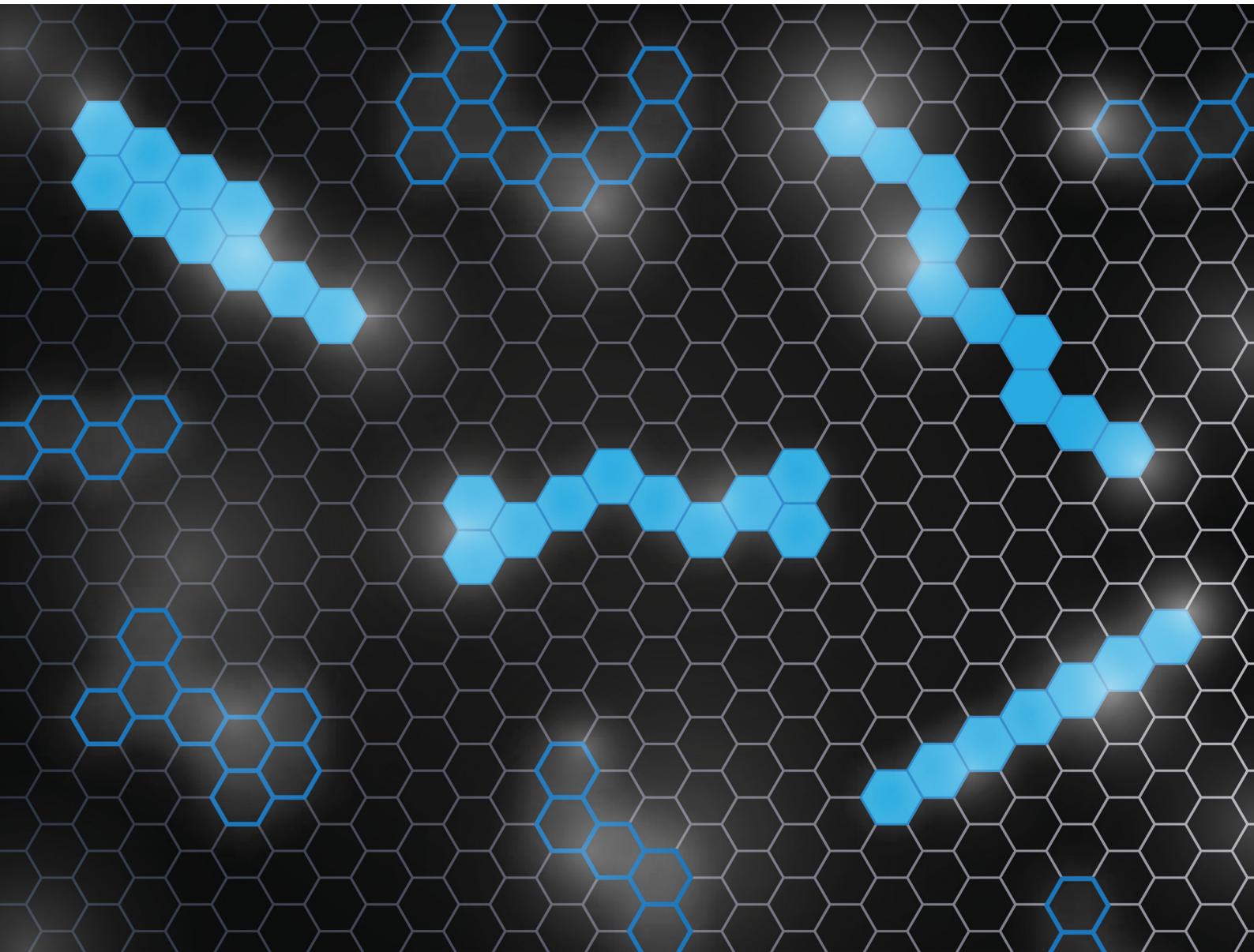


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of polybenzenoid hydrocarbons using xTB fractional
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Accelerated diradical character assessment in large datasets of polybenzenoid hydrocarbons using xTB fractional occupation†

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Polybenzenoid hydrocarbons (PBHs) have garnered significant attention in the field of organic electronics due to their unique electronic properties. To facilitate the design and discovery of new functional organic materials based on these compounds, it is necessary to assess their diradical character. However, this usually requires expensive multireference calculations. In this study, we demonstrate rapid identification and quantification of open-shell character in PBHs using the fractional occupation number weighted electron density metric (N_{FOD}) calculated with the semiempirical GFN2-xTB (xTB) method. We apply this approach to the entire chemical space of PBHs containing up to 10 rings, a total of over 19k molecules, and find that approximately 7% of the molecules are identified as having diradical character. Our findings reveal a strong correlation between xTB-calculated N_{FOD} and the more computationally expensive Yamaguchi y and DFT-calculated N_{FOD} , validating the use of this efficient method for large-scale screening. Additionally, we identify a linear relationship between size and N_{FOD} value and implement a size-dependent threshold for open-shell character, which significantly improves the accuracy of diradical identification across the chemical space of PBHs. This size-aware approach reduces false positive identifications from 6.97% to 0.38% compared to using a single threshold value. Overall, this work demonstrates that xTB-calculated N_{FOD} provides a rapid and cost-effective alternative for large-scale screening of open-shell character in PBHs.

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Introduction

Organic compounds displaying diradical character have long fascinated chemists, as they provide unique platforms for investigating the nature of the chemical bond and other fundamental organic concepts. Classic examples include Tschitschibabin's hydrocarbon, bisphenalenyls, zethrenes, and indenofluorenes (Fig. 1), which have all been extensively investigated both computationally^{1–7} and experimentally.^{2–4,8–12} It is unsurprising that the majority of long-lived, synthetically accessible organic compounds with open-shell character are polycyclic aromatic hydrocarbons (PAHs). In such systems, termed “pro-aromatic” by Wu and coworkers,¹³ the closed-shell resonance structure (RS)

is quinoidal whereas the open-shell RS contains additional disjoint 6- π electrons rings (compare the left and right RSs in Fig. 1; the diradical RSs contain additional Clar sextets). Hence, the increase in aromaticity gained by breaking the double bond stabilizes—and perhaps even drives—the adoption of diradical character.¹⁴

Among such PAHs, polybenzenoid hydrocarbons (PBHs, cutouts of a graphene sheet) have emerged as especially promising for organic electronics. When endowed with significant open-shell character, PBHs demonstrate potential for a broad range of such uses, including light-emitting diodes, batteries, field-effect transistors, and photovoltaics.¹⁵ As the diradical character of a molecule is directly linked to its physical properties (optical, electronic, and magnetic), the identification and quantitative evaluation of this character is crucial to understanding and predicting molecular behavior, as well as to designing new functional compounds.

