

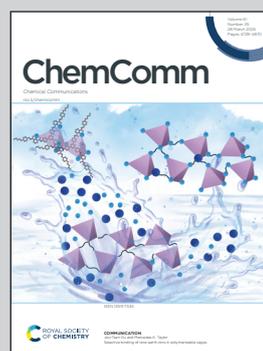
Showcasing research from Professor Griesbeck's laboratory, Department of Chemistry and Biochemistry, University of Cologne, Germany.

Thiolate-CLPG (chemiluminescent protecting groups) based on coumaranone thiolcarbamates

The (quantitative) release of aromatic and aliphatic thiols from coumaranone-thiolurethanes is accompanied by blue or green light chemiluminescence which is long-lasting over the total release period.

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# Thiolate-CLPG (chemiluminescent protecting groups) based on coumaranone thiolcarbamates†

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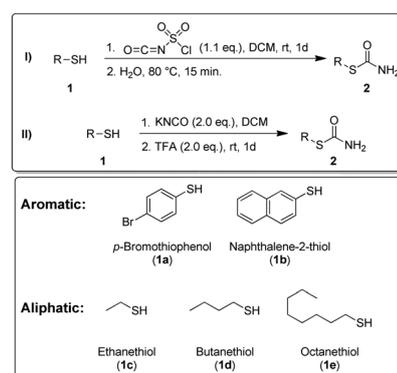
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**The base-induced oxidative release of aromatic and aliphatic thiols from novel coumaranone-thiolurethanes 3a–3e is accompanied by long-lasting blue or green chemiluminescence. Thus, these compounds serve as base/oxidation-labile protecting groups that decompose with strong visible luminescence and subsequent quantitative release of the protected thiolate as leaving group.**

Protecting groups are essential as intermediate blockers in complex organic syntheses but also relevant for the selective release of small molecules in numerous chemical and/or biological applications.<sup>1–5</sup> This release can be initiated by acids, bases, metals, enzymes, redox-active reagents, or visible light. In the latter case, the abbreviation PRPG was introduced to mark these protecting groups as photoreleasable.<sup>6–12</sup> The photon might, however, appear at two parts of the release process: at the beginning as in PRPG or as a consequence of the release process, *i.e.* as chemiluminescence (CL). We have termed the latter PG's as CLPG, chemi-luminescent protecting groups.<sup>13</sup> An advantage of these groups is that the release process can be detected spacially by measuring the emitted light and kinetically evaluated by determining the decay rates. Thus, a spatio-temporal detection of (thermal) release becomes possible with CLPG. We have demonstrated this principle for alkoxides as leaving groups<sup>13</sup> using the coumaranone-carbamates that were investigated intensively by Schramm and coworkers.<sup>14–16</sup> We could show that the leaving group ability is crucial for this release step by comparing urea- with urethane-linked coumaranones: whereas the release of phenolates is fast

and directly connected with CL, aliphatic alkoxides and amide anions are not released. CL can be detected, but decays very fast in the case of the latter.<sup>13</sup> For this study the thiolcarbamate precursors were prepared according to the synthesis approach by Graf *et al.*<sup>17</sup> for aromatic thiols by Loev and Kormendy<sup>18</sup> for aliphatic thiols (Scheme 1). The yields are summarised in Table 1.

The subsequent conversion to the corresponding coumaranones is based on a synthetic protocol by Schramm *et al.*<sup>15</sup> with minor modifications. After condensation with glyoxylic acid monohydrate **4**, TFA was used instead of a mixture of acetic and sulfuric acid and



**Scheme 1** (top) Illustration of the synthesis of thiolcarbamates. Method (I) is according to the adapted synthesis protocol of ref. 17 and method (II) according to ref. 18. (bottom) Substrates used for the synthesis of thiolcarbamates.

