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High-accuracy theoretical rate coefficients for the reaction of H₂S with OH^{†‡}

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The rate coefficient for the reaction of hydrogen sulfide (H₂S) with hydroxyl radicals (OH), which plays an important role in producing sulfuric acid and sulfates in the Earth's atmosphere, was computed by combining the high-accuracy coupled-cluster HEAT-345Q(P) method for energetics, W. H. Miller's semi-classical transition state theory (SCTST) and two-dimensional *E_j*-resolved master equation analysis (2DME) for kinetics. The title reaction proceeds by H-abstraction through a well-skipping mechanism directly yielding products H₂O + SH. Tunneling effects are found to be important at the extremely low temperatures of the interstellar medium, but insignificant above 300 K. The rate coefficient *k₁*(*T*, *P*), calculated over the temperature range *T* = 10–2500 K and a wide pressure range, is found to be effectively independent of pressure for *T* = 200–2500 K, in which range *k₁*(*T*) can be represented by $k_1(T) = 1.44 \times 10^{-16} \times T^{1.674} \times \exp\left(\frac{+257}{T}\right) \text{ cm}^3 \text{ s}^{-1}$. The results agree within 30% with the experimental data available in the *T* range of 230 to 550 K. For the range of 200–500 K of atmospheric interest, we recommend the rate coefficient expression: $k_1(T) = 9.28 \times 10^{-16} \times T^{1.325} \times \exp\left(\frac{+286}{T}\right) \text{ cm}^3 \text{ s}^{-1}$, based on the *ab initio* results of this work and the available experimental determinations.

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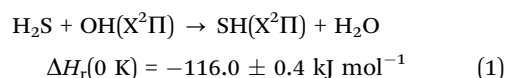
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Introduction

Hydrogen sulfide (H₂S) emitted into the air originates from both natural and anthropogenic sources.^{1–11} In Earth's nature, H₂S mainly stems from oceans and volcanic activities, while decomposition of organic wastes from animal farming and human activities also releases a significant amount of H₂S.^{1–11} In the atmosphere, H₂S is known to convert to sulfuric acid and sulfates, which cause acid rain and the formation of aerosols.^{1–11} In addition, hydrogen sulfide is a common component of natural gas and biomass.¹² Hydrogen sulfide is a significant molecule found in the interstellar medium (ISM)^{13,14} and is often associated with sulfur-rich environments; it is one of the most primitive sulfur-bearing molecules detected in the ISM, appearing in various

regions like cold dark clouds, diffuse clouds, and hot cores.^{13,14}

The lifetime of H₂S in the earth's atmosphere, principally controlled by its reaction with hydroxyl radicals (OH), is *ca.* 2.3 days using *k₁*(298 K) = (5 ± 0.5) × 10^{−12} cm³ s^{−1} and [OH] = 10⁶ cm^{−3}.^{15,16}



Due to its important role in the atmosphere, reaction (1) has been studied extensively.^{1–11} Experimental rate coefficients have been determined for the temperature range of 230–550 K and found to be pressure-independent below 1 atm.^{1–11} Rate constant measurements at room temperature are scattered from 3.5 × 10^{−12} to 5.5 × 10^{−12} cm³ s^{−1},^{1–11} with a recommended *k*(298 K) of (5 ± 0.5) × 10^{−12} cm³ s^{−1}.^{15,16} However, experimental *k₁*(*T*) data are not available for *T* < 200 K as well as for *T* > 600 K, *i.e.* for interesting environments such as the ISM and combustion systems.

On the theoretical side,^{17–20} the reaction mechanism has been characterized in detail using CCSD(T)/aug-cc-pV(5+d)Z single-point energy calculations based on the CCSD(T)/aug-cc-pV(Q+d)Z geometry;¹⁷ however, this work does not report any kinetics analysis. Earlier, *k₁*(*T*) rate constants were computed by direct dynamics (using a CVT/SCT approach) based on a

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potential energy surface constructed using the M06-2X/MG3S method.¹⁸ Quantum dynamics calculations based on a fitted potential energy surface using the UCCSD(T)-F12a/aug-cc-pVTZ level of theory have been recently reported;¹⁹ surprisingly, the calculated rate constants with these advanced techniques significantly differ from experimental values.^{19,20}

In this work, the mechanism of the title reaction is characterized using the high-accuracy HEAT-345Q(P) method,^{21–23} which yields an accuracy of *ca.* ± 0.5 kJ mol^{−1} for relative energies (*i.e.* reaction enthalpies and barriers) as compared to benchmark ATcT values in this case.²⁴ The rate coefficient is computed over wide temperature and pressure ranges using a combination of W. H. Miller's semi-classical transition state theory (SCTST)^{25–29} and a two-dimensional E_J -resolved master equation approach.^{30–32} The calculated $k_1(T)$ values are compared with available experimental data. For the low- T and high- T regimes where there are no experimental results, this work provides high-level theoretical rate coefficients for kinetics modeling of ISM and combustion processes, among others.