Larger PBHs in general, and polyacenes (linearly annulated PBHs) in particular, are considered to have a tendency toward open-shell ground states.^{16,17} In contrast, the majority of small PBHs have closed-shell electron configurations. Nevertheless, open-shell configurations—*i.e.*, radical or diradical character—can occur in three scenarios (Fig. 2): (a) an odd number of hydrogens/carbons (*e.g.*, phenalenyl radical, $C_{13}H_9$, is a

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† Electronic supplementary information (ESI) available: General computational details, further analysis of the effect of temperature on the N_{FOD}^{xTB} value, examples of mistaken classification, the effect of the longest linear stretch in cc-PBHs, the effect of size, comparison of normalization approaches (number of heavy atoms *versus* number of electrons), investigation of y_1 values, and the effect of geometry; tables and figures for the identification of the size-aware thresholds, and analysis on the effect of geometry on N_{FOD}^{xTB} . See DOI: <https://doi.org/10.1039/d4cp04059g>



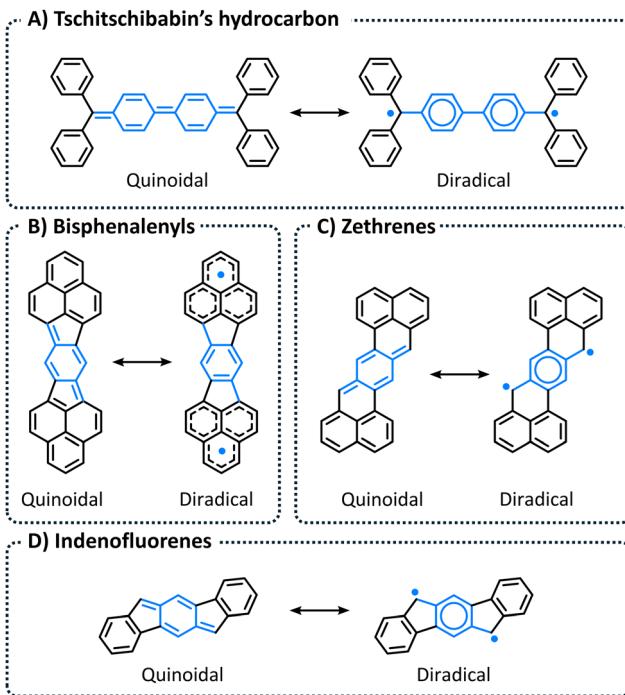


Fig. 1 Examples of PAHs known to have appreciable diradical character: (A) Tschitschibabin's hydrocarbon; (B) Bisphenalenyls; (C) Zethrenes; and (D) Indenofluorenes. In all cases, left: quinoidal RS, right: diradical RS.

three-ring *peri*-condensed (pc)-PBH with a single unpaired electron); (b) non-Kekuléan structures, *i.e.*, PBHs for which no classical closed-shell valence structure can be drawn^{18,19} (*e.g.*, triangulene, $C_{22}H_{12}$, is a non-Kekuléan six-ring pc-PBH with two unpaired electrons in the ground state); and (c) molecules

that possess a closed-shell resonance structure, but have appreciable diradical character (*e.g.*, longer polyacenes, zethrenes, periacycenes).¹⁹

The first two cases are relatively easily identified. It is the third scenario that is most challenging, as the open-shell character is not implied by either the molecular or the structural formula. Thus, reliable methods for identification of open-shell character are required for these molecules.

However, such assessment is not trivial because open-shell PAHs often have multireference (MR) character, meaning they are described by two or more dominant contributing electronic configurations. Hence, to adequately describe the wavefunction of the molecule, MR methods are needed.^{17,20–24} In addition to their high computational cost and their molecular-size limitations, such methods usually require a high level of user expertise and are not easily accessible to non-specialists. Therefore, density functional theory (DFT) remains the most prevalent method, despite the inherent difficulty of choosing the most appropriate functional. To simplify the evaluation of radical and diradicaloid character, Lischka and coworkers recently benchmarked various DFT-derived descriptors and showed that the number of “hot” electrons, N_{FOD} , obtained from the fractional occupation number weighted electron density (FOD), provides a fast and reliable assessment of different classes of PAHs, without the need to perform MR calculations.²⁵ In their report, the authors used thermally-assisted DFT to obtain the N_{FOD} values for a test set of 29 polycyclic aromatic hydrocarbons, which they compared to the numbers of unpaired electrons (N_U) obtained from MR-averaged quadratic coupled-cluster (MR-AQCC). Additional systems were investigated in a subsequent report, published very recently.²⁶ Their work highlights both the ongoing interest in PAHs and the need for inexpensive and user-friendly methods for evaluating diradical character. Indeed, they conclude their report with the statement “This finding opens the possibility of large scale and reliable screening of PAH biradical properties, which is expected to have a significant impact on the PAH research field.”‡

Herein, we build upon the work of Lischka and coworkers in two directions. First, we continue the process of simplification by demonstrating that N_{FOD} values obtained from a semiempirical method (namely, GFN2-xTB; referred to as xTB in this text for conciseness) can be used for this evaluation, rather than DFT. We note that a very recent report²⁶ by the same group also demonstrated the use of xTB N_{FOD} for various PAHs in the context of bond dissociation energies. Second, using this computationally efficient tool, we embark on the large-scale screening they envisioned by assessing the diradical character of all isomers in the chemical space of PBHs containing up to 10 rings.

The work described in this manuscript is detailed in the following sections: (a) the data used in the study, (b) the open-

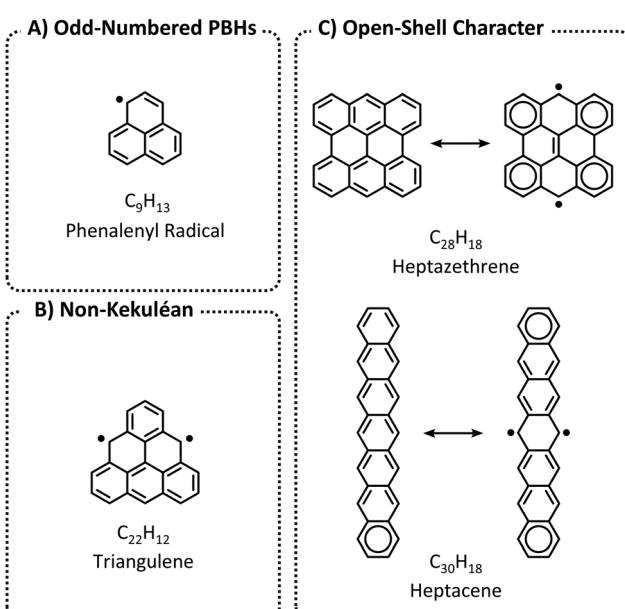


Fig. 2 Three possible scenarios for PBHs to obtain diradical character: (A) Odd-numbered PBHs; (B) Non-Kekuléan PBHs; and (C) PBHs with open-shell character.