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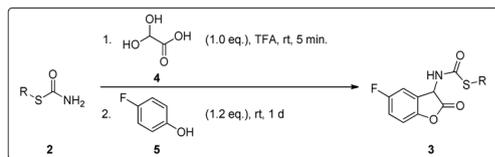
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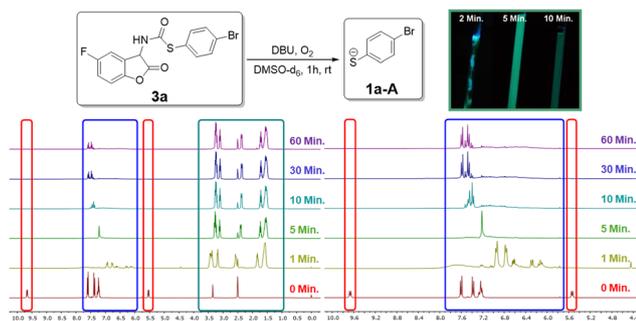
**Table 1** Thiolcarbamates and coumaranones from aromatic and aliphatic thiols

Thiolcarbamate from	Coumaranone from
<i>p</i> -Bromothiophenol ( <b>2a</b> , 83%)	<i>p</i> -Bromothiophenol ( <b>3a</b> , 20%)
Naphthalene-2-thiol ( <b>2b</b> , 67%)	Naphthalene-2-thiol ( <b>3b</b> , 12%)
Ethanethiol ( <b>2c</b> , 35%)	Ethanethiol ( <b>3c</b> , 8%)
Butanethiol ( <b>2d</b> , 55%)	Butanethiol ( <b>3d</b> , 4%)
Octanethiol ( <b>2e</b> , 69%)	Octanethiol ( <b>3e</b> , 17%)





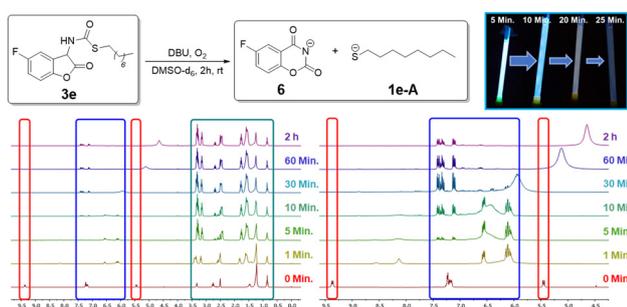
**Scheme 2** Synthesis of coumaranones with a thiolcarbamate substructure according to the adapted synthesis protocol of Schramm *et al.*<sup>15</sup>



**Fig. 1**  $^1\text{H-NMR}$  (in  $\text{DMSO-}d_6$ , 0.1 M) of the base triggered decomposition of **3a**: CL-photographs after the addition of DBU and constant saturation with oxygen. Left: Complete NMR-spectrum, right: enlarged area from 4.8–10.0 ppm. Blue boxes = change in signal intensity and occurrence of new signals; red boxes = disappearance of the NH/ $\alpha$ -CH signals due to deprotonation after 1 min; dark green box = additional signals of DBU.

*p*-fluorophenol **5**, which delivered the highest yields (Scheme 2). The coumaranones **3a–e** could be successfully synthesised in moderate yields, and derivative **3a** with the highest yield (Table 1).

Initially, the decomposition of the coumaranone protected derivatives of *p*-bromothiophenol **3a** (Fig. 1) and octanethiol **3e** (Fig. 2) were investigated *via*  $^1\text{H-NMR}$ . The decomposition of thiolcarbamate–coumaranone **3a** is illustrated in Fig. 1. The signals corresponding to the  $\alpha$ -CH group at 5.54 ppm and the NH group at 9.67 ppm vanish, along with notable changes in the aromatic region between 6.0 and 7.8 ppm after the addition of DBU. Initially, the CL changes from blue to green and gradually fades after more than 10 minutes. During this time, the two signals of *p*-bromothiophenol (**1a-A**) at 7.47 and 7.59 ppm show a significant increase. After

