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Effect of Eu³⁺ ion concentration on optical and magnetic properties of oriented Gd₂O₃/CTAB nanoparticles as multifunctional optical-magnetic probes in biomedicine

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The Gd₂O₃:Eu³⁺ nanoparticles were synthesized using a multi-step chemical method with urea as a reactant to control the ratio of different Eu³⁺ activation centers: 2, 4, 6, 8, 10, and 12 mol% combined with CTAB surfactant to improve surface quality. The study aimed to determine the optimal concentration of Eu³⁺ in the presence of CTAB to increase biocompatibility and achieve the best fluorescence. The structure, surface morphology, optical properties, and magnetic properties of the materials were analyzed through FSEM, XRD, HRTEM, XPS, UV-vis, fluorescence, fluorescence excitation, time-resolved fluorescence, vibrating sample magnetometry (VSM), and magnetic heating measurements. The obtained material had a diameter of 180–280 nm, and it emitted red light with characteristic shifts from ⁵D₀ to ⁷F_J ($J = 0-4$). The strongest emission peak occurred at the transition of ⁵D₀ to ⁷F₂, corresponding to a wavelength of 611 nm. The crystal is in the cubic phase. The highest lifetime of the samples is 2.1 ms, and the highest calculated quantum efficiency is 91% for the Gd₂O₃:8% Eu³⁺ sample. The M–H hysteresis curve revealed that the highest magnetic field obtained was 1.83 emu g⁻¹. Experimental induction heating of samples reached temperatures in the range of 44–49 °C, which is an appropriate temperature range for destroying cancer cells without affecting healthy cells. These findings demonstrate that the material has great potential in cancer diagnosis and treatment.

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1. Introduction

Rare earth elements (RE) have garnered significant interest in both basic and applied research over recent decades due to their unique physical and chemical properties.^{1–4} This interest is reflected in the growing number of applications, as RE elements have become indispensable for important technologies.^{5–7} Nanostructured materials containing RE elements, either as the main component or dopant phase, have opened up new avenues

for various biomedical applications, including bioimaging, biosensors, targeted drug delivery, and other therapies.^{8–11}

The fluorescence properties of inorganic fluorescent materials depend closely on their substrate, dopant ions, size, and morphology.¹² Researchers have thus explored the synthesis of inorganic luminescent materials with different substrates, dopant ions, and shapes.^{12,13} The choice of substrate significantly impacts the luminescent properties of materials, making substrate exploration a focal point for researchers.

Gd₂O₃ (gadolinium sesquioxide) is a well-known material widely used in cathode ray tubes, phosphors, bioimaging, and biosensors.^{1–3} Among various rare earth oxide materials, Gd₂O₃ offers several advantages, including physical, chemical, and thermal stability, low phonon energy, high refractive index, high dielectric constant, paramagnetism, and high density.^{10,14,15} When doped with europium (Eu³⁺), Gd₂O₃ becomes a red luminescent material with potential applications in fluorescent lamps, white light-emitting diodes, plasma display panels, flat screens, cathode ray tubes, MRI contrast agents, biosensors, and bioimaging.¹⁶ The long excitation level of Gd³⁺ ions produces emission lines in the UV region, and luminescence changes occur when other rare

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earth ions are added.¹⁷ Importantly, the ionic radius of Eu³⁺ matches that of Gd³⁺, allowing easy incorporation into the Gd₂O₃ substrate without distorting the crystal structure.^{11,18}

On the other hand, gadolinium (Gd) and Gd³⁺ ions possess seven unpaired electrons, resulting in strong superparamagnetic properties.¹⁹ Organic hybrid compounds containing Gd³⁺ ions are commonly used as contrast agents in magnetic resonance imaging (MRI).²⁰⁻²² In modern medicine, there is a growing emphasis on combining diagnosis and treatment.²³ Therapies such as targeted drug delivery, chemotherapy, thermotherapy, or radiotherapy are increasingly integrated with imaging diagnostics, both *in vitro* and *in vivo*. Various imaging methods, including MRI and fluorescent labeling, can be combined with treatment agents like thermotherapy, chemotherapy, or drug delivery. This combined imaging approach allows for better control, information gathering, and understanding of process dynamics, ultimately enhancing treatment effectiveness.

In our study, we focus on evaluating the luminescent properties of Gd₂O₃ material at different Eu³⁺ doping concentrations, while also investigating the magnetic properties and thermotherapy potential of Gd₂O₃:Eu³⁺. Luminescent Gd₂O₃:Eu³⁺ materials have been synthesized using various methods, such as sol-gel, hydrothermal, co-precipitation, and multi-step chemistry.^{11,24-26} Among these methods, the multi-step chemical synthesis stands out as an easy and cost-effective approach to produce uniform-sized nanoparticles in a shorter time and at lower temperatures. Therefore, we will use this method to synthesize Gd₂O₃:Eu³⁺ material combined with CTAB surfactant to enhance surface quality. Using CTAB not only improves the dispersion of nanoparticles but also enhances their stability in biological environments. This research can lead to improved biocompatibility, making the nanoparticles more effective for *in vivo* applications. Investigating the relationship between optical properties and hyperthermia allows for the design of nanoparticles with tunable heating profiles. This can optimize treatment protocols, providing controlled thermal doses to target tissues while minimizing damage to surrounding healthy cells. The novelty of this research lies in the intersection of improved optical properties, enhanced stability through CTAB, and the potential for real-time imaging during hyperthermia treatments.

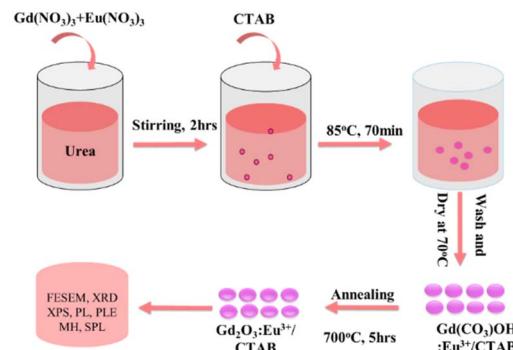
2. Materials and methods

2.1. Materials

Chemicals used in material synthesis Gd₂O₃:Eu³⁺: Gd(NO₃)₃·6H₂O from Sigma-Aldrich, 99.9%, Eu(NO₃)₃·5H₂O from Sigma-Aldrich 99.9%, urea (CO(NH₂)₂) from Sigma-Aldrich, hexadecyltrimethylammonium bromide (CTAB) from Sigma-Aldrich.

2.2. Synthesis process

Prepare “Solution 1” by mixing 0.05 M Eu(NO₃)₃·5H₂O and 0.05 M Gd(NO₃)₃·6H₂O in a 100 mL beaker, adjusting the Eu³⁺:Gd³⁺ ratio to 2%, 4%, 6%, 8%, 10%, or 12%.



Scheme 1 Schematic diagram of preparation process of the Gd₂O₃:Eu³⁺/CTAB.

In a 500 mL flask, prepare a 0.5 M urea solution and gradually add “Solution 1,” stirring for 2 hours. Add 0.01 M CTAB and heat the mixture to 85 °C for 70 minutes to produce a white precipitate of Gd(OH)CO₃·H₂O.

Centrifuge the precipitate with deionized water and ethanol, then dry at 70 °C for 24 hours. Finally, calcine the material at 700 °C for 5 hours (Scheme 1).

2.3. Characterization techniques

The crystal structure of the nanoparticles was analyzed using X-ray diffraction (XRD) with a Bruker D8 Advance instrument with CuK α radiation ($\lambda = 0.154$ nm) at fine steps of 0.02°. The morphology of the synthesized material was analyzed using Field Emission Scanning Electron Microscopy (FESEM) on a Hitachi S-4800 machine. High-resolution transmission electron microscopy (HR-TEM) spectra were measured by a JEM2100 system (Jeol, Japan). X-ray photoelectron spectroscopy (XPS, Nexsa G2) analyzed the chemical bonding configurations. The bandgap energy was estimated using the UV-vis absorption spectrum obtained by JASCO V-750 spectrophotometer. Photoluminescence (PL) spectra were recorded with a Nanolog spectrophotometer (Horiba Jobin Yvon) excited by a 450 W xenon lamp. The saturation magnetization (MS) values were characterized by a vibrating sample magnetometer (VSM, MicroSense EZ9). The inductive heating experiment was conducted using an RDO-HFI device with an output power of 5 kW. All measurements were done in ambient air.

