

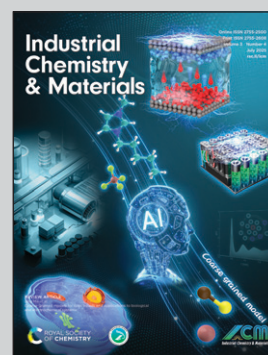
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Institute of High Energy Physics, Chinese Academy of
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Experimental and theoretical progress on the reduction of
 Np(VI) with salt-free reagents in the PUREX process

This review summarizes the reduction reaction of Np(VI)
to Np(V) by hydrazine, hydroxylamine, aldehydes, oximes,
hydroxamic acids and their derivatives as well as the reaction
kinetics, reaction mechanism and electronic structures.

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Experimental and theoretical progress on the reduction of Np(vi) with salt-free reagents in the PUREX process

Xin Huang,^{†a} Xiao-Bo Li,^{†a} Qun-Yan Wu ^{*a} and Wei-Qun Shi ^{*ab}

Effectively controlling the oxidation state of neptunium (Np) is crucial for the separation of Np during the advanced plutonium uranium reduction extraction process. The reduction reactions and kinetics of Np(vi) with salt-free reagents were explored by applying experimental and theoretical studies. This review summarizes the reduction reaction, kinetics, mechanism and electronic structures as well as the potential energy surfaces of Np(vi) to Np(v) using salt-free reagents, such as hydrazine, hydroxylamine, aldehydes, oximes, hydroxamic acids and their derivatives. This review will hopefully serve as a useful resource to inspire further research on the reduction of Np(vi) using salt-free reagents.

Keywords: Reduction kinetics; Reduction mechanism; Np(vi); Salt-free reagents; Theoretical simulation.

1 Introduction

The plutonium uranium reduction extraction (PUREX) process is effectively applied in the commercial scale reprocessing of spent nuclear fuel (SNF). In the PUREX process, U and Pu are separated from HNO₃ solution to the

organic phase using tri-*n*-butyl phosphate (TBP). The extraction performance of U, Np and Pu is significantly sensitive to their oxidation states in the PUREX process. Generally, hexavalent and tetravalent actinide ions exhibit higher extractability compared with trivalent and pentavalent ones. Therefore, the extractability of these actinides can be changed by oxidation or reduction, which achieves independent separation of U, Np and Pu in the different flowsheets. Actinides are very difficult to separate because they are similar in many respects.^{1,2} Thus, an advanced PUREX process is desired to achieve comprehensive control of U, Np, and Pu extraction in a single extraction cycle.²³⁷ Np has an exceptionally long half-life of 2.14×10^6 years³ and

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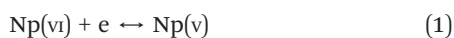


Qun-Yan Wu

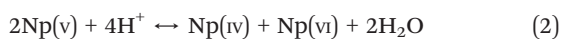
Dr. Qun-Yan Wu received her BE and ME degrees from Qufu Normal University (2006) and PhD from Beijing Institute of Technology (2010). She then pursued postdoctoral research at Tsinghua University. She joined Institute of High Energy Physics, Chinese Academy of Sciences in 2013 as a research assistant and is currently an associate professor. Her scientific interests include the theoretical design of lanthanum/actinium separation ligands, study of the reduction mechanism of Np(vi) by a salt-free reductant and study of actinide multiple bonds and oxidation states.



poses significant environmental hazards and radiotoxicity.⁴ Furthermore, ²³⁷Np serves as a key precursor in the production of ²³⁸Pu, an essential component in the manufacture of nuclear-powered batteries.⁵ Therefore, the separation of Np is very crucial and urgent. Np exhibits three oxidation states during the PUREX process, and their extractability by TBP follows the sequence of Np(vi) > Np(IV) ≫ Np(v).^{6–8} The valence state of Np in aqueous nitric acid depends on the concentration of HNO₂.^{9–11} At low concentrations of HNO₂, Np(v) is rapidly oxidized to Np(vi), while at high concentrations of HNO₂, Np(vi) is reduced to Np(v). HNO₂ exhibits a dual role in equilibrium reactions. At low concentrations, it acts as a catalyst, facilitating the oxidation of Np(v). Conversely, at high concentrations, it acts as a reductant, converting Np(vi) to Np(v). The redox reaction between Np(vi) and Np(v) is quasi-reversible and has simple one-electron transfer characteristic:



In addition, Np(v) slowly disproportionates, generating Np(vi) and Np(IV) in highly acidic solution through the following reaction. The extent of the disproportionation is promoted when the acidity of the solution and the concentration of Np(v) are high:



The valence state of Np ions in the different stages of the PUREX process is displayed in Fig. 1. The most of Np in the co-decontamination step is oxidized to Np(vi) by nitric acid and is extracted to the organic solvent accompanied with U(vi) and Pu(IV). In the U/Pu partition step, Np(vi) is reduced

to Np(IV) or Np(v) by reductants and Np is distributed to the U and the Pu purification cycles. It is essential to isolate Np ions from Pu and U ions before the U/Pu split stream to simplify the PUREX process. In the advanced PUREX, it is necessary to effectively control the oxidation state of Np ions.^{6,12}

Extensive studies have been conducted with the aim of developing an advanced PUREX process involving an efficient precise control of the Np oxidation state. There are several methods to control the Np oxidation state in the flowsheet of the PUREX process: Np can be entirely directed to (1) the high-level liquid waste (HLLW) stream, (2) the Pu purification stage, or (3) the U purification stage before the separation of the U and Pu stream. Np speciation in HNO₃ solution is dominated by the Np(vi)/Np(v) equilibrium, which is a very complex process due to the interactions among reaction kinetics, chemical equilibria, acidity, temperature and the presence of other ions.^{9–11} In particular, the Np(vi)/Np(v) equilibrium is sensitive to the HNO₃:HNO₂ ratio and temperature. For example, the concentration of Np(vi) is enhanced at higher HNO₃ concentrations, lower HNO₂ concentrations and higher temperatures.⁹ Salt-free reagents are considered to be promising for Np(vi) reduction because they can be decomposed into gases, such as N₂, CO₂ and H₂O.¹³ The reduction behaviors of Np(vi) with many salt-free reagents have been explored by the groups of Taylor,^{1,6,14,15} Uchiyama,^{16,17} Koltunov^{18,19} and Ye.^{20,21} A variety of salt-free reductants can effectively reduce Np(vi) to Np(v), such as hydrazine,^{14,22–25} aldehydes,^{16,17} hydroxylamine,^{26–30} oxime,^{31–34} hydroxamic acid^{15,35} and their derivatives. Recently, we systematically explored the Np(vi) reduction using hydrazine, hydroxylamine and its derivatives and elucidated their reduction mechanisms by employing density functional theory (DFT).^{36–45} Koltunov and co-workers reported the reaction kinetics of Np and Pu ions using hydrazine, organic nitriles, hydroxylamine, monoximes, diaziridinylethane and their derivatives.^{18,20,21} Marchenko *et al.* described the characteristics of organic reductants that can keep Np and Pu in a specific oxidation state.⁴⁶ Although much research has focused on Np(vi) reduction by salt-free reagents, there is no review on the subject. Noticeably, some salt-free reagents can reduce Np(vi) to Np(v) and reduce Np(v) to Np(IV). In this review, we mainly elucidate the experimental and theoretical progress on the reduction of Np(vi) to Np(v) by salt-free reagents. We hope that this review will serve as a useful resource to inspire further research on Np(vi) reduction by salt-free reagents in the reprocessing of SNF.



Wei-qun Shi

Professor Wei-qun Shi completed his Bachelor of Engineering degree at Hubei University of Technology in 1999, followed by a Master of Engineering from the China Institute of Atomic Energy in 2002 and a Doctor of Philosophy from Tsinghua University in 2007. Subsequently, he joined Institute of High Energy Physics, Chinese Academy of Sciences, where he worked until 2024. Currently, he serves as a full professor at Shanghai Jiao

Tong University. His ongoing research interests focus on actinide coordination chemistry, as well as the separation of actinides from lanthanides through solvent extraction and molten salt electrolysis techniques. To date, he has co-authored more than 400 scientific articles published in prestigious international peer-reviewed journals, contributing significantly to the relevant academic fields.

2 Experimental progress on the reduction of Np(vi) to Np(v) by salt-free reagents

2.1 Hydrazine and its derivatives

Hydrazine and its derivatives can reduce Np(vi) to Np(v), which enables the separation of Np in SNF.^{14,22–25,47} There are about 20 types of hydrazine and its derivatives studied in



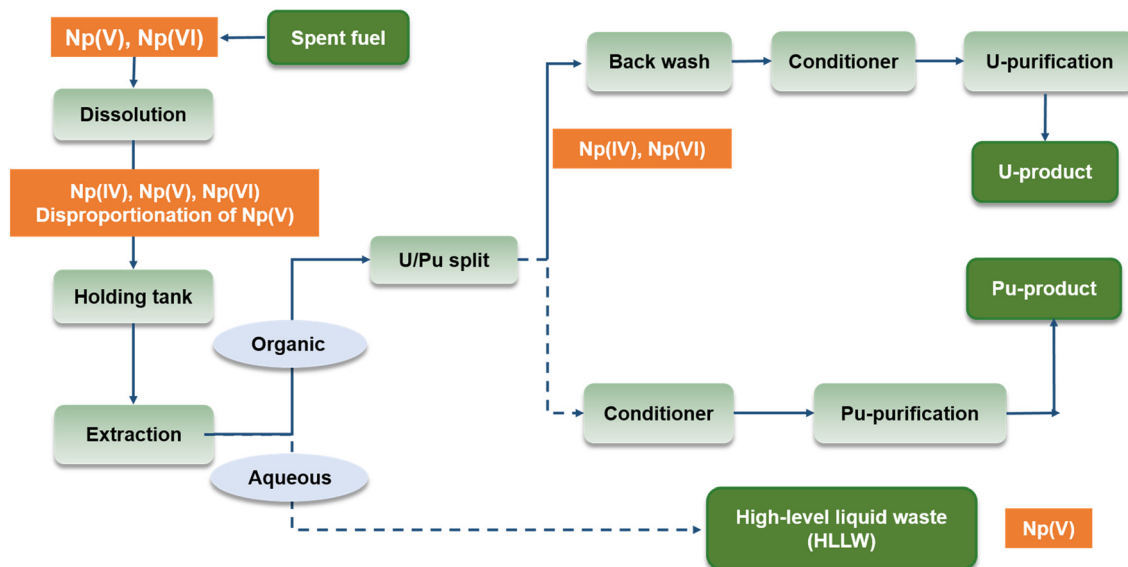


Fig. 1 Distribution of Np(V_{VI},V_{IV}) in the PUREX process.

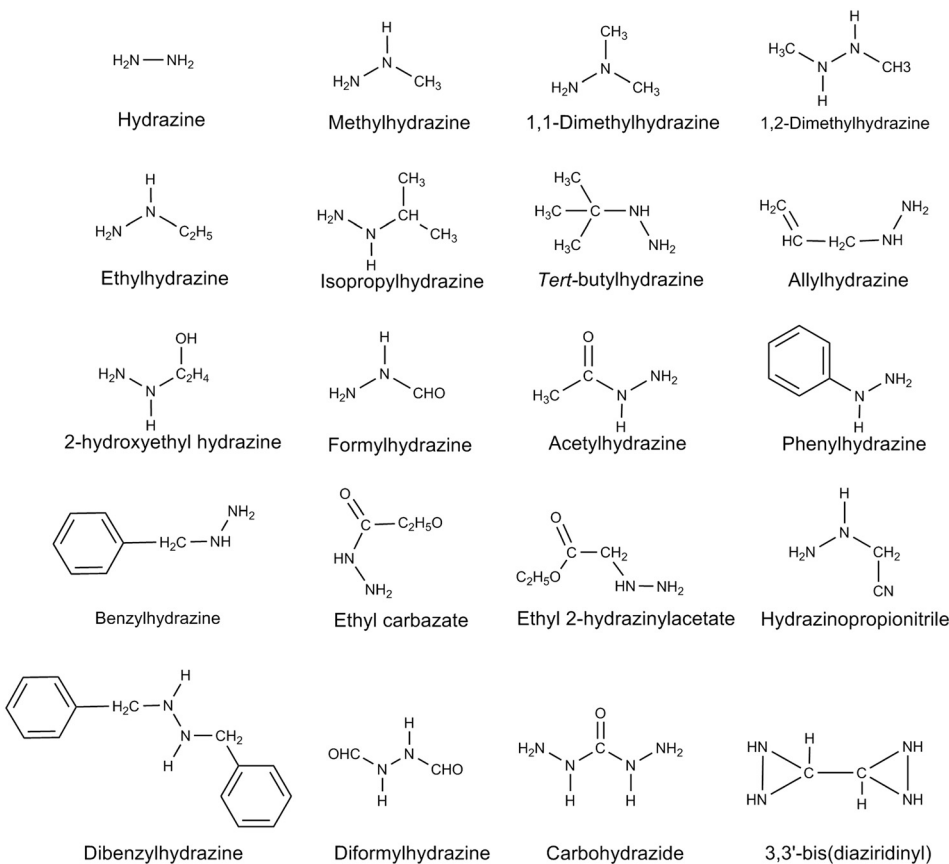
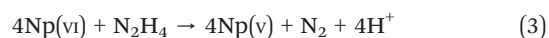
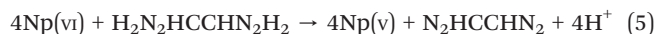
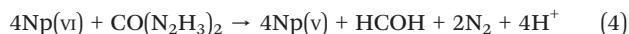


Fig. 2 Structures of hydrazine and its derivatives for Np(vi) reduction studies.

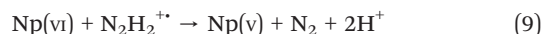
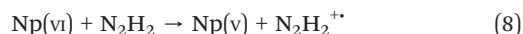
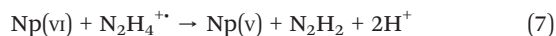
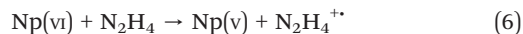
the experiment presented in Fig. 2, some of which can reduce four Np(vi) ions such as hydrazine, carbohydrazide, and 3,3'-bis(diaziridinyl), while the others only enable the reduction of two Np(vi) ions. The total reaction equation for excess

Np(vi) and hydrazine, carbohydrazide and 3,3'-bis(diaziridinyl) is expressed in eqn (3)–(5), respectively.

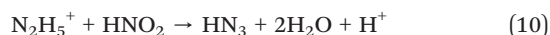




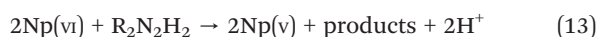
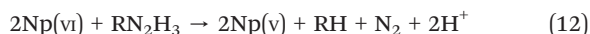
Taking hydrazine as an example, the reduction mechanisms for the excess $\text{Np}(\text{vi})$ and N_2H_4 are shown in eqn (6)–(9).²²



Hydrazine can react with nitric acid and form hydrazinium nitrate, which is often used as a scavenging agent for nitrous acid (HNO_2) in the PUREX process. The reaction rate for the reduction of $\text{Np}(\text{vi})$ is enhanced by increasing the hydrazinium nitrate concentrations.²⁴ The kinetics of the reaction of hydrazinium nitrate and HNO_2 in eqn (10) is very fast. Moreover, the HN_3 product can react with HNO_2 in eqn (11) as soon as all the N_2H_5^+ has disappeared.⁴⁸

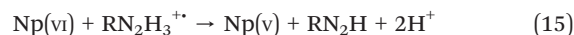
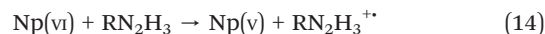


The total reaction equation for $\text{Np}(\text{vi})$ with excess RN_2H_3 and $\text{R}_2\text{N}_2\text{H}_2$ is given in eqn (12) and (13), respectively.

