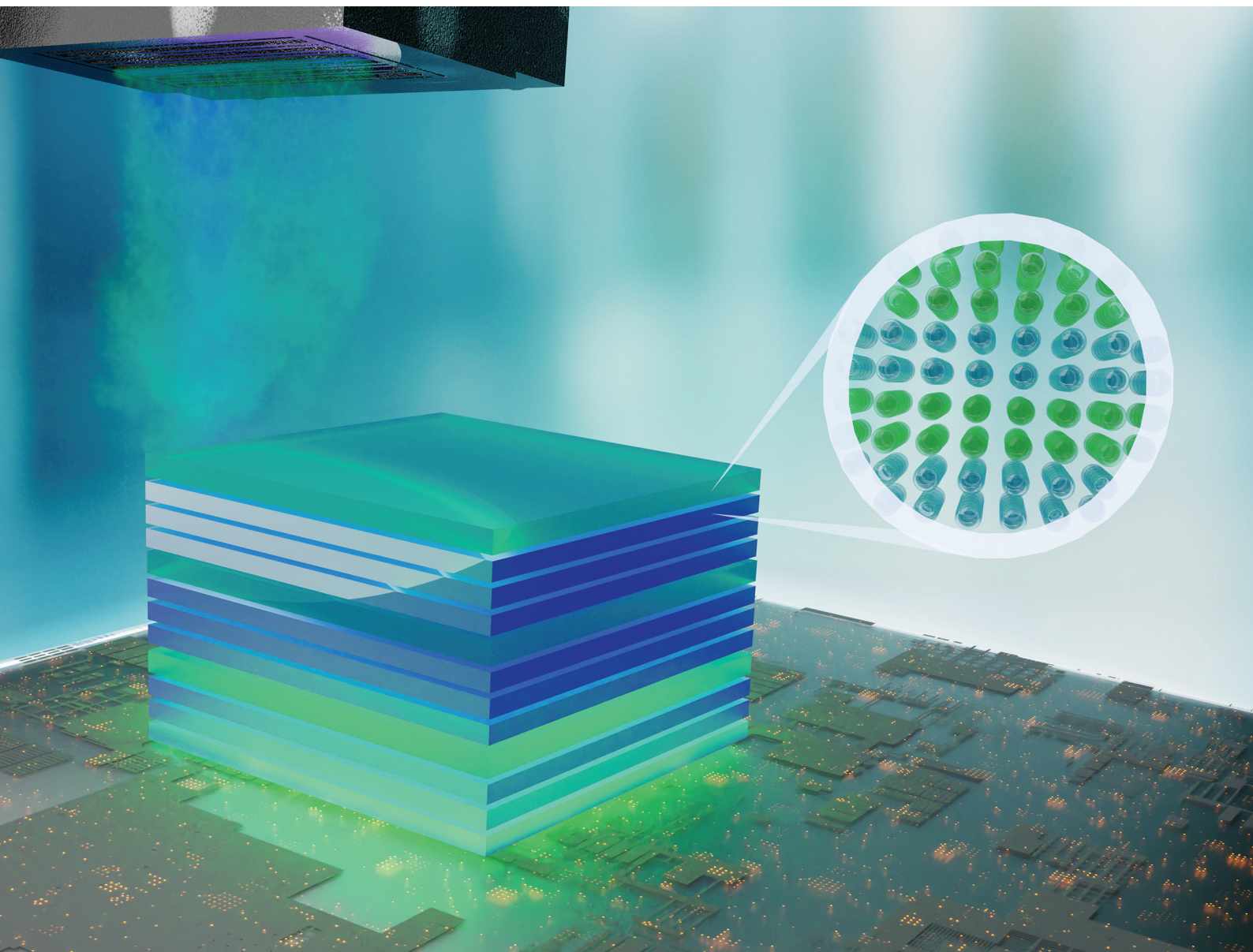


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Multi-head spatial atomic layer deposition: a robust approach for precise doping and nanolaminate fabrication in open-air environments†

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Open-air manufacture of functional nanolaminates demands rapid and scalable methods that maintain nanoscale precision across individual layers. Conventional spatial atomic layer deposition (SALD) employs a single deposition head, and the concurrent exposure to mutually reactive precursors does not provide precise layer thickness accuracy nor sharp interfaces. This work addresses these challenges with an innovative multi-head SALD design. In this study, we introduced a novel multi-head SALD system comprising a uniform head for depositing 5 nm zinc oxide (ZnO) sublayers, while a combinatorial high-throughput head delivered wedge-like thickness gradient in aluminum oxide (Al₂O₃) sublayers, increasing gradually from 0 to approximately 6 nm along the substrate. Eight ZnO/Al₂O₃ bilayers (320 ALD cycles per material) were completed in open air at 200 °C. Transmission electron microscopy recorded highly sharp and reproducible layer-to-layer structures with thickness deviations of 0.3 nm; X-ray photoelectron spectroscopy verified the absence of undesirable aluminum contamination in the reference ZnO regions. With increasing Al₂O₃ thickness, surface roughness was reduced from 0.58 to 0.33 nm, ZnO (002) reflections broadened as crystallite size was confined, and Kelvin probe measurements showed a shift in work function values from 4.75 to 4.35 eV with Al₂O₃ passivation. This study demonstrates that the multi-head SALD method offers a robust platform for achieving precise multilayer control in open-air conditions, thereby paving the way for further innovations in materials research through optimized experimental conditions and combinatorial approaches.

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Introduction

With technological advancement, there is growing demand for time- and cost-effective methods to fabricate high-quality materials. Zinc oxide (ZnO) and aluminum oxide (Al₂O₃) nanolaminates are layered structures that exhibit optical, electrical, and mechanical properties derived from the combined characteristics of both materials. ZnO, known for its wide bandgap, offers semiconducting properties,¹ a wurtzite structure exhibits a built-in polarization and piezoelectric response,² and a high exciton binding energy,³ making it suitable for photonic and electronic devices.⁴ Al₂O₃, on the other hand, is recognized for its high dielectric strength and stability,⁵ thereby enhancing

the protective and insulating capabilities of the composite. Recent studies suggest that engineering the thickness of ZnO and Al₂O₃ layers in ZnO/Al₂O₃ nanolaminates offers great tunability across multiple properties, including mechanical characteristics such as fracture resistance,⁶ hardness and toughness,⁷ dielectric properties,^{8,9} charge storage,¹⁰ work function,¹¹ bandgap,^{12,13} luminescent behaviors and efficacy,^{11,14–16} thermal conductivity,¹⁷ photocatalytic efficiency,¹⁸ gas permeability,⁶ and optical transparency.^{13,19} Therefore, ZnO/Al₂O₃ nanolaminates have been extensively studied for applications such as dielectric barriers, transparent conducting coatings, and optoelectronic components, including solar cells and ultraviolet detectors.^{11,20}

These nanolaminates are often synthesized through atomic layer deposition (ALD), which enables precise control over thickness, composition, and uniformity at the atomic scale.²¹ ALD stands out as a powerful technique due to its ability to achieve nanoscale thickness precision and high conformality. The time-resolved, sequential nature of ALD allows precise control over reaction kinetics, the incorporation of dopants,

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surface treatments, and *in situ* characterization, making it a unique tool in materials research. However, a key limitation of ALD is its slow deposition rate.²² Additionally, conventional ALD systems typically require a chamber and a vacuum system to ensure effective purging, which increases operational costs and complicates maintenance. There is, therefore, a significant gap in the literature regarding open-air direct deposition of nanolaminates. According to our present knowledge, the only reported instance of open-air ZnO/Al₂O₃ deposition, not involving ALD, was conducted by Gonullu utilizing a spray-pyrolysis technique.¹⁸ That work exhibited a single bilayer ZnO/Al₂O₃ with a film thickness variation between 500 nm to 1 μm. It also revealed a considerable diffusion of Al into the ZnO matrix, resulting in a lack of sharp interfaces. This falls far short of reaching the controlled features obtained with ALD.

To overcome these limitations, atmospheric-pressure spatial atomic layer deposition (AP-SALD) has emerged as a promising alternative. Conventional SALD setups employ a spatially sequenced injector configuration, in which gas injectors are integrated into the deposition head. This arrangement enables the sequential exposure of the substrate to different precursors as it moves through the respective zones.^{23,24} Unlike ALD, close-proximity SALD directly targets the substrate with precursors to create localized and reduced reaction zones, which provides two key advantages. First, enhanced mass transport efficiency allows operation in open-air environments without requiring a (vacuum) chamber. This enables continuous deposition processes suitable for large-area applications, such as roll-to-roll (R2R) systems. Second, it reduces the required time for surface saturation and purging per ALD step, which significantly increases the deposition rate—up to 100 times faster than traditional ALD.^{24–26} Additional advantages of reduced reaction zones include lower precursor consumption and a reduced heating zone, making SALD a more sustainable and environmentally friendly gas-phase deposition technique.^{27,28} The efficiency of SALD has drawn considerable interest from both industrial and academic sectors, further reinforcing its appeal for industrial-scale production.

