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## Tailored nonwoven supported non-flammable quasi-solid electrolyte enables an ultra-stable sodium metal battery†

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**An easily scalable approach to developing a nonwoven-supported PVDF-HFP polymer-based quasi-solid-state flexible non-flammable electrolyte shows a conductivity of  $1.16 \times 10^{-4} \text{ S cm}^{-1}$  at RT with a transfer number of 0.68. The  $\text{Na}_3\text{V}_2(\text{PO}_4)_3$  (NVP) cathode-based Na metal cell delivers a discharge capacity of 73 mAh  $\text{g}^{-1}$  at a 2C rate and shows exceptional capacity retention of 98.4% after 2500 cycles.**

Research is focused on electrical energy storage to overcome intermittent power supply issues, rather than just the development of renewable energy sources such as wind, solar, and nuclear power, due to the limited supply of traditional fossil fuels.<sup>1,2</sup> Since their commercialization in 1991, lithium-ion batteries (LIBs) have dominated the market, from portable applications to heavy electric vehicles (HEV), due to their high energy density and stable cycling performance.<sup>3,4</sup> However, they have been unable to meet the ever-increasing market demand for cost-effective energy storage due to resource limitations and price fluctuations in the Li metal supply chain. When considering the potential risks associated with the raw material supply, the potential concerns will have significant implications for battery research and industry.<sup>5</sup> Recently, sodium-ion batteries (SIBs) emerged as an alternative solution to LIBs, as sodium is an earth-abundant element and it offers the same battery chemistry as a LIB.<sup>6–8</sup>

Sodium resources such as  $\text{Na}_2\text{CO}_3$ ,  $\text{Na}_2\text{SO}_4$ , and  $\text{NaCl}$  are readily accessible and affordable from minerals and brines, making them globally available.<sup>9</sup> Further supporting battery research, these highly abundant salts provide reliable sources of sodium. A SIB has a lower energy density due to the higher atomic mass of sodium.  $\text{Na}^+$  has a larger ionic radius which

offers advantages, including increased electrochemical positivity flexibility and reduced polar solvent desolvation energy.<sup>9</sup> The significant difference in ionic radius between  $\text{Li}^+$  and transition metal ions ( $\text{M}^{x+}$ ) often results in rigid material design. The energy density of SIBs can be 1–5 times higher than LIBs, depending on the material chemistry and technology used.<sup>10</sup> Recently, several electrode materials with improved electrochemical performance have been developed that could serve as viable replacements for LIBs. These materials suggest that SIBs could be an excellent alternative to LIBs. The advantages of SIBs over the well-established LIBs include zero energy storage and transport, which are considered to be the ultimate in safety. However, traditional SIBs have been developed with conventional liquid electrolytes that are explosive. The carbonate-based liquid electrolyte reacts with Na metal to form an unstable solid electrolyte interphase (SEI). This eventually reduces the coulombic efficiency (CE) and leads to potential safety risks due to its volatile nature and leakage problems within the battery.<sup>11,12</sup> The practical applicability of lithium-ion, lithium-sulfur, or even sodium-ion batteries is hampered when serious safety issues arise due to overcharge, high-temperature exposure, or internal short-circuiting leading to fire or explosion, particularly during the summer when the temperature may exceed 45 °C.<sup>13</sup>

The safety problem in SIBs is caused by polyolefin-based microporous separators with carbonate-based liquid electrolytes which are highly flammable and volatile, and a high potential of dendrite formation, which affects the cyclability and can cause short-circuiting, a serious safety risk.<sup>14,15</sup> The solution is to use non-flammable electrolytes, as the oxygen generated from the cathode or surrounding air can catalyze organic electrolytes during the explosion, endangering the user. Several strategies have been adopted to overcome the safety issue of batteries, such as adding an ionic liquid and phosphate ester in a carbonate-based electrolyte, incorporating solid-state electrolytes, or improving the cathode, current collector, and battery management system.<sup>16–18</sup> A solid-state electrolyte not only ensures the safety of the battery but increases the energy

