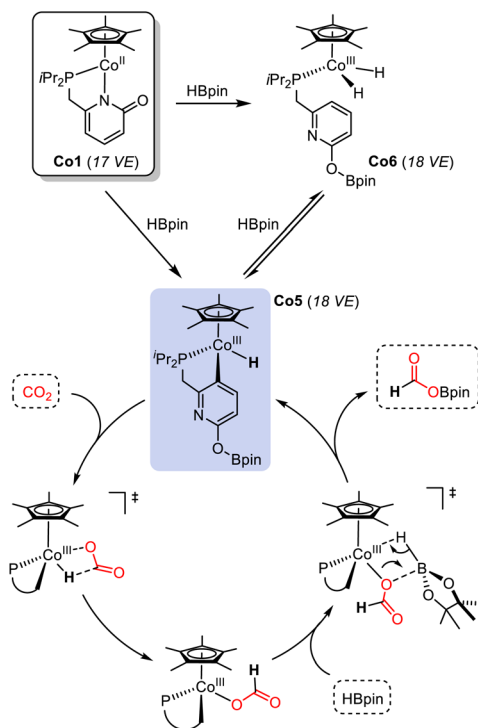
A set of preparative and spectroscopic experiments were performed in order to gain insight into the nature of the catalytically active species (Scheme 4). Similar reaction conditions and catalysts have been applied to the regioselective hydroboration of pyridines, which may operate by a related mechanism.²³ Stoichiometric reaction of the cobalt–pyridonate complex **Co1** with pinacolborane (1.2 equiv.) in THF-d_8 resulted in immediate color change from orange to dark green or even blackish. We postulate the formation of a transient labile cobalt(II) hydride complexes which underwent rapid disproportionation,²⁸ possibly *via* dinuclear hydride or pyridonate-bridged species.²⁸ ^1H NMR monitoring indicated the formation of diamagnetic (and minor amounts of paramagnetic) species: (i) very minor amounts of the borylated ligand **L1Bpin** were observed by ^1H NMR. (ii) We postulate the formation of dinuclear (or higher) complexes of the formula $[(\text{Cp}^*\text{Co})_2\text{H}_n(\text{L1})_m]$ as minor paramagnetic species with Co(I), Co(II), or mixed valence states. ESI-MS spectra showed $m/z = 612.2216$ ($n = 0, m = 1$) and 837.3498 ($n = 1, m = 2$). Similar hydride- and pyridonate-bridged dinuclear complexes were prepared from $[\text{Cp}^*\text{CoCl}]_2/\text{LiAlH}_4$ (ref. 29) and from $[\text{Cp}^*\text{IrCl}(2\text{-pyridonate})]^{30}$ by hydrogen transfer, respectively. (iii) Most interestingly, two distinct cobalt hydride complexes were formed, which differ in the origin of their hydride ligands (Scheme 4, middle and Fig. S13 and S14[†]). The diamagnetic monohydridocobalt(III) complex **Co5** (-16.5 ppm; $^2J_{\text{P,H}} = 87.3$ Hz) formed by C–H activation at the pyridine ligand moiety. The dihydrido phosphine cobalt(III) complex **Co6** exhibited a more upfield hydride signal (-18.0 ppm; $^2J_{\text{P,H}} = 84.4$ Hz), which is in agreement with closely related $\text{CpCo}^{\text{III}}\text{H}(\text{PR}_3)$ complexes.^{31–33} The origin of the hydride ligands was furnished by the same reaction of **Co1**, but with DBpin instead of HBpin: Similar intensity of the ^1H resonance of **Co5** (as from HBpin) was observed but very low intensity of the borane-derived hydride signal of **Co6** (Fig. S15[†]). Complex **Co6** also formed by addition of HBpin to a solution of **Co5**, so that **Co6** can be viewed as an overreduction product of **Co5**. A dinuclear derivative of **Co6** could be isolated by crystallization.²³



Scheme 4 Selected mechanistic studies. Top: synthesis and characterization of potential catalyst intermediates, including the isolation and characterization of key cobalt hydride complexes **Co5** and **Co6**.²³ Bottom: equal catalytic activities of precatalyst **Co1** and proposed catalyst species **Co5** under different conditions.

