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High crystal quality and purity Cu₂O by consecutive *in situ* annealing and thermal oxidation of Cu under H₂ and O₂ at elevated temperatures†

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Cu₂O has been obtained *via* the thermal oxidation of Cu foils with a maximum thickness of 120 μm under Ar and O₂ at 1020 °C for 30 min. Care was taken to suppress the oxidation of Cu during the temperature ramp by using H₂ prior to admitting O₂ at 1020 °C. We suggest that the ruby red Cu₂O crystals are obtained in this way by strain-driven grain growth and consist of ~500 μm grains that have a cubic crystal structure. More importantly we find that the crystal quality and purity of the Cu₂O is improved significantly by pre-annealing the Cu under Ar and H₂ at 1000 °C and 10 mbar for up to 180 min. The Cu₂O exhibited photoluminescence at 6.5 K with a maximum at 2.01 eV corresponding to the direct energy gap in very good agreement with density functional theory calculations of the electronic band structure. No emission due to recombination of bound excitons to copper vacancies V_{Cu} and single V_O or double charged V_O oxygen vacancies was observed consistent with high crystal quality Cu₂O. We find that the Cu₂O crystals obtained in this way are p-type and the highest mobility measured at room temperature was $\mu_p = 82 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ with a carrier density of $7.3 \times 10^{17} \text{ cm}^{-3}$.

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1. Introduction

Cuprous oxide (Cu₂O) is a p-type metal-oxide semiconductor that has a fundamental, direct energy bandgap of 2.1 eV and a cubic crystal structure belonging to the *Pn3m* crystallographic space group with a lattice constant of $a = 4.2696 \text{ \AA}$. The native p-type conductivity of Cu₂O is related to the occurrence of copper vacancies (V_{Cu}) that give rise to acceptor-like states located at $E_A \sim 0.2 \text{ eV}$ above the valence band maximum.

Cu₂O was used in the early 1900's to make rectifiers but has also been suggested to be suitable as a solar cell absorber

for a long time, as it has a high absorption coefficient of $\alpha \sim 10^5$ in the visible.¹ However, so far, device efficiencies have been limited to less than 10% (ref. 2–4) even though the theoretical efficiency of Cu₂O solar cells may reach 20%.⁵ More recently, Cu₂O has been shown to have very good photocatalytic properties^{6–8} so it is an active topic of ongoing investigation for water splitting but also for CO₂ reduction.⁹

On the other hand, Cu₂O is interesting from a fundamental point of view since excitons were observed for the first time in Cu₂O by Gross *et al.*¹⁰ in 1956. More importantly, giant Rydberg excitons with principal quantum numbers up to $n = 25$ have only been observed in Cu₂O (ref. 11) which is interesting due to the possibility of adding long-range interactions to the physics of exciton–polaritons that were recently detected in a SiO₂/Ta₂O₅/Cu₂O/Ta₂O₅/SiO₂ Fabry–Pérot cavity.¹² However, giant Rydberg excitons have only been observed in naturally occurring crystals of Cu₂O. It is desirable then to grow high quality, single crystal Cu₂O comparable to that found in nature which is interesting from a fundamental but also technological point of view. In the past, Cu₂O has been obtained by many different methods such as molecular beam epitaxy (MBE),¹³ atomic layer deposition (ALD),¹⁴ pulsed laser deposition (PLD),¹⁵ electrodeposition (ELD)¹⁶ aerosol assisted chemical vapor

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deposition (AACVD),¹⁷ successive ionic layer adsorption-reaction (SILAR),¹⁸ reactive sputtering¹⁹ *etc.* However, the crystal quality and purity of the Cu₂O obtained using most of the methods listed above is not comparable to naturally occurring single crystals. Interestingly, single crystal Cu₂O was obtained by Toth *et al.*²⁰ back in the 1960's *via* the high temperature oxidation of Cu between 1020 °C and 1040 °C in air followed by annealing at even higher temperatures. The high temperature oxidation of Cu was also investigated in the 1970's by Mrowec *et al.*²¹ and around the same time by Brower *et al.*²² who obtained single crystals of Cu₂O by the floating zone method using Cu₂O feed rods prepared by thermal oxidation of Cu rods with diameters of 6.3 mm, at 1050 °C for 96–100 h. Large single crystals of Cu₂O were also grown in the 1970's by Schmidt-Whitley *et al.*²³ It was much later when Mani *et al.*²⁴ showed that it is possible to obtain Cu₂O *via* the thermal oxidation of Cu on *par* with naturally occurring crystals by observing excitonic absorption up to $n = 5$ of the Rydberg series at 2 K. More recently Lynch *et al.*²⁵ prepared single crystal Cu₂O and observed excitons with principal quantum numbers up to $n = 10$. In particular, Lynch *et al.*²⁵ prepared Cu₂O seed rods *via* the thermal oxidation of 5 mm Cu metal rods at 1100 °C for 40 hours in air. The oxidized rods were then used to grow Cu₂O by the optical float-zone method. The growth of Cu₂O *via* the reaction of Cu with O₂ at elevated temperatures is still an active topic of interest and much more recently Xiao *et al.*²⁶ obtained quasi-single crystal Cu₂O *via* stress-assisted thermal oxidation of 0.2 mm Cu foils at 1040 °C in air for 3 h.