‡ For clarity, please note that we use the terminology introduced in the IUPAC Gold-Book²⁷ and detailed by Abe.²⁸ Accordingly, ‘biradical’ refers to molecules in which the two electrons act independently or nearly independently, whereas ‘diradical’ refers to species in which the magnitude of the dipole-dipole interaction is large enough to produce two spin states, namely, singlet ($S, S = 0$, spin multiplicity = 1) and triplet ($T, S = 1$, spin multiplicity = 3). For PBHs, the latter term is more suitable.



shell character descriptors employed, (c) the benchmarking procedure from which a threshold value was identified and verified, (d) the process developed to make the approach transferable across molecules of various sizes, and (e) comparison of xTB- and DFT-calculated N_{FOD} .

Data and descriptors

Data

Polybenzenoid hydrocarbons (PBHs) are made up of only one type of ring—benzene. Their structural diversity, therefore, stems from variations in the number of fused rings and in the patterns of fusion that are possible. In the context of diradical character, the most extensively studied systems are the extended polyacenes, zethrenes, and periacenes. Several computational investigations have focused on these types of scaffolds, using a variety of methods, including unrestricted DFT,^{1,29} CASSCF,^{21,22} restricted active space configuration interaction,^{20,24} spin-flip techniques,^{20,30} and MR-AQCC.^{17,23} In this work, we sought to extend beyond these recognized structures and explore the chemical space of PBHs more broadly, therefore we investigate the entire chemical space of PBHs comprising up to 10 rings (*i.e.*, all isomers containing $n_{\text{rings}} \leq 10$).

For the purpose of this study, we divided this chemical space into various groups, as depicted in Fig. 3. First, we distinguished between *cata*-condensed (cc-PBHs) and *peri*-condensed (pc-PBHs). In cc-PBHs, a single carbon may only participate in up to two rings; in pc-PBHs, at least one carbon participates in three rings. Second, we further divided each of the subgroups into two groups: cc-PBHs—Groups A and B, pc-PBHs—Groups C and D. For the cc-PBHs, the division was performed using a structural descriptor: the longest linear stretch within the molecule, n_{LL} . Bendikov *et al.*¹⁶ and Yang *et al.*²⁹ independently identified hexacene as the smallest polyacene to show indications of multireference character. This

was corroborated by the experimental observations of instability of the acenes larger than pentacene.^{31–34} Accordingly, we divided our data such that Group A contained molecules with $n_{\text{LL}} < 6$, which were expected to be closed-shell, and Group B contained molecules with $n_{\text{LL}} \geq 6$, which were expected to be diradical. For the pc-PBHs, the division was performed using an $N_{\text{FOD}}^{\text{xTB}}$ -based threshold, which we identified in the course of this study ($N_{\text{FOD}}^{\text{xTB}} = 1.2$, *vide infra*). For this subset of data, Group C contained the molecules with $N_{\text{FOD}}^{\text{xTB}} < 1.2$ and Group D contained the molecules with $N_{\text{FOD}}^{\text{xTB}} \geq 1.2$.

As stated above, in this work we explored the entire chemical space of PBHs with $n_{\text{rings}} \leq 10$. The majority of molecular geometries were extracted from the COMPAS (COMputational database of Polycyclic Aromatic Systems) datasets.^{35,36} The COMPAS datasets were designed to focus on closed-shell molecules and, accordingly, COMPAS-1D closely corresponds to our Group A and COMPAS-3D closely corresponds to our Group C. All additional molecules were calculated in the course of the current investigation to ensure coverage of the entire chemical space.

To remain consistent with the COMPAS datasets, we calculated Groups B and D at the same level of theory. Thus, all geometry optimizations were performed with ORCA version 5.0.3,^{37,38} at the (U)CAM-B3LYP^{39–43}/def2-SVP⁴⁴ level, using Grimme's D3 dispersion correction⁴⁵ with Becke–Johnson damping.^{46,47} Groups A and C were optimized with restricted Kohn–Sham DFT (RKS), while Groups B and D were optimized with unrestricted KS DFT (UKS).

Descriptors

We benchmarked the N_{FOD} values against two widely used methods for assessment of open-shell character: the DFT-derived N_{FOD} and the unrestricted Hartree–Fock (UHF)-derived Yamaguchi y value. Both are described below.

The FOD method. N_{FOD} values were obtained through fractional occupation number weighted electron density (FOD) analysis, a method introduced by Grimme and Hansen⁴⁸ to enable a real-space measure and visualization of static electron correlation effects. The FOD method simplifies the concept behind Chai's thermally-assisted occupation density functional theory (TAO-DFT). TAO-DFT employs a fictitious temperature in the reference system to generate fractional orbital occupancies using the Fermi–Dirac distribution, showing a connection between static correlation and entropy. By disregarding the fictitious temperature-dependent energy functionals, one obtains 'finite-temperature DFT' (FT-DFT). Put simply, the FOD method essentially 'smears' electrons over the molecular orbitals (MOs) in a Fermi–Dirac distribution. The fractional occupation numbers of the electrons in the MOs ($0 \leq f_i \leq 1$) are then summed to give ρ^{FOD} :

$$\rho^{\text{FOD}}(\mathbf{r}) = \sum_i^N (\delta_1 - \delta_2 f_i) |\phi_i^2(\mathbf{r})| \quad (1)$$

where f_i are the fractional occupation numbers ($0 \leq f_i \leq 1$) and the sum is taken over all the molecular spin orbitals ϕ_i . The constants δ_1 and δ_2 are equal to 1 if the energy level is lower

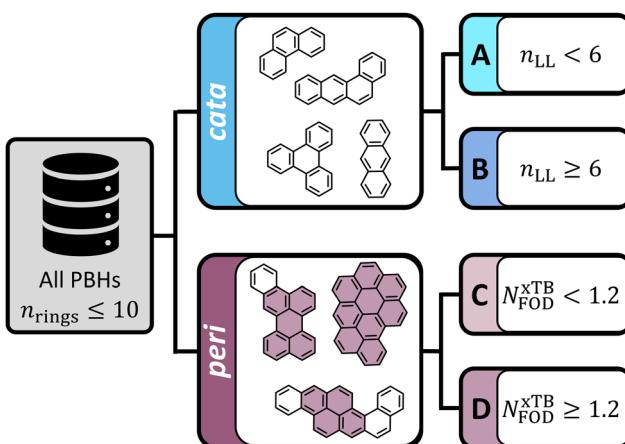


Fig. 3 Overview of the data. All isomers of $n_{\text{rings}} \leq 10$ PBHs were included, divided into *cata*- and *peri*-condensed subsets (representative examples are shown). The cc-PBHs were divided into Groups A and B according to the structural parameter n_{LL} . The pc-PBHs molecules were divided into Groups C and D according to an $N_{\text{FOD}}^{\text{xTB}}$ threshold.



