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First-principles investigations of As-doped tetragonal boron nitride nanosheets for toxic gas sensing applications

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Pristine and arsenic-doped tetragonal boron nitride nanosheets (BNNS and As-BNNS) have been reported as potential candidates for toxic gas sensing applications. We have investigated the adsorption behavior of BNNS and As-BNNS for CO₂, H₂S, and SO₃ gas molecules using first-principles density functional theory (DFT). Both BNNS and As-BNNS possess negative cohesive energies of -8.47 and -8.22 eV, respectively, which indicates that both sheets are energetically stable. Successful adsorption is inferred from the negative adsorption energy and structural deformation in the vicinity of the adsorbent and adsorbate. As-doping results in a significant increase in adsorption energies from -0.094 , -0.175 , and -0.462 eV to -2.748 , -2.637 , and 3.057 eV for CO₂, H₂S and SO₃ gases, respectively. Due to gas adsorption, the electronic bandgap in As-BNNS varies by approximately 32% compared to a maximum of 24% in BNNS. A notable fluctuation in the energy gap and electrical conductivity is seen, with ambient temperature being the point of maximal sensitivity. For SO₃, the maximum charge transfer during adsorption in BNNS and As-BNNS is determined to be $0.08|e|$ and $0.25|e|$, respectively. Due to the interaction with gases, all structures exhibit an extremely high absorption coefficient on the order of 10^4 cm⁻¹ with minimal peak shifting. Additionally, doping an As atom on BNNS' surface remarkably improved its ability to sense CO₂, H₂S, and SO₃ gasses.

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1. Introduction

In today's cities, rapid industrialization has led to substantial pollution from byproducts such as hazardous gasses and heavy metals, among others. Toxic gases and other vapor constituents can be identified with the help of gas-detecting technologies. The gas sensor device applications include pollution control, public and military safety, production, agriculture, tracking the environment, and healthcare diagnosis.¹⁻⁴ Scientists report the presence of very poisonous gases, such as CH₄, CO, SO₂, CO₂, NO, NH₃, CH₃OH, H₂S, O₃, PH₃, and COCl₂.⁵⁻⁸ Among these, H₂S is a poisonous, explosive, and acidic gas that can seriously and irreversibly damage the nervous system.⁹ SO₃ vapor is the most harmful to breathe in and can seriously burn the stomach, esophagus, and mouth.¹⁰ Furthermore, CO₂ impacts human health due to its basic asphyxiant properties and is only slightly harmful when inhaled.¹¹ The atmospheric pollutant CO₂ is

largely created by the ongoing expansion of vehicles and businesses, and it has a detrimental effect on our way of life and health.¹² Different adsorbents have been employed to study the sensing mechanism of CO₂, H₂S, and SO₃ gases.¹³⁻¹⁶

Gas sensing technology is therefore greatly impacted by the search for more affordable, ecologically safe, and effective sensors. Many materials, including conducting, semi-conducting, and nanomaterials like graphene, have been proposed by research groups as possible gas sensors.¹⁷⁻²² Since the discovery of graphene, two-dimensional (2D) materials have drawn a lot of attention. However, because of its poor sensitivity, we must discover new substitute 2D materials for gas sensing.²³ For gas sensing applications, graphene 2D sheets with various crystalline structures—such as hexagonal and tetragonal—are readily accessible and often utilized.²⁴⁻²⁷ Similar to graphene, boron nitride (BN) may exist in phases like the recently created tetragonal BN and hexagonal graphene-type BN (g-BN or h-BN).^{28,29} The h-BN nanosheets, sometimes called white graphene, show great promise due to their ability to sense gases.³⁰⁻³³ BN layers have generally been synthesized in the hexagonal phase and used as gas sensors because of their extraordinary structural and chemical properties, huge band gap, and amazing thermal stability.³²⁻³⁴ Numerous gas molecules, including NO₂, NO, NH₃, CO, CH₄, H₂, and others, have been studied for their adsorption on the surface of h-BN using

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both computational and experimental approaches.^{35–40} However, intrinsic 2D h-BN has limited reactivity for some inert gases.⁴¹

The most common method of improving the sensing performance on BN layers is element doping. Al and Ga-doped BN sheets have shown good halomethane sensitivity, whereas Al and Ga-doped BN nanotubes had substantial contact with NH₃ gas.^{42,43} Higher adsorption energy has been observed in Co and Mn-doped BN layers when they interact more strongly with the gases CH₄, H₂S, NH₃, O₃, PH₃, and SO₂.⁵ For the adsorption of common SF₆ decomposition gases, such as H₂S, SO₂, SOF₂, and SO₂F₂, Sheng-Yuan Xia *et al.* suggested a Rh-doped h-BN (Rh-BN) monolayer as the perfect sensor.⁴⁴ Moladoust *et al.* have created an Al and Si-doped BN nanostructure that exhibited greater adsorption energy for phosgene gas than a pure BN structure.⁴⁵ An As-doped BN monolayer for SO₂F₂ gas molecules demonstrated good adsorption specificity, high sensitivity, and a short recovery time, as demonstrated by Yunfeng Long *et al.*⁴⁶ Tetragonal BN nanosheets (BNNS) are one of the boron nitride allotropes whose physical and adsorption properties have been studied the least. Sakib *et al.* used DFT to study the adsorption properties of BNNS in the tetragonal phase and discovered negative adsorption energy for thioguanine, an anticancer drug.⁴⁷ Tetragonal BNNS has a large semiconducting bandgap, which has been observed, and it has been shown that electrical conductivity changes by 8% under compressive stress.⁴⁸

Significant research attention has been drawn to the novel applications of an emerging class of 2D materials for ecologically hazardous gas sensing applications. We investigated As-doped tetragonal boron nitride nanosheets as a new kind of adsorbent for use in gas sensing. Due to the scarcity of studies on how the adsorption properties are affected by As atom doping in pure BNNS, we have selected this dopant. Using the DFT approach, we have computed the structural, adsorption, and optoelectronic characteristics of As-doped BNNS and pristine BNNS for CO₂, H₂S, and SO₃ gas adsorption.

2. Computational framework

Using DFT and the Cambridge Serial Total Energy Package (CASTEP) algorithm, first-principles calculations of the geometrical, optical, and electronic response of pristine BNNS and As-BNNS during the adsorption of CO₂, H₂S, and SO₃ gas molecules were performed in this study.⁴⁹ Using the Broyden–Fletcher–Goldfarb–Shanno (BFGS) minimization technique, all atoms can be relaxed during the optimization process.⁵⁰ The wave function of valence electrons is simulated by the CASTEP algorithm using a combination of plane wave basis sets with kinetic energy less than the cut-off energy.⁵¹ The exchange–correlation potential is considered in the Perdew–Burke–Ernzerhof (PBE) functional of the Generalized Gradient Approximation (GGA-PBE)⁵² functional.

Optimizing the pristine BNNS structure, the impact of many plane-wave cutoff energies—350 eV, 400 eV, 450 eV, and 500 eV—was investigated; the lowest energy was reached at 500 eV. Subsequently, ultra-soft pseudo-potentials⁵² with a kinetic energy cutoff of 500 eV were employed throughout the

simulation. Here, we have performed Brillouin zone integration for structural optimization of unit cells using a 6 × 6 × 1 mesh Monkhorst–Pack grid.⁵³ To have enough space for adsorption of gas molecules, we built a 3 × 2 × 1 supercell with a 20 Å vacuum space along the BNNS structure's z-direction. This allowed us to sample the Brillouin zone using a 2 × 3 × 1 *k*-point Monkhorst–Pack mesh. In the whole study, Grimme's dispersion correction scheme has been employed to take the long-range van der Waals interaction into account.⁵⁴ For all geometric optimizations, the convergence tolerance conditions are set at fine quality. The self-consistent estimation relies on the self-consistent convergence of the total energy of 1 × 10^{−5} eV per atom, the maximum force on the atom of 0.03 eV Å^{−1}, the maximum ionic displacement within 0.001 Å, and the maximum stress within 0.05 GPa.