When using a close-proximity SALD approach, deposition quality depends strongly on the head geometry design. Many studies have suggested improvements in head design to reduce intermixing and prevent unwanted chemical vapor deposition (CVD) reactions, often guided by computational simulations.^{26,29,30} Despite these efforts, careful optimization of deposition parameters such as substrate speed, deposition temperature, bubbler temperatures, gas flow rates for each SALD head design and the selection of suitable precursor chemistry is still required. Finding the optimal deposition conditions can easily demand considerable time and effort. Consequently, combinatorial synthesis methods have emerged as a highly efficient strategy to deposit samples under varying deposition parameters (*e.g.*, flow rate,³¹ deposition temperature,³² and others³³) on a single sample. Another significant challenge in SALD is the controlled usage of different metal precursors for doping or nanolaminates, which is crucial for

adjusting electrical properties, carrier mobility, optical characteristics such as band gap, and structural features like defect density and grain boundaries.^{34,35} The most common multi-precursor approach in SALD involves co-dosing, where two precursors are mixed in a manifold. While co-dosing is often presumed to enable isotropic mixing without significant precursor competition, in practice, it frequently results in unintended cross-reactions and the formation of undesirable by-products. The precursors for metal oxides and metal nitrides may interact in ways that hinder the desired growth. For example, Lancaster *et al.* highlighted the difficulty of growing Zn(O,S) thin films for tuneable band gaps with controlled sulphur content due to a significant surface exchange reaction between ZnO and H₂S, which makes achieving the desired composition and homogeneity challenging.³⁶ Additionally, temporarily separated co-dosing tends to produce more mixed or alloyed layers, making it difficult to obtain sharp interfaces between layers, such as in dielectric or barrier stacks, thereby rendering this method unsuitable for producing complex multilayer or composite film structures where precursor control over layer composition and thickness is required.³⁷ This mixed exposure of precursors also risks undermining the sequentiality of ALD (one of the key merits of ALD)—namely, the capability to achieve controlled, uniform, and self-limiting film growth. Precursor mixing introduces variability in dopant incorporation and potentially degrades the desired material properties, as also discussed by Macco and Kessels.³⁸ Under these conditions, the super-cycle approach—where precursors are pulsed sequentially—provides a more reliable solution for ALD processes by offering improved control over material composition and interface sharpness. Consequently, this method is frequently employed for the deposition of nanolaminates³⁹ and applications requiring precise doping control, such as aluminium-doped ZnO (AZO).⁴⁰

In this work, we demonstrate a precisely controlled multi-precursor usage through a multi-head SALD set-up in open-air. In comparison with conventional mono-head SALD setups,⁴¹ the proposed configuration provides enhanced thickness control and eliminates undesired cross-contamination. The proof-of-concept is demonstrated by the deposition of ZnO/Al₂O₃ nanolaminates. We demonstrate the versatility of our approach by exploiting two separate deposition heads: the ZnO sublayer is deposited using a standard uniform deposition head, while a custom-designed combinatorial high-throughput (CHT) head is employed for Al₂O₃ layer deposition, thereby generating a wedge-like Al₂O₃ thickness gradient ranging from 0 to approximately 6 nm in each layer.³³ Comprehensive characterizations demonstrate precise control over the bulk properties as the Al₂O₃ sublayer thickness increases, and precise elemental analyses confirm the absence of Al contamination, with structural characterizations revealing highly sharp layer interfaces and reproducible thickness control in a few nanometer precision. Overall, the proposed multi-head SALD approach exhibits significant potential for the facile fabrication of controlled doping and multilayer film structures in open-air, thereby broadening the scope of high-throughput material studies with enhanced controllability.



Experimental methods

Film depositions were conducted using a custom-built AP-SALD system. Diethylzinc (DEZ, $\text{Zn}(\text{C}_2\text{H}_5)_2$) was used as the zinc precursor, while trimethylaluminium (TMA, $\text{Al}(\text{CH}_3)_3$) served as the aluminium precursor. Ultrapure water (H_2O) was used as the co-reactant. Both DEZ, TMA, and H_2O bubblers were contained at room-temperature during the deposition process. Nitrogen gas was employed as both the carrier and barrier gas throughout the deposition. Two distinct deposition heads were utilized: one with a uniform metal flow channel for DEZ and another with a combinatorial channel for TMA. Each head was equipped with one metal channel, two water channels, and six nitrogen channels, the latter functioning as both a barrier and purging mechanism. For ZnO deposition, flow rates per channel were: 20 sccm DEZ (bubbler), 100 sccm N_2 (DEZ dilution), 30 sccm H_2O (bubbler), 60 sccm N_2 (H_2O dilution), and 100 sccm N_2 (barrier). For Al_2O_3 deposition, flow rates per channel were: 10 sccm TMA (bubbler), 100 sccm N_2 (TMA dilution), 30 sccm H_2O (bubbler), 30 sccm N_2 (H_2O dilution), and 300 sccm N_2 (barrier). The substrate-to-head gap was set to 150 μm using a feeler gauge. Highly p-type doped silicon (Si) (100) wafers with an approximately 2 nm native oxide layer (BT Electronics) were used as substrates. Depositions were performed at a temperature of 200 °C. The process involved 40 ALD cycles each for ZnO and Al_2O_3 films, repeated over eight super-cycles, resulting in a total of 320 ALD cycles per material. Substrate speeds were adjusted according to the precursor: 5.23 cm s^{-1} for the ZnO head and 7.85 cm s^{-1} for the Al_2O_3 head.

A custom SALD deposition head was designed with computer-aided tools (Autodesk Fusion) and manufactured with a stereolithography (SLA) printer (Formlabs 2) to facilitate combinatorial high-throughput (CHT) film growth for Al_2O_3 layers and accelerate the optimization of material properties. The substrate was positioned 150 μm below this channel to form a confined space for precursor flow (10 sccm). Computational fluid dynamics (CFD) simulations were performed using COMSOL Multiphysics 5.5 to analyze the gas flow and precursor distribution. A three-dimensional representation of the reactor head was used, and the model of laminar flow and transport of diluted species was applied based on the established low-Reynolds-number conditions inside the SALD reactor. The resulting predictions guided the design by indicating how variations in channel width influence local precursor delivery across the substrate surface. The detailed information for the CFD simulation parameters is given in the ESI.† The thickness of the films was estimated using a FilmSense FS-1 Spectroscopic Ellipsometer. The associated Cauchy model employed for these measurements is provided in the ESI.† For roughness measurements, an atomic force microscope (AFM) was employed. A Bruker Dimension Icon AFM microscope was utilized, operating with platinum-coated silicon probes (PtSi-NCH, from Nanosensors) for the scanning process. Height sensor data were acquired from a 500 nm \times 500 nm area, with the scan speed adjusted to 1 Hz (corresponding to a tip speed

of 1 $\mu\text{m s}^{-1}$) and 256 samples per line scan. Furthermore, the NanoScope Analysis software of Bruker was utilized to extract the three-dimensional AFM images and roughness values.

The crystallinity of the deposited films was examined using a Rigaku Smartlab X-ray Diffractometer (XRD) equipped with a copper X-ray source configured in a grazing-incidence geometry at 0.34° incidence angle. A 2Theta scan was conducted at 45 kV to ensure a sufficiently energetic beam for inducing diffraction while maintaining a parallel beam alignment to reduce geometric aberrations.⁴² The horizontal Soller slit was set at 2.5° to control beam divergence, and a 2.0 mm mask was employed to enhance signal quality and minimize background scattering. The diffracted beam then passed through a vertically parallel slit analyzer (PSA) of 0.5° and a horizontal PSA of 5.0°, thereby refining the scattering profile and improving angular resolution. The morphology of the ZnO films was characterized using a Zeiss GeminiSEM 300 Field Emission Scanning Electron Microscope (FESEM), operated with an 8 kV electron beam source. Energy-selective X-ray microanalysis (EDS) was utilized for initial chemical analysis with 8 kV for different sample regions with a 4 mm working distance. The samples were characterized without surface coating (metallization). A native Bruker software (ESPIRIT) was utilized to calculate the atomic percentage (at%) values from EDS results. For Transmission Electron Microscope (TEM) and Selected Area Electron Diffraction (SAED) analyses, ZnO/Si specimens were prepared in cross-sections using the semi-automated polishing tripod technique with a MultiPrep™ system (Allied High-Tech Products, Inc.). A PIPS II from GATAN was used for the final polishing. TEM and high-resolution TEM (HRTEM) images were recorded with a JEOL JEM 2010 LaB₆ microscope operating at 200 kV with a 0.19 nm point-to-point resolution.

X-ray photoelectron spectroscopy (XPS) was performed using a Thermo Fisher Scientific K-Alpha⁺ spectrometer with an Al K α X-ray source (1486.6 eV). Survey spectra were acquired with a pass energy of 100 eV and a step size of 0.5 eV, while high-resolution (HR-XPS) spectra were collected at a pass energy of 30 eV with a step size of 0.1 eV. All spectra were acquired with a spot size of 400 μm , under fixed flood gun conditions (emission: 100 μA , focus: 25 V, extractor: 40 V, gas cell: 20 V, filament: 3.24 V). High-resolution spectra were analyzed using CasaXPS software,⁴³ applying a Shirley baseline correction and fitting peaks with a combination of asymmetric Lorentzian functions. Binding energy calibration was performed using the adventitious carbon signal in the C_{1s} region, fixing the contribution from C–C and C–H at 284.8 eV. For Zn_{2p} and Al_{2p} HR-XPS spectra, an area ratio of 1:2 and doublet separations of 23.1 \pm 0.1 eV and 0.4 eV, respectively, were considered.

