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Synthesis and characterisation of Mo<sup>IV</sup>-oxo dithiolenes: formate dehydrogenase active site mimetics

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The active sites of metalloenzymes continue to inspire synthetic chemists to create structural models of the intricate structures seen in biology. As well as the fundamental intellectual challenge, this is driven by the potential societal impact that many of these systems offer. The formate dehydrogenases (FDH) hold such potential: formate is one of the key candidates as a hydrogen carrier for future energy transport. Efforts at mimicking the active site of FDH require the synthesis of (functionalised) molybdenum bis(dithiolene) complexes, principally featuring the Mo=O unit. In this review, we give an overview of the synthetic routes used to date in these efforts, along with key infrared and electrochemical data for the full collection of synthetic complexes reported to date.

## 1 Introduction

The industrial application of fossil fuels has provided humanity with an abundant source of materials and energy. Unfortunately, the limited supply and negative environmental effects require alternatives be found. Many exist, for example, wind or solar, but none match the properties of the currently available petroleum infrastructure as well as hydrogen. Hydrogen has a large gravimetric energy density, greater than conventional fossil fuels, but unfortunately a poor volumetric energy density.<sup>1</sup> The oxidation of hydrogen is water and therefore is not as environmentally damaging as the fossil fuel oxidation by-products.

Hydrogen can be generated by the electrolysis of water, but requires electrocatalysts but that themselves often rely on precious metals such as platinum and iridium. The storage and transportation of hydrogen is also an issue stemming from the risks of explosions. This risk is exacerbated as hydrogen, due to its size and reactivity, can penetrate metals, compromising their structural integrity.<sup>2</sup>

To circumvent these issues whilst retaining the benefits of hydrogen as a fuel, carrier molecules are an attractive solution. The ideal carrier molecule will be small, dense, air stable and non-toxic. Water, perhaps the molecule which meets these criteria best, is as already detailed unsuitable for this role. As a result, various other liquid organic hydrogen-carrying molecules (LOHCs) have been explored. The formate anion (HCOO<sup>-</sup>) has a hydrogen density of 53 g/L, 590 times that of gaseous hydrogen. This anion can also serve as a feedstock for other chemicals that

are largely reliant on fossil fuel production, such as methanol. The ability of formate to double as both a hydrogen carrier and a chemical building block makes it a prime candidate in the LOHC space, as is the potential to utilise both formic acid and formate, i.e. either liquid or solid carriers. Unfortunately, producing formate is a thermodynamically unfavourable process, so current methods must be improved towards catalytic means.<sup>3</sup>

The challenge in exploiting the potential of formate is in producing it cleanly and with high energy efficiency. Electroreduction of carbon dioxide is possible, but existing chemical routes sacrifice either energy efficiency (overpotential) or selectivity (dihydrogen release) in order to achieve this. Nature has already solved this conundrum: the formate dehydrogenase (FDH) enzymes can reversibly interconvert CO<sub>2</sub> and formate selectively, without requiring energy-intensive high overpotentials and using only earth-abundant elements.<sup>4</sup> Unfortunately, there are significant downsides to using the enzymes: they are extremely oxygen sensitive and their size limits currents both in solution and when immobilised on surfaces. Whilst this rules out using the FDH enzymes themselves as practical electrocatalysts in real-world devices, the possibility of creating functional formate-producing catalysts by mimicking these natural systems is extremely attractive.

Here, we examine enzymatic structures which provide the inspiration for this work, the challenges in the synthesis and the results to date in producing Mo<sup>IV</sup>=O dithiolene moieties. These structures are the core of the FDH enzyme and access to a variety of such systems is likely a prerequisite for accessing catalytically-competent biomimetics with useful overpotentials.

## 2 Molybdoenzyme active site structure

To date, more than 50 different types of molybdoenzymes have been identified, and they can be classified into four families ac-

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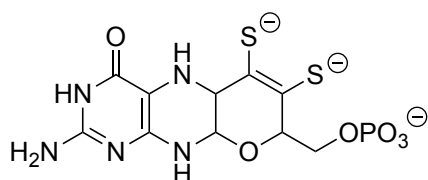
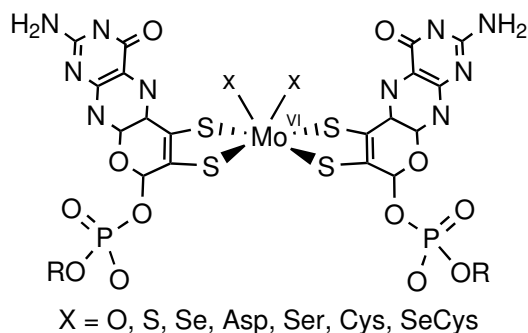


Fig. 1 The pyranopterin cofactor.

