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## Overlooked hydroperoxyl radical reactions in ammonia oxidation under combustion conditions†

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The present work focuses on highlighting bimolecular reactions in the  $\text{NH}_3$  system involving  $\text{HO}_2$ , an important radical at intermediate combustion temperatures. The reaction mechanism generator (RMG) tool was used to identify potentially significant reactions, and the automated rate calculator (ARC) tool was used to automatically compute rate coefficients at the ACCSD(T)/aug-cc-pVTZ-F12//B2PLYP-D3/aug-cc-pVTZ level of theory. Several reactions explored in this work, such as  $\text{N} + \text{HO}_2 \rightleftharpoons \text{NH} + \text{O}_2$ ,  $\text{NH} + \text{HO}_2 \rightleftharpoons \text{NH}_2 + \text{O}_2$ ,  $\text{NNH} + \text{HO}_2 \rightleftharpoons \text{N}_2\text{H}_2 + \text{O}_2$ , and  $\text{HNO}_2 + \text{HO}_2 \rightleftharpoons \text{NO}_2 + \text{H}_2\text{O}_2$ , have not been thoroughly investigated in the existing literature. In particular, the reaction  $\text{HNO} + \text{HO}_2 \rightleftharpoons \text{HNOH} + \text{O}_2$ , though known, lacks a precise rate coefficient in recent chemical kinetic models for ammonia. This study provides computed rate coefficients for 10 hydrogen abstraction and disproportionation reactions involving  $\text{HO}_2$  in the  $\text{NH}_3$  system. The reaction rate coefficients computed here may improve future low- and intermediate-temperature oxidation models of  $\text{NH}_3$ .

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### 1. Introduction

Ammonia has great potential to become a future zero-carbon energy carrier.<sup>1–5</sup> The advantages of ammonia as a fuel include a relatively high power-to-fuel-to-power (PFP) efficiency,<sup>6</sup> a large-scale distribution infrastructure that is already in place, a high-octane rating, and a narrow flammability range, making it relatively safe in terms of explosion risks. On the other hand, ammonia is toxic, emits significant levels of pollutants upon combustion ( $\text{NO}_x$  and  $\text{NH}_3$  residuals) and has a relatively low reactivity as a fuel.<sup>3</sup> A commonly proposed solution to this problem involves mixing ammonia with small amounts of

additional fuels, such as hydrogen. To make ammonia a viable energy carrier, it is crucial to have a reliable and versatile predictive<sup>7</sup> chemical kinetic model for its ignition and oxidation. Although an ammonia oxidation system involves relatively small molecules and substantial research efforts have been made to model it, there is still an ongoing debate regarding its ignition and oxidation chemistry.<sup>8–10</sup>

The main oxygen-containing radicals in the ammonia oxidation system are  $\text{O}$ ,  $\text{OH}$ ,  $\text{O}_2$ ,  $\text{HO}_2$ ,  $\text{NO}$ ,  $\text{NO}_2$ ,  $\text{H}_2\text{NO}$ , and  $\text{HNO}$ .<sup>11</sup> At intermediate temperatures typical of ignition engines,  $\text{HO}_2$  is the main radical chain carrier. This study aims to identify and examine previously unexplored hydrogen abstraction and disproportionation reactions in the ammonia system involving  $\text{HO}_2$ , and to compute the rate coefficients for these reactions.

Dean and Bozzelli<sup>12</sup> estimated the rate coefficient of  $\text{HNO} + \text{O}_2 \rightleftharpoons \text{NO} + \text{HO}_2$  by analogy to other  $\text{RH} + \text{O}_2$  reactions. Klippenstein *et al.*<sup>13</sup> computed the rate coefficient for  $\text{NNH} + \text{O}_2 \rightleftharpoons \text{N}_2 + \text{HO}_2$  at the CASPT2/aug-cc-pVDZ level<sup>14</sup> when exploring the role of  $\text{NNH}$  in  $\text{NO}$  formation. Stagni and Cavallotti<sup>15</sup> computed a rate coefficient for  $\text{H}_2\text{NO} + \text{HO}_2 \rightleftharpoons \text{HNO} + \text{H}_2\text{O}_2$  and  $\text{H}_2\text{NO} + \text{O}_2 \rightleftharpoons \text{HNO} + \text{HO}_2$  at the CASPT2/aug-cc-pVTZ level, while Chavarro Cañas *et al.*<sup>16</sup> computed rate coefficients for the latter reaction on both the doublet and the quartet surfaces at the W3X-L//CCSD/cc-pVTZ level (where the triplet  $\text{HNO}$  is the product on the quartet surface). Klippenstein and Glarborg<sup>17</sup> computed the rate coefficient of  $\text{NH}_2 + \text{HO}_2 \rightleftharpoons \text{NH}_3 + \text{O}_2$  at the CCSD(T)-F12/CBS level. Chavarro Cañas *et al.*<sup>16</sup> again computed rate coefficients for the same reactants on both the singlet and triplet surfaces at the W3X-L//CCSD/cc-pVTZ level (with singlet  $\text{O}_2$  as the product on the singlet surface).

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† Electronic supplementary information (ESI) available: A PDF file with: Table S1: species SMILES representation and T1 diagnostic coefficient; Table S2: transition state geometries, frequencies, and reaction path zero-point energies; details on rate coefficient kinetic computations and statistical mechanics; Fig. S1: the rate coefficients for  $\text{N} + \text{HO}_2 \rightleftharpoons \text{NH} + \text{O}_2$  (reaction R1) computed here and estimated by RMG; Fig. S2: the rate coefficients for  $\text{NH}_2\text{O} + \text{HO}_2 \rightleftharpoons \text{NH}_3\text{O} + \text{O}_2$  (reaction R4) computed here and estimated by RMG; Fig. S3: comparison of rate coefficients for  $\text{NH}_3 + \text{HO}_2 \rightleftharpoons \text{NH}_2 + \text{H}_2\text{O}_2$  (reaction R7); Fig. S4: comparison of rate coefficients for  $\text{N}_2\text{H}_3 + \text{HO}_2$  giving  $\text{H}_2\text{NN}(\text{T}) + \text{H}_2\text{O}_2$  (Reaction R8); Table S3: Chemkin format thermodynamic NASA polynomials for  $\text{NH}_3\text{O}$  and  $\text{H}_2\text{NN}(\text{T})$  and kinetic rate coefficients for the reactions reported in the manuscript; Table S4: Cantera format thermodynamic NASA polynomials for  $\text{NH}_3\text{O}$  and  $\text{H}_2\text{NN}(\text{T})$  and kinetic rate coefficients for the reactions reported in the manuscript; A ZIP file with all species and TS geometries as individual XYZ files; Table S5: the main reference determinants of the CASSCF calculation for each TS. See DOI: <https://doi.org/10.1039/d4cp01761g>



The research described above represents a significant body of work. However, no systematic study on  $\text{HO}_2$  reactivity has been conducted specifically for ammonia oxidation systems. This study thoroughly examines additional significant hydrogen abstraction and disproportionation reactions involving  $\text{HO}_2$  in the  $\text{NH}_3$  oxidation system that remain unexplored.