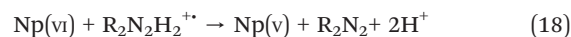
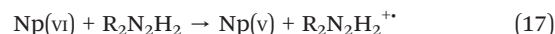


where R is a substituent group of hydrazine.

The reaction mechanism of eqn (12) is shown in eqn (14)–(16)



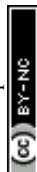
The reaction mechanism of eqn (13) corresponds to eqn (17)–(19)



We summarized temperature, equation of the reaction rate, activation energy (E) and rate constant for the reaction of $\text{Np}(\text{vi})$ reduction with hydrazine and its derivatives in Table 1. The oxidation rate of hydrazine and its derivatives in the HNO_3 solution follows the first-order kinetic equation, while its order with respect to H^+ is about -1 . These results indicate that the reaction rate is not sensitive to the ionic strength. The activation energy for phenylhydrazine and hydrazinopropionitrile is smaller than that of other hydrazine derivatives.^{18,49} Koltunov *et al.* first studied the kinetics of $\text{Np}(\text{vi})$ reduction using different hydrazine derivatives,^{22,47} and revealed that the reduction by these derivatives exhibits a significant difference. Their reduction ability in HNO_3 solution, judged by the reaction rate, follows the order of

Table 1 Temperature (t), kinetic equation, rate constant (k) and activation energy (E) for the $\text{Np}(\text{vi})$ reduction reaction with hydrazine and its derivatives

Reductants (R)	t (°C)	$-\frac{d[\text{Np}(\text{vi})]}{dt} = \frac{k[\text{Np}(\text{vi})][\text{R}]^e}{[\text{H}^+]^n}$	k (min^{-1})	E (kJ mol^{-1})	Ref.
N_2H_4	25	$e = 1, n = 1.24$	14.0	78.7	50
$\text{CH}_3\text{N}_2\text{H}_3$	25	$e = 1, n = 1$	52.7	58.6	50
$(\text{CH}_3)_2\text{N}_2\text{H}_2$	25	$e = 1, n = 0.9$	$55.3 (\text{L mol}^{-1})^{0.1}$	47.2	51
$\text{CH}_3\text{N}_2\text{H}_2\text{CH}_3$	25	$e = 1, n = 0.8$	118	50.2	18, 22
$\text{C}_2\text{H}_5\text{N}_2\text{H}_3$	25	$e = 1, n = 1$	30	61.5	18, 22
$(\text{CH}_3)_2\text{CHN}_2\text{H}_3$	25	$e = 1, n = 0.9$	19.3	69.4	18
$(\text{CH}_3)_3\text{CN}_2\text{H}_3$	25	$e = 0.9, n = 1$	$2.18 \pm 0.19 (\text{L mol}^{-1})^{0.1}$	63.1 ± 2.7	14
$(\text{CH}_3)_3\text{CN}_2\text{H}_3$	25	$e = 0.9, n = 0.75$	$5.4 (\text{L mol}^{-1})^{0.15}$	61.2	52
$\text{CH}_2\text{CHCH}_2\text{N}_2\text{H}_3$	25	$e = 1, n = 1$	46	63.6	18, 22
$\text{HO}(\text{CH}_2)_2\text{N}_2\text{H}_3$	25	$e = 1, n = 1$	391	56.6	53
HCON_2H_3	25	$e = 1, n = 1$	7.95	85.2	18, 22
$\text{CH}_3\text{CON}_2\text{H}_3$	25	$e = 1, n = 1.2$	7.52	76.5	18, 22
$\text{C}_6\text{H}_5\text{N}_2\text{H}_3$	25	$e = 1, n = 0.4$	3000	31.4	18
$\text{C}_6\text{H}_5\text{CH}_2\text{N}_2\text{H}_3$	25	$e = 1, n = 1$	53.2	98.2	18
$\text{C}_2\text{H}_5\text{COON}_2\text{H}_3$	25	$e = 1, n = 1.2$	19.3	65.4	18, 22
$\text{C}_2\text{H}_5\text{COOCH}_2\text{N}_2\text{H}_3$	25	$e = 1, n = 1.1$	97.6	70.1	18
$\text{NCCH}_2\text{N}_2\text{H}_3$	25	$e = 1, n = 1$	$920 \pm 38 (\text{L mol}^{-1})^{0.1}$	38.7 ± 1.3	49
CHONHNHCHO	15	$e = 1.3, n = 1.55$	$10.4 \pm 0.3 (\text{L mol}^{-1})^{0.25}$	85 ± 10	54
$\text{N}_2\text{H}_3(\text{CO})\text{N}_2\text{H}_3$	15	$e = 1.15, n = 1.35$	$39.7 \pm 2.7 (\text{L mol}^{-1})^{0.2}$	85 ± 20	55
$\text{H}_2\text{N}_2\text{HCCHN}_2\text{H}_2$	35.7	$e = 1, n = 1$	132 ± 10	73.1 ± 4.6	56



$\text{CH}_3\text{CON}_2\text{H}_3 < \text{HCON}_2\text{H}_3 < \text{N}_2\text{H}_4 < \text{C}_2\text{H}_5\text{OCON}_2\text{H}_3 < \text{C}_2\text{H}_5\text{N}_2\text{H}_3 < \text{CH}_2\text{CHCH}_2\text{N}_2\text{H}_3 < \text{CH}_3\text{N}_2\text{H}_3 < \text{C}_6\text{H}_5\text{CH}_2\text{N}_2\text{H}_3 < \text{C}_2\text{H}_5\text{OCOCH}_2\text{N}_2\text{H}_3 < \text{CH}_3\text{N}_2\text{H}_2\text{CH}_3 < \text{HOC}_2\text{H}_4\text{N}_2\text{H}_3$. The reduction rate of $\text{Np}(\text{vi})$ is enhanced by some hydrazine derivatives modified by electron-donating groups, which is ascribed to the increase of the electron density around the N atom by introducing electron-donating groups. Oppositely, some hydrazine derivatives modified by electron-withdrawing groups will make the $\text{Np}(\text{vi})$ reduction more difficult. In the reaction process, the free radical ion $\text{RN}_2\text{H}_3^{+\cdot}$ is an intermediate product of the first $\text{Np}(\text{vi})$ reduction by RN_2H_3 , which can result in the second $\text{Np}(\text{vi})$ reduction.

Koltunov reported the rate constants, mechanisms, activation energies, and kinetic equations for $\text{Np}(\text{vi})$ reduction with sixteen kinds of hydrazine derivatives in HNO_3 solution.¹⁸ Xiao *et al.* obtained a quantitative correlation between the molecular structures of hydrazine derivatives with the $\text{Np}(\text{vi})$ reduction rate.²¹ They concluded that the reduction rates are mainly related to the molecular dipole moment, the level of LUMO as well as hydrophobic parameters.

Taylor *et al.*¹⁴ investigated the solvent extraction of Np ions by selective reduction of $\text{Np}(\text{vi})$ using salt-free reagents and found that *tert*-butylhydrazine can be a potential candidate for the selective reduction of $\text{Np}(\text{vi})$ in certain conditions.⁵⁷ The reaction kinetics for $\text{Np}(\text{vi})$ with *tert*-butylhydrazine were explored using spectrophotometry in HNO_3 solution.⁵² The reduction rate of $\text{Np}(\text{vi})$ by *tert*-butylhydrazine in TBP/ HNO_3 solution is similar to that by dibenzylhydrazine.⁵⁸ Subsequently, Koltunov and coworkers explored the kinetic behavior and mechanisms of $\text{Np}(\text{vi})$ reduction by dibenzylhydrazine in TBP/ HNO_3 solution and revealed the final reaction products, kinetic equation, and activation energy with an excess of the reductant.⁵⁹

The separation of Np from Pu was explored using hydrazine as a stripping reductant in a single-stage extraction experiment and the Np/Pu separation efficiency is enhanced by elevating the concentrations of hydrazine, nitric acid, and temperature.²⁴ The feasibility of the Np/Pu separation with methylhydrazine was also assessed by a single-stage extraction device, which suggested that the separation ability of $\text{Np}(\text{vi})/\text{Pu}(\text{iv})$ becomes weaker as the reaction time gets longer, which is similar to hydrazine.^{24,25} The $\text{Np}(\text{vi})$ selective reduction by *N,N*-dimethylhydrazine within a TBP/*n*-dodecane HNO_3 solution including $\text{Pu}(\text{iv})$, $\text{U}(\text{vi})$ and $\text{Np}(\text{vi})$ was explored by spectrophotometry and numerical simulation.²³ The kinetic behavior for the $\text{Np}(\text{vi})$ reduction with $\text{CH}_3\text{N}_2\text{H}_3$ and $\text{HOC}_2\text{H}_4\text{N}_2\text{H}_3$ and their application in U/Np separation were investigated by Zhang *et al.*^{53,60–62} Yin and coworkers also explored the reduction kinetics of $\text{Np}(\text{vi})$ by $(\text{CH}_3)_2\text{N}_2\text{H}_2$, which enables $(\text{CH}_3)_2\text{N}_2\text{H}_2$ as selective reductant of $\text{Np}(\text{vi})$ for U/Np separation.^{51,63} The rate-determining step for $\text{Np}(\text{vi})$ reduction by $\text{HOC}_2\text{H}_4\text{N}_2\text{H}_3$, $(\text{CH}_3)_2\text{N}_2\text{H}_2$, and $\text{CH}_3\text{N}_2\text{H}_3$ in the kerosene-30% TBP/ HNO_3 solution is the back-extraction of $\text{Np}(\text{v})$ from the kerosene-30% TBP into the HNO_3 solution,⁶⁴ which is important to separate Np in the SNF reprocessing. Carbohydrazide is an effective reductant of $\text{Np}(\text{vi})$ in high-burn-up SNF. Volk *et al.*

found that $\text{Np}(\text{vi})$ is quickly reduced to $\text{Np}(\text{v})$ using carbohydrazide, but the reduction rate of $\text{Np}(\text{v})$ to $\text{Np}(\text{iv})$ is very slow.⁶⁵ Subsequently, Zavalina and coworkers examined the kinetic behavior of $\text{Np}(\text{vi})$ reduction to $\text{Np}(\text{v})$ using carbohydrazide in HNO_3 solution,⁵⁵ and Shilov and Fedoseev explored the corresponding reduction in HClO_4 solution.⁶⁶ The reduction rate for $\text{Np}(\text{vi})$ to $\text{Np}(\text{v})$ via carbohydrazide improves with increasing temperature, the concentration of carbohydrazide, and lowering the concentration of HNO_3 . The activation energy for the corresponding reduction reaction is 85 kJ mol^{-1} in HNO_3 solution at $15 \text{ }^\circ\text{C}$ and $86 \pm 5 \text{ kJ mol}^{-1}$ in HClO_4 solution at $25 \text{ }^\circ\text{C}$.

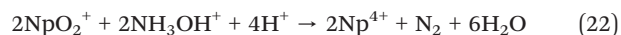
2.2 Hydroxylamine and its derivatives

In the PUREX process, a strategy for controlling Np separation includes the extraction of $\text{U}(\text{vi})$, $\text{Np}(\text{vi})$, and $\text{Pu}(\text{iv})$ together and the separation of $\text{Pu}(\text{iii})$ and $\text{Np}(\text{v})$ from $\text{U}(\text{vi})$. Effective and rapid reduction of $\text{Np}(\text{vi})$ to $\text{Np}(\text{v})$ and $\text{Pu}(\text{iv})$ to $\text{Pu}(\text{iii})$ necessitates the application of a suitable reductant. Hydroxylamine (HA) and its derivatives have been extensively explored for this purpose. Notably, HA enables rapid reduction of $\text{Np}(\text{vi})$ to $\text{Np}(\text{v})$.^{46,57} HA is further able to reduce $\text{Np}(\text{v})$ to $\text{Np}(\text{iv})$ with increasing acidity and temperature.^{1,57} The reduction equations of $\text{Np}(\text{vi})$ with HA are followed as eqn (20)–(22):^{26,67}

In excess $\text{Np}(\text{vi})$,



In excess HA,



The reduction rate constant of $\text{Np}(\text{vi})$ by HA is 3.5 s^{-1} in HNO_3 solution with $\mu = 2$ at $22 \text{ }^\circ\text{C}$.¹⁹ HA is quite stable at low acidity and room temperature. However, it can decompose into N_2O as the HNO_3 concentration and/or temperature increase, which is enhanced under the catalysis of Fe.⁶⁸ Additionally, the reduction of $\text{Np}(\text{vi})$ by HA in 0.5 M HNO_3 solutions with a high U concentration of 850 g L^{-1} at $60 \text{ }^\circ\text{C}$ predominantly produces $\text{Np}(\text{v})$, while in 0.33 M HNO_3 solutions under the same temperature and uranium concentration conditions, the reduced products consist of a mixture of $\text{Np}(\text{v})$ and $\text{Np}(\text{vi})$.²⁰

Most hydroxylamine derivatives exhibit a high reduction rate for $\text{Pu}(\text{iv})$ and $\text{Np}(\text{vi})$ (Fig. 3), making them suitable for the co-extraction of Np and Pu from U.^{19,58} The reaction rates of $\text{Np}(\text{vi})$ reduction by HA and its derivatives are presented in Table 2. HA derivatives avoid additional stabilizers due to their rapid reactivity with HNO_2 , which precludes the production of HN_3 and NH_4NO_3 .⁴⁶ At moderate acidity and room temperature, HA derivatives can reduce $\text{Np}(\text{vi})$ to $\text{Np}(\text{v})$, but they cannot further reduce $\text{Np}(\text{v})$ to $\text{Np}(\text{iv})$ in the absence



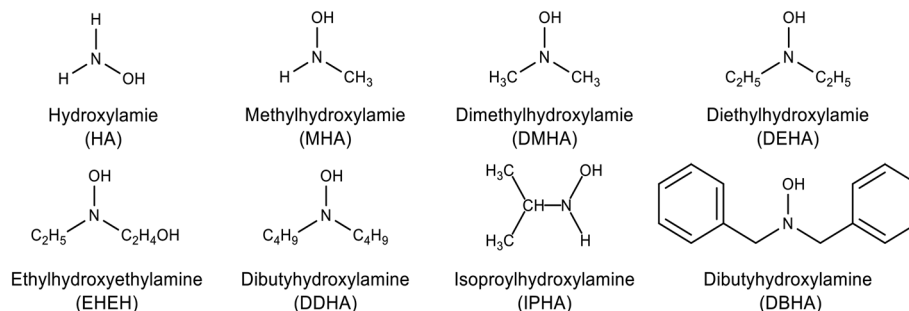


Fig. 3 Structures of hydroxylamine and its derivatives.