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density of the battery. Numerous solid-state electrolytes (SSEs) have been developed as inorganic solid electrolytes (ISE) and solid polymer-based electrolytes (SPE).<sup>19–21</sup> The processing cost of ISE is higher than that of solid polymer electrolyte (SPE). However, poor interfacial contacts in SSEs eventually slow down the kinetics during charge/discharge. Although quasi-solid state polymer electrolytes (QSSEs) ensure a better interface, a large quantity of carbonate/ether-based liquid electrolyte trapped in the polymer matrix, reduces the safety of the electrolyte. Adding non-combustible solvents such as phosphates increases flame retardancy but decreases the mechanical stability of the QSSE. A practical approach has been taken to add inorganic nanofillers to improve mechanical strength, but excessive levels of nanoparticles reduce conductivity. Thus, there should be an optimum balance of all components to make an effective, safe electrolyte for the secondary battery.

Sun *et al.* reported a chloroaluminate ionic liquid electrolyte for sodium-ion batteries, which provides high energy and power density of  $\sim 420 \text{ W h kg}^{-1}$  and  $\sim 1766 \text{ W h kg}^{-1}$ , respectively.<sup>22</sup> A non-flammable electrolyte was reported by Yu *et al.*, consisting of solvents such as phosphate, ether, and fluoroethylene carbonate (termed TMP/F-EPE/FEC) with different Na-salt concentrations, provided 70.8% retention after 500 cycles in pouch cells (NFM||HC; HC): hard carbon.<sup>23</sup> Yang *et al.* demonstrated a Na||NVP cell with a non-flammable quasi-solid-state electrolyte reinforced with bacterial cellulose (BC) with 84.4% capacity retention after 1000 cycles.<sup>24</sup> Zhu *et al.* demonstrated a composite polymer electrolyte comprised of PVDF-HFP and commercial nonwoven fabric, providing conductivity of  $1.38 \times 10^{-3} \text{ S cm}^{-1}$  and excellent electrochemical properties including fast charging at a 20C rate.<sup>25</sup>

Herein, we fabricate a nonwoven-supported, non-flammable quasi-solid-state polymer electrolyte, which we demonstrate as an effective electrolyte for sodium-ion batteries. The cost-effective approach uses a cheap textile (nonwoven mask) coated with poly(vinylidene fluoride-co-hexafluoropropylene) (PVDF-HFP) and soaked in an optimum ratio of sodium-based liquid electrolyte to make it non-flammable and practically applicable for sodium battery applications. The optimized sodium-based quasi-solid-state polymer electrolyte (NaQSSE) shows a room temperature conductivity of  $1.16 \times 10^{-4} \text{ S cm}^{-1}$  and ultra-stable charge/discharge performance with up to 2500 cycles at a 2C rate achieved when coupled with an NVP cathode with a 1.6% capacity loss.

Our simple, scalable, and cost-effective approach to developing QSSE is shown schematically in Fig. 1(a), and the detailed procedure is described in the ESI.† The optical image of the flexible QSSE is shown in Fig. S1 (ESI†). The XRD patterns of the as-prepared electrolytes, P(VDF-HFP), and the nonwoven textile are shown in Fig. 1(b). The characteristic crystal peaks correspond to the crystalline phase of the PVDF-HFP at  $2\theta \approx 17^\circ$  (100),  $19^\circ$  (020),  $26^\circ$  (110), and  $39.3^\circ$  (021).<sup>26–28</sup> The crystalline peaks of the nonwoven textile also show firm peaks, but when coated with PVDF-HFP and soaked in the liquid electrolyte to form a gel, the peak intensity decreases significantly, indicating more amorphous phases, which are eventually needed for

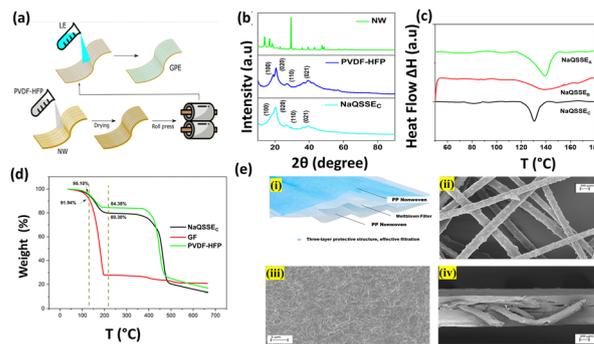


Fig. 1 (a) Schematic representation of the QSSE fabrication. (b) XRD patterns of the as-developed QSSEs and nonwoven (NW) fabric. (c) DSC traces of different QSSEs; (d) TGA analysis of NaQSSE<sub>C</sub> and commercial glass fiber (GF) separator. (e) (i) optical image of nonwoven (mask); SEM images of (ii) nonwoven (NW) fabric, (iii) surface of an NW-SPM, (iv) cross-section of NW-SPM.