## 2. Methods

### Quantum chemical computations and statistical mechanics

*Ab initio* calculations were automatically performed using the open-source automated rate calculator (ARC) software tool,<sup>18</sup> an extensible codebase to automatically calculate thermochemical properties and reaction rate coefficients. Statistical mechanics processing of the electronic structure calculations was performed using Arkane.<sup>19</sup> The transition state theory equation for rate coefficient computation is given in the ESI† (eqn S1). Detailed explanations of the equations used in Arkane are available in the literature.<sup>19</sup> 3D conformer geometry searches were conducted using ARC with a dihedral angle combination approach based on a random set of force field (MMFF94s<sup>20</sup>) conformers generated using RDKit.<sup>21</sup> Up to ten lowest energy force field conformers identified per species were optimized at the  $\omega\text{B97X-D}/\text{Def2-SVP}$ <sup>22,23</sup> level of theory (termed ‘L1’). The lowest energy density functional theory (DFT) conformer was optimized at the double hybrid B2PLYP functional<sup>24</sup> with Grimme’s dispersion correction and Becke–Johnson damping, D3(BJ),<sup>25,26</sup> coupled with Dunning’s correlation-consistent basis set augmented with diffuse functions, aug-cc-pVTZ,<sup>27</sup> level of theory. Harmonic frequencies were subsequently computed at the same level of theory. ARC calculated a frequency scaling factor of 0.995 for B2PLYP-D3/aug-cc-pVTZ (termed ‘L2’) following the method recommended by Truhlar *et al.*<sup>28</sup>

Torsional modes were automatically identified as rotatable single bonds in each species, considering relevant resonance structures<sup>29</sup> and treated with continuous constrained potential energy surface (PES) optimizations with all other internal coordinates relaxed using the ‘L2’ level of theory at a resolution of 10°. If a scan resulted in a lower-energy structure than the original conformer geometry optimized at the same level, ARC identified the former as the new lowest-energy conformer, deleted all other running jobs for the species, and spawned the computations again starting from the updated geometry. The 1D torsional PESs were fitted to truncated Fourier series and used as input to compute energy levels and hence the partition function of the anharmonic mode using Arkane.<sup>19</sup> Quantum tunneling effects were considered using the Eckart correction function.<sup>30</sup>

Single-point energies were calculated using three methods for comparison, one of which is multireference (MR), and the other two are single reference (SR). The MR method is multi-configuration reference configuration-interaction method with the Davidson correction (MRCI + Q).<sup>31–33</sup> The SR approaches comprised the explicitly correlated CCSD(T)-F12a<sup>34</sup> and the left-eigenvector-based ACCSD(T).<sup>35,36</sup> Although formally a single-reference coupled-cluster method, ACCSD(T) exploits

its left-eigenvector-based triples correction to achieve enhanced error cancellation in systems exhibiting moderate static correlation, allowing it to approach MR benchmark accuracy in many near-degenerate situations.<sup>37–39</sup> All ACCSD(T) single-point energies reported here employ an RHF reference determinant; we note that CCSD(T) built on a broken-symmetry UHF reference may yield even lower errors in certain MR-like cases.<sup>37</sup> Nonetheless, ACCSD(T) often exhibits enhanced numerical stability and more consistent error cancellation than RHF-based CCSD(T) under moderate static correlation, making it a robust choice for near-degenerate reaction pathways.<sup>39</sup> All single-point energy calculations employed the aug-cc-pVTZ basis set.<sup>27</sup>

All DFT calculations mentioned above were performed in Gaussian 16.<sup>40</sup> CCSD(T)-F12a computations were performed in Molpro 2022.3.0,<sup>41</sup> with the unrestricted Hartree–Fock (UHF)<sup>42</sup> reference wavefunction. MRCI calculations were performed with Orca 5.0.4,<sup>43</sup> using the complete-active-space self-consistent-field (CASSCF)<sup>44</sup> reference wavefunction (see Table S2 in the ESI† for the number of active electrons and orbitals), and ACCSD(T) calculations were performed with the MRCC software<sup>45</sup> via Psi4’s interface<sup>46</sup> using UHF as a reference wavefunction.

All electronic structure calculations were processed by Arkane<sup>19</sup> for computation of the thermochemical partition functions and reaction rate coefficients. Empirical systematic errors in atomization energies were corrected using atom energy corrections implemented in Arkane<sup>19,47</sup> when available, or computed in the present work for the respective level of theory.

Reaction TS searches were automated using a bimolecular reaction orientation module implemented in ARC.<sup>18</sup> The generated TS guesses are then optimized at the ‘L1’ level. They are clustered by nearly identical internal coordinates after the ‘L1’ geometry optimization, and representative structures from each cluster were ranked by their relative DFT electronic energy. The algorithm then analyzed TS candidate structures in ascending energy order and performed another saddle-point optimization and ro-vibrational analyses at the ‘L2’ level. The algorithm was terminated once an ‘L2’ optimized TS guess passed all relevant checks, *i.e.*, reaction path energy, normal mode displacement analysis, and intrinsic reaction coordinate calculations.<sup>48</sup>

### Reaction identification

The reaction mechanism generator (RMG)<sup>49,50</sup> was used to identify reactions in the subset of interest, *i.e.*, hydrogen abstraction and disproportionation reactions with estimated rate coefficients involving  $\text{HO}_2$  either as an attacking radical or as a result of a radical attack by  $\text{O}_2$ . The model was generated for temperature, pressure and equivalence ratio ranges of 700–1500 K, 1–50 bar, and 0.5–1.5, respectively. A termination rate ratio criterion was set, stopping model generation when the characteristic rate reached 1% of the maximum characteristic rate in the system. The characteristic rate is a function of time and is defined as the root mean square of net rates of production of all the species in the model.

### Chemical kinetic modeling

Two recent models, one by Zhang *et al.*<sup>51</sup> and another by Zhu *et al.*,<sup>52</sup> were used as the basis for simulations in this work. An



updated version of each model was created by incorporating all the reactions and rate coefficients computed in this study into the existing literature models. Four of the reactions reported here already existed in the Sarathy *et al.* model, and three already existed in the Zhu model; their rate coefficients were updated. The models were simulated in a continuously stirred reactor using Cantera v. 3.1.0.<sup>53</sup> Rates of production (ROP) were computed using Cantera, and flux diagrams were generated using features of The Tandem Tool (T3) for automated chemical kinetic model development.<sup>54</sup>

### 3. Results and discussion

Eleven reactions were identified by RMG as belonging to either the hydrogen abstraction or the disproportionation family involving HO<sub>2</sub> and having an estimated rate. Of the eleven reactions, ten belong to the hydrogen abstraction reaction family. In five hydrogen abstraction reactions (R1–R5, Table 1) the hydrogen atom in HO<sub>2</sub> is abstracted and O<sub>2</sub> is formed, and in the other five reactions (R6–R10, Table 1), HO<sub>2</sub> is the abstracting radical, forming H<sub>2</sub>O<sub>2</sub>. One reaction (R11, Table 1) belongs to the disproportionation reaction family.<sup>55</sup> It involves an O<sub>2</sub> molecule abstracting a hydrogen atom from a position adjacent to a radical atom, forming HO<sub>2</sub>.

The T1 diagnostic coefficients<sup>56–58</sup> of the species participating in these reactions, along with their SMILES representations,<sup>59</sup> are given in Table S1 (ESI<sup>†</sup>). The T1 diagnostic coefficients of the TSs are given in Table 1. The inspected energy wells had no imaginary frequencies, and all TSs had only one imaginary frequency as listed in Table S2 (ESI<sup>†</sup>). The TS geometries of all reactions are shown in Fig. 1 and are listed in Table S2 (ESI<sup>†</sup>) together with all the respective vibrational frequencies. The multiplicity of each reaction surface is provided in Table 1, the multiplicity of reaction R2 is discussed in details below. The computed rate coefficients at the ACCSD(T)/aug-cc-pVTZ//B2PLYP-D3/aug-cc-pVTZ level of theory are given in Table 3 and in Chemkin and Cantera formats in Tables S3 and S4 (ESI<sup>†</sup>), respectively.

Some of the species (NO, H<sub>2</sub>NO, NNH, H<sub>2</sub>NN(T), N<sub>2</sub>H<sub>3</sub>, HNOH, NO<sub>2</sub>, and notably HO<sub>2</sub>) and all TSs have a relatively high T1 diagnostic coefficient (Table 1 and Table S1, ESI<sup>†</sup>), which may indicate a strong MR character. Table 1 shows the differences in energy at 0 K including zero-point energy corrections<sup>62</sup> between the wells of the products and reactants of each reaction, while Table 2 shows the deviation of these computed reaction energies from the corresponding enthalpy differences reported by ATct.<sup>60,61</sup> In both cases, values are reported as differences between the product and reactant wells. The high T1 diagnostic coefficients observed for many species and all TSs suggest that MR methods may be necessary.

The MRCI calculations performed in this work exhibit significant deviations compared to the experimentally-based ATct values (Table 2). Specifically, the MRCI method shows a significant average deviation of 10.5 kJ mol<sup>−1</sup> with a standard deviation of 9.2 kJ mol<sup>−1</sup> from the ATct values (Table 2). In contrast, the SR methods performed remarkably well in reproducing the ATct data. Both CCSD(T)-F12 and ACCSD(T) methods exhibit excellent agreement with the ATct values, with average deviations of just 0.2 and −0.6 kJ mol<sup>−1</sup> and standard deviations of 2.0 and 3.7 kJ mol<sup>−1</sup>, respectively (Table 1). For example, the ACCSD(T)/aug-cc-pVTZ and MRCI + Q/aug-cc-pVTZ methods disagree with one another in the computed  $\Delta E_0(P - R)$  by 10.6 kJ mol<sup>−1</sup> on average with a standard deviation of 9.8 kJ mol<sup>−1</sup> (Table 1) after excluding the large deviation observed for reaction R4. When R4 is included, the average difference increases to 21.3 kJ mol<sup>−1</sup>.