than the Fermi level and equal to 0 and -1 , respectively, if the energy level is higher. Under these definitions, only the fractionally-occupied ϕ orbitals contribute to the final value. For these orbitals, the fractional occupation numbers are obtained from the Fermi-Dirac distribution (see ref. 48 for further details). For open-shell singlet cases, the f_i values of each orbital is the sum of α - and β -shell contributions. The integration of ρ^{FOD} over the entire space affords N_{FOD} , a single size-extensive number:

$$N_{\text{FOD}} = \int \rho^{\text{FOD}}(\mathbf{r}) \, d\mathbf{r} \quad (2)$$

The main advantage of N_{FOD} , as highlighted in Lischka and coworkers' recent report,²⁵ is its low cost, which carries an even more substantial benefit when working with large datasets. This advantage can be further exploited by calculating the N_{FOD} value with xTB, a semiempirical method, thereby accelerating calculations by three orders of magnitude. An additional advantage has been pointed out by Chan, who showed that $N_{\text{FOD}}^{\text{xTB}}$ can be used to identify potentially problematic components in a larger system, whereas many MT diagnostics are formulated for small systems and do not provide accurate characterization of large systems.⁴⁹ However, N_{FOD} also has two main disadvantages. The first is that, similarly to most MR diagnostics, there is no well-defined threshold for identifying open-shell character unambiguously. Nevertheless, previous work by Grimme and coworkers⁵⁰ did indeed find a correlation between N_{FOD} and y values on a small test set, demonstrating the possibility to use N_{FOD} when it is benchmarked against a less ambiguous diagnostic. The second disadvantage is that N_{FOD} is size-extensive, which makes it difficult, if not impossible, to define a universal threshold applicable to molecules of varying sizes. In this work, we shed light on the size-dependency of N_{FOD} and provide a framework for alleviating this problem in the context of PBHs.

In this work, we calculated the N_{FOD} values with both DFT ($N_{\text{FOD}}^{\text{DFT}}$) and xTB ($N_{\text{FOD}}^{\text{xTB}}$). For the DFT calculations, we used the B3LYP functional with the def2-SVP basis set and set the temperature to 9000 K, as recommended by Grimme and coworkers,⁵¹ based on the empirical formula $T_{\text{el}} = 20\,000 \text{ K} \times \alpha_x + 5000 \text{ K}$, where α_x is the Fock exchange percentage of the functional ($\alpha_x = 20\%$ for B3LYP). For the xTB calculations, we set the temperature to 5000 K, as recommended by Grimme and coworkers.⁵¹ We note that ref. 26, which was published after completion of this work, recommends a temperature of 5200 K for xTB calculations. See Section S2 of the ESI† for a comparison of the results obtained with the two temperatures.

Yamaguchi y value. The Yamaguchi y value⁵² is given by the following equation:^{52,53}

$$y_i = 1 - \frac{4(n_{\text{HONO}-i} - n_{\text{LUNO}+i})}{4 + (n_{\text{HONO}-i} - n_{\text{LUNO}+i})^2} \quad (3)$$

where $n_{\text{HONO}-i}$ and $n_{\text{LUNO}+i}$ are the occupation numbers of the i th highest occupied natural orbital (HONO) and the i th lowest unoccupied natural orbital (LUNO).

In this work, we used only the y_0 values (see Section S7 of the ESI† for additional discussion on the y_1 values). Hence, for conciseness, we use the simple symbol y throughout.

Although the y value also does not have a well-defined threshold for identifying open-shell character, it is one of the most commonly used diagnostics in the organic chemistry community. This is likely due to the fact that it uses a $[0,1]$ scale. For a 'perfect' closed-shell system, the occupation of the HONO is 2 and the occupation of the LUNO is 0; accordingly, $y_0 = 0$. For a 'perfect' diradical system, the occupation of both the HONO and the LUNO is 1, resulting in $y_0 = 1$. The 0%-to-100% scale makes it very intuitive and straightforward to interpret. Because it is most often applied using the UHF orbital approach, this metric is rather inexpensive. Nevertheless, it requires broken symmetry calculations, which may become prohibitively expensive when applied to large molecules and/or large datasets. For context, in this investigation, we calculated the y values of 19 875 molecules, with a median size of 62 atoms (median of 40 carbon atoms), using the UHF method and the def2-SVP basis set. The cumulative CPU time for these calculations was $\sim 178\,650$ hours. In comparison, $N_{\text{FOD}}^{\text{xTB}}$ calculations required ~ 5 CPU hours and $N_{\text{FOD}}^{\text{DFT}}$ calculations required ~ 7810 CPU hours.

Results and discussion

Benchmarking

Identification of threshold value. To validate the choice of methods and obtain a relevant frame of reference for the studied molecules, we began with a benchmarking procedure. Our benchmarking test set included benzene, naphthalene, and all isomers (with $3 \leq n_{\text{rings}} \leq 10$) of two types of cc-PBHs: (a) polyacenes (molecules containing linearly annulated benzenes), which are known to have significant diradical character starting from $n_{\text{rings}} = 6$,²⁹ and (b) polyphenacenes (molecules containing benzenes annulated in a zig-zag fashion), which are known to not have any significant diradical character, even for the larger systems.⁵⁴ The $N_{\text{FOD}}^{\text{xTB}}$, $N_{\text{FOD}}^{\text{DFT}}$, and y results for all molecules in the benchmarking set are presented in Table 1.

As expected, for every n_{rings} , the acene isomer showed a higher value than the corresponding phenacene, for all three metrics. Additionally, although both families of compounds showed an

Table 1 $N_{\text{FOD}}^{\text{xTB}}$, $N_{\text{FOD}}^{\text{DFT}}$, and y values obtained for the benchmarking molecules, listed by n_{rings}

n_{rings}	Polyacenes			Polyphenacenes		
	$N_{\text{FOD}}^{\text{xTB}}$	$N_{\text{FOD}}^{\text{DFT}}$	y (%)	$N_{\text{FOD}}^{\text{xTB}}$	$N_{\text{FOD}}^{\text{DFT}}$	y (%)
1	0.03	0.11	1	—	—	—
2	0.12	0.27	4	—	—	—
3	0.32	0.53	13	0.18	0.40	5
4	0.58	0.87	26	0.27	0.54	7
5	0.89	1.25	39	0.34	0.68	7
6	1.22	1.65	51	0.42	0.82	9
7	1.56	2.07	61	0.50	0.96	9
8	1.91	2.48	69	0.57	1.11	9
9	2.25	2.90	76	0.65	1.25	9
10	2.60	3.32	82	0.73	1.39	10



increase in open-shell character as n_{rings} increases, this behavior was much more pronounced for the polyacenes. For the polyacenes, open-shell character became dominant (*i.e.*, $y > 50\%$) from $n_{\text{rings}} = 6$ (specifically, $y = 51\%$ for hexacene). In contrast, all polyphenacenes had negligible y values ($\leq 10\%$). For both families of compounds, the results were in agreement with experimental data,^{55,56} which validated our choice of the y value as a benchmarking metric for the N_{FOD} value.

