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Synthesis and characterization of potential CeNiO₃ perovskite for photoelectrochemical water splitting

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Photoelectrochemical (PEC) water splitting offers a sustainable pathway for hydrogen production; however, its practical implementation is often limited by the poor efficiency and stability of photoelectrodes. In this work, porous cerium nickel oxide (CeNiO₃) was synthesized *via* a simple citrate sol–gel method coupled with a hydrothermal approach and employed as a photoanode for PEC water splitting. The structural, morphological, and optical characteristics of the material were comprehensively investigated using XRD, UV–vis spectroscopy, FESEM, EDX, XPS, PL, and FTIR analyses. The optimized CeNiO₃ photoelectrode demonstrated an excellent photocurrent density of 15.14 mA cm^{−2} at 1.4 V vs. RHE. Electrochemical impedance spectroscopy (EIS) revealed enhanced charge transfer kinetics and suppressed recombination of photoexcited charge carriers. The superior PEC activity of CeNiO₃ is attributed to its bimetallic interactions, strong solar light absorption, efficient charge separation, and rapid charge transport. These results highlight the potential of CeNiO₃ as a stable and efficient photoelectrode for solar-driven hydrogen generation.

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1. Introduction

The environmental concerns and interplay of energy have emerged as paramount challenges hindering the sustainable advancement of the global ecosystem. It is imperative to devise an energy reservoir that is storable, environmentally sustainable, cost-effective, and renewable to address the global energy requisites comprehensively.¹ Hydrogen (H₂) energy emerges as an exemplary primary energy vector for establishing a sustainable future, owing to its renewable nature, storage capability, net-zero emission characteristics, and notably high energy density of 143 KJ g^{−1}, surpassing that of gasoline.^{2,3} To date, hydrogen production methods such as steam reforming, methane pyrolysis, ammonia decomposition, and solar thermochemical processes have been explored.^{4,5} Despite their potential, these approaches are limited by rigorous reaction conditions, high energy consumption, and challenges in separating byproducts, especially in terms of H₂ storage and transportation.⁶ Therefore, extensive investigations have been directed towards the development of alternative hydrogen production methods that meet the increasing global hydrogen demand in an environmentally friendly and cost-effective

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manner. In this scenario, photo-electrochemical (PEC) water splitting, utilizing semiconductor materials, remains a subject of global interest due to its ability to directly decompose water into H₂ and O₂ through sunlight absorption. This approach is highly efficient and eco-friendly and holds promise for progress towards a zero-carbon future.^{7,8} The PEC water splitting offers a unique advantage over alternative methods due to its capability to produce H₂ and O₂ at distinct electrodes, mitigating issues related to gas mixing and undesirable back reactions. Moreover, the PEC system exhibits lower overpotentials when juxtaposed with electrochemical water splitting. The external voltage allows the mechanism to accelerate the reaction at the required rate by promoting charge separation and migration, resulting in high efficiency.⁹ In this context, most research efforts have centred on developing suitable semiconductor materials capable of serving as effective photoelectrodes for green hydrogen fuel generation. Efficient PEC water splitting requires that semiconductors have a minimal band gap of 1.23 eV to maximize light absorption, band edge alignment with water redox potentials, sustained stability, and support for effective charge carrier diffusion.^{10,11} Additionally, semiconductor-based photoanodes or photocathodes enable lower operating potentials, facilitating high energy conversion rates in the PEC water splitting, even at low temperatures.

To date, a multitude of photoelectrodes have been explored and optimized for utilization in PEC water splitting, encompassing materials such as TiO₂, Fe₂O₃, ZnO, g-C₃N₄, and WO₃.^{12,13} However, limitations such as rapid recombination, susceptibility to photo-corrosion, wide band gaps, low charge carrier mobility, and suboptimal band edge alignment restrict the applicability of these semiconductor materials as photoelectrodes.^{14,15} Due to their robustness, versatile composition and structure, strong visible light absorption, and adjustable band properties, perovskite oxides (ABO₃) present an attractive choice for PEC water splitting photoelectrodes.¹⁶ Among the various perovskite oxides, CeNiO₃ has emerged as a leading choice in photo-electro-catalysis, characterized by a narrow band gap of 1.51 eV. Its high electrocatalytic efficiency, earth-abundant nature, robust redox cycles, exceptional optical features, environmental sustainability, and the combined effects of photocatalytic and magnetic properties make it highly favourable for hydrogen evolution.¹⁷ Owing to the profusion of oxygen vacancies, CeO₂ metal oxide displays n-type semiconductor properties, with impressive redox reaction capabilities caused by its spatial processes and oxygen-releasing properties.^{18–20} Furthermore, the incorporation of ceria (Ce) and nickel (Ni) into the A and B sites of the perovskite framework enhances its electronic and optical characteristics and improves capacitance values while maintaining robust cycling stability. This enhancement is attributed to structural alterations within the ceria crystalline matrix, primarily involving partial replacement of Ce⁴⁺ ions with Ni²⁺. The substitution of Ni²⁺ for Ce⁴⁺ in the CeNiO₃ lattice induces structural modifications, which can influence defect formation, the electronic structure, and charge carrier dynamics. Specifically, the incorporation of Ni²⁺ can introduce oxygen vacancies and modify the

local coordination environment, enhancing charge separation and transport.^{21,22}

Research has consistently revealed that achieving a large surface area and controlling the morphology in catalysts are additional influential factors that highly influence the charge transfer mechanism and light absorption dynamics, contributing to the efficient production of hydrogen and fabrication of highly efficient photoelectrodes, achievable through suitable synthesis methodologies.²³ Various methodologies have been documented for the development of perovskite materials, encompassing techniques such as electrospinning, solvothermal synthesis, co-precipitation, hydrothermal, and sol-gel methods.^{24,25} Among these techniques, the sol-gel approach is the most beneficial for the manufacture of CeNiO₃ since it is environmentally benign and inexpensive.²⁶ Ming Yang *et al.* developed a TiO₂@SrTiO₃@BiVO₄ photoanode for hydrogen evolution, achieving an impressive photocurrent density of 3.7 mA cm⁻² at 1.23 V *vs.* RHE under AM 1.5 G illumination, along with a peak incident photon-to-current conversion efficiency (IPCE) of 65%.²⁷ Jin Kim and co-workers investigated PEC water splitting using Fe₃O₄ nanoparticles exsolved in SrTiO₃, achieving a photocurrent density of 5.10 mA cm⁻² at 1.23 V *vs.* RHE, along with excellent durability, maintaining 97% of its initial activity over 24 h.²⁸ In their study, Ramesh Reddy and colleagues employed 3D ZnO nanostructures for PEC water splitting, achieving a photocurrent density of 0.6 mA cm⁻² and showing stable performance for up to 5 h under light-on/light-off testing.²⁹ Xie and colleagues developed a CdS nanosphere photoelectrode for PEC water splitting, achieving a photocurrent density of 5.10 mA cm⁻² at 1.23 V *vs.* RHE.³⁰

According to the available literature, there is a significant gap in research on the implementation of a CeNiO₃ perovskite material as a photoelectrode in photoelectrochemical water splitting. This work focuses on cost-effective and stable materials, aligning with sustainable hydrogen production goals, and outlines the synthesis of a CeNiO₃ perovskite material utilising a simple, eco-friendly, and economical citrate sol-gel approach. These samples are then explored as photoelectrode materials for generating hydrogen *via* photoelectrochemical water splitting. The manufactured CeNiO₃ material was thoroughly characterised using cutting-edge equipment to discover the physicochemical attributes of the resulting material. Additionally, the photo-electro-catalytic performance of the CeNiO₃ material has been examined through PEC water-splitting tests. The CeNiO₃ material performed well in PEC water splitting, enabling successful generation of sustainable hydrogen fuel.