Electronic band structure, density of states, electronic density difference, Mulliken charge population, and optical characteristics were all carefully examined to thoroughly characterize optoelectronic responses. The B-2s² 2p¹, N-2s² 2p³, H-1s¹, S-3s² 3p⁴, As-4s² 4p³, C-2s² 2p², and O-2s² 2p⁴ valence electron configurations were employed for pseudo atomic calculations. To compute optical properties, the complex dielectric constant was evaluated using the Kramer–Kronig relationship.⁵⁵ Since it provides access to new optical parameters, the complex dielectric function is a crucial component of optical properties. To study the sensitivity of the nanosheets toward gas molecules, the adsorption energy (E_{ads}) was calculated using the following equation⁵⁶

$$E_{\text{ads}} = E_{\text{BNNS+gas}} - E_{\text{BNNS}} - E_{\text{gas}} \quad (1)$$

where the energy of the gas-adsorbed nanosheet, the nanosheet (without the gas molecule), and the gas molecule is represented, respectively, by the symbols $E_{\text{BNNS+gas}}$, E_{BNNS} , and E_{gas} . The negative adsorption energy³³ confirms the stability of the adsorbed gas on the nanosheets. Nonetheless, the adsorption locator module was utilized to achieve the steady adsorption configuration.^{39–41}

Pristine BNNS and As-BNNS stability was examined by calculating their binding energy or cohesive energy using the following equation⁵⁷

$$E_{\text{coh}} = \frac{E_{\text{(BNNS/As-BNNS)}} - (E_{\text{B}} + E_{\text{N}} + E_{\text{As}})}{n} \quad (2)$$

where $E_{\text{BNNS/As-BNNS}}$, E_{B} , E_{N} , E_{As} , and n are the energy of BNNS or As-BNNS, isolated boron, isolated nitrogen, isolated arsenic atom, and the total number of atoms respectively.

3. Results and discussion

3.1 Geometry analysis

The optimized BNNS, As-BNNS, and gas adsorbed-BNNS structures are shown in Fig. 1 in both side and top views. Table 1 lists the average bond lengths for gas molecules, As-BNNS, and pure BNNS. The B–N bond length in pristine BNNS is stated to be 1.449 Å, which is similar to the earlier finding.⁴⁷ With 48 atoms, the 2D pure BNNS forms 8 hexagons



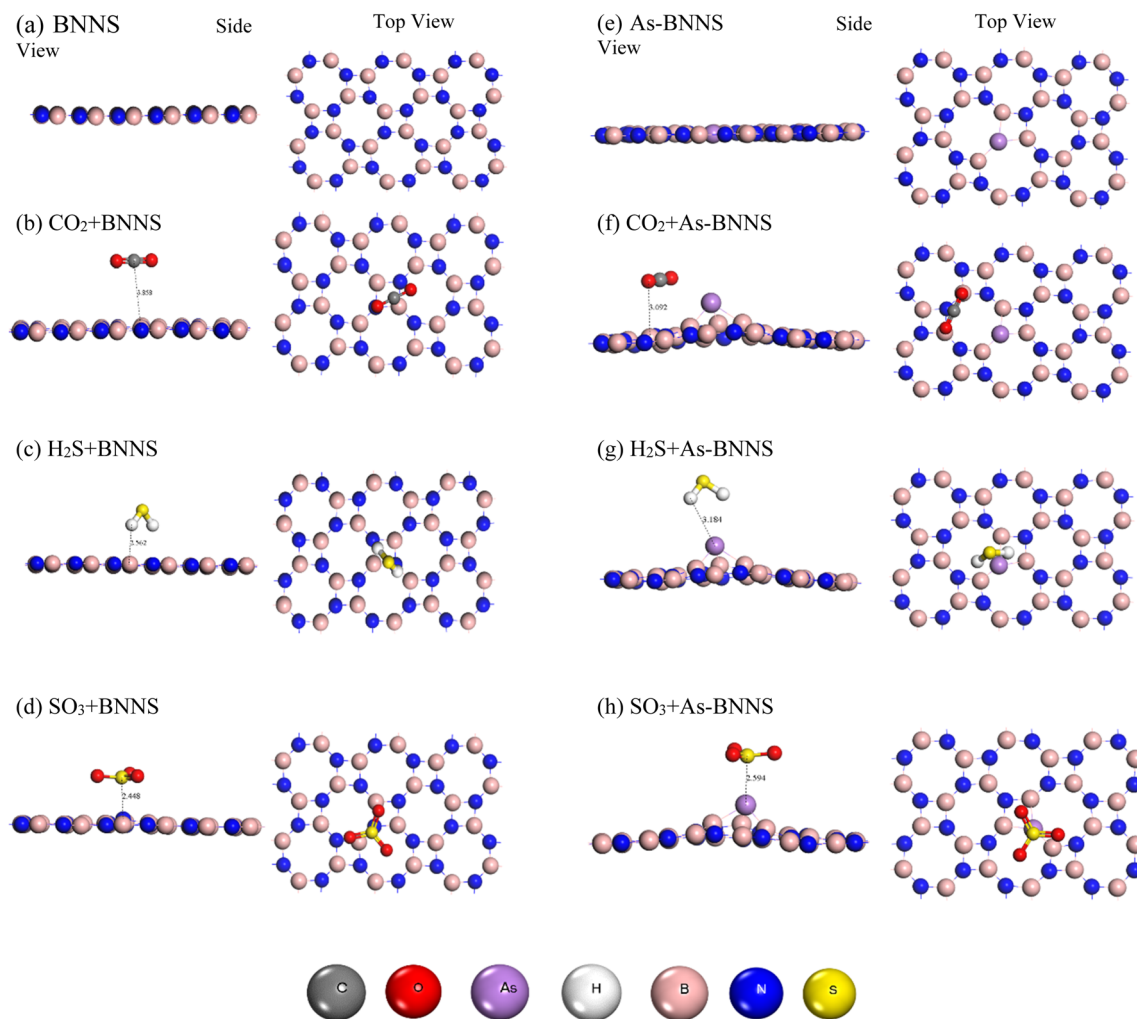


Fig. 1 Optimized structures (side and top views) of pristine BNNS, As-BNNS, and gas adsorbed BNNS and As-BNNS.

Table 1 Bond length (Å) and cohesive energy (E_{coh}) of the optimized structures

Molecules	Bond lengths					E_{coh} (eV)
	B-N	As-B	C=O	H-S	S=O	
BNNS	1.449	—	—	—	—	-8.47
As-BNNS	1.508	1.867, 1.794	—	—	—	-8.22
CO ₂	—	—	1.181	—	—	—
H ₂ S	—	—	—	1.351	—	—
SO ₃	—	—	—	—	1.438	—

across the layers. We observed modifications in the bond length due to arsenic doping instead of nitrogen. The lengths of the As-B bonds are 1.867 Å and 1.794 Å. This might result from the interaction of the arsenic atom with its varied size and charge distribution.⁴⁶ The bond lengths in CO₂, H₂S, and SO₃ optimized gas molecules are 1.181 Å, 1.351 Å, and 1.438 Å for C=O, H-S, and S=O respectively. These findings are consistent with the earlier research.^{12,58-62} We observed a slight alteration in the bond length of both the gas molecules and BNNS following the

gas adsorption. The length of the B-N bond varies from 1.79% to 4.07%, which might be a key sign of how interactive BNNS is with these gas molecules. In CO₂ + BNNS, H₂S + BNNS, and SO₃ + BNNS, the B-N bond lengths are 1.475 Å, 1.477 Å, and 1.508 Å, in that order. The measured modifications in gas molecules are 1.352 Å and 1.442 Å for the H-S and S=O bonds, respectively. The change in the average bond length of the adsorbent and adsorbate signifies the interaction between them.^{62,63} Compared to H₂S and CO₂, a greater interaction with SO₃ is shown by the greater alteration in the B-N bonds caused by SO₃ adsorption. The gas adsorption in the instance of As-BNNS modifies the B-N bond length from 2.01% to 2.12%. This demonstrates that As-BNNS is interactive with gas molecules. However, because of the 8.78% to 9.75% increase in As-B bond length, the interaction is higher in the presence of an arsenic atom when compared to BNNS. We measured the C=O, H-S, and S=O bond lengths in gas-adsorbed As-BNNS to be 1.181 Å, 1.351 Å, and 1.446 Å, respectively.

