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A novel method for the deoxydehydration (DODH) of diols to alkenes has been developed by using hydrazine as a reductant via HOME (hydrazone as an organometallic equivalent) chemistry catalyzed by Ru-PNP complexes. This process avoids harsh reaction conditions, minimizes waste, and can be applied to a variety of diol compounds.

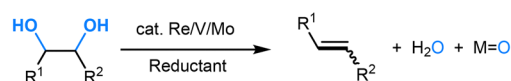
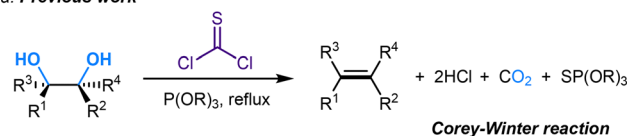
Diol and polyol structures are prevalent in natural compounds and biomolecules.¹ As renewable feedstocks, biomass-derived diols and polyols have become attractive alternatives to fossil fuels in the production of value-added chemicals. Dehydration, hydrogenation, and hydrogenolysis have received great attention for the conversion of polyols.^{2–4} On the other hand, olefin serves as a fundamental building block in organic synthesis.^{5,6} However, methods for directly converting diols into alkenes are very limited. Thus, the selective and efficient conversion of diols into olefins *via* a deoxydehydration (DODH) process holds the potential to facilitate a wide range of transformations in synthetic chemistry. Compared to traditional, well-developed strategies starting from alkene or carbonyl compounds,^{7–10} the olefination derived from abundantly available alcohols *via* DODH offers a cost-effective pathway to achieve C=C linkage from an economic perspective. Currently, DODH methods are quite limited, and existing techniques require toxic and hazardous chemicals (phosphite or phosphine) and harsh reaction conditions.^{11–14} Deoxydehydration (DODH) of diols toward alkenes with the extrusion of one oxygen atom and one equivalent of water molecule was exploited first by Corey *et al.*, in a reaction known as the Corey–Winter reaction.¹⁵ However, the well-established Corey–Winter reaction usually involves prolonged heating at elevated temperatures and the use of phosphorylated reductive reagents (Scheme 1a).¹⁶ In addition to the Corey–Winter reaction, the DODH of vicinal diols and polyols is typically performed using Re,¹⁷ V,¹⁸ or Mo¹⁹ in combination

Ru-catalyzed deoxydehydration (DODH) with hydrazine

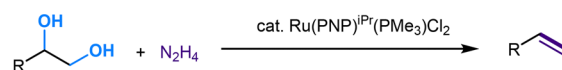
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with phosphine, secondary alcohols including 3-octanol, or H₂ (for certain solid catalysts) as reductants.^{17–19} This process often necessitates rigorous reaction conditions. Furthermore, they exhibit inefficacy in facilitating the DODH of substrates with complex structures (Scheme 1a).¹³ According to HOME-chemistry (hydrazones as organometallic equivalents) and “borrowing hydrogen” strategies,^{20–22} our group has developed a HOME-chemistry-based olefination strategy from two carbonyl compounds, which experienced an E1cb process.²³ Likewise, in 2018, Milstein’s lab revealed a manganese-catalyzed dehydrogenative coupling of alcohols with hydrazine/hydrazone accompanied by a hydrogen transfer process to produce alkenes,²⁴ following the example of the Wittig-type and Julia-type olefination of alcohols.^{25,26} Besides manganese, ruthenium enabled alcohols as carbonyl surrogates.²⁷ Moreover, both ruthenium and titanium demonstrated high activity in the conversion of alcohols to alkenes and alkanes as well.²⁸ Inspired by these works, we proposed that Ru could be a

a. Previous work



b. This work



- Broad substrate scope - Milder reaction conditions - Benign waste

Scheme 1 Methods for the direct conversion of diols into olefins.

