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Investigation of electronic parameters, carrier transport mechanisms *via* the correlated barrier hopping model, electrothermal NTCR effects, and polarization contributions to the dielectric response of $\text{Ni}_3(\text{PO}_4)_2$ orthophosphates synthesized by the sintering process

Iheb Garoui,^a Sourour Ben Yahia,^{ID *a} Noweir Ahmad Alghamdi,^b Iskandar Chaabane,^a Abderrazek Oueslati^{ID a} and Bassem Louati^a

Nickel orthophosphate, $\text{Ni}_3(\text{PO}_4)_2$, was prepared through a conventional solid-state synthesis route and its crystal structure was verified by X-ray diffraction, confirming a monoclinic phase with space group $P2_1/c$. Optical studies using UV-Vis spectroscopy revealed strong absorption in the ultraviolet region, consistent with its semiconducting nature. The dielectric and electrical behaviors were systematically examined as functions of both temperature and frequency using impedance and electric modulus approaches, which allowed a clear distinction between grain and grain-boundary contributions, relaxation processes, and charge transport mechanisms. The temperature-dependent variation of the power-law exponent $s(T)$ showed a decreasing trend, pointing to the correlated barrier hopping (CBH) mechanism as the dominant conduction process within the investigated temperature range (393–633 K). Based on this model, the maximum barrier energy was estimated as $W_M = 0.66$ eV. Analysis of the impedance and modulus spectra confirmed a non-Debye type relaxation with well-resolved grain and grain-boundary contributions. The impedance data were successfully fitted using an equivalent $R//C//CPE$ circuit, and the variation in resistance further confirmed the NTCR behavior. The material displays a notably high dielectric constant at low frequencies ($\epsilon' \approx 1.58 \times 10^4$ at 633 K), reflecting effective polarization processes and thermally activated charge carrier motion. At higher frequencies, it preserves stable dielectric characteristics accompanied by moderate energy losses. Overall, these results establish a strong correlation between relaxation phenomena and hopping-based charge transport, highlighting $\text{Ni}_3(\text{PO}_4)_2$ as a potential candidate for high-frequency dielectric and microelectronic device applications.

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1. Introduction

Phosphate-based compounds have become a focal point of research due to their exceptional structural robustness, chemical adaptability, and versatile functional properties, making them suitable for a wide range of technological applications. These include solid-state electrolytes, high-performance dielectric components, energy storage devices, catalytic systems, and advanced communication technologies.^{1–5} Their crystal frameworks are built from PO_4^{3-} tetrahedra, where

strong covalent P–O bonds impart excellent thermal and chemical stability. This rigidity, combined with tuneable ionic and electronic transport properties, allows phosphate materials to be tailored for diverse functionalities, ranging from ceramic capacitors⁶ and microwave resonators^{7,8} to photovoltaic systems^{9,10} and emerging high-frequency devices for 5G/6G communications.¹¹ Within this family, transition-metal phosphates stand out due to the synergy between the transition-metal cation's electronic configuration and the phosphate network. The $\text{A}_3(\text{PO}_4)_2$ family, where A represents a divalent transition metal, has demonstrated remarkable potential, with examples including $\text{Mg}_3(\text{PO}_4)_2$,¹² $\text{Co}_3(\text{PO}_4)_2$,¹³ and $\text{Cu}_3(\text{PO}_4)_2$.⁷ It is well recognized that the integration of distinct compounds possessing superior electronic characteristics can give rise to advanced composite materials with enhanced functional performance. Such composites have garnered considerable

^a Laboratory of spectroscopic characterization and optical materials, Faculty of Sciences, University of Sfax, B.P. 1171, 3000 Sfax, Tunisia.

E-mail: sourourbenyahia09@gmail.com

^b Departement of Physics, Faculty of Science, Al-Baha University, Alaqiq, 65779, Saudi Arabia



technological interest in recent years due to their tenable physical properties and potential applications in electronics, energy storage, and catalysis.¹⁴ The introduction of a secondary phase can markedly improve key electronic parameters such as electrical conductivity, dielectric stability, and charge transport efficiency by promoting interfacial interactions and synergistic effects between the constituent phases.

In recent years, combining inorganic compounds with carbon-based materials such as graphene, carbon nanotubes, and carbon nanospheres has proven to be an effective approach for improving the electronic, dielectric, and catalytic properties of functional composites.¹⁵ The intrinsic high conductivity and large surface area of these carbon matrices enhance charge transport and strengthen interfacial polarization, thereby boosting the overall electronic performance of the hybrid systems. For example, the $\text{Na}_{1.25}\text{Cr}_{0.25}\text{Ti}_{1.75}(\text{PO}_4)_3/\text{C}$ anode material for sodium-ion batteries exhibited markedly enhanced electrochemical behavior and improved charge transport mechanisms.¹⁶ These findings highlight the considerable potential of carbon-based hybrid composites for the development of materials with optimized electronic and functional properties for advanced technological applications. The electronic properties of oxide- and transition-metal-based compounds are highly sensitive to stoichiometry, particularly the oxygen content.¹⁷ Deviations from ideal stoichiometry, including oxygen excess or deficiency, can alter the oxidation state of cations, thereby significantly affecting charge transport, dielectric response, and overall electronic behavior.

Nickel-based compounds have been widely studied due to their outstanding structural flexibility and diverse functional properties.¹⁸ For example, Ni_2FeO_4 ¹⁹ and NiCo_2O_4 ²⁰ exhibit remarkable electrochemical performance and are extensively utilized in supercapacitors, gas sensors, and catalytic systems, attributed to their mixed-valence nature and high electrical conductivity. Nevertheless, metallic materials, hydroxides, and their alloys are often constrained by their limited resistance to aggressive environments, including elevated temperatures, oxidative atmospheres, and electromagnetic radiation. In contrast, oxide-based compounds display superior thermal and chemical stability, preserving their structural integrity and performance even at temperatures approaching 1000 °C.²¹ Moreover, metals, hydroxides, and their alloys typically suffer from substantial energy losses induced by eddy currents under alternating electromagnetic fields, which further restricts their use in high-frequency applications.²²

Similarly, nickel phosphates such as $\text{Ni}_3(\text{PO}_4)_2$,²³ $(\text{NaNi}_4(\text{PO}_4)_3/\text{GF})$,²⁴ $\text{K}_2\text{NiP}_2\text{O}_7$,²⁵ and LiNiPO_4 ²⁶ have attracted interest as potential candidates for dielectric, magnetic, and energy-storage applications. These compounds combine the stability of the phosphate framework with the redox activity of Ni^{2+} ions, favouring efficient charge transport and polarization phenomena. Furthermore, the incorporation of nickel ions into phosphate lattices can modulate the electronic structure and enhance relaxation processes through hopping or localized polaron mechanisms.^{23,27,28} Consequently, the study of $\text{Ni}_3(\text{PO}_4)_2$ contributes not only to understanding the intrinsic dielectric and

conduction behavior of nickel phosphates but also to exploring their potential for high-frequency and energy-related applications. Nickel orthophosphate ($\text{Ni}_3(\text{PO}_4)_2$) crystallizes in the monoclinic space group $P2_1/c$ with lattice parameters $a = 5.8283(2)$ Å, $b = 4.6931(1)$ Å, $c = 10.0998(4)$ Å, and $\beta = 91.1(2)^\circ$. In its crystal structure, Ni^{2+} ions are coordinated by oxygen atoms from phosphate groups, forming an extended three-dimensional framework.²⁹ This arrangement offers potential pathways for polaron hopping and localized charge migration, making it an attractive candidate for both conduction and dielectric relaxation studies. Previous investigations on $\text{Ni}_3(\text{PO}_4)_2$ have largely focused on its crystal chemistry, thermodynamic properties, and catalytic activity. However, comprehensive studies examining its dielectric performance, impedance behavior, and conduction mechanisms remain limited. Given the increasing demand for lead-free, thermally stable, and environmentally benign dielectric materials, a deeper understanding of the electrical transport and polarization mechanisms in $\text{Ni}_3(\text{PO}_4)_2$ is both timely and essential. In particular, the influence of grain boundaries, defect states, and thermal activation on its AC and DC transport properties requires systematic exploration.

In phosphate frameworks containing PO_4 units, electrical conduction often occurs *via* hopping mechanisms involving electron transfer between mixed-valence states such as P^{4+} and P^{5+} .^{2,30,31} This process generates localized distortions in the lattice, giving rise to polarons quasiparticles formed by an electronic charge coupled with lattice deformation. The dynamics of these polarons are governed by factors such as temperature, bond length variations, lattice vacancies, and electron–phonon interactions, all of which influence the overall conductivity and dielectric relaxation behavior. Complex impedance spectroscopy (CIS) is a particularly powerful technique for probing such transport phenomena. By analysing frequency- and temperature-dependent impedance responses, CIS can separate the contributions of grains, grain boundaries, and electrode interfaces.^{32,33} It also allows for the determination of key dielectric and transport parameters such as relaxation frequencies, relaxation times, dielectric constants, conductivity, dielectric losses, and electric modulus and enables the construction of equivalent circuit models to simulate resistive and capacitive elements in polycrystalline ceramics under alternating electric fields.^{32–35}

Numerous synthetic approaches have been reported for the preparation of oxide precursors, each providing control over the composition, particle size, morphology, and defect concentration of the final product.³⁶ The choice of synthesis route plays a crucial role in determining the phase purity and functional performance of complex oxides, as demonstrated in systems such as $\text{Nd}_{1-x}\text{Ba}_x\text{MnO}_{3-\delta}$.³⁷ Phosphate ceramics can be synthesized through various routes, including solid-state reactions,³⁵ sol-gel processing,³⁸ hydrothermal synthesis,³⁹ precipitation, and combustion methods.⁴⁰ Among these, the conventional solid-state route is the most widely employed due to its simplicity, scalability, and cost-effectiveness. It relies on high-temperature diffusion between precursor powders to achieve complete phase formation, though it often requires extended heating to ensure full reaction and densification.^{39–41} In this

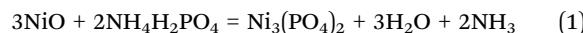


work, $\text{Ni}_3(\text{PO}_4)_2$ ceramic was synthesized *via* the solid-state reaction method and sintered at 900 °C. The optical band gap was determined using UV-Visible spectroscopy. A comprehensive investigation of its dielectric and transport properties was carried out over a wide temperature range (393–633 K) employing impedance (Z^*), permittivity (ε^*), electric modulus (M^*), and electrical conductivity (σ^*) formalisms. These findings offer valuable insights into the structure–property correlations in $\text{Ni}_3(\text{PO}_4)_2$, underscoring its promise as a functional ceramic for energy storage, sensing, and frequency-dependent device applications.