A Kelvin Probe Station (KP Technology, Model KP020) was employed to measure the contact potential difference (CPD) across the sample surfaces. The samples were placed on a grounded stainless steel stage and secured with copper arms that contacted the sample surface to ensure electrical grounding, while the vibrating Kelvin probe tip (3 mm in diameter) was positioned close to the sample surface to generate a paral-



lel-plate capacitor signal. In order to account for the differences in contact potential across the studied samples, the distance between the vibrating tip and the surface was fixed at a gradient force value of 300 (arbitrary unit). All experiments were performed under ambient conditions (relative humidity of $40 \pm 8\%$ and temperature of $25 \pm 3\text{ }^\circ\text{C}$) within a Faraday/optical enclosure. 500 data points were collected for each measurement. The values were averaged, and the standard deviation was noted as noise. The calibration was completed using a reference gold (Au) sample (KP Technology) to provide a reliable reference point for tip and sample work function calculations. The calculation details are provided in the ESI.†

Lastly, ImageJ⁴⁴ software was employed for image processing, while data analysis was performed using OriginLab⁴⁵ and Fityk.⁴⁶ In addition, Inkscape (version 0.92.4) was used for creating high-quality vectorial illustrations.

Results

Nanolaminate film deposition with super-cycle SALD

For the deposition of the combinatorial ZnO/Al₂O₃ nanolaminate film, a novel multi-head SALD setup was utilized. A photograph of the multi-head SALD setup is presented in Fig. 1a. Each ZnO/Al₂O₃ bilayer was repeated 8 times with an increasing Al₂O₃ layer, as shown in Fig. 1b. The majority of the characterization focused on three regions, hereafter referred to as R1, R2, and R3 for convenience. R1 refers to the ZnO region where minimal TMA exposure occurs; ideally, no detectable Al content is expected. In the combinatorial film, an increasing Al₂O₃ layer thickness is observed in R2, and at R3, the thickest Al₂O₃ sublayer in the ZnO/Al₂O₃ nanolaminate film is attained. This spatial gradient in Al₂O₃ thickness enables the characterization of the implications of increasing Al₂O₃ layer thickness in the ZnO/Al₂O₃ multilayer structure. The two deposition heads—one uniform and one combinatorial—are spatially separated to simulate a super-cycle condition. The substrate completed linear oscillations under deposition heads (Fig. 1c). The uniform head was used for ZnO deposition to maintain a fixed sublayer thickness, while a combinatorial head was used for the deposition of an increasing Al₂O₃ sublayer. The corresponding regions (R1, R2, and R3) are shown in Fig. 1d. Finally, the setup was utilized with different ZnO and Al₂O₃ substrate speeds (corresponding to pulse time), allowing highly flexible deposition conditions (Fig. 1e). The multi-head SALD setup enabled precise control of the number of oscillations under each head as needed. In this study, 40 ALD cycles (20 motor oscillations) were completed per head within 8 super-cycles, corresponding to 320 ALD cycles for each material.

Chemical composition

EDS is a convenient method for obtaining semi-quantitative chemical content information. Fig. 2 shows the EDS results of the film for the R1, R2, and R3 regions of the ZnO/Al₂O₃ nano-

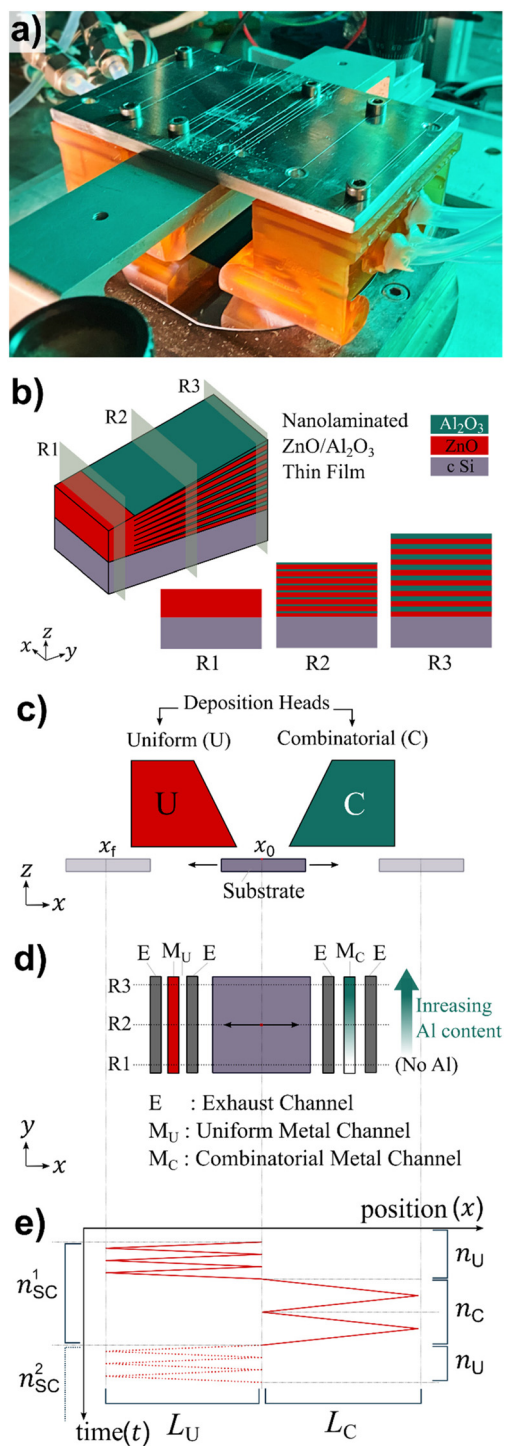


Fig. 1 A concise overview of the multi-head SALD setup. (a) Is the photograph of the configuration used in this work. (b) Illustrates ZnO/Al₂O₃ nanolaminate films with a gradient increase in the Al₂O₃ layer thickness. Regions R1, R2, and R3 are indicated. (c) Depicts both uniform (U) and combinatorial (C) heads along with a designated point (x_0) on the substrate surface that oscillates along the x axis. (d) Provides a bottom view of the setup, highlighting the uniform metal (M_U) and combinatorial metal channels (M_C), where regions R1, R2, and R3 correspond to varying levels of Al contribution. (e) Plots the position of x_0 as a function of time (t). n_U and n_C represent the number of substrate oscillations under uniform and combinatorial heads, respectively. n_{SC} is the number of super-cycles. L_U and L_C are the oscillation distance.



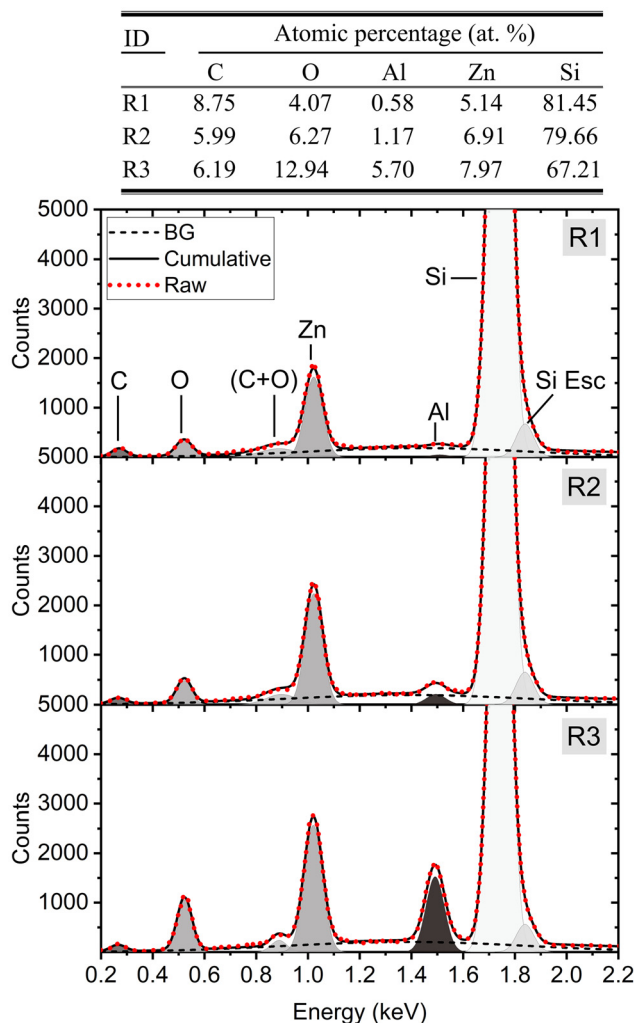


Fig. 2 Deconvoluted EDS spectra for three different regions in combinatorial ZnO/Al₂O₃. Inset table; EDS spectra calculated atomic percentage values for each region (R1, R2, and R3). Cumulative and Raw refer to the envelope of the fitted components and the raw EDS signal, respectively. C + O refers to the pile-up peak from the C K α and O K α emissions. Silicon escape (Si Esc) refers to the artifact caused by interactions between incoming X-rays and the silicon material in the EDS detector.

