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CORRECTION

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Correction: Electrochemical nitrogen fixation and utilization: theories, advanced catalyst materials and system design

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Correction for 'Electrochemical nitrogen fixation and utilization: theories, advanced catalyst materials and system design' by Wenhan Guo *et al.*, *Chem. Soc. Rev.*, 2019, **48**, 5658–5716, https://doi.org/10.1039/C9CS00159J.

The authors regret that the pK_a of hydrazine was incorrectly given in the original article. The correct value should be 7.93. This means that eqn (8) and (9) and Fig. 3 were incorrect. The correct versions are shown below. This also applies to the sentence beginning "For hydrazine, the case is similar...", where the boundary should be given as 7.93.

$$N_2(g) + 5H^+ + 4e^- \rightleftharpoons N_2H_5^+(aq.)\ (pH < 7.93) \quad E^0 = -0.214\ V\ vs.\ RHE\ at\ pH = 0 \tag{8}$$

$$N_2(g) + 4H_2O(l) + 4e^- \rightleftharpoons N_2H_4(aq.) + 4OH^-(aq.)\ (pH \geq 7.93) \quad E^0 = -0.332\ V\ vs.\ RHE \tag{9}$$

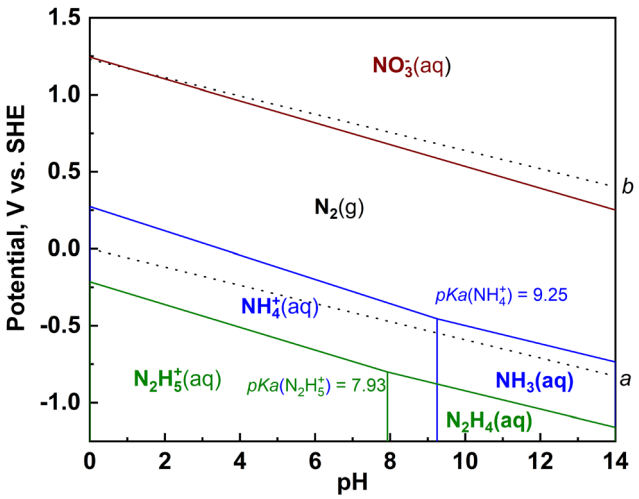

Fig. 3 Partial Pourbaix diagram of the N_2–H_2O system including N_2, NH_3, N_2H_4 and NO_3^-. Region between dotted lines a (HOR/HER) and b (OER/ORR) correspond to the condition of water stability. Reproduced with standard electrode potentials in water at 298.15 K from ref. 48.

The Royal Society of Chemistry apologises for these errors and any consequent inconvenience to authors and readers.

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