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Dual-functionality of CZTS nanoflakes: as an anode material for lithium-ion batteries and as a counter electrode in DSSCs – a DFT and experimental investigation[†]

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This research work aims to develop a new dual-functional electrode material suitable for both lithium-ion batteries (LIBs) and dye-sensitized solar cells (DSSCs). Nanostructured Cu₂ZnSnS₄ (CZTS) was synthesized through the solvothermal method. Structural properties analysed through the X-ray diffraction pattern (XRD) and Raman spectra reveal the formation of the CZTS with kesterite structure. The stoichiometry and the oxidation states of CZTS have been analyzed using X-ray photoelectron spectroscopy (XPS). The core level XPS spectra of Cu 2p, Zn 2p, Sn 3d, and S 2p confirm the presence of the constituent elements in the required oxidation states (Cu⁺, Zn²⁺, Sn⁴⁺, S²⁻). The surface morphology of the CZTS nanoparticles showed a nanoflake-like structure with a surface area of 34.20 m² g⁻¹. The geometrical optimization, electronic, and optical properties were calculated using DFT calculations. The semiconducting material CZTS is electrochemically active toward Li, which can be used as an alternative anode material for lithium-ion batteries offering potential improvements in cycling stability and specific capacity. The electrochemical studies of the CZTS nanoflakes exhibited a specific capacity of 1141.08 mA h g⁻¹ and 350 mA h g⁻¹ at 0.1C and 1C rates respectively. The cycling stability of CZTS at a high scan rate of 1C, and the specific capacity of 220 mA h g⁻¹ over 70 cycles with 73% coulombic efficiency, suggest it to be a promising alternative anode material in the next-generation lithium-ion batteries. The performance of CZTS as a counter electrode in dye-sensitized solar cells was also explored. The DSSC constructed with CZTS as the counter electrode showed an efficiency of 5.9%.

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Introduction

Technological advancements today have increased the need for energy-efficient and smart materials with multiple utilities.¹ On the other hand, rapidly expanding CO₂ discharges and dependence on non-renewable energy source assets have become

worldwide energy concerns. Technologies to meet the energy requirements of the modern world with negligible pollution to the environment remain the top priority of research today. Among the different technologies, solar cells, lithium-ion battery technologies and electrochemical supercapacitors^{2,3} have emerged as the most promising choices for energy harvesting and storage. Compared to traditional lead-acid batteries, lithium-ion batteries have significant advantages in terms of cycle life, power and energy density, efficiency design, flexibility, and large-scale applications.⁴⁻⁸

The lithium-ion battery energy storage and supply mechanism involves inserting and extracting lithium ions within the electrode materials during charge and discharge cycles. The performance of the battery is primarily determined by its energy and power densities. To enhance these properties and cycle life, careful selection of electrode materials and electrolytes is essential. Cathode materials store energy through intercalation or conversion reactions, while anode materials engage in intercalation, conversion reactions, or alloying/dealloying for energy storage. Depending on the specific electrode material, one or more of

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these mechanisms can occur, directly impacting the battery's overall performance.^{9,10} Thus, the characteristics of the anode material play a crucial role in determining the electrochemical performance of lithium-ion batteries.

Graphite has been widely used as a primary anode material in commercial lithium-ion batteries owing to its advantageous characteristics such as conductivity, good reversibility, and relatively low cost. However, its energy density is limited because the six carbon atoms can only accommodate one lithium atom. Silicon has emerged as a promising alternative with approximately ten times more capacity than graphite.¹¹⁻¹³ However, the significant volume expansion of silicon during lithium insertion and extraction leads to issues such as pulverization, the formation of an unstable solid-electrolyte interphase (SEI), and the loss of electrical contact, resulting in capacity degradation and limited cycle life.¹⁴

Extensive research efforts have focused on finding alternative materials that can offer superior energy capacity and cycling stability compared to graphite and silicon.¹⁵ Tin-based lithium storage materials, known for their high theoretical capacity, have emerged as a promising substitute for graphite in the next generation of lithium-ion batteries.¹⁶ Among these materials, tin-sulfide stands out with its superior cycling stability and significantly larger reversible capacities, making it one of the most prospective candidates for high-performance anode materials in lithium-ion batteries.¹⁷⁻²¹ However, the long-term stability of tin-sulfide materials and their suitability within specific voltage ranges have not met the desired standards. To enhance the capacity characteristics of tin-based anodes without compromising energy density, numerous strategies have been investigated. One such strategy involves designing nanostructured electrode materials with significant advantages,¹⁶ including a short diffusion path of lithium ions, easy penetration of electrolytes, and highly available charge storage sites due to the high surface area. An alternative approach involves creating composite structures by combining tin-based materials with conducting materials, introducing elements such as (Cu, Zn, S) into Sn-based materials, improving the electrical conductivity, and serving as a buffer matrix during Li insertion.^{22,23} Other materials such as porous TiO₂ wires,²⁴ microspheres,²⁵ and porous MnO₂/C nanospheres²⁶ have also been utilized as anode materials in Li-ion batteries.

Various researchers have used different methods to synthesize nanostructured CZTS; Zhou *et al.*²³ synthesized flower-like Cu₂ZnSnS₄ nanostructures *via* a solvothermal technique, demonstrating stable lithium-ion reversibility, but with a lithium storage capacity of 150 mA h g⁻¹. Li *et al.*¹⁷ synthesized kesterite Cu₂ZnSnS₄ nanocrystals using a cost-effective wet chemical process, achieving an initial 540 mA h g⁻¹ discharge capacity. Meanwhile, Yang *et al.*²⁷ synthesized CZTS nanocrystals with a microwave-assisted approach, but with a reversible capacity of 288 mA h g⁻¹. Therefore, a reasonable design of electrode materials with cycling ability and high-rate performance is needed.

