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## ARTICLE

# Copper Nanoparticles Exsolution from $\text{Sr}(\text{Ti, Fe})\text{O}_3$ Perovskites: Material Tuning and Probing (Electro)catalytic Applicability

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Copper (Cu) is a recyclable, abundant, and promising catalyst for energy transition reactions like electrochemical conversion of nitrate ( $\text{NO}_3\text{RR}$ ) and  $\text{CO}_2$  electroreduction. However, conventional Cu-based electrocatalysts struggle with activity, selectivity, and durability, especially under harsh electrochemical conditions. Exsolution—the in-situ generation of metallic nanoparticles on oxide supports in a single step—enables tightly anchored, size-controlled particles, enhancing stability and performance. Incorporating Cu into  $\text{Sr}_{1-a}(\text{Ti, Fe})\text{O}_{3-y}$  perovskites, an earth-abundant system with promising ionic-electronic conductivity and adequate oxygen vacancies, overcomes the limitations of traditional  $\text{Sr}(\text{Ti, Fe})\text{O}_3$  in facilitating nanoparticle exsolution. This work demonstrates controlled Cu nanoparticle exsolution from  $\text{Sr}_{0.95}\text{Ti}_{0.5}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$  perovskites at temperatures as low as 400°C, notably milder than conventional exsolution conditions. By systematically varying reduction parameters, we achieve control over nanoparticle size (13–38 nm) and population density (118–650 particles/ $\mu\text{m}^2$ ). Electrochemical characterisation using nitrate reduction as a probe reaction demonstrates how exsolution conditions directly influence surface reactivity, establishing these materials as tuneable platforms for (electro)catalytic applications.

## Introduction

Developing sustainable technologies to ensure efficient energy conversion and storage is crucial to address the challenges associated with the long-term reliance on fossil fuels. In this context, copper-based catalysts have garnered significant attention due to their abundance, recyclability, and versatility in various applications. These include electrochemical nitrate reduction ( $\text{NO}_3\text{RR}$ ) for water remediation and ammonia production, hydrogen evolution reaction for green hydrogen production, and  $\text{CO}_2$  reduction for value-added molecules and platform chemicals, pivotal for energy transition and environmental sustainability.(1–6) However, conventional Cu-based catalysts often lack activity, selectivity, and durability, particularly under the harsh conditions of electrochemical and thermochemical reactions. Innovative strategies that enhance catalyst stability, tunability, and performance are essential to overcome these limitations.

One promising approach involves incorporating copper (ions) into perovskite oxides, facilitating the exsolution of well-

dispersed Cu (metal) nanoparticles at the surface.(7) This technique allows for precise control over catalytic activity.

It addresses the challenges related to particle agglomeration and detachment during operation, paving the way for more robust and efficient catalytic systems. Moreover, the rational design of these catalysts is achievable by controlling key exsolution parameters such as temperature, time, and atmosphere, enabling the fine-tuning of nanoparticle size, shape, and surface distribution to meet specific application requirements.(8)

Perovskite oxides are materials with the  $\text{ABO}_3$  structure, where A and B represent cations while O represents an anion. This material has been used for several applications, mainly due to its tuneable structures, robust stability, and flexible doping capabilities, which allow fine-tuning of electronic and structural properties to optimise catalytic performance.(6) Achieving Cu exsolution—a process where Cu ions segregate from the perovskite lattice to the surface to form well-anchored metallic nanoparticles—addresses common issues in catalysis by creating uniformly distributed stable nanoparticles that resist agglomeration and detachment. Adjusting reduction conditions, such as temperature and duration, allows the nanoparticles' size, shape, and density to be systematically controlled, offering significant advantages over traditional catalyst preparation methods.(9–12)

The success of enhanced exsolution is tied to selecting perovskites with desirable properties such as ionic and electronic conductivity, thermochemical stability, and the ability to form oxygen vacancies, all of which are critical for (electro)catalytic applications.(13–15) Among non-rare-metal-based perovskites, STF oxides have shown

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particular promise. They offer mixed ionic and electronic conductivity along with suitable electrochemical stability.(16–18) In particular, the  $\text{Sr}_{1-x}\text{Ti}_{0.7}\text{Fe}_y\text{O}_3$  system has demonstrated chemical and structural robustness and adaptability for stoichiometric variation, allowing targeted doping with catalytically active elements like Cu.(19–22)

This study explores the rational design of  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$  by leveraging A-site deficiency and Cu-doping to achieve controlled exsolution of Cu nanoparticles at relatively low temperatures. This approach enabled the formation of well-dispersed Cu nanoparticles at reduction temperatures as low as 400 °C within 1 hour. By systematically varying the exsolution temperature and time, we demonstrate how these parameters influence nanoparticle properties and catalytic activity, providing a tunable platform for tailored applications. As a proof-of-concept and surface characterisation tool indicative of electrochemical application suitability, we applied the materials to  $\text{NO}_3\text{RR}$ , illustrating how the rational design of exsolution conditions can tune the electrochemical properties of the catalyst.

## Result and Discussion

### Perovskite System Design and Synthesis Conditions

The  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$  series of perovskites with varying Cu-doping ( $x = 0, 0.05, 0.1, 0.15, 0.2$  mol) were designed to be A-site deficient, resulting in an A-to-B site ratio below 1, which promotes exsolution capability, as reported previously.(23,24) The Cu doping range ( $x = 0\text{--}0.2$ ) was selected to explore the limits of Cu substitution in this perovskite system. While higher Cu content would provide more potentially active sites, it would also introduce excessive oxygen vacancies due to charge compensation with host cations ( $\text{Ti}^{4+}$  and  $\text{Fe}^{3+}/^{4+}$ ) and, given Cu's position as the most reducible first-row transition metal ion, would make the overall lattice more reducible.(25) This could lead to uncontrolled structural decomposition rather than the desired controlled exsolution of nanoparticles. **Table 1** presents the stoichiometry of the respective perovskites.

The perovskites were synthesised using the solid-state synthesis method, with the key steps shown in **Fig. 1(a)**. This involved calcinating dried precursor mixtures at 1000 °C, pellet formation and sintering in the air at 1000–1100 °C. The synthesis temperature was selected based on the thermal behaviour of the essential precursors— $\text{Fe}(\text{NO}_3)_3\cdot 9\text{H}_2\text{O}$ ,  $\text{SrCO}_3$ , and  $\text{TiO}_2$ , as well as the dry mixture of these precursors, which was investigated using thermogravimetry analysis (TGA) and differential scanning calorimetry (DSC) (**Fig. 1(b)**).

Dehydration and nitrate decomposition resulted in a sharp mass loss for  $\text{Fe}(\text{NO}_3)_3\cdot 9\text{H}_2\text{O}$  at 200 °C, whereas the observed mass loss around 800 °C corresponded to carbonate decomposition for  $\text{SrCO}_3$ . The thermal analysis reveals a multi-step transformation process of the precursor material mixture up to 1000 °C, characterised by three main events: an initial dehydration around 200 °C, gradual

decomposition between 300 °C and 600 °C, and a final crystallisation event at approximately 900 °C. The DSC curve and mass loss data indicate that the material undergoes complete transformation by 900 °C, resulting in a thermally stable product, which likely represents the formation of the incipient perovskite phase. Hence, the selection of 1000 °C as the temperature at which the precursors' decomposition products would likely undergo the solid-state reaction to form the perovskite phase.

XRD analysis (**Fig. 1(c)**) revealed that the primary perovskite phase was successfully formed, accompanied by some secondary phases, including  $\text{TiO}_2$  and  $\text{CuO}$ , after calcination. However, sintering at 1000–1100 °C reduced the  $\text{TiO}_2$  secondary phase, resulting in a perovskite with a minimal  $\text{CuO}$  secondary phase and an improved structural composition.

### Crystal Structure and Composition of the $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$ Systems

The structural investigation revealed that all  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$  perovskite compositions exhibited a cubic structure (space group:  $\text{Pm}\bar{3}\text{m}$ ). This result (**Fig. 2(a-d)**) shows that the addition of Cu-doping up to  $x = 0.2$  does not alter the cubic structure of  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$ , despite the presence of a  $\text{CuO}$  secondary phase, indicating incomplete incorporation of the Cu dopant.

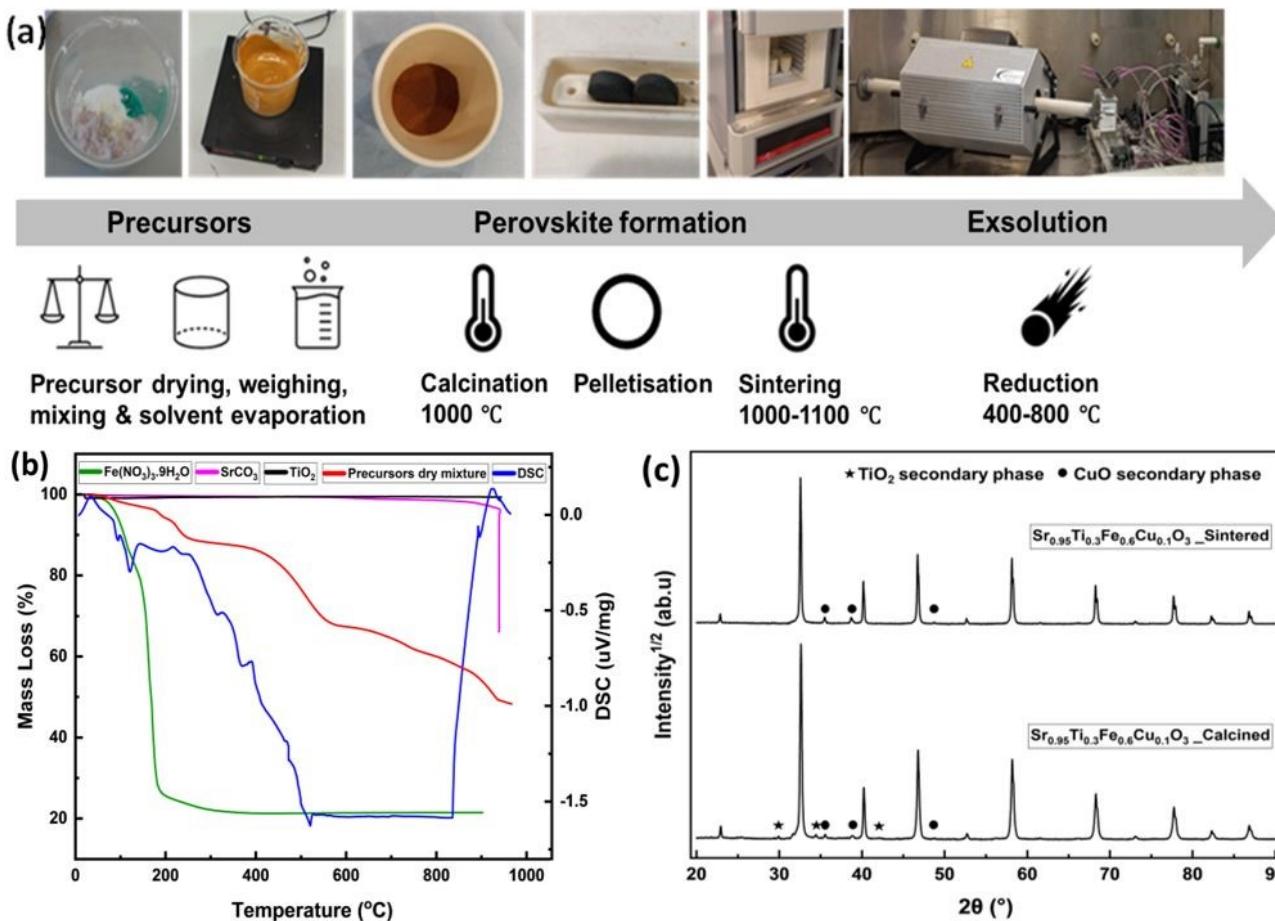
The Rietveld refinement of the sintered  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$  ( $x: 0\text{--}0.1$ ) perovskite structure (**Fig. 2(c-d)** and **Table 2**) demonstrates good agreement between the observed and calculated patterns, confirming the successful synthesis of the target perovskite phase. The cell parameter 'a' shows a slight overall decreasing trend from 3.8860 Å to 3.8788 Å with increasing Cu content (x), while the crystallite size initially increases sharply from 0.11 µm in the undoped sample to ~0.58–0.63 µm upon Cu doping but then remains relatively stable. Notably, the actual Cu incorporation ( $x^*$ ) is significantly lower than the nominal content at low doping levels. Still, it approaches the nominal value as the Cu content increases, suggesting a limit to Cu substitution in the structure at lower doping levels. The actual Cu-doping amount and resulting perovskites' stoichiometry (**Table 2**) were determined based on the  $\text{CuO}$  secondary phase weight fractions in the perovskite after synthesis, as explained in Note S1. The data show that forming the  $\text{CuO}$  secondary phase limited the Cu-doping incorporated into the perovskite structure, altering the designed stoichiometry and the envisaged A-site deficiency. Hence, we suspected the perovskite's exsolution capability would be negatively impacted. However, incorporating Cu into the perovskite lattice is expected to introduce additional oxygen vacancies, as reported for related perovskite structures.(26,27) At the same time, Cu ions are among the most reducible in the first-row transition metals series (most negative Gibbs energy for the reduction), which would represent a significant driving force for exsolution to occur.(28,29) Therefore, as a net result, the deviation from the intended A-site deficiency would likely not significantly limit the perovskite's exsolution capability.