2. Experiment and methodology

2.1. Materials

Citric acid monohydrate (C₆H₈O₇·H₂O), cerium nitrate hexahydrate [Ce(NO₃)₃·6H₂O], nickel nitrate hexahydrate [Ni(NO₃)₂·6H₂O], ammonia water (NH₃·H₂O) and other solvents were sourced from Sigma-Aldrich Pvt. Ltd and chosen as precursors for the preparation of CeNiO₃ catalysts. Deionized water was



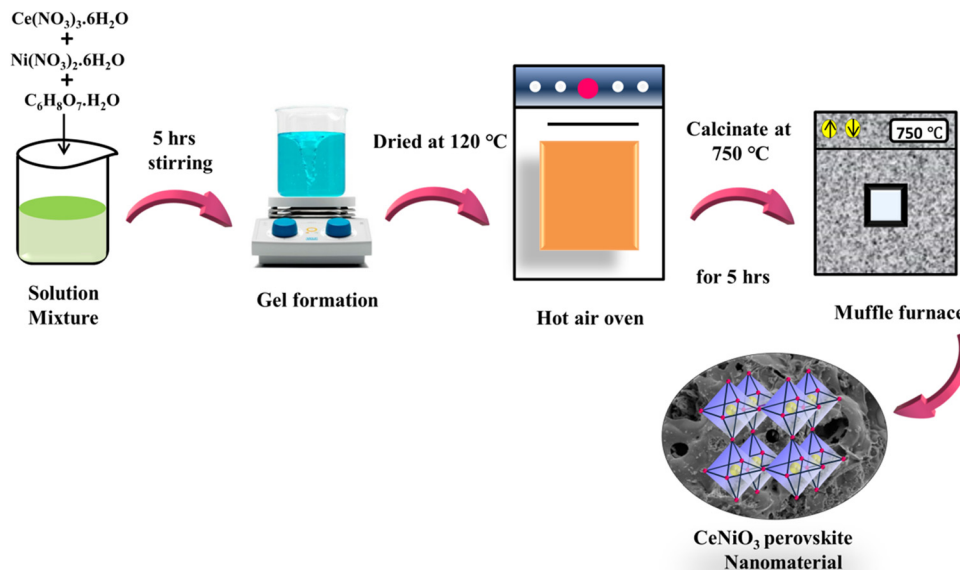


Fig. 1 Visual illustration of the CeNiO₃ perovskite material preparation.

employed as the dissolution solvent throughout the experimental process and all the reagents used in this study were of high analytical purity.

2.2. Synthesis of the CeNiO₃ catalyst

The CeNiO₃ catalytic material was prepared using the universally accepted citrate sol-gel technique and the synthesis techniques employed are depicted in Fig. 1. A stoichiometric volume of cerium nitrate hexahydrate, citric acid monohydrate and nickel nitrate hexahydrate (Ce:C₆H₈O₇·H₂O:Ni = 1:3:1) was dissolved in deionized Millipore water (50 mL) and the obtained mixture solution was continuously stirred for 30–35 minutes. The pH value was adjusted between 5 and 6 approximately by adding an ammonia solution sequentially. The above mixture was kept stirring maintaining at a temperature of 90 °C until a green gel formed. The as-obtained homogenized gel was initially dried at 120 °C for 7–10 hours in order to eliminate any remaining nitrates. The ground powder was subjected to calcination at 750 °C for 5 hours to synthesize the target perovskite phase, CeNiO₃.

2.3. Fabrication and preparation of the CeNiO₃ photoelectrode

A working CeNiO₃ photoelectrode was prepared by coating the fabricated photocatalyst as a thin film onto an FTO glass plate. A total of 5 mg of the specimen was thoroughly mixed into 500 mL of Nafion-ethanol during the first 30 minutes of ultrasonication in an effort to create a catalyst-loaded photoelectrode. Thereafter, 20 μL of the suspension was drop cast onto FTO glass measuring 1 × 1 cm² and dried at 150 °C for 3 hours in an oven and later subjected to photoelectrochemical measurement.

2.4. Characterization of CeNiO₃

The designed CeNiO₃ catalyst underwent a comprehensive array of advanced examinations, delving into its fundamental

physicochemical and innate characteristics. Powder X-ray diffraction (XRD) characterization (Shimadzu, Japan) was employed for crystallinity and lattice configuration studies. The photoluminescence (PL) emission was investigated utilizing an FL-1039/40 fluorescence spectrometer (Horiba Jobin Yvon), specifically to examine the charge separation kinetics of the specimen. FTIR was conducted for examination of the vibrational patterns about the stretching and bending modes of the synthesized material (Bruker, ALPHA, 200619, Germany). Field emission scanning electron microscopy (FE-SEM) was carried out using a JEOL-AV JSM-7100F (Singapore) to explore the surface morphology and microstructural features of the CeNiO₃ material. Energy-dispersive X-ray (EDX) assessment was carried out to discern the chemical constituents and elemental spatial distribution of the sample employing Super X (JSMIT300, JEOL, Singapore). The elemental composition of the material was elucidated through X-ray photoelectron spectroscopy (XPS) using a PHI 5000 VersaProbe III. The optical attributes of the specimen were assessed utilizing UV-visible absorption spectroscopy using a UV 2600 spectrophotometer (Shimadzu, Japan).

2.5. Photo-electro-chemical characterization

Photo-electro-chemical investigations were conducted utilizing a conventional three-electrode cell configuration (CHI760E) with CeNiO₃ photocatalysts serving as the working electrode, where an Ag/AgCl electrode was employed as the reference electrode and a platinum (Pt) rod functioned as the counter electrode. An aqueous solution containing 0.5 M Na₂SO₄ (sodium sulfate) served as the electrolytic medium, maintaining a pH of 5. The photoelectrodes underwent irradiation at an intensity of 100 mW cm⁻², employing a 350 W xenon lamp equipped with an AM 1.5G filter. Linear sweep voltammetric (LSV) polarization profiles were acquired within the voltage range of 0 to 1.4 V. The potentials were referenced to the



reversible hydrogen electrode (RHE) by applying the equation $E_{\text{RHE}} = E_{\text{Ag/AgCl}} + E^0_{\text{Ag/AgCl}} + 0.059 \cdot \text{pH}$. Mott–Schottky curves were generated under illumination conditions at a frequency of 1000 Hz. Electrochemical impedance spectroscopy (EIS) was explored across a frequency range of 1 to 100 kHz. The electrodes underwent durability examination *via* chronoamperometric (CA) studies.

3. Results and discussion

3.1. XRD analysis

XRD assessment was conducted on perovskite CeNiO_3 particles to discover their crystalline structure and phase purity. The outcomes are depicted in Fig. 2. XRD analysis of the pristine CeNiO_3 catalyst has elucidated the presence of an orthorhombic phase. The XRD pattern exhibited distinct d peaks at 2θ angles of 28.45° , 33.08° , 37.16° , 43.16° , 47.48° , 56.34° , 59.12° , 62.92° , 69.51° , 76.75° , and 79.20° , which are attributed to the lattice planes of (111), (002), (112), (221), (040), (321), (123), (331), (242), (412), and (161), respectively, indicating the crystallinity of the CeNiO_3 sample. These outcomes match well with the reference data (ID number: mp-776207) and previously stated research work.²¹ The peaks perceived at 33.08° , 47.48° , 59.12° , 69.51° , 76.75° , and 79.20° correlate to the (002), (040), (123), (242), (412) and (161) crystallographic planes of cerium dioxide (CeO_2), as indicated by JCPDS No: 81-0792. The XRD pattern of cubic NiO exhibits strong signals at 2θ angles of 37.16° , 43.16° , and 62.92° analogous to the crystallographic planes (112), (221), and (331) respectively, as designated by JCPDS No: 75-0197.³¹ The XRD pattern, showing no additional diffraction peaks, supports the excellent crystallinity and phase homogeneity of the synthesized material.

3.2. FE-SEM analysis

The surface morphologies of as-prepared CeNiO_3 perovskite materials are scrutinized through the utilization of FE-SEM.

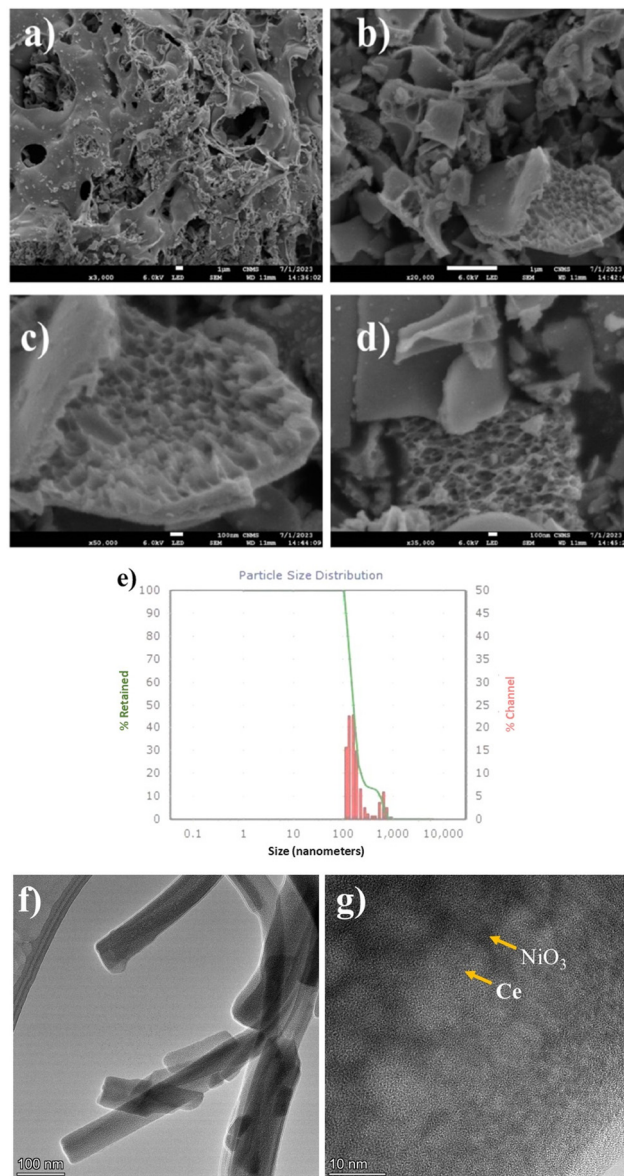


Fig. 3 (a)–(d) FE-SEM micrographs, (e) DLS results of the CeNiO_3 perovskite photocatalyst and (f) and (g) TEM images.