Exploring a variety of applications of the materials is essential as it will be useful in the future for the large-scale production of similar materials for various applications. Dye-sensitized

solar cells (DSSC), which fall under the group of third-generation solar cells, have attracted interest for their variety of advantages, such as their capacity to work under diffused light conditions, and lightweight and semi-transparent nature, which makes them an attractive candidate for window applications.²⁸ In DSSCs, mesoporous TiO₂ sensitized with a dye forms the photoanode layer,²⁹ and a thin layer of platinum coated over conducting substrates has been used as the counter electrode. However, to reduce the manufacturing cost towards commercialization, various other counter electrode materials have also been tested. Among them, CZTS has also been reported as an efficient counter electrode material, and it also has the unique edge of being made of earth-abundant material that can be relatively cheaper than platinum-based counter electrodes.^{30,31}

In the present work, we developed nanostructured Cu₂ZnSnS₄ (CZTS) for dual applications, as an alternative anode material for lithium-ion batteries and as a counter electrode in DSSCs, by taking advantage of our earlier experience in the fabrication of widely recognized CZTS solar cell absorber materials in thin film solar cells (TFSCs).²⁷ In this study, kesterite CZTS nanoflakes were synthesized through a simple solvothermal method, without using expensive vacuum techniques.

Experimental

(a) Synthesis of the kesterite CZTS

All the chemicals employed in this study were of analytical reagent (AR) grade and used without additional purification. Cu₂ZnSnS₄ nanoparticles were synthesized *via* the solvothermal method without using catalysts, capping agents, or surfactants. To achieve this, stoichiometric amounts of (Cu (NO₃)₂), (Zn (NO₃)₂), (SnCl₄), and thiourea were dissolved in a mixture of 2-methoxy ethanol and ethanolamine at room temperature. The resultant solution was stirred for 45 minutes using a magnetic stirrer, producing a homogeneous, clear, and transparent solution. Subsequently, the solution was transferred into an autoclave and maintained at 90 °C for 12 hours. Afterward, the resulting precipitates were centrifuged at 5000 rpm for 10 minutes, followed by several rinses with deionized water. The precipitates were dried at 70 °C for 24 hours and then annealed at 200 °C for 1 hour.

(b) Characterization

The sample's structure and crystallinity were assessed using an X-ray diffractometer (Bruker D8 Advance) equipped with a Cu K α radiation source ($\lambda = 0.15406$ nm) operating at a voltage of 40 kV and a current of 40 mA. A Renishaw instrument with a wavelength of 514.5 nm was employed to analyze the Raman spectrum. A field electron scanning electron microscope (CARL ZEISS-SIGMA) was used to examine the CZTS nanoparticles for surface morphology analysis. The microstructural properties of the CZTS nanoparticles were investigated using a high-resolution transmission electron microscope (HRTEM) with the JEOL JEM-ARM200F instrument. For the characterization of the surface and chemical states, X-ray photoelectron spectroscopy (XPS) was employed, using the Omicron ESCA system manufactured by



Oxford Instruments Germany, which is a powerful technique for surface analysis providing elemental information and chemical state analysis (electron spectroscopy for chemical analysis (ESCA)). The surface area and pore size distribution of the CZTS nanoparticles have been investigated using the Quanta chrome Nova 2200 instrument. *I-V* measurements for the fabricated cells were carried out using a Keithley 2450 source meter under the illumination of 1 Sun with an ORIEL Solar Simulator under standard conditions. Incident photo to current conversion efficient (IPCE) curves of the DSSCs were performed using a system consisting of a Xe lamp (300 W) with an Oriel Newport Corp., monochromator.

(c) Electrochemical measurements

CR2032 coin-type cells were employed in an automatic battery tester system (Bio-logic SAS, France) for charge/discharge studies. The working electrode was prepared by a mixture of 70 wt% as-prepared powders (active material), 20 wt% carbon black, and 10 wt% poly (vinylidene fluoride) (PVDF) binder and NMP (*N*-methyl pyrrolidone) used as a solvent to make a slurry. The homogeneous slurry was coated on the copper foil and dried at 70 °C overnight in a vacuum oven. Lithium chips served as the counter electrode, and the electrolyte was 1 M LiPF₆ dissolved in a mixture of ethylene carbonate (EC) and dimethyl carbonate (DEC) in a 1:1 ratio (v/v). All test cells were assembled inside an argon-filled glove box with oxygen and moisture levels below 1 ppm each to maintain an oxygen and moisture-free environment. The cells underwent discharge and charge cycles at a rate of 0.1C, covering a voltage range of 0.01–3.00 V (vs. Li⁺/Li) at room temperature. The typical masses of electrode materials used in the experiments ranged from 2 to 3 mg. Cyclic voltammetry was also conducted using the Bio-logic SAS system within the voltage range of 0 and 3.0 V (vs. Li/Li⁺) at a scan rate of 0.5 mV s⁻¹ at room temperature.