Table 2 Temperature (t), kinetic equation, rate constant (k), activation energy (E) for Np(vi) reduction by hydroxylamine and its derivatives at $\mu = 2$

Reductants (R)	t (°C)	The kinetic equation		k (min ⁻¹)	E (kJ mol ⁻¹)	Ref.
		$-\frac{d[\text{Np(vi)}]}{dt} = \frac{k[\text{Np(vi)}][\text{R}]^e}{[\text{H}^+]^n}$				
NH ₂ OH (HA)	15	$e = 1, n = 1$		92.1 ± 1.0	82.00	26
CH ₃ NHOH (MHA)	25	$e = 0.7, n = 0.4$		35.0 ± 0.9 (L mol ⁻¹) ^{0.3} min	63.6 ± 3	19
(CH ₃) ₂ NOH (DMHA)	25	$e = 1, n = 0.65$		24.2 ± 11 (L mol ⁻¹) ^{0.35} min	60.03 ± 0.02	70
(C ₂ H ₅) ₂ NOH	25.2	$e = 1, n = 1$		23.0 ± 1.8	69.1	28
C ₂ H ₅ (HOC ₂ H ₄)NOH (EHEH)	25.6	$e = 1, n = 0.8$		334 ± 12 (L mol ⁻¹) ^{0.2} min	42.3 ± 2.7	27
CH ₃ CH(CH ₃)NHOH (IPHA)	25.8	$e = 1, n = 0.8$		41.8 ± 1.0 (L mol ⁻¹) ^{0.2} min	58.20 ± 0.03	71
(C ₄ H ₉) ₂ NOH (DDHA)	25.5	$e = 1, n = 1$		15.6 ± 1.1	45.8 ± 0.8	72
(C ₆ H ₅ CH ₂) ₂ NOH (DBHA)	25	$e = 1, n = 1$		1.8	—	58

of catalysts,^{30,69} and the reaction products are mainly alcohols and aldehydes.¹⁹

The stoichiometry of the Np(vi) reaction with HA derivatives is determined by the starting concentrations of the reactants. In cases where there is an excess of Np(vi) ions, the stoichiometric coefficient for Np(vi):HA derivatives is (6–8):1. However, the corresponding stoichiometric coefficient decreases to (0.5–2):1 when the HA derivatives are in excess.⁵⁷

The reaction order relative to hydrogen ion ranges from 0.4 to 1.0.¹⁸ These reactions occur through different mechanisms, which contain several parallel pathways with R' RNOH and R'RNHOH.^{26,70–72} The mechanisms of Np(vi) reduction by mono- and disubstituted hydroxylamines are different. For the disubstituted hydroxylamines, the rate-determining step is the decomposition reaction of the activated complex including the nitroxyl radical, which undergoes further oxidation to nitron, then forms aldehyde and monoalkyl hydroxylamine, or nitrosoalkyl and alcohol. In the case of monosubstituted hydroxylamine, it forms nitroxide radical, which oxidizes to alcohol and nitrosoalkyl, with the nitrosoalkyl further forms HNO₃.

The reduction rates of Np(vi) by most hydroxylamine alkyl derivatives are slower compared with that of HA,^{57,67} except for *N,N*-ethyl(hydroxyethyl)hydroxylamine (EHEH) and *N,N*-dimethylhydroxylamine (DMAH). The reduction rate of Np(vi) by EHEH is much faster than that by *N,N*-diethylhydroxylamine (DEHA).⁵⁷ τ_{99} for the Np(vi) reduction by HA,²⁶ DMHA,¹⁹

DEHA¹⁹ and EHEH⁷³ is 0.12, 0.18, 0.15, and 0.14 min in 1.0 M HNO₃ and 0.01 M Np(vi) solution at 25 °C, respectively. Koltunov *et al.*^{57,67,74} studied the reduction kinetics of Np(vi) by *N*-methylhydroxylamine (MAH) and DMAH in 2.0 M HNO₃ at 22 °C, and the τ_{99} values are 1.13 and 0.38 min, respectively, while the corresponding τ_{99} values change to 0.65 min and 0.18 min in 2.0 M HNO₃ at 25 °C. These values indicated that the reduction rate of Np(vi) by DMAH is much faster compared with that of MAH under the same conditions.^{57,67,68,70}

Np(vi) is quickly reduced to Np(v) by DEHA in HNO₃ solution,²⁸ τ_{99} value for the reaction of DEHA and HNO₂ is 0.15 min at 1 M HNO₃ and 0.1 M DEHA solution.⁷⁵ DEHA is a better stabilizer compared with HA, The presence of U(vi) (up to 0.5 M) does not affect Np(vi) reduction by DEHA.^{19,28} That is, DEHA is a more effective reductant than HA as an antinitrite. Zhang *et al.* applied DEHA to isolate Np and Pu from U for the U(vi) purification cycle step in the PUREX process and suggested that DEHA can quickly reduce over 99% of Np(vi) to Np(v).²⁹ Moreover, the cascade experiment and the cascade extraction separation demonstrate that DEHA is a promising reductant for the PUREX process. The combination of DEHA with hydrazine derivatives like 2-hydroxyethylhydrazine or 1,1-dimethylhydrazine enables the reprocessing process to be carried out without the presence of salts.⁷⁶

The kinetics of Np(vi) by extraction mass transfer and reversible redox reaction by DEHA were examined. Zhang *et al.* discovered that the rate of Np(vi) transfer is obviously



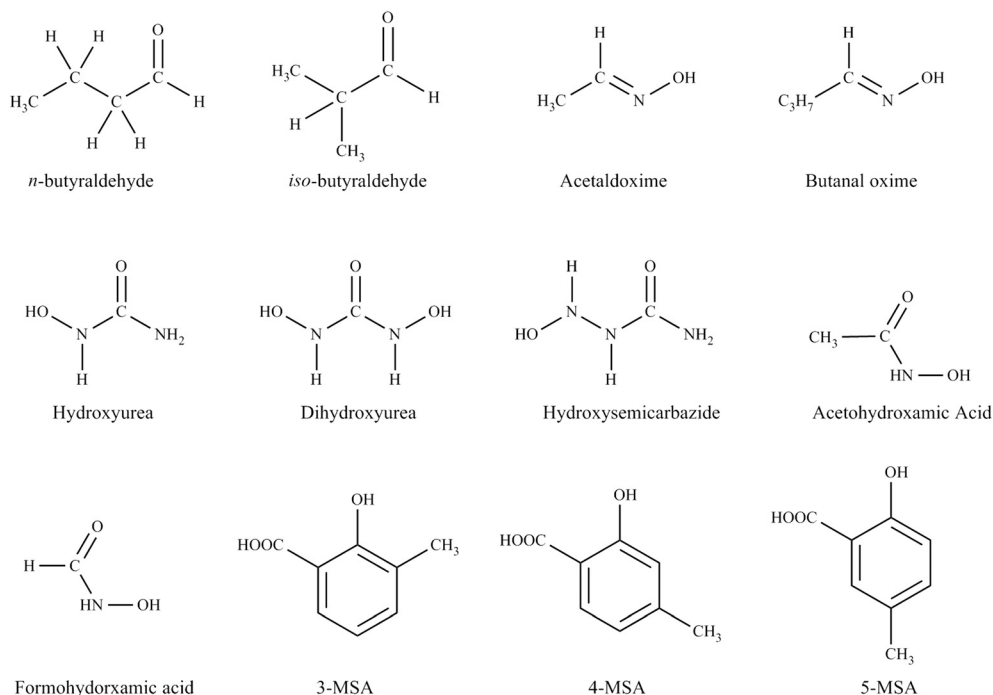


Fig. 4 Some structures of aldehydes, oxime, hydroxyurea, hydroxamic acid and methylsalicylic acids.

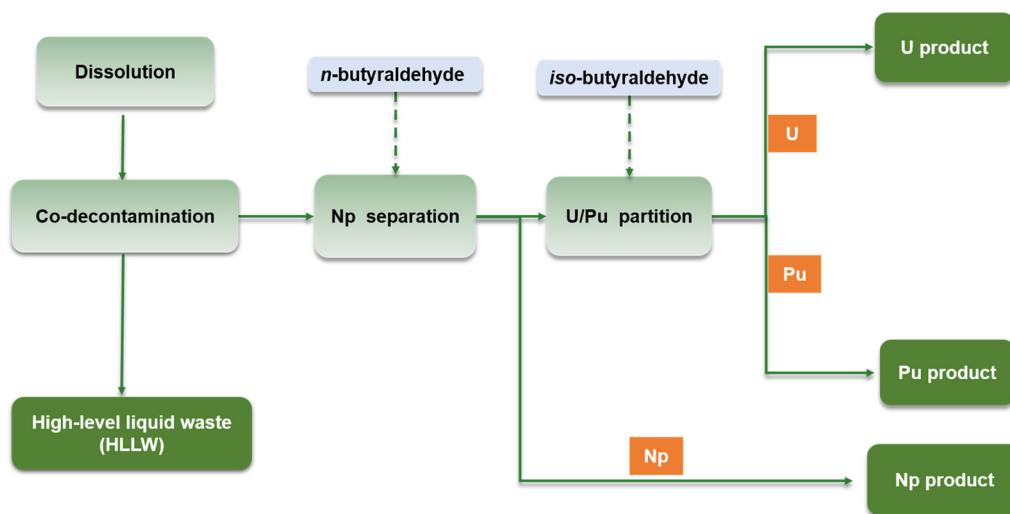
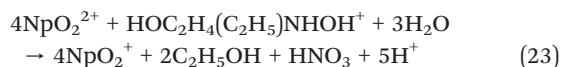


Fig. 5 U, Np, and Pu separation using *n*- and *iso*-butyraldehydes.

faster than that of the redox reaction.⁷⁶ Li's group further determined that the reduction rate of Np(vi) depends on the rate of Np(v) transfer from the kerosene-30% TBP organic phase to HNO₃ solution.⁷ EHEH is DEHA modified by a hydroxyl group, which can enhance the reduction rate of Np(vi).¹⁹ The reaction of Np(vi) with EHEH is:^{64,73}



This rate constant is much larger than that of Np(vi) reduction by DEHA (Table 2). EHEH exhibits adequate

stability in less than 3–4 M HNO₃ solution, even in the presence of Tc(vii) ions.²⁷ Moreover, the reaction of EHEH with HNO₂ is very slow, which eliminates the need for an additional scavenging agent.¹⁹ Therefore, the rapidity of the reduction renders EHEH an ideal candidate for application in the U/Pu split process.

2.3 Aldehyde

Uchiyama *et al.* introduced an innovative method for separating Np/Pu/U by utilizing *n*- and *iso*-butyraldehyde, which are illustrated in Fig. 4, capitalizing on their distinct



reduction capabilities.¹⁷ Both *n*- and iso-butyraldehydes are favorable reductants for selective Np(vi) reduction in the presence of Pu(IV) and/or U(vi). Moreover, *n*-butyraldehyde has much higher selectivity and a much lower reduction rate compared with iso-butyraldehyde. Therefore, *n*-butyraldehyde can selectively reduce Np(vi), while iso-butyraldehyde is utilized to reduce both Np(vi) and Pu(IV). According to the different reduction abilities of *n*- and iso-butyraldehyde, the selective Np, Pu and U separation process in the PUREX is proposed (Fig. 5). *n*-Butyraldehyde can extract and reduce about 99.98% Np in the Np separation process, and iso-butyraldehyde is employed for reducing Pu(IV) during the U/Pu split stage.

Subsequently, Uchiyama *et al.* conducted a more in-depth study on the reduction kinetics of Np(vi) by both iso- and *n*-butyraldehyde in HNO₃ solution.¹⁶ Iso-butyraldehyde shows superior performance for reducing Np(vi) compared with *n*-butyraldehyde. Np(vi) is totally converted to Np(v) in 0.069 M iso-butyraldehyde and 3 M HNO₃. This work showed that *n*-butyraldehyde appears to be a potential reagent for Np separation in the PUREX process. Subsequently, the reduction kinetics of Np(vi) with *n*-butyraldehyde in *n*-dodecane/30% TBP solution were examined with spectrophotometry.⁷⁷ The rate constant is about 10⁻¹⁰ smaller in *n*-dodecane than that in aqueous solution, suggesting that the Np(vi) reduction process with *n*-butyraldehyde primarily occurs in aqueous solution at 294 ± 1 K.

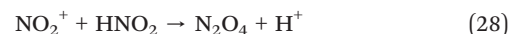
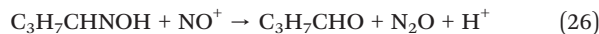
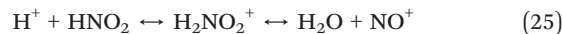
2.4 Oxime

Monomimes are known to be weakly acidic and basic, which have strong reducing abilities towards Np(vi) ions owing to the C=N double bond (Fig. 4). Acetaldoxime is an example of the monoximes, which is considered a valuable reagent for Np/U separation in the advanced PUREX process. Koltunov and coworkers investigated the reduction rate for the reaction of Np(vi) and acetaldoxime in HNO₃ solution:

$$-d[\text{Np(vi)}]/dt = k[\text{Np(vi)}][\text{CH}_3\text{CHNOH}]/[\text{HNO}_3] \quad (24)$$

where $k = 254 \pm 10 \text{ min}^{-1}$ with $\mu = 2$ at 26.0 °C as well as the activation energy of $62.6 \pm 2.6 \text{ kJ mol}^{-1}$.³¹ They also suggest that Np(vi) reduction by acetaldoxime may proceed *via* an intermediate CH₃CHNO radical and the hydrolysis product. Extraction studies show that Np(vi) is reduced to Np(v), which is stripped from 30% TBP/*n*-dodecane solution bearing U(vi). After studying the Np(vi) reduction with acetaldoxime, they show the advantages of acetaldoxime in that it reduces Np(vi) to Np(v) but does not form Np(IV).³⁴ Subsequently, Koltunov *et al.* reported that the reduction rate of Np(vi) by butanal oxime is $230 \pm 15 \text{ min}^{-1}$ with $\mu = 2$ at 25.0 °C and the activation energy in 1 M HNO₃ is $69.4 \pm 12.4 \text{ kJ mol}^{-1}$.³³ The results indicated that Np(vi) reduction by butanal oxime and acetaldoxime is a first-order reaction. The reduction kinetics of Np(vi) by butanal oxime were further studied in undiluted TBP in the presence of HNO₃. The reaction rate constant is $0.058 \pm 0.007 \text{ L}^{2.5} \text{ mol}^{-2.5} \text{ min}^{-1}$ between 0.01 to 0.27 M HNO₃

solution with the activation energy of $79 \pm 9 \text{ kJ mol}^{-1}$ at $T = 25.0 \text{ °C}$.³² Noticeably, oxime can act as a scavenging agent for HNO₂. For example, butanal oxime can react with HNO₂ and obtain N₂O and N₂ as shown in eqn (25)–(27). The NO₂⁺ formed in eqn (27) can be quickly eliminated by reaction with HNO₂ as shown in eqn (28).⁷⁸



2.5 Hydroxyurea

Hydroxyurea (HU) presented in Fig. 4 is also a potential salt-free reagent for U/Np separation. Zhu and coworkers reported that Np(vi) is reduced to Np(v) with HU, in which Np(v) is effectively transferred from TBP to the aqueous phase.⁷⁹ Noticeably, the effectiveness of Np stripping is weakened in HNO₃ solution since the Np(vi) reduction is impeded at higher acidity.

Spectrophotometric analysis was applied to study the Np(vi) reduction by dihydroxyurea (DHU).⁸⁰ The reduction rate of Np(vi) by DHU is quick, which makes it suitable for the U/Np separation. The rate constant for the above reaction is 1.86 s^{-1} in $7.5 \times 10^{-2} \text{ M}$ DHU and 0.44 M HNO₃ solution at 4 °C. Under the given experimental conditions, 98% of Np(vi) was quickly stripped to the aqueous phase. The stripping ability decreases with increasing [HNO₃], while it increases with increasing [DHU], so DHU can be used to separate Np in the advanced PUREX process. The reduction rate of Np(vi) by hydroxysemicarbazide (HSC) is also quick with a rate constant of $1037 \pm 60 \text{ M}^{-1.40} \text{ s}^{-1}$ at 4.0 °C,⁸¹ it is positively influenced by increasing [HSC] or temperature and decreasing [HNO₃] or ionic strength.