laminate film. The deconvoluted EDS spectra with cumulative characteristic X-ray energies provide important qualitative information about the chemical compounds present. Instead of removing the broad background signal (BG), it has been fitted with Gaussian curves. The peak at 0.88 keV is attributed to a pile-up peak from C K α and O K α and is denoted as (C + O); this is a common artifact that appears when two or more X-ray photons strike the EDS detector simultaneously or within an extremely short time interval.⁴⁷ The most noticeable change is the increase in Al content in Fig. 2 (inset table), which is 0.58 at% in R1, gradually increasing to 1.17 at% in R2 and 5.70 at% in R3. Accordingly, an increase in oxygen with an increased amount of Al₂O₃ is observed, with 4.07% recorded in R1. The carbon ratio remains between 6.19 and 8.75 at%, a

value that is higher than expected due to ambient carbon contamination. Due to the high information depth of EDS at 8 kV, the bulk Si substrate provides the most intense peak. It is also observed that Zn shows an increasing trend with increasing film thickness, a phenomenon attributed to the nature of the X-ray scattering mechanism in EDS; to understand this in detail, Monte Carlo (MC) simulations have been performed. In the MC simulation, the R1, R2, and R3 regions were modeled, and X-ray absorption values were plotted in Fig. S2† with Phi-Rho-Z (PRZ) distributions, where Phi (ϕ) represents the X-ray intensity, Rho (ρ) denotes the density of the material, and Z represents the depth within the sample, thereby providing information about the X-ray absorption by the specimen. In the PRZ distributions, the total absorbed intensities are plotted for Zn and Si, with Zn elemental compositions of 120.7, 121.8, and 123.7 for R1, R2, and R3, respectively, and Si elemental compositions of 1574.4, 1480.82, and 1417.7 for R1, R2, and R3, respectively. While Zn absorption shows a slight increase with increasing thickness from R1 to R3, Si exhibits a greatly decreasing trend, which correlates with the EDS results. As a semi-quantitative method, EDS suffers from potential error margins and limited detection sensitivity, and it is not entirely reliable for quantitative analysis without corroboration from complementary characterization methods.

XPS survey analysis in Fig. S3† reveals both photoelectron and Auger electron signals, as well as surface compositions corresponding to the expected elements from the SALD deposition process, specifically Zn and O, across all analyzed surfaces. In samples R2 and R3, significant and progressively increasing amounts of Al are evident following SALD deposition (Table S1†), while all surfaces exhibit the characteristic C_{1s} signal attributed to adventitious carbon.

Fig. 3 presents the deconvoluted HR-XPS spectra for the C_{1s}, Zn_{2p}, Al_{2p}, and O_{1s} regions for all surfaces, with Table 1 summarizing the fitting parameters and assignments. Notably, the highest C_{1s} ratio is observed in R1 at 42 at%, whereas both R2 and R3 display ratios of 24 at% (Table S1†). As presented in Fig. 3a, the C_{1s} spectra were deconvoluted into four characteristic components of adventitious carbon at 284.8, 286.3, 287.8, and 288.8–289.0 eV, corresponding to C–C, C–H; C–O; C=O; and O–C=O, respectively.⁴⁸ The HR-XPS spectrum in the Zn_{2p} region (Fig. 3b) displays a well-defined, intense doublet at 1021.9 and 1045.0 eV (Zn¹), consistent with the formation of ZnO.⁴⁹ In R2 and R3, the Zn_{2p} region shows the ZnO doublet similar to that observed in R1 but with gradually lower intensity. As shown in Fig. 3c, while no signal is obtained in the Al_{2p} region for R1, indicating the absence of Al in this region, R2 exhibits a well-defined signal deconvoluted into a doublet at 74.2 and 74.6 eV (Al¹), confirming the presence of Al₂O₃.⁵⁰ Similarly, R3 also shows an Al₂O₃ doublet with a slightly higher intensity than in R2. However, it is worth noting that, in R3, the binding energies of both the ZnO and Al₂O₃ doublets (Zn¹ and Al¹) in the Zn_{2p} and Al_{2p} regions are slightly shifted towards higher values (0.2–0.4 eV) (Table 1).

Correspondingly, the HR-XPS spectrum in the O_{1s} region for R1 exhibits a signal with two maxima (Fig. 3d), which is



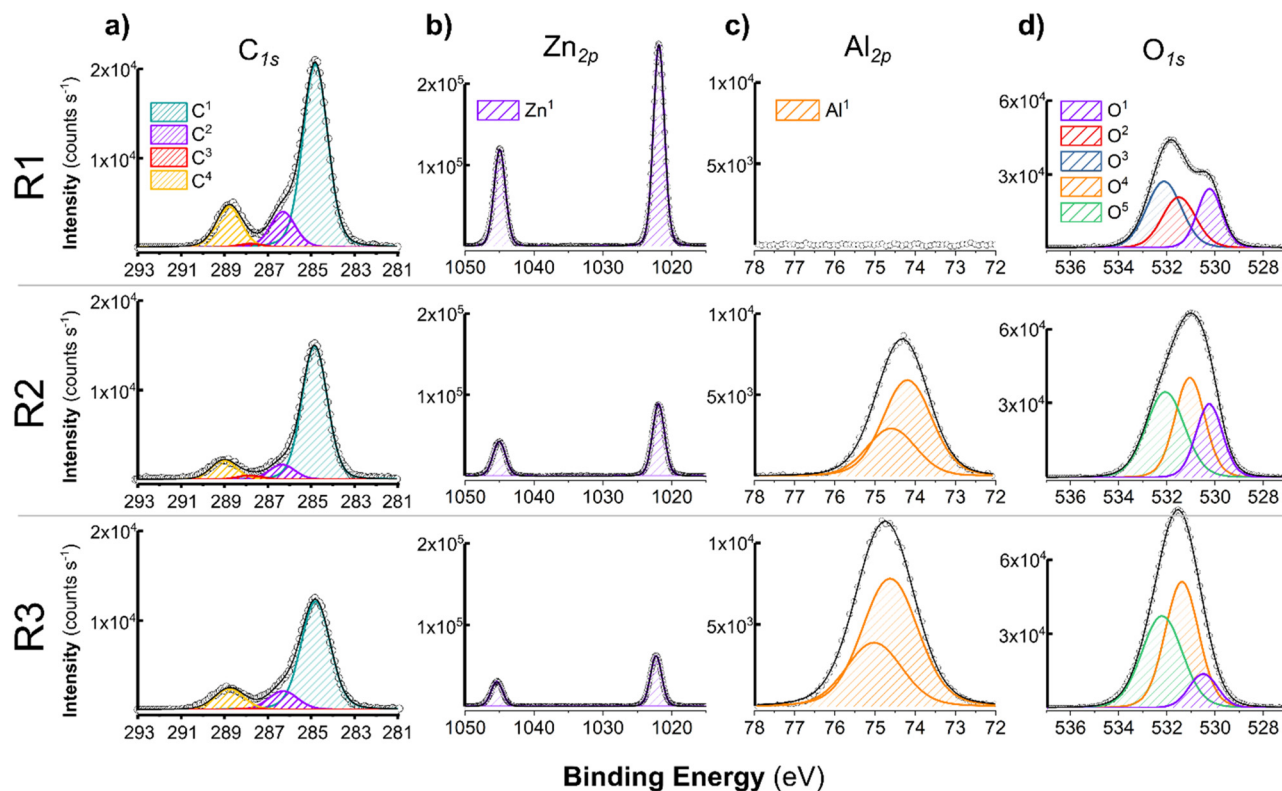


Fig. 3 HR-XPS spectra with fitted contributions; (a) C_{1s} , (b) Zn_{2p} , (c) Al_{2p} , and (d) O_{1s} regions obtained for SALD-deposited films. The black dots represent experimental data, while the solid black line corresponds to the fitted envelope.