3. Results and discussion

3.1. Characterization of the Gd₂O₃ nanoparticles

3.1.1. Morphology. When Eu³⁺ is doped at varying molar ratios of Eu³⁺/Gd³⁺ (from 2% to 12%) in combination with CTAB, uniform spherical nanoparticles are produced. At a low Eu³⁺ doping concentration of 2% (Fig. 1a), the resulting nanospheres have a diameter of 180–200 nm. Increasing the Eu³⁺ concentration to 4% (Fig. 1b) results in nano-spheres measuring 200–220 nm (Fig. 1c). At a doping level of 6% (Fig. 1c), the nano-spheres expand to diameters of 210–230 nm. With 8% Eu³⁺, the spheres measure 220–230 nm (Fig. 1d), while a 10% doping level produces spheres with diameters of 230–



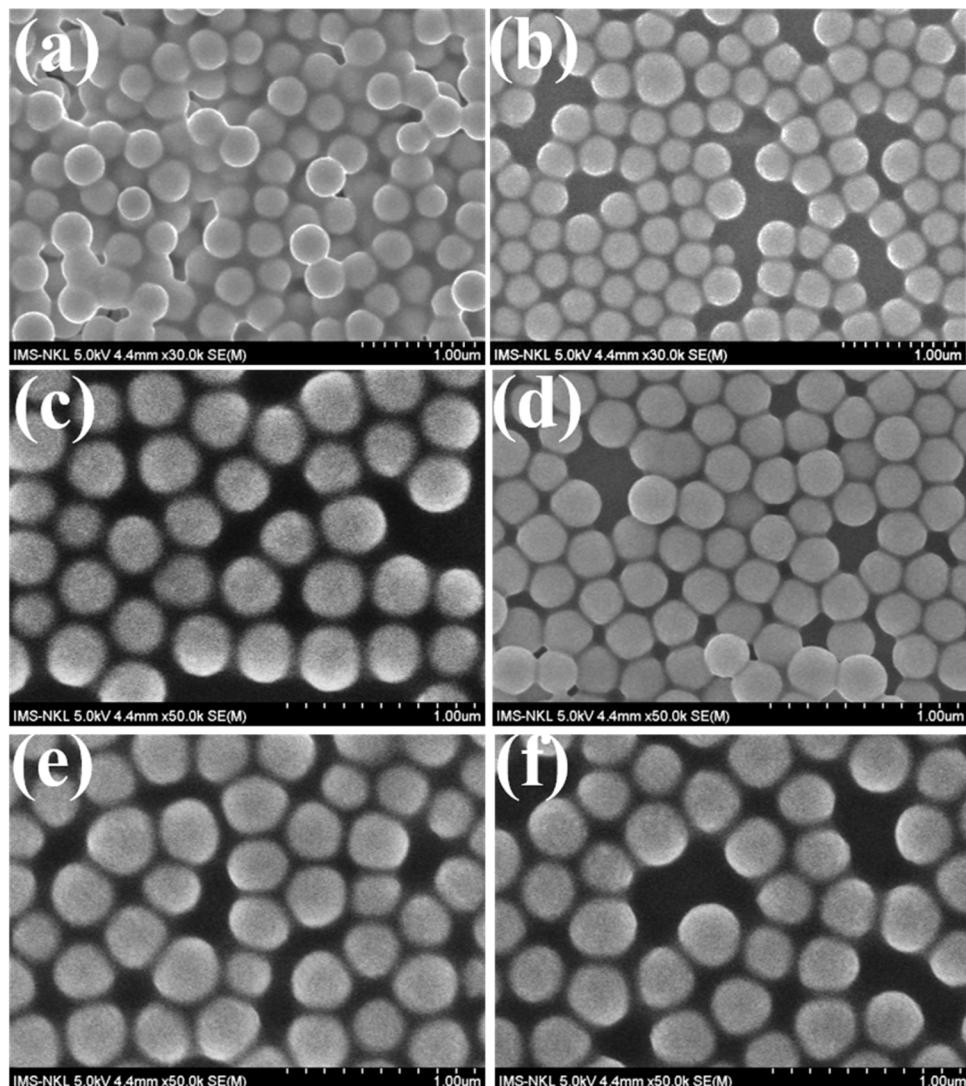


Fig. 1 FESEM images of Eu-doped Gd_2O_3 /CTAB with concentrations of 2% (a), 4% (b), 6% (c), 8% (d), 10% (e) and 12% (f).

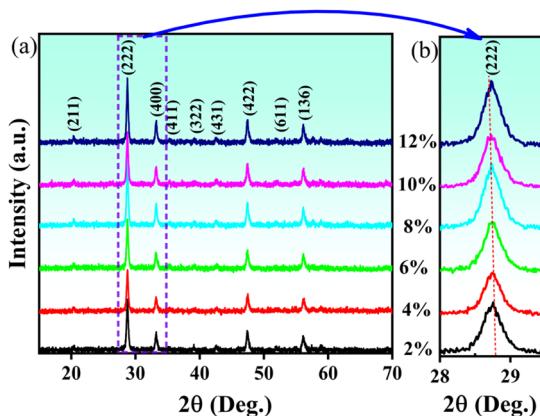


Fig. 2 X-ray diffraction patterns of Eu-doped Gd_2O_3 /CTAB samples with concentrations ranging from 2 to 12% in the range of 15 to 70 degrees (a) and magnification of the (222) peak (b).

240 nm (Fig. 1e). Notably, even at a high doping level of 12% (Fig. 1f), the synthesized spheres retain their spherical shape, with diameters ranging from 220–280 nm.

These results indicate that variations in the Eu^{3+} doping ratio significantly affect both the shape and size of the material, with a clear trend of increasing size as the Eu^{3+} doping ratio rises.

3.1.2. XRD patterns. The X-ray diffraction patterns of Eu^{3+} -doped Gd_2O_3 /CTAB nanoparticles with Eu^{3+} ion content ranging from 2 to 12% in Fig. 2 show distinct diffraction peaks at 2θ values of 20.12° , 28.59° , 33.13° , 38.94° , 42.51° , 47.42° , and 56.29° . These peaks correspond to the crystal planes (211), (222), (400), (332), (431), (422), and (662), which are characteristic cubic structures of Gd_2O_3 , as indicated by the XRD reference pattern (JCPDS card 00-012-0797).¹² The magnification results from 28 to 29.5° with the (222) lattice plane show that the position of this diffraction peak almost does not change when the Eu concentration increases, which shows that when Eu replaces Gd in the Gd_2O_3 cubic lattice, it does not distort the crystal lattice of the material. This result is believed to be due to

the Eu^{3+} ion radius being similar to the Gd^{3+} radius, so the replacement process does not greatly affect the material's crystal lattice. At the same time, the X-ray diffraction pattern of the $\text{Gd}_2\text{O}_3/\text{CTAB}$ material samples doped with Eu^{3+} ions does not show any new X-ray diffraction peaks (especially the diffraction peak of the Eu_2O_3 structural phase) when the Eu concentration changes from 2 to 12%. Based on the X-ray diffraction pattern, we have determined the average crystallite size of the Eu-doped $\text{Gd}_2\text{O}_3/\text{CTAB}$ nanoparticles with different concentrations using the Scherrer formula eqn (1):

$$D = \frac{K\lambda}{\beta \cos \theta} \quad (1)$$

where D is the average crystallite size, K is the dimensionless constant value (0.89), $\lambda = 1.540 \text{ \AA}$ is the wavelength of the irradiation and β is the full width at half maximum measured in radian. Calculating according to eqn (1), we obtained the average crystallite sizes of the materials as 25.40, 27.73, 27.04, 26.93, 26.43 and 26.93 nm corresponding to the Eu-doped $\text{Gd}_2\text{O}_3/\text{CTAB}$ samples with concentrations of 2, 4, 6, 8, 10 and 12%, respectively.

3.1.3. High-resolution transmission electron microscopy.

To determine the structure of Eu-doped $\text{Gd}_2\text{O}_3/\text{CTAB}$ material more precisely, we measured the HRTEM of $\text{Gd}_2\text{O}_3/\text{CTAB}$ sample doped with 8% Eu^{3+} ions and obtained the results in Fig. 3. The results observed in Fig. 3a show that the particles have an average diameter of about 280 nm with a typical cubic structure of Gd_2O_3 material with characteristic lattice planes (211), (222), (400), (420), (521), (600), and (662) (Fig. 3b) with the distance d between the lattice planes (211) shown in Fig. 3b is 3.44 \AA . The HRTEM analysis results are consistent with the XRD results, confirming the consistency in the structure analysis by HRTEM and XRD.

3.1.4. XPS spectra. To further determine the chemical bonding state of $\text{Gd}_2\text{O}_3/\text{CTAB}$ and $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}/\text{CTAB}$ materials, we conducted XPS measurements of $\text{Gd}_2\text{O}_3/\text{CTAB}$ and $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}/\text{CTAB}$ samples doped with 8% Eu^{3+} ions. Fig. 4 shows the XPS spectra of two samples, $\text{Gd}_2\text{O}_3/\text{CTAB}$ and $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}/\text{CTAB}$, doped with 8% Eu^{3+} ions, $\text{Gd} 3d$ (b), $\text{O} 1s$ (c) and $\text{Eu} 3d$ (d).