It is important to emphasize that in the past the growth of single crystal Cu₂O *via* the thermal oxidation of Cu at elevated temperatures has always been carried out in air that contains N₂, O₂, H₂O *etc.* Here we have carried out an investigation into the thermal oxidation of Cu under controlled conditions at elevated temperatures between 1020 °C and 1040 °C using Ar and O₂. We show that high crystal quality and purity Cu₂O can be obtained by dry oxidation at high temperatures. In addition, we show that the crystal quality and phase purity of the Cu₂O depends strongly on that of Cu which was improved by pre-annealing the Cu under H₂ to eliminate all oxides and promote grain growth. We describe the structural and optical properties of the ruby, red crystals of Cu₂O obtained in this way in conjunction with theoretical calculations of the electronic band structure and in terms of the Cu₂O band profile and CuO/Cu₂O band line-up.

2. Methods

A high purity Cu (99.999%) sheet with a thickness of ~20 or 120 μm was cut into rectangular 10 mm × 20 mm pieces after which the surface was cleaned with isopropanol at room temperature, dried with N₂ and straightened out between two glass slides. The Cu was carefully bent at its four corners to form ~2 mm high stands and subsequently placed inside a corundum boat. The Cu foil was pre-annealed under a flow of Ar and H₂ in a 1" hot wall, chemical vapor deposition

reactor capable of reaching 1200 °C, that was fed by a manifold consisting of four mass flow controllers connected to Ar, NH₃, O₂ and H₂. The reactor was initially purged with Ar and H₂ at room temperature for 10 min to remove traces of air after which the temperature was ramped up to 1000 °C at 30 °C min⁻¹. Upon reaching 1000 °C the same flow of Ar and H₂ was maintained for 60 to 180 min to reduce all oxides and promote grain growth. Finally cool down was carried out by maintaining the flow of Ar and H₂ all the way down to room temperature. A constant flow of 100 ml min⁻¹ Ar and 100 ml min⁻¹ H₂ at 10 mbar was maintained throughout the entire process.

The as received or pre-annealed Cu was oxidized under Ar: O₂ in a different 1" hot wall chemical vapor deposition system capable of reaching 1500 °C, that was fed by a manifold also consisting of four mass flow controllers connected to Ar, NH₃, O₂ and H₂. This was initially purged for 10 min using a flow of 100 ml min⁻¹ Ar and 50 ml min⁻¹ H₂ that was maintained for 10 min at 1 atm to remove air at room temperature after which the temperature was ramped up to 900 °C at 30 °C min⁻¹ by maintaining the same flow of Ar and H₂. Upon reaching 900 °C the ramp rate was reduced to 10 °C min⁻¹ to prevent a temperature overshoot and meltdown of Cu, without changing the flow of Ar and H₂. Cu has a melting point of 1085 °C. At 1020 °C the same flow of Ar and H₂ was maintained for a further 30 min after which the flow of H₂ was interrupted and a flow of 100 ml min⁻¹ Ar was maintained for 15 min to purge the H₂ and admit O₂. Subsequently the Cu foil was oxidized for different times *i.e.*, 30, 45 and 60 min under a flow of 100 ml min⁻¹ Ar and 25 ml min⁻¹ O₂. After this the temperature was ramped up to 1040 °C under a flow of 100 ml min⁻¹ Ar at 1 °C min⁻¹. The Cu₂O was annealed at 1040 °C under 100 ml min⁻¹ Ar for 60 min. Cu₂O has a melting point of 1235 °C. Finally cool down occurred in controlled fashion from 1040 °C down to 1000 °C at -1 °C min⁻¹ under a flow of 100 ml min⁻¹ Ar. When the temperature fell below 1000 °C the Cu₂O was transferred to the cold end on the up-stream side in less than 10 s under the same flow of Ar to prevent air from reaching the Cu₂O from the down-stream end that was opened to insert the transfer rod. The Cu₂O was removed and exposed to air at room temperature.

The Cu₂O crystals obtained in this way were initially inspected by optical transmission microscopy. The surface morphology of the Cu₂O was investigated by scanning electron microscopy (SEM) and the crystal structure and its composition determined X-ray diffraction (XRD) using a Rigaku Miniflex in the Bragg-Brentano θ -2 θ geometry using a Cu K α source with a wavelength of 1.54 Å and 0.3 mm slit size at scan rate of 1° min⁻¹. The structural properties of the Cu₂O crystals were also investigated by transmission electron microscopy (TEM) and high-resolution TEM (HRTEM).

Finally, the optical properties of the Cu₂O were investigated by photoluminescence (PL) spectroscopy. In this case a polished sapphire (Al₂O₃) wafer was utilized as a holder for the brittle Cu₂O crystal that was covered by solid



transparent wax. This was heated in a VT 6060P Thermo Scientific vacuum drying oven at 100 °C and 100 mbar to remove bubbles that form during the melting of the wax. The wax encapsulated the Cu₂O crystal, thereby creating a clear protective layer on top of the Cu₂O which upon cooling fixed the crystal in position. The excess wax was removed by using the Allied Opti Prep Precision Polishing System with a diamond lapping film of 0.5 µm until the Cu₂O surface was exposed. The Cu₂O was then placed inside a closed-cycle cryostat at 5 K. We obtained the PL emission of Cu₂O by using 1.5 mW of continuous wave diode laser excitation at 3.1 eV with a 5 µm spot. The PL emission was collected by a 0.42 NA objective lens and guided onto the slit of a Princeton instruments spectrometer 500, while a long pass filter was utilized in the collection path to filter the excitation laser beam.