This suggests that a greater variation in the S=O bond length causes As-BNNS to interact more with SO₃. One of the primary issues was the geometrical stability following the



arsenic doping in the pure BNNS layer. Using eqn (2), we have determined the cohesive energy of the adsorbents to elucidate the stability. The energetic stability of BNNS and As-BNNS is shown by the negative values of E_{coh} , which were determined to be -8.47 and -8.22 eV, respectively. Hence, the tetragonal BNNS remains stable after As-doping with a slight distortion in the geometrical structure.⁴⁶

A traditional method for investigating the adsorption performance and longer lifespan of any sensor device is the deformation study of the adsorbent layer. As a result, we use the following equation (eqn (3)) to determine the deformation of the BNNS and As-BNNS layers caused by gas adsorption by measuring the energy of the adsorbents that are deformed and undeformed.

$$E_{\text{def}} = \frac{E_{\text{Deformed}} - E_{\text{Undeformed}}}{n} \quad (3)$$

where $E_{\text{Undeformed}}$ is the energy of the adsorbent layer and E_{Deformed} is the energy of the deformed adsorbent after gas adsorption. After gas adsorption we removed the gas molecules and the energy of the deformed adsorbent has been calculated. A sensor device with a longer lifespan is more suitable the less the adsorbent deforms when there is an adsorbate present. When CO_2 is adsorbed, the pristine BNNS shows the least amount of deformation. When H_2S and SO_3 are adsorbed, the deformation tends to increase and is greatest for SO_3 . This suggests that, in comparison to alternative options, using BNNS as a CO_2 gas sensor is adequate. The SO_3 adsorption in the As-BNNS instance demonstrates the least amount of base layer deformation, with a deformation energy of 0.051 eV. The As-BNNS experiences a comparable deformation with a deformation energy of 0.053 eV when CO_2 and H_2S are adsorbed. This suggests that As-BNNS is a suitable and longer-lasting SO_3 gas sensor. It is observed that in the presence of gas molecules, the deformation of BNNS is smaller and varies more than that of As-BNNS.

3.2 Adsorption of CO_2 , H_2S and SO_3 gases

When it has a negative value, adsorption energy describes the direct interaction of base layers with gas molecules.⁶⁴ All of the gas molecules are physically adsorbed without creating covalent bonds with the adsorbent, as shown in Fig. 1 and Table 2. We have obtained the adsorption energy and adsorption length as shown in Table 3. All of the gas molecules show negative

Table 3 Adsorption energy (E_{ad}) and adsorption length of the BNNS and As-BNNS sheets for gas molecules

Adsorbed gas	E_{ad} (eV)		Adsorption length (Å)	
	BNNS	As-BNNS	BNNS	As-BNNS
CO_2	-0.094	-2.748	3.858	3.092
H_2S	-0.175	-2.637	2.562	3.184
SO_3	-0.462	-3.057	2.448	2.594

adsorption energy in BNNS, suggesting that they have successfully adsorbed on the material.

Following adsorption, the BNNS planar structure is retained with less deformation. These findings are comparable to research on deformation energy. Comparatively, SO_3 with a maximum adsorption energy of -0.462 eV shows that SO_3 gas is adsorbed more strongly than the others. The adsorption energy finding is followed by the minimum adsorption length for SO_3 . For CO_2 , H_2S , and SO_3 gas molecules, however, the comparable adsorption energy and adsorption length indicated that BNNS is an effective adsorbent.

The BNNS layer's planar structure is somewhat distorted when one arsenic atom is doped. Furthermore, the adsorption of gas molecules enhances this deformation, brought about by Table 2's higher deformation energy. Table 3 demonstrates how arsenic doping improved the adsorption characteristics in As-BNNS relative to BNNS, with greater adsorption energy. As-BNNS, like BNNS, shows maximum adsorption for SO_3 gas at a minimum adsorption length of 2.594 Å and a higher negative value of E_{ad} of -3.057 eV. Overall, the BNNS and As-BNNS are suitable adsorbents for CO_2 , H_2S , and SO_3 gas molecules.

We have investigated the recovery time by following the equation for BNNS and As-BNNS after the CO_2 , H_2S , and SO_3 gas adsorption.⁶⁵

$$\tau = \frac{1}{f_0} e^{-\frac{E_{\text{ad}}}{KT}} \quad (4)$$

where K , T , and f_0 stand for Boltzmann's constant, temperature and UV radiation frequency (10^{12} to 3×10^{14} Hz), respectively. One of the most important factors in enhancing the effectiveness of gas-sensing devices is quick recovery. After the gas adsorption, the adsorbent is typically recovered by chemical exposure and irradiation with high-intensity electromagnetic radiation. It is possible to recover the weak or intermediate-level

Table 2 Bond length (Å) in optimized gas adsorbed BNNS and As-BNNS, and deformation energy (E_{def}) per atom

Elements	Bond length					E_{def} (eV per atom)
	B–N	As–B	C=O	S–H	S=O	
CO_2 + BNNS	1.475	—	1.181	—	—	2.6×10^{-6}
H_2S + BNNS	1.477	—	—	1.352	—	9.75×10^{-6}
SO_3 + BNNS	1.508	—	—	—	1.442	3.08×10^{-3}
CO_2 + As-BNNS	1.476	2.049, 1.990	1.181	—	—	0.053
H_2S + As-BNNS	1.476	2.042, 1.984	—	1.351	—	0.053
SO_3 + As-BNNS	1.477	2.031, 1.971	—	—	1.446	0.051



Table 4 Recovery time τ (s) for different adsorbates in the BNNS and As-BNNS

Adsorbed gas	BNNS		As-BNNS	
	$f_o = 10^{12}$	$f_o = 3 \times 10^{14}$	$f_o = 10^{12}$	$f_o = 3 \times 10^{14}$
CO ₂	3.94×10^{-11}	1.31×10^{-25}	3.03×10^{34}	1.01×10^{20}
H ₂ S	9.03×10^{-10}	3.01×10^{-24}	3.92×10^{32}	1.30×10^{18}
SO ₃	6.40×10^{-5}	2.13×10^{-19}	5.02×10^{39}	1.67×10^{25}

interaction between gas molecules using UV light. Chemical reactions involving reagents including HCl, NaOH, HNO₃, and NaHCO₃ work better for heavy adsorption in recovering the adsorbent for use in subsequent processes.^{66–69} Table 4 shows that while the As-BNNS layer has a longer recovery time than the BNNS layer, the greater adsorption results in a longer recovery time. Sensing CO₂, H₂S, and SO₃ gases will be simpler with BNNS as a sensor than As-BNNS. However, the chemical reaction can be employed in the event of hard adsorption and extended recovery time for As-BNNS.

3.3 Electronic properties

The produced electronic band structure, which is based on the accumulated adsorbate molecules in the adsorbed system, describes the clear variations in the energy gap between the valence band and conduction band. To gain a comprehensive understanding of their electrical response, we carried out electronic band structure (EBS) analysis for the pristine and As-doped BNNS structures as well as the gas-adsorbed pristine and As-doped BNNS structures which shed light on their optical properties. To examine the exchange–correlation potential for electronic structure analysis using CASTEP, we concentrated on GGA-PBE schemes. A range of physicochemical factors, such as the kind of conductor, the electrochemical potential of various species, and the activity of surface reactions, may be explained by the EBS concerning the Fermi level.⁷⁰ Fig. 2 provides a detailed depiction of the adsorbent and adsorption system's EBS series of energetic lines along high symmetry points as $G-X-M-Y-G$. Due to their symmetry with spin polarization, all of the band structures are only depicted for up-spin circumstances.

Pure BNNS and gas-adsorbed BNNS structures have an indirect band gap, with the valence band maximum (VBM) at point M in K -space and the conduction band minimum (CBM) between the G and X points. The calculated bandgaps along with Fermi energy for the investigated structures are tabulated in Table 5. The prohibited passage of electrons from the valence band to the conduction band is described by the enormous bandgap of 4.098 eV seen in pristine BNNS, which is comparatively smaller than that of the h-BN monolayer,⁷¹ making the tetragonal phase more suitable for numerous optoelectronic studies. The CO₂, H₂S, and SO₃ adsorbed pristine BNNS had electronic bandgaps of 4.088 eV, 4.035 eV, and 3.114 eV, respectively. We have reported the minimum and maximum variation of bandgaps between pristine and gas-adsorbed BNNS

is found to be 0.2% and 24% respectively. This might indicate that gas molecules have been adsorbed, causing an insulator-to-semiconductor transition. The dispersion of the energy band and the change in bandgap energy during gas adsorption demonstrate the high interaction of pure BNNS to gas molecules.

In the case of As-BNNS, we observed an indirect semiconducting bandgap of 2.708 eV with the VBM between G and X points and CBM at M points. However, all the gas-adsorbed As-BNNS exhibit a direct band gap nature with the CBM and VBM at M point. The bandgaps for CO₂, H₂S, and SO₃ adsorbed As-BNNS are found to be 3.576 eV, 3.567 eV, and 3.066 eV respectively. In this case, the minimum and maximum variations are 13% and 32% respectively. The semiconducting bandgap of As-BNNS increases with the adsorption of gas molecules. This indicates the opposite interaction of gas molecules with As-BNNS compared to the pure BNNS layer. This may result from the increase of bandgap in the presence of gas molecules.