Complex **Co5** was crystallized from diethylether and *n*-hexane (20% crystal yield). The single crystal structure analysis showed a three-legged piano stool complex with terminal hydride and a 5-membered metallocycle with PNC-coordination of the pyridonate ligand. The 2-oxo position of the ligand bears a Bpin substituent, which supports the notion of a dual activation mode in this hydroboration reaction by the cobalt center and the basic oxygen site.²³

In the catalytic hydroboration of CO_2 , complex **Co5** was equally active as pre-catalyst **Co1** under the standard conditions, which may indicate the role of **Co5** as catalytically active species (Scheme 4, bottom). Full conversion of HBpin to the borylformate was observed after reaction at room temperature



Scheme 5 Postulated reaction mechanism via Co5 as key catalytic intermediate. A closely related mechanism was postulated for a cobalt-catalyzed hydroboration of pyridines.²³

for 5 min with 1 mol% Co5. Co5 is stable toward higher excess amounts of borane. Addition of up to 4 equiv. HBpin did not result in any detectable shift or disappearance of the ¹H NMR resonance at −16.5 ppm. Further evidence that Co5 is a competent catalyst was derived from its instantaneous reaction with CO₂ which was followed by ¹H NMR spectroscopy. The ¹H signal of the Co5 hydride ligand at −16.5 ppm disappeared instantaneously while the −18.0 ppm hydride resonance of Co6 remained unchanged for several hours before it only slowly disappeared overnight. We therefore postulate a reaction mechanism that involves rapid conversion of Co1 into Co5. This key catalytic cobalt hydride intermediate directly transfers a hydride onto CO₂. Borane activation may proceed via a σ-bond metathesis event that releases the formyl borate HCO₂Bpin and regenerates the active catalyst Co5 (Scheme 5).²³

Conclusions

The hydroboration of CO₂ was realized under very mild conditions (1 bar CO₂, 20 °C, 5 min) with low catalyst loading (0.1 mol%) of a simple cobalt catalyst. Perfect chemoselectivity toward the borylformate (100% yield) and very high catalyst activity (TON 1000, TOF 12 000 h^{−1}) were observed. Mechanistic studies revealed two distinct modes of hydride complex formation under the reaction conditions: ligand-centered CH activation to the monohydride cobalt(III) complex Co5 and borane-induced formation of the dihydrido-cobalt(III) complex Co6. The modular composition of Co1 and identification of Co5 as active catalyst will prompt further studies into the generation

of active metal hydride complexes. This concept has already proven successful in a new pyridine reduction method.²³ Further applications to the wide space of hydrofunctionalization and hydrogenation reactions are easily foreseen.

Data availability

The data supporting this article have been included as part of the ESI.†

Author contributions

A. F.: conceptualization, investigation, validation, writing draft. L. L.: investigation, validation. A. J. v. W.: conceptualization, draft review and edit, visualization, supervision, administration, funding acquisition.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by the European Research Council (ERC-CoG 683150, A. J. v. W.). We thank Dr Dieter Schaarschmidt (University of Hamburg) for excellent technical assistance.

Notes and references

- Q. Liu, L. Wu, R. Jackstell and M. Beller, *Nat. Commun.*, 2015, **6**, 5933.
- J. Artz, T. E. Müller, K. Thenert, J. Kleinekorte, R. Meys, A. Sternberg, A. Bardow and W. Leitner, *Chem. Rev.*, 2018, **118**, 434.
- S. Kostera, M. Peruzzini and L. Gonsalvi, *Catalysts*, 2021, **11**, 1.
- Y. C. A. Sokolovicz, O. Nieto Faza, D. Specklin, B. Jacques, C. S. López, J. H. Z. Dos Santos, H. S. Schrekker and S. Dagorne, *Catal. Sci. Technol.*, 2020, **10**, 2407.
- M. R. Espinosa, D. J. Charboneau, A. Garcia De Oliveira and N. Hazari, *ACS Catal.*, 2019, **9**, 301.
- P. Sánchez, M. Hernández-Juárez, N. Rendón, J. López-Serrano, E. Álvarez, M. Paneque and A. Suárez, *Dalton Trans.*, 2018, **47**, 16766.
- C. K. Ng, J. Wu, T. S. A. Hor and H. K. Luo, *Chem. Commun.*, 2016, **52**, 11842.
- For main group element catalyzed hydroborations to formates, see for example: (a) L. J. Morris, P. Mahawar and J. Okuda, *J. Org. Chem.*, 2023, **88**, 5090; (b) B.-X. Leong, J. Lee, Y. Li, M.-C. Yang, C.-K. Siu, M.-D. Su and C.-W. So, *J. Am. Chem. Soc.*, 2019, **141**, 17629.
- R. Shintani and K. Nozaki, *Organometallics*, 2013, **32**, 2459.
- X. Wang, K. Chang and X. Xu, *Dalton Trans.*, 2020, **49**, 7324.
- R. Chamenahalli, R. M. Bhargav, K. N. McCabe, A. P. Andrews, F. Ritter, J. Okuda, L. Maron and A. Venugopal, *Chem.-Eur. J.*, 2021, **27**, 7391.