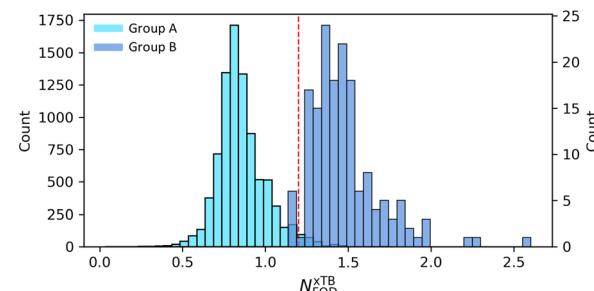
We set $y = 50\%$ as the transition point between predominant closed-shell and predominant open-shell character, which made hexacene the first in the series to cross that threshold, as expected. The N_{FOD} values for hexacene were $N_{\text{FOD}}^{\text{XTB}} = 1.22$ and $N_{\text{FOD}}^{\text{DFT}} = 1.65$. Therefore, we set the $N_{\text{FOD}}^{\text{XTB}}$ threshold at 1.2, meaning that molecules with $N_{\text{FOD}}^{\text{XTB}} \geq 1.2$ were considered to be predominantly open-shell.

Verification of threshold on large data. To verify the generality of this benchmark, we proceeded to calculate the $N_{\text{FOD}}^{\text{XTB}}$ values for the entire chemical space of cc-PBHs with $n_{\text{rings}} \leq 10$. Based on the insight from the benchmarking analysis, we used the longest linear stretch as a structural parameter, dividing this chemical space into two groups around the threshold of $n_{\text{LL}} = 6$. As detailed in the Data section, Group A comprised molecules containing longest linear stretches of $n_{\text{LL}} < 6$ rings (expected to be closed-shell), and Group B comprised molecules with $n_{\text{LL}} \geq 6$ rings (expected to have dominant diradical character). Plotting the distributions of the $N_{\text{FOD}}^{\text{XTB}}$ values for these two populations (Fig. 4A) revealed that the two groups were clearly distinct from one another. Specifically, the molecules of Group A were indeed predominantly closed-shell (centered around $N_{\text{FOD}}^{\text{XTB}} = 0.83$, with a maximal value of $N_{\text{FOD}}^{\text{XTB}} = 1.74$), the majority (97.52%) had $N_{\text{FOD}}^{\text{XTB}} < 1.2$, and only a small number of molecules (212 out of 8556, 2.48%) had $N_{\text{FOD}}^{\text{XTB}} \geq 1.2$. In contrast, Group B showed pronounced open-shell character (centered around $N_{\text{FOD}}^{\text{XTB}} = 1.44$, with a maximal value of $N_{\text{FOD}}^{\text{XTB}} > 2.6$), and all 161 molecules had $N_{\text{FOD}}^{\text{XTB}} \geq 1.2$. Overall, these results validated the use of the structural parameter $n_{\text{LL}} > 6$ as an indicator of diradical character. They also strongly implied that, for cc-PBHs, open-shell character is (only) achievable by incorporating long linear stretches. However, it was also apparent that using only the structural parameter led to mistaken classification of some diradical molecules. Representative examples of molecules wrongly identified as closed- or open-shell are presented in Section S3 of the ESI.[†]

We then plotted distributions of the same $N_{\text{FOD}}^{\text{XTB}}$ values and the same groupings (Fig. 4B; Group A in light blue, Group B in dark blue), this time against n_{rings} . For each n_{rings} , we denoted the values for the respective acene (circle) and phenacene (triangle). Interestingly, we observed that these two were the 'bookends' of the distributions, with the acene consistently marking the highest $N_{\text{FOD}}^{\text{XTB}}$ value and the phenacene marking the lowest. The colored areas between these min and max values denoted the range of $N_{\text{FOD}}^{\text{XTB}}$ values for each n_{rings} and the bold lines within the areas denoted the mean values. This plot clearly showed that the diradical character increases with molecular size. It also provided additional validation that our

Threshold verification on cc-PBHs

A) Histograms of $N_{\text{FOD}}^{\text{XTB}}$ values



B) Distributions of $N_{\text{FOD}}^{\text{XTB}}$ values vs. n_{rings}

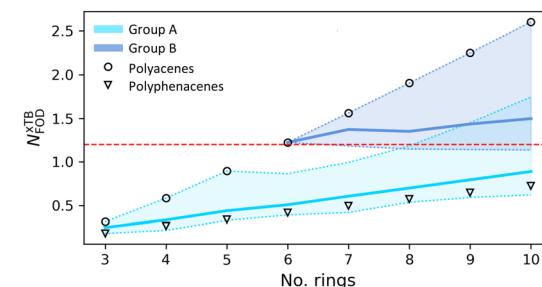


Fig. 4 Threshold verification on cc-PBH data. (A) Histograms of $N_{\text{FOD}}^{\text{XTB}}$ for the entire chemical space of cc-PBHs, divided into Group A ($n_{\text{LL}} < 6$, 8556 molecules, light blue) and Group B ($n_{\text{LL}} \geq 6$, 161 molecules, dark blue). Each histogram is normalized such that $\sum \text{Area} = 1$. (B) Distributions of Groups A and B versus n_{rings} . The colored areas show the ranges of the $N_{\text{FOD}}^{\text{XTB}}$ values (min to max). The bold lines denote the respective mean values. The values for the polyacenes (circles) and polyphenacenes (triangles) are marked.

choice of benchmarking molecules was appropriate. Moreover, it is yet further indication that the linear stretches are the source of diradical character (further discussion and analysis of the role of the linear stretches is provided in Section S4 of the ESI[†]).

As a final evaluation, we plotted a histogram of the y values for Group A. As expected, Fig. 5A showed that all compounds in this group had $y \leq 0.5$, which served as our ground truth threshold value.