Overall, the change in the EBS suggested that the pristine BNNS and As-BNNS have strong interaction with the CO₂, H₂S, and SO₃ gas molecules. However, this interaction is stronger in the case of As-BNNS resulting from the large variation in bandgaps than BNNS. In addition to EBS, a material's DOS and PDOS are employed to determine its electronic nature. The final clue as to whether a material is an insulator, conductor, or semiconductor is the presence of electron states close to the Fermi level E_F . The contribution and interactions of the orbital electrons determine the type and creation of EBS. With the significance of DOS and PDOS in mind, we computed the total DOS and PDOS for gas-adsorbed BNNS and As-BNNS structures as well as BNNS and As-BNNS structures throughout the energy range of -10 eV to 10 eV, as shown in Fig. 3.

We observed that DOS is absent in BNNS throughout a wide energy range indicating BNNS's broad bandgap compared to others. The bandgap was decreased and the valence and conduction bands moved in the direction of the Fermi level as a result of the gas adsorption in BNNS. We have found that in BNNS, B-2p and N-2p orbital electrons predominate in the valence band and conduction band, following the PDOS. Along with B-2p and N-2p, the O-2p orbital electrons in CO₂-BNNS contribute significantly to the valence and conduction bands, compared to C-2p. The addition of S-3p orbital electrons mostly drove the energy bands toward the Fermi level in the H₂S-BNNS instance, reducing the bandgap simultaneously. More than half-occupied O-2p and S-3p orbital electrons have been shown to contribute significantly to the greatest shift of the valence band and conduction band observed in SO₃-BNNS.

Doping allowed us to see the semiconducting bandgap and electron states that were very close to the Fermi level. However, energy bands in gas-adsorbed As-BNNS moved out from the Fermi level, creating a larger bandgap. The As-4p orbital electrons predominate in the VBM and CBM, but all of the gas molecules in As-BNNS and their valence orbitals follow almost identical laws to those in BNNS. BNNS and As-BNNS, however, exhibit high interaction for the studied gas molecules,



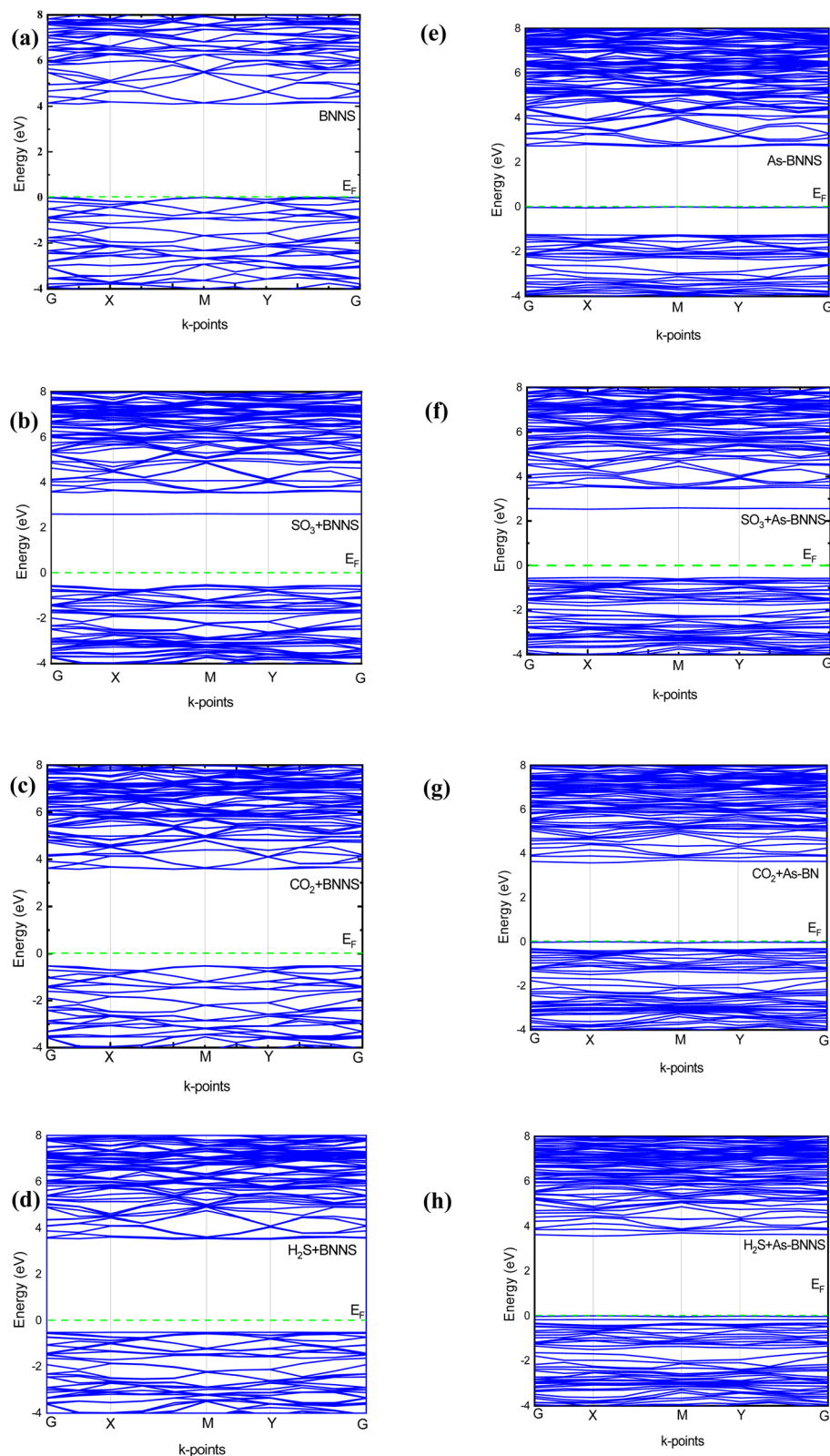


Fig. 2 Electronic band structures of BNNS, As-BNNS, gas adsorbed BNNS and As-BNNS structures.

according to the electronic response revealed by the study of DOS and PDOS. This interaction is increased by doping the As atom in pure BNNS.

The variation in the bandgap energy of any material influences the electrical conductivity. In semiconducting materials, the electrical conductivity reduces with increasing electronic



Table 5 Calculated bandgap (E_g), Fermi energy (E_F), work function (ϕ), electrical conductivity (σ) and sensitivity of the adsorbents and their gas-adsorbed complexes

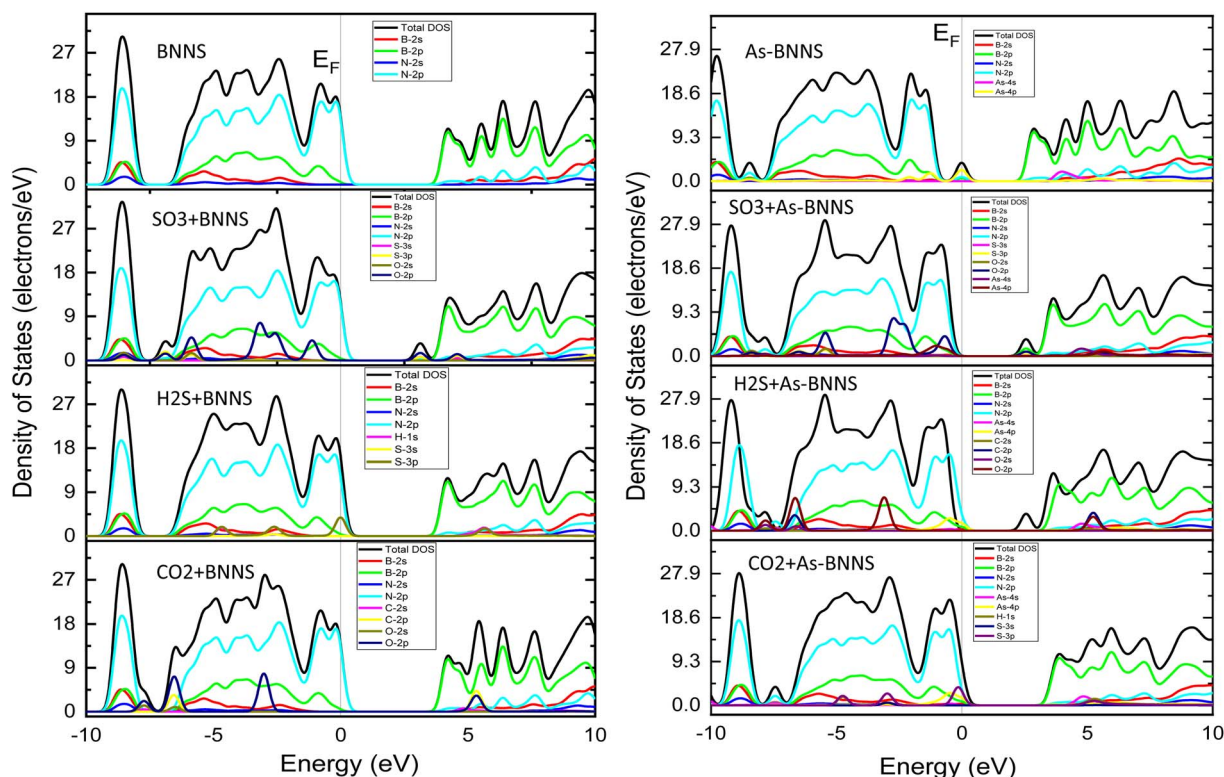
Elements	E_g (eV)	E_F (eV)	ϕ (eV)	$\Delta\phi$ (eV)	σ ($A \times \text{ohm}^{-1} \text{m}^{-1}$)		
					300 K	400 K	500 K
BNNS	4.098	-1.131	1.1313	—	0.924	0.942	0.954
CO ₂ + BNNS	4.088	-4.129	4.1288	2.997	0.924	0.942	0.954
H ₂ S + BNNS	4.035	-4.230	4.2302	3.099	0.925	0.943	0.954
SO ₃ + BNNS	3.114	-2.163	2.1628	1.032	0.942	0.956	0.965
As-BNNS	2.708	-1.375	1.3752	—	0.949	0.961	0.969
CO ₂ + As-BNNS	3.576	-1.350	1.3501	0.025	0.933	0.949	0.959
H ₂ S + As-BNNS	3.567	-1.496	1.4962	0.121	0.933	0.950	0.959
SO ₃ + As-BNNS	3.066	-2.523	2.5227	1.148	0.942	0.956	0.965