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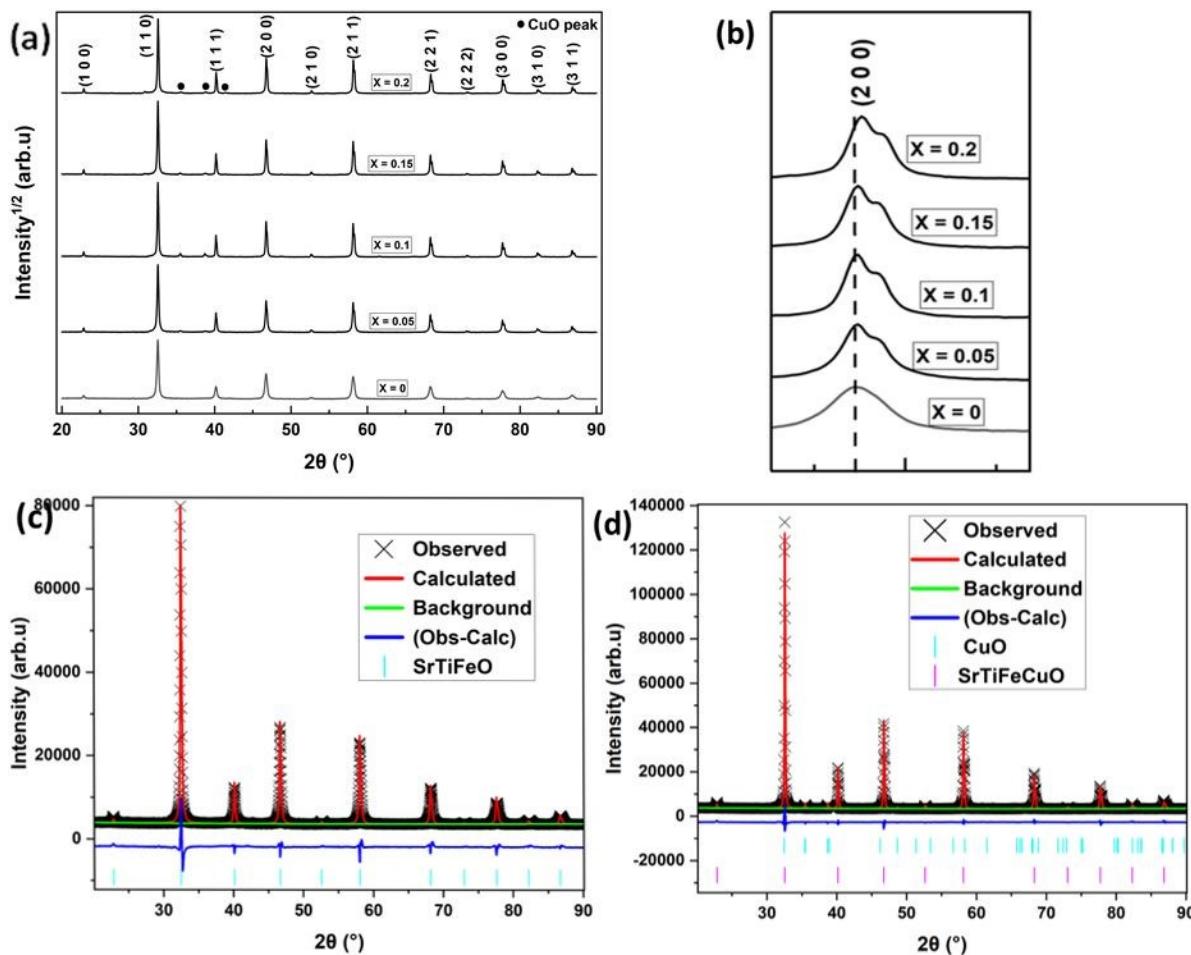
**Table 1.** Composition of the  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-\gamma}$  perovskite based on nominal Cu-doping amount in mols.

Stoichiometry composition	Nominal Cu-doping amount	Nominal substitution/comment
$\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7}\text{O}_{3-\gamma}$	0	5% A-site deficiency, B-site: 30% Ti, 70% Fe
$\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.65}\text{Cu}_{0.05}\text{O}_{3-\gamma}$	0.05	5% A-site deficiency, B-site: 30% Ti, 65% Fe, 5% Cu
$\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-\gamma}$	0.1	5% A-site deficiency, B-site: 30% Ti, 60% Fe, 10% Cu
$\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.55}\text{Cu}_{0.15}\text{O}_{3-\gamma}$	0.15	5% A-site deficiency, B-site: 30% Ti, 55% Fe, 15% Cu
$\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.2}\text{O}_{3-\gamma}$	0.2	5% A-site deficiency, B-site: 30% Ti, 50% Fe, 20% Cu



**Fig. 1** Perovskite synthesis process development and preliminary phase analysis results. a) A schematic of the perovskite synthesis process; b) a thermogram of  $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$  (green),  $\text{TiO}_2$  (black) and  $\text{SrCO}_3$  (pink), the precursors' mixture dry powder (red), and the differential scanning calorimetry (DSC) curve (blue). c) XRD peaks of  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-\gamma}$  showing improvement in phase purity after sintering.

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**Fig. 2** Phase analysis of the synthesised perovskite systems,  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$ : (a) Room temperature XRD pattern, highlighting the additional peaks resulting from doping (b) the  $(2\ 0\ 0)$  peaks showing the slight shift to a higher angle, with increasing  $x$ ; (c-d) Rietveld refinement of sintered samples of (c)  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7}\text{O}_{3-y}$  and d)  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-y}$ , showing the primary phase, and the secondary phase, CuO.

**Table 2** Structural and compositional parameters of  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$  systems ( $x = 0 - 0.2$ , space group:  $\text{Pm}\overline{3}\text{m}$ ) calculated from Rietveld refinement. Stoichiometries are reported in mols per mol of perovskite.

Nominal $x^*$ (mol)	0.00	0.05	0.1	0.15	0.2
Cell Parameter $a$ ( $\text{\AA}$ , $\pm 5\text{E}^{-4}$ )	3.8870	3.8850	3.8820	3.8810	3.8780
Crystallite Size ( $\mu\text{m}$ , $\pm 1\text{E}^{-1}$ )	0.11	0.58	0.63	0.54	0.60
Amount of CuO (mol, $\pm 1\text{E}^{-2}$ )	-	0.046	0.083	0.085	0.107
Actual $x^{[a]}$ (mol, $\pm 5\text{E}^{-2}$ )	-	0.010	0.020	0.070	0.110
Perovskite Stoichiometry	$\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7}\text{O}_{3-y}$	$\text{SrTi}_{0.32}\text{Fe}_{0.68}\text{Cu}_{0.01}\text{O}_{3-y}$	$\text{Sr}_{1.04}\text{Ti}_{0.33}\text{Fe}_{0.65}\text{Cu}_{0.02}\text{O}_{3-y}$	$\text{Sr}_{1.04}\text{Ti}_{0.33}\text{Fe}_{0.6}\text{Cu}_{0.07}\text{O}_{3-y}$	$\text{Sr}_{1.06}\text{Ti}_{0.34}\text{Fe}_{0.56}\text{Cu}_{0.11}\text{O}_{3-y}$
Goodness of Fit (a.u.)	3.05	1.39	2.70	2.86	4.48

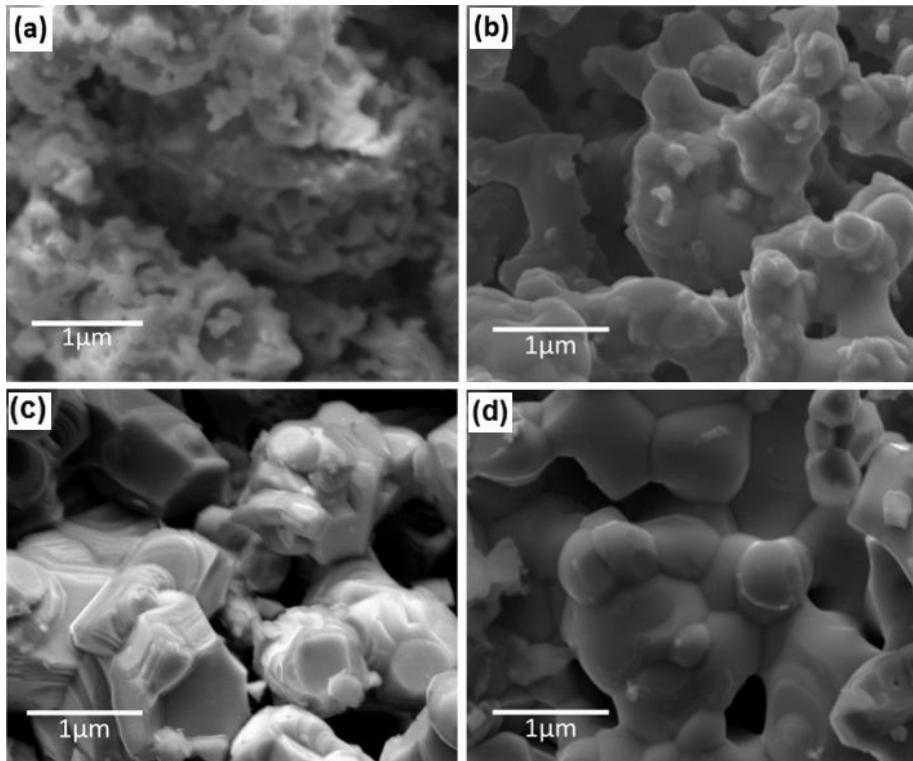
\*Cu-doping amount

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**Microstructure of the  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$  Systems**

The SEM analysis revealed well-formed grains with distinct grain boundaries and a better surface texture for the sintered  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$  ( $x = 0 - 0.2$ ) perovskite compared to the calcined samples (Fig. 3(a-d)). The sintered sample of the  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$  ( $x = 0$ ) perovskite contained interconnected grains with uniform porosity and grain sizes from  $\sim 0.2$  to  $1.5 \mu\text{m}$ . On

the other hand, sintered samples of the Cu-doped  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$  ( $x = 0.05-0.2$ ) perovskite showed evidence of enhanced sintering and larger grain sizes between  $\sim 0.5$  to  $3 \mu\text{m}$ . The observed microstructural evolution suggests that the sintering and Cu incorporation promotes densification and enhance grain boundary development, resulting in significant grain growth.(30)



**Fig. 3** SEM image showing the microstructure of various perovskites. (a)  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7}\text{O}_{3-y}$  ( $x: 0$ ) after calcination at  $1000^\circ\text{C}$ , (b)  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7}\text{O}_{3-y}$  ( $x: 0$ ) after sintering at  $1000^\circ\text{C}$  and (c)  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-y}$  ( $x: 0.1$ ) after calcination at  $1000^\circ\text{C}$ , and (d)  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-y}$  ( $x: 0.1$ ) after sintering at  $1000^\circ\text{C}$ .

**Reduction Kinetics of the  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$  Perovskite System**

The exsolution of metallic nanoparticles from the perovskite lattice occurs through several coupled steps: removing oxygen ions to create vacancies, which simultaneously introduces electrons into the lattice, essential for reducing B-site cations (likely Cu), ultimately leading to their nucleation as metallic particles on the surface. Therefore, quantifying oxygen removal kinetics and the total extent of oxygen loss provides crucial insights into both the rate-limiting steps of the exsolution process and the amount of metal ions being reduced, enabling rational design of reduction conditions for controlled nanoparticle formation. The kinetics and extent of oxygen

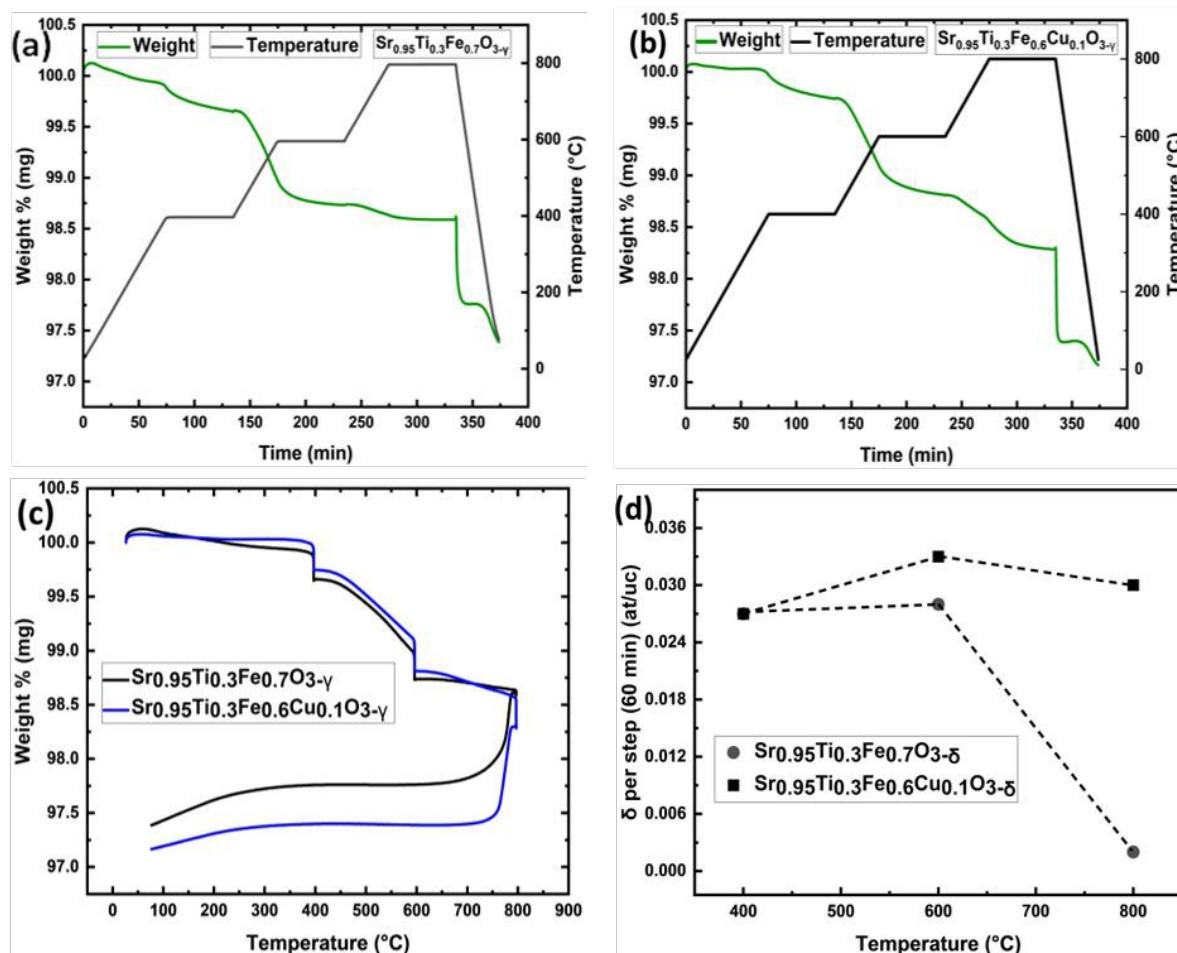
loss can be precisely monitored by tracking the material's weight changes during reduction using thermogravimetric analysis under hydrogen ( $\text{H}_2$ -TGA), where controlled temperature steps allow us to systematically investigate the oxygen transport and reduction processes that underpin exsolution.(31) The reduction kinetics of  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$  ( $x = 0$  and  $0.1$ ) perovskite was investigated using thermogravimetric analysis under a hydrogen atmosphere ( $\text{H}_2$ -TGA). The reduction mechanism was studied at three isothermal temperature steps:  $400^\circ\text{C}$ ,  $600^\circ\text{C}$ , and  $800^\circ\text{C}$ , with detailed comparisons of weight loss and oxygen deficiency (Fig. 4(a-d)). At  $400^\circ\text{C}$ , a weight loss of approximately  $0.23\%$  was observed for both  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7}\text{O}_{3-y}$  and  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-y}$ . **Table 3**

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summarises the H<sub>2</sub>-TGA result and the oxygen deficiencies for the Sr<sub>0.95</sub>Ti<sub>0.3</sub>Fe<sub>0.7</sub>O<sub>3- $\gamma$</sub>  and Sr<sub>0.95</sub>Ti<sub>0.3</sub>Fe<sub>0.6</sub>Cu<sub>0.1</sub>O<sub>3- $\gamma$</sub>  perovskite. The mass losses are attributed to the lattice oxygen loss accompanying the reduction of Fe<sup>4+</sup>→Fe<sup>3+</sup> and the exsolution of Cu, as well as the reduction of the CuO secondary phase to Cu metal. CuO a reduction temperature as low as 180 °C,(31) hence, is expected to reduce first. At 600 °C, both samples exhibited additional weight losses, with the

Cu-doped perovskite showing a slightly higher loss (0.28% compared to 0.24% for the undoped system). This stage is associated with further oxygen loss and the exsolution of Cu nanoparticles from the Cu-doped sample. As shown in the next section, the Cu-doped sample showed more significant and well-dispersed nanoparticle formation at this stage, verified by SEM imaging.