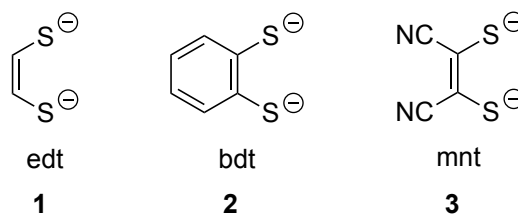
Fig. 2 The DMSO Reductase Active Site.<sup>15</sup>

cording to their protein sequences and the structure of the active site: xanthine oxidases, sulfite oxidases, dimethyl sulfoxide (DMSO) reductases and nitrogenases.<sup>5</sup> Other than the nitrogenases (which is not considered further in this review), these enzyme families exhibit significant active site similarity and a wide scope of roles.<sup>6–10</sup> In particular, they feature a single molybdenum atom in their active site ligated by at least one dithiolene, which is found as part of a pyranopterin cofactor (Figure 1).

## 2.1 Formate dehydrogenases

The formate dehydrogenases (FDH) belongs to the DMSO reductase family of enzymes. The FDH subgroup can be further divided into two categories: metal-dependent and metal-independent (NAD-dependent). In common with the rest of the DMSO reductase family, the FDH subgroup is identifiable by having two pyranopterin ligands leading to a square-based pyramidal geometry at the metal. One or two oxygen, sulfur, or selenium atoms complete the coordination sphere (Figure 2). There are also tungsten-containing FDH enzymes, which exhibit similar structures.<sup>11</sup> In total, three formate dehydrogenase enzymes have been characterised by X-ray crystallography: two molybdenum-based,<sup>12,13</sup> and one tungsten-based.<sup>14</sup>

As well as the intrinsic synthetic challenge and potential technological applications, synthetically modelling the FDH active site is attractive as it may help in unravelling the reaction mechanism employed by the enzyme. To date, five possible mechanisms have been suggested, and so far no consensus has been reached on key aspects.<sup>12,14,16–19</sup> Structures containing the Mo=O functionality are of particular interest due to the demonstration of catalysis in a synthetic mimic containing this metal structure.<sup>20</sup>

Fig. 3 The early ligands of Donahue and Holm.<sup>21</sup>

## 3 Synthesis of biomimetic Mo=O structures

The synthesis of biomimetic molybdenum dithiolenes began with Donahue and Holm in the 1990s with their succinctly named, edt (1), bdt (2) and mnt (3) ligands with a variety of analogous complexes based on changing the functionality at the terminal position (Figure 3).<sup>21</sup> Like most inorganic synthesis, the creation of these biomimetic complexes can be broken down into two parts: the prolignand synthesis, with an emphasis on dithiolene insertion, and subsequent complexation.

### 3.1 Dithiolene precursor synthesis

Many methods exist to achieve the introduction of dithiolenes into coordination sphere. However, this review focuses on those methods suitable for use for systems exceeding a few carbons, for example, Holm's landmark edt, bdt and mnt ligands, which are built from feedstock chemicals such as NaCN and CS<sub>2</sub>.<sup>21–23</sup> These methods focus on creation of suitable precursor molecular structures in the first stage. Given the general sensitivity of the ene-1,2-dithiolate moiety, most research groups have protected this unit until coordination to the metal is attempted, either by retaining in the form of a dithione derivative or using an ester protecting group. The key methods used are summarised in Figure 4.