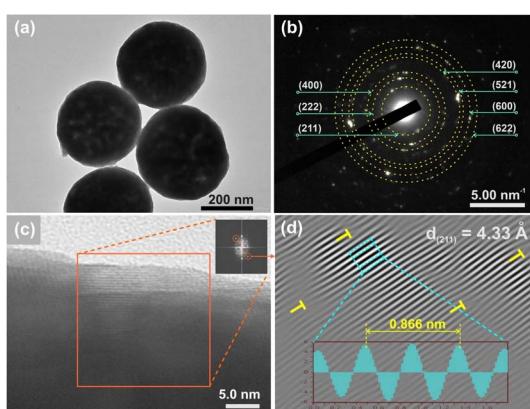


Fig. 3 (a) TEM, (b) electron diffraction, (c and d) HRTEM images of $\text{Gd}_2\text{O}_3/\text{CTAB}$ sample doped with 8% Eu^{3+} ions.

and $\text{Eu} 3d$ (d). The measuring device corrected the survey data by $\text{C} 1s$ (284.7 eV). Fig. 4a shows the characteristic peaks of Gd , O in both $\text{Gd}_2\text{O}_3/\text{CTAB}$ and $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}/\text{CTAB}$ samples. However, with the $\text{Gd}_2\text{O}_3/\text{CTAB}$ sample supplemented with Eu element, the $\text{Eu} 3d$ peak is located at about 1135 eV. For the $\text{Gd}_2\text{O}_3/\text{CTAB}$ sample, in the high energy region of the 3d state of the Gd^{3+} ion, there are two characteristic $\text{Gd} 3d$ binding peaks at the positions 1188.11 eV and 1219.98 eV (ref. 27) with a separation of 31.87 eV, indicating the presence of $\text{Gd} 3d_{5/2}$ and $\text{Gd} 3d_{3/2}$. However, in Fig. 4b, it is observed that these two $\text{Gd} 3d$ peaks are almost insignificantly shifted in both the $\text{Gd}_2\text{O}_3/\text{CTAB}$ and $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}/\text{CTAB}$ samples.

However, the half-width of the peak was broadened, proving that when Eu was doped into the Gd_2O_3 lattice, it affected the local crystal field at the position where the Eu^{3+} ion replaced the Gd^{3+} ion. The effect of this substitution also changed the bonding of the O atom with other ions. Fig. 4c shows the bonding state of $\text{O} 1s$, and the result shows that when there is a circular shoulder of the Eu^{3+} ion in the lattice, the half-width of the $\text{O} 1s$ peak narrows towards low energy. This result is believed to be due to the Eu^{3+} ion radius and its electronegativity (1.0) being smaller than the Gd^{3+} ion radius and electronegativity (1.2), narrowing the half-width towards this low binding energy region. To compare this phenomenon, we conducted a survey of the high-resolution XPS spectrum of the $\text{O} 1s$ state and fitted the peaks corresponding to the characteristic bonds in Fig. 5. At the same time, the high-resolution XPS spectrum in the range from 1125 eV to 1160 eV in Fig. 4d shows that in the $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}/\text{CTAB}$ sample, peaks appeared at 1134.25 eV and 1155.15 eV corresponding to the $\text{Eu} 3d_{5/2}$ and $\text{Eu} 3d_{3/2}$ states with a separation of 20.9 eV of the Eu^{3+} ion, while in the $\text{Gd}_2\text{O}_3/\text{CTAB}$ sample, these two bond peaks were not present.

3.2. Optical studies

In this study, the band gap of $\text{Gd}_2\text{O}_3/\text{CTAB}$ and $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}/\text{CTAB}$ materials was deduced from the UV-vis spectrum according to the Kubelka–Munk eqn (2):

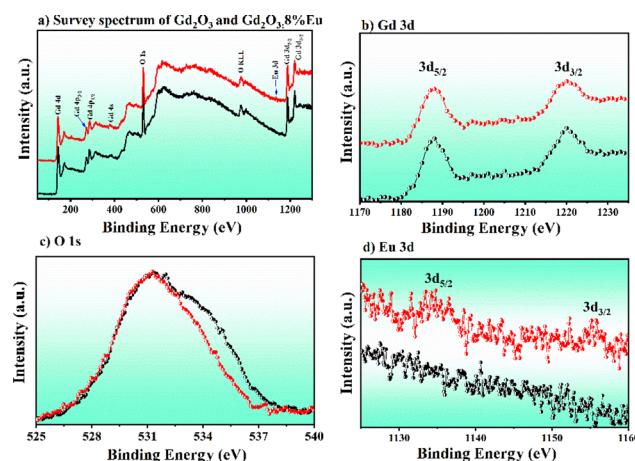


Fig. 4 XPS spectrum of Gd_2O_3 and $\text{Gd}_2\text{O}_3:8\%$ Eu nanorods: (a) survey spectrum, (b) $\text{Gd} 3d$, (c) $\text{O} 1s$ and (d) $\text{Eu} 3d$ levels.



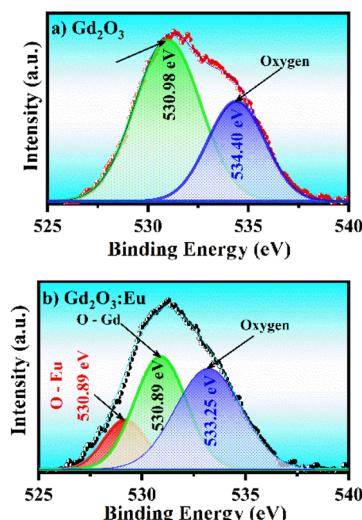


Fig. 5 High-resolution XPS spectra of Gd_2O_3 (a) and Gd_2O_3 :8% Eu (b) nanoparticles corresponding to the deconvoluted O 1s level.

$$(F(R_\infty)h\nu)^\gamma = A(h\nu - E_g) \quad (2)$$

where $(F(R_\infty)h\nu)^\gamma$ and $h\nu$ are the reflection coefficient and photon energy, R_∞ in the % reflectance obtained. A is a characteristic constant of the specific material, respectively. The value of γ is dependent on the origin of transition in a semiconductor ($\gamma = 1/2$ and 2 for indirect transition, respectively and allowed direct transition). In our case, $\gamma = 1/2$ because Gd_2O_3 /CTAB and $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}$ /CTAB are indirect transition semiconductors. Hence, E_g value of all samples could be calculated by plotting $(F(R_\infty)h\nu)^{1/2}$ versus $h\nu$, as illustrated in Fig. 6. The

calculated band gap energies of Gd_2O_3 /CTAB and $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}$ /CTAB in Fig. 6 for samples with Eu^{3+} ion concentrations of 2, 4, 6, 8, 10 and 12% are 3.59, 3.51, 3.46, 3.44, 3.46, and 3.48 eV respectively. This result shows that when the concentration of Eu^{3+} ions doped into the matrix increases, the band gap of the Gd_2O_3 material decreases, and the band gap of the Eu-doped Gd_2O_3 /CTAB material reaches its smallest value at a doping concentration of 8%. When the concentration increases above 8%, the material's band gap tends to increase. The result of this phenomenon is that when the doping concentration increases above a certain threshold, these doping ions cluster together and form clusters and escape from the matrix, leading to a decrease in the density of Eu^{3+} ion emission centers when the doping concentration increases above 8%. This phenomenon will affect the optical properties of the material.

3.3. Optical properties

To study the optical properties of the $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}$ /CTAB material, we measured the fluorescence spectrum of the $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}$ /CTAB sample doped with 8% Eu^{3+} ions. The results obtained in Fig. 7 show that the material emits strongly in the red light region with characteristic emission peaks of Eu^{3+} ions corresponding to the transition from the ${}^5\text{D}_0$ state to the ${}^7\text{F}_J$ state ($J = 0-4$): ${}^5\text{D}_0 \rightarrow {}^7\text{F}_0$ (580 nm), ${}^5\text{D}_0 \rightarrow {}^7\text{F}_1$ (588–600 nm), ${}^5\text{D}_0 \rightarrow {}^7\text{F}_2$ (607–620 nm), ${}^5\text{D}_0 \rightarrow {}^7\text{F}_3$ (620–632 nm) and ${}^5\text{D}_0 \rightarrow {}^7\text{F}_4$ (701–712 nm) in which the emission peak at 611 nm has the most vigorous intensity.^{27–29} This emission peak is a transition from the ${}^5\text{D}_0$ state to the ${}^7\text{F}_2$ state, where the parity rule does not forbid the electric dipole state of the Eu^{3+} ion. The characteristic properties of this fluorescence spectrum once again demonstrate that the Gd_2O_3 matrix has a cubic structure.