Table 1 HR-XPS spectra fitting parameters and assignments in the C_{1s} , O_{1s} , Zn_{2p} , and Al_{2p} regions

Regions and components		Binding energy (FWHM/%Area) ^a			Assignment	Ref.
		R1	R2	R3		
C_{1s}	C^1	284.8 (1.4/69)	284.8 (1.4/78)	284.8 (1.6/73)	C-C, C-H	48 and 51
	C^2	286.3 (1.4/14)	286.3 (1.4/9)	286.3 (1.6/12)	C-OH, C-O-C	48 and 51
	C^3	287.8 (1.4/1)	287.8 (1.4/2)	287.8 (1.6/1)	C=O	48 and 51
	C^4	288.8 (1.4/16)	289.0 (1.4/11)	288.8 (1.6/14)	O-C=O	48 and 51
O_{1s}	O^1	530.2 (1.2/27)	530.2 (1.3/24)	530.5 (1.3/11)	ZnO lattice oxygen	49 and 52
	O^2	531.5 (1.6/31)	—	—	ZnO-H ₂ O; Zn-OH	49
	O^3	532.1 (1.7/42)	—	—	Physisorbed oxygen-containing species	49 and 52
	O^4	—	531.1 (1.5/36)	531.4 (1.5/47)	Al_2O_3 lattice oxygen	50 and 53
	O^5	—	532.1 (1.9/40)	532.2 (1.9/42)	Al-OH, physisorbed oxygen-containing species	50 and 53
Zn_{2p}	Zn^1	1021.9 (1.9/67)	1021.9 (2.1/67)	1022.3 (2.0/67)	ZnO	49
		1045.0 (2.0/33)	1045.1 (2.2/33)	1045.4 (2.0/33)		
Al_{2p}	Al^1	—	74.2 (1.5/67)	74.6 (1.6/67)	Al_2O_3	50
		—	74.6 (1.5/33)	75.0 (1.6/33)		

^a Binding energy and full width at half maximum (FWHM) values expressed in eV.

interpreted as three contributions at 530.2 eV (O^1), 531.5 eV (O^2), and 532.1 eV (O^3), attributed to ZnO lattice oxygen, chemisorbed water molecules on the ZnO surface (ZnO-H₂O) or hydrogen-passivated surface oxygen (Zn-OH), and physisorbed oxygen-containing species (primarily from adventitious carbon), respectively.⁵² In R2, as well as in R3, the O_{1s}

region displays increased intensity and a change in shape from two maxima to a single maximum. This O_{1s} signal is deconvoluted into three contributions at 530.2 eV (O^1), 531.1 eV (O^4), and 532.1 eV (O^5) for R2, whereas in R3, these peaks appear at slightly higher binding energies (530.5 eV, 531.4 eV, and 532.2 eV, respectively) with an overall increase in O_{1s}



intensity. Since ZnO and Al₂O₃ layers are detected in both regions, these contributions are attributed to ZnO lattice oxygen, Al₂O₃ lattice oxygen, and hydrogen-passivated surface oxygen on Al₂O₃ (Al–OH) or physisorbed oxygen-containing species, respectively.^{50,53}

Finally, based on the elemental quantitative analysis presented in Table S1† and the HR-XPS C_{1s} spectra analysis (Fig. 3a and Table 1), the contribution of oxygen from adventitious carbon in each sample can be estimated following the approach proposed by Henderson *et al.*⁵¹ The values obtained using this method exhibit excellent agreement with those derived from the difference between the total atomic percentage of oxygen and the fraction attributed to oxides, assuming the presence of ZnO and Al₂O₃ (see in ESI†). Moreover, the atomic percentages of Zn and Al, together with the oxygen contribution from adventitious carbon, are in good concordance with the area percentages determined for each component in the HR-XPS O_{1s} spectra (Table 1).

Thickness measurements and CFD simulations

Fig. 4a presents the nanolaminate film image, which exhibits a changing color due to the increased film thickness. The film thickness was measured using ellipsometry at 10 distinct points along a 5 mm distance. In Fig. 4b, the ellipsometry measurements reveal that the film on the ZnO side begins at 38.6 nm and increases to 99.7 nm on the multilayer side. By assuming that the ZnO layer thickness remains constant, the Al₂O₃ component is observed to present a film thickness of up to 60 nm over a 50 mm spatial distance. This observation

corresponds to growth per cycle (GPC) values of 0.125 nm for ZnO and 0.187 nm for Al₂O₃. However, when individual depositions were performed for ZnO and Al₂O₃ films, the GPC values were determined to be 0.128 nm for ZnO and 0.155 nm for Al₂O₃. These findings are in agreement with the expected values reported in previous studies for ZnO^{54,55} and Al₂O₃.^{56,57} Moreover, a completely linear increase in film thickness was not observed along the combinatorial sample, which suggests that the deposition rate of Al₂O₃ may not exhibit a linear increase along the substrate, as has been the case in previous studies.³³ CFD simulations were conducted to ensure the obtained gradient is determined by the combinatorial head design used.

Fig. 4b presents the CFD simulations of the combinatorial metal channel. These simulations enable a comprehensive understanding of the mass transportation phenomena through the analysis of the velocity, concentration, and pressure distributions on the substrate surface. The velocity profile begins to increase in the 15 mm region, reaching a maximum value of approximately $30 \times 10^{-4} \text{ m s}^{-1}$ between the 30- and 50 mm regions. The concentration of TMA within the combinatorial channel increases to a maximum of $34 \times 10^{-7} \text{ mol m}^{-3}$, with the highest concentration predominantly observed between 30 and 50 mm, which suggests variations in precursor availability across the substrate surface. In the pressure profile, the highest values, around 3.6 atm, are obtained between the same 30–50 mm region, whereas the minimum pressure, approximately 1 atm, is associated with very low concentration values. Overall, the region from 0 to

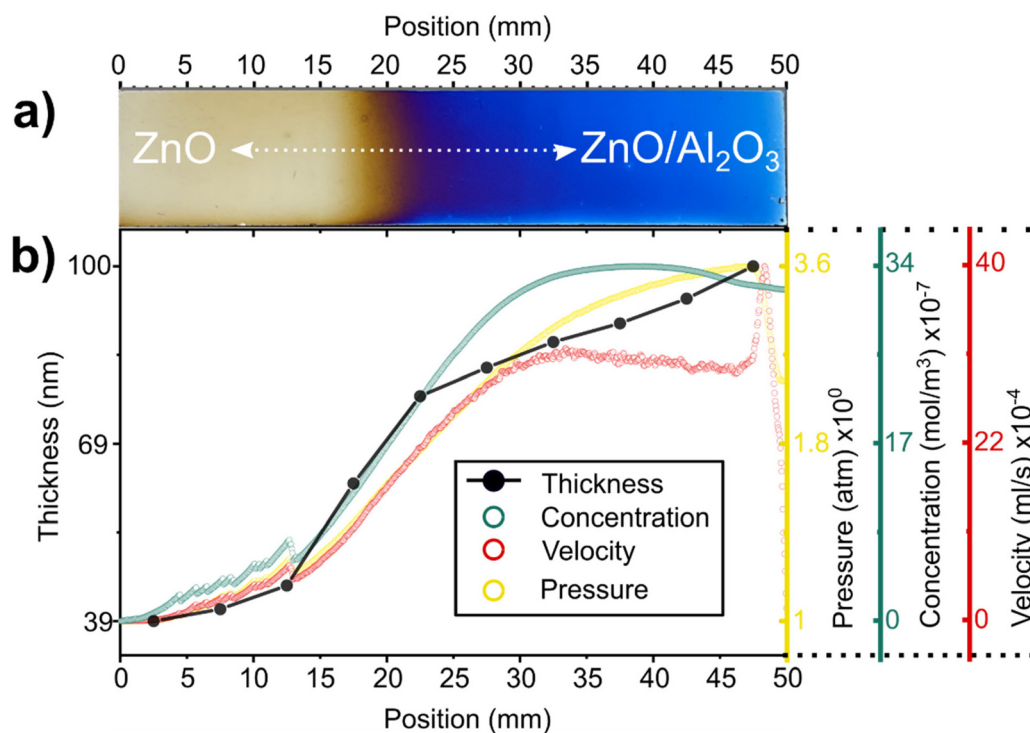


Fig. 4 (a) The image of combinatorial nanolaminate film ZnO/Al₂O₃. (b) Ellipsometry thickness measurements (solid dots) and CFD simulation averaged profiles (hollow dots) on the substrate surface sequentially; velocity, concentration and pressure profiles, as a function of sample position.



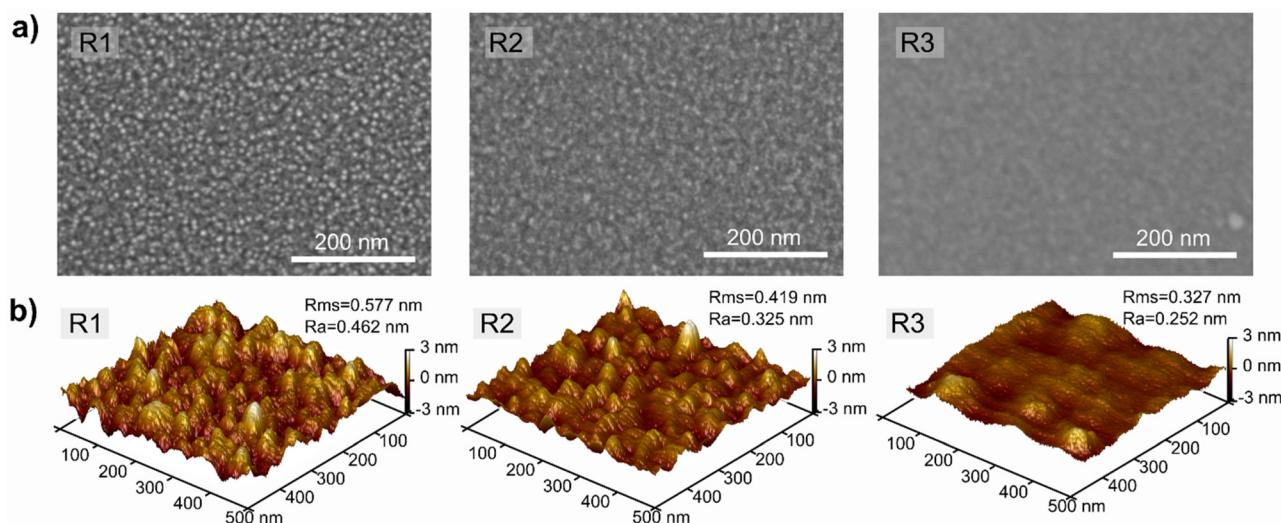


Fig. 5 (a) Top-view SEM images showing the surface morphology of nanolaminate ZnO/Al₂O₃ film from different regions R1, R2, and R3. (b) AFM images displaying the surface roughness over 500 × 500 nm² areas of the same regions. Both RMS and *R_a* roughness values are included.