Although MRCI captures MR effects, it falls short in accurately describing the energetics of key stationary points and is thus likely to produce inaccurate predictions for these reaction systems. This discrepancy can be attributed to limitations in capturing dynamic electron correlation, an area where CCSD(T) and ACCSD(T) are known to excel, but where MR methods often struggle. Furthermore, the lack of size extensivity in MRCI + Q further undermines its suitability for modeling bimolecular reaction energetics.

Coupled-cluster theory is widely recognized for its high accuracy and broad applicability, and it is considered the

**Table 1** Zero-point energy comparisons of the differences between product and reactant wells,  $\Delta E_0(P - R)$ , for the studied reactions at several levels of theory: CCSD(T)-F12a, ACCSD(T) and MRCI + Q

No.	Reaction	<i>m</i> <sup>a</sup>	TS T1 diagnostic coefficient <sup>b</sup>	<i>P</i> − <i>R</i> <sup>c</sup> CCSD(T) <sup>d</sup>	<i>P</i> − <i>R</i> <sup>c</sup> ACCSD(T) <sup>e</sup>	<i>P</i> − <i>R</i> <sup>c</sup> MRCI <sup>f</sup>
R1	N + HO <sub>2</sub> ⇌ NH + O <sub>2</sub>	5	0.046	−125.5	−119.3	−139.0
R2	NH + HO <sub>2</sub> ⇌ NH <sub>2</sub> + O <sub>2</sub>	2, 4 <sup>g</sup>	0.043	−183.4	−182.7	−190.3
R3	NO + HO <sub>2</sub> ⇌ HNO + O <sub>2</sub>	3	0.045	4.1	1.7	−7.2
R4	H <sub>2</sub> NO + HO <sub>2</sub> ⇌ NH <sub>3</sub> O + O <sub>2</sub>	3	0.053	13.5	−14.3	−32.6
R5	NNH + HO <sub>2</sub> ⇌ N <sub>2</sub> H <sub>2</sub> + O <sub>2</sub>	3	0.043	−60.3	−64	−77.5
R6	NH <sub>2</sub> + HO <sub>2</sub> ⇌ NH + H <sub>2</sub> O <sub>2</sub>	3	0.047	24.2	23.5	18.3
R7	NH <sub>3</sub> + HO <sub>2</sub> ⇌ NH <sub>2</sub> + H <sub>2</sub> O <sub>2</sub>	2	0.041	81.8	81.8	77.1
R8	N <sub>2</sub> H <sub>3</sub> + HO <sub>2</sub> ⇌ H <sub>2</sub> NN(T) + H <sub>2</sub> O <sub>2</sub>	3	0.047	−6.8	−6.5	−12.1
R9	HNO <sub>2</sub> + HO <sub>2</sub> ⇌ NO <sub>2</sub> + H <sub>2</sub> O <sub>2</sub>	2	0.040	−70.4	−67.5	−72.0
R10	N <sub>2</sub> H <sub>2</sub> + HO <sub>2</sub> ⇌ NNH + H <sub>2</sub> O <sub>2</sub>	2	0.049	−98.6	−95.2	−97.0
R11	HNO + HO <sub>2</sub> ⇌ HNOH + O <sub>2</sub>	2	0.040	−18.5	−26.0	−52.8

<sup>a</sup> The reaction multiplicity. <sup>b</sup> The T1 diagnostic coefficient<sup>56–58</sup> of the corresponding TS is based on the CCSD(T)-F12a computation. <sup>c</sup> Energy differences in kJ mol<sup>−1</sup> between the corresponding products well and reactants well in the direction the reaction is written here. <sup>d</sup> As computed at the CCSD(T)-F12a/aug-cc-pVTZ-F12//B2PLYPD3/aug-cc-pVTZ level of theory. <sup>e</sup> As computed at the ACCSD(T)/aug-cc-pVTZ//B2PLYPD3/aug-cc-pVTZ level of theory. <sup>f</sup> As computed at the MRCI + Q/aug-cc-pVTZ//B2PLYPD3/aug-cc-pVTZ level of theory. <sup>g</sup> See text for a detailed discussion of the multiplicity of reaction R2.



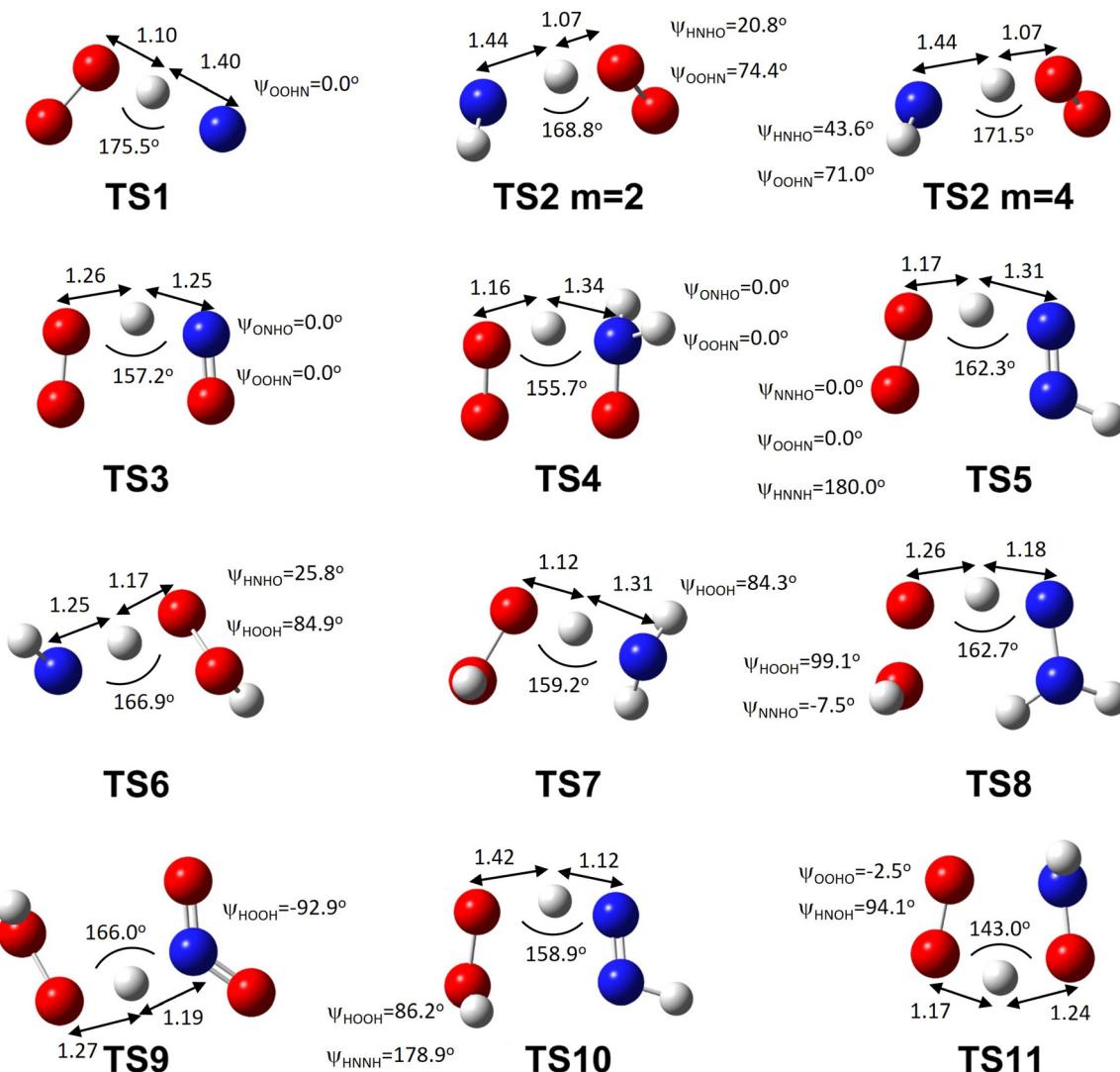


Fig. 1 TS geometries of the explored reactions. For reaction R2 (Table 1), two TS geometries are given at the two relevant multiplicities of 2 and 4. Distances are reported in Angstrom, angles and dihedrals are reported in degrees. Dihedral labels refer to consecutive atoms. Element legend: red: oxygen, blue: nitrogen, white: hydrogen. The Cartesian geometry is reported in Table S2 (ESI†).