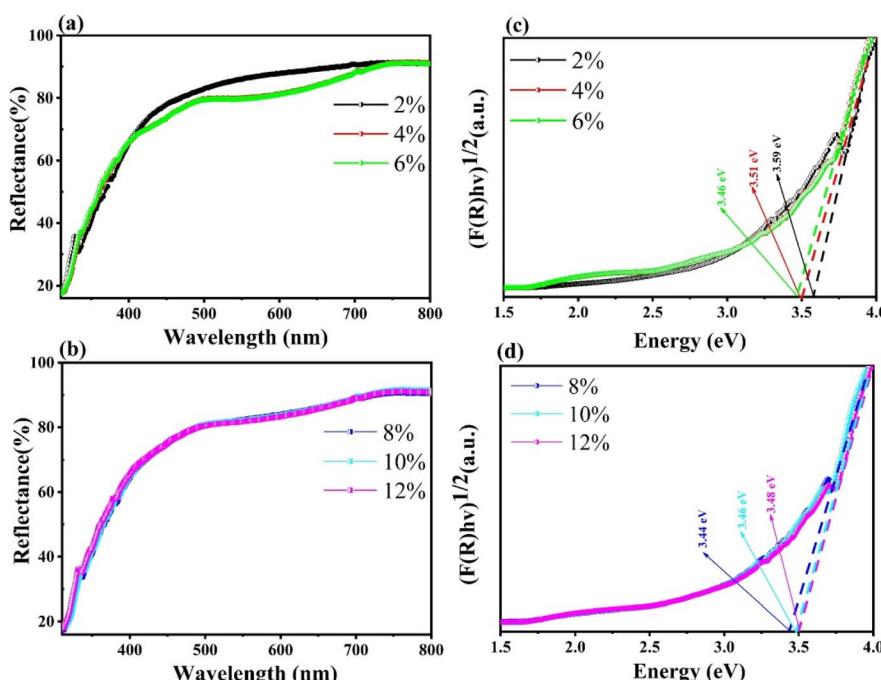
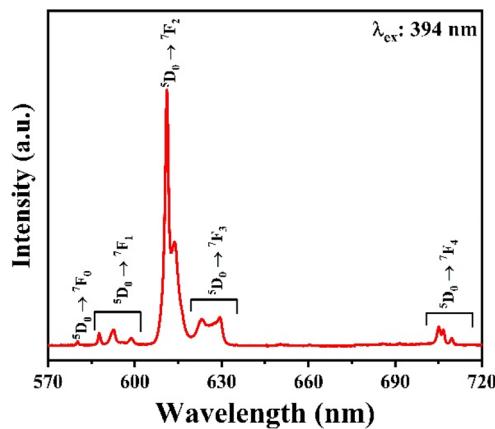
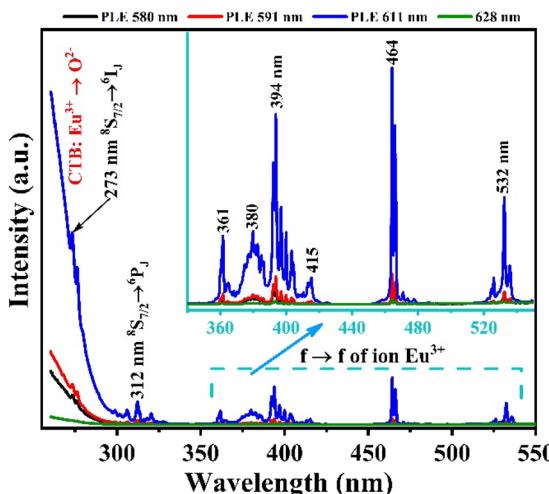
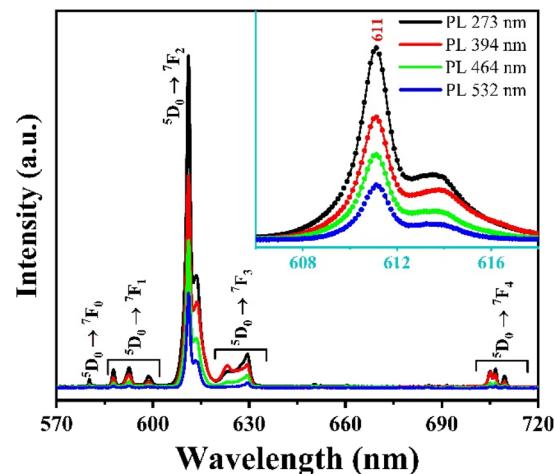


Fig. 6 UV-vis spectra of $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}$ with different mol concentrations (a and b) and energy band gap of $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}$ for different dopant concentrations (c and d) using K-M theory.



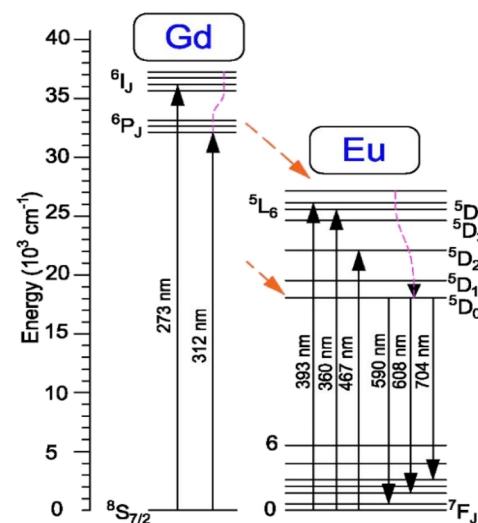
Fig. 7 Fluorescence spectrum of Gd_2O_3 :8% Eu^{3+} /CTAB.

Based on the fluorescence analysis results of the material, we measured the fluorescence excitation spectrum of the Gd_2O_3 /CTAB material sample doped with 8% Eu^{3+} ions with different emission peaks at 580, 591, 611, and 628 nm. The results obtained in Fig. 8 show that the material strongly absorbs in the ultraviolet and visible regions with excitation peaks at 252, 273, 312, 360, 395, 464, and 532 nm. The strong absorption band in the UV region at 252 nm is attributed to the charge transfer transition (CTB) between $\text{Eu}^{3+} \rightarrow \text{O}^{2-}$. The absorption peaks at 273 nm and 312 nm are the characteristic absorption peaks of Gd^{3+} ions corresponding to the $^8\text{S} \rightarrow ^6\text{I}$ and $^8\text{S} \rightarrow ^6\text{P}$ transitions, respectively.²⁷⁻²⁹ Meanwhile, the excitation peaks at 360, 394, 464, and 532 nm are the characteristic absorption peaks of Eu^{3+} ions corresponding to the f-f transition. Among these absorption bands, the CTB band between $\text{Eu}^{3+} \rightarrow \text{O}^{2-}$ is the strongest, and the emission peak for fluorescence excitation has the highest intensity at 611 nm. This result shows that when excited at different wavelengths, the energy level transition of Eu^{3+} ions from the $^5\text{D}_0$ state to the $^7\text{F}_2$ state has the highest transition probability.

Fig. 8 Fluorescence excitation spectrum of Gd_2O_3 :8% Eu^{3+} /CTAB with different emission peaks.Fig. 9 Fluorescence excitation of Gd_2O_3 :8% Eu^{3+} /CTAB.

To evaluate the emission ability of the material with different excitation wavelengths obtained in Fig. 8. We measured the fluorescence spectrum of Gd_2O_3 :8% Eu^{3+} /CTAB with excitation wavelengths of 273, 394, 464, and 512 nm.

The results obtained in Fig. 9 show that the positions of the characteristic emission peaks of Eu^{3+} ions in the Gd_2O_3 matrix do not change, but only the intensity of the peaks changes, and no strange peaks are emitted when excited at different wavelengths. This shows that the emission process of the material only includes the energy level transitions of Eu^{3+} ions from the excited state $^5\text{D}_0$ to the state $^7\text{F}_J$ without including the emission process of Gd^{3+} ions. With this different excitation wavelength, the material emits best when excited at 273 nm (corresponding to the energy level transition of Gd^{3+} ion from ^8S - ^6I state), followed by 394 nm. This shows that the absorption process of the Gd_2O_3 matrix and Gd^{3+} , Eu^{3+} ions, when moving to high energy excited states, all tend to shift without emission to the $^5\text{D}_0$ state of Eu^{3+} ion before shifting to the $^7\text{F}_J$ state for characteristic

Fig. 10 The simplified energy level diagram of Gd^{3+} and Eu^{3+} ions.

emission of Eu^{3+} ion. The energy transfer process from the ^8S excited state of Gd^{3+} ion to Eu^{3+} ion does not lead to an emission process that is worth studying. The mechanism of these energy transfer processes is shown in Fig. 10.

With the results of this study, we will investigate the effect of Eu^{3+} doping concentration on the energy transfer mechanism between Gd^{3+} ions and Eu^{3+} ions by measuring the fluorescence spectrum depending on Eu concentration with an excitation wavelength of 273 nm.