Methodologies

High-accuracy coupled-cluster calculations

Under C_s symmetry, the reactants OH($X^2\Pi$) + H₂S can correlate with the products SH($X^2\Pi$) + H₂O on both $^2A'$ and $^2A''$ electronic state potential energy surfaces (PES). All relevant stationary points on these PESs were optimized using the all-electron correlation coupled-cluster method including single, double, and non-iterative triple excitations (ae-CCSD(T))^{33–35} in conjunction with a basis set of cc-p(w)CVQZ³⁶ (*i.e.* cc-pwCVQZ for the sulfur atom, cc-pCVQZ for the oxygen atom, and cc-pVQZ for hydrogen atoms). Total energies were then computed using the composite HEAT-345Q(P) method,^{21–23} which includes the contributions of different terms of energy:

$$E_{\text{HEAT}} = E_{\text{SCF},\infty} + \Delta E_{\text{CCSD(T)},\infty} + \Delta E_{\text{T-(T)},\infty} + \Delta E_{\text{Q(P)-T}} + \Delta E_{\text{ZPE}} + \Delta E_{\text{DBOC}} + \Delta E_{\text{scalar}} + \Delta E_{\text{SO}} \quad (2)$$

where $E_{\text{SCF},\infty}$ is the SCF electronic energy extrapolated to the complete basis set limit (CBS) using aug-cc-p(w)CVXZ (where $X = 3, 4, 5$) basis sets; $\Delta E_{\text{CCSD(T)},\infty}$ is the electron correlation energy calculated using the CCSD(T) method and extrapolated to the CBS limit using the aug-cc-p(w)CVXZ (where $X = 4, 5$) basis set; $\Delta E_{\text{T-(T)},\infty}$ is the full triple excitation correction at the CBS limit; $\Delta E_{\text{Q(P)-T}}$ is the Q(P) correction calculated as the difference between $E[\text{fc-CCSDTQ(P)/cc-pVDZ}]$ and $E[\text{fc-CCSDT/cc-pVDZ}]$; ΔE_{DBOC} is the diagonal Bohr–Oppenheimer correction; ΔE_{scalar} is the scalar relativity effect; ΔE_{SO} is the spin–orbit correction, which includes electronic–rotational interactions, yielding -38.2 cm^{−1} for OH and -170.3 cm^{−1} for SH;³⁷ ΔE_{ZPE} is the anharmonic zero-point vibrational energy, which is computed using second-order vibrational perturbation theory (VPT2).³⁸ We used harmonic ZPEs calculated at the ae-CCSD(T)/aug-cc-p(w)CVQZ level of theory and anharmonic

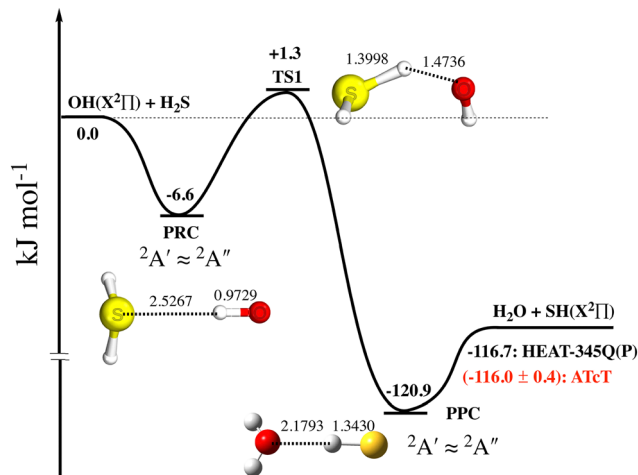


Fig. 1 Unscaled reaction energy profile of OH($X^2\Pi$) + H₂S → H₂O + SH($X^2\Pi$) constructed using the HEAT-345Q(P) method (see text). The benchmark ATcT value²⁴ of the 0 K reaction enthalpy (in red) is included for comparison.

constants obtained from ae-CCSD(T)/aug-cc-p(w)CVTZ calculations to compute anharmonic ZPEs.

There is a hindered internal rotation (HIR) with a low vibrational frequency of 182 cm^{−1} in the H-abstraction transition structure (TS1, see Fig. 1), corresponding to the rotation of the H-atom of the HO group around the S–O axis. This HIR is assumed to be separable from the other vibrations, and it is independently treated as a 1D-HIR. A torsional PES was constructed using the ae-CCSD(T)/aug-cc-p(w)VQZ level of theory, and the 1D-Schrodinger equation was solved to obtain a vector of eigenvalues for this 1D-HIR using the LAMM code of the Multiwell software package (see Fig. S1 and S2 in the ESI†).³⁹ The resulting, directly counted torsional quantum states of the 1D-HIR were combined with the normal vibration states to obtain the overall quantum density and sum of “vibration” states for TS1.

The CFOUR quantum chemical program was used for all CCSD(T) and CCSDT calculations.⁴⁰ The CCSDTQ(P) calculations were carried out using the MRCC program,⁴¹ which is interfaced with CFOUR.