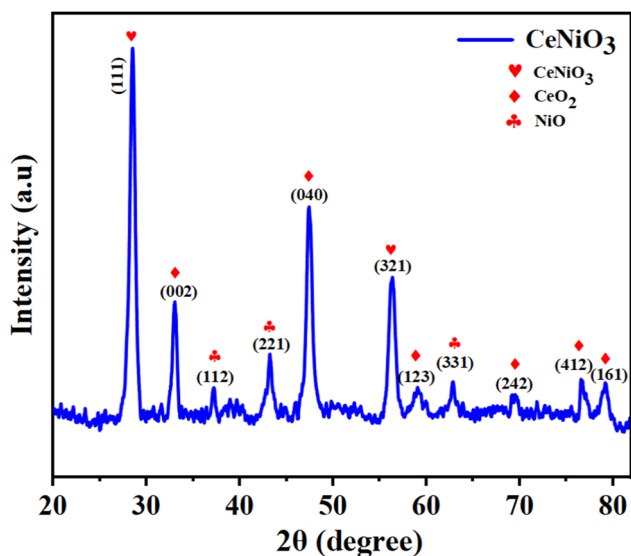


Fig. 2 XRD spectrum of the CeNiO_3 perovskite photocatalyst.

Fig. 3 shows a surface morphology micrograph depicting the synthesized CeNiO_3 material. These FESEM photographs indicate the formation of the CeNiO_3 material with rough surface morphology. FESEM pictures of CeNiO_3 unveiled a myriad of irregularly shaped pores exhibiting varying diameters and the material showcased a notably elevated surface area, highlighting its potential for enhanced catalytic activity. The porous architecture inherent in the CeNiO_3 material is widely recognized for its capacity to improve the photo adsorption performance, specifically in the segregation of electron-hole pairs.^{32,33}

3.3 TEM analysis

The rod-like shape of the synthesised CeNiO_3 nanostructures with uniform dimensions and smooth surfaces is clearly visible



in the TEM image (Fig. 3(f)), indicating that a well-defined crystalline phase was successfully formed. Effective oversight throughout the synthesis process is suggested by the nanorods' apparent well-dispersed state and lack of noticeable aggregation. The high crystallinity and purity of the prepared material are confirmed by the uniform contrast and distinct edges that were observed.

Additionally, the lattice structure and elemental dispersion within the CeNiO_3 matrix can be obtained from the image (Fig. 3(g)). The successful integration of Ni into the Ce lattice structure is confirmed by the distinct lattice fringes that match the interplanar spacings of the Ce and NiO_3 phases. Good crystallinity and close contact between the Ce and NiO_3 domains are indicated by the presence of clearly resolved fringes, which can promote effective charge transfer across the interface. These structural characteristics are anticipated to foster synergistic interactions between Ce and Ni species, improving the CeNiO_3 nanocomposite's catalytic and electrochemical performance.³⁴

3.4. EDX analysis

EDX analysis was conducted to scrutinize the degree of purity, elemental composition and distribution within the 100–300 nm range of the fabricated CeNiO_3 photocatalyst. Fig. 4(a) depicts the obtained EDX outcomes. The EDX assessment verifies the existence of ceria (Ce), nickel (Ni), and oxygen (O) within the sample, providing substantial evidence for the successful formation of the CeNiO_3 photocatalyst. The exceptional purity of the CeNiO_3 structures synthesized *via* sol-gel methodologies was substantiated by the absence of discernible intense peaks corresponding to any additional constituents or impurities within the materials. DLS results of the CeNiO_3 perovskite photocatalyst indicated that the particle size distribution varied in the 100–220 nm range (Fig. 3(e)). Additionally, the elemental mapping evaluation depicted in Fig. 4(b)–(e) revealed a spectrum comprising three different components, delineating the existence of Ce (red), Ni (blue), and O (yellow).

3.5. XPS analysis

XPS has been applied to evaluate the chemical states and elemental constituents of the CeNiO_3 perovskite photocatalyst. Fig. 5(a) confirms the homogeneity of the CeNiO_3 material, as evidenced by the observation of Ce 3d, Ni 2p, and O 1s signals, with no additional elemental constituents detected during the FE-SEM and EDX analyses. The determination of the binding energy of oxygen atoms provides insights into the oxygen's chemical state and enables the identification of oxygen vacancies that are induced within the material's crystalline lattice. The metal cations located at either the A-site or the B-site in the perovskite structure demonstrate ratios of A^{n+1}/A^n and B^{n+1}/B^n , respectively, showcasing their catalytic redox potential.³⁵ The Ce 3d spectrum prominently manifests distinct peaks corresponding to Ce^{3+} 3d_{3/2} and Ce^{4+} 3d_{5/2} electron orbitals, along with the observable satellite signals (Fig. 5(b)). Binding energies for Ce^{3+} 3d_{3/2} are well recorded at 916.98 eV, 905.39 eV, and 900.15 eV. Ce^{4+} 3d_{5/2} demonstrates distinctive binding energies

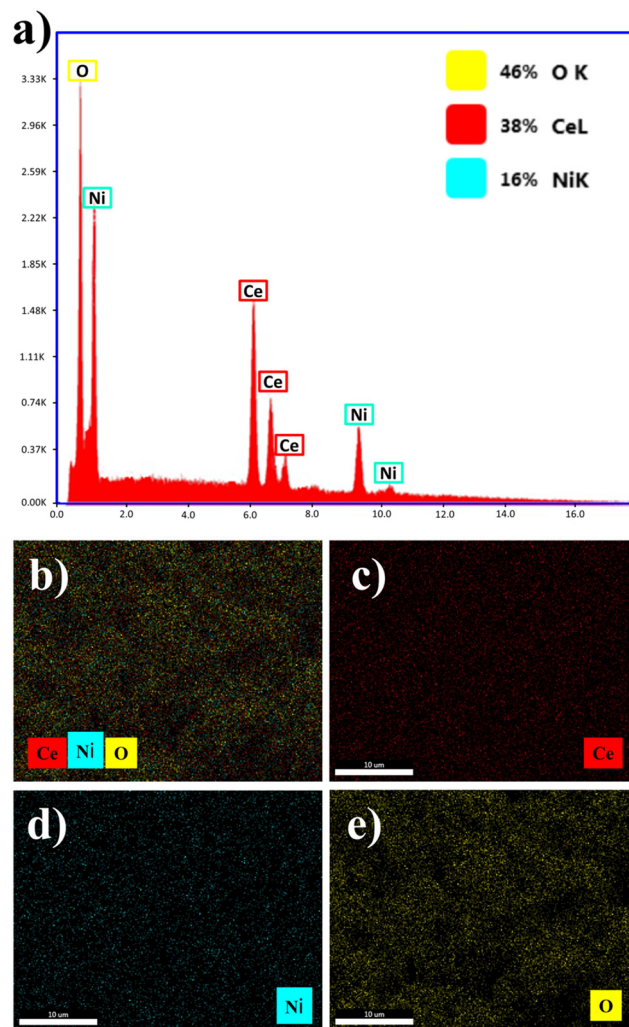


Fig. 4 (a). EDX profiles of the CeNiO_3 specimen and (b)–(e) elemental mapping analysis of the CeNiO_3 perovskite photocatalyst.

at 897.53 eV, 888.48 eV, and 882.77 eV, respectively.³⁶ Fig. 5(c) illustrates the Ni 2p XPS patterns, revealing spin-orbit doublets of Ni 2p_{3/2} and Ni 2p_{1/2} at 855.19 eV and 872.98 eV respectively. Additionally, there are two observable satellite spikes positioned at 861.12 eV and 883.53 eV. The dual peaks observed in the Ni 2p spectrum, positioned at 855.19 eV and 861.12 eV, are attributed to the Ni^{3+} oxidation state. Concurrently, the spikes positioned at 883.53 eV and 872.98 eV are assigned to the Ni^{2+} states.^{37,38} The binding energies observed in the O 1s XPS spectra, specifically the oxygen vacancies (O_v) at 531.33 eV and lattice oxygen (O_L) at 528.91 eV, as demonstrated in Fig. 5(d), serve as conclusive evidence of the existence of O_2^- species within the crystalline framework of the CeNiO_3 material.³⁹ Higher levels of oxygen vacancies facilitate rapid ionic diffusion during the charge transfer mechanism. The XPS outcomes illustrate that the CeNiO_3 material exhibits advantageous characteristics associated with two redox pairs ($\text{Ce}^{3+}/\text{Ce}^{4+}$ and $\text{Ni}^{2+}/\text{Ni}^{3+}$), along with the presence of oxygen vacancies within its structural framework. These attributes contribute to an enhanced capacity for charge storage at a rapid rate.