The analysis results in Fig. 11, when excited at 273 nm with $\text{Gd}_2\text{O}_3/\text{CTAB}$ samples doped with Eu^{3+} ions at concentrations from 2 to 12%, show that when the Eu^{3+} ion concentration is at 2%, the intensity of the fluorescence peaks is very low. This shows that with low Eu concentrations, when excited at 273 nm, the absorbed electrons move to the $^6\text{I}_{7/2}$ state and recombine without emission. Theoretically, Gd^{3+} ions can transfer energy to Eu^{3+} ions when absorbed to a high energy level. However, in the fluorescence spectra of $\text{Gd}_2\text{O}_3/\text{CTAB}$ samples doped with Eu^{3+} ions at different concentrations, we did not observe the emission band of Gd^{3+} ions at the 312 nm emission peak corresponding to the transition of the ^6P excited state to the ^8S state. This result can be explained by the fact that the Gd^{3+} ion has a stable electron structure in the $^4\text{f}_7$ electron configuration.³⁰ Therefore, the host environment has almost no effect on the energy level of Gd^{3+} ions. In addition, considering the very high excitation energy of Gd^{3+} ion, the instantaneous energy transfer from Gd^{3+} ion to Eu^{3+} ion is almost impossible in $\text{Gd}_2\text{O}_3/\text{CTAB}$ material doped with low concentration Eu^{3+} ion. This leads to the lifetime of $\text{Gd}_2\text{O}_3/\text{CTAB}$ samples doped with low-concentration Eu^{3+} ions being usually longer than that of high-concentration doped samples. When the Eu concentration increases, the charge transfer process between Gd^{3+} ions and Eu^{3+} ions increases, causing the fluorescence intensity in the emission band of Eu^{3+} ions from the $^5\text{D}_0$ state to the $^7\text{F}_J$ state to increase and reach a maximum value at a doping concentration

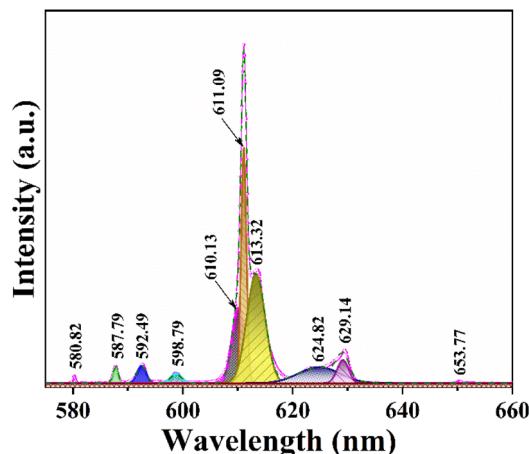


Fig. 12 Fluorescence spectrum of $\text{Gd}_2\text{O}_3:8\%$ Eu^{3+} /CTAB excited at 273 nm fits at different peak positions.

of 8%. The fluorescence quenching phenomenon occurs when the doping concentration increases above 8% in the cubic structure of Gd_2O_3 nanoparticles. This fluorescence quenching result at high Eu concentration is attributed to the fact that Eu^{3+} ion has a similar radius to Gd^{3+} ion, and at the same time, Gd^{3+} ion acts as a photobleach that enhances the luminescence of Eu^{3+} ion when replacing Gd^{3+} ion in general substrate lattices. The concentration of Eu^{3+} ion doping in this Gd_2O_3 substrate lattice is much higher than that in some other substrate lattices such as $\text{Sr}_6\text{P}_5\text{BO}_{20}$,³¹ $\text{Sr}_5(\text{PO}_4)_3\text{Cl}$,³² or with monoclinic Gd_2O_3 structure.

To investigate the crystal symmetry, we analyzed peak splitting in the $\text{Gd}_2\text{O}_3:8\%$ Eu^{3+} sample (Fig. 12). The transition from $^5\text{D}_0$ to $^7\text{F}_1$ is identified as an electric dipole transition that is unaffected by local crystal field symmetry. In cubic Gd_2O_3 , two symmetry positions exist: C_2 and S_6 , occurring in a ratio of 3 : 1. The C_2 position lacks inversion symmetry, while the S_6 position possesses it. When Eu ions occupy the C_2 positions, electric dipole transitions from $^5\text{D}_0$ to $^7\text{F}_2$ follow the selection rule $\Delta J = 2$. In contrast, when occupying S_6 positions with inversion symmetry, the electric dipole transition from $^5\text{D}_0$ to $^7\text{F}_1$ adheres to $\Delta J = 1$.¹⁴ To determine this issue, we have fitted the fluorescence spectra of the $\text{Gd}_2\text{O}_3:8\%$ Eu^{3+} sample at different emission peak positions.

Between 575 nm and 660 nm, several prominent peaks are observed in the luminescence spectrum of Gd_2O_3 samples doped with Eu^{3+} . These peaks correspond to transitions between energy states, specifically from $^5\text{D}_0$ to $^7\text{F}_0$, $^5\text{D}_0$ to $^7\text{F}_1$, $^5\text{D}_0$ to $^7\text{F}_2$, $^5\text{D}_0$ to $^7\text{F}_3$, and $^5\text{D}_0$ to $^7\text{F}_4$. Here's a detailed explanation of the key transitions:

- $^5\text{D}_0$ to $^7\text{F}_1$ transition (around 587.79, 592.49, and 598.77 nm): in highly symmetric crystal fields, the electric dipole transition from $J = 1$ typically does not split into further sublevels. However, in these samples with a monoclinic crystal structure and reduced symmetry, the $J = 1$ state of $^7\text{F}_1$ splits into three sublevels. This splitting is likely influenced by the concentration of impurities.

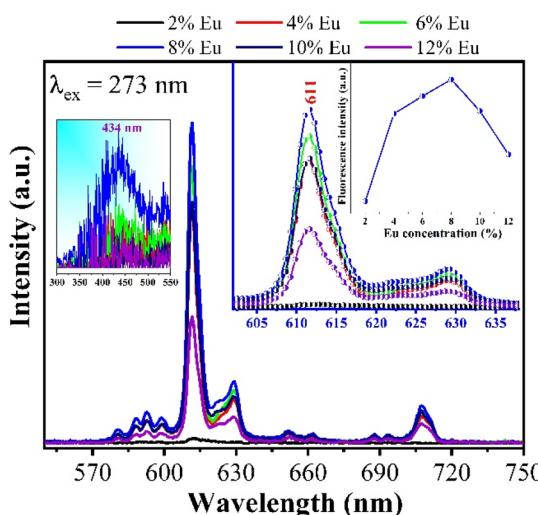


Fig. 11 Fluorescence spectra of $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}$ /CTAB at different concentrations with an excitation wavelength of 273 nm, measured at room temperature.



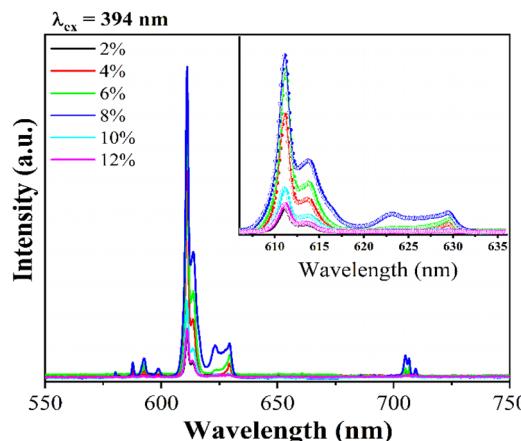


Fig. 13 Fluorescence spectra of $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}/\text{CTAB}$ at different concentrations with an excitation wavelength of 394 nm, measured at room temperature.

• ${}^5\text{D}_0$ to ${}^7\text{F}_2$ transition (around 610.13, 611.09, 613.32, 624.82, and 629.14 nm); the characteristic red luminescence of Eu^{3+} arises from this transition, typically occurring between 610 and 630 nm. Due to the C_s monoclinic symmetry positions of Eu^{3+} , the $J = 2$ state of ${}^7\text{F}_2$ splits into five sublevels. These sublevels are represented by five distinct peaks, indicating transitions from ${}^5\text{D}_0$ to these ${}^7\text{F}_2$ sublevels.

To compare the difference between the energy transfer mechanism from Gd^{3+} ions to Eu^{3+} ions, we also investigated the emission ability of $\text{Gd}_2\text{O}_3/\text{CTAB}$ samples doped with Eu^{3+} at an excitation wavelength of 394 nm, corresponding to the preferential transition of Eu^{3+} ions from the ground state ${}^7\text{F}_0$ to the state ${}^5\text{L}_6$.

The results shown in Fig. 13 show that the material emits strongly in the red light region with the position of the characteristic emission peaks of Eu^{3+} ions from the excited state ${}^5\text{D}_0$ to the state ${}^7\text{F}_j$ almost unchanged compared to when excited at 273 nm. However, we observed that with the 2% doped sample, the intensity of the 611 nm emission peak is relatively large compared to this sample when excited at 273 nm.