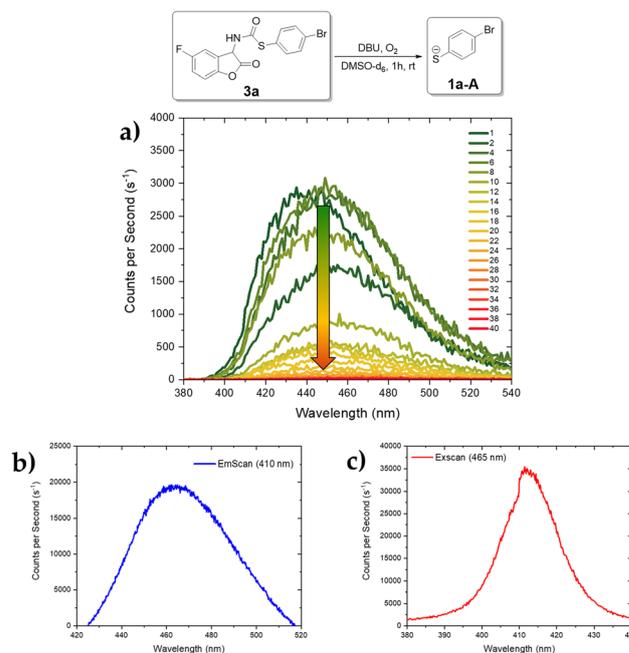


**Fig. 2**  $^1\text{H-NMR}$  (in  $\text{DMSO-}d_6$ , 0.1 M) of the base triggered decomposition of **3e**: CL-photographs after the addition of DBU and constant saturation with oxygen. Left: Complete NMR-spectrum, right: enlarged area from 4.8–10.0 ppm. Blue boxes = change in signal intensity and occurrence of new signals; red boxes = disappearance of the NH/ $\alpha$ -CH signals due to deprotonation after 1 min; dark green box = additional signals of DBU.

30 minutes, no further changes are observed, confirming the complete deprotection of **3a**.  $^1\text{H-NMR}$  analysis of **3e** also confirms the successful deprotection of octanethiol (**1e-A**) (Fig. 2). As with **3a**, the  $\alpha$ -CH and NH signals disappear following DBU addition, and a distinctive aromatic byproduct appears in the range between 7.0 and 7.5 ppm within 5 minutes. The chemiluminescence of **3e** lasts for 30 minutes and exhibits a light blue colour. Once the CL fully decays, no further changes are detected in the  $^1\text{H-NMR}$  spectrum. Subsequent 2D NMR spectra identified fully deprotected octanethiol (**1e-A**) and compound **6** as the main aromatic byproduct of the reaction. Overall, both experiments verified the complete removal of the CLPG.

The photophysical properties of all five thiolcarbamate coumaranones have been characterised as well. For each compound, a stock solution with a concentration of  $c = 10^{-2} \text{ mol l}^{-1}$  in acetonitrile was prepared. To initiate the reaction, 50 equivalents of DBU were added to 1 ml of the respective stock solution in a cuvette. After brief and thorough mixing, the progress was monitored over several scans, with each one requiring 45 seconds. The cuvette itself was not sealed to reduce oxygen deficiency. Overall, three distinctive sets of experiments were conducted: (a) duration and intensity of CL immediately after the addition of DBU without external excitation, (b) investigation of the new emitting species with external excitation after CL was finished and (c) excitation spectroscopy for the previously observed photoluminescence.

Since the NMR results of **3a** and **3e** indicate the consistent release of the corresponding thiolate from the CLPG, similar decomposition patterns and potent CL were anticipated for all derivatives. Fig. 3 presents the luminescence spectra of **3a**.



**Fig. 3** Emission and excitation spectra of **3a** in acetonitrile (every scan took 45 s): (a) emission scans of CL (*i.e.* without external excitation, slit: 0.4), the arrow indicates the decrease in CL; (b) emission scan after CL was finished (external excitation  $\lambda_{\text{ex}} = 410 \text{ nm}$ , slit: 0.7); (c) excitation scan after CL was finished ( $\lambda_{\text{em}} = 465 \text{ nm}$ , slit: 0.7).



Initially, the CL exhibited an emission peak at 435 nm, which shifts constantly to 455 nm with significant intensity loss after the 10th scan (Fig. 3a), aligning with the NMR experiment. Subsequent excitation at 410 nm revealed a fluorescent species with an emission peak at 468 nm (Fig. 3b) and the corresponding excitation scan (Fig. 3c,  $\lambda_{em} = 465$  nm) showed strong absorption at  $\lambda_{max} = 411$  nm.

In the case of the protected aliphatic thiol derivative **3e**, the CL maximum was centered around 446 nm (Fig. 4a, top). After the oxidation process is finished, a fluorescent species is produced emitting at 452 nm and showing a corresponding absorption at 407 nm (Fig. 4b and c, top). In a large-scale experiment (Fig. 4, bottom), the strong blue CL could be easily detected even in an illuminated room. After 5 minutes, the CL decreased and the solution appeared pink and eventually leads to an orange-coloured solution after 7 minutes. Fig. 5 displays the luminescence spectra of **3b–d**. The CL emission of **3b** peaks between 450 and 455 nm and was perceived to be weaker both visually and in terms of counts per second (Fig. 5a). This might originate from a CL quenching contribution from the aromatic thiolurethane group, which competes with the fluorescence and reduces the total CL but not the release rate of the thiolate