The carrier density and mobility of three different Cu₂O crystals with a thickness of ~30 µm were measured at room temperature in a four terminal configuration by the van der Pauw method and Hall effect using a GMW 3470 electromagnet with a magnetic field of 0.2 T. Square 10 mm × 10 mm Cu₂O crystals were prepared as described above and the surface layer of CuO was removed in dilute HCl (aq) and HNO₃ (aq) after which it was rinsed in H₂O and dried rapidly with N₂. Ag contacts were formed at the four corners of the 10 mm × 10 mm Cu₂O crystals.

3. Results and discussion

In the past, the growth of Cu₂O *via* the thermal oxidation of Cu at elevated temperatures has always been carried out in air that contains N₂, O₂, H₂O *etc.*^{20–26} For instance, Cu₂O was obtained *via* the high temperature oxidation of metallic Cu in air by Toth *et al.*²⁰ in horizontal and vertical furnaces between 1020 °C and 1050 °C for 60 min. According to Toth *et al.*²⁰ the Cu was positioned on Pt and after oxidation the Cu₂O was annealed from 5 up to 150 h. An average annealing temperature for 0.76 mm thick Cu was ~1085 °C while that used for 0.38 mm Cu was 1115 °C. The Cu₂O was retracted rapidly in less than 10 s at the end of annealing and the CuO that formed on top of the Cu₂O during cool down was removed by dilute HNO₃. We tried to obtain Cu₂O *via* the thermal oxidation of Cu between 1020 °C and 1040 °C in air as described in detail in the ESI† (S1) but not successfully. Instead, we obtained CuO not Cu₂O. This was attributed to the thermal oxidation and reaction of Cu with an excess of O₂ at low and intermediate temperatures during the temperature ramp which in turn leads to a transition across the phase boundary between Cu₂O and CuO in accordance with Schmidt-Whitley *et al.*²³ It is necessary then to continually adjust the pressure of O₂, *p*(O₂), during the temperature ramp. However, this is difficult as the *p*(O₂) must be adjusted within a very narrow range *e.g.* between 10^{−4} and 10^{−5} mbar in the temperature range 700 °C to 800 °C. In contrast it is easier to prevent a transition across the phase boundary between Cu₂O and CuO and obtain high

crystal quality and purity Cu₂O at high temperatures greater than 1000 °C *e.g.* by adjusting the *p*(O₂) to fall between 10^{−1} and 10^{−1} mbar at 1040 °C. However, it is still necessary to suppress the thermal oxidation of copper during the temperature ramp and eliminate all together any oxides on the surface and in the bulk of Cu to begin with. One way of achieving this is by using H₂ as a reducing agent which is effective in eliminating all oxides on the surface and inside the bulk of Cu but also promotes grain growth and improves the crystallinity of Cu by removing grain boundaries and lowering the energy of the solid.²⁷ It should be noted here that reducing the O₂ pressure is adequate for bulk reduction, but H₂ is required for surface reduction. It has been shown that the CuO surface is reduced to Cu₂O by H₂ at approximately 60 °C and complete reduction from Cu₂O to metallic Cu occurs at 480 °C.²⁸ A large flow of H₂ is very effective in reducing CuO into metallic Cu.²⁹ In addition, it has been shown that H₂ penetrates all metals and stabilizes vacancies, thereby increasing their equilibrium concentration which in turn promotes the rearrangement of metal atoms during annealing and its conversion into a single crystal. In fact, polycrystalline Cu, Ni, and Co foils cannot be converted into single crystals during annealing at high temperatures under Ar without H₂ as shown by Jin *et al.*²⁹ Consequently we pre-annealed the Cu under Ar and H₂ to eliminate all oxides in the Cu and promote grain growth. For completeness, it should be noted that the Cu will react with H₂ leading to the formation of CuH which is the first metal hydride discovered but this is not stable and dissociates into Cu and H₂ according to 2CuH → 2Cu + H₂ at high temperatures.

A typical SEM image of the Cu as received is shown in Fig. 1(a) and (b). A higher magnification SEM image is shown in Fig. 1(c). One may observe the existence of ridges due to the metal rolling. The Cu exhibited weak but nevertheless clear and well resolved peaks in the XRD corresponding to the face centered cubic crystal structure of Cu. We did not observe any peaks corresponding to CuO or Cu₂O in the XRD but traces of CuO and Cu₂O do exist on the surface and inside the bulk at grain boundaries due to the reaction of the Cu with O₂ in air. The 10 mm × 20 mm rectangular Cu acquired a bright metallic appearance after annealing under Ar and H₂ at 1020 °C for 30 min. A typical SEM image is shown in Fig. 1(d) and (e). The Cu consisted of large interlocking grains up to ~5 mm² after annealing for 180 min or 3 hours. Subsequently we carried out thermal oxidation under Ar and O₂ at 1020 °C in a different atmospheric pressure reactor by employing Ar: H₂ during the temperature ramp as described in detail above.