This result indicates that, for the energy transfer phenomenon between Gd^{3+} ions and Eu^{3+} ions to occur, the Eu concentration must be large enough to receive this energy transfer process. To compare the above results, we have established the ratio between the 611 nm peak intensity of samples with different concentrations, with the 8% doped $\text{Gd}_2\text{O}_3/\text{CTAB}$

Table 1 Intensity ratio I_x/I_0 (%) (I_0 is the 611 nm peak intensity of the $\text{Gd}_2\text{O}_3/\text{CTAB}$ sample doped with 8% Eu^{3+} ions)

Eu^{3+} concentration	Excitation 273 nm	Excitation 394 nm
2	1.16	14.40
4	52.35	67.08
6	73.57	92.12
8	100	100
10	48.61	25.39
12	33.01	16.23

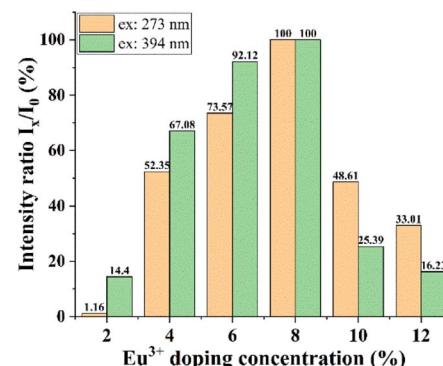


Fig. 14 I_x/I_0 intensity ratio of $\text{Gd}_2\text{O}_3/\text{CTAB}$ samples with different Eu concentrations, with excitation wavelengths of 273 nm and 394 nm.

sample giving the most vigorous intensity; the results are shown in Table 1.

Based on the results obtained in Table 1, we have drawn a graph showing this intensity ratio in Fig. 14. The analysis results show that with low doping concentrations, the energy transfer process between Gd^{3+} ions and Eu^{3+} ions is more complex, making the intensity ratio of the 611 nm peak of the samples compared to the sample with the highest intensity when excited at 273 nm. At the same time, when excited at 394 nm, the intensity ratio of this 611 nm peak to the sample with the highest intensity is stronger. However, when the phenomenon of fluorescence quenching due to concentration occurs, when excited at 273 nm, the quenching process occurs more slowly. This shows that the energy transfer process from Gd^{3+} ions to Eu^{3+} ions reduces the non-radiative recombination process in the material when the doping concentration is high. To supplement this study, we measured the time-resolved fluorescence spectra of $\text{Gd}_2\text{O}_3/\text{CTAB}$ samples doped with Eu^{3+} ions with different doping concentrations, with an excitation wavelength of 273 nm and an emission wavelength of 611 nm.

3.4. Effect of Eu concentration on PL lifetime

To investigate the influence of the concentration of Eu^{3+} ions doped into the matrix of the material, we also measured the

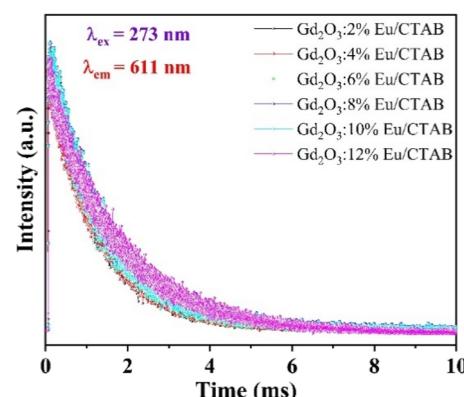


Fig. 15 Decay curves of $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}/\text{CTAB}$ with different mol concentrations.



time-resolved fluorescence spectra of $\text{Gd}_2\text{O}_3/\text{CTAB}$ samples doped with Eu^{3+} ions with doping concentrations from 2 to 12% corresponding to the fluorescence excitation peak at 273 nm and the emission peak at 611 nm (Fig. 15) The results showed that the curves fit a quadratic, exponential function, indicating two separate origins of the emission process in $\text{Gd}_2\text{O}_3/\text{CTAB}$ materials doped with Eu^{3+} ions eqn (3):

$$I(t) = A_0 + A_1 \exp\left(-\frac{t}{\tau_1}\right) + A_2 \exp\left(-\frac{t}{\tau_2}\right) \quad (3)$$

where t is the time, $I(t)$ is the luminescence intensity at time t , A_1 and A_2 are constants, and τ_1 , τ_2 are exponential components of the decay time. The value of the average lifetime τ^* can be calculated using the following formula eqn (3) and (4):

$$\tau^* = \frac{A_1 \tau_1^2 + A_2 \tau_2^2}{A_1 \tau_1 + A_2 \tau_2} \quad (4)$$

Table 2 shows the average lifetime of $\text{Gd}_2\text{O}_3:x\% \text{Eu}^{3+}/\text{CTAB}$ ($x = 2\text{--}12\%$) samples calculated by formula (3).

With the results of fluorescence lifetime analysis of the materials listed in Table 2, it can be seen that, as the Eu concentration increases, the average lifetime of electrons in the excited state tends to increase from 1.18 ms with the 2% Eu doped sample and reaches a maximum of 2.1 ms with the 8% Eu doped $\text{Gd}_2\text{O}_3/\text{CTAB}$ sample and then tends to decrease. The results of the lifetime analysis of these Eu-doped $\text{Gd}_2\text{O}_3/\text{CTAB}$ samples with different concentrations are consistent with the fluorescence analysis results in Section 3.3, as we have analyzed the energy transfer mechanism shown in Fig. 10. The energy transfer mechanisms in Eu^{3+} -doped $\text{Gd}_2\text{O}_3/\text{CTAB}$ materials when the material is excited with a wavelength of 273 nm can include: T1($\text{O}^{2-} \rightarrow \text{Eu}^{3+}$) related to the direct energy transfer between the CTB band to the Eu^{3+} ions; T2($\text{Gd}^{3+} \rightarrow \text{Eu}^{3+}$) related to the energy transfer between the Gd^{3+} ions in the excited state $^6\text{I}_J$ to the Eu^{3+} ions; T3($\text{Eu}^{3+} \rightarrow \text{Eu}^{3+}$) related to the direct energy transfer of the Eu^{3+} ions in the excited state to each other. This is one of the most important mechanisms for concentration-dependent fluorescence quenching. Finally, T4($\text{Eu}^{3+} \rightarrow \text{O}^{2-}$) is the back transfer between Eu^{3+} ions to O^{2-} ; in this process, the back transfer of energy of Eu^{3+} ions from the excited state to the CTB band. During the fluorescence excitation process at 273 nm, the T1 transition is limited because this energy is only enough to excite Gd^{3+} ions to the $^6\text{I}_J$ state. With low-concentration Eu^{3+} doped $\text{Gd}_2\text{O}_3/\text{CTAB}$ samples, the T2

process is less likely to occur than with high-concentration samples because the density of Eu^{3+} ion emission centers in the material is low, so the energy transfer phenomenon between Gd^{3+} ions in the $^6\text{I}_J$ excited state is less likely to occur. We analyzed this result in Section 3.3 when comparing the fluorescence intensity of this sample with the sample with the highest intensity, which is only 1.16%.

Meanwhile, when excited at 394 nm, this ratio is 14.40%. When the Eu doping concentration increases, the density of the emission center increases, this process increases the non-radiative recombination when the Gd^{3+} ion transfers energy from the excited state $^6\text{I}_J$ to the excited state $^5\text{D}_J$ of the Eu^{3+} ion, increasing the lifetime of electrons in the excited state, while increasing the electron density in the $^5\text{D}_0$ state. This process increases the transition of electrons in the $^5\text{D}_0$ state to the $^7\text{F}_J$ state, increasing the fluorescence intensity when the concentration increases. When the Eu concentration increases, the T3 energy level transfer process increases, reducing the electron density in the excited state by the reverse energy transfer process between Eu^{3+} ions, reducing the fluorescence lifetime and intensity due to concentration-dependent fluorescence quenching. The results of the fluorescence lifetime analysis of the material with an excitation wavelength of 273 nm, with an emission peak of 611 nm, are very consistent with the fluorescence analysis results developed in Section 3.3.