2. Methodology

2.1 Solid-state process

Various synthesis methods have been developed for oxide precursors, enabling precise control over composition, particle size, morphology, and defect concentration.³⁶ The selected synthesis route critically influences the phase purity and functional properties of complex oxides.³⁷ In this study, $\text{Ni}_3(\text{PO}_4)_2$ was synthesized using the conventional solid-state method. Eqn (1) illustrates the chemical reaction leading to the formation of the $\text{Ni}_3(\text{PO}_4)_2$ compound.



Analytical-grade NiO and $\text{NH}_4\text{H}_2\text{PO}_4$ powders were accurately weighed in stoichiometric proportions and blended for 3 hours at ambient temperature using an agate mortar to ensure uniform mixing. The homogenized mixture was compacted into pellets with a diameter of 13 mm under a uniaxial load of 3 tons per cm^2 , followed by preheating on a hot plate at 350 °C for 10 hours to eliminate volatile species. The preheated material

was subsequently ground for 1 hour and re-compacted into 8 mm pellets. These were sintered at 900 °C for 3 hours in an alumina crucible to facilitate phase formation and improve densification. The final sintered products reached approximately 97% of their theoretical density. A schematic representation of the synthesis steps is provided in Fig. 1.

2.2 Instrumentation

Powder X-ray diffraction (PXRD) is a highly effective technique for assessing the crystallinity of the sample. Measurements were performed at room temperature using an 8D-Bruker Advance X-ray diffractometer, operating over a 2θ range of 10–80°, and equipped with a copper anode ($\text{CuK}\alpha$, $\lambda = 0.15406 \text{ \AA}$).

A Shimadzu UV-3101PC spectrophotometer was employed to record the UV-Vis absorption spectrum of the powdered sample.

Electrical and dielectric measurements were performed using a Solartron 1260 impedance analyzer, covering a broad frequency range from 0.1 Hz to 1 MHz and a temperature range of 393–633 K. The dielectric characterization aimed to investigate grain and grain boundary contributions to the electrical properties of the prepared $\text{Ni}_3(\text{PO}_4)_2$ pellets. The samples, fabricated in disc form with an 8 mm diameter and 1 mm thickness, were coated with a thin conductive layer (gold or silver) on both surfaces to ensure optimal and stable electrical contact. The pellets were placed between copper or platinum electrodes inside a temperature-controlled chamber equipped with a precision closed-loop feedback system, achieving a temperature resolution of 0.1 K and stability better than ± 0.5 K. Measurements were carried out under standard atmospheric pressure with an applied AC voltage of 0.5 V, ensuring

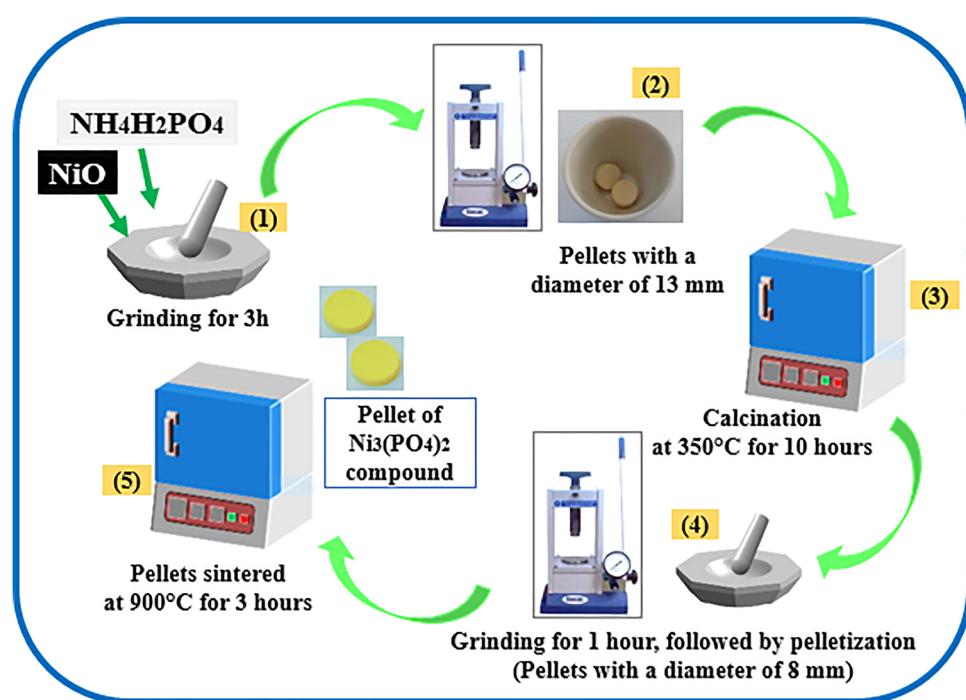


Fig. 1 Schematic illustration of the $\text{Ni}_3(\text{PO}_4)_2$ synthesis process *via* the solid-state reaction method.



reproducible results over the entire frequency and temperature ranges.

3. Results with discussion

3.1. Experimental and theoretical structural studies

X-ray powder diffraction (XRPD) analysis confirmed that $\text{Ni}_3(\text{PO}_4)_2$ crystallizes in the monoclinic system, adopting the $P2_1/c$ space group. The diffraction pattern in Fig. 2 exhibits sharp and well-defined Bragg peaks, all of which are correctly indexed and precisely modelled, demonstrating the robustness of the structural refinement. Rietveld analysis of the room-temperature data yielded lattice parameters $a = 5.843 \text{ \AA}$, $b = 4.692 \text{ \AA}$, $c = 10.088 \text{ \AA}$, and $\beta = 90.676^\circ$, in excellent agreement with literature values.^{29,42,43} The refinement quality, indicated by the agreement factors ($R_P = 5.46\%$, $R_{WP} = 7.33\%$, $R_{EXP} = 2.62\%$, Table 1), confirms a high degree of correlation between the observed and calculated patterns. These results verify the successful synthesis of phase-pure $\text{Ni}_3(\text{PO}_4)_2$.

The crystal structure of $\text{Ni}_3(\text{PO}_4)_2$, shown in Fig. 3 and obtained from the Crystallographic Information File (CIF) of ICSD card no. 4269,⁴³ consists of interconnected polyhedral units. Table 1 presents the fractional atomic coordinates of the atoms in the crystal structure. The positions of the oxygen atoms (O1–O4) indicate the arrangement of oxygen ions forming polyhedral units around Ni and P atoms, defining the overall crystal geometry, which consists of NiO_6 octahedra and PO_4 tetrahedra. These well-defined atomic positions also confirm the excellent stoichiometry of the $\text{Ni}_3(\text{PO}_4)_2$ structure.⁴⁴ Each $[\text{NiO}_6]$ octahedron shares oxygen edges with adjacent $[\text{NiO}_6]$ octahedra and is further linked to $[\text{PO}_4]$ tetrahedra through two common oxygen edges. Such connectivity may introduce O–O repulsion, which can weaken the $[\text{NiO}_6]$ – $[\text{PO}_4]$ interactions and influence the overall chemical stability of the lattice. However, the presence of strong covalent P–O bonds impart considerable structural integrity, ensuring that $\text{Ni}_3(\text{PO}_4)_2$ remains chemically stable.

The crystallite size and its distribution typically increase with rising annealing temperature as a result of thermally

Table 1 Refined structural parameters of $\text{Ni}_3(\text{PO}_4)_2$ compound at room temperature

Formula	$\text{Ni}_3(\text{PO}_4)_2$		
Crystal system	Monoclinic		
Space-group	$P2_1/c$		
Cell parameters	$a = 5.843 \text{ \AA}$		
	$b = 4.692 \text{ \AA}$		
	$c = 10.088 \text{ \AA}$		
	$\beta = 90.676^\circ$		
Cell volume	276.566 \AA^3		
Z	2		
Atoms	x	y	z
Ni(1)	0.5000	0	0
Ni(2)	0.2416	–0.0169	0.2751
P	0.2507	0.4200	0.0932
O(1)	0.2590	0.7416	0.1054
O(2)	0.2521	0.1964	0.4509
O(3)	0.0464	0.3013	0.1734
O(4)	0.4618	0.2671	0.1613
R -factors			
$R_P/R_{WP}/R_{EXP} (\%)$	5.46/7.33/2.62		
χ^2	1.06		

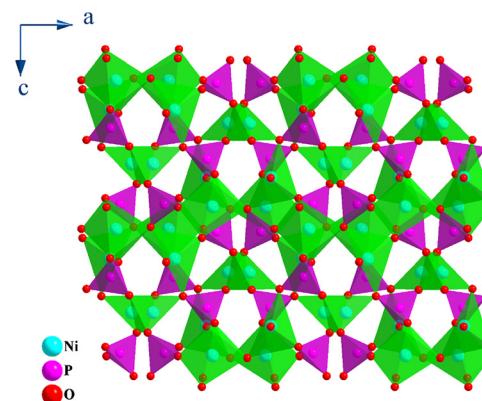


Fig. 3 Crystal structure of $\text{Ni}_3(\text{PO}_4)_2$ projected along the b -axis.

activated grain growth.⁴⁵ The enlargement of grains enhances intergranular connectivity, thereby facilitating charge transport through improved electronic and ionic pathways and consequently enhancing the capacitance behavior. In the present work, $\text{Ni}_3(\text{PO}_4)_2$ was synthesized under optimized annealing conditions to achieve a single-phase structure with a homogeneous crystallite size distribution. Generally, elevated annealing temperatures promote further grain coarsening, which can improve transport characteristics; however, excessive thermal treatment may also trigger oxygen diffusion and mass transport, leading to the formation of vacancies or grain boundary defects that degrade dielectric and electrical performance.^{46–48}

The average crystallite size of the $\text{Ni}_3(\text{PO}_4)_2$ sample was determined from the most intense diffraction peak ($2\theta = 29.7959^\circ$) (Fig. 4(a)) using the Debye–Scherrer equation:⁴⁶

$$D_{sc} = \frac{0.9\lambda}{\beta \cos(\theta)} \quad (2)$$

Here, λ denotes the X-ray wavelength (1.54059 \AA), θ the diffraction angle, and β the full width at half maximum (FWHM).