In Method 1, a trithiocarbonate precursor was synthesised by reaction of a symmetrical alkyne with ethylene trithiocarbonate under reflux (this can also be performed using ultraviolet light).<sup>31,32</sup> The trithiocarbonate was then oxidised using mercury acetate and acetic acid to give the oxo-dithiocarbonate/dithione.<sup>24</sup>

In a variant of this approach, Ried and coworkers demonstrated that the addition of CS<sub>2</sub> and elemental sulfur to an alkyne generated a trithiocarbonate moiety; subsequent oxidation to the dithione was encouraged through addition of sulfuric acid (Method 2).<sup>25</sup>

Fontecave and coworkers developed a route to their novel protected dithiolene ligand by subjecting a bromovinyl triflate to a double palladium-catalysed cross-coupling reaction with two equivalents of HSCH<sub>2</sub>CH<sub>2</sub>CO<sub>2</sub>Et (Method 3).<sup>20,26</sup>

Hortmann and coworkers developed a method to achieve the dithione moiety starting from the addition of bromoacetyl analogues to potassium isopropyl xanthate where eventual cyclisation was encouraged by a strong acid (Method 4).<sup>27</sup>

Although a method that has not been employed to introduce dithiolene ligands in their prolignand form to molybdenum-oxo dithiolene complexes, Yves and coworkers shows that through



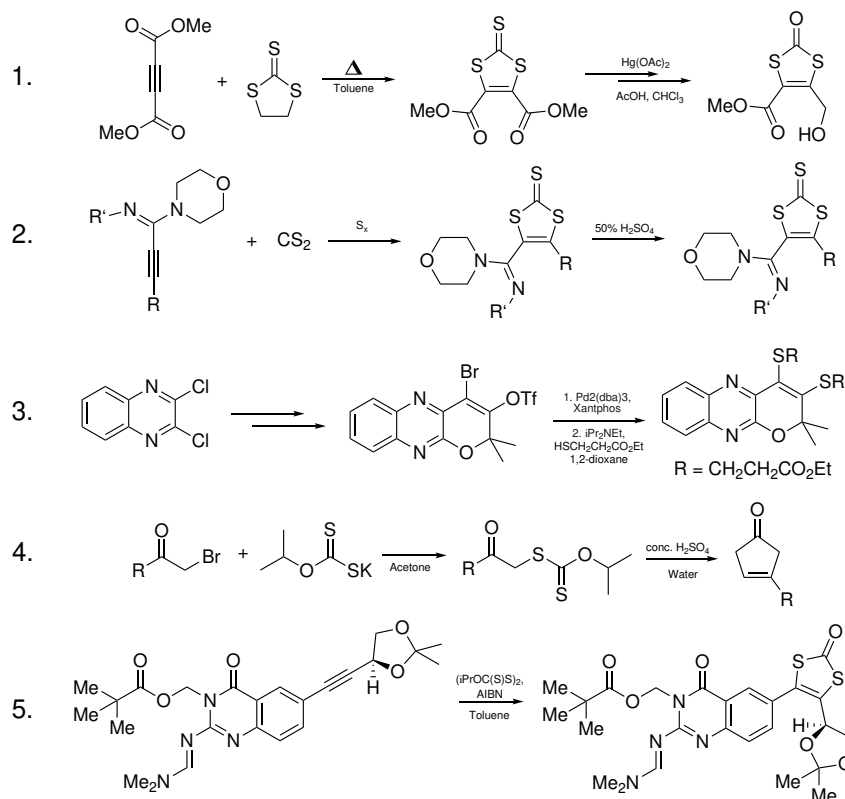
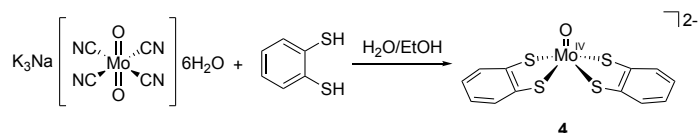


Fig. 4 The synthetic routes developed by various groups to access the dithione/dithiolene moiety. Methods 1,<sup>24</sup> 2,<sup>25</sup> 3,<sup>26</sup> 4,<sup>27</sup> 5.<sup>28–30</sup>

radical cyclised addition of diisopropyl xanthogen disulfide an alkyne functionalised moiety, the dithione group can be inserted.<sup>28</sup> This insertion method was utilised by Garner and Joule to develop a prolignand that after further synthesis gave a ligand of remarkable similarity to the pyranopterins ligand (Method 5).<sup>29,30</sup> It does not appear to have been complexed with molybdenum to generate the biomimetic species.