3.5. Calculation of Judd–Ofelt parameters

The probability of electric dipole transitions from the $^5\text{D}_0$ state to the $^7\text{F}_J$ state (where $J = 2, 4, 6$) is determined by the following formula eqn (5):¹⁹

$$A_{\text{ED}}(^5\text{D}_0 \rightarrow ^7\text{F}_J) = \frac{64\pi^4 \nu_J^3}{3h(2J+1)} \left[\frac{n(n^2+2)^2}{9} \right] \sum_{\lambda=2,4,6} \Omega_{\lambda} \|U^{(\lambda)}\|^2 \quad (5)$$

Here, ν_J represents the energy of the $^5\text{D}_0$ to $^7\text{F}_J$ transition, e is the electron charge, and $\|U^{(\lambda)}\|^2$ angle denotes the reduced matrix element of the unit tensor operator of rank $\lambda = 2, 4, 6$. The Judd–Ofelt parameter Ω_{λ} can be calculated by analyzing the intensity ratios of the $^5\text{D}_0$ to $^7\text{F}_J$ transitions for $J = 2, 4, 6$ as follows eqn (6):

$$\begin{aligned} \frac{\int I_J d\nu}{\int I_1 d\nu} &= \frac{A(^5\text{D}_0 \rightarrow ^7\text{F}_{2,4,6})}{A(^5\text{D}_0 \rightarrow ^7\text{F}_1)} \\ &= \left[\frac{e^2}{S_{\text{md}_1}} \right] \left[\frac{\nu_J}{\nu_1} \right]^3 \left[\frac{n(n^2+2)^2}{9} \right] \sum_{\lambda=2,4,6} \Omega_{\lambda} \|U^{(\lambda)}\|^2 \end{aligned} \quad (6)$$

For the transitions $^5\text{D}_0 \rightarrow ^7\text{F}_2$, the matrix elements are $U^{(2)} = 0.0033$, $U^{(4)} = U^{(6)} = 0$. For the $^5\text{D}_0 \rightarrow ^7\text{F}_4$ transitions, $U^{(2)} = 0$, $U^{(4)} = 0.0023$, and $U^{(6)} = 0$. Lastly, for the $^5\text{D}_0 \rightarrow ^7\text{F}_6$ transitions, $U^{(2)} = U^{(4)} = 0$ and $U^{(6)} = 0.003$. The total area of the absorption bands for $^5\text{D}_0$ to $^7\text{F}_J$ (with $J = 2, 4, 6$) and $^5\text{D}_0 \rightarrow ^7\text{F}_1$ is also considered.

The intensity parameters Ω_{λ} provide valuable insights into the local environment surrounding the Eu^{3+} ion. The parameter Ω_2 is particularly sensitive to changes in ligand asymmetry and the covalency of the Eu^{3+} -ligand bond: a high Ω_2 value indicates

Table 2 Average lifetimes of $\text{Gd}_2\text{O}_3:x\% \text{Eu}^{3+}/\text{CTAB}$ ($x = 2\text{--}12\%$) phosphors excited at 273 nm and monitored at 611 nm

Eu ³⁺ concentration	Average lifetime, τ^* (ms)	Chromaticity diagram (x, y)	
2	1.18	0.61	0.31
4	1.29	0.62	0.34
6	1.36	0.59	0.33
8	2.1	0.64	0.34
10	1.42	0.63	0.33
12	1.38	0.64	0.33



Table 3 Calculation of $\mathcal{Q}_{2,4,6}$ parameters for $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}$ at various molar concentrations

Samples	$\mathcal{Q}_2 (\times 10^{-20} \text{ cm}^2)$	$\mathcal{Q}_4 (\times 10^{-20} \text{ cm}^2)$	$\mathcal{Q}_6 (\times 10^{-20} \text{ cm}^2)$
$\text{Gd}_2\text{O}_3:2\% \text{ Eu}^{3+}$	6.63	1.15	0
$\text{Gd}_2\text{O}_3:4\% \text{ Eu}^{3+}$	8.1	3.22	0
$\text{Gd}_2\text{O}_3:6\% \text{ Eu}^{3+}$	12.2	3.29	0
$\text{Gd}_2\text{O}_3:8\% \text{ Eu}^{3+}$	14.2	4.01	0
$\text{Gd}_2\text{O}_3:10\% \text{ Eu}^{3+}$	11.2	3.94	0
$\text{Gd}_2\text{O}_3:12\% \text{ Eu}^{3+}$	11	3.95	0

significant ligand asymmetry and high covalency in the Eu^{3+} –ligand bond. On the other hand, \mathcal{Q}_4 reflects the rigidity of the environment embedding the rare-earth ion; a high \mathcal{Q}_4 value corresponds to lower environmental rigidity (Table 3).¹⁸

The transition probability from the excited state J to a lower state J' determines the fluorescence intensity of the J to J' transition eqn (7):

$$A(J, J') = A_{\text{ed}} + A_{\text{md}} \\ = \frac{64\pi^4 \nu^3}{3hc^3(2J+1)} \left[n \left(\frac{n^2+2}{3} \right)^2 S_{\text{ed}} + n^3 S_{\text{md}} \right] \quad (7)$$

Total transition probability and lifetime of the excited state J eqn (8) and (9):

$$A_T(J) = \sum_{J'} A_{JJ'} \quad (8)$$

$$\tau_{\text{cal.}}(J) = \frac{1}{A_T(J)} \quad (9)$$

Branching ratio: used to predict the relative intensity of a fluorescence band from an excited state. The theoretical branching ratio is calculated using the formula eqn (10):

$$\beta_{\text{cal.}}(J \rightarrow J') = \frac{A_{JJ'}}{A_T(J)} \quad (10)$$

Quantum efficiency is determined using the following formula eqn (11):

$$\eta = \frac{\tau_{\text{exp.}}}{\tau_{\text{cal.}}} \quad (11)$$

Quantum efficiency calculations of Eu-doped $\text{Gd}_2\text{O}_3/\text{CTAB}$ materials show that the quantum efficiency ranges from 40% to 91% depending on the Eu doping concentration. The sample with the lowest quantum efficiency is the $\text{Gd}_2\text{O}_3/\text{CTAB}$ sample doped with 2% Eu^{3+} ions; the quantum efficiency gradually increases and reaches a maximum value of 91% with the $\text{Gd}_2\text{O}_3:8\% \text{ Eu}^{3+}$ sample and tends to decrease when the doping concentration is above 8%. The results obtained based on the theoretical model are consistent with the fluorescence survey results. When the Eu concentration is low, the energy transfer phenomenon between Gd^{3+} ions and Eu^{3+} ions is more

Table 4 Branching ratio parameters, lifetime, and quantum efficiency of $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}/\text{CTAB}$ at various molar concentrations

Samples	$\beta_{\text{exp.}} (\%)$	$\beta_{\text{cal.}} (\%)$	A_T	$\tau_{\text{cal.}} (\text{ms})$	$\tau_{\text{exp.}} (\text{ms})$	$\eta (\%)$
$\text{Gd}_2\text{O}_3:2\% \text{ Eu}^{3+}$	76.2	76	334.569	2.9	1.18	40.1
$\text{Gd}_2\text{O}_3:4\% \text{ Eu}^{3+}$	80.1	80	587.111	1.7	1.29	75.8
$\text{Gd}_2\text{O}_3:6\% \text{ Eu}^{3+}$	76.5	75	678.170	1.5	1.36	90.1
$\text{Gd}_2\text{O}_3:8\% \text{ Eu}^{3+}$	76.5	72	426.650	2.3	2.1	91
$\text{Gd}_2\text{O}_3:10\% \text{ Eu}^{3+}$	77.2	77	561.172	1.7	1.42	83
$\text{Gd}_2\text{O}_3:12\% \text{ Eu}^{3+}$	75.1	74	527.372	1.8	1.38	76

challenging, so the material absorbs and leads to large non-radiative recombination. When the concentration of Eu^{3+} ions increases, the density of emission centers increases, and the energy transfer process between Gd^{3+} ions and Eu^{3+} ions is more effective, so the quantum efficiency increases. When the doping concentration reaches 8%, the fluorescence quenching phenomenon occurs, so the quantum efficiency tends to decrease. The result obtained for the highest quantum efficiency is 91%, which is larger than the result we obtained.³³ When synthesizing this material in the environment, only urea solution and TEOS or TOPO were used (Table 4).

3.6. Magnetic properties

The magnetic properties of the $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}$ material with $[\text{Eu}^{3+}]/[\text{Gd}^{3+}]$ ratios of 2, 4, 6, 8, 10, and 12 mol% were analyzed using a vibrating sample magnetometer (VSM) on a MicroSense EZ9 (USA), as shown in Fig. 16. The magnetism of the samples was measured at room temperature under an applied magnetic field of 20 kOe using the VSM system.