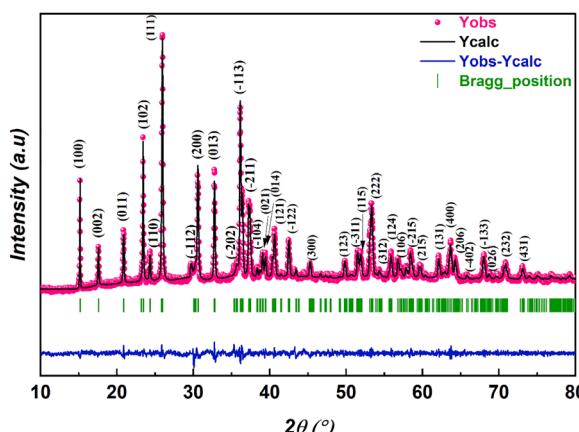


Fig. 2 Room-temperature Rietveld refinement of the X-ray powder diffraction pattern of $\text{Ni}_3(\text{PO}_4)_2$ compound.



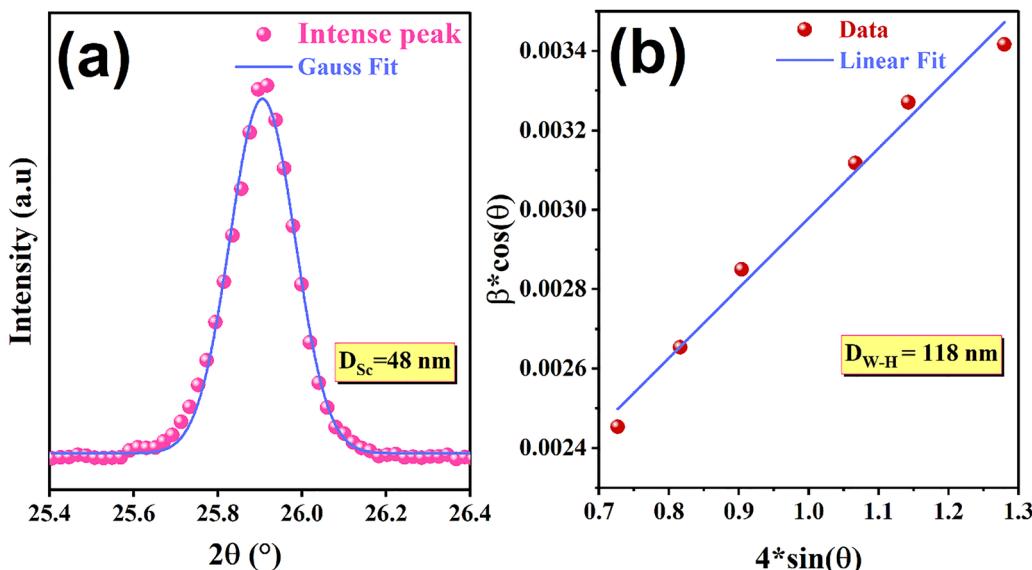


Fig. 4 (a) The Debye–Scherrer plot and (b) the Williamson–Hall plot for $\text{Ni}_3(\text{PO}_4)_2$ compound.

Based on the above equation, the mean crystallite size is calculated to be around 48 nm. Furthermore, the Williamson–Hall (W–H) analysis was performed (Fig. 4(b)). The procedure for constructing the W–H plot is described by eqn (3):⁴⁷

$$\beta \times \cos(\theta) = \frac{0.9\lambda}{D_{\text{W-H}}} + 4\varepsilon \times \sin(\theta) \quad (3)$$

In the Williamson–Hall (W–H) method, $4 \sin(\theta)$ was plotted along the x -axis and $\beta \cos(\theta)$ along the y -axis. The average crystallite size obtained from the W–H plot is approximately 118 nm. It is noteworthy that the size calculated using the Scherrer approach is smaller than that estimated by the W–H model. This discrepancy arises from the fact that the Scherrer approach neglects the strain contribution (ε), suggesting that the line broadening in the compound is influenced not only by crystallite size but also by lattice deformation.⁴⁸

3.2. Optical response by UV-Visible spectroscopy

The UV-Vis absorption and reflectance spectra of $\text{Ni}_3(\text{PO}_4)_2$ (Fig. 5(a)) exhibit well-defined bands centred at 430, 474, 706, and 810 nm, along with a pronounced absorption edge below 300 nm in the ultraviolet range. This high-energy feature originates from charge transfer processes between O^{2-} and Ni^{2+} ions, a phenomenon frequently observed in transition-metal phosphate systems.^{49,50} The peaks observed in the visible to near-infrared region are linked to intra-d electronic transitions of Ni^{2+} (3d⁸ configuration) under an octahedral crystal field. More precisely, the absorptions at ~ 430 and 474 nm correspond to the ${}^3\text{A}_2\text{g}(\text{F}) \rightarrow {}^3\text{T}_{1\text{g}}(\text{P})$ transition, whereas those near 706 and 810 nm are attributed to the ${}^3\text{A}_2\text{g}(\text{F}) \rightarrow {}^3\text{T}_{1\text{g}}(\text{F})$ and ${}^3\text{A}_2\text{g}(\text{F}) \rightarrow {}^3\text{T}_{2\text{g}}(\text{F})$ transitions, respectively.^{51–53} These optical signatures validate the octahedral environment of Ni^{2+} ions in the $\text{Ni}_3(\text{PO}_4)_2$ lattice and align with reported spectra of related nickel-containing materials.⁵⁴ The visible absorption is also responsible for the typical green

appearance of nickel phosphates and highlights their promise for applications in optoelectronics and catalysis.

The optical band gap is a fundamental parameter in many scientific domains, including photovoltaics, lasers, photocatalysis, light-emitting diodes, and photoluminescence.⁵⁵ It represents the electronic transition at the absorption edge, corresponding to the energy difference between the highest occupied state of the valence band and the lowest unoccupied state of the conduction band.⁵⁶ The band gap energy was evaluated using the Marotti method,⁵⁷ which is based on plotting $dR/d\lambda$ as a function of wavelength (λ). The band gap corresponds to the maximum point of this curve, as illustrated in Fig. 5(b), and was estimated using the relation $E_g = 1240/262 \approx 4.73$ eV. To more accurately determine both the value and the nature of the optical band gap, the Schuster–Kubelka–Munk (SKM) model is often employed. This approach is widely used to analyse the absorption characteristics of semiconductors based on their diffuse reflectance spectra, as expressed by the following relation:^{58,59}

$$F(R) = \frac{(1 - R)^2}{2R} \quad (4)$$

The optical band gap of the examined material can be estimated using Tauc's method, which is defined by the following equation:⁶⁰

$$(F(R)h\nu)^{\frac{1}{n}} = A(h\nu - E_g) \quad (5)$$

where " $h\nu$ " represents the photon energy, "A" denotes the optical absorption coefficient, and "n" takes the value of "1/2" for direct transitions or "2" for indirect transitions. Fig. 5(c) presents the plots of $(F(R)h\nu)^2$ and $(F(R)h\nu)^{1/2}$ as a function of photon energy ($h\nu$). From these curves, the optical band gap was extracted, yielding $E_{\text{gd}} = 4.75$ eV for the direct transition and $E_{\text{gi}} = 4.36$ eV for the indirect transition. To further verify the nature of the optical band gap (direct or indirect), eqn (6) was rearranged, leading to the following expression:



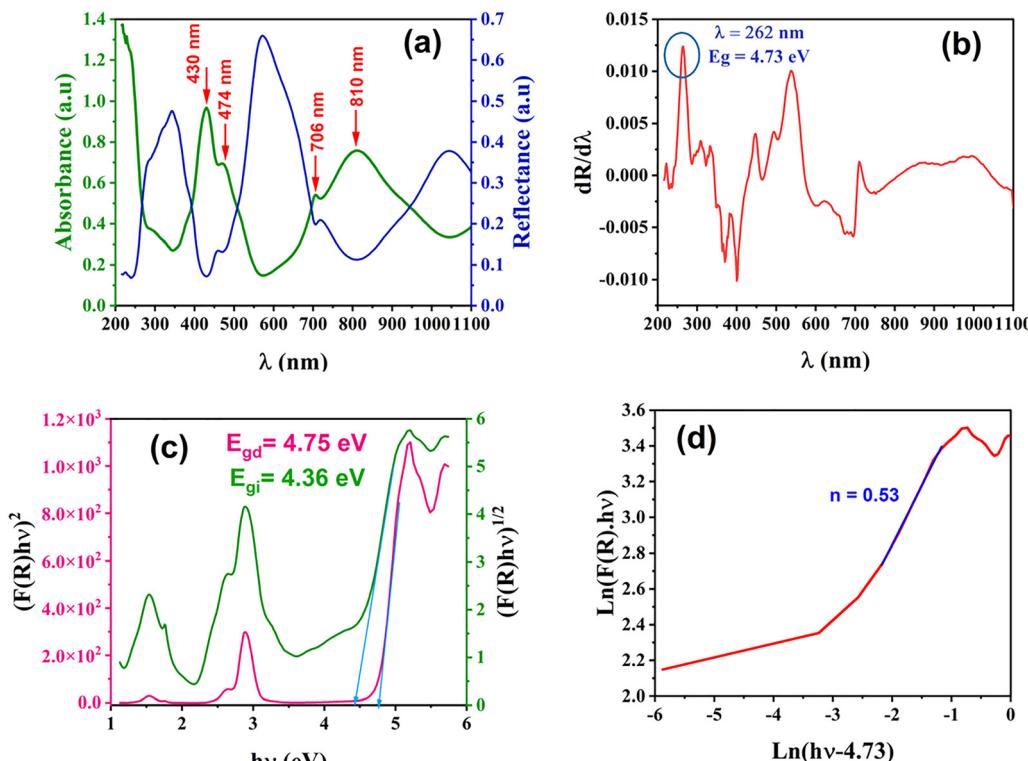


Fig. 5 (a) UV-Visible absorption and reflectance spectra of $\text{Ni}_3(\text{PO}_4)_2$. (b) Variation of the derivative of reflectance ($dR/d\lambda$) with respect to wavelength (λ). (c) Plots of $(F(R)h\nu)^{1/2}$ and $(F(R)h\nu)^2$ as functions of photon energy ($h\nu$). (d) Dependence of $\ln(\alpha h\nu)$ on $\ln(h\nu - 4.73)$.

$$\ln(F(R)h\nu) = \ln(A) + n \ln(h\nu - E_g) \quad (6)$$

From the linear fitting of the curve, a slope value of 0.53 was obtained (see Fig. 5(d)), corresponding to $n \approx 1/2$, which confirms the direct nature of the optical transition in the $\text{Ni}_3(\text{PO}_4)_2$ compound. It is clear that there is only a slight difference between the band gap energy determined from the derivative method and that obtained from the Tauc plot, which can be attributed to the uncertainty ΔE_g . These results further confirm that the compound possesses a well-defined direct band gap of 4.74 ± 0.01 eV. This nature of the band gap is consistent with those reported for other phosphate-based materials, such as α -KYP₂O₇⁶¹ and K₂Ba₃(P₂O₇)₂⁶².