The reaction of Cu with O₂ at 1020 °C for 30 min led to the complete oxidation of Cu into Cu₂O. A typical image of the Cu₂O obtained in this way as seen under an optical microscope in transmission or under direct sunlight is shown in Fig. 1(g). The Cu₂O crystal is semi-transparent and has a ruby, red color like naturally occurring Cu₂O shown in Fig. 1(f). However, the surface is mirror-like with a metallic silver-grey look when viewed at an angle due to the formation of CuO on Cu₂O during



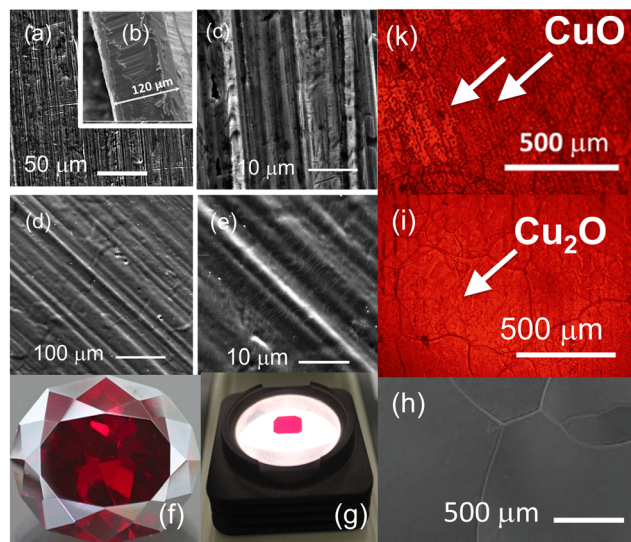


Fig. 1 (a) SEM image of the Cu as received and (b) its section (c) higher magnification SEM of as-received Cu (d) and (e) SEM images of Cu after annealing under Ar and H₂ at 1020 °C for 30 min at 1 atm, without oxidation (f) image of naturally occurring Cu₂O (g) Cu₂O crystal obtained from Cu under Ar and O₂ at 1020 °C for 30 min as viewed under transmission (h) SEM image of the corresponding Cu₂O showing grain boundaries (i) image in transmission showing the Cu₂O crystal obtained from Cu under Ar and O₂ at 1020 °C after pre-annealing under H₂ at 1000 °C and 10 mbar for 60 min (j) image in transmission showing Cu₂O crystal obtained from Cu in the same way but with no pre-annealing under H₂ at 1000 °C and 10 mbar.

cool down. This was not readily observed in the Cu₂O obtained from ~20 μm thick Cu. The 10 mm × 20 mm ruby red crystals of Cu₂O are brittle but easy to handle and have a very flat surface but are bent in one direction suggesting that they are strained. A typical SEM image of the Cu₂O is shown in Fig. 1(h) from which one may clearly observe grains with sizes of 500 × 500 μm² while an image in transmission is shown in Fig. 1(i). We find that the thermal oxidation of Cu into Cu₂O can be completed in 30 min. The Cu₂O exhibited very strong, narrow peaks in the XRD as shown in Fig. 2 corresponding to the (110)

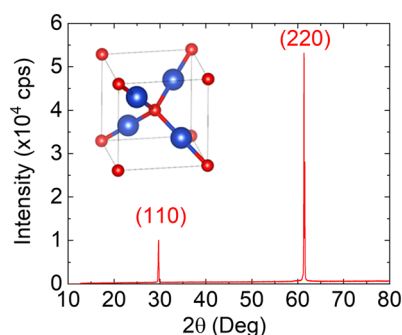


Fig. 2 XRD of the Cu₂O crystal obtained from Cu under Ar: O₂ at 1020 °C showing two narrow, high intensity peaks corresponding to the (110) and (220) crystallographic planes of the cubic crystal of Cu₂O; inset shows the cubic unit cell of Cu₂O, Cu in blue and O in red.

and (220) crystallographic planes of the cubic crystal structure of Cu₂O. The lattice spacing extracted from the XRD is 3.01 Å along the (110) which is very close to that determined by TEM described below and also in very good agreement with 3.02 Å in JCPDS 00-005-0667. It is important to emphasize that the Cu₂O obtained after the thermal oxidation of 120 μm Cu as received, without pre-annealing under Ar and H₂ at 1000 °C, was not so clear as shown in Fig. 1(k) and has a significant content of black inclusions of CuO which is higher than that shown in Fig. 1(i). This was observed systematically and shows that the crystal quality and phase purity of the Cu₂O is strongly dependent on the crystal quality and purity of the Cu.