(d) Computational studies

The CZTS DFT studies of the structural optimization, electronic structure, and optical properties of the materials were obtained using DFT calculation. The GGA-PBE function was employed in electronic structure computations. All the calculations were carried out using the periodic boundary conditions. A 3 × 3 × 2 Monkhorst-Pack (MP) grid was applied for Brillouin zone integration, and the cutoff energy was set to 498 eV. Specific criteria were determined to attain geometric convergence, including maximum force, most significant shift, maximum stress, and most remarkable energy alteration, which were assigned values of 1.0×10^{-5} , 0.03 eV Å⁻¹, 0.05 GPa, and 0.001 Å, correspondingly.

Results and discussion

(a) Structural study

The structural properties of the CZTS nanoparticles have been studied using the X-ray diffraction technique. The X-ray diffraction pattern of the CZTS nanoparticles annealed at 200 °C is compared with their simulated pattern, as shown in Fig. 1. All the predominant

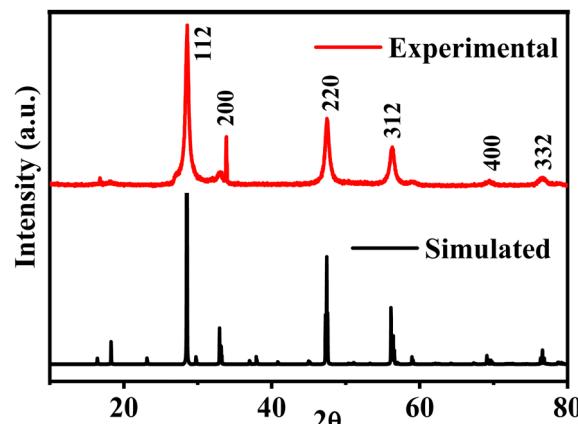


Fig. 1 A comparison of the X-ray diffraction pattern of the experimental CZTS nanoparticles with the corresponding simulated pattern.

peaks of the experimental pattern are comparable to the simulated pattern. The diffraction peaks observed at the 2θ positions of 28.3, 33.6, 47.3 and 55.2 were indexed to the (112), (200), (220) and (312) planes of CZTS with the tetragonal kesterite structure of Cu₂ZnSnS₄ (JCPDS, card no. 26-0575). Additionally, to comprehend the crystal structure, the Rietveld refinement analysis was executed to obtain full structural parameters. The refinement results of the synthesized sample show the $\bar{I}42m$ (121) space group. In the refined structure, the 4d Wyckoff position is particularly occupied by Zn and Cu, the 2a is engaged by Cu, the 2b position by Sn, and the 8i position is occupied by S. The calculated lattice parameters *a*, *b*, and *c*, and cell volume (*V*) are 5.43, 5.43, 10.80, and 318.8, respectively. Refined crystal structure data and the structure are shown in the ESI,† S1 and Table ST1, respectively. In the kesterite structure S1, each sulfur anion forms bonds with four cations, and each cation is bonded to four sulfur anions. The crystallographic *c*-direction exhibits alternating layers of cations (CuZn/SS/CuSn/SS) and sulfur anions.³²

As a quaternary compound, CZTS often contains various binary and ternary phases, making it challenging to control its stoichiometry precisely. The lack of sulfur during the formation of CZTS leads to several secondary phases, including ZnS, CuS, Sn₂S₃, SnS₂, and Cu₂SnS₃. These secondary phases have crystal structures similar to CZTS, resulting in overlapping peaks in the X-ray diffractogram, making it difficult to distinguish CZTS from the secondary phases. Moreover, numerous reports indicate that CZTS tends to lose its anion (sulfur) during the annealing process due to the high volatility of sulfur. Consequently, when CZTS is annealed without a sulfur atmosphere, it exhibits the presence of secondary phases.³³ Raman spectroscopy provides a more reliable indication of the presence of secondary phases in CZTS, as each phase in CZTS exhibits distinct peak positions in Raman scattering which is more distinct than in the X-ray diffraction pattern. The Raman spectrum of CZTS nanoparticles is shown in Fig. 2. The vibration frequency at 331 cm⁻¹ confirms the formation of kesterite CZTS. The strong peak position at 331 cm⁻¹ corresponds to the A1 mode of CZTS. The A1 modes in CZTS nanoparticles represent pure anion vibrations, specifically involving sulfur atoms



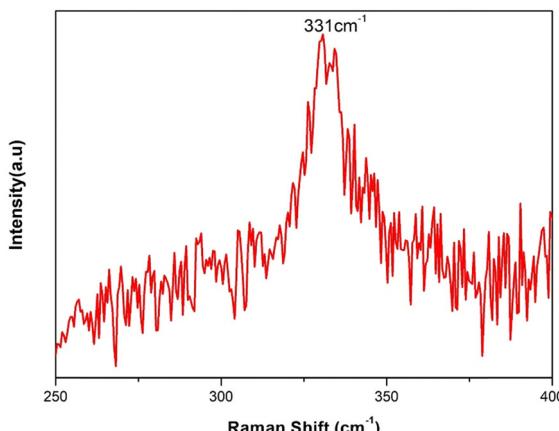


Fig. 2 Raman spectrum of kesterite phase CZTS nanoparticles.

surrounded by neighbouring atoms that remain motionless.³⁴ The broadening of these modes indicates the phonon confinement effects within the nanoparticles.

(b) Morphological and compositional study

Fig. 3(a) shows the surface morphology of the CZTS nanoparticles *via* the field emission scanning electron microscope (FESEM) images. The FESEM image reveals the formation of CZTS nanoflake-like structures. It can be seen that the CZTS nanoflakes are composed of dense and homogeneous structures that are uniform in nature.