Two-dimensional E_J -resolved master equation calculations

As shown in Fig. 1, the reaction of OH with H₂S proceeds *via* a pre-reaction intermediate (PRC) before yielding products. The PRC, once formed in a specific E_J state through the variational entrance transition state TS0 (not shown in Fig. 1), can either re-dissociate back into the reactants through TS0, or react to form products through TS1, or lose/gain energy (E) and/or angular momentum (J) by collisions with the bath gas M. Therefore, the title reaction rate is, in principle, affected by pressure and a master equation needs to be solved to obtain rate constants, which depend on both temperature and pressure. The two-dimensional E_J -resolved master equation that describes the time evolution of the E_J -states of the PRC by its competing unimolecular reactions back to reactants and forward to products and by collisional E_J -transfer processes



(see Fig. 1)—each with E, J -dependent rates—can be expressed as:^{30–32,42–46}

$$\begin{aligned} \frac{\partial C_{\text{PRC}}(E_i, J_i)}{\partial t} = & - (k_{-0}(E_i, J_i) + k_1(E_i, J_i)) \cdot C_{\text{PRC}}(E_i, J_i) \\ & - \omega_{\text{LJ}}[\text{M}] \cdot C_{\text{PRC}}(E_i, J_i) \\ & + \sum_{J_k=0}^{J_{\text{max}}} \int_{E_k=0}^{E_{\text{max}}} \omega_{\text{LJ}}[\text{M}] \cdot P(E_i, J_i | E_k, J_k) \\ & \times C_{\text{PRC}}(E_k, J_k) dE_k + k_{0,\infty}(T) \\ & \times F_{\text{CA}}(E_i, J_i) [\text{OH}] [\text{H}_2\text{S}] \end{aligned} \quad (3)$$

where $C_{\text{PRC}}(E_i, J_i)$ is the population of the energized PRC[†] in a specific state, (E_i, J_i) ; ω_{LJ} is the Lennard-Jones collisional rate (in $\text{cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$) and $[\text{M}]$ (in molecules cm^{-3}) is the concentration of the bath gas; $P(E_i, J_i | E_k, J_k)$ is the transfer probability upon collision with M from an initial state, (E_k, J_k) , to a final state, (E_i, J_i) ; $k_{-0}(E_i, J_i)$ is the microcanonical variational rate constant for the barrier-less re-dissociation of an energized PRC[†] back to initial reactants OH and H₂S (see eqn (4)); $k_1(E_i, J_i)$ is the microcanonical rate constant for the H-abstraction step, computed using W. H. Miller's semi-classical transition state theory (SCTST)^{25–29,47,48} (see eqn (5)). It should be mentioned that multi-dimensional tunneling through the TS1 barrier is automatically accounted in the SCTST by the anharmonicity of the fully coupled vibration frequencies and the anharmonic ZPE. $k_{0,\infty}(T)$ is the bimolecular rate constant for the capture of OH by H₂S to form PRC[†] at the high-pressure limit (HPL) leading to 100% thermalized PRC (see eqn (6)) and $F_{\text{CA}}(E_i, J_i)$ is the nascent energy distribution of the formation of the resulting PRC[†] (see eqn (7)).

The variational transition state TS0 is characterized using variational RRKM theory^{49,50} (data provided in Table S1 in the ESI[†]), *i.e.* the $k_{-0}(E_i, J_i)$ values for PRC re-dissociation are computed by minimizing the chemical flux through the dividing surface:

$$k_{-0}(E, J) = \frac{\sigma_0}{h} \cdot \frac{\text{Min}(G_0(E, J))}{\rho_{\text{PRC}}(E, J)} \quad (4)$$

$$k_1(E, J) = \frac{\sigma_1}{h} \cdot \frac{G_1(E, J)}{\rho_{\text{PRC}}(E, J)} \quad (5)$$

$$k_{0,\infty}(T) = \frac{\sigma_{\text{bi}}}{h} \cdot \frac{Q_{\text{e}}^{\#} Q_{\text{t}}^{\#}}{Q_{\text{OH}} \cdot Q_{\text{H}_2\text{S}}} \cdot \sum_{J=0}^{J_{\text{max}}} (2J+1) \quad (6)$$

$$\begin{aligned} F_{\text{CA}}(E, J) = & \frac{(2J+1) \cdot k_{-0}(E, J) \cdot \rho_{\text{PRC}}(E, J) \cdot \exp\left(-\frac{E}{RT}\right)}{\sum_{J=0}^{J_{\text{max}}} (2J+1) \int_{E=0}^{E_{\text{max}}} k_{-0}(E, J) \cdot \rho_{\text{PRC}}(E, J) \cdot \exp\left(-\frac{E}{RT}\right) dE} \\ & \times \int_{E=0}^{E_{\text{max}}} \text{Min}(G_0(E, J)) \cdot \exp\left(-\frac{E}{RT}\right) dE \end{aligned} \quad (7)$$

Here h is Planck's constant. σ_i is the reaction path degeneracy. $Q_{\text{e}}^{\#}$ and $Q_{\text{t}}^{\#}$ are the electronic and translational partition functions for TS0, respectively. Q_{OH} is the complete partition function of the OH radical including electronic-rotation coupling interactions.³⁷ $Q_{\text{H}_2\text{S}}$ is the complete partition function of H₂S. E_{max} and J_{max} are the maximum internal energy and total angular momentum, respectively; $E_{\text{max}} = 70\,000 \text{ cm}^{-1}$ and $J_{\text{max}} = 300$ were chosen to ensure that the calculated rate constants converge (better than 1%) for a wide temperature range of 10 K to 2500 K. $\rho_{\text{PRC}}(E, J)$ is the density of rovibrational states for PRC, while $G_1(E, J)$ and $G_0(E, J)$ are the sums of rovibrational states for TS1 and TS0 at the given (E, J) , respectively. Vibrational $\rho_{\text{PRC}}(E_{\text{vib}})$ and $G_{\text{TS1}}(E_{\text{vib}})$ are computed using the BDENS and SCTST codes of the Multiwell program.³⁹ Assuming that all stationary points are symmetric top, the rovibrational states can be computed using a J -shifting approximation.^{51,52}