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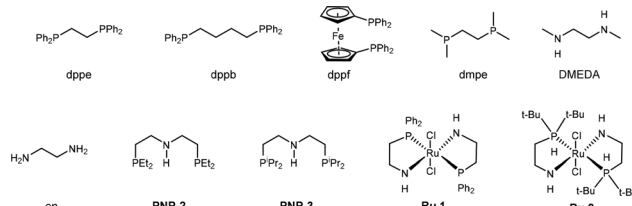
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suitable catalyst for DODH with the assistance of hydrazine due to its versatile applications in coupling processes that involve alcohols. By leveraging the hydrogen transfer process, we developed a Ru-catalyzed DODH reaction enabled by hydrazine. Our approach offers the advantage of being operable under mild conditions with minimal environmental impact. It requires low temperatures and eliminates the need for gas, resulting in only N₂, H₂, and water as byproducts (Scheme 1b). Aromatic-substituted diols and aliphatic diols exhibit good performance within our method. Mechanistically, our approach involves a Ru-catalyzed dehydrogenation, followed by hydrazone formation and an intramolecular E1cb elimination. We started by investigating the reaction conditions of aliphatic diols and hydrazine. A set of ruthenium(II) catalysts was tested first in view of their high activity and good functional group tolerance in hydrazone chemistry involving alcohols as carbonyl surrogates in deoxygenation and Grignard-type reactions.^{27,29,30} We utilized 1,2-dodecanediol as our model substrate, attributed to its simple structure and high boiling point. It should be noted that no desired product was detected in the reaction system catalyzed by Ru(PPh₃)₃Cl₂ with K₃PO₄ as the base at 100 °C (Table 1, entry 1). The combination of a ruthenium catalyst and an electron-rich phosphine ligand typically resulted in better performance in HOME chemistry.³¹ Therefore, after the replacement of Ru(PPh₃)₃Cl₂ with more electron-rich Ru(PMe₃)₄Cl₂ or more stable Ru(dppb)PPh₃Cl₂ as a catalyst, 9% and 7% yields were observed, respectively (Table 1, entries 2 and 3). Then, different bidentate ligands were tested in the presence of [Ru(*p*-cymene)Cl₂]₂ (Table 1, entries 4, 5 and 6). Compared with dppe and DMEDA, the catalyst with more electron-rich and less sterically hindered dmpe showed better activity (Table 1, entry 5). In contrast, ruthenium catalysts with other bidentate ligands proved ineffective in this reaction system (Table 1, entries 7–10). Since the dehydrogenative coupling of alcohols and hydrazine has shown high efficiency with transition-metal pincer catalysts,^{24,32,33} we investigated the efficiency of Ru, Mn and Ir-pincer catalysts (Table 1, entries 11–15).³⁴ The yield was tremendously attenuated with Ru(PNP-2)Cl₂ or Ru(PNP-3)Cl₂ (Table 1, entries 11 and 12). It was reported that the ruthenium-pincer complex with a monodentate phosphine ligand significantly depressed some potential side reactions, including hydrazone dimerization.³⁵ Thus, the highest yield was obtained by treating aliphatic diols with Ru(PNP-3)(PMe₃)Cl₂ (Table 1, entry 13). Organic bases such as DBU failed to furnish the final product (Table 1, entries 16). In contrast, K₃PO₄ (Table 1, entry 13) was the most effective base (Table 1, entries 17 and 18). Temperature has a significant influence on productivity as well. The yield demonstrated an increasing trend against temperature from 60 to 100 °C (Table 1, entries 13, 19 and 20) but decreased above 100 °C (Table 1, entry 21), probably due to the competing Wolff–Kishner reduction. Notably, when the catalyst loading of Ru(PNP-3)(PMe₃)Cl₂ was increased to 10 mol%, an even higher yield was detected at 80 °C (Table 1, entry 23) *versus* 100 °C (Table 1, entry 22). For aromatic diols, an alkyl side product was observed due to the hydrogen transfer ability of the