2.6 Hydroxamic acids

Hydroxamic acids with the general formula RCONHOH illustrated in Fig. 4 are excellent chelating reagents, which provides great benefits in separating Np in advanced PUREX. Furthermore, hydroxamic acids are easily decomposed to gases and consume HNO₂ without an additional reagent. Formohydroxamic acid (FHA) and acetohydroxamic acid (AHA) are hydrophilic ligands that do not enter into TBP. Hydroxamic acids undergo hydrolysis in less than 3 M HNO₃ and result in the formation of hydroxylamine and carboxylic acid. Taylor and co-workers found that Np(vi) is reduced to Np(v) by AHA and FHA in nitric acid.¹⁵ They verified that Np(vi) is effectively reduced and removed from the solvent phase bearing U(vi) by AHA and FHA in laboratory scale experiments.³⁵ Subsequently, further investigation of the reduction kinetics of Np(vi) by FHA in HNO₃ solution shows that it is rapidly reduced within a few seconds and is first



order with respect to $\text{Np}(\text{vi})$ by stopped-flow spectrophotometry with a rate constant of $1.17 \times 10^3 \text{ M}^{-1} \text{ s}^{-1}$ in 2 M HNO_3 solution at 22 °C.⁶⁷

Chung and Lee investigated the reduction of $\text{Np}(\text{vi})$ by AHA in HNO_3 solution.⁸² The reduction reaction is first-order relative to the concentration of $\text{Np}(\text{vi})$ and AHA with the rate constant of $191.2 \pm 11.2 \text{ M}^{-1} \text{ s}^{-1}$ at 25 ± 0.5 °C and 1.0 M HNO_3 solution. Investigation of the reduction stripping behavior of $\text{Np}(\text{vi})$ by AHA from kerosene/30% TBP organic phase to HNO_3 solution⁸³ shows that the stripping rate is dominated by the kinetic process at the interface. In addition, the reduction kinetics of $\text{Np}(\text{vi})$ with AHA was studied with infrared spectroscopy in 1 M HClO_4 media.⁸⁴ When AHA is in excess, the reduction rate equation of $\text{Np}(\text{vi})$ is shown below:

$$-d[\text{Np}(\text{vi})]/d_t = k[\text{Np}(\text{vi})][\text{AHA}] \quad (29)$$

where $k = 2.57 \times 10^3 \text{ M}^{-1} \text{ s}^{-1}$ at 10 °C.

When $\text{Np}(\text{vi})$ is in great excess relative to AHA, the mechanism of $\text{Np}(\text{vi})$ reduction is controlled by two different reactions. The reaction rate is 3.7×10^{-4} and $1.0 \times 10^{-3} \text{ s}^{-1}$ for the first order mechanism at 10 °C and 22 °C, respectively.

2.7 Others

Rao and Choppin explored the reduction kinetics of $\text{Np}(\text{v})$ with three methylsalicylic acids (MSA) as shown in Fig. 4 using spectrophotometry.⁸⁵ Their rate constants decrease in the order of 5-MSA > 3-MSA > 4-MSA, which is attributed to the electronic effect of substituents. Humic acid (HA) is one of the significant species in groundwater. The rate constant for $\text{Np}(\text{vi})$ reduction by HA is $59 \pm 6 \text{ M}^{0.1} \text{ min}^{-1}$ at 10 °C and $\mu = 0.41$.⁸⁶ Additionally, the reduction kinetics of $\text{Np}(\text{vi})$ by H_2O_2 was studied in a perchloric acid–sodium perchlorate solution.⁸⁷ The stoichiometric equation was established to be $2\text{Np}(\text{vi}) + \text{H}_2\text{O}_2 = 2\text{Np}(\text{v}) + 2\text{H}^+ + \text{O}_2$. The respective rate of $\text{Np}(\text{vi})$ reduction by H_2O_2 is $(2.19 \pm 0.01) \times 10^3$ in 0.05 M Na_2CO_3 at 25 °C.⁸⁸ In addition, ethylenediaminetetraacetic acid (EDTA) is a common organic chelating reagent that was studied to reduce $\text{Np}(\text{vi})$ in low ionic strength media.⁸⁹ EDTA can effectively reduce $\text{Np}(\text{vi})$ to $\text{Np}(\text{v})$ in the presence of organic complexes.

3 Theoretical progress on the reduction of $\text{Np}(\text{vi})$ to $\text{Np}(\text{v})$ by salt-free reagents

Simulation for accurate prediction of electronic structures and properties as well as the reaction mechanics of actinide systems is one of the most challenging issues in computational chemistry. Nevertheless, significant advances were achieved in the last few decades owing to the development of electronic structure methods for actinides and improvement in computational speed. Our group has theoretically explored the extraction behaviors of the actinide

systems and the corresponding electronic structures as well as the bonding characters.^{90–109} Recently, we have focused on exploring the mechanisms of $\text{Np}(\text{vi})$ reduction by some salt-free reagents using scalar-relativistic DFT calculations. All calculations of $\text{Np}(\text{vi})$ systems were performed using the B3LYP hybrid functional^{110,111} within the Gaussian 16 program.¹¹² The scalar-relativistic effective core potentials (RECPs)¹¹³ and the ECP60MWB-SEG valence basis set^{114,115} were applied for Np. The solvation effect was considered by implicit solvation models, solvation model density (SMD)¹¹⁶ or the conductor-like polarizable continuum model (CPCM) approach with Klamt's radii.^{117,118} Transition states (TSs), initial complexes (ICs) and intermediates (INTs) were confirmed by the harmonic frequency calculations. The calculations of the intrinsic reaction coordinate were carried out to verify the ICs and INTs. The bonding change of the reduction process was analyzed using the quantum theory of atoms-in-molecules (QTAIM),^{119,120} electron localization function (ELF),^{121–123} and localized molecular orbitals (LMOs)¹²⁴ using the Multiwfn.^{125,126} Analysis of spin density uncovered the oxidation state of the metal atom for ICs, TSs, and INTs, which can elucidate the nature of the reduction reaction.

3.1 Introduction to transition state theory and method

The studies on the mechanisms of $\text{Np}(\text{vi})$ reduction by salt-free reagents essentially depend on the transition state theory, which has been successfully applied to actinide systems.^{38,127–131} Here, we briefly introduce transition state theory. The concepts of potential energy surface and saddle point were originally introduced by Pelzer and Wigner.¹³² Later, Eyring and Polanyi proposed the transition state theory (TST) based on quantum mechanics and statistical mechanics when they studied organic reactions from reactants to products in 1935.^{133,134} TST includes classic and generalized TSTs. Classic TST is a statistical-mechanical

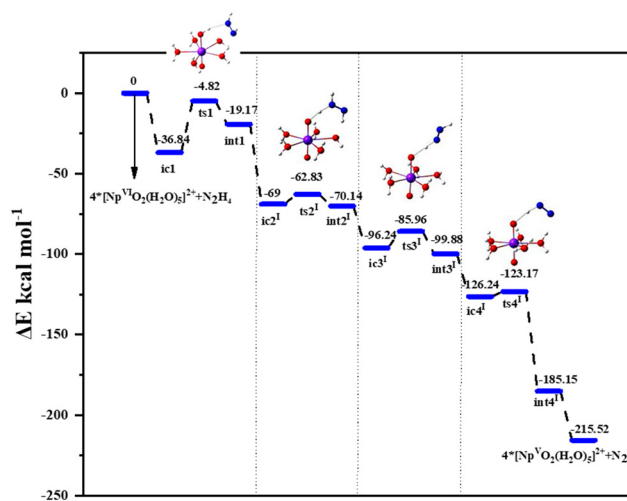


Fig. 6 PEP and four TS structures for $[\text{Np}^{\text{VI}}\text{O}_2(\text{H}_2\text{O})_5]^{2+}$ reduction by N_2H_4 .



theory of chemical reaction rate according to the no-return assumption and the quasi-equilibrium assumption. That is, transition states are in local equilibrium with reactant molecules and do not return to the reaction region once they fall into the product region. The generalized TST defines the transition state as a hypersurface in phase space and no longer localizes the transition state to the saddle point of the reaction pathway.

3.2 Hydrazine and its derivatives

Hydrazine and its derivatives reduce Np(VI) to Np(V) in the PUREX process, while their reduction mechanisms remain unclear. Therefore, studies on their reduction mechanism of Np(VI) to Np(V) using theoretical approaches are needed. The reduction mechanism of Np(VI) by hydrazine was explored theoretically and three reaction pathways were obtained.³⁶ A free radical mechanism is thermodynamically and kinetically feasible based on the results for energy barrier and thermodynamic energy (Fig. 6). Np(VI) reduction by N_2H_4 accompanies dissociation of the N–H bond and formation of the $\text{O}_{\text{yl}}\text{–H}$ bond. To investigate the influence of substituent effects of hydrazine on the reducing ability of Np(VI) , the mechanisms of Np(VI) by CHON_2H_3 , $\text{HOC}_2\text{H}_4\text{N}_2\text{H}_3$ and $\text{CH}_3\text{–N}_2\text{H}_3$ were explored.⁴³ Their reduction mechanisms are similar, based on the potential energy profiles (PEPs), the order of the energy barrier follows $\text{HOC}_2\text{H}_4\text{N}_2\text{H}_3$ (15.15) < $\text{CH}_3\text{N}_2\text{H}_3$ (16.57) < CHON_2H_3 (25.73 kcal mol^{-1}) for the free radical ion mechanism, which agrees with the rates of reduction observed in the experiments.

Carbohydrazide ($\text{CO(N}_2\text{H}_3)_2$) appears significant for potential applications in reducing Np(VI) to Np(V) in high-burn-up SNF. The mechanism of Np(VI) reduction by $\text{CO(N}_2\text{H}_3)_2$ is shown in Fig. 7.⁴⁵ The calculated energy barrier agrees with the experimental activation energy.⁶⁶ The value of spin density (Fig. 8) shows that the reduction nature is outer-sphere electron transfer and hydrogen transfer.

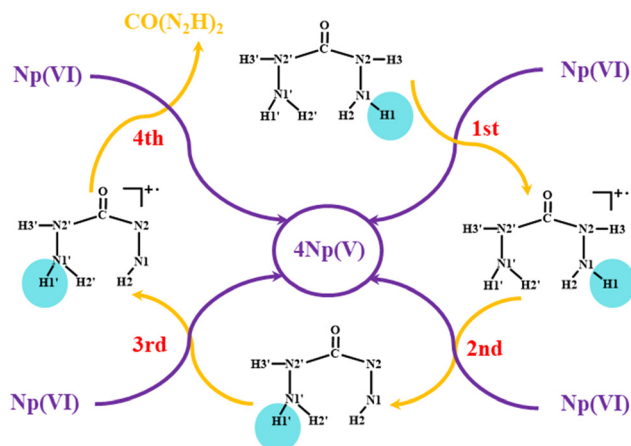


Fig. 7 Mechanism of Np(VI) reduction by $\text{CO(N}_2\text{H}_3)_2$. Reprinted with permission from ref. 45. Copyright 2023, American Chemical Society.

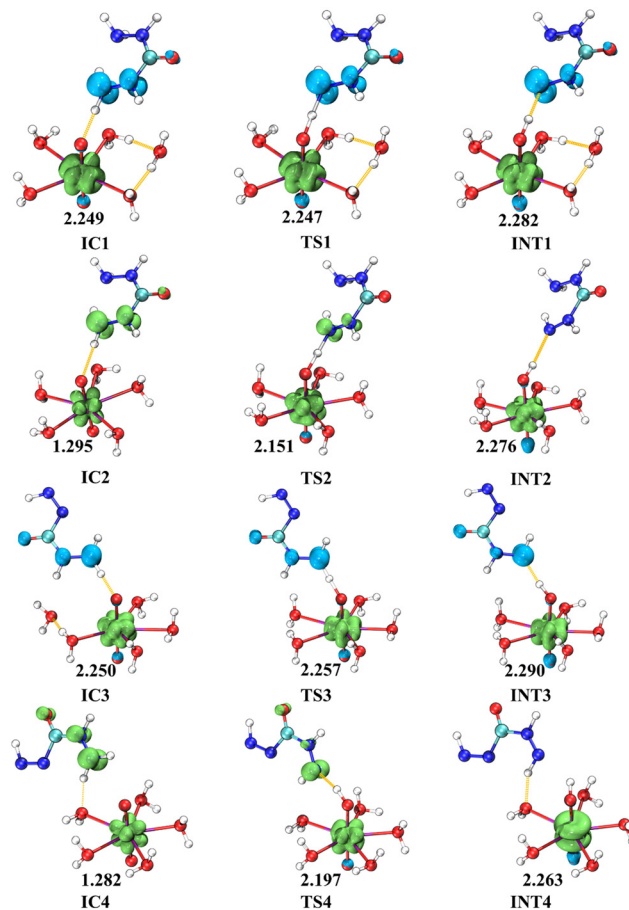


Fig. 8 Diagrams and value of spin density on the Np atom for the ICs, TSs and INTs of Np(VI) reduction by $\text{CO(N}_2\text{H}_3)_2$. Yellow dashed lines depict H-bonds. Reprinted with permission from ref. 45. Copyright 2023, American Chemical Society.

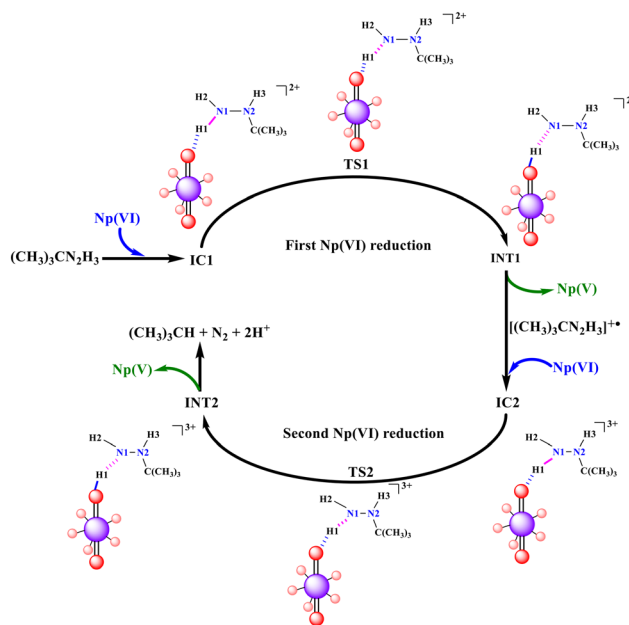


Fig. 9 Reduction processes of Np(VI) ions by $(\text{CH}_3)_3\text{CN}_2\text{H}_3$. Reprinted with permission from ref. 41. Copyright 2024, Springer Nature.