### 3.2 Complexation

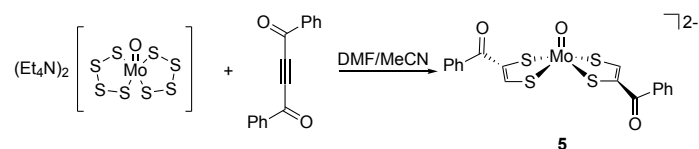
Complexing these ligands to molybdenum can follow a variety of routes from direct coordination to a molybdenum precursor (for example  $\text{MoO}_2(\text{CN})_4^{3-}$  or  $\text{MoOCl}_5^{2-}$ ) (Scheme 1) or by transmetalation from a nickel intermediary. Obtaining the latter complexes is relatively straightforward when the ligand scaffold is unfunctionalised, as in the case of edt (1) and mnt (3). Synthetic difficulty increases according to the complexity of the ligand.<sup>21</sup>



Scheme 1 The synthesis of  $[\text{MoO}(\text{bdt})_2]^{2-}$  (4) by Holm and Donahue.<sup>21</sup>

Some groups have found success implementing a post-coordination functionalisation approach in which the dithiolene scaffold is constructed first and then, by the addition of a substituted alkyne, additional functionality is deposited on to the com-

plex (Scheme 2).<sup>33–35</sup> However, this method is highly dependent on the activating nature of the alkyne.



Scheme 2 The post-coordination functionalisation synthesis of complex 5 employed by Sarakar and coworkers.<sup>34</sup>

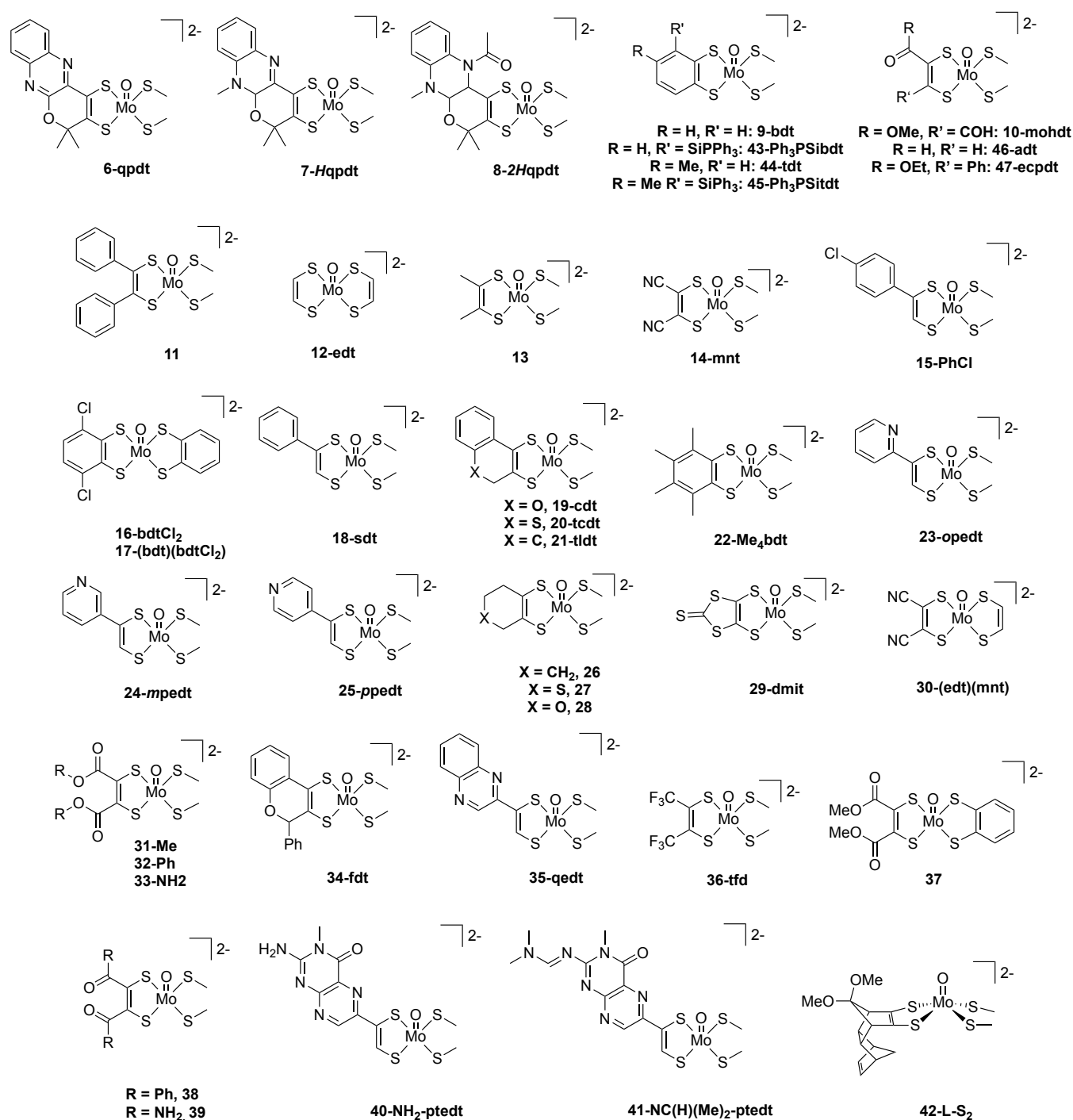
Figure 5 summarises all molybdenum-oxo bis(dithiolene) complexes synthesised to date. Most of these structures are limited in biomimetic features beyond the primary coordination sphere. Therefore, further synthesis directed toward individual components of the cofactor is likely to be productive in identification of the most important features for the natural system.