Also, the energy dispersive X-ray analysis (EDAX) has been used to study the elemental composition of the CZTS nanoparticles. The EDAX spectra of the synthesized CZTS nanoparticles are shown in Fig. 3(b). EDAX analysis reveals that the constituent elements Cu, Zn, Sn, and S of CZTS are present in the sample, and it shows the formation of CZTS with almost stoichiometric composition. The chemical composition of the CZTS particles was found to be Cu 22.4 at%, Zn 12.2 at%, Sn 15.8 at%, and S 49.6 at%, respectively, within the nanoparticles.

High resolution (TEM) has been used to analyze the micro-structure of the CZTS nanoparticles. For high-resolution transmission electron microscopy analysis, the prepared CZTS nanoparticles were dispersed in acetone, and the solution was sonicated for 10 minutes.

The high-resolution transmission electron microscope (HRTEM) images of the CZTS nanoparticles annealed at 200 °C are shown in

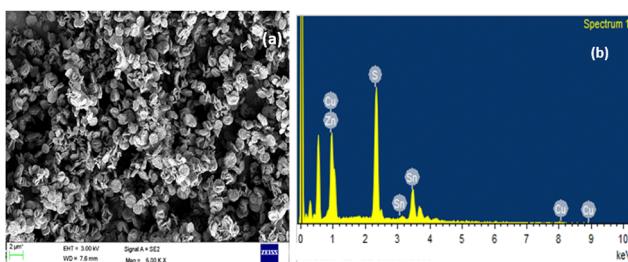


Fig. 3 (a) FESEM image of CZTS ($\text{Cu}_2\text{ZnSnS}_4$) nanoparticles and (b) EDAX spectrum of CZTS nanoparticles.

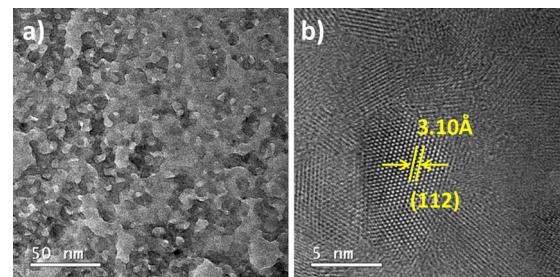


Fig. 4 (a) and (b) High-resolution transmission electron microscope images of CZTS nanoparticles.

Fig. 4. The HRTEM image shows the lattice fringes, indicating the formation of nanocrystalline CZTS nanoparticles. The HRTEM image exhibits lattice fringes with a d spacing of 3.10 Å corresponding to the (112) reflection of the kesterite CZTS.

(c) Surface study

X-ray photoelectron spectroscopy (XPS), a surface-sensitive technique, was utilized to examine the stoichiometry and oxidation states of the constituent elements in the CZTS nanoparticles. Fig. 5 shows the core level spectra of Cu 2p, Zn 2p, Sn 3d, and S 2p obtained from the XPS analysis. The peak of Cu 2p is split into two peaks (Fig. 5b) and is observed at 932.1 eV ($2p_{3/2}$) and 951.9 eV ($2p_{1/2}$) with a splitting energy of 19.8 eV, indicating the formation of Cu^+ . The 2p peak of Zn (Fig. 5a) was observed at the binding energies of 1022.01 ($2p_{3/2}$) and 1044.9 eV ($2p_{1/2}$) with a splitting energy of 22.9 eV corresponding to $\text{Zn}^{(II)}$. This peak separation was due to the formation of Zn^{2+} . The peak of Sn 3d (Fig. 5 c) split into 486.3 ($2p_{5/2}$) and 494.7 eV ($2p_{3/2}$) with a splitting energy of 8.4 eV, and this is attributed to Sn(IV) due to the so-called final state effects. The 2p peak of S (Fig. 5d) was located at 161.35 eV and 162.55 eV with splitting energy of 1.20 eV, which is consistent with the 160–164 eV range expected for S in sulfide phases,³⁵ which also indicates that sulfur is present in the S^{2-} state.³⁶ The S 2p spectra show the region which belongs to sulfur emission. The peaks fitted at 161.35 and 162.55 are

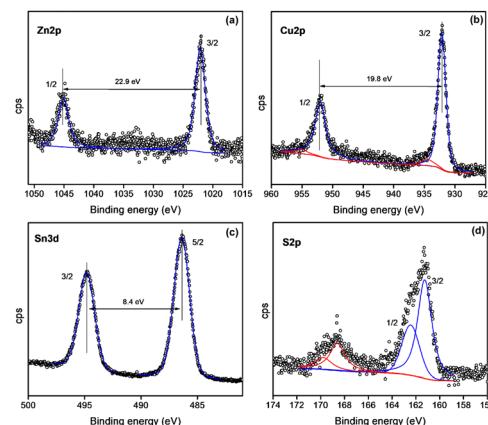


Fig. 5 X-ray photoelectron spectra of (a) Zn 2p, (b) Cu 2p, (c) Sn 3d, (d) S 2p of CZTS nanoparticles.