Tert-butylhydrazine ((CH₃)₃CN₂H₃) has exceptional selectivity in the reduction of Np(vi) in experiments, and the reduction reaction of Np(vi) with (CH₃)₃CN₂H₃ was explored using scalar-relativistic DFT.⁴¹ During Np(vi) reduction, the free radical ion [(CH₃)₃CN₂H₃]^{•+} intermediate can achieve another Np(vi) reduction (Fig. 9). The first Np(vi) reduction in pathway I and the second Np(vi) reduction in pathway III is kinetically optimal based on the energy barrier of the pathways. The reduction nature of the first and second Np(vi) reduction for the three pathways is outer-sphere electron transfer and hydrogen transfer, respectively. Moreover, the optimal reduction process from IC2 to TS2 is probably due to the existence of water-mediated proton transfer.

The short τ_{99} for the Np(vi) reduction by phenylhydrazine (C₆H₅N₂H₃) and hydrazinopropionitrile (NCCH₂N₂H₃) indicates that they quickly reduce Np(vi) to Np(v), which was confirmed by theoretical investigations.^{37,44} Pathways I and II are the free radical ion mechanism and pathway III is the free radical mechanism for Np(vi) reduction by C₆H₅N₂H₃ (Fig. 10(a)). The first Np(vi) reduction by C₆H₅N₂H₃ is the rate-determining step with an energy barrier of 12.89 kcal mol⁻¹. LMOs of the structures for pathway II reveal N–H bond dissociation and O_{y1}–H bond formation. The reduction kinetics of Np(vi) by NCCH₂N₂H₃ was studied. Each NCCH₂N₂H₃ molecule can reduce two Np(vi) ions and the second Np(vi) reduction is the rate-determining step, with an energy barrier of 14.36 kcal mol⁻¹. This is attributed to water-mediated proton transfer.³⁷ The fast reduction kinetics of Np(vi) by NCCH₂N₂H₃ and C₆H₅N₂H₃ is attributed to the σ - π hyperconjugation effect and delocalized π -orbital, respectively. The spin density of the Np atom indicates that the reduction of Np(vi) by NCCH₂N₂H₃ arises from outer-sphere electron transfer and hydrogen transfer.

The stable Np(vi) species are [Np^{VI}O₂(H₂O)₅]²⁺, [Np^{VI}O₂(NO₃)(H₂O)₃]⁺ and [Np^{VI}O₂(NO₃)₂(H₂O)] in 0–14 M HNO₃

solution.¹³⁵ The reduction mechanisms of [Np^{VI}O₂(NO₃)(H₂O)₃]⁺ by CH₃N₂H₃¹³⁶ and [Np^{VI}O₂(NO₃)₂(H₂O)] by N₂H₄, CHON₂H₃, and HOC₂H₄N₂H₃⁴² were also explored. The reduction of [Np^{VI}O₂(H₂O)₅]²⁺ by CH₃N₂H₃ has a lower energy barrier compared with that of [Np^{VI}O₂(NO₃)(H₂O)₃]⁺ due to the coordination of nitrate ions (Fig. 10(b)). To further explore the influence of nitrate ions on the energy barrier of Np(vi) reduction, the reactions for [Np^{VI}O₂(NO₃)₂(H₂O)] with N₂H₄, CHON₂H₃, and HOC₂H₄N₂H₃ were also investigated. [Np^{VI}O₂(NO₃)₂(H₂O)] reduction by the three reductants needs to overcome the higher energy barrier. Overall, the reduction of Np(vi) becomes more difficult when the nitrate ion coordinates Np(vi).

3.3 Hydroxylamine and its derivatives

Hydroxylamine derivatives are promising salt-free reductants that can reduce Np(vi) to Np(v) in nitric acid solution and their kinetic behavior of Np(vi) to Np(v) were studied experimentally. The reduction mechanism of Np(vi) by diethylhydroxylamine (DEHA) in aqueous solution was explored using scalar-relativistic DFT.³⁹ The whole reduction process of Np(vi) to Np(v) by DEHA occurs *via* four stages (Fig. 11). For the first stage, the Ho atom of DEHA contacts the O_{y1} atom of [Np^{VI}O₂(H₂O)₅]²⁺ and forms Np(v) and free radical (C₂H₅)₂NO. Subsequently, the H_C atom of free radical (C₂H₅)₂NO contacts the O_{y1} atom of [Np^{VI}O₂(H₂O)₅]²⁺ and obtains Np(v) and (C₂H₅)₂N(O)C₂H₄ for the second stage. Then, (C₂H₅)₂N(O)C₂H₄ is quickly hydrolyzed to mono-substituted C₂H₅NHOH, which can further reduce Np(vi) *via* two reaction pathways due to the different participating position of the H atom. The largest energy barrier among the four Np(vi) reduction processes is 12.23 kcal mol⁻¹ for the third stage of pathway I, which indicates that the reduction rate of Np(vi) by DEHA is rapid, agreeing with the

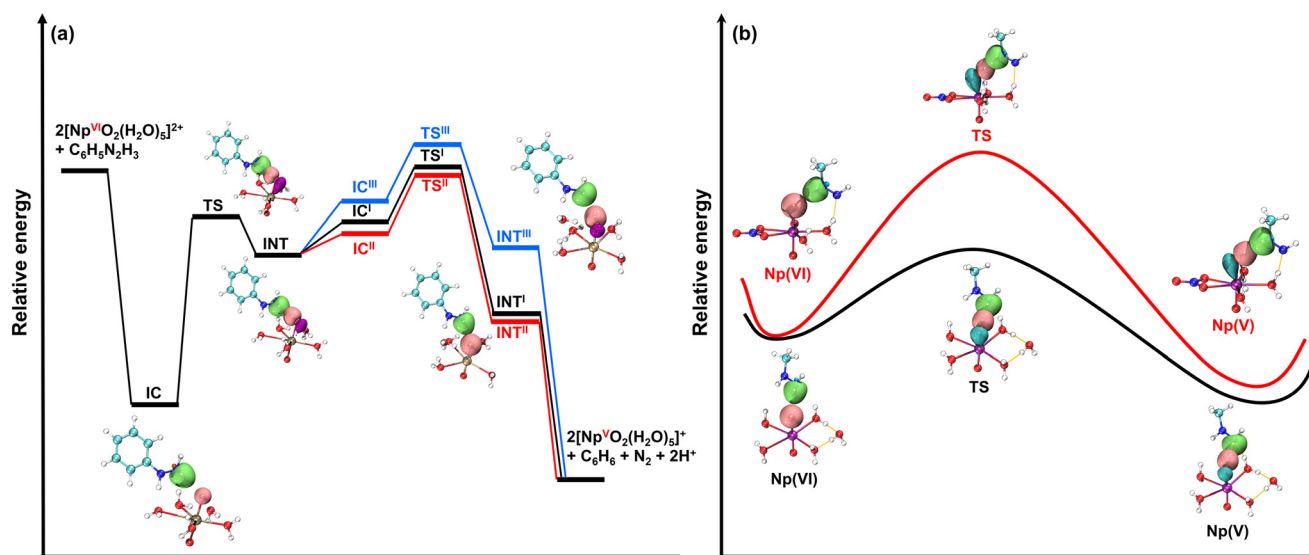


Fig. 10 Three PEPs for Np(vi) reduction with C₆H₅N₂H₃ and LMOs for the structures of pathway II (a). Reprinted with permission from ref. 45. Copyright 2023, American Chemical Society. PEPs for [Np^{VI}O₂(H₂O)₅]²⁺ and [Np^{VI}O₂(NO₃)(H₂O)₃]⁺ reduction by CH₃N₂H₃ (b).



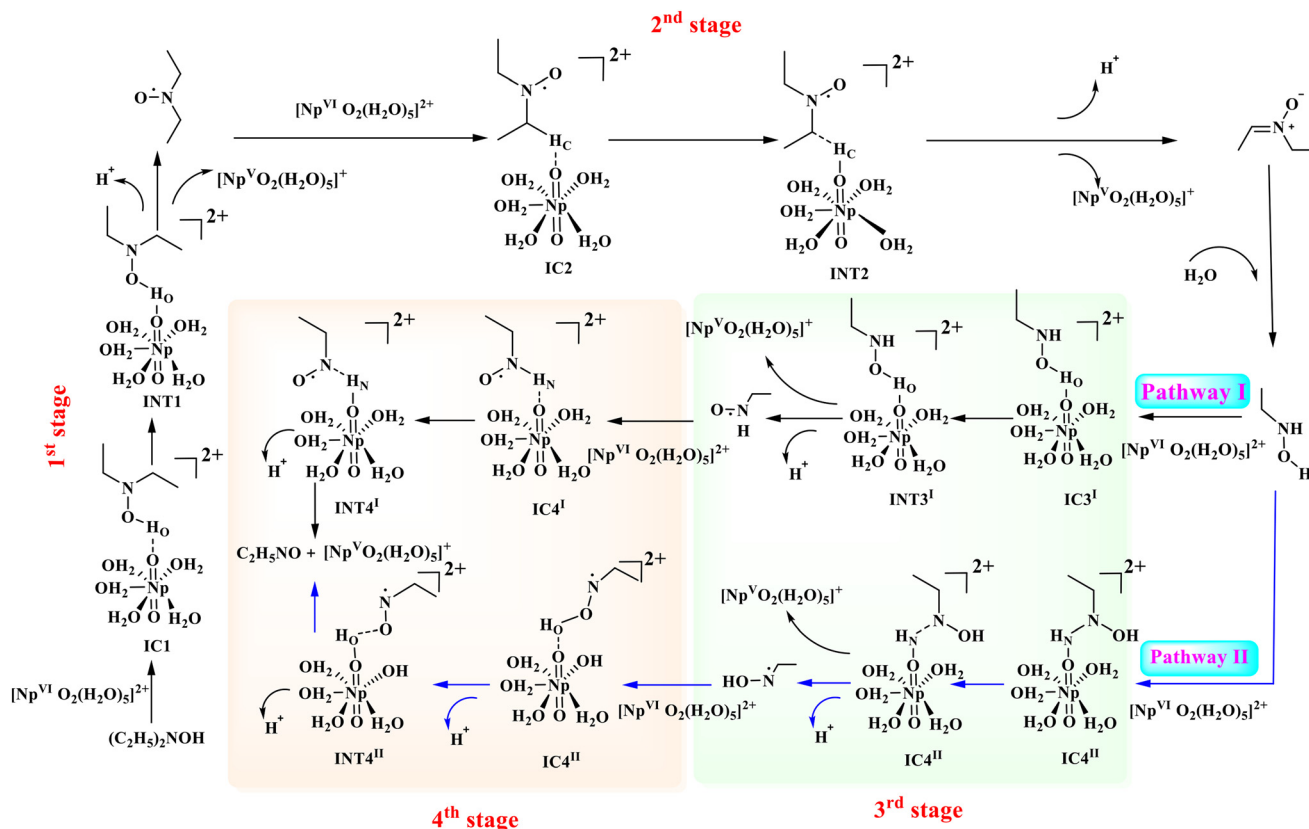


Fig. 11 Reaction of Np(vi) with DEHA. Reprinted with permission from ref. 39. Copyright 2024, The Royal Society of Chemistry.

experimental observations. The Np–O_{y1} bond distances and spin density on Np elucidate the reduction essence, including outersphere electron transfer or hydrogen atom transfer.

The reduction rates of Np(vi) by hydroxylamine (HA), monomethyl substituted *N*-methylhydroxylamine (MHA) and dimethyl substituted *N,N*-dimethylhydroxylamine (DMHA) are different. To evaluate the impact of methyl substitution on the reduction mechanism, the Np(vi) reduction reactions by HA, MHA and DMHA were theoretically investigated (Fig. 12).⁴⁰ The electron density results indicate that the nature of Np(vi) reduction by HA and the first Np(vi) reduction by methyl-

substituted MHA and DMHA are hydrogen atom transfer processes. Meanwhile, the second Np(vi) reduction is outersphere electron transfer, probably due to the electronic effect of methyl substitution. The rate-determining step for MHA and DMHA is the first Np(vi) reduction. Moreover, the energy barrier for the Np(vi) reduction by DMHA is lower than that by MHA, which reflects that the reaction rate of the former is faster than that of the latter. This work provided a kinetic insight into the effect of methyl substitution on the reduction of Np(vi) by hydroxylamine.

3.4 Other related reactions of Np(vi)

As far as we know, there are no theoretical investigations on the Np(vi) reduction by other reductants, except for hydrazine and hydroxylamine and their derivatives as discussed above. Some relevant Np(vi) reactions were explored by applying theoretical approaches, including the water exchange mechanism of neptunyl(vi) aqua ion^{38,130} and the hydrolysis mechanism of $[(\text{CH}_3)_3\text{Np}^{\text{VI}}\text{O}_2(\text{O}_2\text{C}-\text{CH}_3)_2]^-$.¹²⁹ The water exchange of the $[\text{Np}^{\text{VI}}\text{O}_2(\text{H}_2\text{O})_5]^{2+}$ complex can be achieved by dissociative-, associative-, and interchange mechanisms.³⁸ The reaction pathway and the corresponding structures are presented in Fig. 13(a and b). As for the dissociative water exchange mechanism, the water molecule in the second shell gradually deviates from the neptunyl ion in the structure of IC to INT. Concurrently, a water molecule coordinated to

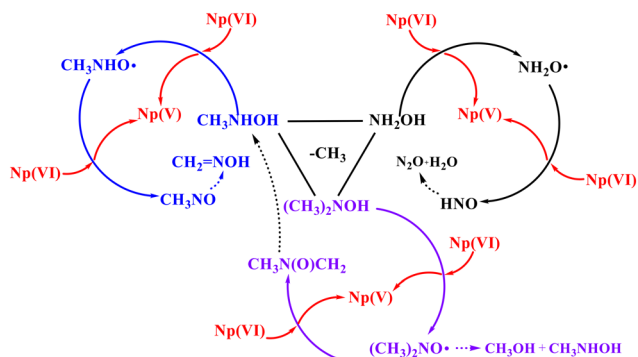


Fig. 12 Redox process of Np(vi) with HA, MHA and DMHA. Reprinted with permission from ref. 40. Copyright 2025, The Royal Society of Chemistry.