“gold standard” in quantum chemistry. Indeed, CCSD(T) performed exceptionally well for all species, even those with MR characteristics, compared to the experimentally-based data (Table 2). However, it may introduce deviations relative to the ground truth when applied to TSs because of their strong MR characteristics. Although ACCSD(T) remains formally a single-reference method, it has been shown to handle systems with multireference character, such as predictive singlet-triplet gaps, very effectively when supplied with a suitable reference wavefunction.<sup>37,63</sup> Consequently, ACCSD(T) offers greater numerical stability in non-equilibrium geometries than standard CCSD(T). High-spin TSs, such as TS1, are therefore expected to benefit substantially from the left-eigenvector-based triples correction built into ACCSD(T). Accordingly, we adopt the ACCSD(T) values for computing the rate coefficients in the present work. As shown in Table S5 (ESI†), all TS wavefunctions except TS2 at  $m = 2$  are dominated by a single

CASSCF configuration (leading weight  $\geq 0.889$ ), confirming that single-reference methods are appropriate for those barriers. By contrast, TS2's largest configuration weight for  $m = 2$  is only 0.417, placing it firmly in the strong MR regime. Accordingly, the reaction R2 results presented here should be regarded as preliminary, and a more rigorous MR treatment is required for quantitative accuracy.

Reactive collisions between NH(T) and O<sub>2</sub>(T) can lead to the formation of NO(D) + OH(D), HNO(S) + O(S), HNO(T) + O(T), and NO<sub>2</sub>(D) + H(D) when reacting unimolecularly and to N(D) + HO<sub>2</sub>(D) or N(Q) + HO<sub>2</sub>(D) either unimolecularly or bimolecularly (where S, D, T, and Q refer to singlet, doublet, triplet, and quartet spin states, respectively). Talipov *et al.*<sup>64</sup> computed rate coefficient values for the production of NO + OH and HNO + O, while Baulch *et al.*<sup>65</sup> commented that experimental and theoretical studies suggest that the formation of NO<sub>2</sub> + H in this system is negligible. Talipov *et al.*<sup>64</sup> studied the unimolecular

**Table 2** Deviation between reaction enthalpies from the ATcT and computed zero-point corrected reaction energies for various levels of theory. Both ATcT and computed values correspond to the energy difference between product and reactant wells in the direction written. Reported values are defined as  $[\Delta H_{f,0}^{\circ} \text{K}(\text{P} - \text{R})]_{\text{ATcT}} - [\Delta E_0(\text{P} - \text{R})]_{\text{calc}}$

No. Reaction	ATcT <sup>a</sup> – CCSD(T) <sup>b</sup>	ATcT <sup>a</sup> – ACCSD(T) <sup>c</sup>	ATcT <sup>a</sup> – MRCI <sup>d</sup>
R1 $\text{N} + \text{HO}_2 \rightleftharpoons \text{NH} + \text{O}_2$	–1.4	–7.6	12.1
R2 $\text{NH} + \text{HO}_2 \rightleftharpoons \text{NH}_2 + \text{O}_2^e$	–1.5	–2.2	5.4
R3 $\text{NO} + \text{HO}_2 \rightleftharpoons \text{HNO} + \text{O}_2$	–0.1	2.5	11.4
R4 $\text{H}_2\text{NO} + \text{HO}_2 \rightleftharpoons \text{NH}_3\text{O} + \text{O}_2$	–1.0	–0.2	18.1
R5 $\text{NNH} + \text{HO}_2 \rightleftharpoons \text{N}_2\text{H}_2 + \text{O}_2$	0.2	3.9	17.4
R6 $\text{NH}_2 + \text{HO}_2 \rightleftharpoons \text{NH} + \text{H}_2\text{O}_2$	1.1	1.8	7.0
R7 $\text{NH}_3 + \text{HO}_2 \rightleftharpoons \text{NH}_2 + \text{H}_2\text{O}_2$	1.2	1.2	5.9
R8 $\text{N}_2\text{H}_3 + \text{HO}_2 \rightleftharpoons \text{H}_2\text{NN}(\text{T}) + \text{H}_2\text{O}_2^f$	—	—	—
R9 $\text{HNO}_2 + \text{HO}_2 \rightleftharpoons \text{NO}_2 + \text{H}_2\text{O}_2$	–0.5	–3.4	1.1
R10 $\text{N}_2\text{H}_2 + \text{HO}_2 \rightleftharpoons \text{NNH} + \text{H}_2\text{O}_2$	–5.3	2.3	29.0
R11 $\text{HNO} + \text{HO}_2 \rightleftharpoons \text{HNOH} + \text{O}_2$	–0.9	–4.3	–2.5

<sup>a</sup>  $\Delta H_{f,0}^{\circ} \text{K}$  values are taken from the active thermochemical tables (ATcT, version 1.130),<sup>60,61</sup> all energy differences are reported in  $\text{kJ mol}^{-1}$ . <sup>b</sup> As computed at the CCSD(T)-F12a/aug-cc-pVTZ-F12//B2PLYPD3/aug-cc-pVTZ level of theory. <sup>c</sup> As computed at the ACCSD(T)/aug-cc-pVTZ//B2PLYPD3/aug-cc-pVTZ level of theory. <sup>d</sup> As computed at the MRCI + Q/aug-cc-pVTZ//B2PLYPD3/aug-cc-pVTZ level of theory. <sup>e</sup> TS2 exhibits a strong MR characteristic (Table S5, ESI), therefore the calculated rate coefficient for R2 carries a correspondingly higher uncertainty. <sup>f</sup> The active thermochemical tables version 1.130 did not have values for  $\text{H}_2\text{NN}(\text{T})$ , hence no comparison to ATcT is given for this reaction.

$\text{NH} + \text{O}_2$  PES. However, to our knowledge, there is no relation in the literature for reaction R1,  $\text{NH}(\text{T}) + \text{O}_2(\text{T}) \rightleftharpoons \text{N}(\text{Q}) + \text{HO}_2(\text{D})$  (listed in Table 1 in reverse for context). No rate, measured or computed, is available for this reaction, and none of the recent  $\text{NH}_3$  oxidation models<sup>11,51,52,66–81</sup> considered it. The rate coefficient of reaction R1 is plotted in Fig. S1 (ESI†).

Here, we calculated a bimolecular rate coefficient for the reaction  $\text{NH}(\text{T}) + \text{O}_2(\text{T}) \rightleftharpoons \text{N}(\text{Q}) + \text{HO}_2(\text{D})$  in both directions, and neglected the  $\text{N}(\text{D}) + \text{HO}_2(\text{D})$  products since  $\text{N}(\text{D})$  is  $\sim 300 \text{ kJ mol}^{-1}$  higher than  $\text{N}(\text{Q})$  (the ground state), making the electronically-excited well negligible under typical combustion conditions. The rate coefficient comparison in Fig. 2 shows that the pathway leading to the  $\text{N} + \text{HO}_2$  products only becomes significant above 2000 K, and that the main products of  $\text{NH} + \text{O}_2$  are expected to be  $\text{HNO} + \text{O}$ . The rate coefficient in the reverse direction,  $\text{N} + \text{HO}_2 \rightleftharpoons \text{NH} + \text{O}_2$  as written in Table 1, is still relatively low:  $\sim 3.3 \times 10^5 \text{ cm}^3 (\text{mol s})^{-1}$  at 1000 K (Fig. S1, ESI†). In this direction (as in Table 1) reaction R1 is exothermic by  $\sim 120 \text{ kJ mol}^{-1}$  with an energy barrier of  $\sim 140 \text{ kJ mol}^{-1}$  (Table S2, ESI†). Due to the relative importance of the N and  $\text{HO}_2$  radicals in this system, it is recommended to consider reaction R1 in future models.