Since the $\text{Ni}_3(\text{PO}_4)_2$ compound shows direct band gap, we estimated the conduction band minimum (E_{CB}) and valence band maximum (E_{VB}) using the empirical formula:⁵⁹

$$E_{CB} = \chi - E_e - \frac{1}{2}E_g \quad (7)$$

$$E_{VB} = E_{CB} + E_g \quad (8)$$

In this relation, χ denotes the absolute electronegativity of the compound, whereas E_e is a fundamental constant corresponding to the free-electron energy on the hydrogen scale, commonly taken as 4.5 eV. The electronegativity of the compound was estimated using the following relation:⁵⁹

$$\chi = [\chi(A^\alpha)\chi(B^\beta)\chi(O^\gamma)]^{\frac{1}{(\alpha+\beta+\gamma)}} \quad (9)$$

In this framework, α , β , and γ correspond to the atomic multiplicities of Ni, P, and O, respectively. From our calculations, the absolute electronegativity of the sample was found to be 6.5 eV. By combining this value with the calculated E_{VB} , E_{CB} , and E_{gd} , a schematic band structure was constructed, as shown in Fig. 6(a).

The disorder in the material was evaluated by calculating the Urbach energy E_u , which is expressed as follows:⁶³

$$\ln(\alpha) = \ln(\alpha_0) + \frac{h\nu}{E_u} \quad (10)$$

Based on the $\ln(\alpha)$ versus $h\nu$ plot (Fig. 6(b)), the Urbach energy (E_u) was found to be 0.47 eV. This relatively low value suggests that the material has minimal structural disorder.

3.3. Impedance spectral response analysis

3.3.1. Frequency dependence of the real and imaginary components of impedance. Impedance spectroscopy (IS) is a powerful, non-destructive analytical method for examining phosphate-based electrolyte materials by monitoring their current response over a range of frequencies and temperatures. Since its first application in electrical engineering by Oliver Heaviside in 1880,⁶⁴ IS has evolved to serve diverse domains such as dielectric, electrochemical, and bioelectrical analysis, with applications spanning biomaterials, photovoltaic devices, sensing, frequency-dependent device applications, and electrochemical systems.^{65,66} The technique enables separation of the electrical contributions from different regions of a material,



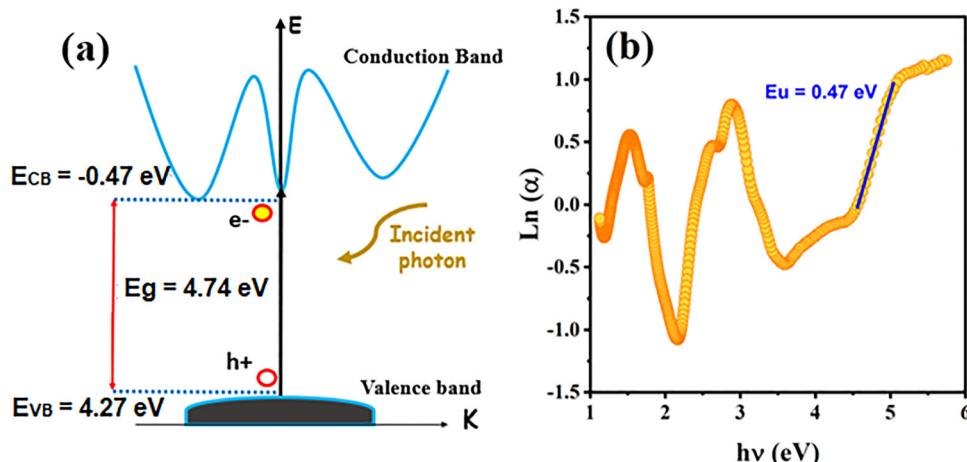


Fig. 6 (a) Schematic representation of the band structure. (b) Variation of $\ln(\alpha)$ with incident photon energy.

including grain interiors (bulk), grain boundaries, and electrode–material interfaces, each characterized by distinct relaxation behaviors and resistive–capacitive properties.³² These features are identified through analysis of impedance spectra, which are often plotted in the complex plane to derive parameters such as resistance, capacitance, and relaxation time. Such parameters are determined using established theoretical models and impedance relations that link the real (Z') and imaginary (Z'') components of the complex impedance vector (Z^*):

$$Z^* = Z' + jZ'' \quad (11)$$

Fig. 7(a) presents the variation of the real component of the complex impedance (Z') with angular frequency for the $\text{Ni}_3(\text{PO}_4)_2$ material over the temperature range of 393–633 K. The frequency response reveals two distinct regions. In the low-frequency range, Z' remains almost constant, signifying DC conduction dominated by long-range charge migration, where carriers successfully hop between adjacent sites.^{2,67,68} At higher frequencies, Z' displays a pronounced dispersion, which is associated with AC conduction arising from localized carrier motion and incomplete hopping events. This behavior indicates the coexistence of two conduction mechanisms: extended-range charge transport at low frequencies and localized relaxation processes at high frequencies, which contribute to AC conductivity.

As the temperature increases, the frequency at which Z' begins to deviate from its low-frequency plateau shifts toward higher values from around 18 rad s⁻¹ at 393 K to approximately 800 rad s⁻¹ at 633 K implying a thermally activated relaxation process. In the high-frequency region, Z' approaches a minimum, nearly temperature-independent value, suggesting that thermal energy reduces barriers to charge movement. Conversely, in the low-frequency regime, Z' decreases noticeably with rising temperature, reflecting a negative temperature coefficient of resistance (NTCR). This behavior is characteristic of semiconductors and confirms the material's NTC thermistor nature.^{32,35}

The imaginary part of the impedance spectrum serves as a key indicator of the dielectric relaxation characteristics of the

$\text{Ni}_3(\text{PO}_4)_2$ material, which arise from either charge carrier motion or the reorientation of dipolar groups when subjected

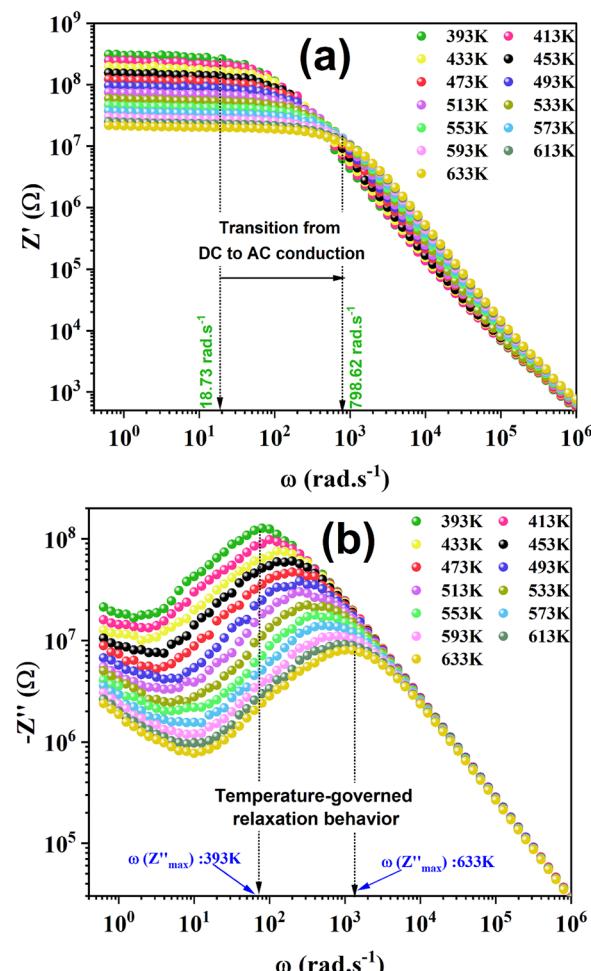


Fig. 7 (a) Temperature-dependent variation of the real component of complex impedance (Z') with angular frequency in the range of 393–633 K. (b) Temperature-dependent variation of the imaginary component of complex impedance ($-Z''$) with angular frequency in the range of 393–633 K.

to an external electric field. In Fig. 7(b), the variation of $-Z''$ with angular frequency is presented for temperatures between 393 K and 633 K, illustrating how relaxation dynamics evolve with increasing thermal energy. Each curve displays a distinct peak, marking the transition region between DC and AC conduction, consistent with the features observed in the Z' spectra (Fig. 7(a)). This peak, known as the relaxation frequency, represents the crossover from extended charge transport to more localized carrier motion. The progressive broadening of these peaks with temperature suggests a deviation from the ideal Debye relaxation behavior,⁶⁹ highlighting the intricate interplay between polarization and conduction mechanisms in the system. At frequencies below the peak, charge transport is dominated by long-range hopping processes, while frequencies above the peak are primarily associated with localized relaxation effects.

Fig. 8(a) presents the variation of relaxation time (τ), extracted from the Z'' data, as a function of temperature. A noticeable decrease in τ with increasing temperature reveals a thermally activated relaxation mechanism. This behavior is primarily associated with ionic conduction, likely governed by polaron hopping facilitated by local lattice vibrations, often referred to as "rattling" motions. The reduction in τ reflects enhanced charge carrier mobility and progressive delocalization.^{70,71} Moreover, this evolution is accompanied by a tendency toward more uniform relaxation, approaching Debye-like behavior at higher temperatures, indicative of a narrowing relaxation time distribution.