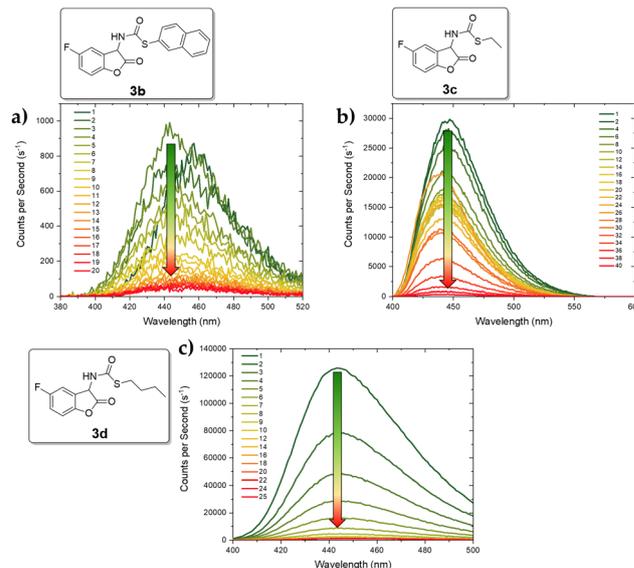


Fig. 5 CL spectra of **3b–d** in acetonitrile (every scan took 45 s): (a) emission scans of CL of **3b** (*i.e.*, without external excitation, slit: 0.2); (b) emission scans of CL of **3c** (*i.e.*, without external excitation, slit: 0.8); (c) emission scans of CL of **3d** (*i.e.*, without external excitation, slit: 0.2). The arrow indicates the decrease of CL.

group. The fluorescent species in this reaction has similar properties to **3a**, emitting at 467 nm and an excitation scan showed an absorbance with a maximum at 409 nm (see ESI†).

Comparison among the protected aliphatic thiol derivatives (**3c** to **3e**) revealed a decrease in CL duration with longer alkyl side chains. As mentioned above, **3e**'s CL fades after 5 scans/4 minutes, while that of **3d** remained for 7 scans/5 minutes (Fig. 5c). Derivative **3c** exhibited the longest emission of CL lasting for 36 scans/27 minutes (Fig. 5b). In contrast, the CL maxima of all three derivatives are very similar, ranging from 443 to 446 nm. Additionally, emission and excitation scans of the fluorescent byproducts post-CL decay exhibited similar photophysical properties across all three derivatives as well.

Table 2 summarises the results of all luminescence experiments, exhibiting strong CL between 443 and 455 nm for each derivative, with variations in duration influenced by alkyl side chain length or aromatic system complexity. The generated fluorescent species demonstrated similar photophysical properties, with **3c** showing the most variation in absorption and fluorescence maxima at 387 nm and 443 nm, respectively.

Following the release of the protected substrate, it is generally anticipated that a benzoxazinedione derivative (*e.g.* **6-A**)

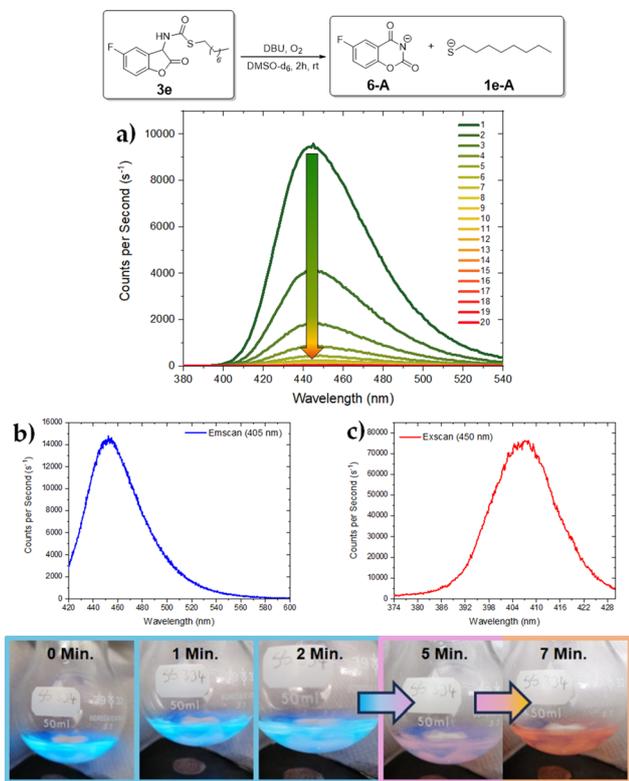
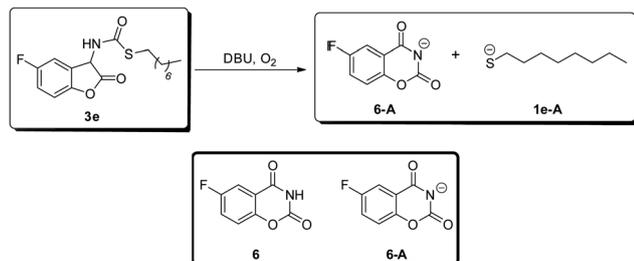