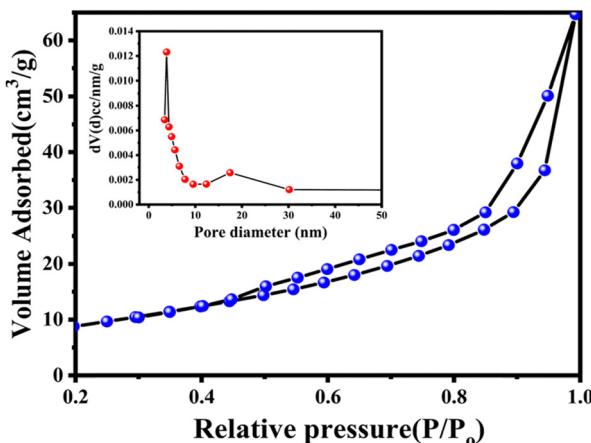


Fig. 6 BET surface area analysis of a $\text{Cu}_2\text{ZnSnS}_4$ nanoparticle. The inset depicts the Barrett–Joyner–Halenda pore size distribution curves.

assigned to S 2p_{3/2} and S 2p_{1/2} respectively, which indicates that the sulfur is in the S²⁻ oxidation state.³⁶

The textural properties, such as surface area and the total pore volume of the CZTS nanoparticles, were evaluated based on N_2 adsorption–desorption isotherms shown in Fig. 6. The inset depicts the pore size distribution of CZTS nanoparticles. Based on the Brunauer–Deming–Deming Teller (BDDT) classification, physisorption isotherms can be categorized into six types, and the hysteresis loop shapes are used to determine the specific pore structure. CZTS nanoparticles correspond to type-IV isotherms and type-H3 hysteresis loops. The observed BET surface area of CZTS is $34.20 \text{ m}^2 \text{ g}^{-1}$. The Barrett–Joyner–Halenda (BJH) pore size distribution curve (inset) indicates the high degree of uniformity of the pores in the range of 3.3 nm for CZTS.

(d) DFT study

We executed the structural optimization, electronic structure, and optical property simulations through DFT calculations employing the Cambridge Serial Total Energy Package (CASTEP) code. After the structural optimization, we executed electronic structure calculations to obtain the density of states and band structure. Fig. 7 and Fig. S2 (ESI†) correspondingly exhibit the band structure and the site-projected partial density of states (PDOS). A dotted line in the electronic structures is denoted as the Fermi level. According to the band structure, the valence band maximum (VBM) and the conduction band

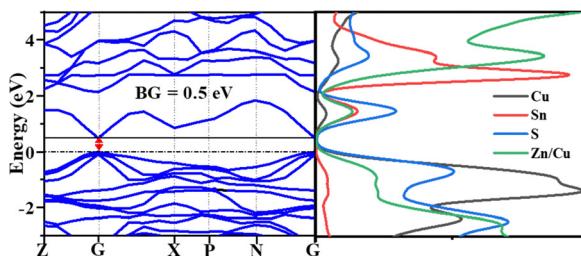


Fig. 7 The band structure and density of states of the CZTS nanoparticles.

minimum (CBM) appear near the G-position with a 0.5 eV narrow bandgap of 0.5 eV. This value is underestimated compared to reported structures because of the use of the GGA-PBE functional owing to computational constraints. However, the features and impact of the band and density of states remain comparable to those found using the HSE functional, offering a valuable understanding of the material properties. From the band structure, the position of the VBM and the CBM indicates that the material is a narrow band p-type direct semiconductor.

The density of states (DOS) indicates that the electron from Cu dominates the VBM, and the CBM is by the S atoms. Furthermore, to identify in-depth, the site-projected DOS was observed for the Wyckoff positions (4d-Cu/Zn, 2a-Cu, 2b-Sn, 8i-S), as shown in ESI,† S2.

Notably, the CBM region is predominantly influenced by the Cu-3d electron from the 2a Wyckoff position. Similarly, in the same region, there is a significant contribution from the S 3p electrons, indicating the Cu–S bonding in the sample. Equally, in the CBM, Sn 5s and 5p-electrons and the S 3p dominate. The mixed 4d Zn/Cu position does not show a significant contribution at the Fermi level, but it showed a vital contribution below and above the Fermi level.

Fig. S3 (ESI†) depicts the calculated optical absorption spectra compared with the standard solar spectrum within the 200–800 nm range. The CZTS nanoparticles show a wide range of absorption of the UV-Vis spectral region. Notably, several absorption peaks were observed due to the electronic transitions between the VB and CB, which includes inter-band transition occurring in the ultraviolet-visible (UV-Vis) region. The transition includes the Cu 3d to S 3p electrons and Cu 3d to Sn 4s and 4p electronic transition from VB to CB. The identified optical absorption peaks match the standard solar spectrum range, supporting the sample's potential as a highly promising photoactive material, especially for photovoltaic applications.

(e) Electrochemical study

To study the lithium storage mechanism, cyclic voltammetry (CV) was carried out for the CZTS electrode, covering the potential range from 0.01 V to 3 V at a scan rate of 0.5 mV s^{-1} . The CV profiles depicted in Fig. 8a and b illustrate the specific behaviour during the first 3 cycles. During the cathodic scan upon the initial lithiation, the cyclic voltammogram reveals four broad reduction peaks consistent with typical metal sulfide electrode behaviour. Notably, the four initial reduction peaks occurring around 2.5 V correspond to the lithiation of $\text{Li}-\text{Cu}_x\text{S}$ alloys, while two additional peaks around the voltage range (1.4–1.6 V and 1.1 V to 0.8 V) are attributed to the involvement of Sn and Zn species in multistep reactions, as well as the formation of a solid electrolyte interface (SEI layer formation).