The structural properties of the Cu₂O obtained from 20 μm thick Cu that was pre-annealed under Ar and H₂ at 1000 °C and 10 mbar were also investigated by SEM, TEM and HRTEM. A typical SEM image of the Cu₂O is shown in Fig. 3(a) and (b). A bright field TEM image is shown in Fig. 3(d) and (e) from which one may observe that the Cu₂O surface is covered by a thin layer of CuO consisting of nanocrystals with an average size of ~10 nm. The surface roughness of the Cu₂O is estimated to be about 3 nm. A HRTEM image of the Cu₂O/CuO interface is shown in Fig. 4(a). The {10-1}, and {110} lattice fringes of the nanocrystals correspond to the monoclinic structure CuO whose unit cell is shown in Fig. 3(c) while the FFT pattern of the Cu₂O square region shown in Fig. 4(a) confirms the cubic crystal structure of Cu₂O; the frequencies correspond to the (11-1), (-110) and (02-1) atomic planes of Cu₂O phase. More importantly, the high crystal quality of the bulk Cu₂O was confirmed by HRTEM as shown in Fig. 4(b).

While we did not observe any bulk inclusions of CuO in the Cu₂O by TEM and HRTEM we expect traces of CuO to exist at grain boundaries or as inclusions elsewhere in the

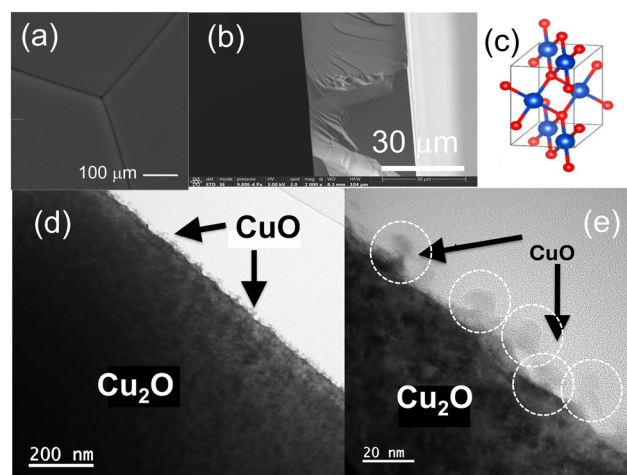


Fig. 3 (a) and (b) SEM image of the surface and section of the 30 μm thick Cu₂O obtained from Cu under Ar and O₂ at 1020 °C (c) unit cell of CuO (d) low magnification, bright field TEM image of the Cu₂O showing CuO nanocrystals on the surface of Cu₂O (e) higher magnification TEM image of the Cu₂O indicating a surface roughness of ~3 nm.



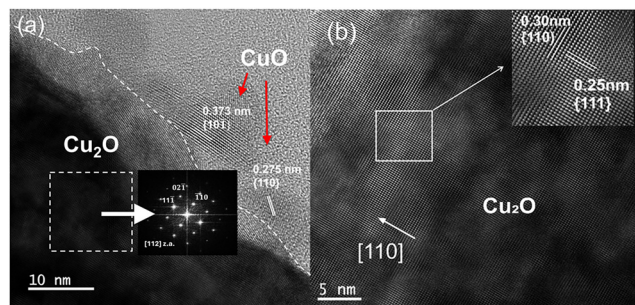


Fig. 4 (a) High resolution TEM image of the $\text{Cu}_2\text{O}/\text{CuO}$ interface showing the $\{110\}$ and $\{10\bar{1}\}$ crystallographic planes of the CuO nanocrystals on the surface of Cu_2O ; inset shows the FFT of the Cu_2O square region (b) HRTEM of Cu_2O showing the $[110]$ direction; inset shows the $\{110\}$ and $\{111\}$ crystallographic planes of the bulk Cu_2O .

Cu_2O and further investigations are required to quantify the phase purity of the Cu_2O that can be improved greatly by pre-annealing the Cu under H_2 as discussed above.

In the past the growth of Cu_2O *via* the thermal oxidation of Cu at elevated temperatures has always been carried out in air that contains O_2 as well as H_2O but the level of humidity or its effect on the growth of Cu_2O was not considered.^{20–26} In fact, no one has investigated the dry or wet oxidation of Cu above 1000 °C. It is known that the addition of traces of H_2O may promote grain growth in accordance with the findings of Liu *et al.*¹⁷ who investigated how the humidity influences the properties of Cu_2O obtained at low temperatures. One might argue then that the thermal oxidation of Cu under Ar and O_2 at 1020 °C and the growth of high crystal quality Cu_2O described here is related to H_2O obtained *via* the reaction of traces of H_2 with O_2 . To clarify this, we investigated the thermal oxidation of Cu under Ar and O_2 at 1040 °C and 1 mbar without using H_2 and find that Cu_2O may be obtained by dry oxidation, see ESI† S2 but its crystal quality and purity are not comparable to that obtained by annealing the Cu under H_2 . On the other hand, we find that Cu_2O cannot be obtained by wet oxidation of Cu at 1040 °C as described in ESI† S3. An excess of H_2O alone resulted in the full oxidation of Cu into CuO . It has been shown that the wet oxidation of Cu at low and intermediate temperatures is faster compared to dry oxidation and follows a parabolic law, so it is difficult to control the rate of oxidation especially at very high temperatures. In short, high crystal quality Cu_2O can be obtained in a controllable fashion by dry oxidation and the reaction of Cu with O_2 but not using only H_2O .