faster ion transport. To determine the melting temperature and to correlate the crystallinity of the electrolyte with our X-ray diffraction, the thermal properties were evaluated by differential scanning calorimetry (DSC) (Fig. 1(c)). The melting temperature ( $T_m$ ) and the melting enthalpy of the PVDF-HFP were found to be  $141^\circ \text{C}$  and  $104.7 \text{ J g}^{-1}$ , respectively.<sup>29</sup> The degree of the crystalline phase ( $X_C$ ) was calculated using the following equation:

$$X_C = \frac{\Delta H_m}{\Delta H_{\text{PVDF-HFP}}} \quad (1)$$

, where  $\Delta H_m$  and  $\Delta H_{\text{PVDF-HFP}}$  are the melting enthalpies of the NaQSSEs and PVDF-HFP-based electrolyte. It is noted that the melting temperature,  $T_m$ , decreased to  $130.3^\circ \text{C}$  in NaQSSE<sub>C</sub> but is still high enough for high-temperature applications. The melting enthalpy of NaQSSE<sub>C</sub> is  $9.48 \text{ J g}^{-1}$  and the corresponding crystalline phase ( $X_C$ ) is 9%, lower than that of NaQSSE<sub>A</sub> (20.8%) and NaQSSE<sub>B</sub> (21%). The electrolyte's lower crystalline phase indicates a faster ion conduction process and, ultimately, an increase in ionic conductivity. Thermogravimetric analysis of NaQSSE<sub>C</sub>, compared to commercial glass fiber electrolytes soaked in the same liquid electrolyte, is shown in Fig. 1(d). The organic solvents begin to evaporate above  $70^\circ \text{C}$ , but a little weight loss is observed in the electrolytes up to  $90^\circ \text{C}$ . Compared to NaQSSE<sub>C</sub> at  $200^\circ \text{C}$ , the solvents evaporate almost completely in the GF electrolyte. The TGA curve of the pure PVDF-HFP film immersed in a commercially available liquid electrolyte is added for comparison; the mass loss due to solvent evaporation is less than that of NaQSSE<sub>C</sub> because the latter has more trapped liquid. The sudden mass loss at  $450^\circ \text{C}$  is due to the thermal decomposition of PVDF-HFP, while the gradual mass loss is due to the evaporation of bound molecules from the electrolyte. Only 20% weight loss is seen for NaQSSE<sub>C</sub>, an indication that most of the organic liquids are present in the membrane for enhanced safety. The surface morphology of the PVDF-HFP coated nonwovens is shown in Fig. 1(e). It can be observed in Fig. 1(e)-(ii) that the pore size of the fibrous matrix is too large, so it could not act as



a separating agent; instead, after being coated with PVDF-HFP, the surface becomes smooth and uniform, as shown in Fig. 1(e)-(iii). The nonwoven fibers are embedded in the polymer matrix, providing a transport channel for ion movement, as shown in the cross-sectional images in Fig. 1(e)-(iv).