**Fig. 4** Thermograms and oxygen deficiency ( $\delta$ ) plots from the H<sub>2</sub>-TGA of Sr<sub>0.95</sub>Ti<sub>0.3</sub>Fe<sub>0.7-x</sub>Cu<sub>0.1</sub>O<sub>3- $\gamma$</sub>  ( $x = 0 - 0.1$ ) perovskites showing mass loss and temperature vs time plot of (a) Sr<sub>0.95</sub>Ti<sub>0.3</sub>Fe<sub>0.7</sub>O<sub>3- $\gamma$</sub>  and (b) Sr<sub>0.95</sub>Ti<sub>0.3</sub>Fe<sub>0.6</sub>Cu<sub>0.1</sub>O<sub>3- $\gamma$</sub> ; (c) comparison of the plot weight loss vs temperature for Sr<sub>0.95</sub>Ti<sub>0.3</sub>Fe<sub>0.7</sub>O<sub>3- $\gamma$</sub>  (black) and Sr<sub>0.95</sub>Ti<sub>0.3</sub>Fe<sub>0.6</sub>Cu<sub>0.1</sub>O<sub>3- $\gamma$</sub>  (blue); and (d) oxygen deficiency vs temperature plot of Sr<sub>0.95</sub>Ti<sub>0.3</sub>Fe<sub>0.7</sub>O<sub>3- $\gamma$</sub>  and Sr<sub>0.95</sub>Ti<sub>0.3</sub>Fe<sub>0.6</sub>Cu<sub>0.1</sub>O<sub>3- $\gamma$</sub>  perovskite material. The straight, dotted line serves as a guide for the eye.

**Table 3** Weight losses and calculated oxygen deficiency ( $\delta$ ) from the H<sub>2</sub>-TGA of Sr<sub>0.95</sub>Ti<sub>0.3</sub>Fe<sub>0.7</sub>O<sub>3- $\gamma$</sub>  ( $x = 0$  and  $x = 1$ ) perovskite. Only oxygen deficiency calculated from mass loss recorded at the isothermal temperature step was considered, as proposed in the literature.(32)

Temperature (°C)	Weight Loss (%)		Cumulative Weight Loss (%)		Oxygen Deficiency $\delta$ (at/uc)		Cumulative $\delta$ (at/uc)	
	$x = 0$	$x = 0.1$	$x = 0$	$x = 0.1$	$x = 0$	$x = 0.1$	$x = 0$	$x = 0.1$
400	0.23	0.23	0.23	0.23	0.026	0.027	0.026	0.027
600	0.24	0.28	0.47	0.51	0.028	0.032	0.054	0.059
800	0.02	0.26	0.49	0.77	0.002	0.030	0.056	0.089

The final temperature step at 800 °C resulted in a minimal weight loss of 0.02% for the undoped sample and an additional weight loss of 0.26% for the Cu-doped samples. This minimal change for the

undoped sample suggests its reduction process was nearly complete by 600 °C. In contrast, the weight loss at 800 °C indicates a possibility for additional exsolved nanoparticle growth in the Cu-doped

perovskites. Cu-doped perovskite displayed a cumulative oxygen deficiency reaching 0.089 at/uc, vs 0.056 for the undoped composition (Table 3), the highest observed value, indicating greater reducibility due to Cu doping, consistent with one of the premises of the work. At the same time, the limited increase in oxygen deficiency at lower temperatures for the Cu-doped samples compared to the undoped samples is consistent with the observation that not all Cu substitutes in the perovskite as intended. As corroborated in the literature,(32) in this work, the reduction behaviour of the undoped and Cu-doped  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$  perovskites was modelled using a double exponential decay function, following Equation (1):

$$m(t) = f_0 + f_1 e^{-k_1 t} + f_2 e^{-k_2 t} \quad (1)$$

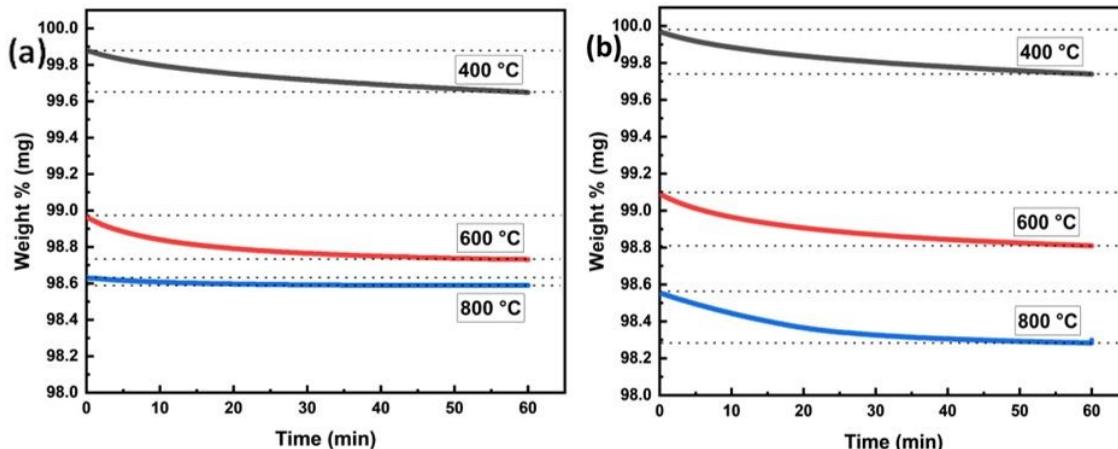
Where  $m(t)$  is the perovskite's mass at time  $t$  during isothermal reduction;  $f_0$  the sample mass at the isothermal step;  $f_1$  and  $f_2$  represent the fractional contributions of the two distinct kinetic processes;  $k_1$  and  $k_2$  are the respective time constants (kinetic parameters) associated with each process. This reflects two parallel, rate-controlling processes: (i) Surface exchange kinetics (fast,  $k_1$ ) and (ii) bulk oxygen diffusion (slow,  $k_2$ ).<sup>(32,33)</sup>

**Fig. 5 (a-b)** compares the reduction kinetics of the undoped and Cu-doped systems. Our fitting revealed that Cu doping enhances the bulk diffusion component ( $k_2$ ), consistent with a modification of the oxygen vacancy energy landscape.

**Table 4** presents the calculated parameters for the two perovskites at different isothermal temperature steps. For the perovskites studied here, the kinetic parameter,  $k_1$ , had the highest values at 400 °C, the early stage of the reduction process.

A striking feature in the Cu-doped sample is the dramatic shift in the  $k_2/k_1$  ratio from 0.14 at 400 °C to 5.67 at 600 °C, before returning to 1.0 at 800 °C. This shift reveals a complex evolution of rate-determining steps with temperature. This non-monotonic behaviour suggests that different reduction mechanisms dominate at various temperature regimes, with bulk diffusion becoming particularly enhanced at intermediate temperatures. Nanoparticle exsolution and its impact on the bulk crystal and surface structure could also be contributing factors. At 600 °C, Cu-doping appears to fundamentally alter the reduction kinetics, significantly accelerating bulk processes ( $k_2 = 0.17$ ) while simultaneously slowing surface processes ( $k_1 = 0.03$ ) compared to the undoped sample ( $k_2 = 0.04$ ,  $k_1 = 0.19$ ). This inverse effect on surface versus bulk kinetics suggests that Cu-doping may modify oxygen transport pathways by affecting the material's vacancy formation and migration energies. However, it could also be related to competition with reactions such as exsolution, which has surface and bulk components.

The undoped sample shows an unusual evolution in the relative contributions of surface and bulk processes, with the  $f_1/f_2$  ratio shifting from ~0.17 at 400 °C to 1.0 at 800 °C. This indicates the initial dominance of bulk oxygen transport that eventually equilibrates with



**Fig. 5** The reduction kinetics at the three isothermal temperature steps (400, 600, and 800 °C), based on mass loss due to structural oxygen loss from (a)  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7}\text{O}_{3-y}$  and (b)  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-y}$  perovskite. TGA was run in a 5%  $\text{H}_2$ /95% Ar stream.

**Table 4** Fitting parameters for the reduction kinetics model of  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7}\text{O}_{3-y}$  and  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-y}$  perovskite material

Fitting Parameters	Perovskite at Different Temperatures					
	$\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7}\text{O}_{3-y}$			$\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-y}$		
	400°C	600°C	800°C	400°C	600°C	800°C
$f_0$ (mg)	99.54	98.72	98.59	99.66	98.77	98.27
$f_1$ (mg)	0.05	0.08	0.02	0.05	0.24	0.15
$K_1$ (min <sup>-1</sup> )	0.15	0.19	0.08	0.14	0.03	0.06
$f_2$ (mg)	0.29	0.17	0.02	0.26	0.08	0.15
$K_2$ (min <sup>-1</sup> )	0.02	0.04	0.08	0.019	0.17	0.06

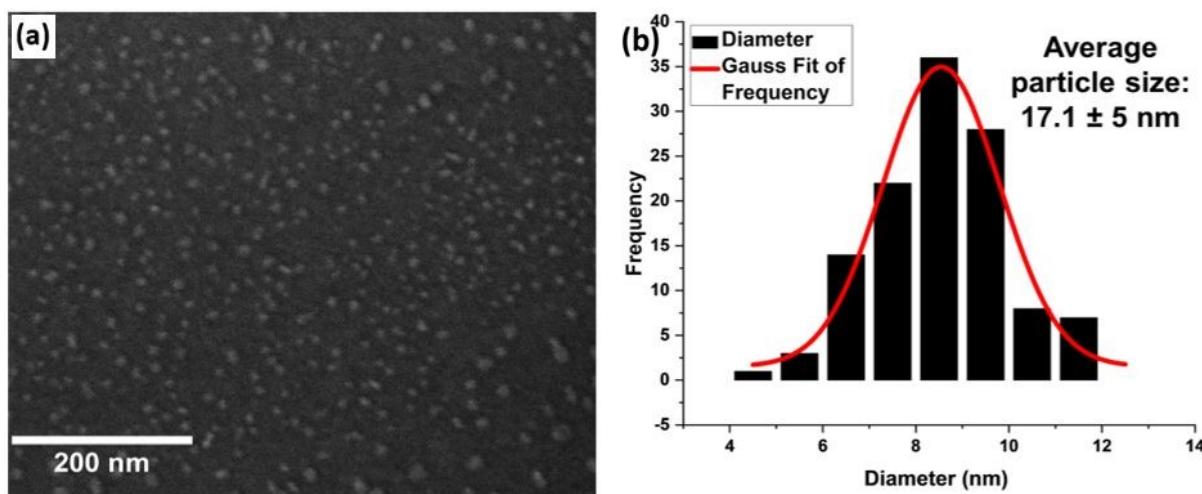
## ARTICLE

surface processes. In contrast, the Cu-doped sample maintains consistently higher  $f_2$  values throughout the temperature range, suggesting that Cu-doping promotes sustained bulk-mediated reduction processes even at higher temperatures. This result indicates that Cu-doping enhances the reducibility and bulk lattice oxygen transport capability of  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-\gamma}$ , suggesting enhanced exsolution capability. It also demonstrates that  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-\gamma}$  is a promising candidate for applications requiring high oxygen vacancy concentrations and reducibility.

#### Exsolution Capability of the $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-\gamma}$ Perovskite System

The exsolution behaviour of the  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-\gamma}$  perovskite system ( $x: 0-0.1$ ) was investigated, focusing on the influence of Cu-

doping, reduction temperature, and time on the material's structure and surface properties. Samples were reduced in 5%  $\text{H}_2$  at temperatures of 400, 600, and 800 °C and for durations ranging from 1 to 4 hours. **Fig. 6(a-b)** shows the SEM image and particle size analysis result of the undoped perovskite,  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7}\text{O}_{3-\gamma}$ , reduced at 600 °C in 1 hour. Under these conditions, the observed particles display an average size of  $\sim 17$  nm. XRD and XPS analysis (Fig. 7(a-d) and later discussion) revealed important distinctions about particle formation in the undoped perovskite at different temperatures. At 400°C, no particle formation occurred (Figure S1). At 600°C, while particles were observed (Fig. 6a), our data indicate these are iron oxide nanoparticles formed by segregation rather than true metallic iron exsolution. Genuine Fe metal nanoparticle exsolution only occurred at the higher temperature of 800°C (Fig. 7(a-d)).