We suggest that the high crystal quality Cu_2O is obtained due to the strain-energy-driven grain growth in accordance with Xiao *et al.*²⁶ The oxidation of Cu with O_2 initially proceeds with the formation of small nucleation centers of Cu_2O on the surface of Cu after which growth occurs as the oxidation front moves inwards *via* the diffusion of oxygen. The lattice constant of Cu_2O is $a = 4.2696$ Å and that of Cu is 3.615 Å. This difference in the lattice constants leads to the generation of compressive growth strains. According to Xiao *et al.*²⁶ Cu_2O grains undergo rapid diffusion and move

preferentially in the vertical growth direction to release the strain through the generation of Frank partial dislocations which grow by absorbing V_{Cu} and penetrate through the bulk. The Frank partial dislocations tend to slip along the lattice plane with the densest atoms and the crystal rotates on the sliding surface, triggering grains to orient in a specific direction. Xiao *et al.*²⁶ showed that $\{111\}$ -oriented Cu_2O with small grains are obtained by surface energy-dominated growth while the $\{110\}$ oriented Cu_2O is obtained by strain-energy-driven grain growth. The Cu_2O crystals obtained in this work consist of $\{110\}$ and $\{220\}$ high index grains like those of Xiao *et al.*²⁶ but we did not impose a strain on the Cu. We suggest that the strain-energy-driven grain growth occurs in the Cu_2O obtained here probably due to the incorporation of H_2 which moves easily and segregates at grain boundaries, voids *etc.* resulting in strain.

The optical properties of the Cu_2O were investigated by steady state photoluminescence (PL) spectroscopy. A typical image of the Cu_2O crystal used in this case is shown in Fig. 5(a). The Cu_2O exhibited strong PL at 6 K as shown in Fig. 5(b) from which one may observe a peak at 2.01 eV and a side peak at 2.04 eV. A smaller, weak but discernible peak exists at 1.95 eV. To obtain a better understanding of the optical properties it is useful to consider the conduction and valence band potential profile of Cu_2O in Fig. 6(a) showing the surface depletion and downward band bending which is usually attributed to the occurrence of oxygen vacancy (V_{O}) donor-like states that are positively charged. It has been suggested that the Fermi level (E_{F}) of Cu_2O is actually pinned at these surface states related to V_{O} which reside energetically in the upper half of the band gap ~ 0.4 eV below the conduction band minimum. In thermodynamic equilibrium, the Fermi level will be constant throughout the $p\text{-Cu}_2\text{O}$. The

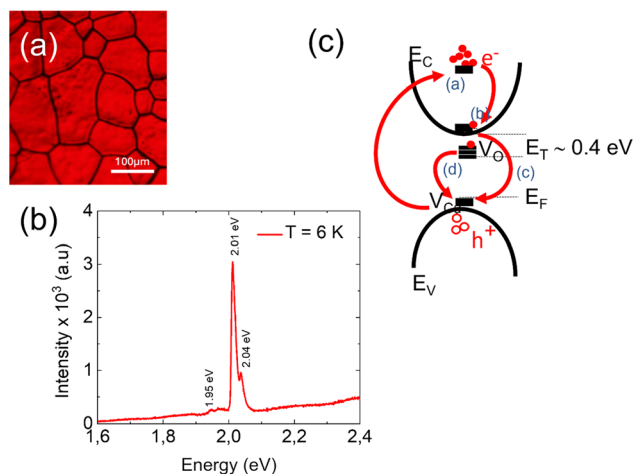


Fig. 5 (a) A typical image of the Cu_2O obtained after polishing for PL (b) PL spectrum of the 30 μm Cu_2O at 6 K showing a large peak at 2.01 eV and side peak at 2.04 eV; a much smaller side peak exists at 1.95 eV (c) schematic energy band diagram showing the generation of photo excited electrons and hole as well as the direct gap radiative recombination and that occurring *via* donor like states close to the conduction band.