$$\rho_{\text{PRC}}(E, J) = \sum_{K=-J}^{K=+J} \rho_{\text{vib}}^{\text{PRC}}(E - E_{\text{rot}}^{\text{PRC}}(J, K)) \quad (8)$$

$$G_{\text{TS1}}(E, J) = \sum_{K=-J}^{K=+J} G_{\text{vib}}^{\text{TS1}}(E - E_{\text{rot}}^{\text{TS1}}(J, K)) \quad (9)$$

$$E_{\text{rot}}(J, K) = \langle B \rangle \cdot J \cdot (J+1) + (A - \langle B \rangle) \cdot K^2 \quad \text{with } -J \leq K \leq +J \quad (10)$$

All parameters including the energy bin, angular momentum bin, collisional parameters between PRC and the bath gas (air), and other data needed for the master equation simulations can be found in Table S2 (see the ESI[†]).

The rate coefficient $k_1(T, P)$ can be obtained by solving the master equation (eqn (3)) using two different (deterministic) methods, which necessarily yield the same result. In the first method, proposed by Barker,⁵³ the yield of products Y_{HS} as opposed to that of the regenerated reactants ($= 1 - Y_{\text{HS}}$) is determined by solving eqn (3) without the source term, but starting from an initial energized PRC[†] population $C_{\text{PRC}}(E, J)(t=0)$ taken equal to the nascent PRC formation distribution $F_{\text{CA}}(E, J)$ in eqn (7), which then evolves in time by re-dissociation to reactants and reaction to form the products HS + H₂O in competition with the collisional energy transfer processes until the PRC population C_{PRC} vanishes.³⁷ The rate constant $k_1(T, P)$ follows from eqn (11):

$$k_1(T, P) = Y_{\text{SH}} \times k_{0,\infty}(T), \text{ cm}^3 \text{ s}^{-1} \quad (11)$$

In the second method, suggested by Pilling,^{54,55} the ME simulation is started with the initial reactants OH + H₂S, but with a very large excess concentration of H₂S (*i.e.* $[\text{H}_2\text{S}]_0 \gg [\text{OH}]_0$) to simplify the bimolecular reaction kinetics to a (pseudo)first-order decay of OH. The ME equation (eqn (3)) is solved iteratively to find the lowest 1st eigenvalue (λ_1) using the



ARPACK numerical software library.⁵⁶ The rate constant is obtained using eqn (12):

$$k_1(T, P) = \frac{|\lambda_1|}{[\text{H}_2\text{S}]_0}, \text{ cm}^3 \text{ s}^{-1} \quad (12)$$

In this work, we used both methods, with identical results indeed.⁵⁷

For the two extreme conditions of pressure approaching zero or infinity, the rate constants can be calculated using the analytical formulas given below. These analytical solutions, which are equivalent to the two-transition state (2-TS) kinetics model,^{58–60} were used to validate the numerical solutions of eqn (3).

For the zero-pressure limit (LPL):

$$k_1(T)_{P=0} = \frac{\sigma_{\text{bi}}}{h} \cdot \frac{Q_{\text{e}}^{\ddagger} Q_{\text{t}}^{\ddagger}}{Q_{\text{OH}} \cdot Q_{\text{H}_2\text{S}}} \cdot \sum_{J=0}^{J_{\text{max}}} (2J+1) \times \int_{E=0}^{E_{\text{max}}} G_{\text{eff}}(E, J) \cdot \exp\left(-\frac{E}{RT}\right) dE \quad (13a)$$

with

$$\frac{1}{G_{\text{eff}}(E, J)} = \frac{1}{\text{Min}(G_0(E, J))} + \frac{1}{G_{\text{TS1}}(E, J)} \quad (13b)$$

For the infinity-pressure limit (HPL):

$$k_1(T)_{P=\infty} = k_{0,\infty}(T) \times \frac{k_{1,\infty}^u(T)}{k_{-0,\infty}^u(T) + k_{1,\infty}^u(T)} \quad (14a)$$

in which $k_{1,\infty}^u(T)$ and $k_{-0,\infty}^u(T)$ are thermal rate coefficients in the high-pressure limit of the unimolecular reaction to products and re-dissociation of the PRC, respectively.

$$k_{1,\infty}^u(T) = \frac{\sigma_1}{h} \cdot \frac{1}{Q_{\text{PRC}}} \cdot \sum_{J=0}^{J_{\text{max}}} (2J+1) \times \int_{E=0}^{E_{\text{max}}} G_{\text{TS1}}(E, J) \cdot \exp\left(-\frac{E}{RT}\right) dE \quad (14b)$$

$$k_{-0,\infty}^u(T) = \frac{\sigma_0}{h} \cdot \frac{1}{Q_{\text{PRC}}} \cdot \sum_{J=0}^{J_{\text{max}}} (2J+1) \times \int_{E=0}^{E_{\text{max}}} \text{Min}(G_0(E, J)) \cdot \exp\left(-\frac{E}{RT}\right) dE \quad (14c)$$

where Q_{PRC} is the rotation–vibration partition function of PRC.

The pTS, pTDME, and fTDME codes of the Multiwell software package³⁹ were used for the chemical kinetics analysis.