bandgap. Semiconducting materials have a positive temperature coefficient of conductivity. Such variation of electrical conductivity with bandgap and temperature can be observed in the following equation:⁷²

$$\sigma = Ae^{-\frac{E_g}{2K_B T}} \quad (5)$$

where A is a constant term which has not been determined yet since the structures are not synthesized experimentally. Hence only the exponential term was calculated to show the variation of electrical conductivity with respect to E_g . Therefore, the unit of σ has been written as $A \times \text{ohm}^{-1} \text{m}^{-1}$. Table 5 lists all of the computed electrical conductivity values for the compositions under investigation. We have observed that as bandgap energy decreases, electrical conductivity rises exponentially. When gas

adsorption occurs on BNNS, the minimum bandgap and maximum conductivity are displayed by SO₃ + BNNS. The adsorbent layer As-BNNS exhibits the lowest bandgap and highest conductivity during gas adsorption, and conductivity also decreases after gas adsorption. Overall, BNNS's electrical conductivity rises with gas adsorption, but As-BNNS exhibits the opposite behavior. Table 5 shows the opposite trend to the bandgap fluctuation based on the temperature change of electrical conductivity. Electrical conductivity rises as we go from 300 K to 500 K temperature. The As-BNNS structure's highest conductivity is discovered at 500 K. By displaying changes in electrical conductivity and resistivity when gas is present near the adsorbent, the gas sensor can identify the presence of gas molecules.

**Fig. 3** The density of states (DOS) and projected density of states (PDOS) of BNNS, As-BNNS, gas-adsorbed BNNS and As-BNNS structures.

Any adsorbent's fluctuation in electrical conductivity is determined by the work function Φ , which gives a probability that charge carriers will be transported from the adsorbent surface. As a result, we noticed that the work function of adsorbent materials varied depending on whether gas molecules were present or not as follows:⁷³

$$\Phi = |V_{\infty} - E_F| \quad (6)$$

where V_{∞} refers to the potential at an infinite distance from the adsorbent surface which is considered as zero. Table 5 presents the work function-derived values in an ordered manner.

In comparison to the As-BNNS structure, the BNNS structure has a maximum charge escape probability and a minimum Φ . Following gas adsorption, As-BNNS is responsible for CO₂ adsorption while BNNS has a minimal work function during SO₃ adsorption. The gas adsorption resulted in a significant change in Φ which can be utilized to build work function-based sensors, where Φ of the complexes can be measured experimentally *via* the Kelvin method based on these variations in work function.⁶⁰ The sensor can be calibrated with reference values of Φ , which can not only sense the presence of selected gases but also identify the type of toxic gas. The mathematical relation between the variation of conductivity and Φ is the Richardson–Dushman relation⁷⁴ as follows:

$$J_R = B_R T^2 e^{-\frac{\Phi}{k_B T}} \quad (7)$$

where B_R and J_R stand for the Richardson constant and the density of released charge carriers, respectively. According to the aforementioned equation, the number of escaping charge carriers rises exponentially as Φ decreases. Based on Table 5's observation of Φ fluctuation, the BNNS layers have a very high surface conductivity.

The sensitivity of any adsorbent for certain gas molecules can be observed by detecting the variation of electrical conductivity in the presence and absence of gas molecules, which can be formulated as:⁷⁵

$$\text{Sensitivity} = \frac{|\sigma_{\text{adsorbent}} - \sigma_{\text{adsorbent+gas}}|}{\sigma_{\text{adsorbent}}} \quad (8)$$

where $\sigma_{\text{adsorbent}}$ and $\sigma_{\text{adsorbent+gas}}$ are the conductivity of the adsorbent layers and adsorbent with gas respectively. The variation of sensitivity with the variation of gas molecules and temperature is illustrated in Fig. 4. The As-BNNS has the highest sensitivity to SO₃ at 300 K out of all the gas adsorption structures. Furthermore, when the temperature rises from 300 K to 500 K, the sensitivity falls. Consequently, the As-BNNS may be effectively employed as a sensor medium for the adsorption of SO₃ gas under ambient conditions.

We conducted Mulliken bond population analysis (MPA),^{76,77} a widely used method for learning more about a molecule's bonding activity. The calculated Mulliken charges in BNNS and gas-adsorbed BNNS are given in Table 6. As seen in Fig. 5, electron density difference (EDD) mapping provides a clear image of the fluctuation in electron density caused by gas adsorption. The yellow zone denotes an electron shortage, while the blue region indicates an electron buildup. The partially

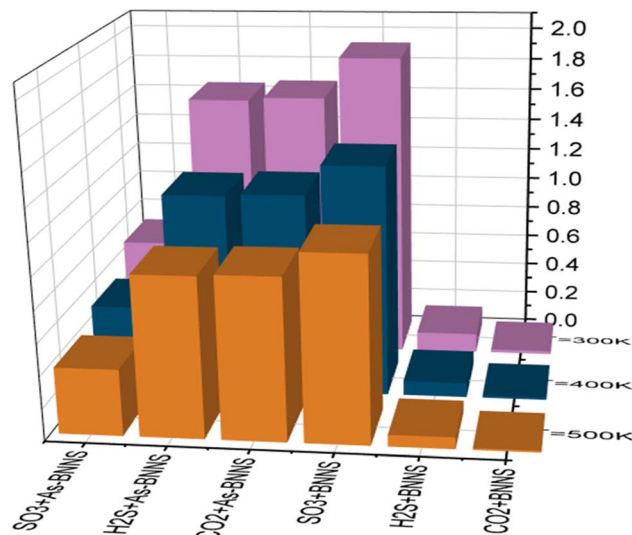


Fig. 4 Sensitivity of gas adsorbed BNNS and As-BNNS structures at different temperatures.

ionic nature of the B–N bonds in the pristine BNNS is indicated by charge analysis using MPA, which shows that an average of 0.87|e| charge is transferred from the B atoms to their neighboring N atoms within the sheet due to the N atom's higher electronegativity.

Following Mulliken charges in Table 7, there is only a little shift in the charges of B and N. H and C exhibit their characteristic electropositive nature in the gas adsorption structures, whereas O exhibits an electronegative nature. On the other hand, S has an electropositive character when O is present and an electronegative character when H is present.

This may result from the S-3p⁴ orbital's behavior to receive or leave electronic charge to make a stable configuration. The comparison shows that SO₃ > H₂S > CO₂ is the overall charge transfer from gas molecules to the adsorbent based on the total Mulliken charge of both the gas and the adsorbent. The EDD mapping result from the dispersion of charge density towards gas molecules follows this outcome. Though the quantity of charge transfer in As-BNNS is more than that of BNNS, charge transfer between the gas molecules and adsorbent still occurs in the same sequence.