Overall, these analyses affirmed our choice of the $N_{\text{FOD}}^{\text{XTB}} = 1.2$ threshold value and provided insight into the structure–property relationship governing the open-shell character of cc-PBHs.

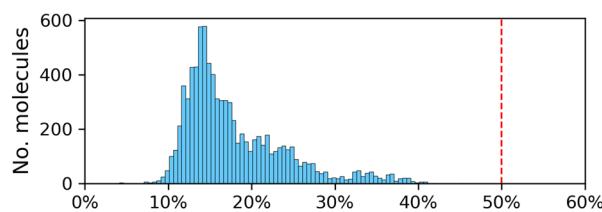
Transferability of the approach

Having confirmed the validity of our approach and the $N_{\text{FOD}}^{\text{XTB}}$ threshold value for the cc-PBHs, we next turned to testing their transferability by applying them to the pc-PBH ($n_{\text{rings}} \leq 10$) dataset. Analogously to the cc-PBH division, we divided the pc-PBH dataset into two groups, as well, this time using the threshold value we had identified for cc-PBHs, rather than a structural parameter. Accordingly, Group C comprised pc-PBHs with $N_{\text{FOD}}^{\text{XTB}} < 1.2$ and Group D comprised pc-PBHs with $N_{\text{FOD}}^{\text{XTB}} \geq 1.2$. Based on this division, Group C molecules were expected to be closed-shell. To test this, we calculated the y values for all



Validation with y values

A) Group A (cc-PBHs)



B) Group C (pc-PBHs)

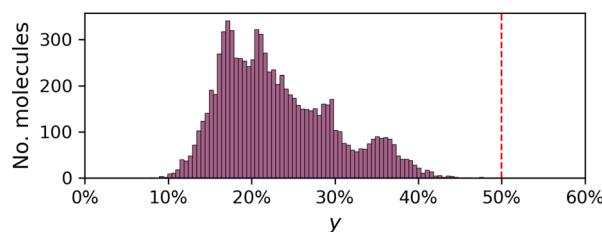


Fig. 5 Histograms of the y values for: (A) Group A (8556 molecules); (B) Group C (8701 molecules).

8701 molecules in this group. As seen in the resulting histogram (Fig. 5B), the values were all below 0.5, corroborating that these were indeed closed-shell molecules.

Thus, the application to the pc-PBHs suggested that our $N_{\text{FOD}}^{\text{XTB}}$ threshold successfully filtered out molecules with appreciable open-shell character, for both types of PBHs. However, it did not indicate whether this threshold was too conservative. In other words, perhaps by setting the threshold too low, molecules with $y < 0.5$ had been mistakenly identified as predominantly open-shell. Conversely, if the threshold were too high, some molecules had been mistakenly identified as closed-shell. To obtain a more comprehensive evaluation of the transferability of the method, we needed to account for four different scenarios (assuming the y values as the ground truth):

1. False negatives (FN): molecules that $N_{\text{FOD}}^{\text{XTB}}$ identifies as not open-shell, but which are diradical according to y .
2. True negatives (TN): molecules that both $N_{\text{FOD}}^{\text{XTB}}$ and y identify as closed-shell.
3. True positives (TP): molecules that both $N_{\text{FOD}}^{\text{XTB}}$ and y identify as open-shell.
4. False positives (FP): molecules that $N_{\text{FOD}}^{\text{XTB}}$ identifies as open-shell, but which are closed-shell according to y .

Plotting y versus $N_{\text{FOD}}^{\text{XTB}}$ for the entire pc-PBH chemical space ($n_{\text{rings}} \leq 10$) allowed us to see each of these scenarios. As shown in Fig. 6A, the majority of the data were identified correctly as TP (quadrant 1) and TN (quadrant 3). There were no cases of FNs (quadrant 2), however a certain percentage of the molecules (6.97%) were falsely identified as open-shell, *i.e.*, FP (quadrant 4). Meaning, our $N_{\text{FOD}}^{\text{XTB}}$ threshold of 1.2 overestimated the open-shell character of these molecules. For context, the percentage of TP molecules in the entire chemical space is 7.23%; hence, the identification of 6.97% of the molecules as FP almost doubled the number of molecules identified as open-shell according to $N_{\text{FOD}}^{\text{XTB}}$.

Threshold Transferability

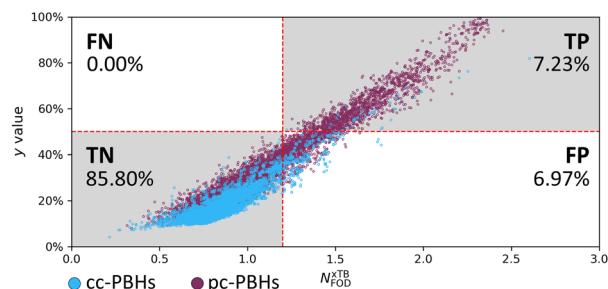
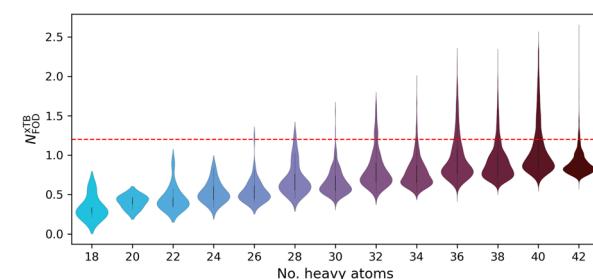
A) $N_{\text{FOD}}^{\text{XTB}}$ vs. y B) $N_{\text{FOD}}^{\text{XTB}}$ vs. n_{HA} 

Fig. 6 (A) Scatter plot of y vs. $N_{\text{FOD}}^{\text{XTB}}$ values for all PBHs (cc: blue, pc: burgundy; $n \leq 10$ rings). (B) Violin plots of the KDE distributions of all PBHs ($n_{\text{rings}} \leq 10$ rings) grouped by n_{HA} .

Thus, we concluded that although the approach had merit, it was not trivially transferable between different classes of molecules, even ones as similar as cc-PBHs and pc-PBHs! We hypothesized that the source of this issue may be the size-extensive nature of the N_{FOD} metric, as observed in both our benchmarking set and in the analysis on the cc-PBH chemical space. Therefore, we envisioned it may be possible to provide a solution by taking the molecular size into account.

The effect of size

To solve the transferability problem, we began by elucidating the relationship between molecular size and the $N_{\text{FOD}}^{\text{XTB}}$ metric. To this end, we plotted kernel density estimates (KDE) distributions of the $N_{\text{FOD}}^{\text{XTB}}$ values for all molecules (Groups A–D combined), grouped by the number of heavy atoms, n_{HA} (the numbers of molecules in each group are detailed in Table S3 in Section S5 of the ESI†). We observed a clear upward trend, whereby larger molecules had higher $N_{\text{FOD}}^{\text{XTB}}$ values (Fig. 6B), and further noted that the majority of the data were below the chosen threshold of $N_{\text{FOD}}^{\text{XTB}} = 1.2$.