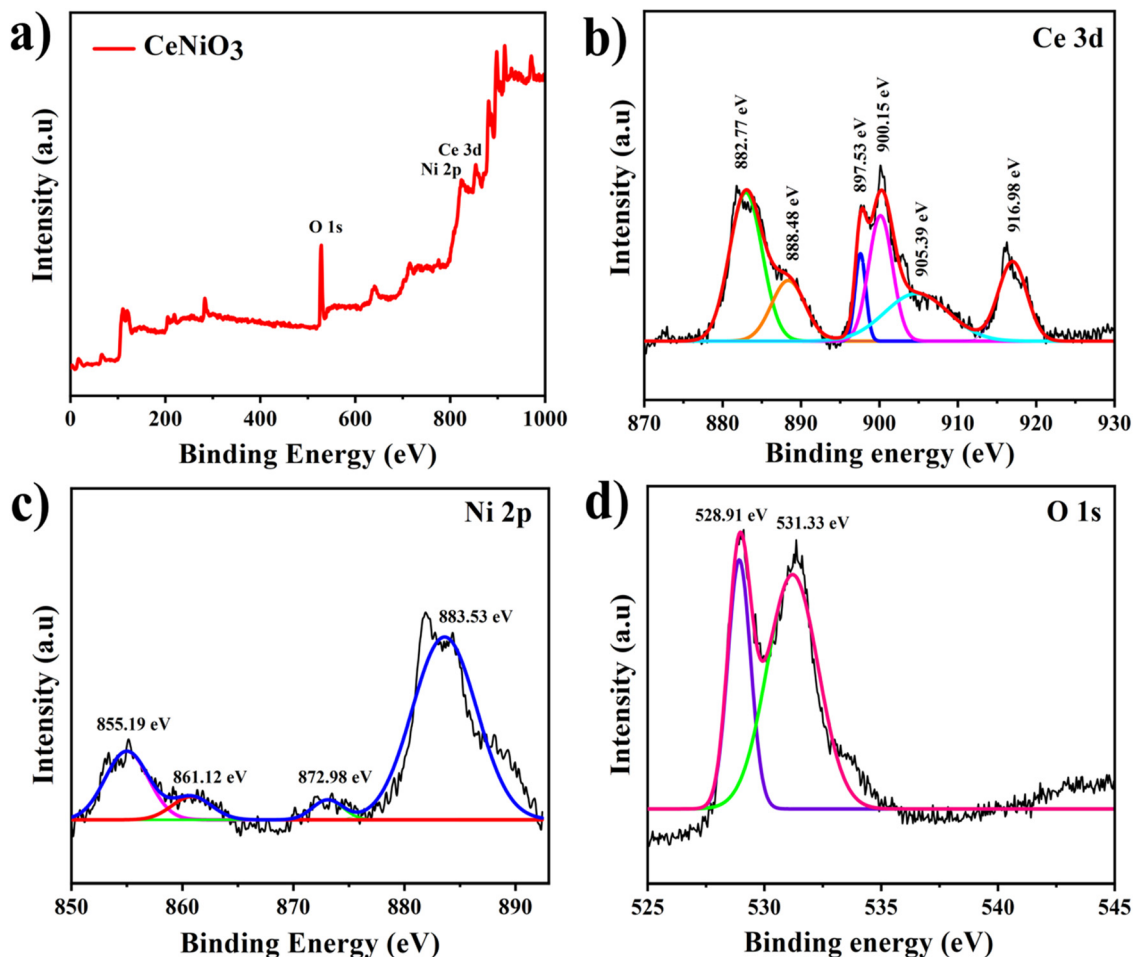


Fig. 5 (a) XPS full scan spectra of the CeNiO₃ material, (b) Ce 3d, (c) Ni 2p, and (d) O 1s.

3.6. UV-vis spectra

Band-gap energy and the light-scattering characteristics of the manufactured specimen were evaluated using UV-visible spectroscopy. Fig. 6(a) and (b) illustrate the UV-visible

spectroscopic data, as well as the corresponding Tauc plot for the given experimental conditions. The CeNiO₃ sample shows a red shift in its absorption onset, broadening into the visible light region, which enhances visible light absorption.⁴⁰ The

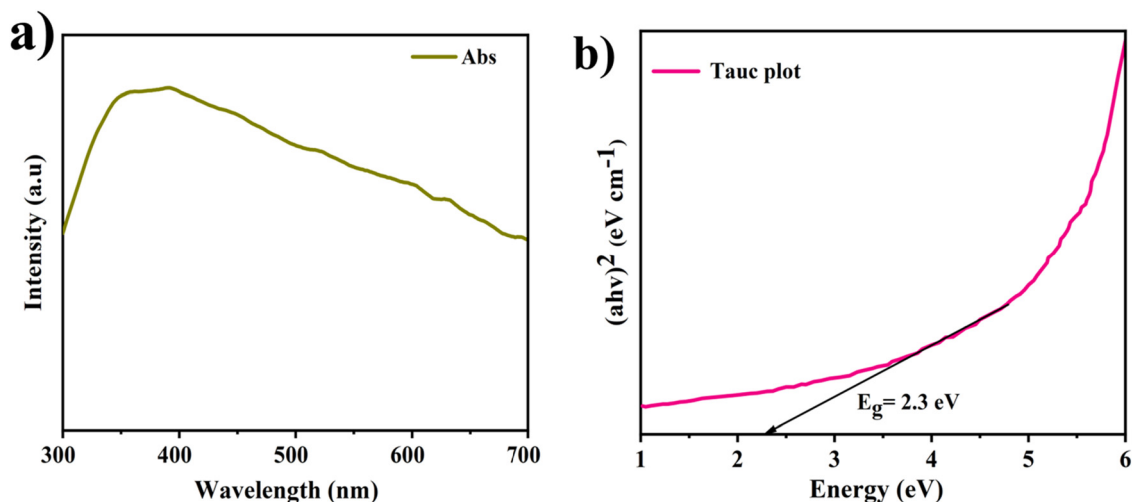


Fig. 6 (a) UV-visible absorbance spectra of the CeNiO₃ perovskite photocatalyst. (b) Calculated Tauc plot graph of the CeNiO₃ perovskite photocatalyst.



band-gap energy (E_g) of the fabricated CeNiO₃ specimen was determined employing Tauc's equation (eqn (1)), and the outcomes are presented in Fig. 6(b).

$$(xh\nu)^2 = A(h\nu - E_g) \quad (1)$$

The CeNiO₃ catalyst exhibited $E_g = 2.3$ eV, which implies that the material possesses advantageous optical characteristics, rendering it well-suited for engaging in photo-electro-catalytic applications using visible light.

Theoretical computation of the semiconductor's VB and conduction band CB potentials can be achieved by applying the Mulliken electronegativity and the semiconductor's band-gap through the subsequent formula.⁴¹

$$E_{VB} = X - E_e + 0.5E_g \quad (2)$$

$$E_{CB} = E_{VB} - E_g \quad (3)$$

where E_{VB} and E_{CB} symbolize the VB and CB potentials, respectively, E_g represents the band gap energy (2.3 eV), and E_e signifies the energy level of free electrons calibrated on the hydrogen scale, maintaining a constant value of 4.5 eV. Furthermore, X stands as the absolute electronegativity value of CeNiO₃.⁴² The detailed calculation of the conduction band and valence band of the CeNiO₃ material is as follows:

The first ionization energy of the cerium element (Ce):

$$I_1 = 534.4 \text{ kJ mol}^{-1}$$

The first electron affinity of the cerium element (Ce):

$$E_1 = 50 \text{ kJ mol}^{-1}$$

The absolute electronegativity of the cerium element (Ce):

$$X = \frac{1}{2}(I_1 + E_1) = \frac{1}{2}(534.4 + 50) = 292.2 \text{ kJ mol}^{-1}$$

Converting into eV,

$$\frac{292.2}{96.48} = 3.02 \text{ eV}$$

For nickel (Ni): $I_1 = 737.1 \text{ kJ mol}^{-1}$ and $E_1 = 111.65 \text{ kJ mol}^{-1}$; therefore,

$$X = \frac{1}{2}(I_1 + E_1) = \frac{1}{2}(737.1 + 111.65) = 424.37 \text{ kJ mol}^{-1}$$

Converting into eV,

$$\frac{424.37}{96.48} = 4.39 \text{ eV}$$

For oxygen (O): $I_1 = 1313.9 \text{ kJ mol}^{-1}$ and $E_1 = 140.97 \text{ kJ mol}^{-1}$; therefore,

$$X = \frac{1}{2}(I_1 + E_1) = \frac{1}{2}(1313.9 + 140.97) = 727.43 \text{ kJ mol}^{-1}$$

Converting into eV,

$$\frac{727.43}{96.48} = 7.53 \text{ eV}$$

The geometric mean of the absolute electronegativity for CeNiO₃ is calculated as follows:

$$X = \sqrt[3]{3.02 \times 4.39 \times 7.53^3} = 5.63 \text{ eV}$$

The band gap of CeNiO₃ was obtained by the UV-vis measurements and determined to be 2.3 eV (Fig. 6b).