Given the greater dispersion of the isosurfaces toward gas molecules, the EDD mapping is consistent with these

Table 6 Average Mulliken charge distribution of pristine BNNS and gas-adsorbed BNNS

Elements	BNNS	BNNS + CO ₂	BNNS + H ₂ S	BNNS + SO ₃
B	0.87	0.869	0.867	0.873
N	−0.87	−0.87	−0.868	−0.868
H	—	—	0.17	—
C	—	0.97	—	—
O	—	−0.49	—	−0.80
P	—	—	—	—
S	—	—	−0.36	2.32



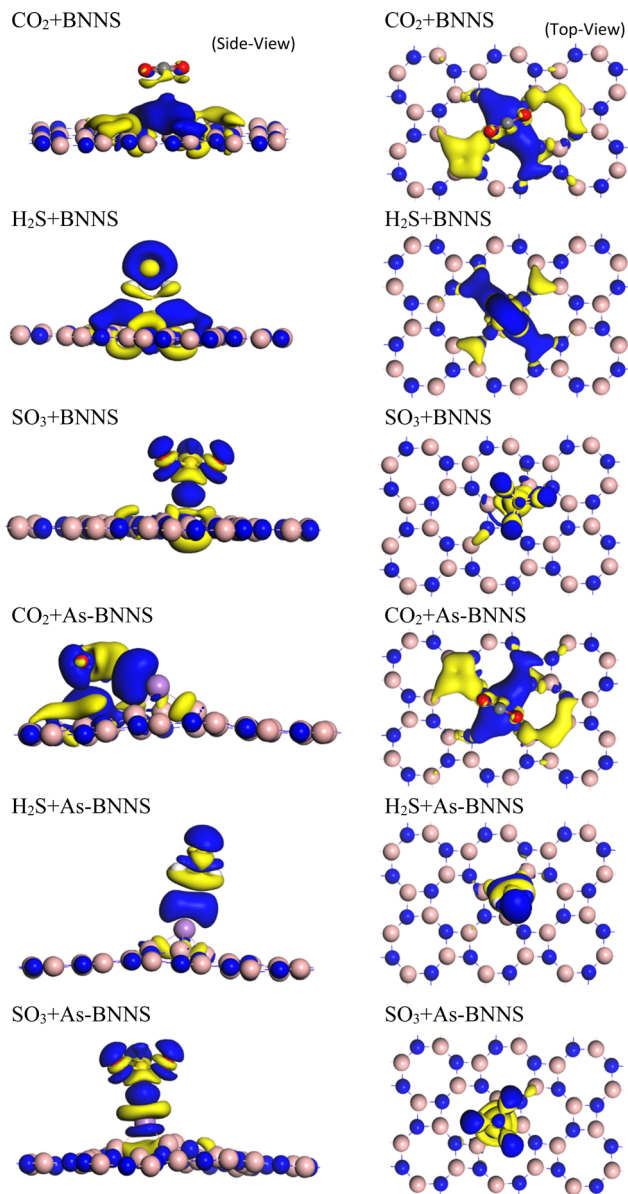


Fig. 5 EDD mapping of gas adsorbed BNNS and As-BNNS structures at different temperatures.

Table 7 Mulliken charge distribution of As-BNNS and gas adsorbed BNNS

Elements	As-BNNS	As-BNNS + CO ₂	As-BNNS + H ₂ S	As-BNNS + SO ₃
B	0.808	0.820	0.820	0.828
N	-0.854	-0.861	-0.861	-0.861
As	0.27	0.12	0.16	0.19
H	—	—	0.16	—
C	—	0.94	—	—
O	—	-0.48	—	-0.83
P	—	—	—	—
S	—	—	-0.35	2.24

observations. It is noted that the As-BNNS and gas-adsorbed As-BNNS structures exhibit comparatively greater changes in the partial charges of B and N. As, an electropositive element, exhibits partial positive charges in the structures of As-BNNS and gas-adsorbed As-BNNS. Following gas adsorption, the As atom primarily loses its charge to the BN sheet and the electronegative components of gas molecules. This being said, there is consistency between the EDD analysis and Mulliken charge. The main determinant of material interaction is charge transportation. This idea can help explain how gas molecules interact with the adsorbent layers. Fig. 6 illustrates the charge transfer between the adsorbent and gases. It is evident that during gas adsorption, the adsorbent layer leaves charge carriers behind, which act as donors and gas molecules as receivers. A maximum charge transfer of $-0.08e$ is seen in the SO₃ + BNNS structure. The As-BNNS structure has demonstrated similar outcomes in the charge transfer pattern. It can be shown from Fig. 6 that the doping of the As atom enhances the charge transfer from the adsorbent to gas. Overall, the interaction between the adsorbent layer and the gas molecules was improved by As atom doping.

3.4 Optical properties

3.4.1. Dielectric constants. In gas sensing research, the study of the pure layer's and gas-adsorbed layer's optical properties aids in determining a layer's sensitive features. DFT is a valuable theoretical technique applied to many-body systems using the principles of quantum mechanics. All the frequency-dependent optical parameters of pure and gas adsorbed layers are depicted in Fig. 7–9. The optical characteristics mentioned above are mainly ascribed to the computation of the complex dielectric function $\epsilon(\omega) = \epsilon_1(\omega) + i\epsilon_2(\omega)$, where the imaginary part $\epsilon_2(\omega)$ represents the energy lost through electronic absorption, and the real part $\epsilon_1(\omega)$ is related to the dielectric constant and describes the electronic-polarizability of the material.⁷⁶ The polarized optical constants along the [100] direction of the plane have been estimated with Gaussian smearing of 0.5 eV, and Drude damping of 0.05 eV. The

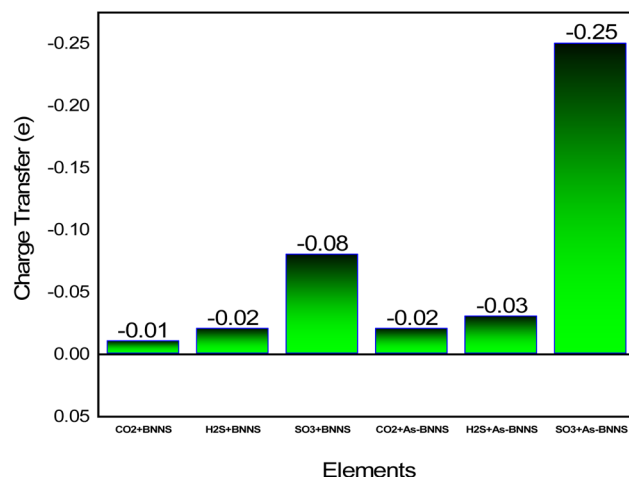


Fig. 6 Charge transfer between the host layers and gas molecules.



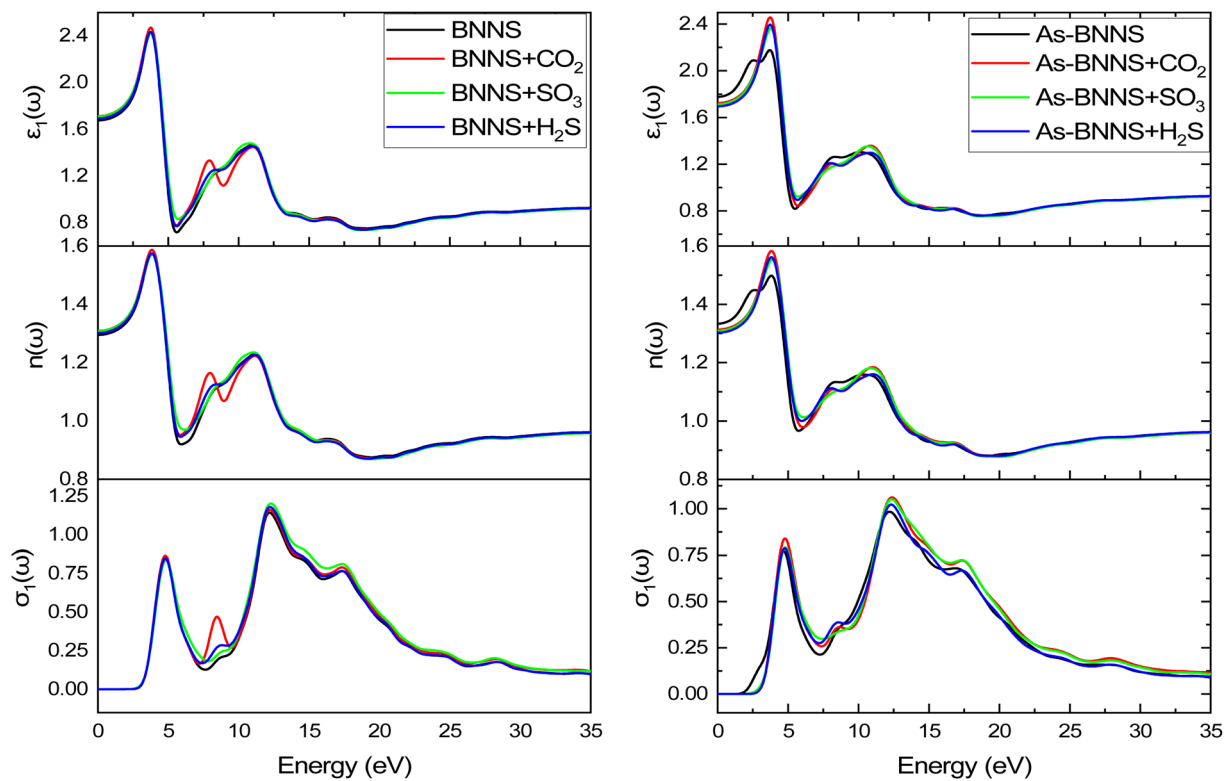


Fig. 7 The real part of dielectric constant $\epsilon_1(\omega)$, refractive index $n(\omega)$, and conductivity $\sigma_1(\omega)$ from top to bottom.