Based on this, we concluded that a different threshold should be identified for each heavy-atom group, *i.e.*, a *size-aware* threshold. To identify the most suitable threshold for each group, we plotted y against $N_{\text{FOD}}^{\text{XTB}}$ and fit individual regression lines for the data within each respective group (Fig. 7A; panel B shows a zoomed-in section of the plot for additional clarity). Groups containing fewer than 3 molecules



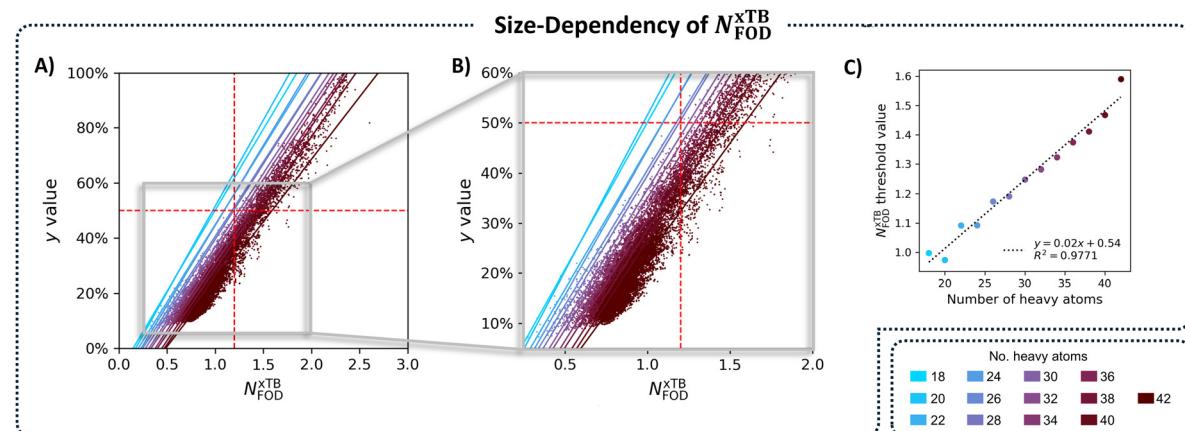


Fig. 7 (A) Scatter plot of $N_{\text{FOD}}^{\text{xTB}}$ against y , colored by the number of heavy atoms, n_{HA} . Color-coded regression lines are shown for each grouping. (B) A zoomed-in portion of the scatter plot in panel (A), corresponding to the area in the gray rectangle. (C) Scatter plot of the individual size-aware thresholds for each grouping plotted against n_{HA} , color-coded with the same colors as panels (A) and (B).

were omitted (namely, the groups of $n_{\text{HA}} = 6, 10, 14$, and 16 , which account for 5 molecules altogether). Using the respective fitting equations (see Fig. S6 in Section S5 of the ESI,[†] for the individual fitting equations), we calculated the $N_{\text{FOD}}^{\text{xTB}}$ value corresponding to $y = 0.5$ for each group (shown pictorially in Fig. 7A and B as the intercepts of the various regression lines with the horizontal $y = 0.5$ line). We plotted the calculated thresholds against n_{HA} (Fig. 7C) and found an excellent linear fitting, confirming the suspected size effect. This indicated that the size effect should not be neglected, in particular when characterizing increasingly larger molecules.

With these results in hand, we repeated our previous evaluation, identifying the TP, TN, FN, FP percentages for each n_{HA} group using its respective size-aware threshold. Fig. 8 displays the confusion matrices obtained with the initial single threshold approach of $N_{\text{FOD}} = 1.2$ (panel A) and with the size-aware approach (panel B). Using the size-aware method resulted in an 18-fold decrease in the number of FP identifications (from 6.97% to 0.38%). We also noted an increase in the percentage of FN cases (from 0.00% to 0.51%), which suggested that the new thresholds were slightly more permissive (allowing diradical molecules into the so-called closed-shell group). Considering that the previous threshold of 1.2 was overly conservative (especially for the larger molecules), this was not surprising. Nevertheless, we emphasize that, cumulatively, FN and FP accounted for fewer than 1% of the molecules when using the size-aware approach. Overall, the results indicated that incorporating molecular size allows for identification of appropriate thresholds and renders the approach transferable across the chemical space of PBHs.

Two important comments should be made at this point. First, we note that in the original publication of the GFN2-xTB method, Grimme and co-workers suggested normalization of the N_{FOD} value using the number of electrons.⁴⁸ For a comparison between the two approaches, we refer the reader to Section S6 of the ESI.[†] Second, as we mentioned above, the Yamaguchi y value used in this work only measures diradical character, whereas the N_{FOD} value measures total open-shell character.

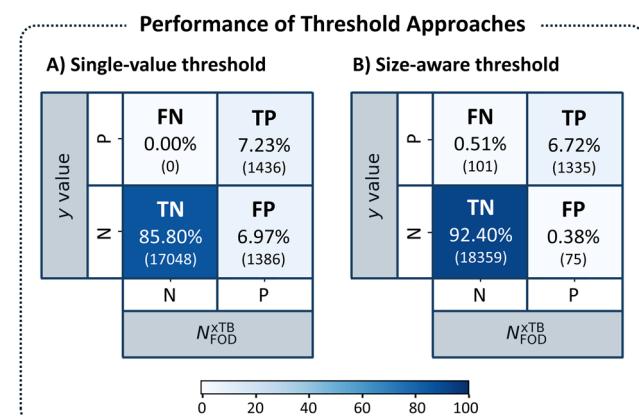


Fig. 8 Confusion matrices assessing the performance of the threshold approaches: (A) the single-value threshold approach and (B) the size-aware threshold approach. Each quadrant identifies the type of identification (TP, FN, FP, and TN, respectively) and details the percentage and number (in parentheses) of molecules identified. The color corresponds to the percentage. The five molecules in groups with $n_{\text{HA}} < 18$ were omitted.

Therefore, the two quantities are not directly comparable and the latter may be biased by partial tetraradical character. The good linear agreement shown in Fig. 6A suggests that this is not the case for the molecules studied herein. Nevertheless, to ensure this, we investigated the y_1 values, as well. These are presented along with additional discussion in Section S7 of the ESI.[†]

xTB vs. DFT

As we mentioned in the introduction, Lischka and coworkers²⁵ originally demonstrated that N_{FOD} values enable accurate estimation of the diradical character of PAHs at a fraction of the computational cost of MR-AQCC. In this work, we demonstrated even more cost-effective characterization by using xTB-generated N_{FOD} , instead of DFT, for a large dataset. As a final aspect of this investigation, we were interested in comparing the xTB and DFT N_{FOD} values.