15 mm exhibits low mass transportation, the region from 15 to 30 mm displays a significant increase in values, and the region from 30 to 50 mm forms a plateau by the highest values; this trend correlates closely with the measured film thickness values and with have been obtained previously.³³

Film morphology and structure

The scanning electron microscopy results in Fig. 5a clearly reveal significant variations in the film surface morphology as the thickness of the Al₂O₃ layer increases. In R1, where the ZnO film is deposited without an Al₂O₃ interlayer, distinct ZnO nanostructures are observed. This textured appearance is predominantly indicative of *c*-axis oriented ZnO nanostructures.⁵⁸ By increasing the Al₂O₃ thickness from R1 to R3, the film morphology exhibits a noticeably softer contour.

Following the observed morphological changes in SEM results, a topographic AFM characterization was performed to obtain quantitative roughness information. The nanolaminate sample presented more defined vertically aligned nanostructures in R1, and a progressively smoother surface was observed in R2 and R3 as the thickness of the Al₂O₃ layer increased. The roughness information for R1, R2, and R3 was obtained by AFM from 500 nm × 500 nm area scan and is plotted in Fig. 5b. The distinct vertical ZnO grains, which are clearly discernible in R1, become gradually broader in R2 and R3. From these measurements, the arithmetic average (*R_a*) and root mean square (RMS) of the profile height variations were derived, yielding RMS [*R_a*] values of 0.577 [0.462] nm, 0.419 [0.325] nm, and 0.327 [0.252] nm for R1, R2, and R3, respectively, which indicate a gradual decrease in surface roughness.

GI-XRD analysis was employed to elucidate the crystallinity of the films. Gaussian fitting of the (002) peaks yielded full widths at half maximum of 0.62°, 1.49°, and 1.42° for R1, R2, and R3, respectively. Furthermore, a subtle yet distinct sharp

Si (311) reflection was observed at 56.2° in R2 and R3, as denoted by an asterisk in Fig. 6. The calculated penetration depth at a 0.34° incidence angle revealed a significant disparity between the ZnO and Al₂O₃ materials due to their differing absorption coefficients; specifically, the penetration depths were determined to be 57.47 nm for ZnO and 263.84 nm for Al₂O₃, as illustrated in Fig. S1.†

TEM analysis provides a distinct formation of multilayers in regions R2 and R3. Fig. 7a shows cross-sectional images of

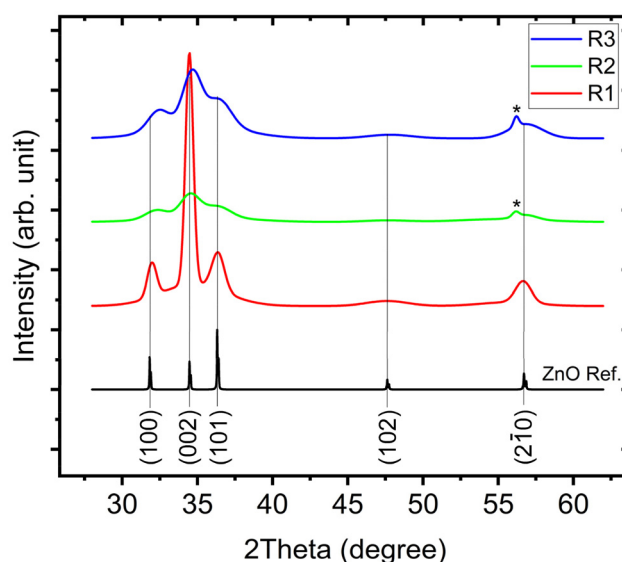


Fig. 6 GI-XRD patterns of the thin film sample at incident angles of 0.34°. The diffraction peaks correspond to the crystalline phases present in the surface layer, with a notable decrease in the (002) peak of ZnO by increasing Al₂O₃ composition, indicating a reduction of crystalline quality. PDF Card – 04-020-0364 used as a reference ZnO spectra the intensity multiplied fifty times.



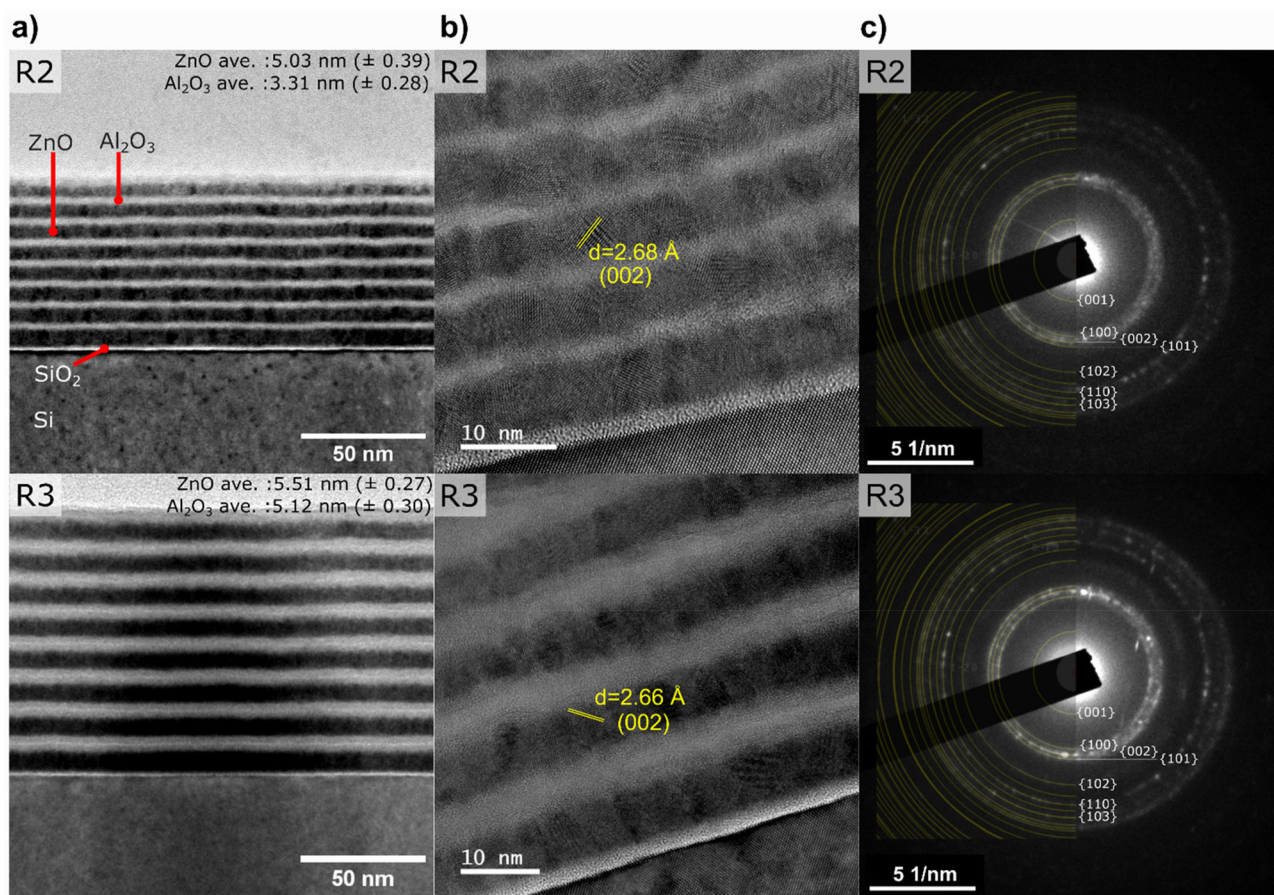


Fig. 7 TEM analysis of ZnO/Al₂O₃ nanolaminate films. (a) TEM cross-sectional image shows the ZnO/Al₂O₃ bilayer deposited with eight super-cycle. ZnO and Al₂O₃ sublayers have been annotated (b) HR-TEM image showing lattice fringes with a measured spacing of 2.7 Å for ZnO film. (c) SAED pattern indexed to the ZnO (PDF Card – 04-020-0364).