For further insight, the imaginary part of the electric impedance, $Z''(\omega, T)$, was normalized by its maximum value, Z''_{\max} , and plotted against the logarithm of the normalized angular frequency, $\log(\omega/\omega_{\max})$, where ω_{\max} corresponds to the peak-loss frequency. Here, T denotes temperature. The convergence of impedance loss curves obtained at different temperatures into a single master curve suggests that the relaxation time distribution is temperature-independent. As depicted in Fig. 8(b), the nearly perfect overlap of the curves confirms that the relaxation dynamics remain unchanged over the investigated temperature range.⁷²

3.3.2. Nyquist plot representation analysis ($-Z''$ vs. Z'). The Nyquist plot, serves as an effective tool for exploring the electrical response of materials. It enables the identification

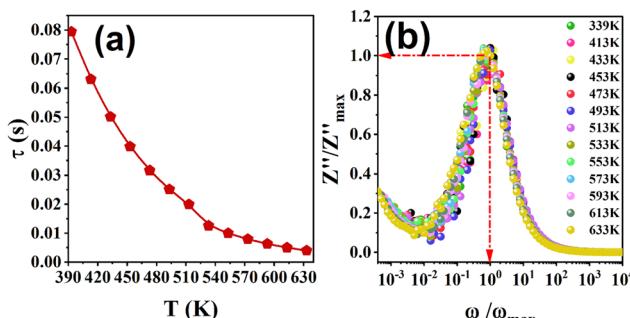


Fig. 8 (a) Temperature dependence of the relaxation time (τ). (b) Normalized variation of Z''/Z''_{\max} as a function of ω/ω_{\max} at different temperatures for the $\text{Ni}_3(\text{PO}_4)_2$ compound.

of surface layer formation due to undesirable chemical processes such as oxidation or hydration, and it reveals interfacial effects at the electrode–material interface. Furthermore, Nyquist analysis provides valuable information on the type and mobility of charge carriers, as well as on the separate contributions from the bulk (grain) and interfacial (grain boundary) regions of the material.³² Fig. 9(a) presents the complex impedance spectra in Nyquist form, displaying depressed and asymmetric semicircular arcs whose diameters decrease progressively with increasing temperature. This trend reflects non-Debye relaxation, associated with a distribution of relaxation times arising from heterogeneous dipolar environments within the sample.⁷³ A single arc in the plot may appear in two situations: (i) when grain boundaries are highly conductive, resulting in negligible impedance and no distinct arc, or (ii) when the bulk phase exhibits high conductivity, minimizing its impedance contribution and leaving only the grain boundary response visible. The shift of the real-axis intercept and the peak frequency of the semicircular arcs toward the origin with temperature rise indicates enhanced electrical conductivity.

To quantitatively assess the grain boundary contribution, the depressed semicircular arcs were analyzed using an equivalent circuit model in Z_{view} software,^{61,62} as illustrated in Fig. 9(b). This figure presents both the sample geometry used for complex impedance spectroscopy (CIS) measurements and the corresponding equivalent circuit, comprising a resistor (R), a capacitor (C), and a constant phase element (CPE) in parallel. This configuration effectively models the non-ideal dielectric behavior, allowing a clear distinction between bulk and grain boundary contributions. During the Z_{view} fitting process, the

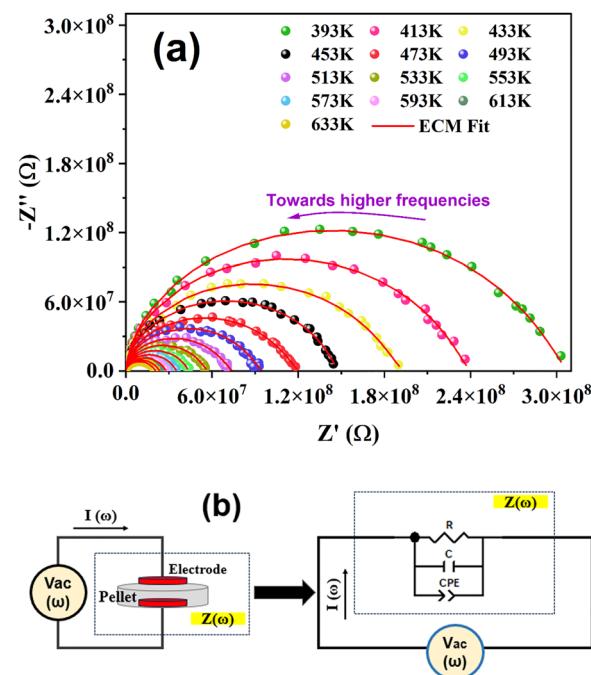


Fig. 9 (a) Nyquist plots (Z'' vs. Z') for $\text{Ni}_3(\text{PO}_4)_2$ compound at various temperatures. (b) Schematic illustration along with the sample geometry for CIS measurement and the equivalent circuit model used to fit the data.



Table 2 Electrical parameters extracted from Z_{view} fitting at selected temperatures

T (K)	R (10^7 Ω)	C (10^{-11} F)	CPE- T (10^{-10} F)	CPE- P
393	30.75	3.418	1.063	0.60456
413	23.99	3.424	1.091	0.57810
433	19.33	3.433	1.359	0.54744
453	14.82	3.441	1.711	0.61646
473	12.04	3.449	1.942	0.54243
493	9.492	3.466	2.146	0.53375
513	7.425	3.477	2.474	0.51353
533	5.710	3.488	2.699	0.53197
553	4.397	3.498	2.809	0.58448
573	3.545	3.512	3.145	0.56844
593	2.831	3.531	3.538	0.55779
613	2.365	3.546	4.052	0.54754
633	2.083	3.56	4.462	0.54283

constant phase element (CPE) is typically characterized by two parameters, CPE- T (Q) and CPE- P (α), and is mathematically defined as:

$$Z_{\text{CPE}} = \frac{1}{[(\text{CPE-}T)(j\omega)]^{\text{CPE-}P}} \quad (12)$$

Here, CPE- T is expressed in units of capacitance, while the angular frequency (ω) is given in radians per second. The exponent CPE- P takes values of 1, 0, and 0.5, corresponding to ideal capacitive, resistive, and Warburg-type behaviors, respectively.⁷⁴ The parameters obtained from the equivalent circuit fitting are summarized in Table 2, which lists the calculated resistance (R), capacitance (C), and constant phase element (CPE) values for the $\text{Ni}_3(\text{PO}_4)_2$ compound. The high grain boundary resistance is likely due to structural defects such as voids or micro-gaps in these regions. The measured capacitance values, on the order of $\sim 10^{-11}$ F for C and $\sim 10^{-10}$ F for Q , suggest that at higher frequencies, conduction is predominantly confined to the grains, whereas at lower frequencies, it is mainly governed by grain boundary effects. The increase of both parameters with temperature suggests enhanced dielectric polarization and improved charge carrier mobility within the material.

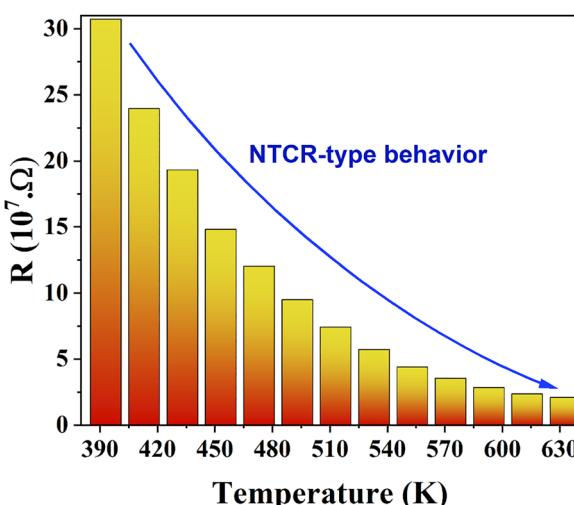


Fig. 10 Evolution of the resistivity against the temperature.

Fig. 10 illustrates the temperature dependence of the electrical resistivity of the prepared solid $\text{Ni}_3(\text{PO}_4)_2$ over a wide temperature range (393–633 K). The results reveal an exponential decrease in resistivity with increasing temperature, confirming the semiconducting nature of the sample. This behavior is attributed to thermally activated charge carriers gaining higher mobility at elevated temperatures, which is characteristic of a negative temperature coefficient of resistance (NTCR) and reflects the thermal sensitivity of both resistive and capacitive components.³²

3.4. Electrical conduction: from DC to AC regimes

Analyzing the electrical conductivity (σ) as a function of frequency and temperature is crucial for understanding the sample's electrical response and identifying the dominant conduction mechanism. The conductivity is evaluated using the following relation:

$$\sigma = \frac{e}{s} \frac{Z'}{Z'^2 + Z''^2} \quad (13)$$

Where “ e ” and “ s ” refer to the thickness and area of the pellet, respectively.

Fig. 11(a) illustrates the frequency-dependent conductivity spectra of $\text{Ni}_3(\text{PO}_4)_2$ measured over the temperature range 393–633 K. The spectra exhibit two main regions: a frequency-independent plateau at low frequencies and a dispersive region extending up to ~ 100 rad s^{-1} . Beyond this range, a slope-changing dispersive regime emerges, progressively shifting toward higher frequencies as the temperature increases. The low-frequency plateau reflects the long-range transport of charge carriers through hopping processes, corresponding to the DC conductivity of the material. In this regime, carrier motion can be approximated as hopping within an extended lattice of equivalent potential wells. The dispersive region up to 100 rad s^{-1} is attributed to short-range hopping between localized sites, confined to a double-well potential separated by large barriers, enabling forward-backward hopping.^{75,76} Under the influence of an AC electric field, this process becomes perturbed, resulting in an exponential rise in conductivity.⁷⁶ At frequencies exceeding 100 rad s^{-1} , the response is dominated by localized carrier dynamics, most likely associated with polaron hopping coupled with anharmonic lattice vibrations in the $\text{Ni}_3(\text{PO}_4)_2$ framework. With increasing frequency, the conductivity continues to rise. At elevated temperatures, the DC plateau in the low-frequency region becomes more dominant compared to lower temperatures. The critical frequency at which the conductivity deviates from the DC plateau and begins to increase is identified as the hopping frequency (ω_H). Hence, the conductivity response across the studied temperature interval can be accurately modeled by Jonscher's universal power law:⁷⁵

$$\sigma(T, \omega) = \sigma_{\text{dc}}(T) + \sigma_{\text{ac}}(\omega) = \sigma_{\text{dc}} \left[1 + \left(\frac{\omega}{\omega_H} \right)^s \right] \quad (14)$$

In this relation, the frequency exponent s lies within the range $0 < s < 1$,⁷⁵ and its value is influenced by the degree of interaction among mobile ions. According to Jonscher's model,