Fig. 2(a) shows the conductivity obtained from the complex impedance analysis, which shows that the highest conductivity of  $1.16 \times 10^{-4} \text{ S cm}^{-1}$  was obtained for a suitable ratio of propylene carbonate/trimethyl phosphate (PC/TMP) and fluoroethylene carbonate (FEC). The electrolyte composition NaQSSE<sub>C</sub> (PC : TMP : FEC; 85 : 5 : 10), which includes more FEC than the other two compositions (NaQSSE<sub>A</sub> (PC : TMP : FEC; 85/10/5) and NaQSSE<sub>B</sub> (PC : TMP : FEC; 85 : 7 : 8)), has been observed. NaQSSE<sub>C</sub> has a composition of PC : TMP : FEC in the ratio 85 : 5 : 10. This is because the high dielectric constant of FEC (110) leads to more substantial sodium-ion solvation power in the FEC-containing electrolytes compared to the other combinations. In Fig. 2(b), the interface stability between the developed QSSE and Na metal anodes was verified using cyclic voltammetry, in the voltage range  $-0.70 \text{ V}$  to  $2 \text{ V}$  at a scan rate of  $5 \text{ mV s}^{-1}$ . A significant redox peak was observed in the NaQSSE<sub>C</sub> electrolyte shown in Fig. 2b, attributed to sodium's plating and stripping processes at about  $0 \text{ V}$  on the stainless-steel electrode. Fig. 2b shows no significant redox reactions except for the plating/stripping of sodium from  $-0.70$  to  $2 \text{ V}$ . The absence of anodic current indicates that the electrolyte is stable and compatible with the sodium anode. During the first cycle, sodium deposition occurs on a particular electrode surface area. This process leads to polarization to low potentials, as surface species and side reactions develop. In later cycles, as the deposited sodium acts as a nucleation center for further deposition, the deposition of sodium results in higher measured currents than in the first cycle. The anodic currents of the sodium stripping process decrease, resulting in decreasing CV peaks due to the side reactions that the sodium deposits inevitably experience. Anodic peaks are observed at higher

potentials than successive cycles due to the presence of surface films on the electrodes that impede sodium ions. For high-energy-density batteries, the antioxidant ability of the electrolyte is a critical parameter. As shown in Fig. 2(c), there was a  $4.77 \text{ V}$  stability window, indicating that high-voltage cathodes could be used.

The transfer number  $T_{\text{Na}^+}$  is the contribution of the  $\text{Na}^+$  ions to the conductivity of the electrolyte. Since  $\text{Na}^+$  is the charge carrier involved in the redox process of the battery,  $T_{\text{Na}^+}$  is one of the key factors in the electrolyte. To calculate the transfer number ( $T_{\text{Na}^+}$ ), an AC impedance and DC polarization technique was used, as shown in Fig. 2(d). For the NaQSSE<sub>C</sub>, it is found that the  $T_{\text{Na}^+}$  is 0.68 using eqn (2),<sup>30</sup> which is significantly higher than that of the conventional liquid electrolyte ( $\leq 0.5$ ),

$$T_{\text{Na}^+} = \frac{I_s(\Delta V - I_0 R_1^0)}{I_0(\Delta V - I_s R_1^s)} \quad (2)$$

where  $I_0$ ,  $I_s$ ,  $R_1^0$  and  $R_1^s$  are the initial currents, steady-state currents, initial resistance, and final resistance, of the electrolyte before and after polarizing, respectively;  $\Delta V$  is the applied direct current potential over the symmetric  $\text{Na} \parallel \text{NaQSSE}_C \parallel \text{Na}$  cell.

The Na plating/stripping cycle tests ( $\text{Na} \parallel \text{electrolyte} \parallel \text{Na}$ ) were carried out with NaQSSE<sub>C</sub> and commercial glass fiber (GF) fiber soaked in conventional electrolyte ( $1 \text{ M NaClO}_4$  in EC/PC) at a constant current density of  $0.1 \text{ mA cm}^{-2}$ , as shown in Fig. 2(e). Initially, the overpotential is  $\sim 5.5 \text{ mV}$ , which is increased to  $\sim 8.7 \text{ mV}$ , indicating a stable voltage profile over 300 h. Whereas, the overpotential for the GF-based separator increased from  $345 \text{ mV}$  to  $133 \text{ mV}$ . The Nyquist plot (inset) for both cells indicates a lower interfacial resistance with a value of  $465 \text{ Ohms}$  ( $\text{Na} \parallel \text{NaQSSE}_C \parallel \text{Na}$ ) compared to  $877 \text{ Ohms}$  ( $\text{Na} \parallel \text{GF} \parallel \text{Na}$ ). This, in turn, results in an increase in the initial overpotential. The results confirm that NaQSSE<sub>C</sub> has good compatibility with sodium metal. Related to the above, SEM analysis of the Na metal was carried out to understand the change in the microstructure after 300 h of plating/stripping. The microstructural change (Fig. 2(f)) reveals a dense, compact, and smooth surface morphology for the NaQSSE<sub>C</sub> electrolyte where needle morphology with commercial GF-based cells is reported elsewhere.<sup>31</sup>