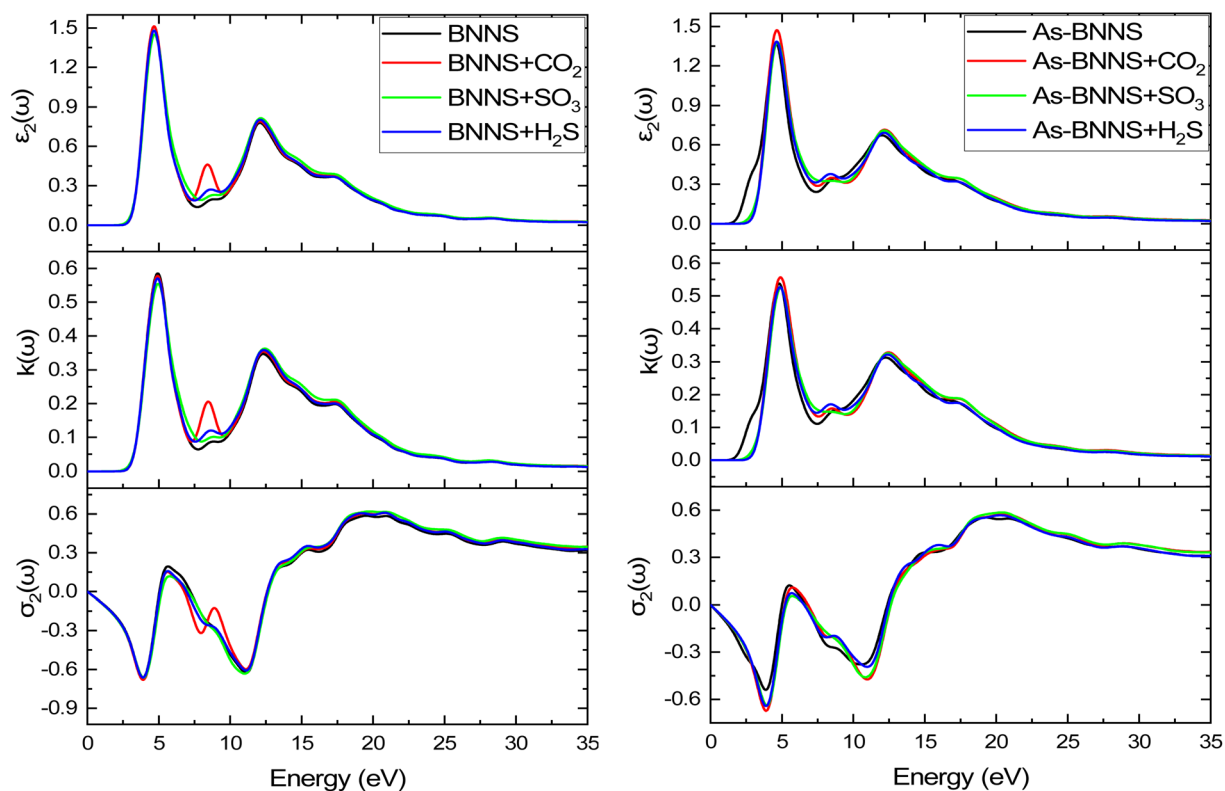


Fig. 8 The imaginary part of dielectric constant $\epsilon_2(\omega)$, refractive index $k(\omega)$, and conductivity $\sigma_2(\omega)$ from top to bottom.



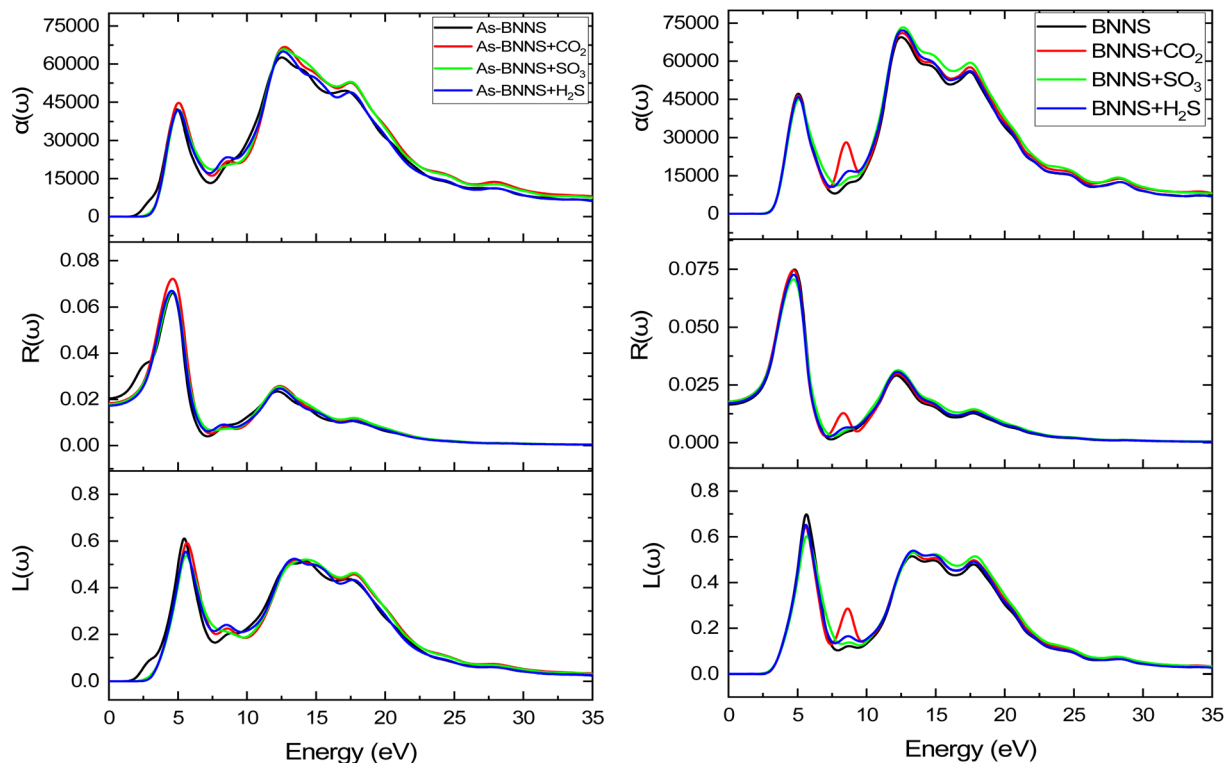


Fig. 9 The calculated absorption coefficient $\alpha(\omega)$, reflectivity $R(\omega)$, and loss function $L(\omega)$.

observed real and imaginary components of the optical dielectric constant are displayed in the uppermost layers of Fig. 7 and 8, respectively.