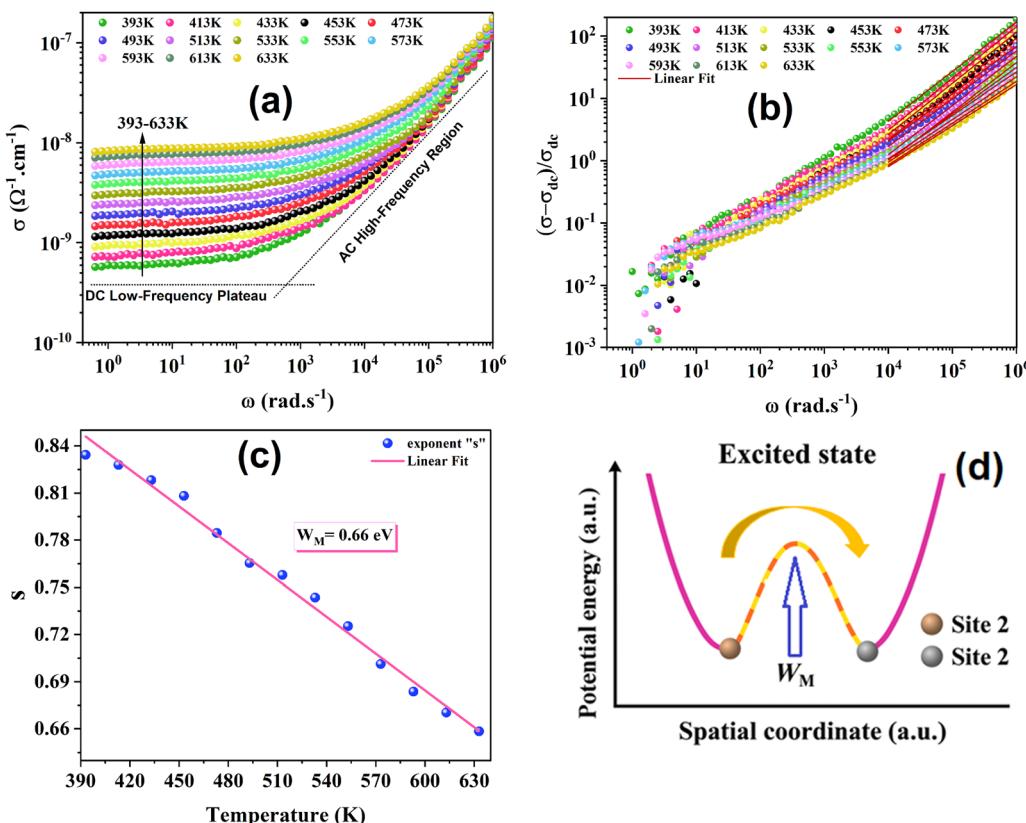


Fig. 11 (a) Frequency-dependent variation of the electrical conductivity of $\text{Ni}_3(\text{PO}_4)_2$ at different temperatures. (b) Frequency dependence of $(\sigma - \sigma_{dc})/\sigma_{dc}$ vs. frequency plots at different temperatures. (c) variation of frequency exponent "s" against temperature. (d) Schematic model of the AC conduction mechanism in $\text{Ni}_3(\text{PO}_4)_2$.

the frequency dependence of conductivity originates from relaxation processes driven by mobile charge carriers. When a charge carrier hops from one site to another, it resides transiently between two potential energy minima before completing the transition. To gain a deeper understanding of AC conduction, it is essential to analyze how the frequency exponent "s" varies with both temperature and frequency. Various theoretical models such as Quantum Mechanical Tunneling (QMT), Correlated Barrier Hopping (CBH), Small Polaron Tunneling (SPT), and Overlapping Large Polaron Tunneling (OLPT) have been proposed to interpret these conduction processes.⁷⁷⁻⁸⁰ In their work, Nowick and co-workers⁸¹ suggested employing a normalized conductivity *versus* frequency plot to minimize DC conductivity contributions when extracting the Jonscher exponent "s". The resulting plot (Fig. 11(b)) demonstrates a clear linear region with slope "s", which characterizes the degree of interaction among charge carriers participating in the polarization and transport mechanisms.

Fig. 11(c) illustrates the temperature dependence of the frequency exponent s for $\text{Ni}_3(\text{PO}_4)_2$. A gradual decrease in s is observed within the 393–633 K range, indicating that the conduction mechanism in this compound is consistent with the Correlated Barrier Hopping (CBH) model.^{79,82} Within the framework of this model, the exponent "s" is expressed by the following relation:⁸³

$$s = 1 - \frac{6K_B T}{W_M} \quad (15)$$

The self-trapping energy (W_M), determined from eqn (15) in Fig. 11(c), is found to be 0.66 eV. The electrical transport in $\text{Ni}_3(\text{PO}_4)_2$ is mainly controlled by a small-polaron hopping process, in which charge carriers become localized on Ni sites and move between adjacent Ni ions through successive $\text{Ni}^{2+}/\text{Ni}^{3+}$ redox transitions. This hopping is accompanied by a local lattice distortion within the phosphate network, characteristic of polaronic conduction. Fig. 11(d) depicts the carrier hopping process in $\text{Ni}_3(\text{PO}_4)_2$, which is predominantly governed by the Correlated Barrier Hopping (CBH) mechanism. In this model, electronic conduction originates from thermally activated hopping of localized electrons between neighboring Ni cation sites, with the phosphate framework playing a crucial role in modulating the potential barriers that control carrier motion. At high frequencies, the charge transport becomes increasingly restricted to short-range polaron hopping, reflecting localized carrier dynamics. As the temperature rises, the effective potential barriers decrease, thereby enhancing carrier mobility and facilitating more efficient charge transport. The dominance of the CBH mechanism in $\text{Ni}_3(\text{PO}_4)_2$ aligns well with observations in related phosphate systems such as $\text{K}_2\text{NiP}_2\text{O}_7$,²⁵ $\text{LiNi}_{1.5}\text{P}_2\text{O}_7$,⁸⁴ $\text{Na}_{3.6}\text{Ni}_{2.2}(\text{P}_2\text{O}_7)_2$,⁸⁵ and TiFeP_2O_7 ² (see Table 3), confirming that thermally activated hopping between localized states governs the conduction behavior consistent with the non-Debye relaxation character of these materials.

To assess the activation energy in $\text{Ni}_3(\text{PO}_4)_2$, Fig. 12 displays the plot of $\ln(\sigma_{\text{dc}} \times T)$ against $1000/T$. The activation energy (E_a) was extracted from this relationship using the following equation:⁸³

$$\sigma_{\text{dc}} = \frac{\sigma_0}{T} \exp\left(\frac{-E_a}{K_B T}\right) \quad (16)$$

Here, E_a represents the activation energy, k_B is the Boltzmann constant, and T is the absolute temperature in kelvin. The activation energy determined for $\text{Ni}_3(\text{PO}_4)_2$ is 0.28 eV, which is markedly lower than that of other Ni-based phosphate compounds, including LiNiPO_4 ,⁸⁶ $\text{LiNi}_{1.5}\text{P}_2\text{O}_7$,⁸⁴ and $\text{K}_2\text{NiP}_2\text{O}_7$,²⁵ and Ni-based other compounds listed in Table 3. This relatively small value indicates that charge carriers in $\text{Ni}_3(\text{PO}_4)_2$ require significantly less thermal energy to surmount the potential barriers governing conduction. In alkali-containing nickel phosphates, the presence of Li^+ or K^+ ions introduces additional electrostatic interactions and structural constraints, which can restrict carrier mobility and raise the energy barriers for hopping or migration. By contrast, the simpler framework of $\text{Ni}_3(\text{PO}_4)_2$, composed only of Ni^{2+} and phosphate groups, offers more open conduction pathways with reduced potential barriers, thus favoring charge transport at lower thermal energies.

3.5. Frequency-dependent dielectric polarization analysis

In recent years, inorganic phosphate materials have attracted considerable interest due to their remarkable dielectric properties.^{26,88,89} In this work, the dielectric behavior of the synthesized $\text{Ni}_3(\text{PO}_4)_2$ ceramic was investigated by analyzing the effects of frequency and temperature on the real part of the dielectric permittivity (ϵ') and the dielectric loss factor ($\tan \delta$). As shown in Fig. 13(a), ϵ' exhibits a strong dependence on both frequency and temperature. At lower frequencies, the permittivity values are significantly higher and increase with temperature, reaching around 1.58×10^4 at 633 K. In this work, the real part of the dielectric permittivity (ϵ') was calculated from the complex impedance measurements using the following relation:³⁵

$$\epsilon' = \frac{Z''}{\omega C_0(Z'^2 + Z''^2)} \quad (17)$$

The capacitance (C_0) is determined using the relation ($C_0 = (\epsilon_0 A)/d$), where (ϵ_0) is the permittivity of free space, “ A ” is the electrode area, and “ d ” is the thickness of the pellet.

The real part of the dielectric constant (ϵ') indicates how effectively a material can store electric energy under an applied field. As shown in Fig. 13(a), ϵ' is relatively high at low frequencies but gradually decreases as the frequency increases. This behavior is attributed to the combined effects of various polarization processes, including interfacial (space charge) and dipolar polarization, which govern the material's electrical response.^{90,91} At higher frequencies, slower polarization mechanisms cannot follow the rapidly oscillating field, leading to a reduction in ϵ' . At low frequencies (Region I, $< 10^2 \text{ rad s}^{-1}$), interfacial polarization dominates due to charge accumulation at grain boundaries and interfaces. At higher frequencies (Region II), dipolar polarization becomes more significant. Both contributions decrease as frequency continues to rise.⁹² Additionally, ϵ' decreases with decreasing temperature because

Table 3 Reported activation energies (E_a), dielectric constants (ϵ'), and conduction models of phosphate-based compounds

Compound	Temperature range (K)	Activation energy (eV)	ϵ'	Conduction model	Ref.
$\text{K}_2\text{NiP}_2\text{O}_7$	603–728	1	$\approx 10^3$	CBH	25
LiNiPO_4	423–573	0.68	$\approx 2 \times 10^4$	—	26 and 86
$\text{LiNi}_{1.5}\text{P}_2\text{O}_7$	633–723	0.85	—	CBH	84
$\text{Na}_{3.6}\text{Ni}_{2.2}(\text{P}_2\text{O}_7)_2$	594–710	1.03	—	CBH	85
SrNiP_2O_7	609–728	0.97	—	OLPT	87
$\text{NaMn}_4(\text{PO}_4)_3$	300–573	—	≈ 90	—	10
TlFeP_2O_7	473–673	0.52–1.3	$\approx 3 \times 10^4$	CBH	2
CsFeP_2O_7	453–673	0.13–0.44	$\approx 2 \times 10^3$	—	35
$\text{Ni}_3(\text{PO}_4)_2$	393–633	0.28	1.58×10^4	CBH	This work