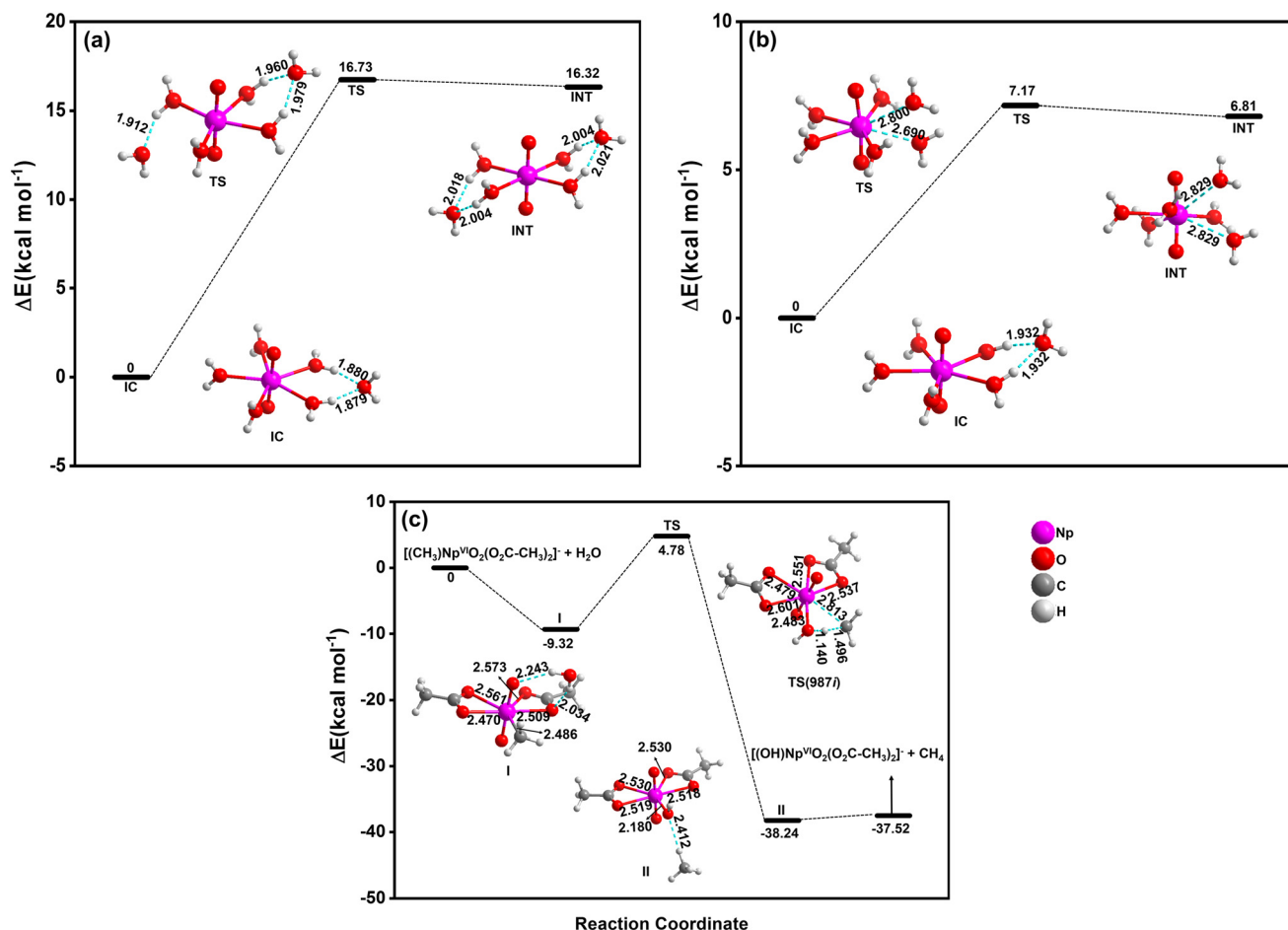


Fig. 13 Reaction pathway and the corresponding structures for the dissociative (a) and associative (b) water exchange mechanisms for the $[\text{Np}^{\text{VI}}\text{O}_2(\text{H}_2\text{O})_5]^{2+}$ complex. Reprinted with permission from ref. 38. Copyright 2004, American Chemical Society. PEP and the corresponding structures for the $[(\text{CH}_3)\text{Np}^{\text{VI}}\text{O}_2(\text{O}_2\text{C}-\text{CH}_3)_2]^-$ hydrolysis reactions (c). Reprinted with permission from ref. 130. Copyright 2018, American Chemical Society.

$\text{Np}(\text{VI})$ shifts out of the initial coordination sphere and forms two hydrogen bonds in the INT with four-coordinated $\text{Np}(\text{VI})$ (Fig. 13(a)). As for the associative water exchange mechanism in Fig. 13(b), the six-coordinated $\text{Np}(\text{VI})$ complex was formed. The energy barrier for the dissociative water mechanism is $16.73 \text{ kcal mol}^{-1}$, which is obviously larger than that for the associative one ($7.17 \text{ kcal mol}^{-1}$), indicating that the latter is more feasible kinetically. The hydrolysis reaction of $[(\text{CH}_3)\text{Np}^{\text{VI}}\text{O}_2(\text{O}_2\text{C}-\text{CH}_3)_2]^-$ was investigated in Fig. 13(c).¹²⁹ $[(\text{CH}_3)\text{Np}^{\text{VI}}\text{O}_2(\text{O}_2\text{C}-\text{CH}_3)_2]^-$ with one water molecule forms structure I by two hydrogen bonds, which is characterized as the exothermic reaction with the energy of $9.32 \text{ kcal mol}^{-1}$. Structure I overcomes the energy barrier of $14.10 \text{ kcal mol}^{-1}$ and forms structure II. In this process, the H atom from the H_2O molecule transfers to the CH_3^- fragment coordinated with $\text{Np}(\text{VI})$, which leads to the coordination of the hydroxyl group (OH^-) with $\text{Np}(\text{VI})$ and formation of the methane molecule. Essentially, the hydrolysis reaction of $[(\text{CH}_3)\text{Np}^{\text{VI}}\text{O}_2(\text{O}_2\text{C}-\text{CH}_3)_2]^-$ was achieved by the dissociation of the water molecule, leading to the exchange of ligands from the CH_3^- to OH^- group.

4 Conclusions and outlook

An important aim of the future advanced PUREX process is to create a single cycle reprocessing flowsheet that streamlines plant size, simplifies operations, and minimizes waste. The effective control of the valence state of Np is an important factor in the single cycle design. Although many salt-free reagents have been studied for the $\text{Np}(\text{VI})$ reduction, it is not easy to accomplish within the simplified flowsheets favored for an advanced reprocessing plant. Hydrazine and its derivatives, as well as *n*-butyraldehyde effectively reduce $\text{Np}(\text{VI})$ to $\text{Np}(\text{V})$ and do not reduce $\text{U}(\text{VI})$ and $\text{Pu}(\text{IV})$ under specific experimental conditions, which are suitable for the Np separation. Hydroxylamine and its derivatives, as well as isobutyraldehyde efficiently and quickly reduce $\text{Pu}(\text{IV})$ to $\text{Pu}(\text{III})$ and $\text{Np}(\text{VI})$ to $\text{Np}(\text{V})$, so they are applied to separate Np and Pu from U. The reduction ability of $\text{Np}(\text{VI})$ with different salt-free reagents are suitable for different flowsheets, but no single reductant has been applied for the $\text{Np}(\text{VI})$ reduction in a single cycle reprocessing flowsheet.



Therefore, it is essential to develop effective reductants for the U/Np and Np/Pu separation during SNF reprocessing.

Generally, experimental techniques for the investigation of Np(VI) reduction by salt-free reagents are difficult or dangerous; therefore, theoretical simulation seems to be an effective method to obtain reliable predictions. However, computational approach for accurate prediction of electronic structures and mechanisms of Np(VI) reduction remains a scientific challenge due to their relativistic and electron correlation effects. It requires intensive algorithms and available computational power for large actinide systems. Fortunately, with the rapid development of artificial intelligence, machine learning (ML) and the improvement of computational power, *ab initio* molecular dynamics (AIMD) and quantum mechanics/molecular mechanics (QM/MM) is applied for Np(VI) reduction by salt-free reagents. ML facilitates high-throughput screening and can obtain the key descriptors between molecular structures and the reduction rate of Np(VI). AIMD can investigate the effect of temperature, solvent and nitric acid concentration on the reduction rate of Np(VI) in the water/organic phase. The QM/MM approach can evaluate local electron transfer between Np(VI) and salt-free reagents. These methods synergistically uncover electronic structure, bonding evolution, thermodynamics and kinetics of Np(VI) reduction by salt-free reagents.

In summary, research on the reduction of Np(VI) by salt-free reagents still faces many challenges. (1) Finding stable efficient salt-free reagents could be achieved by conducting an in-depth study on the reduction mechanism and kinetics of Np(VI) by salt-free reagents, and accurately controlling the amount of salt-free reagents and reaction conditions such as temperature, pH value, reaction time, *etc.* for quantitative reduction of Np(VI). (2) Np(VI) may form various coordinated complexes in different solutions, which increases the complexity of the Np(VI) reduction process. (3) Np is radioactive: special radioactive analysis techniques and equipment are required. Although the research on the reduction of Np(VI) by salt-free reagents faces many challenges, with the rapid development of science and technology, (1) it is expected to develop more salt-free reagents with excellent performance to achieve accurate and efficient reduction of Np(VI); (2) with the help of advanced spectroscopy technology and actinide quantum chemical calculations, the reaction mechanism of Np(VI) reduction by salt-free reagents is studied in depth, revealing the electron transfer and chemical bond evolution; (3) the environmentally friendly process of Np(VI) reduction by salt-free reagents is optimized to reduce the amount of radioactive substances. In conclusion, the design and development of a new and environmentally friendly process for Np(VI) reduction by salt-free reagents is one of the future directions for the nuclear fuel cycle. In conclusion, the design, synthesis and development of salt-free reagents that specifically reduce Np(VI) is one of the future directions for the nuclear fuel cycle.

Data availability

The authors confirm that the data supporting the findings of this study are available within the article.

Author contributions

Xin Huang and Xiao-Bo Li: writing – original draft. Qun-Yan Wu: writing – original draft, supervision. Wei-Qun Shi: supervision.

Conflicts of interest

There are no conflicts to declare.

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References

- R. J. Taylor, I. S. Denniss and A. L. Wallwork, Neptunium control in an advanced Purex process, *Nucl. Energy*, 1997, **36**, 39–46.
- J. E. Birkett, M. J. Carrott, O. D. Fox, C. J. Jones, C. J. Maher, C. V. Roubé, R. J. Taylor and D. A. Woodhead, Recent developments in the Purex process for nuclear fuel reprocessing: Complexant based stripping for uranium/plutonium separation, *Chimia*, 2005, **59**, 898–904.
- Z. Yoshida, S. G. Johnson, T. Kimura and J. R. Krsul, Neptunium, in *The Chemistry of The Actinide and Transactinide Elements*, Springer, Dordrecht, 2008, pp. 699–812.
- R. C. Thompson, Neptunium-The neglected actinide: A review of the biological and environmental literature, *Radiat. Res.*, 1982, **90**, 1–32.
- M. A. Prelas, C. L. Weaver, M. L. Watermann, E. D. Lukosi, R. J. Schott and D. A. Wisniewski, A review of nuclear batteries, *Prog. Nucl. Energy*, 2014, **75**, 117–148.
- R. J. Taylor, C. R. Gregson, M. J. Carrott, C. Mason and M. J. Sarsfield, Progress towards the full recovery of neptunium in an advanced PUREX process, *Solvent Extr. Ion Exch.*, 2013, **31**, 442–462.
- H. Zhang, Z. Y. Liu, X. M. Zhou and L. Li, The complex reaction kinetics of neptunium including redox and extraction process in 30% TBP-nitric acid system, *J. Radioanal. Nucl. Chem.*, 2017, **312**, 173–180.
- P. K. Verma and P. K. Mohapatra, Fate of neptunium in nuclear fuel cycle streams: State-of-the art on separation strategies, *Radiochim. Acta*, 2022, **110**, 527–548.
- C. Gregson, C. Boxall, M. Carrott, S. Edwards, M. Sarsfield, R. Taylor and D. Woodhead, Neptunium (V) oxidation by nitrous acid in nitric acid, *Procedia Chem.*, 2012, **7**, 398–403.
- B. J. Mincher, M. Precek, S. P. Mezyk, G. Elias, L. R. Martin and A. Paulenova, The redox chemistry of neptunium in



- γ -irradiated aqueous nitric acid, *Radiochim. Acta*, 2013, **101**, 259–266.
- 11 H. Chen, R. J. Taylor, M. Jobson, D. A. Woodhead, C. Boxall, A. J. Masters and S. Edwards, Simulation of neptunium extraction in an advanced PUREX process—Model improvement, *Solvent Extr. Ion Exch.*, 2017, **35**, 1–18.
 - 12 V. A. Drake, Predicting the behavior of neptunium during nuclear-fuel reprocessing, *Nucl. Energy*, 1987, **26**, 253–258.
 - 13 S. Li, Y. Gao and Y. Ouyang, A review of the kinetics of oxidation and reduction of sale-free reagents in the U/Pu separation process, *Nucl. Tech.*, 2012, **35**, 929–935.
 - 14 R. J. Taylor, I. May, V. S. Koltunov, S. M. Baranov, V. I. Marchenko, E. A. Mezhov, V. G. Pastuschak, G. I. Zhuravleva and O. A. Savilova, Kinetic and solvent extraction studies of the selective reduction of Np(VI) by new salt-free reducing agents, *Radiochim. Acta*, 1998, **81**, 149–156.
 - 15 R. J. Taylor, I. May, A. L. Wallwork, I. S. Denniss, N. J. Hill, B. Y. Galkin, B. Y. Zilberman and Y. S. Fedorov, The applications of formo- and aceto-hydroxamic acids in nuclear fuel reprocessing, *J. Alloys Compd.*, 1998, **271**, 534–537.
 - 16 G. Uchiyama, S. Hotoku, S. Fujine and M. Maeda, Reduction of neptunium(VI) by butyraldehyde isomers in nitric acid solution, *Nucl. Technol.*, 1998, **122**, 222–227.
 - 17 G. Uchiyama, S. Fujine, S. Hotoku and M. Maeda, New separation process for neptunium, plutonium, and uranium using butyraldehydes as reductants in reprocessing, *Nucl. Technol.*, 1993, **102**, 341–352.
 - 18 V. S. Koltunov, Kinetics and mechanism of redox reactions of Np and Pu Ions with several organic reductants, *J. Nucl. Sci. Technol.*, 2002, **39**, 347–350.
 - 19 V. S. Koltunov and S. M. Baranov, Organic derivatives of hydrazine and hydroxylamine in future technology of spent nuclear fuel reprocessing, *Trans. Am. Nucl. Soc.*, 1993, **67**, 1.
 - 20 X. M. Zhou, G. A. Ye, H. Zhang, L. Li, F. X. Luo and Z. K. Meng, Chemical behavior of neptunium in the presence of technetium in nitric acid media, *Radiochim. Acta*, 2014, **102**, 111–116.
 - 21 S. T. Xiao, G. A. Ye, L. Li, X. C. Liu, H. Yang, H. B. Li and Z. W. Yuan, The investigation of the quantitative structure-activity relationships between the structure of hydrazine derivatives and the reduction of Np(VI), *J. Radioanal. Nucl. Chem.*, 2017, **311**, 1565–1575.
 - 22 V. S. Koltunov and S. M. Baranov, Kinetics and mechanism of Np and Pu reactions with organic derivatives of hydrazine, *Inorg. Chim. Acta*, 1987, **140**, 31–34.
 - 23 Y. Ban, T. Asakura and Y. Morita, Reduction kinetics of Pu(IV) and Np(VI) by N,N-dimethylhydrazine, and its potential application in nuclear fuel reprocessing, *J. Radioanal. Nucl. Chem.*, 2009, **279**, 423–429.
 - 24 H. Yang, H. Zhang, L. Li, X. H. Huang and R. T. Wang, Separation Np from Pu based on reduction-stripping by using hydrazine nitrate as reductant, *Nucl. Tech.*, 2016, **39**, 1–8.
 - 25 Z. Y. Liu, H. Zhang, R. T. Wang and Y. Z. Ning, Separation Np from Pu based on reduction-stripping by using methyl hydrazine as reductant, *Nucl. Tech.*, 2017, **40**, 1–7.
 - 26 V. S. Koltunov and M. F. Tikhonov, Reduction kinetics of actinoids with hydroxylamine. I. Neptunium (VI) reduction in nitric acid solutions, *Radiochemistry*, 1977, **19**, 611–619.
 - 27 V. S. Koltunov, R. J. Taylor, S. M. Baranov, E. A. Mezhov and I. May, The reduction of plutonium (IV) and neptunium (VI) ions by N,N-ethyl (hydroxyethyl) hydroxylamine in nitric acid, *Radiochim. Acta*, 1999, **86**, 115–122.
 - 28 A. Y. Zhang and Y. Liu, Hydroxylamine derivative in Purex process III. The kinetics of oxidation-reduction reaction between N,N-diethylhydroxylamine and neptunium(VI), *J. Radioanal. Nucl. Chem.*, 2000, **245**, 357–361.
 - 29 A. Y. Zhang, J. X. Hu, X. Y. Zhang and F. D. Wang, Hydroxylamine derivative in the Purex process – Part V. The single-stage reduction extraction and back extraction of neptunium with N,N-diethylhydroxylamine, *J. Radioanal. Nucl. Chem.*, 2002, **253**, 107–113.
 - 30 Z. Anyun, H. Jingxin, Z. Xianye and W. Fangding, Hydroxylamine derivative in Purex process, *J. Radioanal. Nucl. Chem.*, 2002, **252**, 565–571.
 - 31 V. S. Koltunov, R. J. Taylor, S. M. Baranov, E. A. Mezhov, V. G. Pastuschak and I. May, The reduction of plutonium and neptunium ions by acetaldoxime in nitric acid, *Radiochim. Acta*, 2000, **88**, 65–70.
 - 32 V. S. Koltunov, V. I. Marchenko, G. I. Zhuravleva and O. A. Savilova, Kinetics of redox reactions of U, Pu, and Np in TBP solutions: VII. kinetics of reduction of Pu(IV) and Np(VI) with butanal oxime in undiluted TBP, *Radiochemistry*, 2001, **43**, 334–337.
 - 33 V. S. Koltunov, E. A. Mezhov and S. M. Baranov, Kinetics of Np(VI) reduction with butanal oxime, *Radiochemistry*, 2001, **43**, 342–345.
 - 34 V. S. Koltunov, R. J. Taylor, S. M. Baranov, E. A. Mezhov, V. G. Pastuschak and G. V. Koltunov, Oxidation reactions of acetaldoxime in nitric acid, *J. Nucl. Sci. Technol.*, 2014, **39**, 878–881.
 - 35 R. J. Taylor and I. May, The reduction of actinide ions by hydroxamic acids, *Czech. J. Phys.*, 1999, **49**, 617–621.
 - 36 Z. P. Cheng, Q. Y. Wu, Y. H. Liu, J. H. Lan, C. Z. Wang, Z. F. Chai and W. Q. Shi, The redox mechanism of Np^{VI} with hydrazine: A DFT study, *RSC Adv.*, 2016, **6**, 109045–109053.
 - 37 X. B. Li, Q. Y. Wu, C. Z. Wang, J. H. Lan, M. Zhang, J. K. Gibson, Z. F. Chai and W. Q. Shi, Reduction of Np(VI) with hydrazinopropionitrile via water-mediated proton transfer, *Phys. Chem. Chem. Phys.*, 2022, **24**, 17782–17791.
 - 38 V. Vallet, T. Privalov, U. Wahlgren and I. Grenthe, The mechanism of water exchange in AmO₂(H₂O)₅²⁺ and in the isoelectronic UO₂(H₂O)₅⁺ and NpO₂(H₂O)₅²⁺ complexes as studied by quantum chemical methods, *J. Am. Chem. Soc.*, 2004, **126**, 7766–7767.
 - 39 X. Huang, X. B. Li, Q. Y. Wu, C. Z. Wang, J. H. Lan, H. Q. Wang and W. Q. Shi, Uncovering the reduction mechanism of Np(VI) with N, N-diethyl hydroxylamine: a scalar-relativistic DFT investigation, *Phys. Chem. Chem. Phys.*, 2024, **26**, 27395–27405.