### 4 Spectral and electrochemical properties

Table 1 enumerates the key data for the  $\text{Mo}=\text{O}$  complexes where electrochemical data have been reported, along with the  $\text{Mo}=\text{O}$  stretching frequency where available. The infrared data are readily compared directly from the primary literature. Notably, almost all of the values lie around  $900\text{ cm}^{-1}$  independent of the (formal) oxidation state of the metal centre. This is unaffected by ligation (see below) and shows no pattern in the values.

However, for the electrochemical potential, a direct comparison is more challenging due to variations in solvent and internal reference in the original reports. Work by Addison and coworkers al-



Fig. 5 Mo<sup>IV</sup>O(dithiolene)<sub>2</sub> complexes reported to date.<sup>20,21,23,24,26,33–44</sup>



allows one to convert between the different referencing methods;<sup>45</sup> this adjustment is determined in acetonitrile, and therefore work in other solvents is likely subject to more uncertainty.

Table 1 illustrates how strongly ligation affects the electrochemistry of these molybdenum-oxo dithiolene complexes. The Mo<sup>IV</sup>/Mo<sup>V</sup> redox couple is well documented and is reversible for all complexes except **38**. In contrast, the Mo<sup>V</sup>/Mo<sup>VI</sup> redox couple is much more poorly documented, with a range of reported behaviours. It has been reported as quasi-reversible for six systems and reversible for seven. Looking at the ligands and comparing them with each other offers some commonalities in the scaffolds (Figure 5). The ligands supporting complexes with reversible behaviour are usually 'simple': alkyl and aryl substituents which lack adjacent functionality. In contrast, those with quasi-reversible behaviours feature more intricate ligands which likely undergo more significant rearrangement in solution on oxidation.

There are two general trends linking ligand structure and redox values. First, the alkyl/aryl-substituted systems that lack electron withdrawing groups and thus exhibit higher redox potentials. Secondly, the more reducing a ligand is, the more negative the redox potential, seen for example in the work of Fontecave and coworkers.<sup>20,26</sup>

## 5 Catalysis

Of the above biomimetic structures of the FDH active site only the complexes created by Fontecave and coworkers, **5–7**, have exhibited an ability to photocatalytically reduce carbon dioxide to formate, following the methodology presented by Ishanti and coworkers.<sup>20,26,47,48</sup> Compound **6** had the highest cumulative TONs followed by **8** and then **7**. Interestingly, the reverse is true for complexes with the most CO<sub>2</sub> derived reduction products, that being **8**, **7**, and **6**.