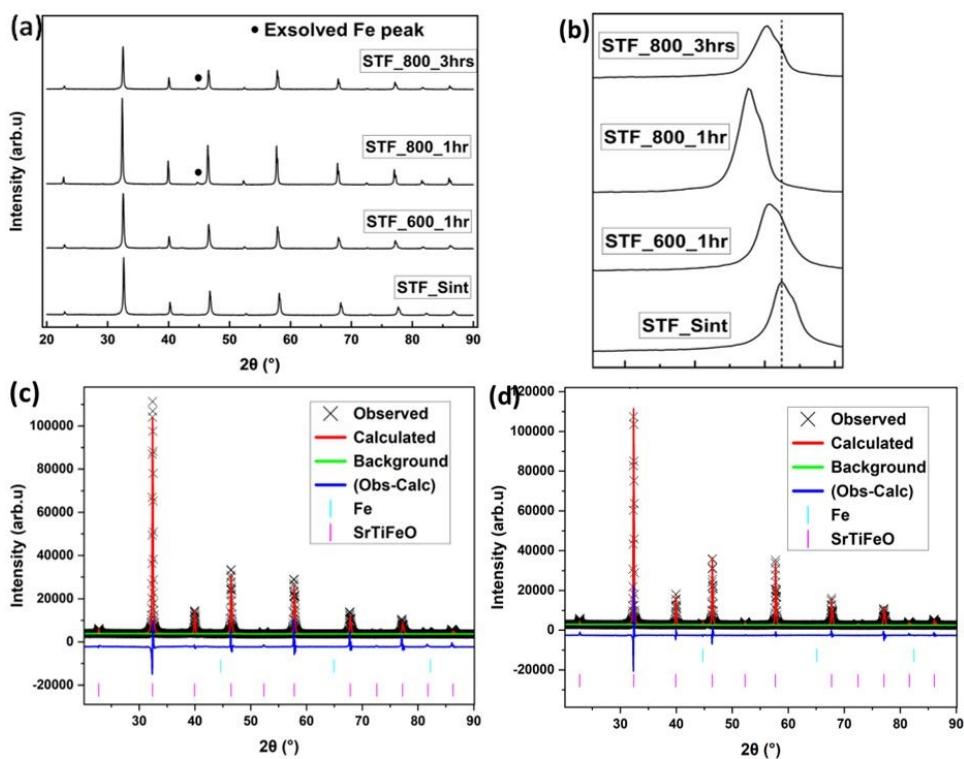
**Fig. 6** Micrograph image and the corresponding particle size analysis for  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7}\text{O}_{3-\gamma}$  ( $x = 0$ ) perovskite, showing the structure following reduction at 600 °C. (a) SEM of observed exsolved nanoparticles, (b) corresponding particle size distribution.

At 400 °C, this likely indicates an insufficient driving force for Fe exsolution from the undoped material. At 800 °C, the exsolution may have been suppressed, possibly due to surface restructuring or A-site enrichment, which is known to impact exsolution in perovskites.(34-36) The oxygen loss measured at 600°C correlates with particle formation observed in SEM images. However, XRD and XPS analysis (detailed later) confirm that iron remains in an oxidized state rather than forming metallic nanoparticles. This indicates that while the undoped perovskite undergoes structural changes and forms oxide particles at moderate temperatures, it cannot achieve true metallic exsolution below 800°C. This temperature limitation restricts the

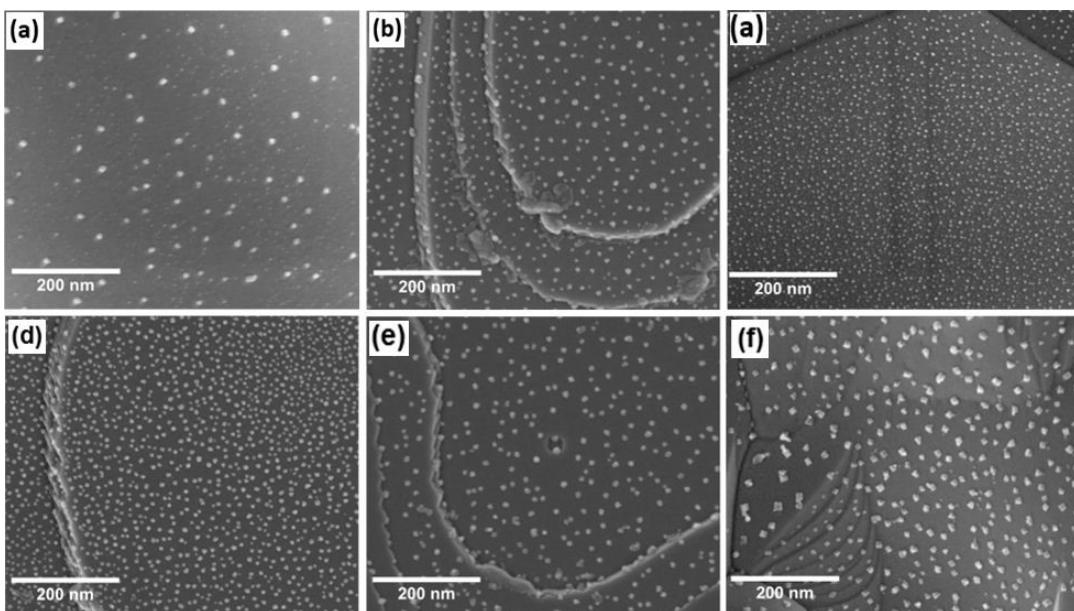
material's usefulness for applications requiring low-temperature processing.

**Fig. 8(a-f)** show the SEM image while **Fig. 9(a)** shows the particle size analysis results for  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6-x}\text{Cu}_{0.1}\text{O}_{3-\gamma}$  perovskite at various reduction conditions. See Fig. S1(d) for an SEM image of the  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-\gamma}$  pristine sample. This perovskite exhibited pronounced nanoparticle exsolution, particularly at temperatures of 600 °C and 800 °C. At 400 °C, relatively spherical nanoparticles were observed, with a bimodal distribution centred at approximately 13 and 24 nm, suggesting that these temperature and time conditions may be insufficient for achieving size equilibration.

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**Fig. 7** Phase analysis of  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7}\text{O}_{3-\gamma}$  perovskite system reduced at different temperatures. (a) Room temperature XRD pattern, highlighting the exsolved Fe peak clearly observed after reduction at  $800\text{ }^\circ\text{C}$ ; (b) the (2 0 0) peaks showing the slight shift to a lower angle; (c-d) Rietveld refinement of the samples reduced at  $600\text{ }^\circ\text{C}$  and  $800\text{ }^\circ\text{C}$ ; (c) sample reduced at  $600\text{ }^\circ\text{C}$  in 1 hour and (d) sample reduced at  $800\text{ }^\circ\text{C}$  in 1 hour, confirming Fe exsolution.



**Fig. 8** Micrograph image of  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-\gamma}$  perovskite after reduction at  $400\text{ }^\circ\text{C}$ ,  $600\text{ }^\circ\text{C}$ , and  $800\text{ }^\circ\text{C}$ . SEM of observed exsolution at (a)  $400\text{ }^\circ\text{C}$  in 1 hour; (b – e)  $600\text{ }^\circ\text{C}$  in reduction time of 1, 2, 3, and 4 hours, respectively; and (f)  $800\text{ }^\circ\text{C}$  in 1 hour. Nanoparticle sizes distribution at (g)  $400\text{ }^\circ\text{C}$  in 1 hour; (h – k)  $600\text{ }^\circ\text{C}$  in reduction time of 1, 2, 3, and 4 hours, respectively; and (l)  $800\text{ }^\circ\text{C}$  in 1 hour.



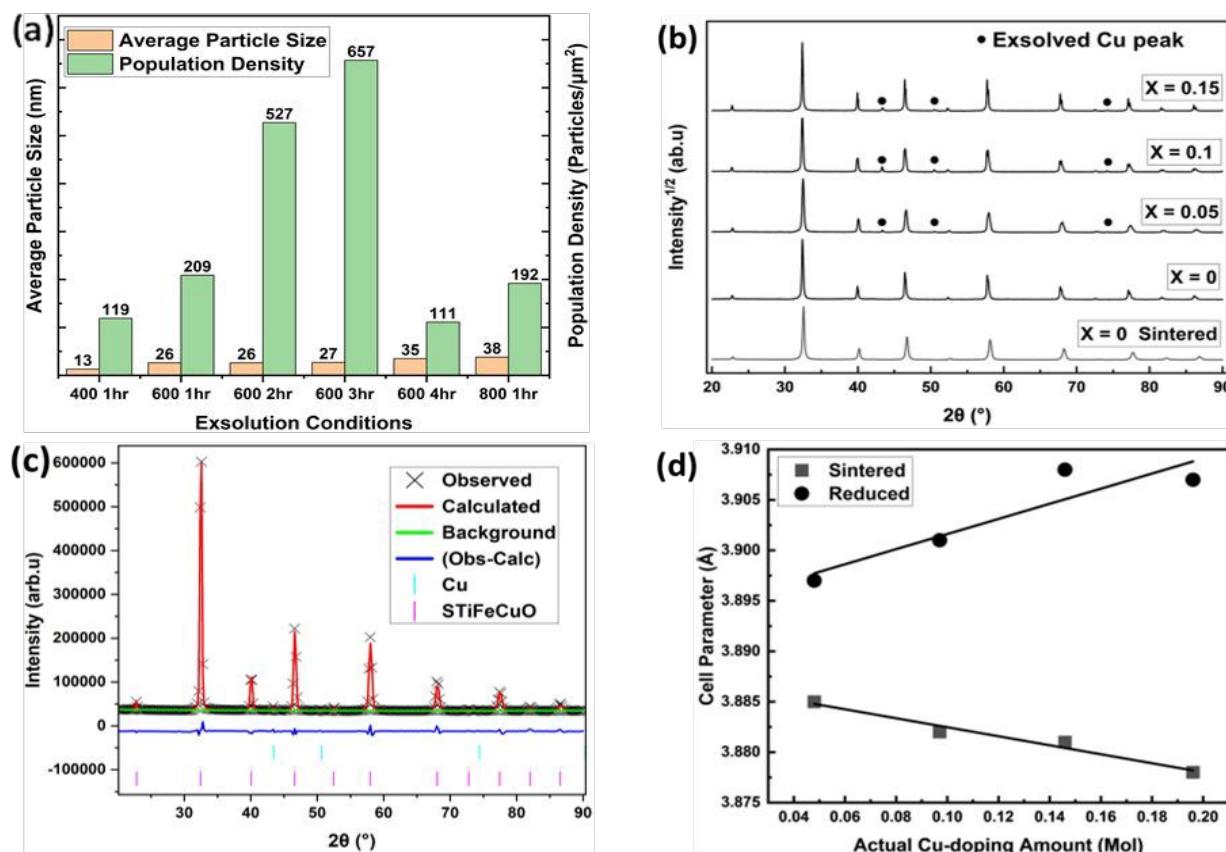
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Increasing the temperature to 800 °C resulted in multifaceted particle shapes with an average diameter of ~38 nm. The influence of the reduction time was also investigated, specifically at 600 °C, where nanoparticle density increased from  $207 \pm 2$  particles/ $\mu\text{m}^2$  at 1 hour to  $650 \pm 7$  particles/ $\mu\text{m}^2$  at 3 hours. At 800 °C, the elevated temperature promoted further growth and morphological evolution of the nanoparticles, resulting in a decrease in the nanoparticle population to approximately  $188 \pm 4$  particles/ $\mu\text{m}^2$ , compared to the  $207 \pm 2$  particles/ $\mu\text{m}^2$  obtained at 600 °C in 1 hour. The quoted error in the population density was determined, as explained in Note S2. See Table S1 for the summary of particle size analysis for the  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7}\text{O}_{3-y}$  and  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-y}$ .

XRD and Rietveld refinement (Fig. 9(b–c)) indicated successful exsolution driven by the incorporated Cu-dopant in the perovskite system. See Figure S2 and Table S2 for the pattern and summary of the SEM-EDS result and the determined stoichiometry of the  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-y}$  perovskite's surface composition.

The structural stability of the perovskite remained largely unaffected by the exsolution process, with only some changes in lattice parameters, as confirmed by the Rietveld refinement result. Fig. 9(d) shows that the cell parameters decreased with the Cu-doping amount for the sintered samples (3.886–3.878 Å) and increased for the reduced samples, reaching a maximum of 3.908 Å. See Table S3 for the cell parameters of the perovskites derived from the Rietveld refinement. The increase in lattice parameters was attributed to the reduction of  $\text{Fe}^{4+}$  to  $\text{Fe}^{3+}$  (which increases the average B-site cation size), oxygen vacancy formation, and the exsolution of Cu nanoparticles, which contributed to charge compensation via the removal of  $\text{Cu}^{2+}$  cations from the lattice.

The findings presented here demonstrate that the  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$  perovskite system, particularly with Cu doping, despite limited substitutions, has strong potential for the controlled exsolution of nanoparticles, making it a promising candidate for catalysis applications.



**Fig. 9** Particle size analysis result of  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-y}$  perovskite at different reduction conditions; and XRD and Rietveld refinement result of  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$  perovskite system reduced at 600 °C in 1 hour. (a) Particle sizes and population density at different reduction conditions (400 °C, not fully represented as a bimodal distribution on the plot); (b – c) XRD patterns: (b) showing peaks of Cu nanoparticles after reduction, (c) Rietveld refinement profile identifying the exsolved nanoparticles as metallic Cu, and (d) variation of perovskites' cell parameter with actual Cu-doping amount for the as-synthesized and reduced perovskites.