Table 1 Investigation of the reaction conditions on aliphatic diols

Entry	Catalyst	Ligand	Catalyst loading (mol%)	Base	T (°C)	2a (%)
1	Ru(PPh ₃) ₃ Cl ₂	—	5	K ₃ PO ₄	100	0
2	Ru(PMe ₃) ₄ Cl ₂	—	5	K ₃ PO ₄	100	9
3	Ru(dppb)PPh ₃ Cl ₂	—	5	K ₃ PO ₄	100	7
4	[Ru(<i>p</i> -cymene)Cl ₂] ₂	dppe	5	K ₃ PO ₄	100	0
5	[Ru(<i>p</i> -cymene)Cl ₂] ₂	dmpe	5	K ₃ PO ₄	100	36
6	[Ru(<i>p</i> -cymene)Cl ₂] ₂	DMEDA	5	K ₃ PO ₄	100	0
7	Ru1	—	5	K ₃ PO ₄	100	7
8	Ru2	—	5	K ₃ PO ₄	100	0
9	Ru2	dmpe	5	K ₃ PO ₄	100	20
10	Ru(dppf)(en)Cl ₂	—	5	K ₃ PO ₄	100	7
11	Ru(PNP-2)Cl ₂	—	5	K ₃ PO ₄	100	11
12	Ru(PNP-3)Cl ₂	—	5	K ₃ PO ₄	100	18
13	Ru(PNP-3)(PMe ₃)Cl ₂	—	5	K ₃ PO ₄	100	64
14	Mn(CO) ₂ (PNP-3)Br	—	5	K ₃ PO ₄	100	28
15	Ir(PNP-3)H ₃ Cl	—	5	K ₃ PO ₄	100	6
16	Ru(PNP-3)(PMe ₃)Cl ₂	—	5	DBU	100	0
17	Ru(PNP-3)(PMe ₃)Cl ₂	—	5	K ₂ PO ₃	100	49
18	Ru(PNP-3)(PMe ₃)Cl ₂	—	5	KO ^t Bu	100	20
19	Ru(PNP-3)(PMe ₃)Cl ₂	—	5	K ₃ PO ₄	60	42
20	Ru(PNP-3)(PMe ₃)Cl ₂	—	5	K ₃ PO ₄	80	61
21	Ru(PNP-3)(PMe ₃)Cl ₂	—	5	K ₃ PO ₄	120	51
22	Ru(PNP-3)(PMe ₃)Cl ₂	—	10	K ₃ PO ₄	100	58
23	Ru(PNP-3)(PMe ₃)Cl ₂	—	10	K ₃ PO ₄	80	71

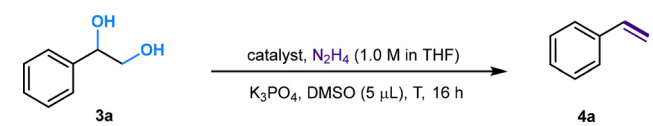


Reaction conditions: **1a** (40.5 mg, 0.2 mmol), Ru(PNP-3)(PMe₃)Cl₂ (11.1 mg, 0.02 mmol, 10 mol%), K₃PO₄ (31.8 mg, 0.15 mmol, 75 mol%), 1 M N₂H₄ solution in THF (0.5 mL, 0.5 mmol, 2.5 equiv.), 24 h at 80 °C, under N₂.

Ru catalyst. Thus, further investigation into the reaction conditions was necessary, and 1-phenyl-1,2-ethanediol was chosen as the model substrate. It should be noted that an 88% yield was detected at 70 °C when we shortened the reaction time to 16 h, decreased the catalyst loading to 5 mol% and the amount of base to 25 mol%, and added 5 μL DMSO as an additive (Table 2, entry 7). When the reaction was charged with the identical catalyst loading and base quantity used for aliphatic diols, the resulting yield was diminished in comparison (Table 2, entries 1 and 2). In addition, a decreased yield was noted when the reaction was carried out at temperatures either above or below 70 °C. (Table 2, entries 5 and 8). With other conditions unchanged, a control experiment was performed without DMSO, leading to a diminished yield (Table 2, entry 9).³⁰ We then applied the optimized reaction conditions to various compounds containing diols, giving rise to moderate to high yields in most cases. In the scope of aliphatic diols, aside from 1,2-dodecanediol affording a satisfactory 70% isolated yield (Table 3, **2a**), a phenyl ring (Table 3, **2b**), morpholine