Several of the remaining molybdenum-oxo bis(dithiolene) complexes have had their oxygen-atom-transfer catalytic ability investigated. Oku and coworkers determined the influence sterics has on **tdt** and **bdt** coordinated systems through their novel silyl substituted dithiolene ligands.<sup>23</sup> Upon addition of Me<sub>3</sub>NO to both **9** and **43**, the bulkiness of the Ph<sub>3</sub>Si group significantly decreased the rate constant of MoO<sub>2</sub> formation. Schulzke and coworkers investigated the oxygen atom transfer (OAT) catalysis of their novel complex **10** by using the model OAT reaction developed by Berg and Holm in 1985 to investigate MoO<sub>2</sub>(LNS<sub>2</sub>)-type systems.<sup>24,49</sup> The reaction involves the OAT from dimethylsulphoxide (DMSO) to the catalyst and then to an acceptor such as triphenylphosphine (PPh<sub>3</sub>) to give triphenylphosphine oxide (PPh<sub>3</sub>O). The reaction proceeded very slowly with a maximum conversion of 93 % over approximately 60 hours. When trimethylamine N-oxide (Me<sub>3</sub>NO), a stronger oxidiser, was used instead of DMSO, there was a 37 % conversion of PPh<sub>3</sub>O in 15 hours. Through the use of UV/VIS spectroscopy, Schulzke reasoned that the inert Mo<sup>V</sup><sub>2</sub>O<sub>3</sub> species was forming, resulting in the slow times recorded, possibly encouraged by the presence of hydrogen bonding groups present on **10**.

Drawing mechanistic conclusions from the above catalytic experiments to aid future ligand design is difficult because of the lack of further testing of the complete catalogue of complexes.

It is important to bear in mind that other research groups have found more interest in characterising OAT ability of the oxidised analogue of their complexes (Mo<sup>VI</sup>O<sub>2</sub>) rather than the reduced species (Mo<sup>IV</sup>O).<sup>38,43,50–52</sup> One could assume that the lack of catalytic results implies that those systems were in fact not active; this is especially true with regards to carbon dioxide reduction. For OAT catalysis, some primary conclusions can be drawn: steric bulk impedes MoO<sub>2</sub> formation; hydrogen bonds encourage dimerisation through a bridging oxygen.

## 6 Summary and outlook

Over the roughly 40 years since the first recorded deliberate synthesis of a molybdenum-oxo bis(dithiolene) complex, progress in understanding these biomimetic systems has been slow, likely because of how synthetically challenging access to these systems is. However, this and similar areas of research will see increased interest in the coming years as the financial and social incentive to procure viable fossil fuel alternatives becomes more pressing. Early Mo<sup>IV</sup>O(dithiolene)<sub>2</sub> structures focused around the core dithiolene unit, lacking any real functionality on the vinyl carbons. Like with all things, advances in synthetic organic chemistry as a function of time have allowed for more interesting ligand scaffolds with recent designs resulting in the first series of complexes to show photocatalytic carbon dioxide reduction to formate. With a greater ease of procuring new novel proligands, and subsequent small-molecule mimics, one's ability to better understand the catalytic mechanism behind FDH grows with each passing year. However, incomplete characterisation, and a lack of standardisation of techniques between research groups, makes comparing catalytically-successful complexes difficult.

Very recent work on related tungsten systems suggests that there is significant untapped potential in this area.<sup>53</sup> Even unfunctionalised ligands may offer the potential for catalysis, and one may anticipate that as more complexes architectures become available, advances in electrocatalysis using Mo=O dithiolenes will increase markedly.

## 7 Author contributions

JAW conceptualised the research topic. IF carried out investigation to collect the literature data. IF and JAW wrote the paper jointly.

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Table 1 Spectroscopic and electrochemical properties of molybdenum-oxo bis(dithione) complexes. <sup>a</sup>The redox couples have been standardised to the Fc/Fc\* redox couple. <sup>b</sup>Quasi-reversible wave. <sup>c</sup>Irreversible wave.