Results and discussion

Reaction mechanism

Key stationary points on the lowest-lying doublet electronic state potential energy surface (PES) displayed in Fig. 1 were characterized using a composite HEAT-345Q(P) method. As seen there, the capture of OH by H₂S leads to the formation of a van der Waals complex (PRC), which has a binding energy of 6.6 kJ mol^{−1}. PRC has a C_s symmetry with two (nearly degenerate) electronic states of ²A'' and ²A'. When formed, PRC can re-dissociate back to the initial reactants *via* a loose, variational

Table 1 Individual contributions (kJ mol^{−1}) of the various terms to the relative energy of TS1, the PRC and the reaction enthalpy at 0 K using the HEAT-345Q(P) method

Term	TS1	PRC	H ₂ O + SH
δE _{SCF} →∞	59.91	−6.64	−78.81
δE _{CCSD(T)} →∞	−59.32	−6.74	−45.30
δE _{CCSDT} →∞	−1.32	0.01	0.37
δE _{CCSDTQ(P)}	−0.58	−0.04	−0.40
δE _{scalar}	−0.35	−0.16	−0.59
δE _{ZPE}	2.54	6.38	9.93
δE _{DBOC}	−0.04	0.18	−0.28
δE _{spin-orbit}	0.46	0.46	−1.58
HEAT	1.30	−6.55	−116.66 ± 0.5
ATcT			−116.0 ± 0.4 ^a

^a Heats of formation at 0 K of OH (37.278 ± 0.022 kJ mol^{−1}), H₂S (−17.36 ± 0.18 kJ mol^{−1}), H₂O (−238.903 ± 0.022 kJ mol^{−1}), and SH (142.81 ± 0.18 kJ mol^{−1}) are taken from ATcT (ref. 24).

TS0 (not shown in Fig. 1) or can undergo a H-abstraction *via* a tight TS1 leading to a (highly vibrationally excited) pre-product complex (PPC[‡]), which then decomposes quickly into products, H₂O + SH. TS1 has no symmetry and lies 1.3 kJ mol^{−1} above the initial reactants. Given that overcoming (or tunneling through) TS1 is the rate determining step, it is important to understand the contributions of the various terms to the energy barrier in the HEAT calculations. As shown in Table 1, the SCF energy is the most important, and it significantly overpredicts the effective barrier (*i.e.* relative to the reactants) at +59.9 kJ mol^{−1}. The CCSD(T) electron correlation is the second most important, and it brings the barrier down to +0.6 kJ mol^{−1}. The ZPE correction is the third most important, and it increases the barrier again by 2.6 kJ mol^{−1}. It should be mentioned that the 1D-HIR treatment increases the barrier by +0.36 kJ mol^{−1} due to the change of ZPE. The higher-level corrections (HLC) that go beyond the CCSD(T) lower the barrier by −1.9 kJ mol^{−1}. Finally, the scalar relativity effects and spin–orbit correction also have important contributions, but they have opposite signs.

The title reaction was theoretically studied earlier.^{17–20} A comparison of relative energies obtained in this work with other values reported in the literature is summarized in Table 2. As seen there, the agreement between different levels of theory is good, *i.e.* within 3 kJ mol^{−1}. It should be noted that the energies reported in the literature^{17–20} do not include the higher-level corrections (such as CCSDT and CCSDTQ(P)) as well as other corrections (anharmonic correction, spin–orbit, and scalar relativity effects). As compared to the benchmark ATcT value²⁴ for the reaction enthalpy, the HEAT-345Q(P) value is in excellent agreement, within 0.66 kJ mol^{−1}. To the best of our knowledge, the HEAT-345Q(P) method used in this work is the highest level of theory that has been applied to the title reaction. Surprisingly, the DFT calculations with the M06-2X/MG3S method¹⁸ yield an effective barrier of 1.05 kJ mol^{−1}, which is only 0.25 kJ mol^{−1} lower than the HEAT-345Q(P) value.

Reaction rate coefficients

The experimental k_1 results (at $T = 230$ – 550 K) are not indicative of pressure dependence for $P < 1$ atm.^{1–11} To check the



Table 2 A comparison of energies^a (kJ mol⁻¹) relative to initial reactants (OH + H₂S) calculated using the HEAT-345Q(P) method with values reported in the literature

Method	TS1	H ₂ O + SH	PRC
HEAT-345Q(P)	1.30 (-1.24)	-116.66 ± 0.5 (-126.59)	-6.55 (-12.93)
CCSD(T)/aV(5+d)Z ^b	3.14 (0.46)	-113.22 (-123.47)	-6.03 (-13.81)
MCG3/3//MC-QCISD/3 ^c	(1.72)	(-131.38)	N/A
M06-2X/MG3S ^c	1.05 (-1.00)	(-123.43)	N/A
UCCSD(T)-F12a/aVTZ ^d	2.37 (-0.43)	-113.70 (-124.20)	-7.72 (-14.00)
ATcT ^e	N/A	-116.0 ± 0.4	N/A

^a The values given in parentheses are exclusive of ZPE correction.