Table 2 Investigation of the reaction conditions on aromatic diols

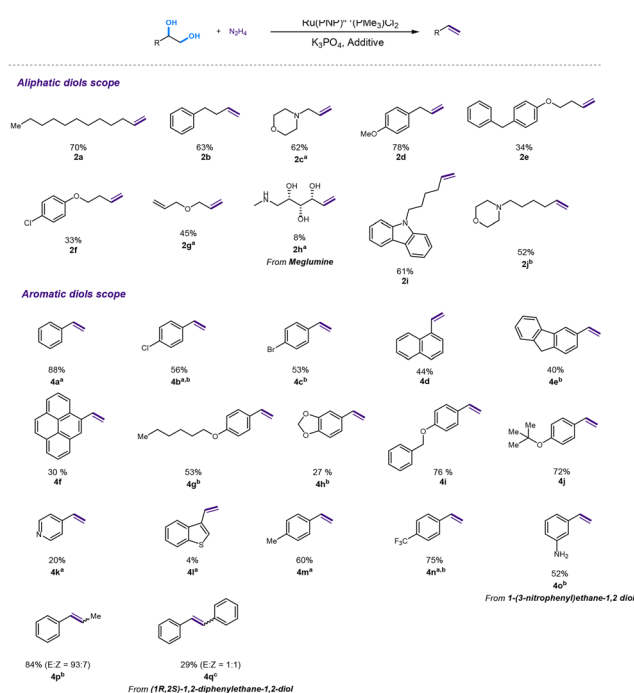


Entry	Catalyst	Catalyst loading (mol%)	K ₃ PO ₄ (mol%)	T (°C)	DMSO (μL)	4a (%)
1	Ru(PNP-3)(PMe ₃)Cl ₂	10	75	80	5	34
2	Ru(PNP-3)(PMe ₃)Cl ₂	5	75	80	5	51
3	Ru(PNP-3)(PMe ₃)Cl ₂	5	100	80	5	43
4	Ru(PNP-3)(PMe ₃)Cl ₂	5	50	80	5	53
5	Ru(PNP-3)(PMe ₃)Cl ₂	5	25	80	5	58
6	Ru(PNP-3)(PMe ₃)Cl ₂	5	10	80	5	40
7	Ru(PNP-3)(PMe ₃)Cl ₂	5	25	70	5	88
8	Ru(PNP-3)(PMe ₃)Cl ₂	5	25	50	5	23
9	Ru(PNP-3)(PMe ₃)Cl ₂	5	25	70	—	68

Reaction conditions: **3a** (27.6 mg, 0.2 mmol), Ru(PNP-3)(PMe₃)Cl₂ (5.54 mg, 0.01 mmol, 5 mol%), K₃PO₄ (10.6 mg, 0.05 mmol, 25 mol%), 1 M N₂H₄ solution in THF (0.5 mL, 0.5 mmol, 2.5 equiv.), DMSO (5 μL, 0.07 mmol, 35 mol%), 16 h at 70 °C, under N₂.