- 40 X. Huang, Q.-Y. Wu, C.-Z. Wang, J.-H. Lan, H.-Q. Wang and W.-Q. Shi, Theoretical study on the kinetic behavior of Np(VI) reduction by hydroxylamine and its derivatives: substituent effect, *Phys. Chem. Chem. Phys.*, 2025, **27**, 6014–6023.
- 41 X.-B. Li, M. Zhang, C.-Z. Wang, J.-H. Lan and Q.-Y. Wu, Theoretical insights into the reduction reaction of Np(VI) by tert-butylhydrazine, *J. Radioanal. Nucl. Chem.*, 2025, **334**, 995–1006.
- 42 Z. P. Cheng, X. B. Li, Q. Y. Wu, Z. F. Chai and W. Q. Shi, Theoretical insights into the reduction mechanism of neptunyl nitrate by hydrazine derivatives, *Radiochim. Acta*, 2022, **110**, 471–480.
- 43 X. B. Li, Q. Y. Wu, C. Z. Wang, J. H. Lan, S. Y. Ning, Y. Z. Wei, Z. F. Chai and W. Q. Shi, Theoretical study on the reduction mechanism of Np(VI) by hydrazine derivatives, *J. Phys. Chem. A*, 2020, **124**, 3720–3729.
- 44 X. B. Li, Q. Y. Wu, C. Z. Wang, J. H. Lan, M. Zhang, Z. F. Chai and W. Q. Shi, Theoretical insights into the reduction mechanism of Np(VI) with phenylhydrazine, *J. Phys. Chem. A*, 2021, **125**, 6180–6188.
- 45 X. B. Li, Q. Y. Wu, C. Z. Wang, J. H. Lan, M. Zhang, Z. F. Chai and W. Q. Shi, Insights into the reduction mechanisms of Np(VI) to Np(V) by carbohydrazide, *J. Phys. Chem. A*, 2023, **127**, 4259–4268.
- 46 V. I. Marchenko, K. N. Dvoeglazov and V. I. Volk, Use of redox reagents for stabilization of Pu and Np valence forms in aqueous reprocessing of spent nuclear fuel: Chemical and technological aspects, *Radiochemistry*, 2009, **51**, 329–344.
- 47 V. S. Koltunov, G. I. Zhuravleva, V. I. Marchenko and M. F. Tikhonov, Features of kinetics and mechanism of some oxidation-reduction reactions of Np, Pu and U, *J. Inorg. Nucl. Chem.*, 1976, 189–196.
- 48 L. Venault, P. Moisy, P. Blanc and C. Madic, Kinetics of hydrazinium nitrate decomposition in nitric acid solutions under the effect of power ultrasound, *Ultrason. Sonochem.*, 2001, **8**, 359–366.
- 49 V. S. Koltunov, S. M. Baranov, E. A. Mezhev and V. G. Pastushchak, Kinetics of reactions of Np and Pu ions with hydrazine derivatives. 14. Hydrazinopropionitrile, *Radiokhimiya*, 2000, **42**, 117–120.
- 50 A. Y. Zhang, J. X. Hu, X. Y. Zhang and F. D. Wang, Progress of reaction kinetics between organic reductant and Np(VI), Pu(IV), *At. Energy Sci. Technol.*, 2001, **35**, 83–90.
- 51 D. G. Yin, X. Y. Zhang and J. X. Hu, Kinetic study of Np(VI) reduction with 1,1-dimethylhydrazine, *He Huaxue Yu Fangshe Huaxue*, 1997, **19**, 23–27.
- 52 W. Q. Shi, H. B. Tang, Y. X. Ye, J. X. Hu and X. Y. Zhang, Kinetic study of the Np (VI) reduction by tert-butyl hydrazine, *He Huaxue Yu Fangshe Huaxue*, 2002, **24**, 134–137.
- 53 X. Y. Zhang, Z. L. Huang, S. T. Xiao and J. X. Hu, Reduction of Np (VI) with 2-hydroxyethylhydrazine I. Studies on the reaction kinetics, *At. Energy Sci. Technol.*, 1997, **32**, 433–437.
- 54 K. N. Dvoeglazov, E. Y. Pavlyukevich and P. V. Mitrikas, Kinetics of Np(VI) reduction with diformylhydrazine in nitric acid, *Radiochemistry*, 2018, **60**, 581–587.
- 55 O. A. Zavalina, K. N. Dvoeglazov, E. Y. Pavlyukevich and S. I. Stepanov, Kinetics of Np(VI) reduction with carbohydrazide in nitric acid, *Radiochemistry*, 2017, **59**, 453–457.
- 56 V. S. Koltunov, E. A. Mezhev and S. M. Baranov, Kinetics of reaction between neptunium (VI) and 3, 3'-bis (diaziridinyl), *Radiochemistry*, 2000, **42**, 65–68.
- 57 V. I. Marchenko, V. N. Alekseenko and K. N. Dvoeglazov, Organic reductants of Pu and Np ions in wet technology for spent nuclear fuel reprocessing, *Radiochemistry*, 2015, **57**, 366–377.
- 58 V. I. Marchenko, V. S. Koltunov, O. A. Savilova and G. I. Zhuravleva, Redox kinetics of U, Pu, and Np in TBP solutions VI. Reactions of neptunium and plutonium with organic reducing agents determination of the final oxidation states and reaction rates, *Radiochemistry*, 2001, **43**, 276–283.
- 59 V. S. Koltunov, K. M. Frolov and Y. V. Isaev, Kinetics of redox reactions of U, Np, and Pu in TBP solutions: IX. Reduction of Np(VI) with dibenzylhydrazine, *Radiochemistry*, 2002, **44**, 121–126.
- 60 Z. X. Ye, Y. G. An, X. S. Tao, Y. D. Guang and H. J. Xin, Reduction of Np(VI) with monomethylhydrazine I. Studies on reaction kinetics, *At. Energy Sci. Technol.*, 1997, **31**, 193–198.
- 61 Z. X. Ye, Y. G. An, X. S. Tao, Y. D. Guang and H. J. Xin, Reduction of Np(VI) with monomethylhydrazine II. Studies on partition of U-Np in Purex process, *At. Energy Sci. Technol.*, 1997, **31**, 315–320.
- 62 X. Y. Zhang, Z. L. Huang, S. T. Xiao and J. X. Hu, Reduction of Np (VI) with 2-hydroxyethylhydrazine II. Studies on separation of U-Np in Purex process, *At. Energy Sci. Technol.*, 1999, **33**, 8–11.
- 63 D. G. Yin, X. Y. Zhang, J. X. Hu and S. T. Xiao, Separation of neptunium from uranium in contactor 1A with 1,1-dimethylhydrazine, *He Huaxue Yu Fangshe Huaxue*, 1998, **20**, 146–151.
- 64 C. Li, T. Yan, S. Xiao and W. Zheng, Rate determining step of reduction Np(VI) by organic reagents from 30%TBP/Kerosene to nitrate solution, *Radiochim. Acta*, 2020, **108**, 839–846.
- 65 V. I. Volk, V. I. Marchenko, K. N. Dvoeglazov, V. N. Alekseenko, S. I. Bychkov, E. Y. Pavlyukevich, V. V. Bondin and A. S. D'yachenko, Reduction of Pu (IV) and Np (VI) with carbohydrazide in nitric acid solution, *Radiochemistry*, 2012, **54**, 143–148.
- 66 V. P. Shilov and A. M. Fedoseev, Reduction of Np(VI) with carbohydrazide in a perchloric acid solution, *Radiochemistry*, 2019, **61**, 309–311.
- 67 B. J. Colston, G. R. Choppin and R. J. Taylor, A preliminary study of the reduction of Np(VI) by formohydroxamic acid using stopped-flow near-infrared spectrophotometry, *Radiochim. Acta*, 2000, **88**, 329–334.
- 68 V. I. Marchenko, G. I. Zhuravleva, K. N. Dvoeglazov and O. A. Savilova, Behaviors of plutonium and neptunium in



- nitric acid solutions containing hydrazine and technetium ions, *Theor. Found. Chem. Eng.*, 2008, **42**, 733–739.
- 69 V. S. Koltunov and S. M. Baranov, Organic derivatives of hydrazine and hydroxylamine and future reprocessing of irradiated nuclear-fuel, *Radiochemistry*, 1993, **35**, 622–630.
- 70 V. S. Koltunov, S. M. Baranov, T. P. Zharova and E. V. Abramina, Reaction-kinetics of Np and Pu ions with hydroxylamine derivatives. 2. Reaction of Np(VI) with N, N-dimethylhydroxylamine, *Radiochemistry*, 1993, **35**, 408–412.
- 71 V. S. Koltunov, S. M. Baranov and T. P. Zharova, Reaction-kinetics of Np and Pu ions with hydroxylamine derivatives. 1. Reaction of Np(VI) with isopropylhydroxylamine, *Radiochemistry*, 1993, **35**, 402–407.
- 72 V. S. Koltunov and S. M. Baranov, Kinetics of reactions of Np and Pu ions with hydroxylamine derivatives: VIII. Reduction of Np(VI) with N, N-dibutylhydroxylamine, *Radiochemistry*, 2000, **42**, 236–241.
- 73 S. M. Baranov, V. S. Koltunov, R. J. Taylor and I. May, Nuclear fuel reprocessing using hydrophilic substituted hydroxylamines, *U. S. Pat.*, US6444182B1, 2002.
- 74 V. S. Koltunov, S. M. Baranov, R. J. Taylor and T. P. Zharova, Kinetics of reactions of Np and Pu ions with hydroxylamine derivatives. V. Reaction of Np(VI) with N-methylhydroxylamine, *Radiochemistry*, 1993, **35**, 71–78.
- 75 A. Y. Zhang, H. X. Hu, Z. Y. Zhang and F. D. Wang, Hydroxylamine derivatives in Purex Process I. Study on the kinetics of redox reaction between N,N-diethylhydroxylamine and nitrous acid, *J. Radioanal. Nucl. Chem.*, 1998, **230**, 235–239.
- 76 A. Y. Zhang, J. X. Hu, X. Y. Zhang and F. D. Wang, Hydroxylamine derivative in Purex process. VI. Study on the partition of uranium/neptunium and uranium/plutonium with N,N-diethylhydroxylamine in the purification cycle of uranium contactor, *Solvent Extr. Ion Exch.*, 2001, **19**, 965–979.
- 77 Y. Ban, T. Asakura and Y. Morita, Reduction kinetics of Np(VI) by n-butyraldehyde in tributyl phosphate diluted with n-dodecane, *Radiochim. Acta*, 2004, **92**, 883–887.
- 78 A. Costagliola, L. Venault, A. Deroche, G. Garaix, J. Vermeulen, R. Omnee, F. Duval, G. Blain, J. Vandenborre, M. Fattahi-Vanani and N. Vigier, Radiation chemical behavior of aqueous butanal oxime solutions irradiated with helium ion beams, *Radiat. Phys. Chem.*, 2016, **119**, 186–193.
- 79 Z. W. Zhu, J. Y. He, Z. F. Zhang, V. Zhang, J. M. Zhu and W. F. Zhen, Uranium/plutonium and uranium/neptunium separation by the Purex process using hydroxyurea, *J. Radioanal. Nucl. Chem.*, 2004, **262**, 707–711.
- 80 T. H. Yan, W. F. Zheng, C. Zuo, L. Xian, Y. Zhang, X. Y. Bian, R. X. Li and Y. Di, The reduction of Np(VI) and Np(V) by tit dihydroxyurea and its application to the U/Np separation in the PUREX process, *Radiochim. Acta*, 2010, **98**, 35–38.
- 81 Y. Wang, S. Xiao, T. Lan, X. Liu, Y. Ouyang and H. Li, Kinetics of reaction between Np(VI) and hydroxysemicarbazide in nitric acid solution, *He Huaxue Yu Fangshe Huaxue*, 2018, **40**, 359–365.
- 82 D. Y. Chung and E. H. Lee, The reduction of Np(VI) by acetohydroxamic acid in nitric acid solution, *Bull. Korean Chem. Soc.*, 2005, **26**, 1692–1694.
- 83 C. Li, T. Yan, C. Zuo and W. Zheng, Kinetics of reductive stripping of Np(VI) by acetohydroxamic acid from 30% TBP/kerosene to nitrate medium using a high-speed stirred cell, *Radiochim. Acta*, 2015, **103**, 627–634.
- 84 B. S. Matteson, M. Precek and A. Paulenova, *Presented in part at the IOP Conference Series: Materials Science and Engineering*, San Francisco, CA, Jul 12–17, 2010.
- 85 L. F. Rao and G. R. Choppin, Kinetics and mechanism of the reduction of neptunium(VI) by methylsalicylic acids, *Radiochim. Acta*, 1991, **54**, 21–24.
- 86 Y. Z. Chen, B. M. Tan and Z. J. Lin, A kinetic-study of the reduction of Np(VI) with humic-acid, *Radiochim. Acta*, 1993, **62**, 199–201.
- 87 A. J. Zielen, J. C. Sullivan, D. Cohen and J. C. Hindman, A kinetic study of the reduction of neptunium (VI) by hydrogen peroxide, *J. Am. Chem. Soc.*, 1958, **80**, 5632–5635.
- 88 M. E. Thompson, K. L. Nash and J. C. Sullivan, Complexes of hydrogen-peroxide with dioxoactinide(VI) species in aqueous carbonate and bicarbonate media-formation of An(VI)-H₂O₂ Complexes, *Isr. J. Chem.*, 1985, **25**, 155–158.
- 89 D. T. Reed, D. G. Wygmans, S. B. Aase and J. E. Banaszak, Reduction of Np(VI) and Pu(VI) by organic chelating agents, *Radiochim. Acta*, 1998, **82**, 109–114.
- 90 J. H. Lan, W. Q. Shi, L. Y. Yuan, Y. L. Zhao, J. Li and Z. F. Chai, Trivalent actinide and lanthanide separations by tetradentate nitrogen ligands: A quantum chemistry study, *Inorg. Chem.*, 2011, **50**, 9230–9237.
- 91 C.-Z. Wang, J.-H. Lan, Y.-L. Zhao, Z.-F. Chai, Y.-Z. Wei and W.-Q. Shi, Density functional theory studies of UO₂²⁺ and NpO₂⁺ complexes with carbamoylmethylphosphine oxide ligands, *Inorg. Chem.*, 2013, **52**, 196–203.
- 92 Q.-Y. Wu, Y.-T. Song, L. Ji, C.-Z. Wang, Z.-F. Chai and W.-Q. Shi, Theoretically unraveling the separation of Am(III)/Eu(III): insights from mixed N,O-donor ligands with variations of central heterocyclic moieties, *Phys. Chem. Chem. Phys.*, 2017, **19**, 26969–26979.
- 93 X.-W. Chi, Q.-Y. Wu, C.-Z. Wang, J.-P. Yu, K. Liu, R.-A. Chi, Z.-F. Chai and W.-Q. Shi, A theoretical study of unsupported uranium–ruthenium bonds based on tripodal ligands, *Organometallics*, 2022, **41**, 1304–1313.
- 94 J.-H. Lan, W.-Q. Shi, L.-Y. Yuan, J. Li, Y.-L. Zhao and Z.-F. Chai, Recent advances in computational modeling and simulations on the An(III)/Ln(III) separation process, *Coord. Chem. Rev.*, 2012, **256**, 1406–1417.
- 95 Q.-Y. Wu, J.-H. Lan, C.-Z. Wang, C.-L. Xiao, Y.-L. Zhao, Y.-Z. Wei, Z.-F. Chai and W.-Q. Shi, Understanding the bonding nature of uranyl ion and functionalized graphene: A theoretical study, *J. Phys. Chem. A*, 2014, **118**, 2149–2158.
- 96 Q.-Y. Wu, C.-Z. Wang, J.-H. Lan, C.-L. Xiao, X.-K. Wang, Y.-L. Zhao, Z.-F. Chai and W.-Q. Shi, Theoretical investigation on multiple bonds in terminal actinide nitride complexes, *Inorg. Chem.*, 2014, **53**, 9607–9614.