The paramagnetic properties of $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}/\text{CTAB}$ arise from the presence of seven unpaired electrons in the 4f shell of Gd^{3+} . These unpaired electrons are shielded from the crystal field by the outer $5s^25p^6$ shell electrons.¹⁴ The shape of the hysteresis curve (M–H) varies across all samples with different concentrations of Gd and Eu ions, likely due to changes in the size of the synthesized particles. As shown in Fig. 16, the magnetism of the Gd_2O_3 material reaches a peak value of 1.83 emu g^{-1} . This value is twice as high as that reported by Zhang *et al.*³⁴ and

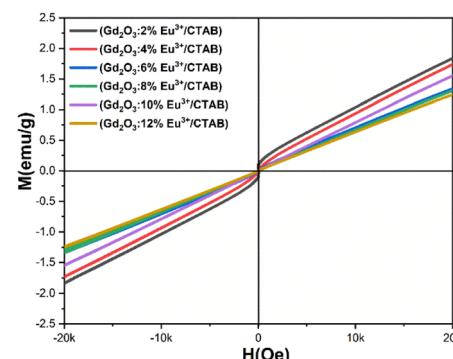


Fig. 16 The hysteresis curve of $\text{Gd}_2\text{O}_3:\text{Eu}^{3+}/\text{CTAB}$ for different dopant concentrations.



comparable to the results obtained by Xu *et al.*³⁵ Although the addition of Eu³⁺ leads to a decrease in magnetic value, the Gd₂O₃:12% Eu³⁺ sample exhibits the lowest magnetism at 1.23 emu g⁻¹, which aligns with the values reported by Zhang *et al.*³⁴ Consequently, Gd₂O₃:Eu³⁺/CTAB material demonstrates significant potential for enhancing the contrast in magnetic resonance imaging (MRI).

To evaluate the potential application of the material's magnetothermal effect in targeting cancer cell destruction, we conducted an investigation of the material's heat generation capability. The inductive heating experiment was conducted in an alternating magnetic field with a frequency of 390 kHz and an intensity of 300 Oe. This magnetic field was generated by an induction coil (7 turns, 3 cm in diameter and 11.5 cm long) connected to a commercial RDO-HFI generator with an output power of 5 kW. The magnetic field intensity was calculated using the formula: $H = nI$, where n is the number of coil turns per unit length, and I is the amplitude of the alternating current flowing through the coil. The samples for measurement were dissolved in a water solution and thermally insulated from the external environment using a vacuum-drawn glass bottle maintained at 10⁻³ to 10⁻⁴ torr. Temperature readings were taken using an optical thermometer (GaAs sensor, Opsens) with an accuracy of ± 0.3 °C within the range of 0 to 250 °C. The specific loss power (SLP) was calculated using the following formula eqn (12):

$$\text{SLP} = C \frac{m_s \Delta T}{m_i \Delta t} \quad (12)$$

where $C = 4.18 \text{ kJ kg}^{-1} \text{ K}^{-1}$ represents the specific heat capacity of the sample system (comprising magnetic particles and the solution), m_s is the total mass of the sample system, and m_i is the mass of the magnetic particles. The initial rate of temperature increase is determined from the tangent of the temperature-time curve at the moment the magnetic field is activated. For an isolated system, the specific loss power (SLP) value is equivalent to the specific absorption rate (SAR).

The results indicate that the samples achieve temperatures ranging from 44 to 49 °C. Physiological studies on cancer cells have demonstrated their limited heat tolerance, identifying a suitable temperature range of 42 to 49 °C for effectively destroying cancer cells without harming healthy ones.³⁶ Therefore, in the research and development of nanomaterials for magnetic heating, it is essential to establish conditions that meet these criteria. With Eu³⁺ doped at molar ratios of 2%, 4%, 6%, 8%, 10%, and 12%, the observed temperatures were 49 °C, 46 °C, 45.3 °C, 45.1 °C, 44.2 °C, and 44 °C, respectively. While increasing the Eu³⁺ doping concentration enhances luminescent properties, it simultaneously reduces the material's magnetic properties. Thus, it is crucial to select an optimal medium doping ratio of Eu³⁺ that balances both luminescent and magnetic properties. These findings confirm that these material systems possess magnetism and can be effectively utilized in magnetic hyperthermia applications (Fig. 17).

To provide a clearer overview, Table 5 presents the parameters from the magnetic induction heating experiment with Gd₂O₃:Eu³⁺/CTAB samples at molar ratios of 2%, 4%, 6%, 8%,

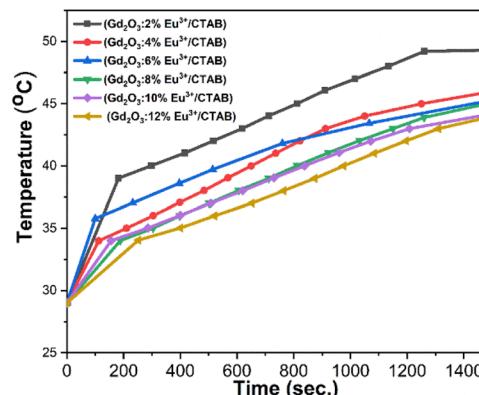


Fig. 17 Heating curves of Gd₂O₃:Eu³⁺/CTAB at varying dopant concentrations.

Table 5 Experimental parameters of magnetic induction heating for Gd₂O₃:Eu³⁺/CTAB at varying dopant concentrations

Sample	(H , Oe)-(f , kHz)	T_s (°C)	dT/dt (°C s ⁻¹)	SAR (W g ⁻¹)
Gd ₂ O ₃ :2% Eu ³⁺	300 Oe	49.3	0.054	11.29
Gd ₂ O ₃ :4% Eu ³⁺		46.0	0.044	18.40
Gd ₂ O ₃ :6% Eu ³⁺		45.3	0.068	9.20
Gd ₂ O ₃ :8% Eu ³⁺		45.1	0.027	5.53
Gd ₂ O ₃ :10% Eu ³⁺		44.2	0.032	6.69
Gd ₂ O ₃ :12% Eu ³⁺		44.0	0.020	4.18

10%, and 12%. The table includes magnetic field intensity, saturation temperature at 1500 seconds, initial heating rate, specific absorption power, and material concentration.

The specific absorption rate (SAR) values for Gd₂O₃ doped with Eu³⁺ at various molar concentrations in the presence of CTAB reveal that SAR increases at lower doping concentrations, while it gradually decreases at higher concentrations. This trend indicates that an increase in the Eu³⁺ ratio within the composite nanoparticles leads to a reduction in the magnetization saturation (MS) value, which subsequently decreases heat generation from magnetic induction. However, the SAR remains adequate to achieve temperatures above 42 °C, ensuring its suitability for magnetic induction heating applications.

4. Conclusions

In this study, Gd₂O₃:Eu³⁺/CTAB material was synthesized by chemical method through many steps in the presence of CTAB. The obtained material has a spherical shape, an average size distribution from 40 to 220 nm, and a typical cubic structure of Gd₂O₃. The material's band gap depends on the doping concentration of Eu³⁺ ions. The band gap tends to decrease when the Eu doping ratio increases and reaches the smallest value of about 3.44 eV when the doping ratio is 8%. Then, the band gap tends to increase when the concentration increases above 8%. This result is because when the Eu concentration

increases, Eu³⁺ ions replace Gd³⁺ ions in the matrix of the material, forming emission centers in the band gap of the Gd₂O₃ material, causing the band gap to decrease. When the Eu ratio increases, these ions tend to cluster together and escape from the Gd₂O₃ material's matrix, causing the material's band gap to increase when the doping ratio increases above 8%.

The Gd₂O₃:Eu³⁺/CTAB material strongly absorbs in the ultraviolet region, giving strong emission in the red light region with characteristic emissions of Eu³⁺ ions from the ⁵D₀ excited state to the ⁷F_J state ($J = 0-4$). The fluorescence spectrum shows that when the Eu doping ratio is low, the energy transfer process between the bright Gd³⁺ ions and the Eu³⁺ ions is low. The fluorescence quenching phenomenon due to the concentration of this material system corresponds to the doping ratio of 8%. With the best-emitting sample, Gd₂O₃:Eu³⁺/CTAB has a quantum efficiency of about 91%.

In addition, the Gd₂O₃:Eu³⁺/CTAB material also exhibits weak ferromagnetic properties, with a maximum magnetic field of 1.83 emu g⁻¹. This result is twice as high as the maximum magnetic field obtained by other groups using different precursors when synthesizing the material instead of CTAB. Magneto-thermal experiments show that the material can reach temperatures from 43 to 49 °C, within the optimal range for effectively killing cancer cells without harming surrounding healthy cells.

Overall, the results obtained for the Eu³⁺ ion-doped Gd₂O₃/CTAB material show that it possesses optical and magnetic properties, making it a promising candidate for multifunctional applications in diagnosis and treatment.

Data availability

All data are presented in the article.

Author contributions

Pham Thi Lien: writing – original draft, methodology, investigation, formal analysis, data curation, conceptualization. Tong Quang Cong: methodology, investigation. Vu Ngoc Phan: methodology, investigation. Vu Thi Hong Ha: writing – original draft, methodology. Ho Thi Anh: methodology, investigation. Pham Duc Thang: methodology, investigation. Le Thi Hoi: methodology, investigation. Dinh Manh Tien: methodology, formal analysis. Nguyen Thanh Huong: methodology, investigation. Hoang Thi Khuyen: methodology, investigation. Lam Thi Kieu Giang: methodology, investigation. Nguyen Vu: methodology, formal analysis. Pham Hong Nam: methodology, formal analysis. Le Tien Ha: writing – review & editing, and editing the final manuscript.

Conflicts of interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

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