Therefore, the valence band of CeNiO₃ is calculated as follows:

$$\begin{aligned} E_{VB} &= X - E_e + 0.5E_g \\ &= 5.63 - 4.5 + 0.5 \times 2.3 \\ &= 2.28 \text{ eV} \end{aligned}$$

The conduction band of CeNiO₃ is calculated as follows:

$$\begin{aligned} E_{CB} &= E_{VB} - E_g \\ &= 2.28 - 2.3 \\ &= -0.02 \text{ eV} \end{aligned}$$

The computed E_{VB} value for CeNiO₃ is 2.28 eV, while its corresponding E_{CB} value is -0.02 eV.

3.7. FTIR spectra

The functional groups and discernible chemical bonds of the synthesised CeNiO₃ catalyst specimens were assessed using FTIR, and the findings are illustrated in Fig. 7(a). The fabricated specimen displayed stretching vibrations originating from crystallized water ($-OH$) molecules, potentially contributing to the broad transmission detected at 3440.01 cm^{-1} and 1773.46 cm^{-1} . The Ni^{2+}/Ni^{3+} cations, resembling the asymmetric sharp bands, were detected in the CeNiO₃ sample within the spectral range of 1374.37 cm^{-1} to 1633.85 cm^{-1} . The transmittance bands detected at 707.12 cm^{-1} are credited to the stretching of Ni-O bonds. Two discrete peaks identified at 1036.01 and 505.20 cm^{-1} are correlated with Ce-O stretching vibrations.²¹

3.8. PL analysis

To comprehend the segregation of the charges generated by light exposure during the PEC process, photoluminescence studies were performed at an excitation wavelength of 365 nm at ambient temperature. Fig. 7(b) displays the emission spectra of the CeNiO₃ specimen. Generally, a decrease in PL emission intensity indicates high charge carrier separation efficacy, ultimately optimizing the photo-electro-catalytic performance of photocatalysts. The PL spectra of the CeNiO₃ material disclosed dual peaks, with a lower ultraviolet-excitonic band at 412 nm and a higher green emission band at 468 nm. The UV-excitonic region is designated for charge



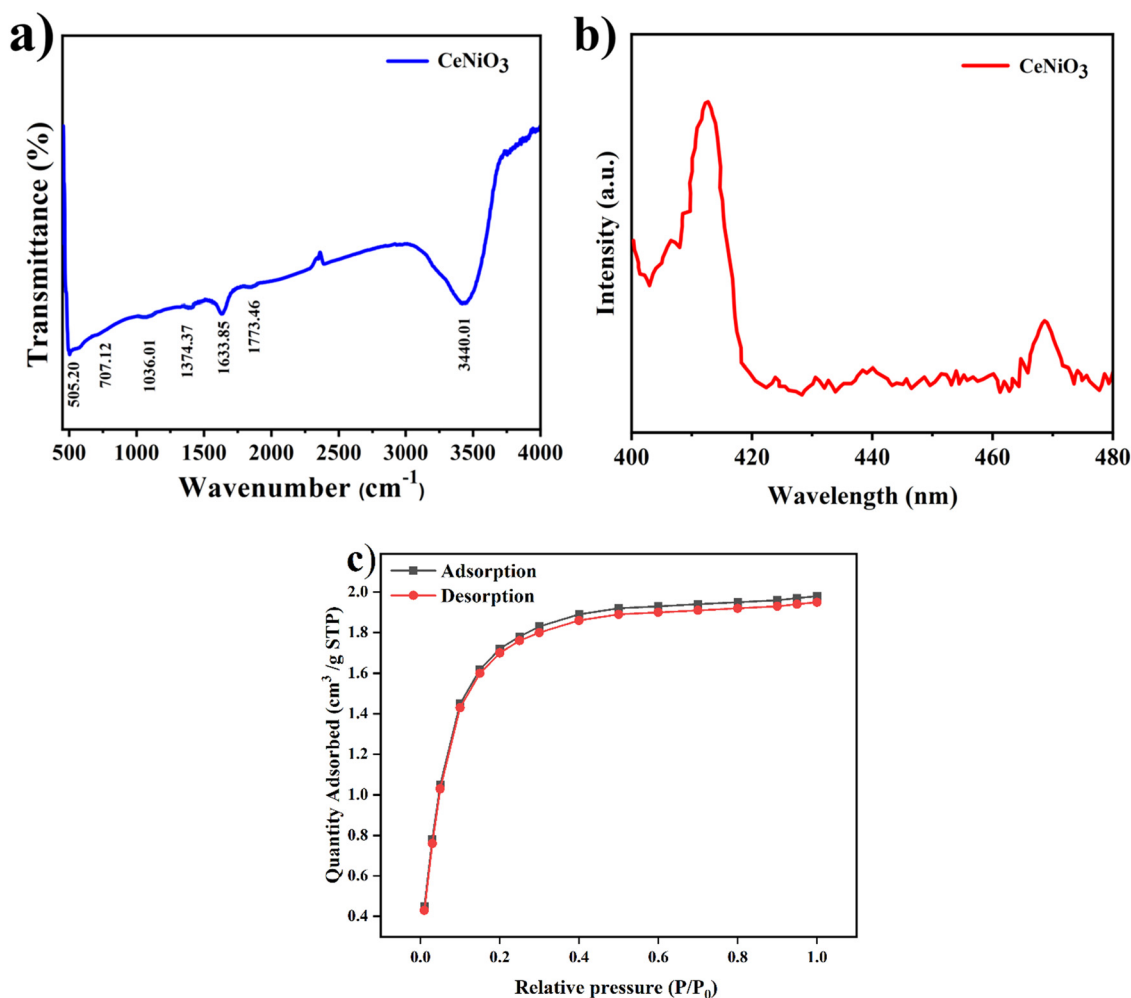


Fig. 7 (a) FTIR spectra, (b) PL spectra of the CeNiO₃ perovskite photocatalyst and (c) BET isotherm.

carrier recombination, while the green emission region is attributed to the occurrence of oxygen vacancies.⁴³ The PL spectra of CeNiO₃ exhibited a considerable decline in intensity, suggesting enhanced effectiveness in transferring and separating electron and hole pairs.⁴⁴

3.9. BET isotherm

According to IUPAC classification, the nitrogen adsorption-desorption isotherm of CeNiO₃ (Fig. 7(c)) displays a Type I(a) profile, signifying a microporous structure. Adsorption inside micropores is confirmed by a plateau after a sharp uptake at a low relative pressure ($P/P_0 < 0.1$). With a pore volume of $0.00414 \text{ cm}^3 \text{ g}^{-1}$ and an average pore diameter of 0.9 nm, the BET surface area was found to be $18.4 \text{ m}^2 \text{ g}^{-1}$. These values correspond to the development of a dense perovskite framework and tiny micropores. The compact configuration of metal-oxygen octahedra results in the low surface area and pore volume that are characteristic of perovskite oxides. Strong reactant interactions and effective charge transfer in catalytic or electrochemical processes can be facilitated by the confined micropores, even due to their limited porosity.