Reaction R2,  $\text{NH} + \text{HO}_2 \rightleftharpoons \text{NH}_2 + \text{O}_2$ , was previously considered to proceed *via* the doublet or quartet surfaces ( $m = 2'$  and ' $m = 4'$ , respectively).<sup>82</sup> Both transition states (Fig. 1) are similar in terms of bond lengths (including the lengths of the reactive bonds) and angles, but differ in the torsional angle formed by the internal NH rotor, *i.e.*, the dihedral angle of the hydrogen atom bonded to N relative to the N–H–O plane (Fig. 1).

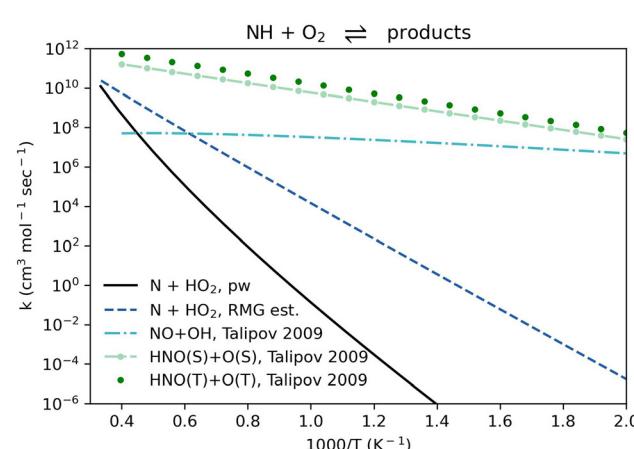
Reaction R2 on the doublet PES ( $m = 2$ ) proceeds through a complex electronic reconfiguration, as depicted in the spin-

**Table 3** Rate coefficients computed at the ACCSD(T)/aug-cc-pVTZ//B2PLYPD3/aug-cc-PVTZ level of theory reported in the modified Arrhenius expression format,  $k = A \cdot T^n \cdot e^{-E_a/RT}$ , fitted between 300–3000 K

No. Reaction	$A (\text{cm}^3 \text{ mol}^{-1} \text{ s}^{-1})$	$n$	$E_a (\text{kJ mol}^{-1})$
R1 $\text{N} + \text{HO}_2 \rightleftharpoons \text{NH} + \text{O}_2$	$1.82 \times 10^{-27}$	9.81	60.5
R2 $\text{NH} + \text{HO}_2 \rightleftharpoons \text{NH}_2 + \text{O}_2^a$	—	—	—
R3 $\text{NO} + \text{HO}_2 \rightleftharpoons \text{HNO} + \text{O}_2$	$1.19 \times 10^{-5}$	5.06	29.9
R4 $\text{H}_2\text{NO} + \text{HO}_2 \rightleftharpoons \text{NH}_3\text{O} + \text{O}_2$	$1.40 \times 10^{-5}$	4.53	8.1
R5 $\text{NNH} + \text{HO}_2 \rightleftharpoons \text{N}_2\text{H}_2 + \text{O}_2$	$6.25 \times 10^{-5}$	4.57	2.2
R6 $\text{NH}_2 + \text{HO}_2 \rightleftharpoons \text{NH} + \text{H}_2\text{O}_2$	$1.97 \times 10^{-5}$	5.23	34.2
R7 $\text{NH}_3 + \text{HO}_2 \rightleftharpoons \text{NH}_2 + \text{H}_2\text{O}_2$	$4.44 \times 10^{-1}$	4.00	75.5
R8 $\text{N}_2\text{H}_3 + \text{HO}_2 \rightleftharpoons \text{H}_2\text{NN}(\text{T}) + \text{H}_2\text{O}_2$	$2.79 \times 10^{-3}$	4.00	7.9
R9 $\text{HNO}_2 + \text{HO}_2 \rightleftharpoons \text{NO}_2 + \text{H}_2\text{O}_2$	$2.49 \times 10^{-3}$	4.52	0.2
R10 $\text{N}_2\text{H}_2 + \text{HO}_2 \rightleftharpoons \text{NNH} + \text{H}_2\text{O}_2$	$7.79 \times 10^{-1}$	3.96	–0.6
R11 $\text{HNO} + \text{HO}_2 \rightleftharpoons \text{HNOH} + \text{O}_2$	$3.08 \times 10^0$	2.98	2.4

<sup>a</sup> A rate coefficient for reaction R2 is not provided because the transition states pronounced multireference character (Table S5, ESI) precludes accurate barrier estimation by the methods used in this study.

conserving paths (A) and (B) in Fig. 3. A key aspect of this reaction is the redistribution of electrons during the transformation of reactants into products, which is directly connected to the spin state of  $\text{O}_2$  and NH. In path (A), triplet NH (arbitrarily assigned spins “up, up”) captures a hydrogen atom with a spin “down” electron, leaving a singlet  $\text{O}_2$  in the products. In path (B), when considered in reverse, ground state triplet  $\text{O}_2$  abstracts a hydrogen atom from  $\text{NH}_2$  with a spin “up” electron if triplet  $\text{O}_2$  is arbitrarily assigned spins “down, down”. In this path, a singlet NH is formed with a  $b^1\Sigma^+$  electronic configuration (Fig. 3), which can quickly convert into the more stable  $\text{NH} a^1\Delta$  configuration (that is, with two pairs of lone electrons). The R2 reaction *via* paths (A) or (B) on the doublet PES ( $m = 2$ ) has a different reactant well (singlet NH) or a different product well (singlet  $\text{O}_2$ ) relative to the quartet PES ( $m = 4$ ) pathway of R2. Therefore, R2m2 and R2m4 are not only different paths (on different spin surfaces) of the same reaction,



**Fig. 2** Comparison of rate coefficients for  $\text{NH} + \text{O}_2$  giving  $\text{N} + \text{HO}_2$  (reaction R1, given in reverse in Table 1) computed here (pw) and compared to RMG's estimate, and giving  $\text{NO} + \text{OH}$  and  $\text{HNO} + \text{O}$  as recommended by Talipov *et al.*<sup>64</sup>



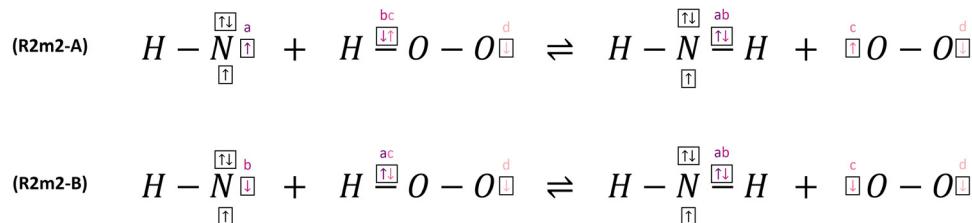


Fig. 3 The electronic configuration of the reactants and products of reaction R2 on the doublet potential energy surface (R2m2). (A) forming singlet  $O_2$   $1\Delta g$ , (B) reacting singlet  $NH$   $b^1\Sigma^+$ . Lowercase letters, arrows, and colors represent specific electron spins in each reaction path separately. The image does not show all bonding electrons and does not show the oxygen lone electron pairs.

but rather different reactions altogether with distinct reactant and product electronic configuration. Reacting  $NH(T) + HO_2$  on the quartet surface yield  $NH_2 + O_2(T)$ , while on the doublet surface the products will be  $NH_2 + O_2(S)$ . The R2m2-A and R2m2-B paths (Fig. 3) are not the same reaction themselves, due to the different electronic configurations of their reactants and products. We identified a single TS on the  $m = 2$  PES (Fig. 1), and the question of whether R2m = 2 path A and R2m = 2 path B (Fig. 3) share the same TS remains open.

Recent  $NH_3$  oxidation models<sup>11,51,52,66–81</sup> consider  $NH_2 + O_2 \rightleftharpoons H_2NO + O$  and  $NH_2 + O_2 \rightleftharpoons HNO + OH$  as the only reactions on the  $NH_2O_2$  PES, *e.g.*, as computed by Klippenstein *et al.*<sup>13</sup> These models do not consider reaction R2,  $NH_2 + O_2 \rightleftharpoons NH + HO_2$  (listed in Table 1 in reverse for context), which might become relatively significant at high temperatures (Fig. 4). Furthermore, reaction R2 could be significant in  $NH_3$  systems in the exothermic (forward) direction, affecting the interconversion of two main radicals in this system,  $NH_2$  and  $NH$ .