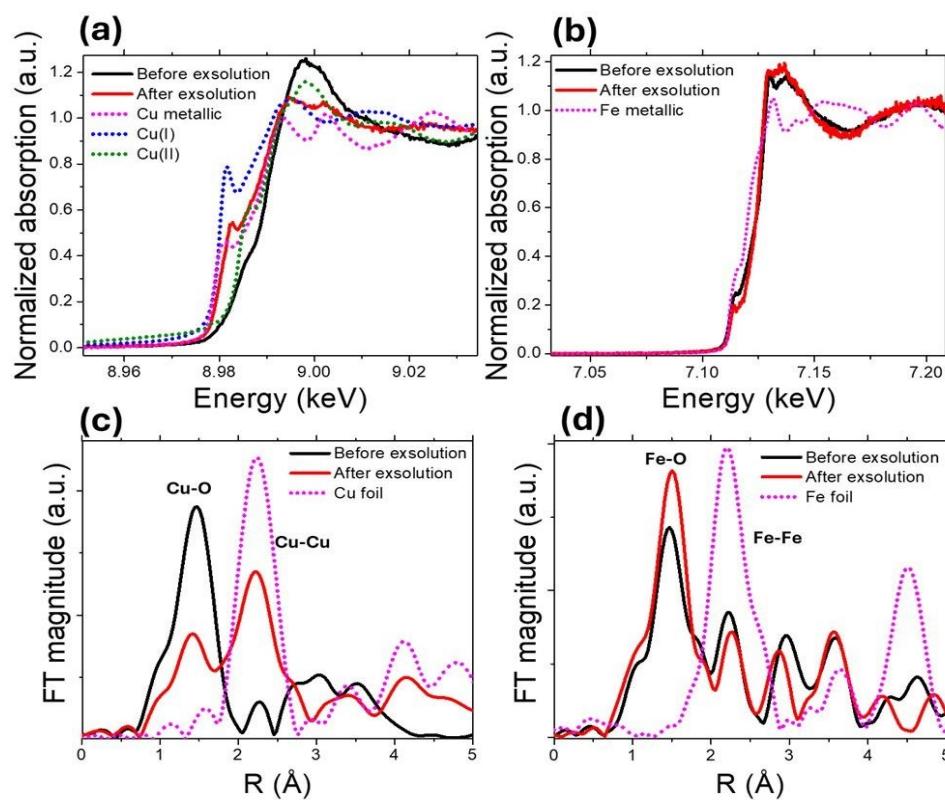
## ARTICLE

**Chemical State Characterization of Exsolved Perovskites by XAS and Surface Characterisation by XPS**

To understand how the thermal treatment affects the oxidation state of Copper and Iron, we performed *in situ* XAS measurements in the Cu and Fe edges before and after the exsolution process at 600 °C for 1 h (Fig. 10 (a-b)). Before exsolution (black curve), the X-ray Near Edge Structure (XANES) profile exhibits an elevated white line characteristic of Cu oxides,(37) in line with all the previous characterizations. After exsolution (red curve), the absorption edge shifts to lower energy values, accompanied by a reduction in white-line intensity, suggesting a reduction of Cu species. The spectrum becomes more similar to that of Cu(I) or metallic Cu (pink curve), again in line with the previous characterizations (XRD) showing the presence of metallic Cu. The linear combination analysis yielded fractions of  $0.28 \pm 0.02$  for Cu<sup>0</sup>,  $0.50 \pm 0.02$  for Cu<sup>1+</sup>, and  $0.22 \pm 0.04$

for Cu<sup>2+</sup> after exsolution, with an R-factor of 0.0018. Therefore, there remains a considerable amount of Cu<sup>1+</sup> after exsolution, which explains why its XANES profile considerably differs from that of metallic Cu.

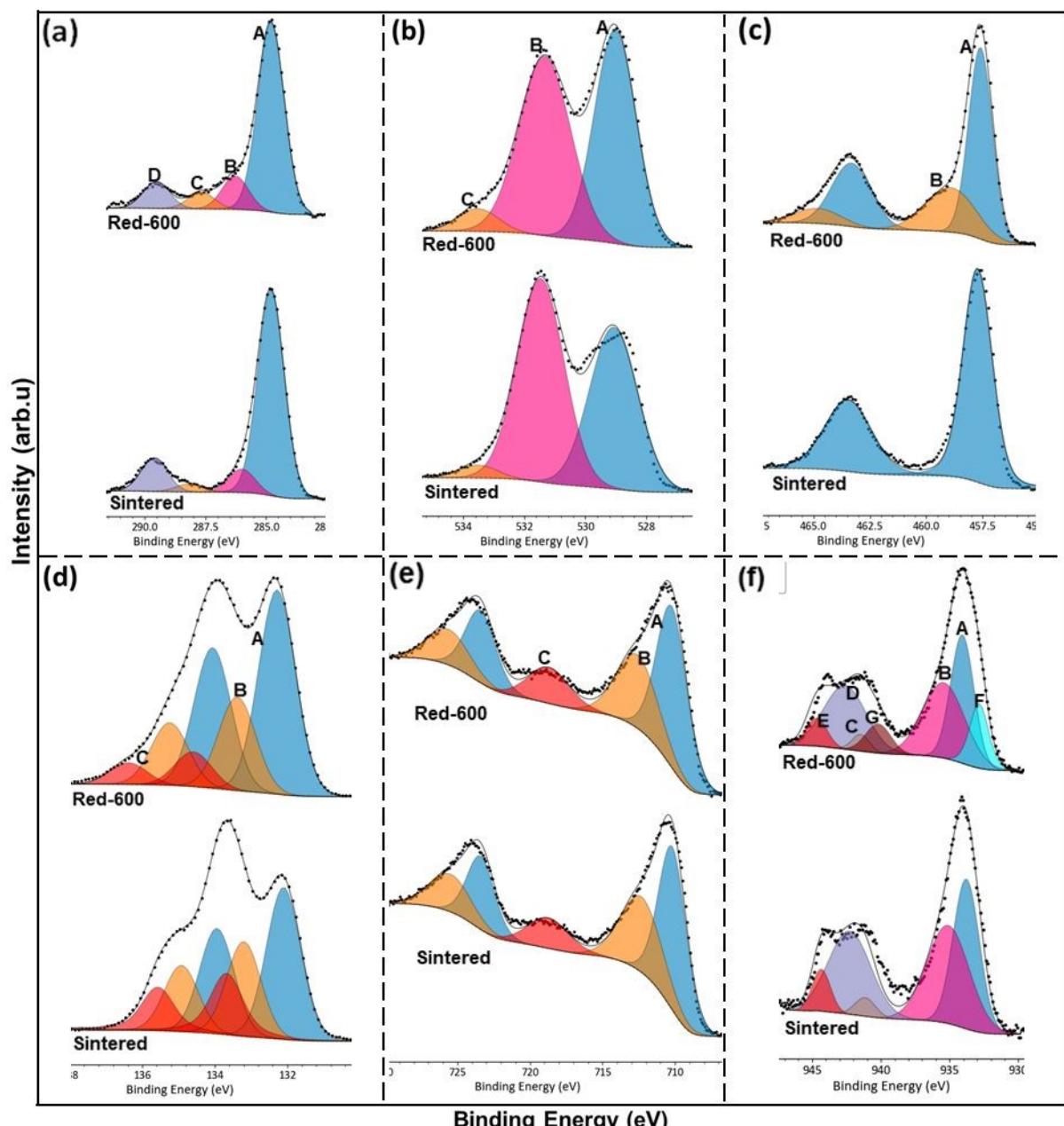
On the other hand, the changes are much less clear in the case of the Fe edge, indicating that Fe atoms undergo some relatively minor changes in comparison with Cu, confirming that the exsolution occurs mainly due to the reduction of Cu, and then the consequent formation of pure metallic Cu nanoparticles. We cannot discard the possibility of some alloying but, if it is the case, the XAS experiments show that the nanoparticles should be highly enriched in Cu. The Extended X-ray Absorption Fine Structure (EXAFS) analysis (Fig. 10 (c-d)) reinforces what is observed with XANES. After exsolution, there is a remarkable increase in the Cu-Cu bond peak, while the Fe-Fe peak barely changes.



**Figure 10.** *In situ* Cu and Fe K-edge XAS spectra of the Cu-doped perovskite collected before (black) and after (red) exsolution at 600 °C for 1 h. (a) Cu K-edge XANES spectrum shows a shift to lower energies and a reduced white-line intensity, indicating partial reduction of Cu species and the emergence of Cu(I)/Cu<sup>0</sup> features after exsolution. (b) Fe K-edge XANES displaying only subtle spectral changes, suggesting that Fe remains largely unaffected during reduction. (c) Fourier-transformed EXAFS magnitude plot showing the Cu-Cu peak markedly increases after reduction, confirming metallic Cu nanoparticle formation. (d) Fourier-transformed EXAFS magnitude plot showing that Fe-O and Fe-Fe contributions show minimal change, indicating limited structural modification of Fe.



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**Fig. 11** Surface analysis by XPS for  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-y}$  perovskites. Core spectra of C1s, O1s, Ti2p, Sr3d, Fe2p and Cu2p, for Sintered and Reduced ( $600\text{ }^{\circ}\text{C}$ , 1hr) sample, after deconvolution. (a) C1s spectra deconvoluted into four components [A = C-C, B = C-O, C = C=C; and D = O-C=O] (b) O1s spectra deconvoluted into 3 components [A = O-Lat, B = O-Surface, and C = O-CO/O-OH]; (c) Ti2p deconvoluted into 1 component [A =  $\text{Ti}^{4+}$  doublet] for both the sintered and reduced sample, and two additional satellite peaks [B =  $\text{Ti}^{4+}$  satellite peaks] for the reduced sample; (d) Sr3d spectra deconvoluted into 3 components [A = Sr-Lat doublet, B =  $\text{SrCO}_3$  doublet, and C =  $\text{SrO}_{(1-x)}$  doublet]; (e) Fe2p spectra deconvoluted into 2 components and a satellite peak [A =  $\text{Fe}^{3+}$  doublet, B =  $\text{Fe}^{4+}$  doublet, and C =  $\text{Fe}^{3+}$  satellite peak]; and (f) Cu2p, deconvoluted into 2 components and 3 satellite peaks for the sintered and reduced sample [A =  $\text{Cu}^{2+}$ , B =  $\text{Cu}^{2+}$  Isolated atom, C - E = Satellite peak 1, 2, and 3], plus additional 1 component and 1 satellite peak [F =  $\text{Cu}^0$ , and G =  $\text{Cu}^0$  satellite peak] for the reduced sample. The different species of all the fitted spectra are summarised on **Table S4 and S9**.



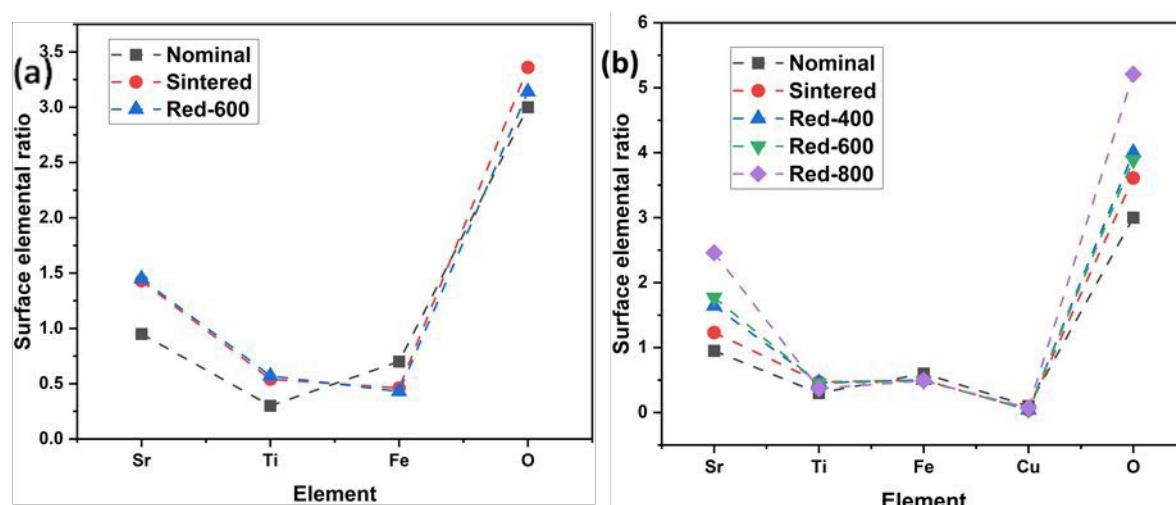
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Furthermore, the XPS core-level spectra (Figs. 11, S3 and S4) collected on both the undoped STF and Cu-STF materials in their as-sintered state and after  $H_2$  reduction (400–800 °C) provide supporting information that confirms metallic Cu as the only exsolved metallic component at 600 °C. Fig. 11 Compares the high-resolution core spectra of C1s, O1s, Ti2p, Sr3d, Fe2p and Cu2p, for  $Sr_{0.95}Ti_{0.3}Fe_{0.7-x}Cu_xO_{3-y}$  Sintered and Reduced (600 °C, 1hr) sample, after deconvolution. The XPS analysis shows that  $Sr_{0.95}Ti_{0.3}Fe_{0.7-x}Cu_xO_{3-y}$  surfaces are Sr-enriched with mixed  $Fe^{3+}/Fe^{4+}$  states and surface oxygen overstoichiometry. Cu doping increases surface oxygen and carbonate adsorption, while reduction drives oxygen vacancy formation, and Cu exsolution. The Cu2p spectra reveal a mixed  $Cu^{2+}/Cu^0$  environment under reducing conditions, with  $Cu^{2+}$  sites believed to aid nitrate adsorption while the  $Cu^0$  nanoparticles enhancing the reduction. Overall, both  $Sr_{0.95}Ti_{0.3}Fe_{0.7}O_{3-y}$  and  $Sr_{0.95}Ti_{0.3}Fe_{0.6}Cu_{0.1}O_{3-y}$  surfaces deviate significantly from bulk stoichiometry, with these changes intensifying upon reduction (Fig. 12 a-b and Table S10 as well as Figure S5).