3.10. Photo-electro-chemical measurements

To interrogate the photo-electro-chemical performance of the CeNiO₃ perovskite photocatalyst, a three-electrode configuration was employed within a 0.5 M Na₂SO₄ electrolytic medium. Linear sweep voltammogram patterns of the CeNiO₃ photoelectrode are illustrated in Fig. 8(a). Under dark conditions, the current density remains minimal. However, upon exposure to solar radiation, a significant augmentation in current density is observed. The elevated photocurrent density signifies the effective sensitivity of the photoelectrode materials to the solar spectrum. This pronounced generation of photocurrent signifies that the semiconductor material possesses a high degree of photoactivity.⁴ The resulting photocurrent densities recorded for CeNiO₃ were established to be 15.14 mA cm^{-2} and 5.1 mA cm^{-2} with the influence of light and in dark mode at the specific potential of 1.4 V vs. the RHE. The exceptional functionality of CeNiO₃ stems from its uniform topological characteristics, porous nature, high surface area, and electronic structure, substantially enhancing the higher light absorption with efficient charge transformation within the semiconductor material.⁴⁵



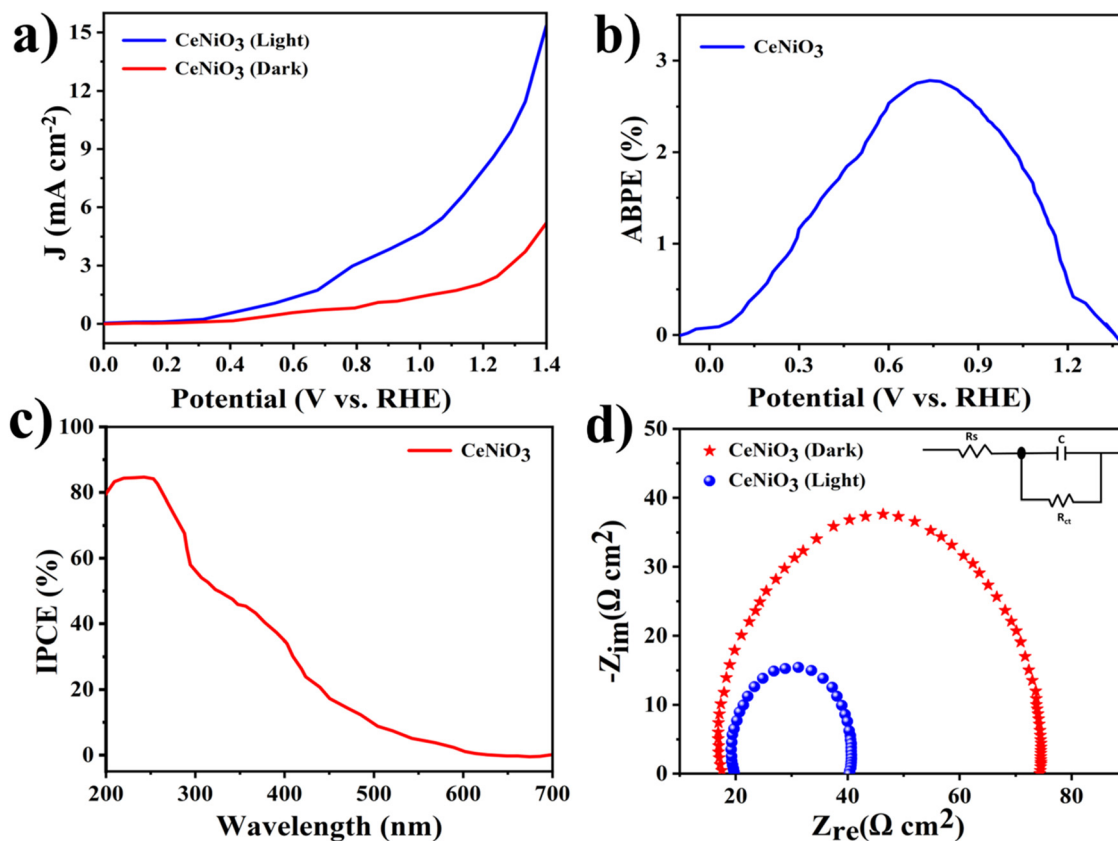


Fig. 8 (a) LSV curves, (b) (ABPE%) plot, (c) (IPCE%) graph, and (d) EIS Nyquist plot of the CeNiO₃ electrode (the inset shows an equivalent circuit model).

Furthermore, achieving elevated photocurrent density under minimal applied bias poses a pivotal obstacle in improving the efficacy of water-splitting approaches. Solar-to-hydrogen (STH) conversion efficiencies (η) of the photoelectrode under the influence of an external bias potential against the RHE have been assessed using the following equation:

$$\eta = \frac{J \times (1.23 \times V_b)}{P_{\text{light}}} \times 100\% \quad (4)$$

In this equation, J stands for the photocurrent density, and 1.23 V corresponds to the equilibrium potential that is required for splitting water into O₂ and H₂. V_b is the applied bias with respect to P_{light} , which is the power density of incident light, and RHE. Fig. 8(b) presents a plot of photoconversion efficiency (η %). The highest efficiency achieved for CeNiO₃ was 2.76%, reflecting its impressive ability to convert incident light into hydrogen fuel.

Additionally, the assessment of incident photon to current conversion efficiency (IPCE) is undertaken to precisely estimate the photoelectric conversion efficacy of the photoelectrode across a spectrum of incident light wavelengths, utilizing the subsequent formula.

$$\text{IPCE (\%)} = \frac{J \times 1240}{\lambda \times P_{\text{light}}} \times 100\% \quad (5)$$

In this study, the photocurrent density is expressed by the variable J , the value of 1240 is a product of the speed of light, P_{light} depicts the luminous power density, and λ indicates the incident wavelength. As presented in Fig. 8(c), a high IPCE of 80.10% is demonstrated by the CeNiO₃ photoelectrode over the 300–700 nm light spectrum. CeNiO₃ demonstrates heightened photoresponsivity, potentially attributed to the provision of additional reactive sites and efficient pathways for charge carrier transmission, facilitating light capture and electron excitation.⁴⁶

EIS is a method for elucidating the dynamics of charge carrier migration and assessing interfacial charge transfer resistance (R_{ct}). Fig. 8(d) showcases the intriguing findings from the EIS analysis conducted on the CeNiO₃ photoelectrode under both dark and illuminated conditions. The simulated circuit model representing the charge transfer mechanisms within photocatalysts is depicted in the inset of Fig. 8(d). The radius of the Nyquist arc collected in EIS is intricately linked to the kinetics of the photo-electro-catalytic reaction occurring on the surface of the photoelectrode. It quantifies the impedance to charge transfer at the photoelectrode–electrolyte interface. A smaller arc diameter signals higher electron transportation, which supports better segregation of light-induced charge carriers, thereby enhancing the proficiency of PEC water splitting.⁴⁷ The R_{ct} value for the CeNiO₃ photoelectrode is listed in Table 1. EIS spectra (Fig. 8(d)) of CeNiO₃ unveil that, in



Table 1 EIS fitting results for the CeNiO₃ photoelectrode

Sample	Condition	(R_{ct}) ($\Omega \text{ cm}^2$)
CeNiO ₃ photoelectrode	Dark	56.94
CeNiO ₃ photoelectrode	Light	20.13

comparison with the dark mode, a smaller semicircle is noticed under irradiation, signifying faster charge carrier mobility due to its highly porous nature and dual-metal synergy observed in the CeNiO₃ photoelectrode in the Na₂SO₄ electrolyte. Notably, there is remarkable inhibition of the charge carrier recombination.^{48,49}