Fig. 4 (top) Emission and excitation spectra of **3e** in acetonitrile (every scan took 45 s): (a) emission scans of CL (*i.e.* without external excitation, slit: 0.2), arrow indicates the decrease in the CL; (b) emission scan of **3e** after CL was finished (external excitation  $\lambda_{ex} = 405$  nm, slit: 0.5); (c) excitation scan of **3e** after CL was finished ( $\lambda_{em} = 450$  nm, slit: 0.5). (bottom) Photographs of the CL of coumaranone **3e** ( $1.5 \times 10^{-2}$  M) in a round-bottom flask (50 mg were dissolved in 10 ml MeCN and 2 equivalents of DBU were added).

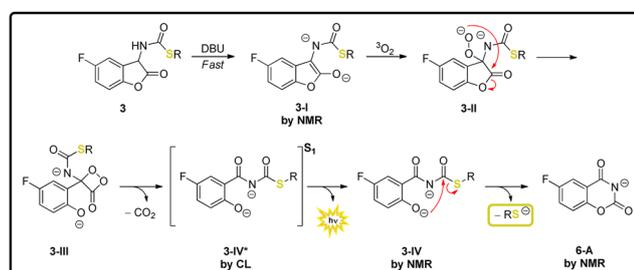
Table 2 Summary of the photophysical properties of thiolcarbamate-coumaranones. The absorption and fluorescence maxima refer to the residual chromophores after oxidation of the respective coumaranone decomposition

Entry	CL $\lambda_{max}$ [nm]	CL intensity/duration	Abs./Fl. $\lambda_{max}$ [nm]
<b>3a</b>	435/453	Bright/5 min	411/468
<b>3b</b>	450–455	Bright/4 min	409/467
<b>3c</b>	443	Very bright/30 min	387/443
<b>3d</b>	444	Very bright/5 min	415/450
<b>3e</b>	446	Very bright/4 min	407/452





**Scheme 3** (top) Decomposition of coumaranone **3e** with oxygen and DBU. The release of **1e-A** and the byproduct of the benzoxazinedione **6-A** could be verified by NMR as (bottom) benzoxazinedione in its protonated **6** and deprotonated state **6-A**.



**Scheme 4** Proposal for the general mechanism of thiolcarbamate-coumaranones and the methods used for the reconstruction.

emerges as the final product (beside the released component), which was also observed in the reaction of thiolcarbamate-coumaranone **3e**: this would lead to the assumption that it serves as the fluorescent species after CL is finished (Scheme 3). Additional theoretical calculations regarding its photophysical properties were conducted (see ESI†). The calculated data indicated that for efficient fluorescence, the terminal benzoxazinedione ought to be protonated.

Moreover, the absorption and emission maxima wavelengths calculated for the  $S_0 \rightarrow S_1$  and  $S_1 \rightarrow S_0$  transitions do not align with those of the fluorescent species detected after the oxidation of thiolcarbamate-coumaranone **3e**. Consequently, it is plausible that a more improbable and different fragmentation process takes place that deviates from the mechanisms established in the literature and could be responsible for the formation of different fluorescent species.

With respect to the deprotection mechanism, the applied analytical methods make it possible to follow and comprehend the relevant steps of the overall process. The initiation *via* deprotonation can be traced *via* NMR, showing that the  $\alpha$ -CH proton disappears as soon as DBU has been added (Scheme 4, **3-I**). It is noteworthy that the signal of the NH-proton vanishes as well. After the formation of the peroxide anion **3-II** and subsequent decarboxylation of the 1,2-dioxetanone intermediate **3-III**, the emitting species **3-IV\*** can be tracked through its bright CL. The rate of thiolate release correlates with the leaving group ability (aromatic thiolates are faster) and also influences CL intensity and wavelength. The weaker CL emission for **3a,b** can be due to an interaction between FG and the fluorophore, resulting in a stronger fluorescence quenching of **3-IV\*** (see Table 2). The degradation of all protected thiols also appears to be

comparable to each other, as the photophysical properties of the detected fluorescent species do not show major differences, with the exception of the absorption maximum of the fluorescent species of **3c**. It remains unknown which exact species is responsible for the obtained absorption and fluorescence values. Compound **6-A** can be discarded as one potential option.