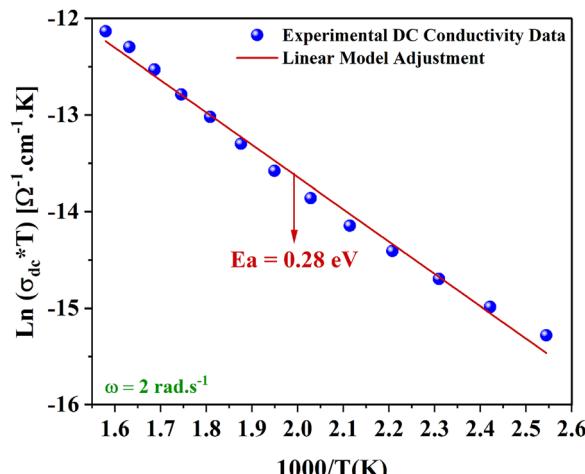


Fig. 12 Plot of $\ln(\sigma_{\text{dc}} \times T)$ versus $1000/T$.

charge carriers become less mobile and respond more slowly to the applied field. The dielectric constant of $\text{Ni}_3(\text{PO}_4)_2$ reaches approximately 1.58×10^4 at 633 K, reflecting its strong ability to resist electric field penetration and sustain polarization.⁶⁹ As shown in Table 3, most phosphate-based compounds exhibit comparably high dielectric constants in the range of 10^3 – 10^4 , confirming their high polarizability and suitability for dielectric applications. $\text{Ni}_3(\text{PO}_4)_2$ stand out for their pronounced dielectric responses, which may arise from efficient polarization processes involving defect dipoles, space charge accumulation, or localized hopping of charge carriers.

The pronounced dielectric response is primarily influenced by interfacial effects, particularly space charge polarization at grain boundaries, in line with the Maxwell–Wagner model.^{93,94} This occurs in heterogeneous materials where conductive grains are separated by less conductive boundaries (inset Fig. 13(a)). Under low-frequency electric fields, charges accumulate at these interfaces, enhancing polarization.⁹⁴ The extent of this effect depends on differences in conductivity and permittivity between grains and boundaries, structural imperfections, and temperature.



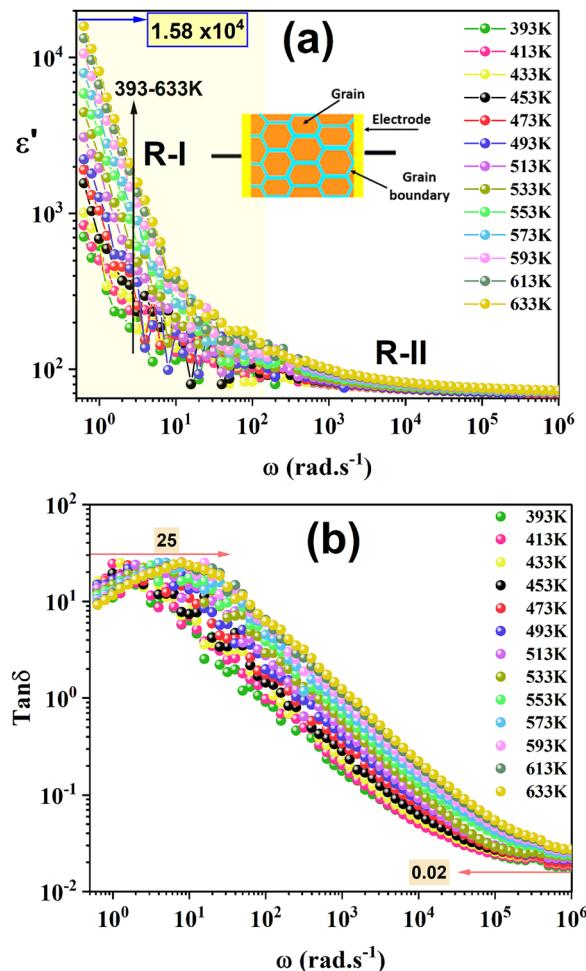


Fig. 13 (a) Frequency dependence of the real dielectric constant (ϵ') for $\text{Ni}_3(\text{PO}_4)_2$, with inset showing the grain and grain boundary structure. (b) Variation of $\tan \delta$ with angular frequency at different temperatures for $\text{Ni}_3(\text{PO}_4)_2$.

As the temperature rises, increased carrier mobility strengthens interfacial polarization, resulting in higher ϵ' . Unlike intrinsic dipolar contributions, this mechanism arises mainly from the microstructural arrangement of the material.

To further examine the stability of the dielectric response at high frequencies, the dielectric properties of the studied compound were analyzed over the temperature range of 393–633 K, as summarized in Table 4. At each temperature, the dielectric constant (ϵ') shows a slight decrease with increasing frequency; for example, at 393 K, ϵ' values are 75.2 (1 kHz), 71.6 (10 kHz), and 69.6 (100 kHz). Similarly, the dielectric loss ($\tan \delta$) diminishes with frequency, indicating that energy dissipation is somewhat higher at lower frequencies but remains low across the high-frequency range. At a fixed frequency, both ϵ' and $\tan \delta$ gradually increase with temperature for instance, at 10 kHz, ϵ' rises from 71.6 at 393 K to 76.7 at 633 K, while $\tan \delta$ increases from 0.0269 to 0.0695. This behavior reflects thermally activated dipolar motion, which enhances dielectric polarization as the temperature rises. Even at the highest measured temperature, dielectric losses remain moderate (e.g., $\tan \delta = 0.309$ at 633 K).

Table 4 Dielectric constant (ϵ') and loss tangent ($\tan \delta$) of the studied compound as a function of temperature at selected frequencies (1, 10, 100 kHz)

Frequency (kHz)	100		10		1				
	Temperature (K)	ϵ'	$\tan \delta$	Temperature (K)	ϵ'	$\tan \delta$	Temperature (K)	ϵ'	$\tan \delta$
393		69.6	0.0182	71.6	0.0269	75.2	0.0633		
413		69.8	0.0189	71.9	0.0268	75.5	0.0631		
433		70.1	0.0198	72.3	0.0291	75.9	0.0709		
453		70.4	0.0206	72.6	0.0303	76.3	0.0804		
473		70.7	0.0210	72.9	0.0327	76.6	0.100		
493		71	0.0219	73.4	0.0350	76.9	0.115		
513		71.3	0.0222	73.7	0.0353	77.6	0.126		
533		71.7	0.0237	73.8	0.0368	78.1	0.151		
553		72.1	0.0248	74.7	0.0476	80.4	0.182		
573		72.4	0.0258	75.2	0.0529	81.4	0.212		
593		72.8	0.0269	75.7	0.0588	82.5	0.246		
613		73.2	0.0281	76.3	0.0647	83.6	0.282		
633		73.5	0.0287	76.7	0.0695	84.4	0.309		

1 kHz and 633 K), indicating good stability of the dielectric response. Overall, the material demonstrates consistent high-frequency dielectric behavior with moderate energy dissipation, suggesting its potential suitability for high-frequency electronic or dielectric applications.

The dielectric loss factor ($\tan \delta$) characterizes the ratio between the energy dissipated (ϵ'') and the energy stored (ϵ') within the material.⁹⁵ A larger $\tan \delta$ value reflects stronger energy dissipation, typically associated with resistive contributions or relaxation processes caused by structural defects. Conversely, a smaller $\tan \delta$ indicates lower dielectric losses, enhancing the suitability of the material for energy storage applications. Analyzing the variation of $\tan \delta$ with frequency and temperature provides essential information about the balance between storage and dissipation mechanisms, thereby supporting the overall dielectric assessment. The dissipation factor can be expressed mathematically as:

$$\tan \delta = \frac{\epsilon''}{\epsilon'} \quad (18)$$

Fig. 13(b) illustrates that the loss tangent ($\tan \delta$) for $\text{Ni}_3(\text{PO}_4)_2$ ranges from about 25 at low frequencies to nearly 0.02 at high frequencies. The variation of $\tan \delta$ with frequency reveals a consistent decrease across all investigated temperatures (393–633 K), accompanied by a dispersion in the low-frequency region. According to Koop's theory, imperfections within the material introduce potential barriers that hinder the mobility of charge carriers.⁹⁶ This behavior is characteristic of thermally activated dielectric relaxation, where the loss peak arises from energy dissipation associated with the dynamic hopping of charge carriers.^{32,87} The dielectric loss increases progressively with temperature due to the higher thermal energy, which enhances both carrier mobility and the population of thermally excited charge carriers. At lower temperatures, the reduced carrier mobility suppresses conduction, leading to weaker dielectric losses. In contrast, at elevated temperatures, the combined effects of intensified hopping transport and interfacial polarization result in higher conductivity and a more pronounced dielectric loss response.^{97,98}



3.6. Analysis of electric modulus spectra and dielectric relaxation behavior

The complex electric modulus is a powerful tool for investigating the detailed electrical and dielectric characteristics of materials, providing deeper understanding of the underlying mechanisms. In phosphate-based systems, many studies have shown that modulus analysis effectively uncovers the key processes governing dielectric relaxation.^{2,4,35} This approach is especially useful for detecting secondary effects that may be difficult to resolve or that overlap in standard impedance spectroscopy due to differences in relaxation time constants.^{82,98} The complex modulus, M^* , can be expressed in terms of the complex dielectric permittivity (ϵ^*) as follows:⁹⁸

$$M^*(\omega) = \frac{1}{\epsilon^*(\omega)} = \frac{1}{\epsilon' - j\epsilon''} = \frac{\epsilon'}{\epsilon'^2 + \epsilon''^2} + j \frac{\epsilon''}{\epsilon'^2 + \epsilon''^2} = M' + jM'' \quad (19)$$

Fig. 14 displays the complex diagrams plots of the complex modulus over the investigated frequency and temperature ranges. Unlike Nyquist plots of complex impedance, which primarily emphasize the most resistive components of the material and thus provide limited sensitivity to the least capacitive contributions, the modulus representation proves particularly effective in separating electrode-related effects. In these plots, the frequency increases from left to right along the real axis (M'). All curves originate at zero, confirming the absence of electrode polarization effects.⁹¹ Across the studied temperature range, the semicircles observed on the right correspond to the high-frequency response of the grains, while those on the left are linked to the low-frequency response of grain boundaries. The appearance of these arcs clearly demonstrates that both grain and grain boundary capacitances play an active role in the conduction process of the material.