- 12 H. Gao, J. Jia, C.-H. Tung and W. Wang, *Organometallics*, 2023, **42**, 944.
- 13 C. Erken, A. Kaithal, S. Sen, T. Weyhermüller, M. Hölscher, C. Werlé and W. Leitner, *Nat. Commun.*, 2018, **9**, 1.
- 14 S. Kostera, M. Peruzzini, K. Kirchner and L. Gonsalvi, *ChemCatChem*, 2020, **12**, 4625.
- 15 G. Jin, C. G. Werncke, Y. Escudié, S. Sabo-Etienne and S. Bontemps, *J. Am. Chem. Soc.*, 2015, **137**, 9563.
- 16 S. Chakraborty, J. Zhang, J. A. Krause and H. Guan, *J. Am. Chem. Soc.*, 2010, **132**, 8872.
- 17 L. J. Murphy, H. Hollenhorst, R. McDonald, M. Ferguson, M. D. Lumsden and L. Turculet, *Organometallics*, 2017, **36**, 3709.
- 18 S. R. Tamang and M. Findlater, *Dalton Trans.*, 2018, **47**, 8199.
- 19 A. Fedulin and A. Jacobi von Wangelin, *Catal. Sci. Technol.*, 2024, **14**, 26.
- 20 X. Zhuang, J. Y. Chen, Z. Yang, M. Jia, C. Wu, R. Z. Liao, C. H. Tung and W. Wang, *Organometallics*, 2019, **38**, 3752.
- 21 G. S. Nichol and W. Clegg, *Acta Crystallogr., Sect. C: Cryst. Struct. Commun.*, 2005, **61**, 383.
- 22 F. Jiang, M. Achard, T. Roisnel, V. Dorcet and C. Bruneau, *Eur. J. Inorg. Chem.*, 2015, 4312.
- 23 For the synthesis, characterization, and a related catalytic application of this complex, see here: F. Hoeeg, L. Luxenberger, A. Fedulin and A. Jacobi von Wangelin, *Chem. Sci.*, 2024, **15**, 5201.
- 24 C. Das Neves Gomes, E. Blondiaux, P. Thuéry and T. Cantat, *Chem.–Eur. J.*, 2014, **20**, 7098.
- 25 I. Knopf and C. C. Cummins, *Organometallics*, 2015, **34**, 1601.
- 26 K. Fujiwara, S. Yasuda and T. Mizuta, *Organometallics*, 2014, **33**, 6692.
- 27 S. Ilic, A. Alherz, C. B. Musgrave and K. D. Glusac, *Chem. Soc. Rev.*, 2018, **47**, 2809.
- 28 U. Koelle, B. Fuss, M. Belting and E. Raabe, *Organometallics*, 1986, **5**, 980.
- 29 J. L. Kersten, A. L. Rheingold, K. H. Theopold, C. P. Casey, R. A. Widenhoefer and C. E. C. A. Hop, *Angew Chem. Int. Ed. Engl.*, 1992, **31**, 1341.
- 30 A. M. Royer, T. B. Rauchfuss and D. L. Gray, *Organometallics*, 2010, **29**, 6763.
- 31 J. A. Zurakowski, B. J. H. Austen, M. C. Dufour, M. Bhattacharyya, D. M. Spasyuk and M. W. Drover, *Dalton Trans.*, 2021, **50**, 12440.
- 32 D. M. Heinekey, A. Liegeois and M. van Roon, *J. Am. Chem. Soc.*, 1994, **116**, 8388.
- 33 U. Koelle and S. Paul, *Inorg. Chem.*, 1986, **25**, 2689.