^b Taken from ref. 17. It was calculated using the fc-CCSD(T)/aug-cc-pV(5+d)Z//fc-CCSD(T)/aug-cc-pV(Q+d)Z level of theory. ^c Reported in ref. 18. ^d From ref. 19. ^e Taken from the benchmark ATcT (ref. 24).

possible effects of pressure on the rate coefficient, we computed the ratio of rate constants at the HPL using eqn (14) and rate constants at the LPL using eqn (13). The results presented in Fig. 2 show that the rate constant is effectively independent of pressure for $T \geq 200$ K, *i.e.* for example, for all applications relevant to atmospheric and combustion environments. The pressure dependence becomes noticeable when $T < 100$ K, and it increases sharply when temperature decreases, for example, to temperatures as in the ISM. However, as the pressure in the ISM is extremely low, the title reaction is beyond doubt in the LPL in this environment. There are three different (combined) causes for the lack of pressure dependence of the rate constant at $T > 200$ K: (i) first, the van der Waals complex PRC lies in a shallow well of only 6.6 kJ mol⁻¹, so its collisional stabilization and thermalization is unlikely to compete effectively with its prompt re-dissociation except at extremely high pressures; (ii) the barrier is very low, only 1.3 kJ mol⁻¹ above the initial reactants, such that tunneling through the barrier is expected

to be a minor effect except at very low temperatures; and (iii) the imaginary vibration frequency along the reaction coordinate of TS1, which is inversely proportional to the width of the barrier, has a moderate value of only 882i cm⁻¹. These three factors combined keep tunneling effects minor for $T > 300$ K (see Fig. 3). As shown in Fig. 3, the tunneling factor increases sharply once $T < 100$ K. It may be noted that quantum mechanical tunneling is the main cause for the pressure-dependent rate of the similar reaction of OH and HNO₃ as shown in a previous study.⁶¹ Note that the results of fall-off curves (*i.e.* $k(T, P)$) at very low temperatures are provided in the ESI† (see Fig. S3, ESI†).

For the following discussion, the results of this work for $k_1(T, P = 0)$ were computed at the LPL, using eqn (13a) and (13b), and denoted as $k_1(T, 0)$. Fig. 4 shows a comparison of our theoretical LPL results with various experimental data for $T = 230$ –550 K. In general, the agreement between theory and experiment is good (within *ca.* 30%). Our *ab initio* $k_1(T, 0)$ data (solid black curve) are in line with the higher experimental values, while when increasing the barrier of TS1 by 0.5 kJ mol⁻¹ (the possible energy error of the HEAT-345Q(P) method), the calculated data (dotted pink curve) agree well with the lower experimental values. Fig. 5 shows a fit of a modified Arrhenius equation to the combined set of our theoretical LPL results and all experimental data in Fig. 4: $k_1(T) = 9.28 \times 10^{-16} \times T^{1.325} \times$

$\exp\left(\frac{+286}{T}\right)$ cm³ s⁻¹ for $T = 200$ –500 K, with equal weights given to our *ab initio* data and the combined experimental data; this fit is represented by the solid red line in Fig. 5. The $k_1(T)$ fit equation above is recommended by this work for atmospheric applications. Two earlier recommendations^{15,16} are also shown in Fig. 5; as can be seen, they differ from our present recommendation either at the higher or at the lower temperatures of this range.

Fig. 6 shows LPL rate coefficients calculated for the extensive temperature range from 10 to 2500 K. The theoretical curves

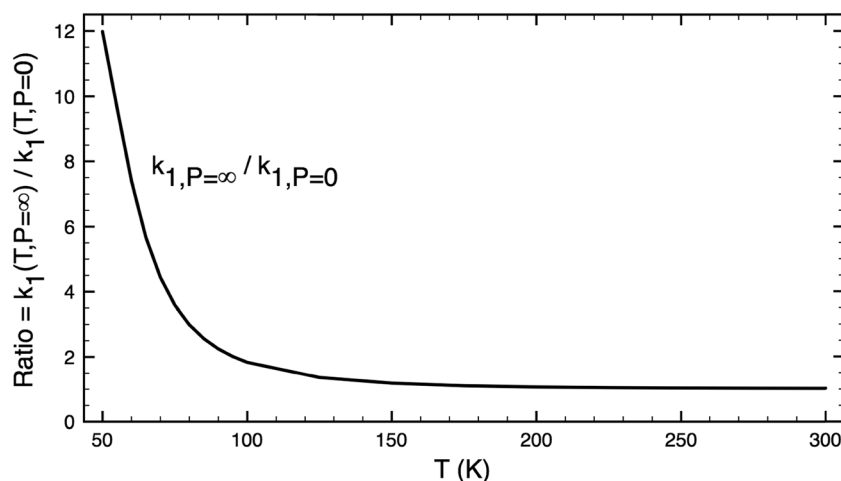


Fig. 2 Ratio of the calculated $k_{1,P=\infty}(T)/k_{1,P=0}(T)$ of the H₂S + OH reaction as a function of temperature. The rate coefficient is effectively independent of pressure for $T > 200$ K.