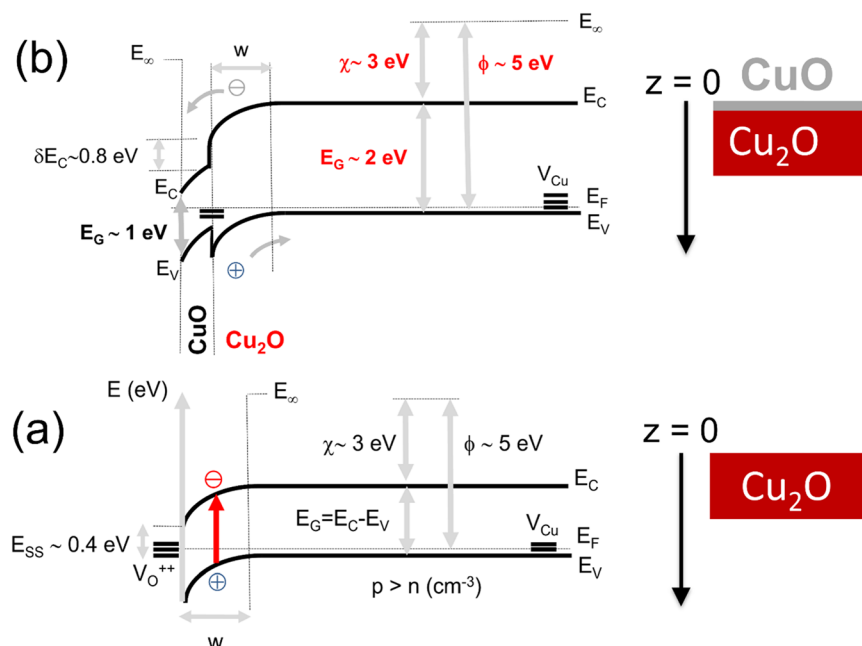


Fig. 6 (a) Conduction and valence band potential profile of Cu₂O showing the surface depletion, downward band-bending, work function ϕ and electron affinity χ as well as the energetic position of donor like surface states (b) corresponding band profile of the Cu₂O/CuO junction, showing the conduction and valence band offsets as well as states related to crystallographic imperfections at the interface.

boundary conditions *i.e.*, energetic position of the Fermi level with respect to the conduction band edge at the surface and in the bulk govern the overall band bending and extent of depletion. Excitation with photons having energies greater than the gap will lead to the generation of an excess density of electrons and holes. The photogenerated electrons in the surface depletion will move towards the surface and holes will move in the opposite direction into the bulk Cu₂O. This will reduce the electric field in the surface depletion until a dynamic equilibrium is reached after which the quasi-Fermi levels will not change with time. However, photoexcited electrons and holes that are generated in the near surface region may recombine *via* surface states related to crystallographic imperfections. In the past PL has been observed at energies lower than the energy band gap of Cu₂O, specifically at 1.47 eV, 1.72 eV, and 1.82 eV which are related to recombination of excitons bound to single charged copper vacancies V_{Cu}¹⁺, single charged oxygen vacancies V_O¹⁺ and double charged oxygen vacancies V_O²⁺, at temperatures as low as 14 K as depicted schematically in Fig. 5(c).^{30–32} We do not observe any PL at these energies consistent with a high crystal quality. It should be noted that photogenerated carriers will recombine in a non-radiative fashion in any CuO that exists at the surface or in the bulk so its effect would be to reduce the overall PL intensity. CuO is an indirect energy gap p-type semiconductor with a band gap of ~1 eV. The CuO/Cu₂O heterojunction has a straddled *i.e.* type I, band line-up and conduction band discontinuity $\Delta E_C(\text{Cu}_2\text{O}/\text{CuO}) \sim 1.0$ eV as shown in Fig. 6(b). In this case electrons and holes will be generated in the near surface region upon photoexcitation and recombine inside the CuO which acts as

a sink for the photogenerated carriers. The lattice constant of Cu₂O is $a = 4.2696$ Å and CuO has a monoclinic crystal structure with lattice constants of $a = 4.6837$ and $b = 3.4226$ Å so this difference in lattice constant will lead to the formation of many crystallographic imperfections at the CuO/Cu₂O. The poor properties of the CuO/Cu₂O interface are well known so eliminating and preventing the formation of CuO in Cu₂O is essential in order to obtain good electrical and optical properties. As we have shown, this was possible by using H₂.

The PL observed at 2.01 eV in Fig. 5 is related to band edge transitions in the Cu₂O in very good agreement with calculations of the electronic band structure of Cu₂O shown in Fig. 7(a) from which we obtained a fundamental, direct energy band gap of 1.96 eV. The first-principles calculations were performed employing Quantum Espresso *ab initio* code.^{33,34} In the self-consistent calculations, an energy cutoff for the plane-wave basis of 350 Ryd was used, while an $8 \times 8 \times 8$ Monkhorst-Pack mesh was assumed for the *k*-space sampling. The structure was geometrically optimized using the PBE exchange–correlation functional³⁵ and the obtained lattice constant, 4.291 Å, is very close to the experimental. Due to the known problem of GGA functionals, like PBE, in reproducing the electronic properties, we employed DFT+U (PBE + U) approximation.³⁶ For the Hubbard parameters, we adopted the values of 11 and 7 eV for Cu-d and O-p orbitals. The combination of these values with the PBE optimized geometry are reported to be accurate for the band gap of Cu₂O.³⁷ In Fig. 7(a), we show the obtained band structure of Cu₂O, while in (b) the corresponding density of states (DOS) is included. It is worth mentioning that optimizing the structure with PBE + U (instead of plain PBE) leads to an