Moreover, the final reduction peak identified near 0.4–0.5 V is associated with the conversion reaction between Sn and Zn. In contrast, during the anodic scan, four oxidation peaks are evident at 0.55, 1.3, 1.9, and 2.5 V. The peak at 0.55 V corresponds to the de-lithiation of Sn and Zn alloy, while those at 1.3 and 1.9 V involve reversible reactions among Sn, Zn, and Li. The broad peak spanning 2.4 to 2.5 V can be assigned to the



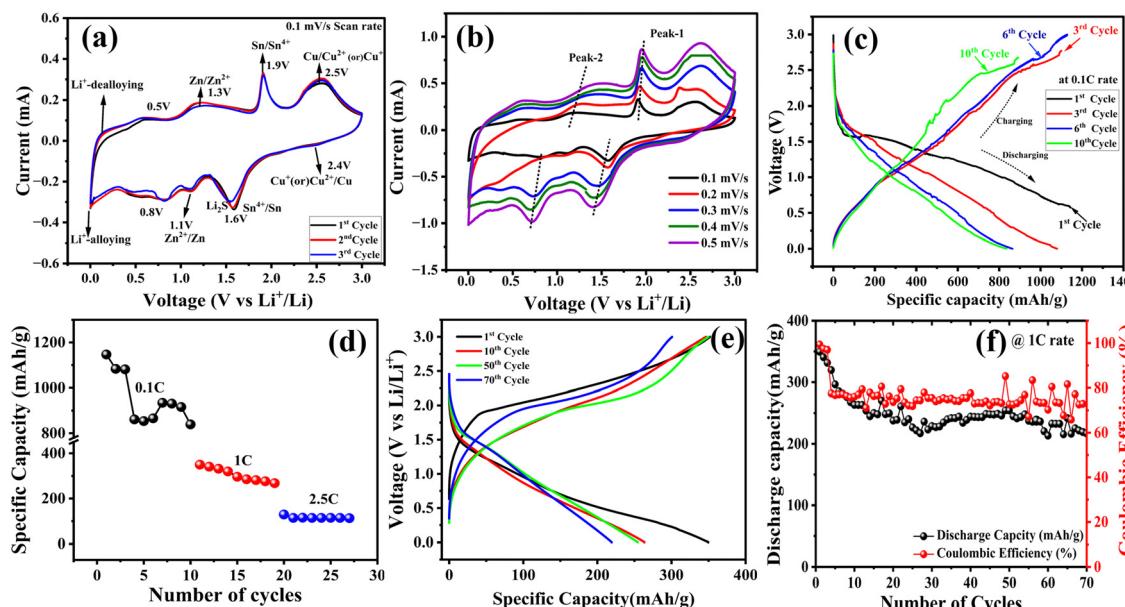


Fig. 8 (a) CV curves for 3 cycles at 0.1 mV s^{-1} scan rate, (b) CV curves at different scan rates, (c) galvanostatic charge/discharge cycles of CZTS nanoparticles at 0.1C rate, (d) rate capability curves at C-rate from 0.1 to 2.5C, (e) galvanostatic charge/discharge at 2.5C rate and (f) cycling performance and coulombic efficiency at 1C rate.

delithiation of $\text{Li}-\text{Cu}_x\text{S}$ alloys and the oxidation of Li_2S to S and lithium ions. Notably, the CV curves of the 2nd and 3rd cycles exhibit significant overlap, confirming the material's stability.^{37–39}

Fig. 8(b) shows the CV at various scan rates ranging from 0.1 to 0.5 mV s^{-1} to understand the nanoparticle rate performance of the CZTS nanoparticles. The peak currents of the oxidation and the reduction peaks increase with the increase in the scan rates, which increases the peak potential difference between them. And it confirms that there is no significant deviation in the shape of the curves.

The electrochemical behaviour of the CZTS electrodes was investigated by galvanostatic charge/discharge cycles over the potential range of 0 – 3.0 V , as shown in Fig. 8(c). The CZTS electrodes showed a 1st cycle discharge capacity of 1143 mA h g^{-1} , which is higher than the theoretical value of 847 mA h g^{-1} and corresponding discharge capacities of the 3rd, 6th, and 10th cycles are 1080 , 864 , and $833.3 \text{ mA h g}^{-1}$. The charge capacities are 1100 , 1128 , and 891 mA h g^{-1} , respectively at 0.1C rate. The high discharge capacity of the 1st cycle is more because of the hierarchical morphology with a large surface area of the CZTS nanoparticles, which will lead to a large amount of irreversible lithium insertion. Also, the extra capacity at the first discharge may be attributed to forming solid electrolyte interphase (SEI) layers on the electrolyte interface.^{23,40} The rate performance, cycling stability, and coulombic efficiency of the CZTS electrodes are illustrated in Fig. 8d–f. Fig. 8d displays the rate studies, showing initial discharge capacities of 1146.7 , 350 , and $129.3 \text{ mA h g}^{-1}$ at different C rates of 0.1C , 1C , and 2.5C , respectively. Regarding cyclic stability, CZTS electrodes initially delivered a capacity of 350 mA h g^{-1} and maintained 221 mA h g^{-1} capacity after 70 cycles, even at a high scan rate of 1C , with a 73% coulombic efficiency (Fig. 8f). These results underline the excellent cycling performance and high-rate capability of the electrode

materials, which are crucial for achieving high-performance lithium-ion batteries.