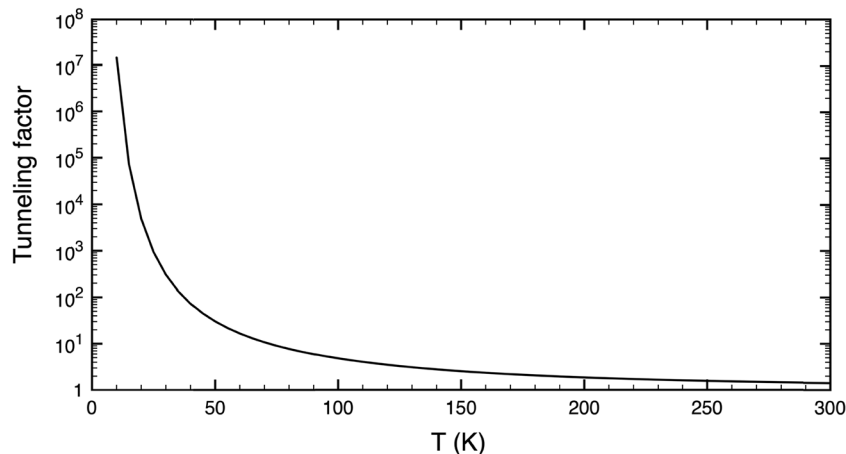


Fig. 3 Tunneling factor of the $\text{H}_2\text{S} + \text{OH}$ reaction from first principles as a function of temperature.

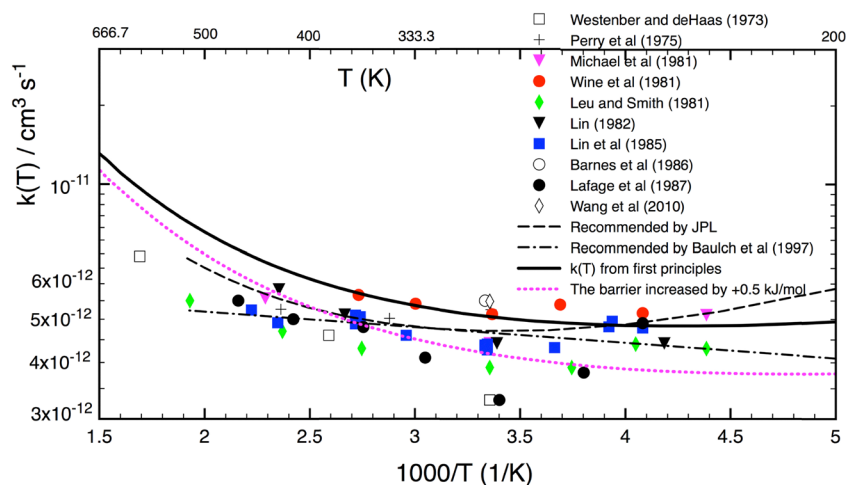


Fig. 4 A comparison of the calculated LPL rate coefficients $k_1(T, P = 0)$ from this work with experimental data (symbols)^{1–11} and two $k_1(T)$ expressions recommended in the literature^{15,16} for the temperature range of 200–550 K. The solid black curve represents our $k_1(T, P = 0)$ data from first principles, while the dotted pink curve is obtained by raising the TS1 barrier by +0.5 kJ mol^{–1} (to 1.8 kJ mol^{–1}), which is the expected possible error of the HEAT-345Q(P) method for this reaction system (see text).

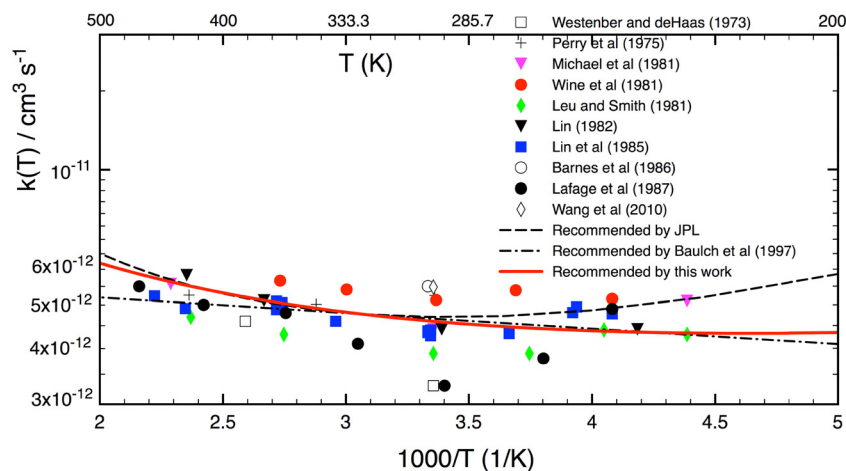


Fig. 5 A comparison of the rate coefficient recommended in this work (solid red curve; see text) $k_1(T) = 9.28 \times 10^{-16} \times T^{1.325} \times \exp\left(\frac{+286}{T}\right)$ in cm³ s^{–1} for $T = 200$ –500 K with the experimental data (symbols)^{1–11} as well as two other recommendations^{15,16} in the literature.