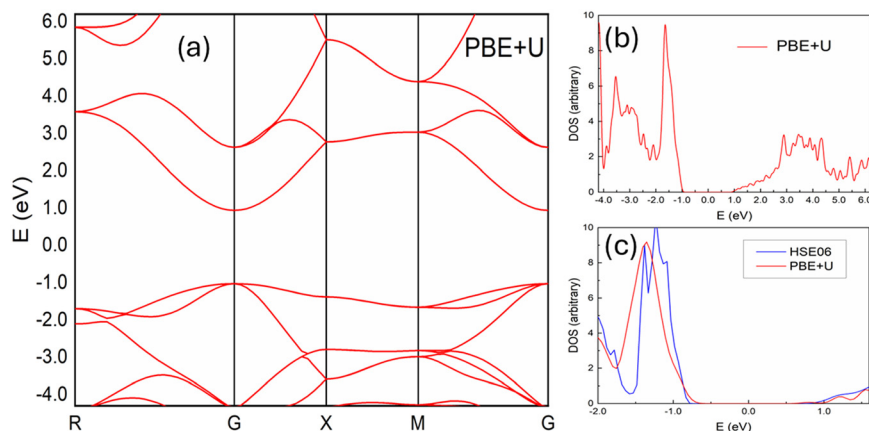


Fig. 7 (a) Electronic band structure of Cu₂O obtained by PBE + U approximation (for PBE optimized structure) showing that the direct gap is 1.96 eV at the Γ symmetry point, (b) the corresponding DOS, and (c) the DOS with PBE + U compared with that of HSE06, for the PBE + U optimized structure.

overestimation of the lattice constant (4.475 Å) and to a significant reduction of the band gap (1.71 eV). In Fig. 7(c), we include the DOS obtained by PBE + U for this lattice parameter, and for comparison we also include the DOS obtained by HSE06 (ref. 38) range-separated hybrid-functional.

Finally, we measured the carrier density and mobility of three different 10 mm × 10 mm Cu₂O crystals with a thickness of ~30 μm after removing the CuO layer on the front and back of the Cu₂O in HCl (aq) and HNO₃ (aq) to allow forming Ag ohmic contacts. We find that all of the Cu₂O crystals are p-type. The largest mobility was found to be $\mu_p = 82 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ at room temperature with a corresponding carrier density of $7.3 \times 10^{11} \text{ cm}^{-3}$. The other two Cu₂O crystals had lower mobilities of $\mu_p = 69 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ and $\mu_p = 61 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ at room temperature with corresponding carrier densities of $8.1 \times 10^{11} \text{ cm}^{-3}$ and $8.8 \times 10^{11} \text{ cm}^{-3}$. These mobilities are smaller than $100 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ measured by Xiao *et al.*²⁶ We suggest that the mobility is limited primarily by grain boundary and defect scattering. The hole mobility can be improved by annealing at even higher temperatures leading to the formation of larger grains and eventually a single crystal of Cu₂O.

We ought to mention here that the bare surface of Cu₂O will be oxidized into CuO upon exposure to ambient conditions over time so we suggest that it is important to remove the CuO on the front and back surface once and for all after which the Cu₂O surface has to be passivated immediately with a thin layer of MgO that has a lattice constant of 4.21 Å which is very close to that of Cu₂O *i.e.* 4.2696 Å. The lattice mismatch between MgO and Cu₂O is ~1%. MgO is a transparent, insulator with a large energy band gap thereby allowing light to reach the underlying Cu₂O. We also have to mention that the Cu₂O crystals obtained by thermal oxidation of Cu foils are brittle and bent. One may obtain circular like, flat wafers by sawing boules of Cu₂O grown by the floating zone method but this takes a significantly longer time and is much more energy consuming. The growth of high quality, polycrystalline and

even single crystal Cu₂O from Cu foils with a thickness of 10–50 μm's is attractive due to the shorter times required for complete oxidation and crystallization by annealing.

It is necessary then to obtain flat Cu₂O crystals from Cu foils with a thickness of a few tens of μm's which may be achieved by controlling the thermal oxidation, heating and cooling rates. These issues must be resolved before moving into the fabrication of devices such as p–n heterojunction solar cells as well as cavities for the observation of exciton-polaritons.

4. Conclusions

We have carried out a detailed investigation into the thermal oxidation of Cu at elevated temperatures between 1020 °C and 1040 °C under Ar and O₂. Care was taken to eliminate oxides and promote grain growth in the Cu by annealing under Ar: H₂ prior to oxidation. We obtain ruby red Cu₂O crystals that consist of ~500 μm² grains with a cubic crystal structure. In addition, we show that high crystal quality and phase purity Cu₂O is obtained by dry not wet oxidation. The crystal quality and phase purity of the Cu₂O was improved significantly by pre-annealing the Cu under Ar: H₂ at 1020 °C and 10 mbar for up to 3 hours. The Cu₂O obtained in this way exhibited clean PL at 6.5 K with a maximum at 2.01 eV corresponding to the direct energy gap in very good agreement with calculations of the electronic band structure. We did not observe any emission related to crystallographic imperfections confirming the high crystal quality and phase purity of the Cu₂O. The Cu₂O crystals obtained in this way were p-type and exhibited mobilities between $\mu_p = 61$ and $82 \text{ cm}^2 \text{ Vs}^{-1}$ which may be improved by annealing at higher temperatures for longer times.

Data availability

The data supporting this article has been included as part of the ESI.†



Author contributions

M. Zervos was responsible for the growth of the Cu₂O crystals, X-ray diffraction, microscopy, hall effect and writing up; I. Paschos and P. Savvidis were responsible for the measuring the optical properties and writing up; N. Florini, K. Koutsokostas and P. Komninou were responsible for the TEM/HRTEM and writing up; N. N. Lathiotakis, P. M. Levendis and S. Marinakis were responsible for the electronic band structure theoretical calculations and writing up.

Conflicts of interest

There are no conflicts of interest to declare.

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