XPS analysis reveals notable deviations of surface composition from the nominal stoichiometry according to XRD Rietveld refinement, as shown on Table S10, for both  $Sr_{0.95}Ti_{0.3}Fe_{0.7}O_{3-y}$  and  $Sr_{0.95}Ti_{0.3}Fe_{0.6}Cu_{0.1}O_{3-y}$ . In all cases, Sr is enriched at the surface, increasing with reduction temperature, particularly in the Cu-doped perovskite (showing Sr/B site ratio of 1.63 → 2.46), indicating

pronounced A-site migration under reducing conditions. Fe ratio is low relative to nominal, for example: decreasing from 0.7 in the nominal to 0.43 in the undoped reduced sample. on the other hand, Ti remains higher than the nominal, indicating Sr segregation driven surface reconstruction. The  $Fe^{3+}$  to  $Fe^{4+}$  ratio increases with reduction, confirming partial  $Fe^{4+} \rightarrow Fe^{3+}$  reduction and corresponding oxygen vacancy formation. Cu ratio remains lower than the nominal value across all conditions, though increasing with reduction temperature, indicating increase extend of exsolution at elevated temperatures. Surface oxygen ratios remain above 3, consistent with oxygen over-stoichiometry arising from Sr segregation and  $CO_3$  and oxide species adsorption.

Across processing conditions, the evolution of Sr enrichment, Fe depletion, and Fe redox transitions reflects a dynamic balance between reduction and cation redistribution. Moderate reduction (400 °C–600 °C) provides the most favourable surface state for catalytic redox activity, characterised by accessible  $Fe^{3+}/Fe^{4+}$  or  $Cu^0$  sites and oxygen vacancies, and moderate Sr segregation. The excessive Sr segregation and oxygen over-stoichiometry at higher temperatures (e.g., Red-800) likely produce insulating Sr-rich surface layers that hinder charge transfer and block active sites. These findings reiterate that Cu incorporation enhances surface stability and redox flexibility, while controlled reduction is essential for maximising catalytic performance and maintaining structural integrity.



**Fig. 12.** Surface stoichiometry by XPS of the  $Sr_{0.95}Ti_{0.3}Fe_{0.7-x}Cu_xO_{3-y}$  perovskite system showing the deviation of different processed samples from the nominal (as-designed) sample. (a) Comparison of the  $Sr_{0.95}Ti_{0.3}Fe_{0.7}O_{3-y}$  perovskite sintered and reduced sample from the nominal sample. (b) Comparison of the  $Sr_{0.95}Ti_{0.3}Fe_{0.7-x}Cu_xO_{3-y}$  perovskite sintered and reduced (400, 600, and 800 °C) sample from the nominal sample.



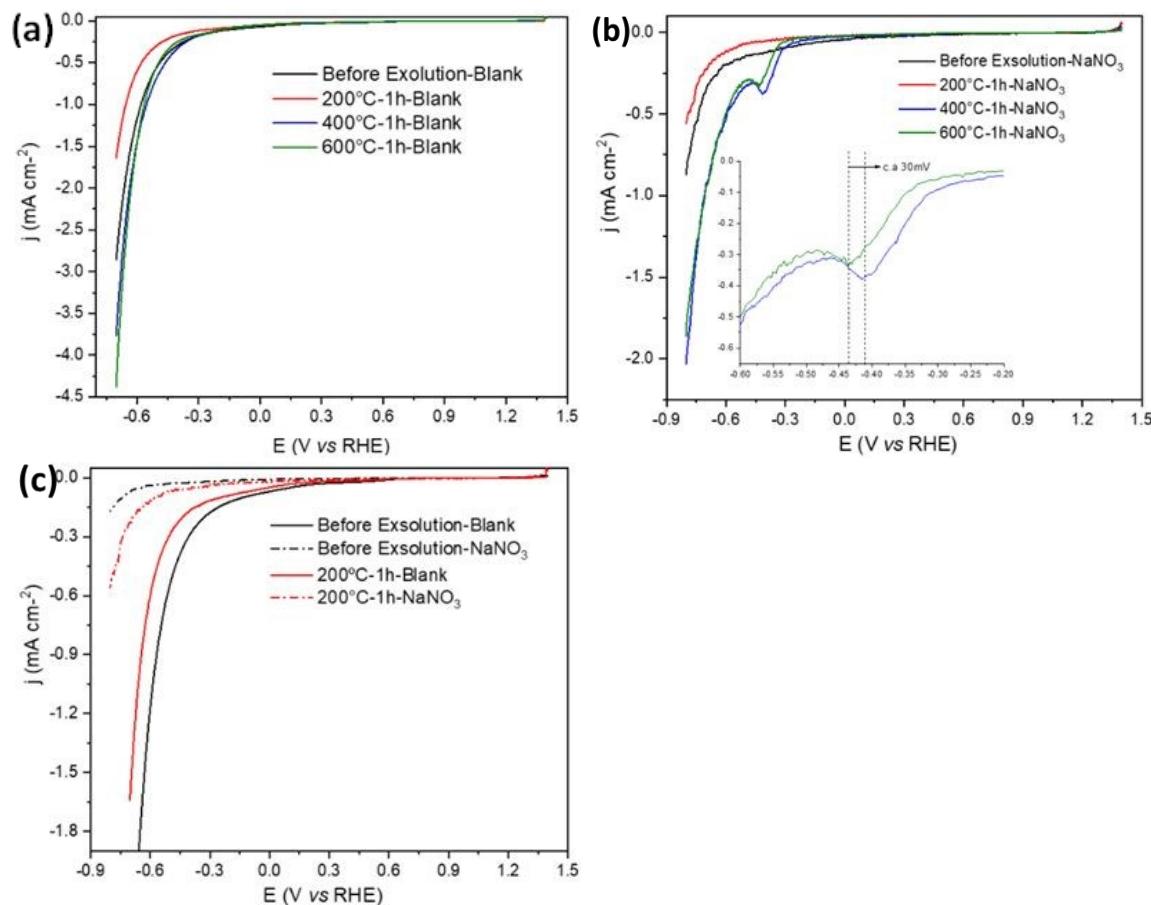
## ARTICLE

## Electrochemical Characterization

To analyse the surface features of  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-\gamma}$  before and after exsolution at different temperatures, we conducted a reduction treatment (5% H<sub>2</sub> and 95% Ar) at 200 °C (where no exsolution is evident) and at intermediate temperatures of 400 °C and 600 °C for 1 hour. The 200 °C treatment was added to reduce the gap (and then have a smoother transition) between the reduced and non-reduced samples.

Even if electrochemical measurements are most commonly used to test the activity and stability of materials, these can be extremely useful for characterising some surfaces.<sup>(38)</sup> Thus, in this work, we have utilised the interaction between nitrate ions and Cu to monitor changes at the surface of Cu-based materials qualitatively. **Fig. 13(a–**

b) shows the cathodic sweep of cyclic voltammograms (CVs) recorded in the absence and presence of NaNO<sub>3</sub>, respectively. The choice of nitrate provides insights into the NO<sub>3</sub>RR (electro)catalytic activity changes before and after exsolution, serving as a surface probe for material characterisation.<sup>(39)</sup> Shifts in the reduction potentials or current densities indicate changes in surface chemistry.<sup>(40)</sup> In the absence of nitrate, notable reduction currents were observed for all samples at more negative potentials than -0.3V  $V_{\text{RHE}}$  (Fig. 13(a)), corresponding to the hydrogen evolution reaction (HER) and which is typical of metallic and metallic oxide surfaces.<sup>(41–43)</sup>



**Fig. 13** The cathodic sweep of CVs for  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-\gamma}$  before exsolution and after exsolution at 200 °C, 400 °C, and 600 °C for 1h in (a) 0.1 M NaOH; (b) In 0.1 M NaOH and 1 mM NaNO<sub>3</sub> at 50 mVs<sup>-1</sup> (inset: NO<sub>3</sub>RR peaks for 400 °C and 600 °C samples); (c) Comparison of  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.6}\text{Cu}_{0.1}\text{O}_{3-\gamma}$  before exsolution and exsolved at 200 °C for 1h in 0.1 M NaOH vs. 0.1 M NaOH with 1 mM NaNO<sub>3</sub>. All the potentials are referred to RHE.

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However, it is worth noting that the sample treated at 200 °C showed notably lower HER activity. We argue that the CuO secondary phase was reduced (as discussed earlier) to a less active Cu fraction, which lacks the enhanced catalytic properties of Cu nanoparticles or the CuO phase existing in the sample before exsolution. This change likely results from the altered distribution of Cu oxidation states and surface oxidation coupled with complex electrochemical conditions. Identifying active sites and understanding the interplay of various parameters influenced by thermal treatment or exposure to the electrochemical environment—such as surface oxidation, oxophilicity, and hydrophilicity—remain challenging.(30) Nonetheless, these results highlight the sensitivity of electrochemical methods as powerful tools for characterising the surface properties of such materials.(41)

For the samples reduced at 400 and 600 °C, in the presence of  $\text{NO}_3^-$ , a peak emerges between  $-0.3$  and  $-0.4$  V<sub>RHE</sub> due to the  $\text{NO}_3\text{RR}$ .(44) A steep increase follows this in the current reduction due to the HER. The hydrogen coverage on the surface of these Cu-based catalysts increases as the potential is scanned in the negative direction. However, while hydrogen adsorption is essential for forming products such as ammonia, excessive hydrogen coverage can block active sites (likely the Cu nanoparticles), thus inhibiting further  $\text{NO}_3\text{RR}$ . This blockage eventually leads to the re-emergence of HER.(45) Fig. 13(b) shows differences for the sample exsolved at 400 and 600 °C. The latter exhibits a lower current peak intensity and a shift of approximately 30 mV towards a more negative potential. The lower intensity is likely due to the larger size of the exsolved nanoparticles at this temperature (as shown in the SEM images in Fig. 8(b)), resulting in a lower surface area and reduced  $\text{NO}_3\text{RR}$  activity. Still, changes in the composition, either of the nanoparticles or in the support near the nanoparticles, cannot be discarded.

Finally, for the samples before exsolution or those reduced at 200 °C for 1 hour (Fig. 13(c)), we did not observe the characteristic peak of the  $\text{NO}_3\text{RR}$ . These materials are inert for this reaction under these conditions, and the currents observed follow the same trend as in the absence of  $\text{NO}_3^-$ , suggesting that in both cases, we are observing the response of the HER. However, a shift toward a more negative HER onset potential when  $\text{NO}_3^-$  ions are present indicates the presence of Cu species on the surface,(46) possibly in the form of Cu ions within the perovskite lattice surface and/or CuO secondary phase.

Furthermore, while the  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7}\text{O}_{3-\gamma}$  support would provide some level of mixed conductivity and oxygen-vacancy sites as earlier discussed, according to Fig. 10 (a – c), it shows no intrinsic  $\text{NO}_3\text{RR}$  activity in the absence of exsolved Cu (onset only at  $-0.20$  V RHE, no  $\text{NO}_3\text{RR}$  peak). Only upon Cu nanoparticle exsolution (400–600 °C) does a clear  $\text{NO}_3\text{RR}$  signal arise (Fig. 10 (b)), with the most positive

onset and highest currents seen for the finest particles (400 °C). Thus, although support interactions may modulate selectivity, the predominant enhancement in nitrate reduction performance is governed by the presence and size of the exsolved Cu<sup>0</sup> nanoparticles. Importantly, the electrochemical results are used here primarily as a proof-of-concept to demonstrate how controlled exsolution can be used to tune material properties rather than to optimize performance for nitrate reduction specifically.

## Experimental

## Material Synthesis

An A-site deficient  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7-x}\text{Cu}_x\text{O}_{3-\gamma}$  perovskite system with varied Cu-doping ( $x = 0, 0.05, 0.1, 0.15, 0.2$  mol) was designed and synthesised using the solid-state synthesis method. The following precursors were used:  $\text{SrCO}_3$  (99.99% purity, Thermo Fisher),  $\text{TiO}_2$  (99.5% purity, Sigma-Aldrich),  $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$  (98%, Alfa Aesar), and  $\text{Cu}(\text{NO}_3)_2 \cdot 2.5\text{H}_2\text{O}$  (98% purity, Thermo Fisher).

The solid-state synthesis (Fig. 1(a)) involved calcining dried precursor mixtures at 1000 °C, pellet formation and sintering in air at 1000 – 1100 °C. The synthesis temperature was selected based on the thermal behaviour of essential precursors— $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ ,  $\text{SrCO}_3$ , and  $\text{TiO}_2$ , as well as the dry mixture of these precursors, using thermogravimetry analysis (TGA) and differential scanning calorimetry (DSC). A Netzsch STA 449F3 equipment was used for the TGA. The TGA was run at 30 – 950 °C for  $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ ,  $\text{SrCO}_3$  and  $\text{TiO}_2$  at a 10 °C /min heating rate. Subsequently, the TGA was run at 30–1010 °C for the dried mixture of the precursors before calcination to  $\text{Sr}_{0.95}\text{Ti}_{0.3}\text{Fe}_{0.7}\text{O}_{3-\gamma}$ .

## Actual Perovskite Stoichiometry Determination

The actual Cu-doping amount,  $y$ , and the resulting perovskite stoichiometry were determined using Equation (2) (Equation (S3) in the Supporting Information).

$$y = \frac{KM_A}{M_B} \quad (2)$$

The derivation and parameters of Equation (S3) are explained in Note S1.

## Reduction of Sample

Reduction samples were prepared by breaking the sintered pellets into small pieces to have samples from the same pellet for different reduction temperatures and times. This ensured that any observed variation was not associated with factors other than the reduction parameters and the material's intrinsic properties. The prepared



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samples were reduced at different temperatures (400, 600, and 800 °C) in 5% H<sub>2</sub>/Ar (95%). All reduction process was done at the same reduction time of 1 hour. To investigate the effect of reduction time, samples were reduced at 600 °C for 1, 2, 3, and 4 hours respectively. All reduction was initiated at a room temperature of 25 °C, with a ramp rate of 5 °C/min.

### Crystal structure study

X-ray diffraction (XRD) analysis was employed to investigate the calcined and sintered perovskite crystal structure. A Rigaku equipment with Cu-K $\alpha$  radiation ( $\lambda$ : 1.5406 Å) was utilised. The scanning was performed with a step size of 0.02°, a scan rate of 3°/min, and a scanning range of 20–90°. The Rietveld refinement of the synthesised and reduced perovskites was done using the GSAS II software. The reference structures used for the refinement were Ti<sub>0.5</sub>Fe<sub>0.5</sub>SrO<sub>0.29</sub> (1525550) and CuO (1011148) for the as-sintered perovskite, while Ti<sub>0.5</sub>Fe<sub>0.5</sub>SrO<sub>0.29</sub> (1525550) and Cu (4105681) were used for the reduced perovskites. The unavailability of A-site deficient (Sr<sub>1-x</sub>Ti<sub>y</sub>Fe<sub>1-y</sub>O<sub>3-y</sub>) perovskite's CIF with Cu-doping on the B-site warranted the choice of the Ti<sub>0.5</sub>Fe<sub>0.5</sub>SrO<sub>0.29</sub> CIF. Also, the phase analysis on the Rigaku Miniflex suggested that the chosen structures could be phase matches for the respective perovskites.