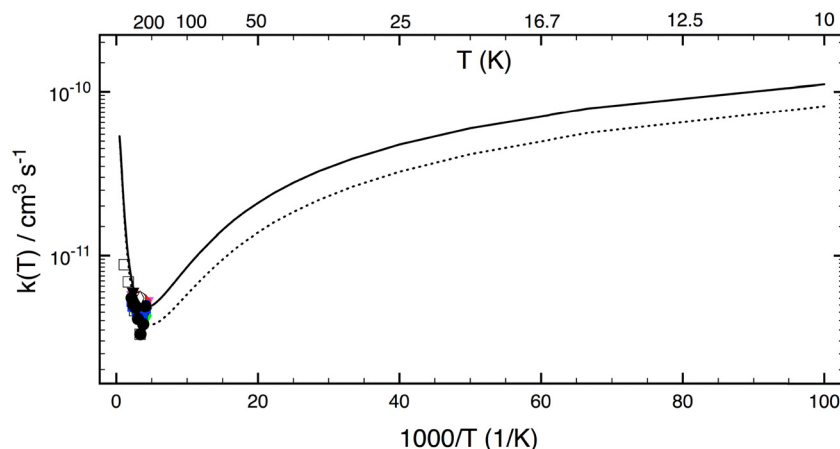


Fig. 6 Rate constants calculated using SCTST/2DME/HEAT approaches (this work) for a wide temperature range of 10–2500 K. The solid curve represents values from first principles. The dotted curve represents values obtained by raising the TS1 barrier by +0.5 kJ mol^{−1}. Experimental data (symbols)^{1–11} are also included for comparison.

(the *ab initio* result and that for TS1 0.5 kJ mol^{−1} higher) have a pronounced concave shape. Starting from a temperature of 10 K, with $k_1(10 \text{ K}, P = 0)$ close to $10^{-10} \text{ cm}^3 \text{ s}^{-1}$, the rate coefficient decreases as the temperature increases, to reach a minimum at about 250 K, and then increases again with temperature. Interestingly, at 10 K—characteristic of the ISM environment— $k_1(P = 0)$ is about 20 times higher than at room temperature ($5 \times 10^{-12} \text{ cm}^3 \text{ s}^{-1}$), mainly due to quantum tunneling effects. For the T range 200–2500 K, our *ab initio* LPL results can be represented in good approximation by the modified Arrhenius expression:

$$k_1(T, P = 0) = 1.44 \times 10^{-16} \times T^{1.674} \times \exp\left(\frac{+257}{T}\right) \text{ cm}^3 \text{ s}^{-1}.$$

Computed rate coefficient data $k_1(T)$ were reported earlier by Ellingson *et al.*¹⁸ using the CVT/SCT method based on a PES constructed using the M06-2X/MG3S DFT-method. In Table 3, the CVT/SCT-based values are compared with our SCTST/2DME $k_1(T)$, showing good agreement, within *ca.* 30%, over a wide

temperature range of 200 to 2400 K. One can expect such a good agreement because the M06-2X barrier of TS1 agrees very well with the HEAT method's value (see Table 2); however, this agreement may be due to a cancellation of errors in the DFT method.

Conclusions

In this work, high-level theoretical calculations were used to readdress the reaction mechanism and kinetics of the H₂S + OH reaction: the high-accuracy HEAT method for energetics and a two-dimensional E, J -resolved master equation analysis for the kinetics. The reaction is shown to proceed by hydrogen abstraction *via* a well-skipping mechanism,⁶² leading to the formation of H₂O and SH. The calculated reaction enthalpy is in excellent agreement with the benchmark ATcT value. The rate coefficient was calculated from first principles over wide temperature (10–2500 K) and pressure ranges. Pressure dependence was found to be limited to the T range below 200 K, where it becomes gradually very important as T decreases. The rate coefficient data agree well with the available experimental data at 200–500 K, within *ca.* 30%, which is similar to the spread of the different measurements. For the range $T = 200$ –500 K, we recommend the rate expression $k_1(T) = 9.28 \times 10^{-16} \times T^{1.325} \times \exp\left(\frac{+286}{T}\right) \text{ cm}^3 \text{ s}^{-1}$ for atmospheric applications. For applications in the ISM environment and in combustion processes, *i.e.* for temperatures where experimental data are not available, we provide high-level theoretical results.

Conflicts of interest

There are no conflicts of interest to declare.

Table 3 A comparison of rate constants calculated using the SCTST/2DME approach in this work with those obtained using the CVT/SCT method and Polyrate software

T (K)	SCTST/2DME (<i>ab initio</i>) ^a	SCTST/2DME (adjusted barrier) ^b	CVT/SCT ^c	Difference ^d (%)
200	4.94×10^{-12}	3.78×10^{-12}	4.38×10^{-12}	−16
250	4.84×10^{-12}	3.87×10^{-12}	4.20×10^{-12}	−9
300	5.08×10^{-12}	4.20×10^{-12}	4.23×10^{-12}	−1
350	5.53×10^{-12}	4.70×10^{-12}	4.50×10^{-12}	4
400	6.15×10^{-12}	5.33×10^{-12}	4.87×10^{-12}	9
600	1.00×10^{-11}	9.12×10^{-12}	7.25×10^{-12}	21
1000	2.24×10^{-11}	2.12×10^{-11}	1.57×10^{-11}	26
1500	3.89×10^{-11}	3.76×10^{-11}	2.61×10^{-11}	31
2400	5.57×10^{-11}	5.45×10^{-11}	6.66×10^{-11}	−22

^a This work, $k_1(T, P = 0)$ calculated from first principles (see text).

^b This work, with the energy of TS1 increased by +0.5 kJ mol^{−1} (see text).

^c Taken from ref. 18 based on a PES constructed using the M06-2X/MG3S level of theory. ^d The difference is defined as $\frac{k_{\text{SCTST}} - k_{\text{CVT/SCT}}}{k_{\text{SCTST}}} \times 100\%$.



Data availability

The data that support the findings of this study are available within the article and its ESI.† Detailed theoretical data that support the findings of this study are available from the corresponding author upon reasonable request.

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