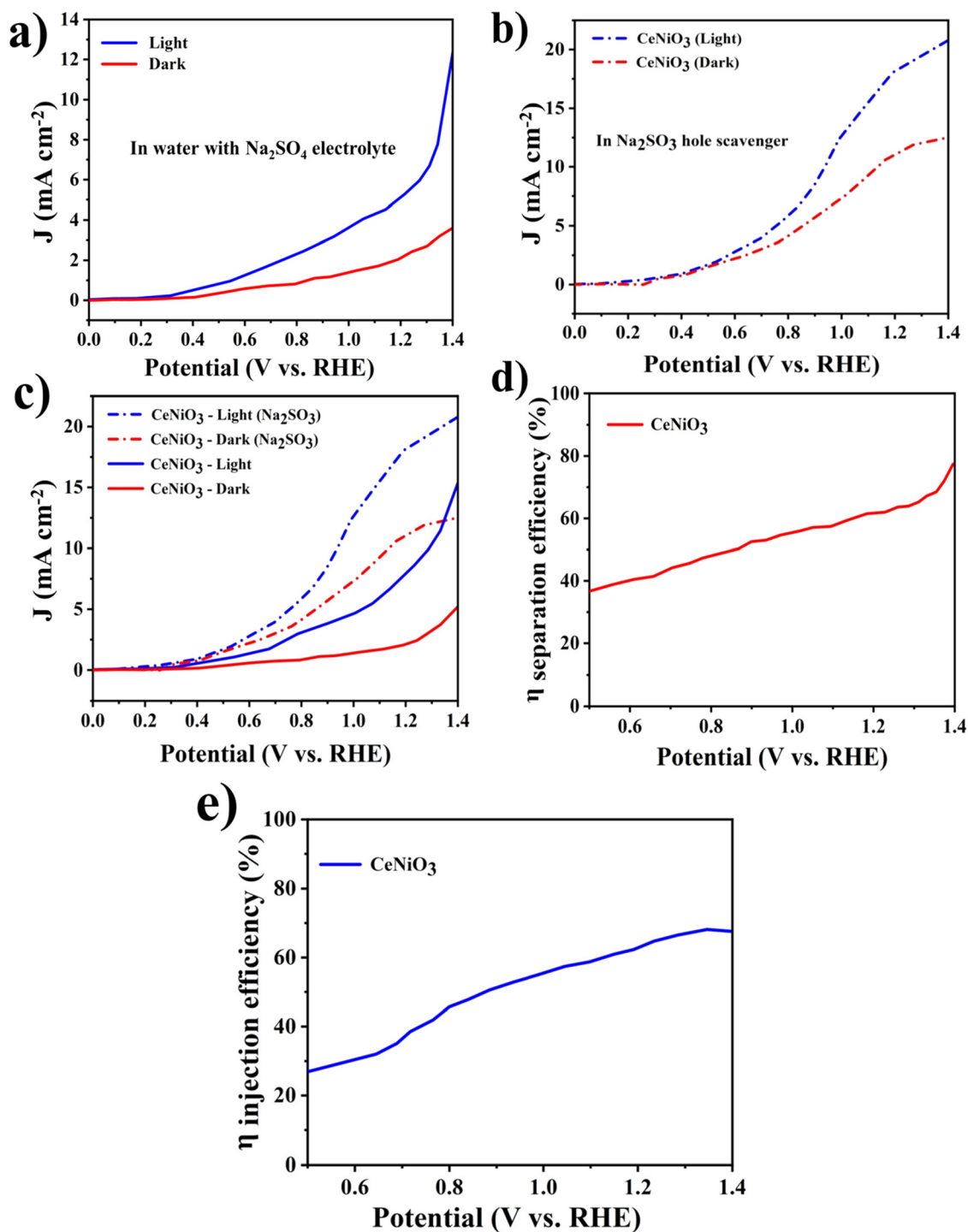


Fig. 9 (a) LSV curves with 0.5 M Na₂SO₄ electrolyte in water, (b) and (c) with addition of 0.5 M Na₂SO₃ solution for sulfite oxidation as a hole scavenger, (d) charge separation efficiency of the CeNiO₃ photoelectrode, and (e) charge injection efficiency of the CeNiO₃ photoelectrode.



To provide insight into the catalytic process, comprehensive evaluation of carrier kinetics is systematically performed. The charge separation efficiency ($\eta_{\text{separation}}$) and charge injection efficiency ($\eta_{\text{injection}}$) of the CeNiO₃ photoelectrode were assessed employing 0.5 M Na₂SO₃ as a hole scavenger with Na₂SO₄ electrolyte, as depicted in Fig. 9(a)–(c), applying the following equation:

$$\eta_{\text{separation}} = \frac{J_{\text{H}_2\text{O}}}{J_{\text{Na}_2\text{SO}_3}} \quad (6)$$

$$\eta_{\text{injection}} = \frac{J_{\text{Na}_2\text{SO}_3}}{J_{\text{abs}}} \quad (7)$$

where $J_{\text{H}_2\text{O}}$ and $J_{\text{Na}_2\text{SO}_3}$ stand for the photocurrent densities acquired for water oxidation and sulfite oxidation, and J_{abs} denotes the photo-current density achieved at 100% internal quantum efficiency, respectively. $\eta_{\text{separation}}$ quantifies the efficacy of separating photo-induced charge carriers to yield holes that reach the surface. $\eta_{\text{injection}}$ signifies the fraction of holes efficiently delivered into the water for oxidation. As shown in Fig. 9(d) and (e), the CeNiO₃ photoelectrode delivers the maximum $\eta_{\text{separation}}$ efficiency and $\eta_{\text{injection}}$ efficiency of 77% and 67% at 1.4 V vs. RHE. The favourable band structure, carrier transport properties, surface attributes, electronic characteristics and catalytic performance of the CeNiO₃ photoelectrode synergistically contribute to its remarkable efficiency in charge separation and injection processes.^{50,51}

Mott–Schottky (M–S) analysis was employed to derive the flat band potential (V_{fb}) of the CeNiO₃ material, aiding in the elucidation of its charge transfer dynamics. The following Mott–Schottky equation is exploited to derive the flat-band potential (V_{fb}).

$$\frac{1}{C^2} = \frac{2}{e\epsilon\epsilon_0 N_d} \left[(V - V_{\text{fb}}) - \frac{kT}{e} \right] \quad (8)$$

where C denotes the capacitance due to space charge, e represents the elementary charge, ϵ stands for the dielectric

constant, k and T represent the Boltzmann constant and temperature, N_d indicates the charge carrier density, V_{fb} signifies the flat band potential, and ϵ_0 symbolizes the permittivity of free space. Fig. 10(a) shows the M–S graph. The V_{fb} can be extrapolated using the X-intercept from the M–S plot. The verified V_{fb} value for the CeNiO₃ sample is recorded as 0.31 V. The reduced V_{fb} observed for CeNiO₃ suggests a diminished onset potential when exposed to light. The upward shift in the CeNiO₃ electrode signifies a lowering in band bending, which promotes the transfer of electrons.⁵²

In order to ensure scalable and reliable hydrogen production, the enduring stability of the photoanode is crucial. Long-term stability of the CeNiO₃ photoelectrode's photocurrent was investigated through chronoamperometric analysis. During these experiments, the photoanodes underwent continuous exposure to irradiation as presented in Fig. 10(b). The CeNiO₃ photoelectrodes demonstrated sustained and durable J (mA cm⁻²) of 11.17 mA cm⁻² at 0.8 V vs. RHE, with minimal decline in density of current observed over an extended period. Even after 10 hours of testing, the photocurrent density remains stable, demonstrating the photoelectrode's remarkable resilience against photo-induced corrosion.⁵³ These findings underscore the photocatalyst's durability under light exposure, validating its potential for efficient use in a photoelectrochemical water-splitting system. The PEC water-splitting performance of the CeNiO₃ photoelectrode was compared with that of previously reported materials, as outlined in Table 2.

The stability of the catalyst during the photoreaction was further evaluated by analyzing the XRD pattern of the used CeNiO₃ sample (Fig. 11(a)). No significant changes were observed in the diffraction peaks, confirming that the crystal structure of CeNiO₃ remained intact after the photoelectrocatalytic reaction. This indicates the excellent structural stability and reusability of the synthesized catalyst for practical hydrogen generation and degradation applications. The post-reaction SEM image of the CeNiO₃ sample is displayed in Fig. 11(b). After the hydrogen evolution process, the morphology almost

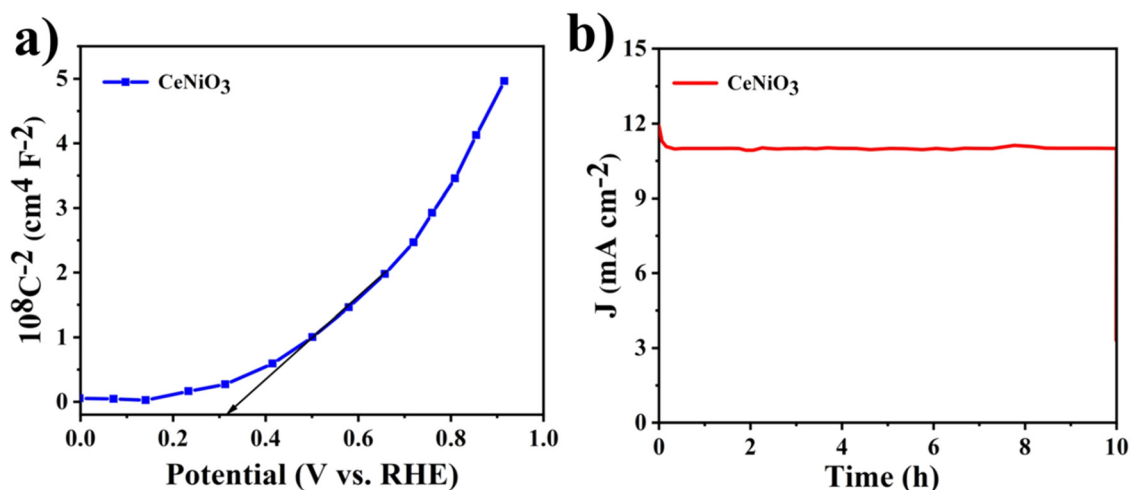


Fig. 10 (a) M–S plot and (b) stability measurements of the CeNiO₃ photoelectrode.



Table 2 A comparative analysis of the PEC efficiency of the CeNiO₃ perovskite against other materials

Materials	Conditions	<i>J</i> (mA cm ⁻²) (V vs. RHE)	Ref.
CeNiO ₃	0.5 M Na ₂ SO ₄	15.14 at 1.4	Present work
CdS	0.1 M Na ₂ SO ₃	5.10 at 1.23	30
O ₃ /W	0.5 M H ₂ SO ₄	1.71 at 1.23	54
TiO ₂ /BiVO ₄	0.1 M KOH	1.5 at 1.23	55
ZnO	0.25 M KOH	0.6 at 2.4	29

remains the same, suggesting that the catalyst's structural integrity is well maintained. The absence of any apparent particle fusion or agglomeration indicates that the CeNiO₃ surface has maintained its dispersed state. Under reaction conditions, the irregularly aggregated and flake-like particles remain distinct and exhibit good morphological stability and resistance to irradiation. The XPS spectra (Fig. 11(c)) show no noticeable shift or change in peak intensities or binding energy positions upon post-analysis, suggesting that the oxidation states of Ce, Ni, and O do not significantly alter following reaction. This confirms the structural integrity and chemical stability of CeNiO₃ by indicating that its surface composition

and chemical environment are stable and that there has not been a noticeable shift in the electronic structure.