### Microstructure study

Scanning electron microscopy (SEM) was used to study the microstructure of the calcined and sintered perovskite samples. All analyses were performed at the Geoanalytical Electron Microscopy and Spectroscopy (GEMS) facility within the School of Geographical and Earth Sciences at the University of Glasgow using Zeiss Sigma VP FEG SEM equipment. Furthermore, SEM analysis was also employed to investigate the reduced perovskites' microstructure and the exsolved nanoparticles' morphology. Sample imaging was performed using a secondary electron and in-lens detector, with an accelerating voltage of 3 kV, an aperture size of 30 μm, and a working distance (WD) of 1.4–3.2 mm. Samples were coated with a thin layer of gold/palladium (80/20 ratio), about 10 nm thick, to prevent charging due to the accumulation of static electric field when exposed to the electron beam. Furthermore, electron dispersive spectroscopy (EDS) was performed on the sample using the same instrument as the SEM. The setting parameters used for the SEM were an accelerating voltage of 20 kV, an aperture size of 60 μm, and a working distance (WD) of 5–9 mm.

### Reduction kinetics study

Thermogravimetry analysis under hydrogen gas (H<sub>2</sub>-TGA) was performed to investigate the reduction behaviour of the Sr<sub>0.95</sub>Ti<sub>0.3</sub>Fe<sub>0.7-x</sub>Cu<sub>x</sub>O<sub>3-y</sub> (x = 0–0.1) system. The samples were subjected to controlled heating at a heating rate of 5 °C/min in a 5% H<sub>2</sub>/Ar (95%) environment at 400, 600, and 800 °C. Each temperature step was maintained for a dwell time of 60 minutes, ensuring a comprehensive analysis. According to the literature, a regression fit of the weight loss data was performed using OriginPro Software to model the kinetic process.(32)

### Exsolved particle size analysis

ImageJ software was used for particle size analysis, with particle size distribution determined by Gaussian curve fitting in Origin Pro software. To determine the exsolved particle population density, particle size measurements were taken from three areas of each

sample, all with the exact dimensions. The population density was then determined by dividing the number of particles by the area and averaging.

### Surface characterisation

The XPS analysis data were collected using the Kratos analytical system, which utilised a monochromatic Al K $\alpha$  source. A filament current of 0.45 A and a filament bias of 1 V were used. Charge compensation was included in the data collection. Deconvolution of the core XPS spectra of the perovskite components was achieved using the KherveFitting software. All spectra deconvolution followed a specific spectra model in the KherveFitting Software library. A Voigt (Area, L/G,  $\sigma$ ) was used for the spectra of C 1s, O 1s, Ti 2p, Sr 3d, while a mixed Gaussian–Lorentzian line shape (GL30) was used for Fe 2p. The GL30 function was adopted in this study because it is widely recommended for oxides in the literature.(47,48) On the other hand, a mixed Gaussian–Lorentzian line shape (GL, as optimised during the fitting) was used for the spectra of Cu 2p. Furthermore, for Cu 2p spectra, the work of Torres-Ochoa et al guided most components and species identification according to BE.(49)

For the deconvolution of the high-resolution spectra (C 1s, O 1s, Ti 2p, Sr 3d, Fe 2p, and Cu 2p), background subtraction was performed using the Multi-region Smart algorithm implemented in the KherveFitting software. The Multi-region Smart combines a Shirley and linear background models to account for both step-like changes in intensity near the photoemission edges and gradual linear drifts across wider energy windows. Although the Tougaard-type backgrounds are often recommended in the literature for Fe and Cu core level spectra, as they can more accurately capture the extended inelastic loss features and multiplet-related structures that dominate these spectra.(50–52) However, the additional parameterisation and assumptions regarding the energy-loss function introduced by Tougaard, particularly the background type, can complicate fitting when multiple overlapping species are present. Hence, the *Multi-region Smart* background was selected for a consistent and reproducible approach across all spectral regions, while effectively minimising residuals and avoiding over-parameterisation. This choice ensured uniform treatment of all core levels analysed, thereby facilitating direct comparison of relative peak areas between elements and oxidation states. Charge referencing or carbon correction was done for all the samples using adventitious C 1s (284.8 eV) as recommended in the literature.(48)

Surface elemental ratio were determined from high-resolution XPS by RSF-corrected quantification of Sr, Ti, Fe, Cu, and O. Carbonate contributions were excluded, oxygen was considered on a Total-O basis (O-Lat + O-Vac + O-Surf), and shake-up satellites were omitted. The resulting atomic amounts were renormalised using a B-site normalisation scheme (Ti + Fe + Cu = 1) to yield elemental ratios relative to the transition-metal sublattice.

### X-ray Absorption Spectroscopy

The XAS analysis was performed at B18 beamline in Diamond Light Source. The spectra were obtained at the Cu and Fe K-edges, with a



varying energy step from 0.25 to 0.35 eV using the average from 5 scans. Data analysis was carried out using the Athena software. XANES spectra were normalized by applying a linear fit to the pre-edge region and a second-order polynomial after the edge. The fractions of Cu species were determined by linear combination analysis -15 eV before and +30 eV after the edge using Cu foil, Cu<sub>2</sub>O, and CuO as reference samples for Cu (0), (I) and (II), respectively, and the sample before exsolution. The EXAFS analysis consisted of  $k^2$ -weighted Fourier transform of the data over the  $k$ -range from 1.9 to 11.5 Å<sup>-1</sup>. The samples consisted of quartz capillary tubes sealed by Kapton wool and filled with the Cu-doped ( $x=0.1$  mol of Cu) perovskite powder under controlled atmosphere (4% H<sub>2</sub>/He) for exsolution at a ramp rate of 5°C/min. After exsolution, the spectra were collected under pure He atmosphere at room temperature without any exposure to air.

### Electrochemical activity study

The electrochemical measurements were performed using a Metrohm Autolab PGSTAT302N electrochemical workstation in a three-electrode glass cell configuration. The working electrode (WE) was prepared by drop-casting a dispersion ink onto a 5 mm diameter glassy carbon (GC) electrode. The ink was prepared by dispersing 5 mg of perovskite oxide powder in a mixture of 250 µL isopropanol, 247.5 µL ultrapure water (Milli-Q, 18.2 MΩ), and 2.5 µL of a 5% Nafion solution. The solution was sonicated for 20 minutes to ensure homogeneity. A 20 µL aliquot of the ink was deposited on the GC electrode and dried in an oven at 60°C for 15 minutes. Cyclic voltammetry experiments were conducted with a platinum mesh as the counter electrode and a reversible hydrogen electrode (RHE) as the reference electrode. The voltammograms were recorded at a sweep rate of 50 mV s<sup>-1</sup> over a potential range from 1.4 V to -0.7 V vs. RHE. Blank profiles were recorded in 0.1 mol L<sup>-1</sup> NaOH (Sigma-Aldrich®, 99.99 %), NO<sub>3</sub>RR CVs were carried out in 0.1 mol L<sup>-1</sup> NaOH with 1 mM NaNO<sub>3</sub> (Labsynth®, 99.0 %) as the nitrate source.

### Conclusion

This study systematically explored the impact of Cu doping on the structure, reduction kinetics, and exsolution behaviour of Sr<sub>0.95</sub>Ti<sub>0.3</sub>Fe<sub>0.7-x</sub>Cu<sub>x</sub>O<sub>3-y</sub> ( $x$ : 0 – 0.2) perovskites, aiming to optimise their properties for catalytic applications. Incorporating Cu significantly enhanced reduction kinetics, facilitating the controlled exsolution of Cu nanoparticles. XRD, Rietveld refinement, SEM, in-situ XAS and XPS analyses confirmed the successful synthesis and exsolution process. SEM and particle size analysis revealed that reduction at 600 °C for 3 hours yielded the highest nanoparticle population density (~650 particles/µm<sup>2</sup>), resulting in a higher density of active sites beneficial for catalytic performance. However, extended reduction times led to larger nanoparticles, presenting a trade-off between active site density and surface area availability. Furthermore, the in-situ XAS, in conjunction with the XPS result, further confirmed that Cu is the primary exsolved species, showing the partial reduction of Cu<sup>2+</sup> to Cu<sup>0</sup> and the emergence of Cu–Cu

coordination after exsolution, while the Fe environment remained largely unchanged. Overall, the combined electrochemical and XPS surface characterisation indicates that the exsolution temperatures significantly influenced the material's catalytic behaviour, with only samples treated at 400 °C and 600 °C showing NO<sub>3</sub>RR activity due to the presence of exsolved Cu nanoparticles. The 600 °C exsolved sample exhibited slightly lower activity than the 400 °C sample, which could be attributed to differences in particle size, morphology, population density, and composition. These results indicate a strong correlation between exsolution conditions and electrocatalytic activity, highlighting the trade-off between nanoparticle features and activity, where desirable performance requires balancing particle formation temperature with preserving a high surface area.

### Author Contributions

Ubong A. Essien: Investigation, Data curation, Formal analysis, Simulation, Visualization, Funding acquisition, Writing – review & editing, Writing – original draft. Swathi P. Raju: Investigation, Formal analysis, Writing – review & editing. Keyla T. Santos: Investigation, Formal analysis, Writing – review & editing. Rafael A. Vicente: Investigation, Formal analysis, Writing–review & editing. Chinyere Ekperechukwu: Methodology. Francisco R. García-García: Methodology, Supervision, Funding acquisition, Writing – review & editing. Pablo S. Fernández: Conceptualization, Methodology, Supervision, Funding acquisition, Writing – review & editing. Dragos Neagu: Conceptualization, Methodology, Supervision, Funding acquisition, Writing – review & editing.

### Conflict of Interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

### Data Availability

The data supporting this article have been included as part of the Supplementary Information. The crystal information files (1525550, 1011148, and 4105681) used for Rietveld refinement analysis in this study is available at crystallographic open database at <https://www.crystallography.net/cod/search>.

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1. Hu X, Sun Y, Guo S, Sun J, Fu Y, Chen S, et al. Identifying electrocatalytic activity and mechanism of  $\text{Ce}_1/3\text{NbO}_3$  perovskite for nitrogen reduction to ammonia at ambient conditions. *Appl Catal B*. 2021 Jan 1;280:119419.
2. Xu X, Chen Y, Zhou W, Zhu Z, Su C, Liu M, et al. A Perovskite Electrocatalyst for Efficient Hydrogen Evolution Reaction. *Advanced Materials* [Internet]. 2016 Aug 1 [cited 2024 Nov 4];28(30):6442–8. Available from: <https://onlinelibrary.wiley.com/doi/full/10.1002/adma.201600005>
3. Ji Q, Bi L, Zhang J, Cao H, Zhao XS. The role of oxygen vacancies of  $\text{ABO}_3$  perovskite oxides in the oxygen reduction reaction. *Energy Environ Sci*. 2020;13(5):1408–28.
4. Hu S, Zhang L, Cao Z, Yu W, Zhang P, Zhu X, et al. Cathode activation process and  $\text{CO}_2$  electroreduction mechanism on  $\text{LnFeO}_3\text{-}\delta$  ( $\text{Ln}=\text{La, Pr and Gd}$ ) perovskite cathodes. *J Power Sources*. 2021 Feb 15;485:229343.
5. Zhang W, Gao C, Zhang P, Feng Q, Zhang L, Zhang X, et al.  $\text{LaNi}_0.2\text{Cu}_0.8\text{O}_3\text{-}\delta$  Perovskites as Electrocatalysts for Electro-synthesis of Ammonia from Nitrogen Oxyanion. *ChemCatChem* [Internet]. 2024 Aug 12 [cited 2024 Nov 2];16(15):e202400185. Available from: <https://onlinelibrary.wiley.com/doi/full/10.1002/cctc.202400185>
6. Yang WJ, Yang LH, Peng HJ, Lv SH, Muhammad Adeel Sharif H, Sun W, et al. Perovskite oxide  $\text{LaMO}_3\text{-}\delta$  ( $\text{M}=\text{Fe, Co, Ni and Cu}$ ) cathode for efficient electroreduction of nitrate. *Sep Purif Technol*. 2022 Aug 15;295:121278.
7. Berger T, Drexler H, Ruh T, Lindenthal L, Schrenk F, Bock J, et al. Cu-doped perovskite-type oxides: A structural deep dive and examination of their exsolution behaviour influenced by B-site doping. *Catal Today* [Internet]. 2024 Jul 1 [cited 2024 May 13];437:114787. Available from: <https://linkinghub.elsevier.com/retrieve/pii/S0920586124002815>
8. Mohammadi A, Hejny C, Kahlenberg V, Gallmetzer JM, Purtscher FRS, Hofer TS, et al. Copper-perovskite interfacial engineering to boost deNOx activity. *Chemical Engineering Journal*. 2024 Oct 15;498:155425.
9. Myung J ha, Neagu D, Miller DN, Irvine JTS. Switching on electrocatalytic activity in solid oxide cells. *Nature* [Internet]. 2016;537(7621):528–31. Available from: <https://doi.org/10.1038/nature19090>;10.1039/D5NA00426H
10. Weber ML, Jennings D, Fearn S, Cavallaro A, Prochazka M, Gutsche A, et al. Thermal stability and coalescence dynamics of exsolved metal nanoparticles at charged perovskite surfaces. 2023 Dec 26 [cited 2024 Mar 20]; Available from: <https://chemrxiv.org/engage/chemrxiv/article-details/6585de6666c1381729cfabb8>
11. Irvine J, Rupp JLM, Liu G, Xu X, Haile S, Qian X, et al. Roadmap on inorganic perovskites for energy applications. *Journal of Physics: Energy* [Internet]. 2021;3(3):31502. Available from: <http://dx.doi.org/10.1088/2515-7655/abff18>
12. Kousi K, Tang C, Metcalfe IS, Neagu D. Emergence and Future of Exsolved Materials. *Small*. 2021;17(21).
13. Kim YH, Jeong H, Won BR, Jeon H, Park CH, Park D, et al. Nanoparticle Exsolution on Perovskite Oxides: Insights into Mechanism, Characteristics and Novel Strategies. *Nanomicro Lett*. 2024;16(1).
14. Ruh T, Berkovec D, Schrenk F, Rameshan C. Exsolution on perovskite oxides: morphology and anchorage of nanoparticles. This journal is Cite this: *Chem Commun* [Internet]. 2023 [cited 2023 Dec 18];59:3948. Available from: <https://www.webofscience.com>
15. Sun Z, Hao C, Toan S, Zhang R, Li H, Wu Y, et al. Recent advances in exsolved perovskite oxide construction: exsolution theory, modulation, challenges, and prospects. *J Mater Chem A Mater* [Internet]. 2023 Aug 29 [cited 2024 Dec 18];11(34):17961–76. Available from: <https://pubs.rsc.org/en/content/articlehtml/2023/ta/d3ta03292b>
16. Zhang SL, Cox D, Yang H, Park BK, Li CX, Li CJ, et al. High stability  $\text{SrTi}_{1-x}\text{Fe}_x\text{O}_3\text{-}\delta$  electrodes for oxygen reduction and oxygen evolution reactions. *J Mater Chem A Mater* [Internet]. 2019;7(37):21447–58. Available from: <http://dx.doi.org/10.1039/C9TA07548H>
17. Yang H, Hanif MB, Zhang SL, Li CJ, Li CX. Sintering behavior and electrochemical performance of A-site deficient  $\text{Sr}_x\text{Ti}_0.3\text{Fe}_0.7\text{O}_3\text{-}\delta$  oxygen electrodes for solid oxide electrochemical cells. *Ceram Int* [Internet]. 2021;47(17):25051–8. Available from: <https://www.sciencedirect.com/science/article/pii/S027284221016357>
18. Shao Z, Halle SM. A high-performance cathode for the next generation of solid-oxide fuel cells. *Nature* 2004 431:7005 [Internet]. 2004 Sep 9 [cited 2024 Jan 9];431(7005):170–3. Available from: <https://www.nature.com/articles/nature02863>
19. Jennings D, Zahler MP, Wang D, Ma Q, Deibert W, Kindelmann M, et al. Grain boundary segregation in iron doped strontium titanate: From dilute to concentrated solid solutions. *Acta Mater*. 2024 Jul 1;273:119941.
20. Kharton V V, Kovalevsky A V, Viskup AP, Jurado JR, Figueiredo FM, Naumovich EN, et al. Transport Properties