(f) Kinetic charge storage mechanism

A combined mechanism of faradaic and double-layer non-faradaic charge storage describes the area under the CV curve, requiring the identification of the dominant mechanism in electrochemical reactions for a complete knowledge of electrode material charge storage.

Dunn and co-workers study provides a quantification technique for the charge storage mechanisms, where the total current is considered as the sum of capacitive and diffusion-controlled currents.^{41,42}

$$i(\nu) = i_{\text{capacitive}} + i_{\text{diffusive}} \quad (1)$$

The relation between the peak current (i) and the scan rate (ν) obeys the power law.⁴³

$$i_p = a\nu^b \quad (2)$$

Both a and b are variables that were adjustable parameters, and i_p represents the Cathodic/Anodic peak current and ν denotes the sweep rate.^{44,45}

For this, we take log on both sides of eqn (2), which gives:

$$\log(i) = b \log(\nu) + \log(i) \quad (3)$$

The b -values are calculated from the slope of the plot of $\log \nu$ vs. $\log i$. There are two well-defined conditions: $b = 0.5$ and $b = 1.0$, whereas the value of b ranges from 0.5 to 1 . If the value of b is around 0.5 , then it is a diffusion-controlled process, and if it is close to 1 , then it is a surface-capacitance-dominated process.^{46,47}



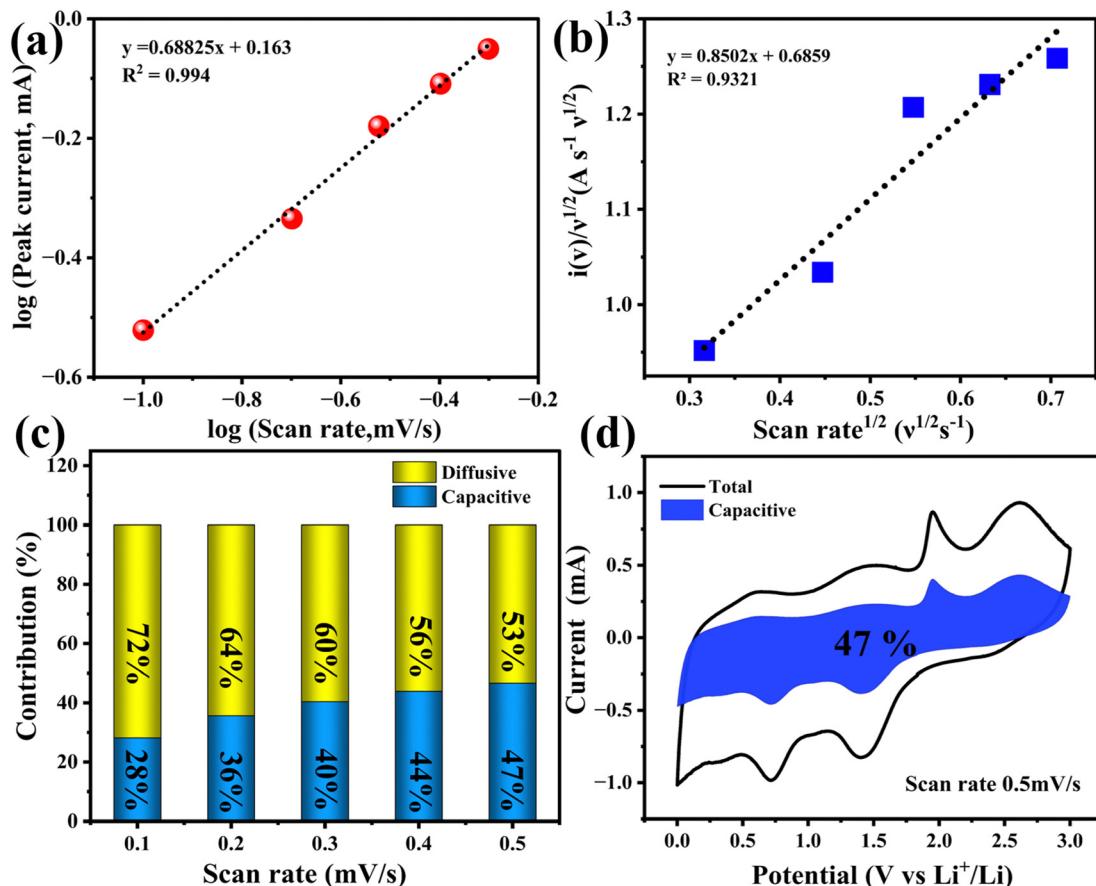


Fig. 9 Kinetics and quantitative capacitive analysis of the lithium storage mechanism in the CZTS. (a) Relationship between logarithmic anodic peak currents and scan rates. (b) Plot of scan rate^{1/2} (V^{1/2} s⁻¹) vs. $i(v)/v^{1/2}$ (A s⁻¹ V^{1/2}) for the anodic sweep. (c) Relative diffusion and capacitive contributions at different scan rates. (d) Pseudocapacitive contribution (shaded region) in the CV curve.

The linear fitting of the logarithmic peak currents and scan rates was performed to obtain the value of parameter b , as shown in Fig. 9(a). The b values for the oxidation peak at 0.5 V and reduction peak at 0.3 V are 0.62 and 0.68 respectively. This value indicates that the lithium storage process of CZTS nanoparticles exhibits a transitive behaviour, which further confirms that the kinetics of the CZTS nanoparticles are a surface controlled capacitive process.