3.11. Detailed reaction mechanism of water splitting in PEC systems

The mechanism by which CeNiO₃ enables water splitting encompasses a series of steps that leverage sunlight to yield H₂ and O₂. As illustrated in Fig. 12, this process begins with CeNiO₃ absorbing photons, leading to electron-hole pair creation. These charge carriers engage in redox reactions with water adsorbed on the CeNiO₃ surface, where holes oxidize water to release O₂, and electrons reduce protons to form H₂. CeNiO₃'s structure aids in efficient charge separation, positioning it as a key material in solar-driven hydrogen production. The reactions at both electrodes are summarized here.

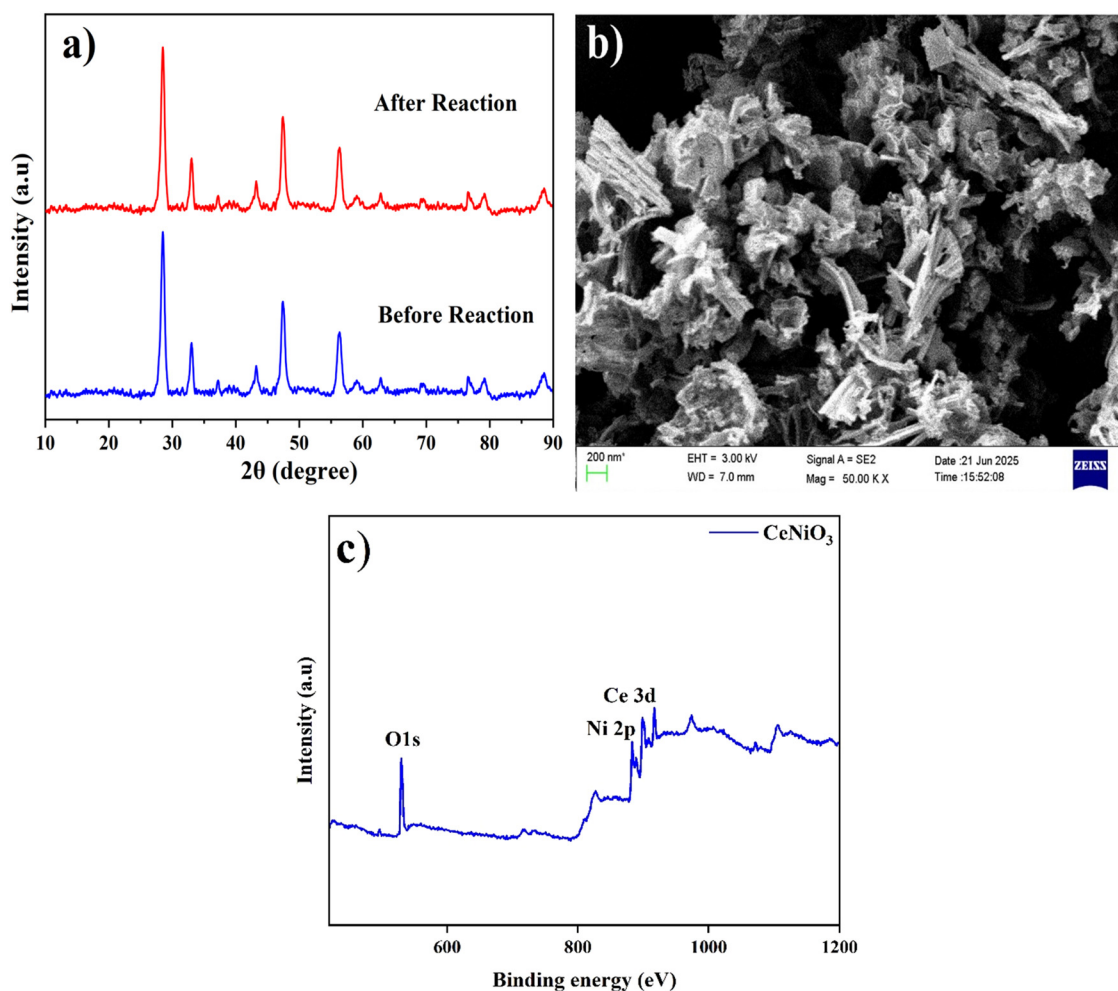
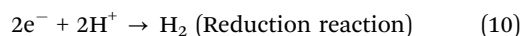


Fig. 11 Post analysis of CeNiO₃: (a) XRD, (b) SEM and (c) XPS.



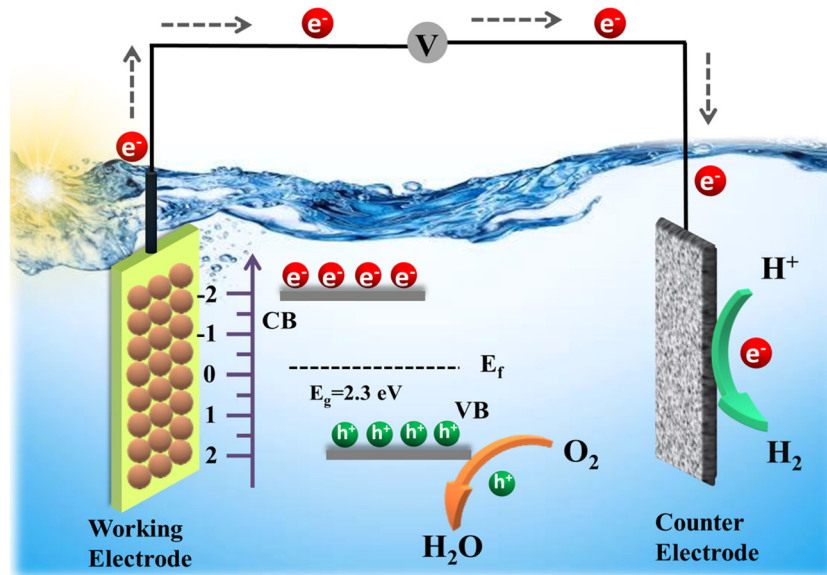


Fig. 12 Visual depiction of the PEC water splitting mechanism over the CeNiO₃ photoelectrode.

4. Conclusion

To summarize, this work presents the fabrication of CeNiO₃ perovskite structures using a citrate sol-gel technique, specifically designed for an enhanced hydrogen evolution reaction under a visible light driven PEC based water splitting process. Comprehensive analyses were performed on the material to investigate its crystalline arrangement, morphological features, elemental makeup, and optical properties. These assessments conclusively validated the fabrication of the CeNiO₃ perovskite with the targeted structural and chemical attributes. XRD scans confirmed the orthorhombic phase in CeNiO₃ and FE-SEM photographs of the specimen uncovered an uneven, porous structure. The CeNiO₃ photocatalyst displayed a narrow band-gap of 2.3 eV, emerging as an ideal candidate for harnessing the power of visible light irradiation. Moreover, the VB and CB potentials of the CeNiO₃ material are calculated to be 2.28 and -0.02 eV, respectively. CeNiO₃ has been deployed as a photoelectrode for PEC measurements under illumination. The cerium nickel oxide photoelectrode exhibits outstanding PEC activity, with a current density of 15.14 mA cm⁻² recorded at 1.4 V and an ABPE% of 2.76%. Remarkably, the CeNiO₃ electrode demonstrates exceptional photoelectrochemical stability over a prolonged duration, with negligible photocurrent density deterioration. The aforementioned results illustrate the immense potential of CeNiO₃ as a stable and efficient photoelectrode material for hydrogen generation. Future objectives may entail further optimisation of synthetic parameters, exploration of CeNiO₃-based tandem PEC systems, and implementation into practical devices for real-world uses, thus promoting the shift towards a carbon-neutral energy economy.

Author contributions

Anusha Hosakote Shankara: examination; data organization, systematic/theoretical analysis, conception, and writing – initial draft. Vinod Divya: visualization and review & editing. V. Rajeshwar, Usha Jinendra, Jagadeep Chandra S, Elayaperumal Sumitha, Basavarajappa Sannappa Hanumanthappa, Kotermame Mallikarjunappa Anilkumar, Peter R. Makgwane, and Honnegowdanahalli Shivabasappa Nagendra Prasad: review, validation, and proofreading. Mohammad Khalid and Shadma Wahab: resource and visualization. Harikaranahalli Puttaiah Shivaraju: conceptualization, methodology, supervision, writing, and proofreading – review & editing.

Conflicts of interest

The authors declare that they have no conflicts of interest.

Data availability

The data supporting the findings of this study are available within the article and its supplementary information (SI). Supplementary information is available. See DOI: <https://doi.org/10.1039/d5ma00914f>.

Additional data that support the study's conclusions are available from the corresponding author upon reasonable request.

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