For BNNS, CO₂ + BNNS, H₂S + BNNS, and SO₃ + BNNS, the real component of the dielectric tensors at zero photon energy $\epsilon_1(0)$ was determined to be 1.676, 1.712, 1.693, and 1.714, respectively. Regarding As-BNNS, $\epsilon_1(0)$ has been determined to be 1.778, 1.725, 1.697, and 1.714 for As-BNNS, CO₂ + As-BNNS, H₂S + As-BNNS, and SO₃ + As-BNNS. It shows that during gas adsorption, $\epsilon_1(0)$ of BNNS and As-BNNS revealed moderate alterations. As-BNNS and BNNS nearly always retain a similar route with distinct values as a function of energy in their depicted lines of $\epsilon_1(\omega)$. For both BNNS and As-BNNS, the highest value of $\epsilon_1(\omega)$ was determined to be 2.43 and 2.18 respectively. Following gas adsorption, we discovered that the $\epsilon_1(\omega)$ values for SO₃ + BNNS, H₂S + BNNS, and CO₂ + BNNS are, respectively, 2.47, 2.43, and 2.42. We also determined $\epsilon_1(\omega)$ to be 2.46, 2.40, and 2.36 for CO₂ + As-BNNS, H₂S + As-BNNS, and SO₃ + As-BNNS, respectively. The sensitive nature of gas molecules is implied by this shift in the maximum values of $\epsilon_1(\omega)$ caused by gas adsorption. The doping of As in the pure BNNS layer improves the gas sensitivity resulting from the greater chance in $\epsilon_1(\omega)$ in As-BNNS. The dielectric tensor $\epsilon_2(\omega)$ in BNNS shows peaks with maxima at around 3.78 eV and stays at zero from 0.0 eV to 1.91 eV when it starts to rapidly rise. This might be explained by charge carriers' intraband transition. Up to 7.5 eV, there is very little movement in the $\epsilon_2(\omega)$ for the gas-adsorbed BNNS. The gas adsorption in As-BNNS has been reported to cause a significant change of $\epsilon_2(\omega)$ of about 1.9 eV. For BNNS,

CO₂ + BNNS, H₂S + BNNS, and SO₃ + BNNS, the highest $\epsilon_2(\omega)$ values are 1.50, 1.51, 1.48, and 1.45 respectively. As-BNNS, CO₂ + As-BNNS, SO₃ + As-BNNS, and H₂S + As-BNNS are calculated to be 1.37, 1.47, 1.38, and 1.38 in As-BNNS, respectively. These results are a prime indicator of a shift in $\epsilon_2(\omega)$ due to gas adsorption by BNNS and As-BNNS.

3.4.2. Refractive index. The spectra of the real component of the refractive index, $n(\omega)$, in both the gas adsorbed and host layers closely resemble the $\epsilon_1(\omega)$ tensor seen in Fig. 7's middle layer for both BNNS and As-BNNS. The highest values of $n(\omega)$ exhibit variation throughout the energy range of 1.9 eV to 4.33 eV, with a decline beginning at 12.0 eV after two-fold peaks. The square root value of the dielectric function $\epsilon_1(\omega)$ is comparable to the term "refractive index $n(\omega)$ ", as we saw, and as a result, the graphs for these two parameters resembled one another. As shown in the top and middle layers of Fig. 8, the imaginary component of the refractive index, $k(\omega)$, and incident photon attenuation consisted of $\epsilon_2(\omega)$. The number of occupied states and photon attenuation are closely correlated. As a result of this process, the number of occupied states increases, increasing the dielectric constant and the attenuation of light photons. A means of identifying the deposited gas might be the fluctuation in refractive index, which indicates that the speed of the photon in the medium varies after gas adsorption.

3.4.3. Absorption coefficient and optical conductivity. To ascertain a 2D layer's sensitivity to gas molecules, the absorption coefficient and optical conductivity analysis may be the most crucial and comprehensible components. The electrical conductivity of a material resulting from electromagnetic



radiation is known as optical conductivity, and the percentage of light intensity attenuation per unit of distance that light travels through a medium is known as the absorption coefficient.

The optical conductivity is represented by the real component σ_1 , while the absorption loss is represented by the imaginary part σ_2 , which is related to $\varepsilon_2(\omega)$. We may achieve a very similar pattern of optical conductivity $\sigma_1(\omega)$ and absorption coefficient $\alpha(\omega)$ as seen in Fig. 7 and 9. The first peaks of $\sigma_1(\omega)$ and $\sigma_2(\omega)$ were made up of the findings from the EBS investigation, and the nonzero values of $\alpha(\omega)$ and $\sigma_1(\omega)$ for a certain energy range indicate the presence of a bandgap as shown in EBS analysis. The As doping and gas adsorption cause a change in the beginning point of $\sigma_1(\omega)$ and $\alpha(\omega)$. These changes are more pronounced in As-BNNS than in pure BNNS and are attributed to gas adsorption. They are comparable to how bandgap analysis has changed. Remarkably, neither conduction nor absorption took place at lower energies as they are not strong enough to excite electrons; instead, they help molecules vibrate. Both $\sigma_1(\omega)$ and $\alpha(\omega)$ show a maximal number of electron transitions at the higher energy range, around 11.4 eV to 17.5 eV, as a result of the wider and bigger spectra. For BNNS/As-BNNS and gas-adsorbed BNNS/As-BNNS, the highest values of $\sigma_1(\omega)$ and $\alpha(\omega)$ are obtained at 12.1 eV and 12.5 eV, respectively. It is evident from the spectra of $\sigma_1(\omega)$ and $\alpha(\omega)$ that the gas adsorption causes a rise in their maximum values. A slight peak shift in the absorption spectra is observed due to gas adsorption which can be utilized to identify the type of associated toxic gas.⁷⁸

3.4.4. Reflectivity. The adsorbed gas type may be identified by examining the variance in the reflected energy spectra.⁷⁸ The middle layer of Fig. 9 shows the reflectivity spectra of virgin and gas-adsorbed structures. As is well known, some photons are absorbed, some are transmitted, and others are reflected when they collide with a substance. Since absorption is weaker in the low-energy zone, all structures had greater reflectivity values. As-BNNS has a maximum loss of 6.6% of photon energy at 4.6 eV, whereas BNNS shows a maximum reflectivity of 7.48% of incident photon energy at 4.77 eV. Primarily the image illustrates that the maximum reflectance shifts as a result of gas adsorption. Because of the strong absorption in the UV zone, reflectance tends to zero there. Low-energy reflectivity spectra show a redshift following gas adsorption in BNNS, whereas high-energy region reflectivity spectra show a blueshift. But gas adsorption in the case of As-BNNS results in a blue shift in the reflectivity spectrum at all energies. The shifting in reflection peaks may result in a slight change in the color of the adsorbent due to gas adsorption which can be calibrated to identify the type of toxic gas present in a chamber.⁶²

3.4.5. Loss function. One of the most significant peaks in the energy loss function spectrum is the plasmon peak. Connected to this peak are the reflectivity peaks and the imaginary component of the dielectric function, $\varepsilon_2(\omega)$. The highest peaks are referred to as plasmon peaks and the matching frequency is referred to as the plasma frequency because of their collective oscillations. The material becomes transparent to light if the incident light's frequency is higher than the plasma frequency.

We noticed a slight blue shift in the plasmon peaks from the bottom layers of Fig. 9 as a result of gas adsorption. Following the dielectric function $\varepsilon_2(\omega)$ the loss peaks were found to be high within the 11.4 eV to 17.5 eV range.

4. Conclusions

A tetragonal BN layer has been designed and doped with an As atom to explore their adsorption properties using DFT calculations. Toxic gas molecules such as CO₂, H₂S, and SO₃ are adsorbed successfully with a significant increase in adsorption energy *via* As-doping. Through geometry optimization and the determination that the cohesive energy is negative, the energetically stable configuration of the doped and undoped BNNS layers has been achieved. Greater variations in the bond lengths of the adsorbent and adsorbate suggest a comparatively strong interaction between As-BNNS and selected gases. Compared to CO₂ and SO₃, the adsorbent layers interact more with SO₃, and As-doping enhances this BNNS interaction significantly. As evidenced by the conductivity research, DOS, and EBS results, pristine BNNS is naturally insulating. Gas adsorption reduces the band gap and promotes conductivity. During SO₃ adsorption, an almost insulator-to-semiconductor transition takes place. In contrast, gas adsorption broadens the bandgap and As-doping turns BNNS into a semiconductor. The adsorbent materials' sensitivity to adsorbates was reinforced by the change in conductivity brought about by the adsorption of gas molecules. The BNNS and As-BNNS exhibit low refractive index and strong UV absorption coefficients, indicating their promise as optoelectronic research materials. These adsorbent layers can function as gas sensors, as demonstrated by the shift in peak positions of reflectivity and absorption throughout the gas adsorption process. A notable alteration in the reflection spectra implies that color changes may be used to identify the adsorbed gases. A thorough examination of the geometrical and optoelectronic characteristics reveals that As-BNNS outperforms BNNS as a gas sensor.

Data availability

The research data is not shared since it is currently being used in an ongoing study.

Conflicts of interest

The authors state that none of the work presented in this study may have been influenced by any known conflicting financial interests or personal ties.

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