ZnO/Al₂O₃ nanolaminate films with varying Al₂O₃ layer thicknesses in both regions. A detailed image analysis was conducted to measure the layer thicknesses: in R2, the average ZnO and Al₂O₃ layer thicknesses were determined to be 5.03 nm and 3.31 nm, respectively. In contrast, in R3, the corresponding values were 5.51 nm and 5.12 nm, with a measurement deviation of approximately 0.3 nm in both regions. Additionally, in both regions, the first ZnO layer is slightly thicker (~6 nm) compared to subsequent ZnO layers, which are consistently 5 nm. The Al₂O₃ layer thickness is controlled at 3 nm in R2 and 5 nm in R3. High-resolution TEM images in Fig. 7b reveal that ZnO exhibits a polycrystalline nature with distinct lattice parameters, whereas Al₂O₃ forms an amorphous structure. Within the ZnO layers, apparent grain boundary defects are observed, with ZnO grain size limited to 5 nm due to interruption by the Al₂O₃ layers. Furthermore, while no diffusion from the Al₂O₃ layer into the well-crystalline ZnO grains is observed, diffusion at the grain boundaries is evident. The lattice distance of the (002) planes is slightly higher than the unstressed ZnO lattice distance (2.6 Å), and the (002) plane orientations appear random, aligning with the XRD results for both R2 and R3.

In Fig. 7c, the SAED pattern of the ZnO/Al₂O₃ film is presented; the Debye–Scherrer rings comprise a combination of rings and spots that indicate the polycrystalline nature of ZnO, and region R3 exhibits slightly more intense spots, which might be attributable to larger ZnO grains in the local SAED characterization zone. The SAED pattern matches well with the ZnO reference (PDF Card – 04-020-0364), and the indexed diffraction patterns indicate stronger reflections from the (100), (002), and (101) planes. In general, the TEM and SAED results collectively demonstrate the properties of well-crystalline ZnO and amorphous Al₂O₃ layers, and the cross-sectional images illustrate highly reproducible layer thicknesses and sharp interfaces along the multilayer structure, thereby highlighting the consistency of super-cycle deposition with multi-head SALD even in open-air conditions.

Surface potential properties

In ZnO/Al₂O₃ nanolaminate films, changing the sublayer thickness enables the tuning of the work function of the film.¹¹ Fig. 8 demonstrates the variations in the work function with increasing Al₂O₃ thickness using a Kelvin probe setup, where the work function values are plotted as a function of measure-



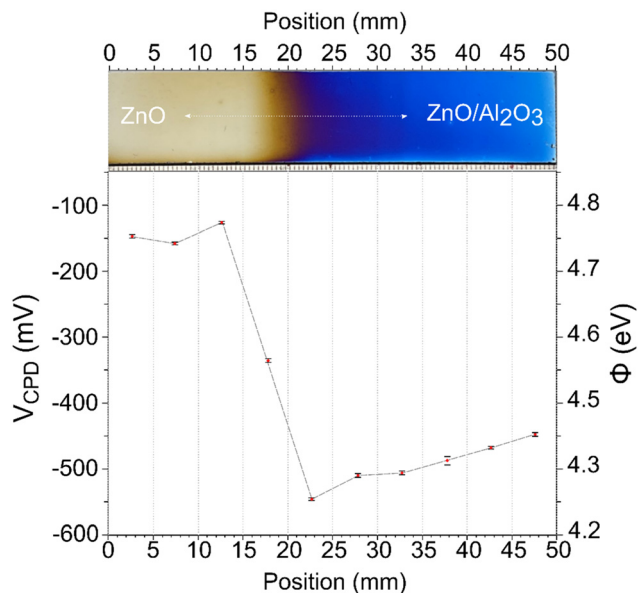


Fig. 8 The contact potential difference, represented on the left y-axis, and the calculated work function values, displayed on the right y-axis, were systematically measured at various points on the investigated samples.

ment location in our gradient film. The Kelvin probe presented a surface potential of 197.55 ± 8.92 mV on a gold (Au) reference sample, which is an expected value.⁵⁹ The calculated work functions (ϕ) relative to the Au reference ($\phi_{\text{Au}} = 5.1$ eV) are presented on the right y-axis. The first three data points (0–15 mm) are relatively high contact potential difference (V_{CPD}) values (-147.6 ± 2.1 mV), corresponding to an average work function of 4.75 ± 0.02 eV. The initial points likely represent regions where the Al_2O_3 coverage is still minimal; in these regions, the surface is dominated by ZnO or is partially covered by Al_2O_3 . The slight variation in these values may be attributed to local doping changes or the initial onset of Al_2O_3 coverage. Between 15–25 mm, a sharp decrease in surface potential to -546.8 ± 2.3 mV ($\phi = 4.35$ eV) is observed. This change likely indicates that sufficient Al_2O_3 is present to alter the near-surface electronic structure, for instance, through band bending in ZnO, passivation of surface states, or screening by the Al_2O_3 .⁶⁰ Above 35 mm, after reaching a threshold coverage of Al_2O_3 , the measured work function begins to recover gradually toward positive values up to a value of -448.0 ± 2.6 mV resulting in $\phi = 4.46$ eV.

Discussion

Spatial film thickness variation and mass transport dynamics

When GPC values were calculated using ellipsometry for nanolaminates and compared with those of reference ZnO and Al_2O_3 films, the ZnO thickness matched expectations. The Al_2O_3 thickness, however, showed discrepancies. The Al_2O_3 films in the nanolaminate had a GPC of 0.187 nm, as opposed

to 1.55 nm per cycle for individual Al_2O_3 films. Consequently, rather than attributing these differences to variations in the ZnO layer thickness, it is plausible that the change arises from the utilization of a four-wave ellipsometry model that was not specifically designed for alternating ZnO and Al_2O_3 layers; the differing refractive indices for each layer may have introduced discrepancies when measuring the bulk nanolaminate thickness. For more precise bulk thickness measurements in nanolaminates, it is recommended that spectroscopic ellipsometry (SE) or X-ray reflectivity (XRR) techniques, combined with rigorous modeling and fitting procedures, be employed as corroborated by other studies.⁶¹ Furthermore, TEM results indicated that the initial ZnO layer was slightly thicker (by less than 1 nm) compared to subsequent ZnO layers. Although this observation might initially be interpreted as a difference in GPC of ZnO on the native SiO_2 layer, it should be noted that at the start of the deposition process the substrate passed under the injection head when returning to the starting position. Thereby undergoing an extra ALD cycle that resulted in a marginally thicker initial ZnO layer. An alternative perspective considers the conversion-etch effect of the TMA precursor on the ZnO film, a phenomenon previously reported by Zywojko *et al.*⁶² for atomic layer etching (ALE) of ZnO. This may also explain the observed differences in Al_2O_3 GPC for individual (0.187 nm) and bulk nanolaminate (0.155 nm). Nevertheless, the remainder of the layers exhibited reproducible thickness with sharply defined interfaces, a characteristic attributed to the inherent nature of the ALD mechanism. Notably, although previous trials with mono-head SALD have been conducted,^{41,63} achieving interfaces of such sharpness under open-air conditions has not been previously reported. In the present study, the application of SALD under open-air conditions *via* a super-cycle approach has yielded interfaces of exceptional sharpness.

Additionally, our CFD simulations underscore a strong correlation between film thickness and the amount of TMA molecule transported for a particular head design. The initial 15–30 mm region appears to exhibit a sharp increase in TMA mass transport, as indicated by the concentration, velocity, and pressure results in Fig. 4b. This behavior strongly correlates with the final film thickness gradient, which manifests as a smooth, step-like transition rather than a linear one. Furthermore, achieving a more controlled thickness gradient necessitates a comprehensive understanding of the interplay between combinatorial head geometry, deposition parameters, and precursor chemistry, which would constitute a significant advancement in the design of more efficient deposition heads for combinatorial film depositions using SALD. Such an investigation may require the development of a mathematical model supported by multiphysics simulations and experimental data, which is beyond the primary focus of the present article.

Structural and morphological properties

In addition, ZnO demonstrates high crystallinity and strong *c*-axis orientation on a native silicon oxide layer even at a



40 nm thickness. However, an increase in the thickness of the amorphous Al_2O_3 layer diminishes the bulk crystalline properties of the nanolaminate, as evidenced by a ZnO (002) diffraction peak that remains apparent but becomes weaker and broader. When the FWHM of the diffraction peaks are calculated, a substantial difference is observed between region R1 and the other regions, whereas the difference between regions R2 and R3 is minimal. This phenomenon can be attributed to the interruption of ZnO growth by each 5 nm increment of the amorphous Al_2O_3 layer, which confines the ZnO crystallite size and results in broader diffraction peaks; however, further increases in Al_2O_3 thickness do not contribute additional broadness to the nanolaminate film. The results further indicate that the crystallinity of the ZnO/ Al_2O_3 system is directly related to the number of ZnO ALD cycles, thereby permitting tunability in crystallinity. Moreover, TEM results (Fig. 7) also indicate that an increased amorphous Al_2O_3 layer does not affect the crystallinity of the individual ZnO layers. Notably, silicon peaks appear exclusively in regions R2 and R3, which is attributed to the increased penetration depth of the Al_2O_3 layer; calculated values in Fig. S1† show a significant disparity in penetration depths between ZnO (57 nm) and Al_2O_3 (264 nm) under identical X-ray incidence conditions, a factor that may contribute to the detection of increased silicon reflection intensities. In contrast, in region R1, the lower ZnO penetration depth may lead to only a minor contribution to the (210) ZnO reflection, a contribution that remains indistinguishable at the resolution of current X-ray diffraction measurements, which results as a broader peak including both ZnO and Si components.