- 97 Q.-Y. Wu, J.-H. Lan, C.-Z. Wang, Y.-L. Zhao, Z.-F. Chai and W.-Q. Shi, Terminal U \equiv E (E = N, P, As, Sb, and Bi) bonds in uranium complexes: A theoretical perspective, *J. Phys. Chem. A*, 2015, **119**, 922–930.
- 98 H. Wu, Q.-Y. Wu, C.-Z. Wang, J.-H. Lan, Z.-R. Liu, Z.-F. Chai and W.-Q. Shi, New insights into the selectivity of four 1,10-phenanthroline-derived ligands toward the separation of trivalent actinides and lanthanides: A DFT based comparison study, *Dalton Trans.*, 2016, **45**, 8107–8117.
- 99 Q. Y. Wu, J. H. Lan, C. Z. Wang, Z. P. Cheng, Z. F. Chai, J. K. Gibson and W. Q. Shi, Paving the way for the synthesis of a series of divalent actinide complexes: A theoretical perspective, *Dalton Trans.*, 2016, **45**, 3102–3110.
- 100 X.-W. Chi, Q.-Y. Wu, Q. Hao, J.-H. Lan, C.-Z. Wang, Q. Zhang, Z.-F. Chai and W.-Q. Shi, Theoretical study on unsupported uranium–metal bonding in uranium–group 8 complexes, *Organometallics*, 2018, **37**, 3678–3686.
- 101 X.-H. Kong, Q.-Y. Wu, C.-Z. Wang, J.-H. Lan, Z.-F. Chai, C.-M. Nie and W.-Q. Shi, Insight into the extraction mechanism of americium(III) over europium(III) with pyridylpyrazole: A relativistic quantum chemistry study, *J. Phys. Chem. A*, 2018, **122**, 4499–4507.
- 102 C.-Z. Wang, T. Bo, J.-H. Lan, Q.-Y. Wu, Z.-F. Chai, J. K. Gibson and W.-Q. Shi, Ultrastable actinide endohedral borospherenes, *Chem. Commun.*, 2018, **54**, 2248–2251.
- 103 Q.-Y. Wu, Z.-P. Cheng, J.-H. Lan, C.-Z. Wang, Z.-F. Chai, J. K. Gibson and W.-Q. Shi, Insight into the nature of M–C bonding in the lanthanide/actinide-biscarbene complexes: A theoretical perspective, *Dalton Trans.*, 2018, **47**, 12718–12725.
- 104 X.-W. Chi, Q.-Y. Wu, J.-H. Lan, C.-Z. Wang, Q. Zhang, Z.-F. Chai and W.-Q. Shi, A theoretical study on divalent heavier group 14 complexes as promising donor ligands for building uranium–metal bonds, *Organometallics*, 2019, **38**, 1963–1972.
- 105 C. Wang, Q.-Y. Wu, X.-H. Kong, C.-Z. Wang, J.-H. Lan, C.-M. Nie, Z.-F. Chai and W.-Q. Shi, Theoretical insights into the selective extraction of americium(III) over europium(III) with dithioamide-based ligands, *Inorg. Chem.*, 2019, **58**, 10047–10056.
- 106 Q.-Y. Wu, C.-Z. Wang, J.-H. Lan, Z.-F. Chai and W.-Q. Shi, Electronic structures and bonding of the actinide halides An(TREN^{TIPS})X (An = Th–Pu; X = F–I): A theoretical perspective, *Dalton Trans.*, 2020, **49**, 15895–15902.
- 107 Z.-R. Ye, Q.-Y. Wu, C.-Z. Wang, J.-H. Lan, Z.-F. Chai, H.-Q. Wang and W.-Q. Shi, Theoretical insights into the separation of Am(III)/Eu(III) by hydrophilic sulfonated ligands, *Inorg. Chem.*, 2021, **60**, 16409–16419.
- 108 X.-F. Luan, C.-Z. Wang, Q.-Y. Wu, J.-H. Lan, Z.-F. Chai, L.-S. Xia and W.-Q. Shi, Theoretical insights on improving amidoxime selectivity for potential uranium extraction from seawater, *J. Phys. Chem. A*, 2022, **126**, 406–415.
- 109 X.-P. Lei, Q.-Y. Wu, C.-Z. Wang, J.-H. Lan, Z.-F. Chai, C.-M. Nie and W.-Q. Shi, Theoretical insights into the substitution effect of phenanthroline derivatives on Am(III)/Eu(III) separation, *Inorg. Chem.*, 2023, **62**, 2705–2714.
- 110 C. T. Lee, W. T. Yang and R. G. Parr, Development of the Colle-Salvetti correlation-energy formula into a functional of the electron density, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1988, **37**, 785–789.
- 111 A. D. Becke, Density-functional thermochemistry. III. The role of exact exchange, *J. Chem. Phys.*, 1993, **98**, 5648–5652.
- 112 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, G. A. Petersson, H. Nakatsuji, X. Li, M. Caricato, A. V. Marenich, J. Bloino, B. G. Janesko, R. Gomperts, B. Mennucci, H. P. Hratchian, J. V. Ortiz, A. F. Izmaylov, J. L. Sonnenberg, F. Ding Williams, F. Lipparini, F. Egidi, J. Goings, B. Peng, A. Petrone, T. Henderson, D. Ranasinghe, V. G. Zakrzewski, J. Gao, N. Rega, G. Zheng, W. Liang, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, K. Throssell, J. A. Montgomery Jr, J. E. Peralta, F. Ogliaro, M. J. Bearpark, J. J. Heyd, E. N. Brothers, K. N. Kudin, V. N. Staroverov, T. A. Keith, R. Kobayashi, J. Normand, K. Raghavachari, A. P. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, J. M. Millam, M. Klene, C. Adamo, R. Cammi, J. W. Ochterski, R. L. Martin, K. Morokuma, O. Farkas, J. B. Foresman and D. J. Fox, *Gaussian 16 Rev. B.01.*, Gaussian Inc., Wallingford, CT, 2016.
- 113 W. Küchle, M. Dolg, H. Stoll and H. Preuss, Energy-adjusted pseudopotentials for the actinides-parameter sets and test calculations for thorium and thorium monoxide, *J. Chem. Phys.*, 1994, **100**, 7535–7542.
- 114 X. Cao, M. Dolg and H. Stoll, Valence basis sets for relativistic energy-consistent small-core actinide pseudopotentials, *J. Chem. Phys.*, 2003, **118**, 487–496.
- 115 X. Y. Cao and M. Dolg, Segmented contraction scheme for small-core actinide pseudopotential basis sets, *J. Mol. Struct.: THEOCHEM*, 2004, **673**, 203–209.
- 116 A. V. Marenich, C. J. Cramer and D. G. Truhlar, Universal Solvation Model Based on Solute Electron Density and on a Continuum Model of the Solvent Defined by the Bulk Dielectric Constant and Atomic Surface Tensions, *J. Phys. Chem. B*, 2009, **113**, 6378–6396.
- 117 M. Cossi, N. Rega, G. Scalmani and V. Barone, Energies, structures, and electronic properties of molecules in solution with the C-PCM solvation model, *J. Comput. Chem.*, 2003, **24**, 669–681.
- 118 V. Barone and M. Cossi, Quantum calculation of molecular energies and energy gradients in solution by a conductor solvent model, *J. Phys. Chem. A*, 1998, **102**, 1995–2001.
- 119 J. Cioslowski and S. T. Mixon, Covalent bond orders in the topological theory of atoms in molecules, *J. Am. Chem. Soc.*, 1991, **113**, 4142–4145.
- 120 R. F. W. Bader, Atoms in molecules, *Acc. Chem. Res.*, 1985, **18**, 9–15.
- 121 A. Savin, R. Nesper, S. Wengert and T. F. Fässler, ELF: The electron localization function, *Angew. Chem., Int. Ed. Engl.*, 1997, **36**, 1808–1832.
- 122 A. D. Becke and K. E. Edgecombe, A simple measure of electron localization in atomic and molecular systems, *J. Chem. Phys.*, 1990, **92**, 5397–5403.



- 123 B. Silvi and A. Savin, Classification of chemical bonds based on topological analysis of electron localization functions, *Nature*, 1994, **371**, 683–686.
- 124 J. Pipek and P. G. Mezey, A fast intrinsic localization procedure applicable for ab initio and semiempirical linear combination of atomic orbital wave functions, *J. Chem. Phys.*, 1989, **90**, 4916–4926.
- 125 T. Lu and F. W. Chen, Multiwfn: A multifunctional wavefunction analyzer, *J. Comput. Chem.*, 2012, **33**, 580–592.
- 126 T. Lu, A comprehensive electron wavefunction analysis toolbox for chemists, Multiwfn, *J. Chem. Phys.*, 2024, **161**, 082503.
- 127 L. Castro, A. Yahia and L. Maron, A DFT study of the reactivity of actinidocenes (U, Np and Pu) with pyridine and pyridine N-oxide derivatives, *Dalton Trans.*, 2010, **39**, 6682–6692.
- 128 M. Sundararajan, R. S. Assary, I. H. Hillier and D. J. Vaughan, The mechanism of the reduction of $[\text{AnO}_2]^{2+}$ (An = U, Np, Pu) in aqueous solution, and by Fe(II) containing proteins and mineral surfaces, probed by DFT calculations, *Dalton Trans.*, 2011, **40**, 11156–11163.
- 129 P. D. Dau, D. Rios, Y. Gong, M. C. Michelini, J. Marçalo, D. K. Shuh, M. Mogannam, M. J. Van Stipdonk, T. A. Corcovilos, J. K. Martens, G. Berden, J. Oomens, B. Redlich and J. K. Gibson, Synthesis and hydrolysis of uranyl, neptunyl, and plutonyl gas-phase complexes exhibiting discrete actinide–carbon bonds, *Organometallics*, 2016, **35**, 1228–1240.
- 130 F. P. Rotzinger, Quantum chemical study of the water exchange mechanism of the neptunyl(VI) and -(V) aqua ions, *Inorg. Chem.*, 2018, **57**, 2425–2431.
- 131 N. Shan, Q. Wang, H. Xiao, L. Wan and T. Gao, Ab initio molecular dynamics study, the reaction mechanism and topological properties of the microscopic interaction of PuO_2 and H_2O , *ChemistrySelect*, 2022, **7**, 202104589.
- 132 H. Pelzer and E. Wigner, Über die Geschwindigkeitskonstante von Austauschreaktionen, *Z. Phys. Chem., Abt. B*, 1932, **15**, 445–471.
- 133 H. Eyring, The activated complex and the absolute rate of chemical reactions, *Chem. Rev.*, 1935, **17**, 65–77.
- 134 M. G. Evans and M. Polanyi, Some applications of the transition state method to the calculation of reaction velocities, especially in solution, *Trans. Faraday Soc.*, 1935, **31**, 0875–0893.
- 135 P. Lindqvist-Reis, C. Apostolidis, O. Walter, R. Marsac, N. L. Banik, M. Y. Skripkin, J. Rothe and A. Morgenstern, Structure and spectroscopy of hydrated neptunyl(VI) nitrate complexes, *Dalton Trans.*, 2013, **42**, 15275–15279.
- 136 X. B. Li, Q. Y. Wu, C. Z. Wang, J. H. Lan, M. Zhang and W. Q. Shi, Theoretical perspectives on the reduction of Pu(IV) and Np(VI) by methylhydrazine in HNO_3 solution: Implications for Np/Pu separation, *Chin. Chem. Lett.*, 2024, **35**, 109359.