Complex	M=O Stretch $\tilde{\nu}/\text{cm}^{-1}$	Supporting Electrolyte	Solvent	Reference	$E_{1/2}/\text{V}^a$	
					Mo <sup>IV</sup> /Mo <sup>V</sup>	Mo <sup>V</sup> /Mo <sup>VI</sup>
6-qpdt <sup>26</sup>	905	Bu <sub>4</sub> NOCl	MeCN	Ag/AgCl/KCl sat	−0.617	0.123 <sup>b</sup>
7-Hqpdt <sup>20</sup>	901	Bu <sub>4</sub> NOCl	MeCN	Ag/AgCl/KCl sat	−0.767	−0.027 <sup>b</sup>
8-2Hqpdt <sup>20</sup>	925	Bu <sub>4</sub> NOCl	MeCN	Ag/AgCl/KCl sat	−0.937	−0.127 <sup>b</sup>
9-bdt <sup>21</sup>	903	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	SCE	−0.77	0.18
10-mohbdt <sup>24</sup>	925	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	Fc/Fc*	−0.62	
11 <sup>46</sup>	903 <sup>b</sup>	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	SCE	−0.89	−0.13
12-edt <sup>21</sup>	917	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	SCE	−0.99	<sup>c</sup>
13 <sup>46</sup>	889	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	SCE	−1.04	−0.27
14-mnt <sup>21</sup>	928	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	SCE	0.1	<sup>c</sup>
16-bdtCl <sub>2</sub> <sup>42</sup>	910	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	SCE	−0.48	<sup>c</sup>
17-(bdt)(bdtCl <sub>2</sub> ) <sup>33</sup>	907	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	SCE	−0.61	
18-sdt <sup>37</sup>	879	Bu <sub>4</sub> NPF <sub>6</sub>	DMF	SCE	−0.86	<sup>c</sup>
19-cdt <sup>36</sup>	905	Bu <sub>4</sub> NPF <sub>6</sub>	—	Fc/Fc*	−0.72	
20-tcdt <sup>36</sup>	—	Bu <sub>4</sub> NPF <sub>6</sub>	—	Fc/Fc*	−0.87	
21-tldt <sup>36</sup>	—	Bu <sub>4</sub> NPF <sub>6</sub>	—	Fc/Fc*	−0.95	
22-Me <sub>4</sub> bdt <sup>21</sup>	—	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	SCE	−0.91	−0.54 <sup>b</sup>
23-opedt <sup>37</sup>	902	Bu <sub>4</sub> NPF <sub>6</sub>	DMF	SCE	−0.8	
24-mpedt <sup>37</sup>	882	Bu <sub>4</sub> NPF <sub>6</sub>	DMF	SCE	−0.77	<sup>c</sup>
25-ppedt <sup>37</sup>	900	Bu <sub>4</sub> NPF <sub>6</sub>	DMF	SCE	−0.73	
26 <sup>43</sup>	896	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	SCE	−0.94	−0.19
27 <sup>43</sup>	899	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	SCE	−0.95	−0.20
28 <sup>43</sup>	902	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	SCE	−1.08	−0.32
29-dmit <sup>40</sup>	930	Bu <sub>4</sub> NOCl	MeCN	SCE	−0.26	0.14 <sup>b</sup>
30-(edt)(mnt) <sup>21</sup>	919	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	SCE	−0.46	0.16 <sup>b</sup>
31 <sup>35</sup>	914	Bu <sub>4</sub> NOCl	MeCN	Ag/AgCl/KCl sat	−0.457	0.393 <sup>c</sup>
34-fdt <sup>41</sup>	1017	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	Fc/Fc*	−1.33	−0.04 <sup>c</sup>
35-qedt <sup>37</sup>	905	Bu <sub>4</sub> NPF <sub>6</sub>	DMF	SCE	−0.66	
36-tfd <sup>21</sup>	—	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	SCE	−0.2	0.26 <sup>b</sup>
37 <sup>33</sup>	908	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	SCE	−0.62	
38 <sup>34</sup>	950	Et <sub>4</sub> NOCl	MeCN	SCE	−1.22 <sup>c</sup>	<sup>c</sup>
40-NH <sub>2</sub> -ptedt <sup>37</sup>	886	Bu <sub>4</sub> NPF <sub>6</sub>	DMF	SCE	−0.73	
42-NC(H)(Me) <sub>2</sub> -ptedt <sup>37</sup>	890	Bu <sub>4</sub> NPF <sub>6</sub>	DMF	SCE	−0.71	
42-L-S <sub>2</sub> <sup>21</sup>	890	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	SCE	−1.09	−0.34
43-PPh <sub>3</sub> Sibdt <sup>23</sup>	915 <sup>b</sup>	Bu <sub>4</sub> NOCl	DMF	SCE	−0.79	
44-tdt <sup>23</sup>	900 <sup>b</sup>	Bu <sub>4</sub> NOCl	DMF	SCE	−0.83	
45-PPh <sub>3</sub> Sitdt <sup>23</sup>	906 <sup>b</sup>	Bu <sub>4</sub> NOCl	DMF	SCE	−0.84	0.14
46-adt <sup>44</sup>	936	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	SCE	−0.87	
47-ecpdt <sup>44</sup>	929	Bu <sub>4</sub> NPF <sub>6</sub>	MeCN	SCE	−0.52	



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No new data are presented in the manuscript: they are all taken from the primary literature

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