Reaction R3,  $NO + HO_2 \rightleftharpoons HNO + O_2$ , is a radical chain-terminating reaction that competes with the pressure-dependent reaction  $NO + HO_2 \rightleftharpoons NO_2 + OH$ . The rate coefficients computed here are in agreement with a previous work by Wang *et al.*<sup>84</sup> performed at the CCSD(T)/CBS//B2PLYPD3/aug-cc-pVTZ level of theory. Using the updated rate coefficient could be crucial for  $NH_3$  oxidation models to properly represent the radical pool,

replacing the relatively high values previously used by recent  $NH_3$  models<sup>11,51,52,66–81</sup> (Fig. 5).

The rate coefficient for reaction R4,  $H_2NO + HO_2 \rightleftharpoons NH_3O + O_2$ , was not reported previously, and it is plotted in Fig. S2 (ESI†). This reaction is exothermic by  $\sim 14$  kJ mol<sup>-1</sup> (Table S2, ESI†). Its rate coefficient is considerably lower than the RMG estimated value (Fig. S2, ESI†). This divergence could likely stem from the zwitterion generated in this reaction (*i.e.*,  $NH_3O$ , SMILES:<sup>59</sup> “[ $NH^{3+}][O^-]”]) and the absence of relevant training reactions within the RMG database. This reaction seems to be relatively slow, and after a rate coefficient refinement by ARC, the generated RMG model did not consider this reaction again for a temperature range of 700–1500 K. The current exercise shows that reaction R4 is probably insignificant for  $NH_3$  oxidation models.$

The dominant radicals that abstract a hydrogen atom from  $N_2H_2$  to form  $NNH$  are  $H$ ,  $NH$ ,  $NH_2$ ,  $O$ ,  $OH$ , and  $HO_2$ , as summarized in Fig. 6. The rate coefficient for  $HO_2$  as the attacking radical computed here (reaction R10) is lower than the rate coefficients of most other attacking radicals. The rate coefficient of reaction R5,  $N_2H_2 + O_2 \rightleftharpoons NNH + HO_2$  (listed in Table 1 in reverse for context), is smaller than the rate

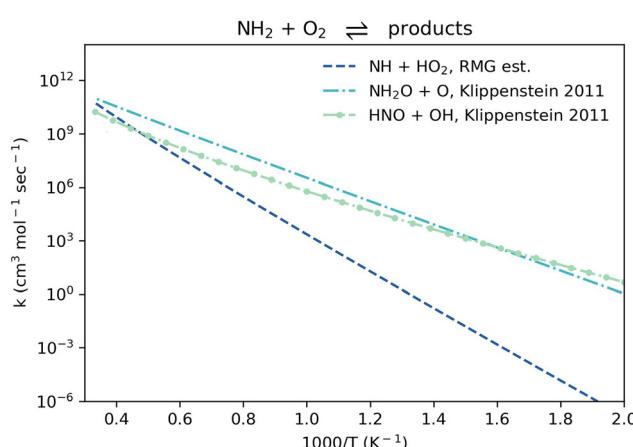


Fig. 4 Comparison of rate coefficients for  $NH_2 + O_2$  giving  $NH + HO_2$  (reaction R2, given in reverse in Table 1) with a comparison of RMG's estimate, and giving  $H_2NO + O$  and  $HNO + OH$  as computed by Klippenstein *et al.*<sup>13</sup>

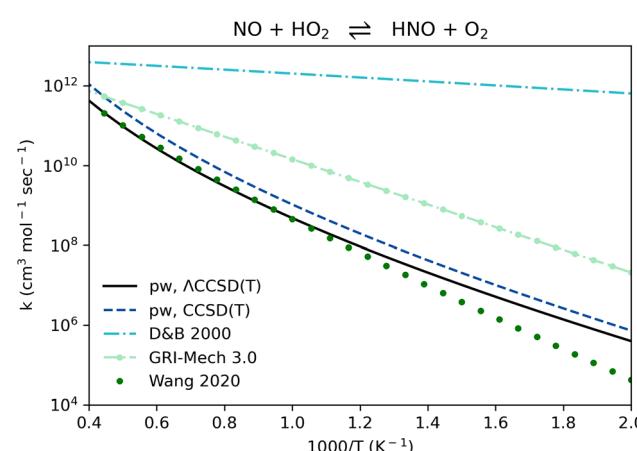


Fig. 5 Comparison of rate coefficients for  $NO + HO_2 \rightleftharpoons HNO + O_2$  (reaction R3, Table 1) computed here (pw) using ACCSD(T)/aug-cc-pVTZ//B2PLYPD3/aug-cc-pVTZ and using CCSD(T)-F12a/aug-cc-pVTZ-F12//B2PLYPD3/aug-cc-pVTZ, compared to an estimation by Dean and Bozzelli 2000,<sup>12</sup> to the GRI-Mech 3.0 value,<sup>83</sup> and to a CCSD(T)/CBS//aug-cc-pVTZ//B2PLYPD3/aug-cc-pVTZ computation by Wang *et al.* 2020.<sup>84</sup>

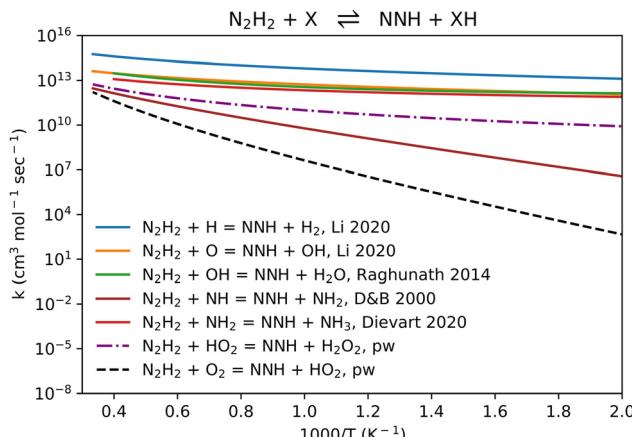


Fig. 6 Comparison of reaction rate coefficients of various radicals abstracting a hydrogen atom from  $\text{N}_2\text{H}_2$  forming NNH. Sources: Li and Sarathy 2020,<sup>85</sup> Raghuanth et al. 2014,<sup>86</sup> Dean and Bozzelli 2000,<sup>12</sup> Diévert and Catoire 2020,<sup>87</sup> and two reactions from the present work (pw). The reactions  $\text{N}_2\text{H}_2 + \text{HO}_2 \rightleftharpoons \text{NNH} + \text{H}_2\text{O}_2$  and  $\text{N}_2\text{H}_2 + \text{O}_2 \rightleftharpoons \text{NNH} + \text{HO}_2$  are shown in Table 1 (reactions R10 and R5 in reverse, respectively).

coefficients of its sister reactions (Fig. 6). Nonetheless, reaction R5 could be important for modeling the ammonia oxidation system: although its rate coefficient at low temperatures is orders of magnitude lower than the parallel pathways in the comparison (Fig. 6), the high concentration of  $\text{O}_2$  is orders of magnitude higher than the concentration of all other attacking radicals in typical  $\text{NH}_3$  oxidation systems. Therefore, the rate of reaction R5 could become significant at temperatures higher than 1000 K, especially in lean fuel or stoichiometric combustion mixtures.

Reaction R6,  $\text{NH}_2 + \text{HO}_2 \rightleftharpoons \text{NH} + \text{H}_2\text{O}_2$ , also only becomes significant in the high temperature range relative to the other consumption channels of  $\text{NH}_2 + \text{HO}_2$  (Fig. 7). Since the  $\text{HO}_2$  radical is mostly important at the low and intermediate temperature ranges, it is reasonable to omit reaction R6 from  $\text{NH}_3$  oxidation models.