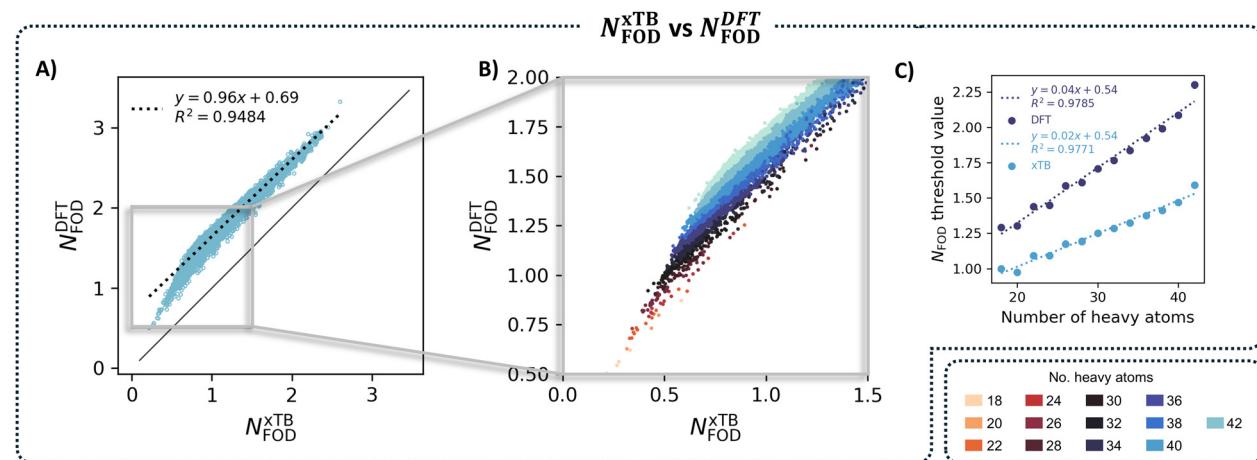


Fig. 9 (A) Scatter plot of $N_{\text{FOD}}^{\text{xTB}}$ against $N_{\text{FOD}}^{\text{DFT}}$ with trendline. (B) A zoomed-in portion of the scatter plot in panel (A), corresponding to the area in the gray rectangle. The data are colored by n_{HA} . (C) Scatter plots of the individual size-aware thresholds for $N_{\text{FOD}}^{\text{xTB}}$ (light blue) and $N_{\text{FOD}}^{\text{DFT}}$ (dark blue) plotted against n_{HA} .

We generated a scatter plot of the two N_{FOD} metrics (Fig. 9A) and observed a rather good correlation between the two methods ($R^2 = 0.93$). Coloring the individual data points according to n_{HA} further revealed a series of parallel correlations, which shift with the molecular size; the larger the molecules, the farther away from the $x = y$ line (Fig. 9B shows a zoomed-in portion of the total plot, color-coded by n_{HA}). Similarly to $N_{\text{FOD}}^{\text{xTB}}$, we extracted the individual threshold values for $N_{\text{FOD}}^{\text{DFT}}$, to compare the size-dependency of the two metrics. Plotting the individual size-aware thresholds obtained for the two N_{FOD} metrics against n_{HA} (Fig. 9C) showed that $N_{\text{FOD}}^{\text{DFT}}$ has a much steeper dependence on size than $N_{\text{FOD}}^{\text{xTB}}$.

Conclusions

In this study, we demonstrated the use of xTB-calculated N_{FOD} values as a computationally efficient approach for identifying open-shell character in polybenzenoid hydrocarbons (PBHs). We computed the Yamaguchi y and $N_{\text{FOD}}^{\text{xTB}}$ values for the entire chemical space of PBHs containing $n_{\text{rings}} \leq 10$ and identified the value of $N_{\text{FOD}}^{\text{xTB}} = 1.2$ as a suitable threshold for distinguishing between closed- and open-shell systems. In general, $N_{\text{FOD}}^{\text{xTB}}$ values show good correlation with the more computationally expensive y values, validating their use as a predictor of open-shell character. Based on this analysis, we found that 6.72% of this chemical space (1335 molecules) are identified as open-shell. However, our findings also highlighted a size-dependency that makes transferability of the approach non-trivial. We elucidated the relationship between $N_{\text{FOD}}^{\text{xTB}}$ and molecular size, finding a linear correlation that allows for easy adjustment of thresholds for molecules of varying sizes. Applying the individual size-aware thresholds for different sized molecules reduced false positive identifications from 6.97% to 0.38%, compared to using a single threshold value. Finally, comparison between xTB- and DFT-calculated N_{FOD} values showed that the former has a much less pronounced size dependency than the latter. As high-throughput calculations with xTB and DFT

methods are becoming more widely used for data-driven research, these findings have important implications for the field of organic electronics and materials science. In this context, we wish to add a note of caution regarding the choice of computational method. To retain consistency between all methods, the $N_{\text{FOD}}^{\text{xTB}}$ values included in this report were obtained for DFT-optimized geometries. However, we also performed a similar analysis using xTB-optimized geometries. We found that the general trends and conclusions remain the same, but the single-value threshold increases to 1.3 and the size-aware thresholds increase accordingly. Furthermore, when using xTB-optimized geometries, a subset of molecules (mostly zethrene derivatives) appeared as outliers on the y vs. $N_{\text{FOD}}^{\text{xTB}}$ scatter plot, whereas well-behaved correlations were obtained when using DFT-optimized geometries. This indicates that the $N_{\text{FOD}}^{\text{xTB}}$ value is sensitive to small geometric changes, and that xTB does not provide accurate geometries for some of the molecules in our dataset (see Section S8 of the ESI† for further details).

In conclusion, we believe that the contribution of the current study is twofold. From a practical perspective, it demonstrates rapid and accurate assessment of open-shell character in large datasets of PBHs, striking a balance between computational efficiency and accuracy. As such, it facilitates the design and discovery of new functional organic materials with tailored electronic properties and can be a valuable tool for researchers in the field of organic chemistry and materials chemistry. From a conceptual perspective, we have shed light on the behavior of the N_{FOD} metric. We have shown that there is a linear relationship between size and N_{FOD} in PBHs.

Data and availability

The data underlying this study are openly available on Gitlab at https://gitlab.com/porannegroup/nfod_xTB and <https://gitlab.com/porannegroup/compas>. The datasets are provided as .csv



files. Further description of the data structure is provided on the GitLab repositories.

Conflicts of interest

There are no conflicts to declare.

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