Fig. 15(a) shows the temperature- and frequency-dependent variation of the real part of the complex modulus (M') for the synthesized $\text{Ni}_3(\text{PO}_4)_2$ compound. At low frequencies, M' remains nearly zero, indicating negligible electrode polarization.

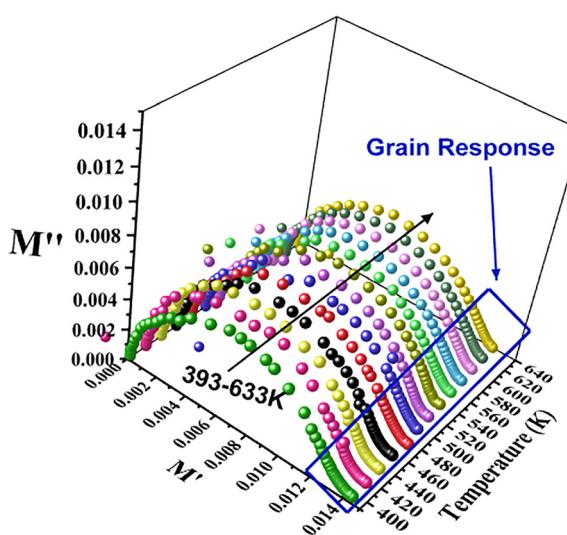


Fig. 14 Complex diagrams of electric modulus, at various temperatures.

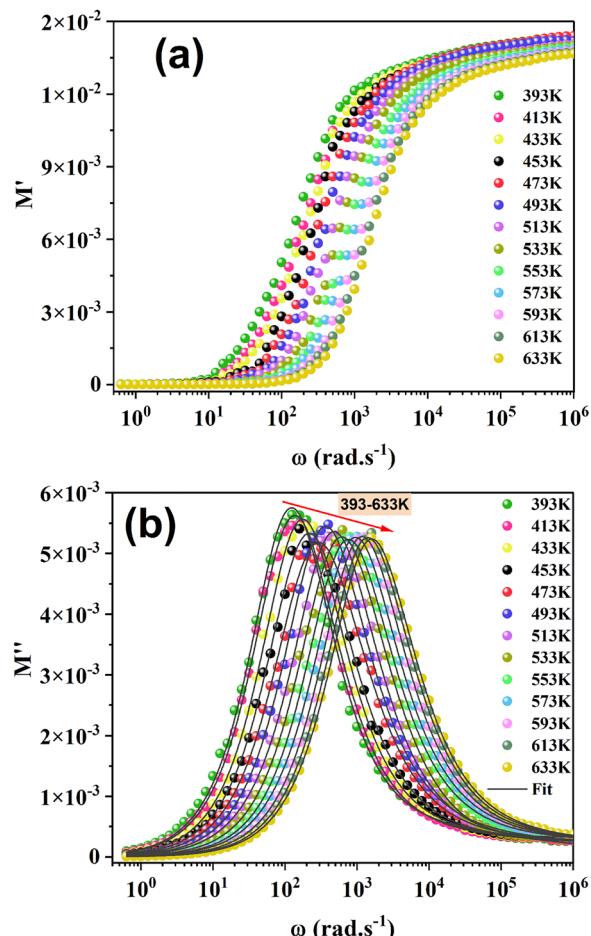


Fig. 15 (a) Evolutions of real part M' of complex electric modulus as a function of angular frequency. (b) Imaginary part (M'') of the electric modulus versus angular frequency at various temperatures.

As the frequency increases, a noticeable dispersion emerges, which is associated with the localized motion of charge carriers.^{91,98} At higher frequencies, M' reaches asymptotic values, reflecting the reduced influence of the external electric field on charge-carrier dynamics.

Fig. 15(b) shows the variation of the imaginary component of the electric modulus (M'') with frequency for $\text{Ni}_3(\text{PO}_4)_2$ ceramic across the studied temperature range. The spectra reveal a distinct relaxation peak (ω_p) at each temperature, which is mainly attributed to grain boundary effects. The occurrence of these peaks confirms the existence of a relaxation process linked to the transition from localized (short-range) to extended (long-range) charge carrier dynamics.⁹⁹ At frequencies lower than ω_p , electrical conduction is governed by long-range hopping of charges, whereas at higher frequencies, carriers are restricted to localized states.¹⁰⁰ The presence of short-range hopping contributions becomes evident above the relaxation frequency. With increasing temperature, the peaks gradually shift toward higher frequencies and display asymmetric broadening, indicating enhanced relaxation processes caused by thermally activated charge carriers. Such behavior highlights the non-Debye nature of relaxation in $\text{Ni}_3(\text{PO}_4)_2$.¹⁰¹



Table 5 Summarizes the parameters extracted from eqn (20) applied to the frequency-dependent modulus (M'') data

T (K)	M''_{\max}	ω_p	β
633	0.00526	1577.46	0.80568
613	0.00522	1467.86	0.80293
593	0.00523	1253.02	0.80410
573	0.00527	995.31	0.79981
553	0.00527	790.60	0.78362
533	0.00538	628	0.77562
513	0.00527	498.83	0.77161
493	0.00547	396.24	0.76025
473	0.00543	314.745	0.75853
453	0.00540	250.011	0.74804
433	0.00548	198.59	0.74204
413	0.00563	157.74	0.73236
393	0.00561	125.30	0.72626

To gain further insight into the dielectric relaxation mechanism from the frequency dependence of M'' , the modulus parameters (β , ω_p , M''_{\max}) was evaluated using the following expression:⁹⁸

$$M''(\omega) = \frac{M''_{\max}(\omega)}{\left[1 - \beta + \left(\frac{\beta}{\beta + 1}\right) \left[\beta \left(\frac{\omega_p}{\omega}\right) + \left(\frac{\omega}{\omega_p}\right)^{\beta} \right] \right]} \quad (20)$$

In this analysis, $M''(\omega)$ denotes the imaginary part of the electric modulus, while M''_{\max} represents its maximum value. The angular frequency is expressed as ω , and ω_p corresponds to the frequency at which $M''(\omega)$ attains its peak. The parameter β describes the degree of broadening in the modulus spectrum and quantifies the deviation from ideal Debye-type relaxation (with $0 < \beta \leq 1$). Physically, the exponent β provides key insight into the relaxation dynamics of the material: for an ideal dielectric, $\beta = 1$, reflecting the absence of significant dipole-dipole interactions, typical of Debye relaxation. When $\beta < 1$, however, the system exhibits notable dipole-dipole interactions, signifying a non-Debye relaxation process. Table 5 presents the temperature dependence of the parameters obtained from fitting the results of eqn (20).

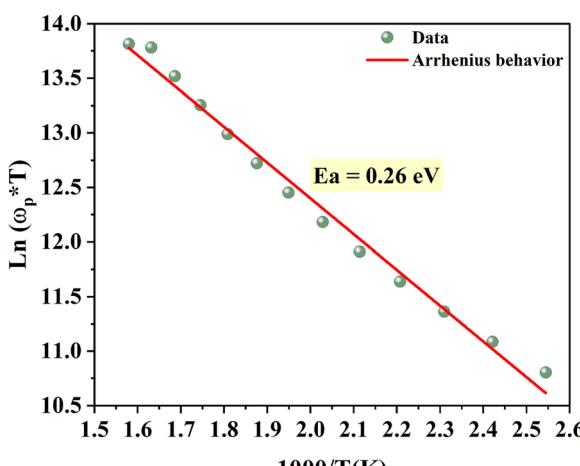


Fig. 16 Variation of $\ln(\omega_p \times T)$ as a function of $1000/T$ for the $\text{Ni}_3(\text{PO}_4)_2$ sample.

For $\text{Ni}_3(\text{PO}_4)_2$, the β parameter remains below 1 throughout the studied temperature range, confirming the non-Debye nature of the conduction mechanism (Table 5). Moreover, β increases gradually with rising temperature, which can be attributed to the thermal activation of charge carriers. As temperature increases, the mobility of these carriers improves and dipole-dipole interactions become less dominant, leading to a narrowing of the relaxation time distribution. This trend reflects a progressive shift toward more uniform relaxation dynamics, although still deviating from the ideal Debye behavior.

To determine the activation energies linked to electrical relaxation, the relaxation peak frequencies (ω_p) were extracted from the high-frequency peaks observed in the modulus spectra. Fig. 16 shows the logarithmic variation of $(\omega_p \times T)$ as a function of $1000/T$, confirming that the (ω_p) peaks obey the Arrhenius law. The activation energy was determined to be $E_a = 0.26$ eV. This value, in close agreement with that obtained from DC conductivity measurements, highlights a strong correlation between the relaxation dynamics and the charge transport mechanism in the $\text{Ni}_3(\text{PO}_4)_2$ system.

4. Conclusion

With the increasing interest in inorganic phosphates, this work explores the structural, optical, electrical, and dielectric properties of $\text{Ni}_3(\text{PO}_4)_2$ synthesized by a conventional solid-state reaction. The synthesis route employed is both cost-effective and scalable, enhancing the material's potential for industrial use. X-ray diffraction (XRD) confirmed the phase purity and revealed a monoclinic structure consistent with crystallographic calculations. Optical analysis revealed a wide band gap of 4.74 ± 0.01 eV, indicative of semiconducting behavior. Impedance spectroscopy demonstrated thermally activated electrical conduction with a non-Debye relaxation process, while DC conductivity followed the Arrhenius law with an activation energy of 0.28 eV, significantly lower than that of related Ni-based phosphates. The decrease of the frequency exponent s with temperature further confirms that charge transport is governed by the Correlated Barrier Hopping (CBH) mechanism. At low frequencies, the dielectric permittivity exhibits very high values ($\epsilon' \approx 1.58 \times 10^4$ at 633 K), which increase with temperature due to interfacial polarization at grain boundaries, consistent with the Maxwell-Wagner model. The material also maintains stable high-frequency dielectric behavior with moderate energy dissipation, highlighting its efficient polarization response. Overall, the combination of structural stability, low activation energy, and strong dielectric performance positions $\text{Ni}_3(\text{PO}_4)_2$ as a promising candidate for high-permittivity components, energy storage devices, and high-frequency electronic applications.

Conflicts of interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.



Data availability

All data relevant to this study are available from the corresponding author upon reasonable request.

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