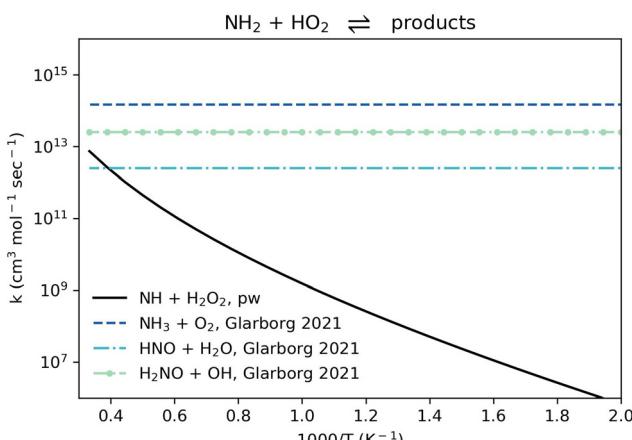


Fig. 7 Comparison of rate coefficients for  $\text{NH}_2 + \text{HO}_2$  giving  $\text{NH} + \text{H}_2\text{O}_2$  (reaction R6) computed here (pw) and giving  $\text{NH}_3 + \text{O}_2$ ,  $\text{HNO} + \text{H}_2\text{O}$ , and  $\text{H}_2\text{NO} + \text{OH}$  as recommended by Glarborg et al. 2021.<sup>88</sup>

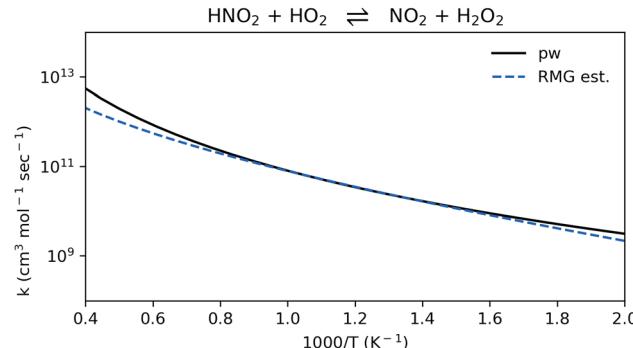


Fig. 8 Comparison of rate coefficients for the reaction  $\text{HNO}_2 + \text{HO}_2 \rightleftharpoons \text{NO}_2 + \text{H}_2\text{O}_2$  (reaction R9) computed here (pw) and compared to RMG's estimate.

Reaction R7,  $\text{NH}_3 + \text{HO}_2 \rightleftharpoons \text{NH}_2 + \text{H}_2\text{O}_2$ , is critical to predict the low- and intermediate-temperature oxidation of ammonia, since it describes the direct interaction between  $\text{NH}_3$  and the main radical carrier in this temperature range,  $\text{HO}_2$ . The endothermicity of reaction R7 is about 80 kJ mol⁻¹ (Table S2, ESI†). Its rate coefficient was recently computed at the CASPT2/aug-cc-pVTZ//M06-2X/aug-cc-pVTZ level of theory,<sup>70</sup> and the rate coefficient computed in the present work is in satisfying agreement with this reported calculation (Fig. S3, ESI†).

Reaction R8,  $\text{N}_2\text{H}_3 + \text{HO}_2 \rightleftharpoons \text{H}_2\text{NN}(\text{T}) + \text{H}_2\text{O}_2$ , has a relatively insignificant rate coefficient throughout the temperature range of interest compared to competing pathways that produce  $\text{N}_2\text{H}_2 + \text{H}_2\text{O}_2$  or  $\text{N}_2\text{H}_4 + \text{O}_2$  (Fig. S4, ESI†). Similar to reaction R6, it is also reasonable to omit reaction R8 from  $\text{NH}_3$  combustion models.

To our knowledge, the rate coefficient of reaction R9,  $\text{HNO}_2 + \text{HO}_2 \rightleftharpoons \text{NO}_2 + \text{H}_2\text{O}_2$ , was not previously reported. The RMG estimation is in good agreement with the computed value (Fig. 8). This new pathway could become important for  $\text{NO}_2$  formation in locations in the flame where the concentration of  $\text{HO}_2$  is significant.

Reaction R11,  $\text{HNOH} + \text{O}_2 \rightleftharpoons \text{HNO} + \text{HO}_2$  (listed in Table 1 in reverse for context), appears in all recent  $\text{NH}_3$  oxidation

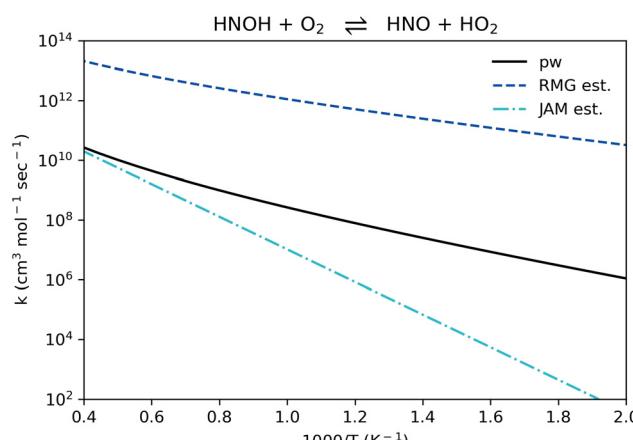


Fig. 9 Comparison of rate coefficients for the reaction  $\text{HNOH} + \text{O}_2 \rightleftharpoons \text{HNO} + \text{HO}_2$  (reaction R11, given in reverse in Table 1) computed here (pw) and compared to RMG's estimate and to J. A. Miller and P. Glarborg's estimate.<sup>89</sup>



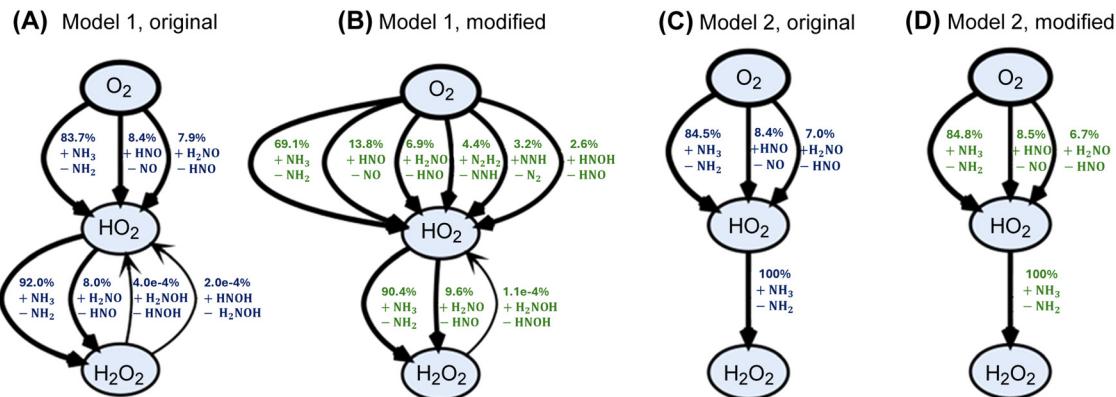


Fig. 10 Partial flux diagrams showing only the dioxigen species ( $O_2$ ,  $HO_2$ , and  $H_2O_2$ ) for "Model 1" by Zhang *et al.*<sup>51</sup> and "Model 2" by Zhu *et al.*<sup>52</sup> (A), and (C) show the fluxes obtained from the original models without modifications, (B), (D) show the fluxes obtained after adopting the rate coefficients recommended by the present work into the corresponding models (see text). Percentages correspond to the relative flux through each pathway for a stoichiometric feed of  $NH_3$  and air simulated in a continuous stirred reactor with volume of  $20\text{ cm}^3$  at  $1200\text{ K}$ ,  $10\text{ bar}$ , and a residence time of 1 second.

models examined here.<sup>11,51,66-81</sup> Its rate coefficient in all of these models was taken from a single source, an estimation made by Miller and Glarborg in 1999.<sup>89</sup> The rate coefficient of reaction R11,  $HNO + HO_2 \rightleftharpoons HNOH + O_2$ , computed in the present work falls between the 1999 estimation and the updated RMG estimation (Fig. 9). It generally follows the same temperature dependence suggested by RMG, yet it is about 3 orders of magnitude lower. The 1999 estimation is adequate only at the very high temperature range ( $> 2000\text{ K}$ ), yet it significantly under-predicts the rate coefficient at the intermediate and low temperature ranges. It is recommended to use the updated rate coefficient provided here for reaction R11 in future ammonia modeling efforts.

Fig. 10 and 11 illustrate the impact of incorporating reactions R1–R11 with the rate coefficients recommended in the present work into two recent ammonia combustion models: "Model 1" by Zhang *et al.*<sup>51</sup> and "Model 2" by Zhu *et al.*<sup>52</sup> The

modified models show significant changes in reaction pathways and species flux distributions compared to their original counterparts. In Model 1, the added reactions alter the dominant formation routes to  $HO_2$  (Fig. 10(A) and (B)). After updating Model 1, the flux through the main pathway  $NH_3 + O_2 \rightleftharpoons NH_2 + HO_2$  decreases, while new minor pathways for  $HO_2$  formation emerge involving  $N_2H_2$ ,  $NNH$ , and  $HNOH$  reactions with  $O_2$ . This change was not observed for Model 2.