Fig. 9(c) displays a histogram of the contribution of capacitance and diffusion behaviour to the total capacity, which was quantitatively studied at various scan rates. With an increase in the scan rate, the diffusion contribution was reduced, and the capacitive contribution increased. This increase in the capacitive contribution is due to the rapid lithium-ion diffusion across the whole cell and the small nanoscale size of the CZTS nanoparticles, which increases the specific surface area of the material and thereby shortens the lithium transport and diffusion path. The contribution of the capacitance to the total capacity at a specific scan rate of 0.5 mV s⁻¹ was quantitatively analysed to be 46.7%, as shown in Fig. 9(d). This storage behaviour is due to the improved kinetics and charge transfer process of CZTS nanoparticles and is also in agreement with the higher value of parameter b denoting a much-enhanced capacitive kinetics.

(g) Counter electrode in DSSCs

For the construction of DSSC counter electrodes (CEs), the prepared CZTS material was mixed with TritonX 100 and was made into a paste-like consistency. The prepared paste was coated over the pre cleaned FTO substrates through the spin coating technique. The coated substrates were annealed at 350 °C under an argon atmosphere for one hour. The coating thickness was maintained to be 3–4 μm. N719 dye was used as a sensitizer for the preparation of the photo anode. The detailed construction of DSSC is given in the ESI.† To assess the functioning of the CZTS as counter electrodes in DSSCs, the Pt counter electrode as a control device was also studied. The constructed devices had an architecture of glass/FTO/compact-TiO₂/porous-TiO₂ with dye/electrolyte (Iodine based)/CZTS (or Pt)/FTO/glass. Fig. 10(a) depicts the J - V curves of the DSSCs constructed with the prepared CZTS and platinum as counter electrodes respectively. The device constructed with the CZTS as the CE exhibited comparable performance with that of platinum-based counter electrodes. The device with CZTS as the CE exhibited an open circuit voltage (V_{OC}) of 0.71 V, which is comparable to the device made with Pt-based CEs with V_{OC} of 0.72 V. The device exhibited a maximum current density of 13.6 mA cm⁻² and fill factor of 61%



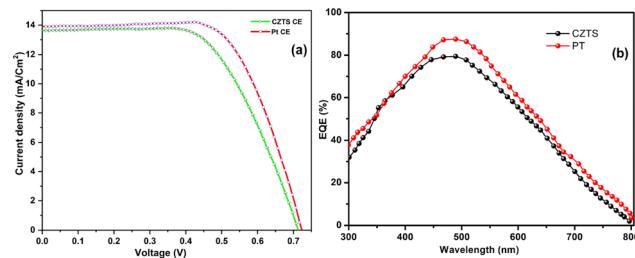


Fig. 10 Photocurrent density–voltage (a) and IPCE (b) curves of the DSSCs prepared with CZTS and Pt-based counter electrodes.

(or 0.61) for CZTS-based CEs, and 13.9 mA cm^{-2} and 64% (or 0.64) for Pt-based CEs. From the J – V graph it can be inferred that the fill factor of the CZTS-based CEs was found to be slightly lower than that of the Pt-based CEs. This could be attributed to the higher series resistance in the device with the CZTS material than that of the Pt-based devices. The efficiency of the devices with CZTS and PT-based CEs was found to be comparable with efficiencies (η) of 5.9 and 6.4%, respectively. Comparing the cost efficiency of the CZTS-based material to noble metals makes it a potential candidate as a DSSC CE. Fig. 10(b) shows the IPCE spectra of the DSSC prepared with Pt and CZTS as counter electrodes. The IPCE response was observed to be in accordance with the J_{sc} values obtained from the J – V results.

Conclusion

In conclusion, this research article explores the electrochemical activity and potential application of quaternary semiconducting material $\text{Cu}_2\text{ZnSnS}_4$ (CZTS), as an alternative anode material for lithium-ion batteries and as counter electrodes in DSSCs. CZTS nanoparticles were synthesized *via* a cost-effective solvothermal method, and their structural properties were confirmed using X-ray diffraction and Raman spectroscopy. X-ray photoelectron spectroscopy was employed to determine the stoichiometry and oxidation states of CZTS. It also exhibited a higher surface area of $34.2 \text{ m}^2 \text{ g}^{-1}$, which led to a large amount of irreversible lithium insertion and a higher degree of uniformity in the range of 3.3 nm. Field-emission scanning electron microscopy revealed the formation of CZTS nanoflakes, and their electrochemical properties showed their potential for high cycling stability and exhibited an initial discharge capacity of $1141.08 \text{ mA h g}^{-1}$, 350 mA h g^{-1} at the scan rate of 0.1C and 1C. The charge storage mechanism revealed the surface dominant behaviour of the CZTS material. This study highlights CZTS as a promising, earth-abundant, and low-cost anode material for next-generation lithium-ion batteries and as a counter electrode for DSSCs.

Author contributions

G. Rajesh – conceptualization; data curation; investigation; methodology, writing – original draft; project administration. Jeyakaruba Palraj – conceptualization; data curation; investigation,

software, writing – original draft. M. R. Venkatraman – investigation, methodology, writing. Ramesh Sivasamy – software, writing review. Sreejith P. Madhusudanan – writing review – editing, Helen Annal Therese – supervision, resources, writing reviewing – editing, Marcos Flores – project administration, supervision, writing review – editing.

Conflicts of interest

There are no conflicts to declare.

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