## Journal Name

## ARTICLE

and Thermal Expansion of  $\text{Sr}_0.97\text{Ti}_{1-x}\text{Fe}_x\text{O}_3-\delta$  ( $x=0.2-0.8$ ). *J Solid State Chem* [Internet]. 2001;156(2):437–44. Available from: <https://www.sciencedirect.com/science/article/pii/S0022459600990194>

21. Xie K, Tao S, Tao S. Conductivity and redox stability of new perovskite oxides  $\text{SrFe}_{0.7}\text{TM}_{0.2}\text{Ti}_{0.1}\text{O}_3-\delta$  ( $\text{TM} = \text{Mn, Fe, Co, Ni, Cu}$ ). *Solid State Ion*. 2017 Mar 1;301:99–105.

22. Kharton V V., Viskup AP, Kovalevsky A V., Jurado JR, Naumovich EN, Vecher AA, et al. Oxygen ionic conductivity of Ti-containing strontium ferrite. *Solid State Ion*. 2000 Aug 1;133(1–2):57–65.

23. Neagu D, Tsekouras G, Miller DN, Ménard H, Irvine JTS. In situ growth of nanoparticles through control of non-stoichiometry. *Nat Chem* [Internet]. 2013;5(11):916–23. Available from: <https://doi.org/10.1038/nchem.1773>

24. Neagu D, Oh TS, Miller DN, Ménard H, Bukhari SM, Gamble SR, et al. Nano-socketed nickel particles with enhanced coking resistance grown in situ by redox exsolution. *Nat Commun* [Internet]. 2015;6(1):8120. Available from: <https://doi.org/10.1038/ncomms9120>

25. Neagu D, Irvine JTS, Wang J, Yildiz B, Opitz AK, Fleig J, et al. Roadmap on exsolution for energy applications. *Journal of Physics: Energy* [Internet]. 2023 Jun 20 [cited 2024 Jan 2];5(3):031501. Available from: <https://iopscience.iop.org/article/10.1088/2515-7655/acd146>

26. Darwish E, Mansouri M, Yilmaz D, Leion H. Effect of Mn and Cu Substitution on the  $\text{SrFeO}_3$  Perovskite for Potential Thermochemical Energy Storage Applications. *Processes* 2021, Vol 9, Page 1817 [Internet]. 2021 Oct 13 [cited 2024 Jul 9];9(10):1817. Available from: <https://www.mdpi.com/2227-9717/9/10/1817/htm>

27. Liu J, Wen Z, Cui Y, Zhao Y, Chen Y. The Spin Glass State and Exchange Bias Effect in  $\text{SrFe}_{1-x}\text{Cu}_x\text{O}_3-\delta$ . *J Supercond Nov Magn* [Internet]. 2022 Nov 1 [cited 2024 Jul 9];35(11):3403–9. Available from: <https://link.springer.com/article/10.1007/s10948-022-06360-z>

28. Colpani GL, Zeferino RCF, Zanetti M, Mello JMM, Silva LL, Fiori MA. Cation-modified photocatalysts. Photocatalytic Systems by Design: Materials, Mechanisms and Applications. 2021 Jan 1;23–53.

29. Mejia-Rodriguez D, Kunitsa AA, Fulton JL, Aprà E, Govind N. GOWO Ionization Potentials of First-Row Transition Metal Aqua Ions. *Journal of Physical Chemistry A* [Internet]. 2023 Nov 23 [cited 2024 Dec 14];127(46):9684–94. Available from: <https://pubs.acs.org/doi/abs/10.1021/acs.jpca.3c04419>

30. Zhong X, Yuan E, Yang F, Liu Y, Lu H, Yang J, et al. Optimizing oxygen vacancies through grain boundary engineering to enhance electrocatalytic nitrogen reduction. *Proc Natl Acad Sci U S A* [Internet]. 2023 Oct 3 [cited 2025 Jan 14];120(40):e2306673120. Available from: <https://www.pnas.org/doi/abs/10.1073/pnas.2306673120>

31. Weckhuysen BM, De Smit E, De Groot FMF, Blume R, Hävecker M, Knop-Gericke A. The role of Cu on the reduction behavior and surface properties of Fe-based Fischer–Tropsch catalysts. *Physical Chemistry Chemical Physics* [Internet]. 2009 Dec 22 [cited 2024 Jul 1];12(3):667–80. Available from: <https://pubs.rsc.org/en/content/articlehtml/2010/cp/b920256k>

32. McColm TD, Irvine JTS. Structural and property investigations of Strontium Galloniobate. *Solid State Ion*. 2002 Dec 1;152–153(153):615–23.

33. Neagu D, Irvine JTS. Structure and properties of  $\text{La}_0.4\text{Sr}_0.4\text{TiO}_3$  Ceramics for use as anode materials in solid oxide fuel cells. *Chemistry of Materials* [Internet]. 2010 Sep 14 [cited 2024 Jan 23];22(17):5042–53. Available from: <https://pubs.acs.org/doi/full/10.1021/cm101508w>

34. Koo B, Kim K, Kim JK, Kwon H, Han JW, Jung W. Sr Segregation in Perovskite Oxides: Why It Happens and How It Exists. *Joule* [Internet]. 2018;2(8):1476–99. Available from: <https://www.sciencedirect.com/science/article/pii/S254235118303271>

35. Bachelet R, Sánchez F, Palomares FJ, Ocal C, Fontcuberta J. Atomically flat  $\text{SrO}$ -terminated  $\text{SrTiO}_3$  (001) substrate. *Appl Phys Lett*. 2009 Jan 1;95(14):141915.

36. Szot K, Speier W, Breuer U, Meyer R, Szade J, Waser R. Formation of micro-crystals on the (100) surface of  $\text{SrTiO}_3$  at elevated temperatures. *Surf Sci*. 2000 Jul 20;460(1–3):112–28.

37. Zhang Y, Xu Y, Chen Z, Zhang Z, Liu X, Xue Z, et al. Self-Assembled Controllable Cu-Based Perovskite/Calcium Oxide Hybrids with Strong Interfacial Interactions for Enhanced  $\text{CH}_4$  Electrosynthesis. *ACS Nano* [Internet]. 2024 Nov 12 [cited 2025 Nov 24];18(45):31466–77. Available from: <https://pubs.acs.org/doi/abs/10.1021/acsnano.4c11459>

38. Climent V, Feliu JM. Thirty years of platinum single crystal electrochemistry. *Journal of Solid State Electrochemistry* 2011 15:7 [Internet]. 2011 Apr 12 [cited 2025 Mar 4];15(7):1297–315. Available from: <https://link.springer.com/article/10.1007/s10008-011-1372-1>

39. Zhang R, Zhang S, Cui H, Guo Y, Li N, Zhi C. Electrochemical nitrate reduction to ammonia using copper-based electrocatalysts. *Next Energy*. 2024 Jul 1;4:100125.

40. Fu X, Zhao X, Hu X, He K, Yu Y, Li T, et al. Alternative route for electrochemical ammonia synthesis by reduction of nitrate on copper nanosheets. *Appl Mater Today*. 2020 Jun 1;19:100620.

41. Farinazzo Bergamo Dias Martins P, Papa Lopes P, Ticianelli EA, Stamenkovic VR, Markovic NM, Strmcnik D. Hydrogen evolution reaction on copper: Promoting water dissociation



## ARTICLE

## Journal Name

by tuning the surface oxophilicity. *Electrochem commun.* 2019 Mar 1;100:30–3.

42. Sheng W, Myint M, Chen JG, Yan Y. Correlating the hydrogen evolution reaction activity in alkaline electrolytes with the hydrogen binding energy on monometallic surfaces. *Energy Environ Sci* [Internet]. 2013 Apr 24 [cited 2024 Nov 11];6(5):1509–12. Available from: <https://pubs.rsc.org/en/content/articlehtml/2013/ee/c3ee00045a>

43. Alom MS, Kananke-Gamage CCW, Ramezanipour F. Perovskite Oxides as Electrocatalysts for Hydrogen Evolution Reaction. *ACS Omega* [Internet]. 2022 Mar 8 [cited 2024 May 31];7(9):7444–51. Available from: <https://pubs.acs.org/doi/full/10.1021/acsomega.1c07203>

44. Pérez-Gallent E, Figueiredo MC, Katsounaros I, Koper MTM. Electrocatalytic reduction of Nitrate on Copper single crystals in acidic and alkaline solutions. *Electrochim Acta*. 2017 Feb 10;227:77–84.

45. Reyter D, Bélanger D, Roué L. Study of the electroreduction of nitrate on copper in alkaline solution. *Electrochim Acta*. 2008 Aug 20;53(20):5977–84.

46. Hu T, Wang C, Wang M, Li CM, Guo C. Theoretical Insights into Superior Nitrate Reduction to Ammonia Performance of Copper Catalysts. *ACS Catal* [Internet]. 2021 Dec 3 [cited 2024 Dec 28];11(23):14417–27. Available from: <https://pubs.acs.org/doi/abs/10.1021/acscatal.1c03666>

47. Ghaffari M, Shannon M, Hui H, Tan OK, Irannejad A. Preparation, surface state and band structure studies of  $\text{SrTi}(1 - x)\text{Fe}(x)\text{O}(3 - \delta)$  ( $x = 0\text{--}1$ ) perovskite-type nano structure by X-ray and ultraviolet photoelectron spectroscopy. *Surf Sci* [Internet]. 2012 Mar 1 [cited 2025 Jun 10];606(5–6):670–7. Available from: <https://www.sciencedirect.com/science/article/pii/S0039602811004821?via%3Dihub>

48. Biesinger MC, Lau LWM, Gerson AR, Smart RSC. Resolving surface chemical states in XPS analysis of first row transition metals, oxides and hydroxides: Sc, Ti, V, Cu and Zn. *Appl Surf Sci*. 2010 Nov 15;257(3):887–98.

49. Torres-Ochoa JA, Cabrera-German D, Cortazar-Martinez O, Bravo-Sanchez M, Gomez-Sosa G, Herrera-Gomez A. Peak-fitting of Cu 2p photoemission spectra in Cu0, Cu1+, and Cu2+ oxides: A method for discriminating Cu0 from Cu1+. *Appl Surf Sci* [Internet]. 2023 Jun 15 [cited 2025 Aug 4];622:156960. Available from: <https://www.sciencedirect.com/science/article/pii/S0169433223006360>

50. Seah MP. Background subtraction: I. General behaviour of Tougaard-style backgrounds in AES and XPS. *Surf Sci* [Internet]. 1999 Jan 20 [cited 2025 Aug 19];420(2–3):285–94. Available from: <https://www.sciencedirect.com/science/article/pii/S0039602898008528>

51. Lee D, Park S. Proper spectroscopic analysis of transition metal oxides using ex-situ X-ray photoelectron spectroscopy. *Current Applied Physics* [Internet]. 2024 Aug 1 [cited 2025 Aug 19];64:25–33. Available from: <https://www.sciencedirect.com/science/article/pii/S156773924000968>

52. Engelhard MH, Baer DR, Herrera-Gomez A, Sherwood PMA. Introductory guide to backgrounds in XPS spectra and their impact on determining peak intensities. *Journal of Vacuum Science & Technology A: Vacuum, Surfaces, and Films* [Internet]. 2020 Dec 1 [cited 2025 Aug 19];38(6):63203. Available from: <https://avs.jva/article/38/6/063203/1024050/Introductory-guide-to-backgrounds-in-XPS-spectra>



## Data Availability Statement

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The data supporting this article have been included as part of the Supplementary Information. The crystal information files (1525550, 1011148, and 4105681) used for Rietveld refinement analysis in this study is available at crystallographic open database at <https://www.crystallography.net/cod/search>.