The most significant changes, however, are observed in the formation of  $NNH$  radicals from  $N_2H_2$  (Fig. 11). In the original Model 1 and Model 2 simulations (Fig. 11(A) and (C)), the  $N_2H_2 \rightleftharpoons NNH$  conversion is dominated by  $NH_2$ -mediated hydrogen abstraction (56% and 86%, respectively). Upon adding R1–R11 (Fig. 11(B) and (D)), the dominant channel shifts to  $O_2$ -mediated hydrogen abstraction, resulting in  $HO_2$  formation and accounting for more than 90% of the flux in both cases. The  $NH_2$ -mediate pathway contributes less than 10% of the flux

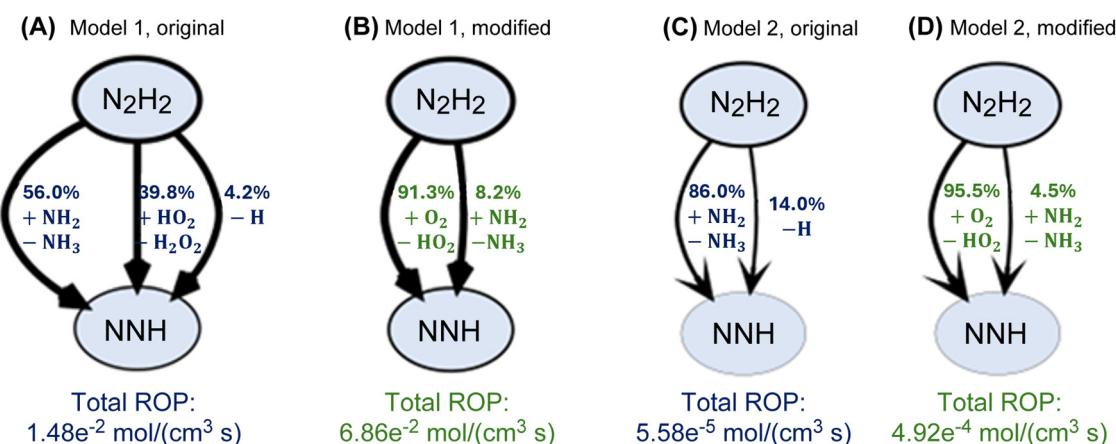


Fig. 11 Partial flux diagrams showing only major  $N_2H_2 \rightarrow NNH$  pathways for "Model 1" by Zhang *et al.*<sup>51</sup> and "Model 2" by Zhu *et al.*<sup>52</sup> (A), and (C) show the fluxes obtained from the original models without modifications, (B), (D) show the fluxes obtained after adopting the rate coefficients recommended by the present work into the corresponding models (see text). Percentages correspond to the relative flux through each pathway for a stoichiometric feed of  $NH_3$  and air simulated in a continuous stirred reactor with volume of  $20\text{ cm}^3$  at  $1200\text{ K}$ ,  $10\text{ bar}$ , and a residence time of 1 second. ROP – net rate of production.

at the examined conditions. The total net rate of production of NNH from  $\text{N}_2\text{H}_2$  increases by  $\sim 5\times$  and  $\sim 10\times$  after updating the two examined models (Fig. 11). This large shift in flux is attributed to reaction R5.

These results demonstrate that the incorporation of reactions R1–R11 enables both models to capture new mechanistic pathways that were previously underrepresented. Since ammonia modeling is an ongoing effort in the community and recent models still require significant updates,<sup>90</sup> we did not perform a thorough kinetic analysis here. The representative conditions chosen here (1200 K, 10 bar, 1 second in Fig. 10 and 11) show the potential impact of incorporating updated  $\text{HO}_2$  kinetics in  $\text{NH}_3$  modeling in future models.

## 4. Conclusions

This work reports the identification of a subset of reactions automatically suggested during chemical kinetic model generation with respective rate coefficients automatically computed using *ab initio* methods. While not all reactions were determined as significant after quantum chemical computations, the potential importance of some of the identified reactions was shown. This automated approach of using software such as RMG to automatically suggest reactions that could be significant has some shortcomings since RMG's rate coefficient estimations could be inaccurate for reactions for which training data is scarce. Nonetheless, this approach was successful in highlighting previously overlooked reactions that might be important for modeling the  $\text{NH}_3$  oxidation system.

Bimolecular reactions involving  $\text{HO}_2$  in the  $\text{NH}_3$  oxidation system, most previously unexplored (reactions R1, R2, R4, R5, R6, R8, R9), were studied and their rate coefficients were computed at the ACCSD(T)/cc-pVTZ-F12//B2PLYP-D3/aug-cc-pVTZ level of theory. This level of theory appears to evaluate the reaction wells to within reasonable accuracy, averaging at about  $-0.6 \text{ kJ mol}^{-1}$ . This method also claims the static correlation to some extent while being an SR method, which could reduce the effects the relatively large T1 diagnostic coefficient indicates. In addition, high spin systems are often misrepresented by CCSD(T), and are better described by post-CCSD(T) methods, such as the one shown in this paper.

Previous literature studies examined the  $\text{NHO}_2$  system primarily from the  $\text{NH} + \text{O}_2$  direction and determined that  $\text{NO} + \text{OH}$  and  $\text{HNO} + \text{O}$  are the major products. In this direction, the formation of  $\text{N} + \text{HO}_2$  is indeed insignificant below 2000 K. However, in the reverse direction,  $\text{N} + \text{HO}_2 \rightleftharpoons \text{NH} + \text{O}_2$ , the rate coefficient of reaction R1 is significant, and the reaction is highly exothermic.

Reaction R2,  $\text{NH} + \text{HO}_2 \rightleftharpoons \text{NH}_2 + \text{O}_2$ , which might be an important channel of  $\text{NH}_2 + \text{O}_2$  reactivity, was not assigned a rate coefficient in this study due to the severe multireference character of the transition state and the resulting inability of our chosen methods to estimate its activation barrier. While  $\text{O}_2$  was found to react relatively slowly with  $\text{N}_2\text{H}_2$  to form NNH, an ROP analysis showed that reaction R5,  $\text{NNH} + \text{HO}_2 \rightleftharpoons \text{N}_2\text{H}_2 + \text{O}_2$ , is

pivotal for the  $\text{N}_2\text{H}_2 \rightarrow \text{NNH}$  transformation. An updated rate coefficient for reaction R11,  $\text{HNOH} + \text{O}_2 \rightleftharpoons \text{HNO} + \text{HO}_2$ , is given in replacement of an estimated value used throughout the recent  $\text{NH}_3$  oxidation models. Incorporating the updated rate coefficient for reaction R11 is expected to result in a shift of the dominant pathway for the conversion of  $\text{HNOH}$  into  $\text{HNO}$ , increasing the overall flux of converting  $\text{HNOH}$  into  $\text{HNO}$ .

The incorporation of the computed rate coefficients into two recent literature models demonstrated the mechanistic relevance of several of these reactions. The dominant  $\text{N}_2\text{H}_2 \rightarrow \text{NNH}$  channel shifted from  $\text{NH}_2$ -mediated to  $\text{O}_2$ -mediated pathways, increasing the rate of NNH formation by up to an order of magnitude, while fluxes across species involved in  $\text{HO}_2$  formation were redistributed in one of the tested models. The flux analysis demonstrates the potential of incorporating these updated  $\text{HO}_2$ -involving reactions in future modeling efforts.

It is recommended to consider at least reactions R1, R3, R5, R7, R9, R10, and R11 in future relevant  $\text{NH}_3$  chemical kinetic oxidation models, and a rate coefficient for reaction R2 should further be computed at an appropriate MR level. Additional reactions from the present work could be relevant if relatively high temperature ranges are of interest.

## Author contributions

K. K.: formal analysis and methodology, M. K.: formal analysis and methodology, A. G. D.: Formal analysis, methodology, conceptualization, funding acquisition, and supervision.

## Conflicts of interest

There are no conflicts of interest to declare.

## Data availability

The data supporting this article have been included as part of the ESI.<sup>†</sup>

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