Overall, thiolcarbamate-coumaranones represent a potent class of CLPG with bright blue/greenish CL that allow fast and efficient release of both aromatic and aliphatic thiols. Aromatic thiols are released faster, possibly due to a literature-described dark side reaction.<sup>16</sup> Further insight was gained into the relationship between the duration and emission maximum of the CL and the potential nucleofuge. Recent applications of thiol-PGs concern the de-protection of thiols (especially cysteine)<sup>1,19</sup> to allow a more facile synthesis of peptides, peptoids and proteins,<sup>20,21</sup> and new chromatic orthogonal PRPGs.<sup>22–24</sup> Considering the importance of thiols for antioxidant properties, the formation of disulfide bonds to stabilise protein structures, applications in chelation therapy and medicinal applications, the release of thiol-protected molecules *via* CLPGs could open up new opportunities for biological and medical research.

## Data availability

The data supporting this article have been included in the ESI.†

## Conflicts of interest

There are no conflicts to declare.

## Notes and references

- 1 T. W. Greene and P. G. M. Wuts, *Protective Groups in Organic Synthesis*, John Wiley & Sons, Hoboken, NJ, 5th edn, 2014.
- 2 S. Vidal, *Protecting Groups: Strategies and Applications in Carbohydrate Chemistry*, Wiley-VCH, Weinheim, Germany, New York, NY, USA, 1st edn, 2019.
- 3 K. Jarowicki and P. J. Kociński, *J. Chem. Soc., Perkin Trans. 1*, 2001, 2109–2135.
- 4 P. J. Kociński, *Protecting Groups: Foundations of Organic Chemistry*, Thieme, Stuttgart, Germany, 3rd edn, 2005.
- 5 M. Schelhaas and H. Waldmann, *Angew. Chem., Int. Ed. Engl.*, 1996, 35, 2056–2083.
- 6 P. Klan, *et al.*, *Chem. Rev.*, 2013, 113, 119–191.
- 7 P. Wang, *Asian J. Org. Chem.*, 2013, 2, 452–464.
- 8 A. P. Pelliccioli and J. Wirz, *Photochem. Photobiol. Sci.*, 2002, 1, 441–458.
- 9 C. G. Bochet, *J. Chem. Soc., Perkin Trans. 1*, 2002, 125–142.
- 10 A. G. Griesbeck, *et al.*, *Synthesis*, 2017, 539–553.
- 11 A. Herrmann, *Photochem. Photobiol. Sci.*, 2012, 11, 446–459.
- 12 D. E. Falvey and C. Sundararajan, *Photochem. Photobiol. Sci.*, 2004, 3, 831–838.
- 13 T. Lippold, *et al.*, *Photochem*, 2023, 3, 373–389.
- 14 S. Schramm, *et al.*, *Arkivoc*, 2015, 2015, 44–59.
- 15 S. Schramm, *et al.*, *Eur. J. Org. Chem.*, 2016, 678–681.
- 16 S. Schramm, *et al.*, *Phys. Chem. Chem. Phys.*, 2017, 19, 22852–22859.
- 17 R. Graf, *Chem. Ber.*, 1963, 96, 56–67.
- 18 B. Loev and M. F. Kormendy, *J. Org. Chem.*, 1963, 28, 3421–3426.
- 19 R. J. Spears, *et al.*, *Chem. Soc. Rev.*, 2021, 50, 11098–11155.
- 20 Y.-K. Qi, *et al.*, *Org. Biomol. Chem.*, 2016, 14, 4194–4198.
- 21 X. Qiu, *et al.*, *RSC Adv.*, 2023, 13, 2483–2486.
- 22 N. Kotzur, *et al.*, *J. Am. Chem. Soc.*, 2009, 131, 16927–16931.
- 23 J. A. Karas, *et al.*, *Chem. – Eur. J.*, 2014, 20, 9549–9552.
- 24 H. Yang, *et al.*, *ChemistrySelect*, 2022, 7, 1–4.