Based on SEM and AFM results, the thickness of the Al_2O_3 layer provides considerable tunability in controlling surface roughness. Given that the preferentially *c*-axis oriented ZnO (R1) exhibits high roughness even at a thickness of 40 nm, lamination with Al_2O_3 (R2 and R3) effectively reduces surface protrusions and smooths the film. These smoother surfaces decrease photon scattering, thereby resulting in higher transmittance and lower absorbance,⁶⁴ which may be advantageous for applications that require highly uniform thin films, such as transparent conducting electrodes or heterostructures for electronic devices. In contrast, rougher surfaces scatter and trap more photons, which increases the likelihood of photon absorption and the generation of electron-hole pairs that facilitate the photocatalytic process.⁶⁵ Additionally, rough surfaces provide more active sites, which improves the molecular adhesion of the photocatalytic adsorbates, it is important to note that excessively rough surfaces can cause coarsening of adsorbates, thereby reducing adsorption efficiency,⁶⁵ which suggests that optimization of surface roughness is crucial. Therefore, by employing a multi-head SALD system to tailor the Al_2O_3 sublayer thickness, one can fine-tune surface roughness to balance high optical transmittance with enhanced photocatalytic activity, thereby optimizing thin film performance across diverse applications.

Surface chemistry and work function

Different from XPS, the EDS analysis revealed a small amount of Al content (0.58 at%), and it has been previously noted that surface roughness plays an important role in the error rates in EDS.⁶⁶ The roughest morphology observed in R1 might lead to increased error rates due to the EDS scattering mechanism. However, it has been clearly observed that the chemical composition ratio of Al increases with increasing Al_2O_3 layer thickness, although the Zn content does not exhibit the same atomic percentage. In EDS, the absorption mechanism of nanolaminates with varying thickness greatly alters the measured chemical composition, as demonstrated by the PRZ distribution obtained from MC simulation (Fig. S2†), which shows that increased nanolaminate thickness significantly impacts the total X-ray absorption of Si, while Zn maintains consistent absorption values across all regions. This effect is attributed to the increased bulk nanolaminate thickness in R2 and R3, where the penetration depth gradually decreases, thereby reducing Si information and resulting in an overall increased signal for the other elements; thus, XPS results may be more convenient for quantitative analysis when only shallow information depth is required.

In XPS results, the Zn intensity gradually decreases with increasing Al_2O_3 due to attenuation by the upper Al_2O_3 layer (Fig. 3b and Fig. 7a). In R2 and R3, the Al_{2p} binding energies exhibit slight shifts that are attributed to the higher Al_2O_3 content. As an insulating material, Al_2O_3 causes the applied flood gun conditions to be insufficient in fully compensating for surface charging during sequential HR-XPS spectrum acquisition, leading to differential shifts in the adventitious carbon peak relative to other elements in the same sample.⁶⁷ In contrast to the EDS results, the HRXPS results show no trace of Al contribution in R1. Furthermore, the gradual increase of Al content in R2 and R3 proves the reliability of the SALD method for precise chemical composition control and the avoidance of unwanted mixing. The O_{1s} peak in R2 and R3 presents a different shape and intensity compared to R1, which is due to the increased Al_2O_3 layer thickness contributing significantly to the Al_2O_3 lattice oxygen (O^4); furthermore, in the O_{1s} spectra of R2 and R3, the binding energies corresponding to ZnO– H_2O and Zn–OH bonds have disappeared owing to the new top layer of Al_2O_3 . In the XPS results, R1 shows the highest C_{1s} atomic percentage (Table S1†) compared to R2 and R3, a phenomenon that can be attributed to the surface roughness since rougher surfaces have a larger effective surface area available for the adsorption of hydrocarbons and carbonaceous contaminants.⁶⁸ A relatively higher O atomic percentage could be due to water and hydroxyls adsorbed on the surface, which often introduce extra oxygen signals; when the surface roughness decreases in R2 and R3, the stoichiometry improves, although the ratios remain O-rich.

Film stoichiometry is a critical factor in open-air deposition methods. Previous studies have reported that open-air deposition techniques can incorporate more oxygen into deposited films compared to vacuum-based methods due to the presence



of atmospheric oxygen in the deposition environment.⁶⁹ Nyugen *et al.* demonstrated the impact of open-air SALD on the carrier mobility of Al-doped ZnO films. Typically, stoichiometric calculations involve sputtering the surface carbon layer with argon gas. However, due to the presence of a thin 3 nm Al₂O₃ top layer, Ar sputtering is avoided to prevent potential damage to this layer. Instead, the model proposed by Henderson *et al.*⁵¹ has been employed to carefully account for the contribution of carbon species and extract calculated stoichiometric information. As shown in Table S2,† the Zn:O ratio at the surface is approximately 1:1, indicating high phase purity and stable material composition. Furthermore, despite being an open-air deposition method, our study determined that the overall metal-to-oxide ratio closely aligns with the expected stoichiometric values across all examined regions (Table S2†).

The variations in surface potential correlated to the varying composition of the investigated surfaces. The hydroxylated intrinsic ZnO surface attributes the initial high values. This yields a higher V_{CPD} compared to the reference Au sample, which is consistent with previous reports.⁷⁰ This suggests that local doping variations are minimal and that the Al₂O₃ coverage is slight. As a highly surface-sensitive method, the Kelvin probe method focuses on the outermost atomic layers about 3 nm.⁷¹ With partial Al₂O₃ coverage on the ZnO surface, the measured V_{CPD} and ϕ progressively shift to lower values. This is in agreement with the characteristic of Al and Al₂O₃.⁷² It also relates to a potential barrier forming at the interface between the sub-layers. Al₂O₃ is well known to passivate surface states on ZnO and other semiconductors, thereby modifying the Fermi-level position at or near the surface.⁷³ A thin layer of Al₂O₃ helps cover the ZnO surface. This reduces the dangling bonds and lowers defect-induced Fermi-level pinning. As a result, the surface potential is minimized. Schuldis *et al.*⁷⁴ demonstrated that a 0.5 nm-thick Al₂O₃ layer provides optimal surface passivation. Our findings support this claim, the region where a thin Al₂O₃ layer leads to a significant accumulation of charges at the ZnO/Al₂O₃ interface. As the Al₂O₃ layer gets thicker, the insulating surface potential weakens the net effect, which raises the work function. Sufficient Al₂O₃ coverage can enhance the electron affinity effect and reduce the measured work function. Thin ZnO layers (under 1 nm) can show the Maxwell-Wagner polarization effect. However, the 5 nm ZnO thickness in this study does not present such significant charge buildup from this effect.⁹

Conclusion

The important contribution of this study is the demonstration of a novel approach to control the usage of multiple precursors in an open-air SALD method, which allows for exceptional dosing control. This multi-head system permits the use of optimized head designs and pulse times (substrate speed) for each material deposition, a feature that is not fea-

ible with conventional mono-head SALD systems operating in a super-cycle approach. Furthermore, the multi-head setup allows the implementation of different custom head designs. In the present study, a combinatorial head was employed to demonstrate the impact of Al₂O₃ sublayer thickness on ZnO/Al₂O₃ nanolaminate films. The use of separated heads prevents issues such as premature surface saturation and competition between different precursors, resulting in more stable growth kinetics and improved stoichiometric control. Consequently, the multi-head SALD method yielded no evidence of unwanted aluminum contamination in regions where TMA was not injected. Cross-sectional TEM images confirmed that the nanolaminate film exhibited high reproducibility in each sublayer and well-defined, sharp interfaces. Furthermore, XPS results revealed no trace of alloying and a favorable elemental ratio, thereby affirming that the key merits of ALD can be realized in open-air using the demonstrated approach. Finally, work function analyses confirm that partial Al₂O₃ coverage passivates ZnO surfaces, leading to controlled electronic properties, underscoring the promise of SALD for functional thin-film fabrication in a precise manner.

While the paper presents comprehensive results on combinatorial ZnO/Al₂O₃ nanolaminates, the multi-head SALD technique shows promising potential for controlled doping and optimization of various multi-material depositions. Precise control of this super-cycle approach, coupled with the incorporation of machine learning into the combinatorial SALD approach, holds significant potential as a rapid and effective technique for experimental studies in materials research. Moreover, combining this multi-head method with surface treatments, such as plasma, ozone, or UV, and *in situ* characterization techniques offers numerous opportunities for further advancements. These advanced approaches collectively establish SALD as an adaptable deposition method with broad potential for adoption in both scientific research and industrial applications.

Author contributions

H. O. undertook the conceptualization, combinatorial head design, material deposition, conducted data curation (with the exception of TEM and XPS), formal analysis, visualization and writing – original draft. M. I. B. conducted XPS formal analysis, validation and visualization. V. S. P. R. performed CFD data curation, and P. J. V. R. contributed to conceptualization and formal analysis of Kelvin probe. G. A. contributed to funding acquisition, while D. M. R. contributed to conceptualization and funding acquisition. All authors contributed to the review and editing of the manuscript.

Conflicts of interest

The authors declare no competing financial interest.



Data availability

The data supporting this article have been included as part of the ESI.†

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