(Table 3, **2c** and **2j**), methoxy group (Table 3, **2d**), alkene (Table 3, **2g**) and carbazole (Table 3, **2i**) were well tolerated. In contrast, lower reactivities were observed on substrates with substituted ethoxy benzenes (Table 3, **2e** and **2f**). Their products, **2e** and **2f**, were mixed with unreacted starting materials. Furthermore, meglumine exhibited a low reactivity, likely attributed to challenges in solubility and the incompatibility of amine functionalities (Table 3, **2h**). It is worth noting that complex mixtures were formed in the presence of an additional terminal hydroxy group or aldehyde functional group, including glyceraldehyde. This could be explained by the possible intermolecular olefination side reaction and other competing reactions. Furthermore, fructose is not suitable for the reaction, indicating that condensation of N₂H₄ did not occur on the internal carbonyl group, which serves as evidence for the mechanism that dehydrogenation likely takes place at the terminal OH group. As for aromatic diols, apart from the 88% yield detected in the conversion of 1-phenyl-1,2-ethanediol (Table 3, **4a**), the reaction is applicable to naphthalene (Table 3, **4d**) and halo-substituted aromatic rings (Table 3, **4b** and **4c**). Substrates bearing sterically hindered polycyclic aromatic hydrocarbons such as fluorene (Table 3, **4e**) and pyrene (Table 3, **4f**) delivered 40% and 30% yields, respectively. Besides, the reactivity was dramatically attenuated in the presence of the methylenedioxy group (Table 3, **4h**) compared to the alkoxy group (Table 3, **4g**). This likely occurs due to the conflicting electronic effect of *meta* and *para*-oxygen on the aromatic ring, slowing down the E1cB elimination process. An alternative explanation could be that oxygen may coordinate with ruthenium, thereby hindering the reaction pathway. In contrast, the substrate with the *para*-electron-donating benzyloxy and *tert*-butoxy groups exhibited high reactivity due to their weaker electronic effect and higher steric hindrance (Table 3, **4i** and **4j**). Other common functional groups were well tolerated (Table 3, **4m** and **4n**), with the exception of heterocycles

Table 3 Substrate scope of hydrazine-assisted DODH



^a ¹H NMR yields were reported using mesitylene as an internal standard. ^b The reaction was conducted at 100 °C. ^c The reaction was conducted at 120 °C. Reaction conditions A for aliphatic diols: diol (0.2 mmol), Ru(PNP-3)(PMe₃)Cl₂ (11.1 mg, 0.02 mmol, 10 mol%), K₃PO₄ (31.8 mg, 0.15 mmol, 75 mol%), 1 M N₂H₄ solution in THF (0.5 mL, 0.5 mmol, 2.5 equiv.), 24 h at 80 °C, under an N₂ atmosphere. Yields of isolated products were reported unless otherwise noted. Reaction conditions B for aromatic diols: diol (0.2 mmol), Ru(PNP-3)(PMe₃)Cl₂ (5.54 mg, 0.01 mmol, 5.0 mol%), K₃PO₄ (10.6 mg, 0.05 mmol, 25 mol%), 1 M N₂H₄ solution in THF (0.5 mL, 0.5 mmol, 2.5 equiv.), DMSO (5 μL, 0.07 mmol, 35 mol%), 16 h at 70 °C, under an N₂ atmosphere. Yields of isolated products were reported unless otherwise noted.

(Table 3, **4k** and **4l**). Furthermore, the compound 1-(3-nitrophenyl)ethane-1,2-diol underwent a conversion to vinylaniline alongside a reduction process (Table 3, **4o**). Aromatic diols that featured an ester group (methyl 4-vinylbenzoate), a cyano group (4-vinylbenzonitrile), or an additional simple alkyl group at the 2-position (prop-1-en-2-ylbenzene) were found to be ineffective (see the SI). The reactions involving certain internal diols derived from their respective alkenes also proceeded effectively (Table 3, **4p** and **4q**). The 1-phenylpropane-1,2-diol-derived product **4p** was obtained in an 84% yield with a 93 : 7 *E* : *Z* ratio; however, **4q**, due to its steric hindrance, gave only a 29% yield with poor stereoselectivity.

In conclusion, we have disclosed for the first time a DODH process of terminal diols toward alkenes with hydrazine by using a ruthenium catalyst. Both aromatic and aliphatic diols were efficiently transformed, resulting in the release of benign water, hydrogen gas, and nitrogen gas as by-products. Compared to traditional DODH methodologies, this protocol exhibits high selectivity of terminal diols under mild reaction conditions. From a sustainability perspective, this DODH method sets up a promising avenue for the selective conversion



of biomass-based feedstocks into high-value olefin products. Therefore, further studies of the mechanism and synthetic utilities of this DODH strategy are ongoing.

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Conflicts of interest

There are no conflicts to declare.

Data availability

The data supporting this article have been included as part of the SI. Supplementary information: General information, experimental procedures, unsuccessful substrates, spectroscopic data, and copies of ^1H and ^{13}C NMR spectra of the products. See DOI: <https://doi.org/10.1039/d5cc04813c>.

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