We tested the applicability of the electrolyte in a half-cell configuration with an NVP cathode for long-term cycling at a high charge/discharge rate of 2C as shown in Fig. 3(a). NVP was tested as the cathode material, being able to produce an extended cycle profile with liquid/gel polymer electrolytes.<sup>31,32</sup> The cell delivers a discharge capacity of  $107 \text{ mAh g}^{-1}$  at a charge/discharge rate of C/10, whereas  $73 \text{ mAh g}^{-1}$  is obtained at a charge/discharge rate of 2C. Battery capacity is directly affected by the chemical composition and potentials of the electrodes and is proportional to the number of ions transferred. If the battery is charged with different C rates for the same voltage, the overvoltage is more significant for the faster-charged battery. Accordingly, the reduced difference in chemical potentials will result in a decrease in the number of

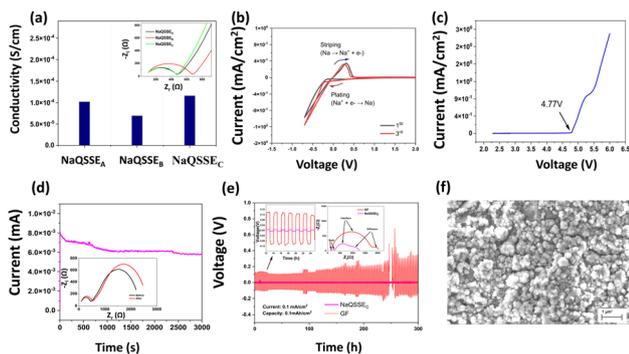
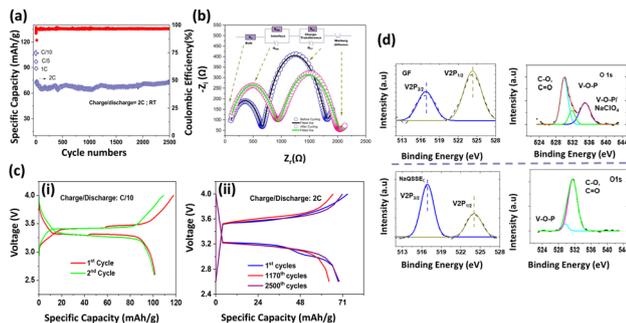


Fig. 2 (a) Conductivity variation of different NaQSSEs at RT ( $30 \text{ }^\circ\text{C}$ ); (b) cyclic voltammetry curve of the  $\text{Na} \parallel \text{NaQSSE}_C \parallel \text{SS}$  cell and (c) linear sweep voltammetry at  $30 \text{ }^\circ\text{C}$ ; (d) current–time response at RT at  $20 \text{ mV}$  dc polarization voltage and Nyquist plots for symmetric cells with NaQSSE<sub>C</sub> (inset); (e) sodium plating/stripping at a current rate of  $0.1 \text{ mA cm}^{-2}$  of the  $\text{Na} \parallel \text{Na}$  symmetric cells fabricated with the NaQSSE<sub>C</sub> electrolyte, plating/stripping time of 60 minutes; (f) SEM image of sodium metal after plating/stripping.





**Fig. 3** (a) Cycling stability of the Na||NaQSSE<sub>C</sub>||NVP cell at different current densities and long cycling at a 2C charge/discharge rate, (b) Nyquist plot of the Na||QSSE<sub>C</sub>||NVP cell before cycling and after cycling, (c) (i) charge/discharge curve at C/10, and (ii) at 2C rate in the cell configuration of Na||QSSE<sub>C</sub>||NVP at room temperature; (d) analysis of V 2p and O 1s from the XPS spectra of the NVP cathode after cycling with the NaQSSE<sub>C</sub> electrolyte and GF soaked in commercial electrolyte (1 M NaClO<sub>4</sub> in EC/PC).

ions transferred and a consequent reduction in capacity. In addition, charge transfer through the bilayers plays a significant role in resistance. High impedance will slow down the rate of ion transfer and increase the overvoltage (Fig. 3(c)) for the same charge rate. Interestingly, the capacity retention is 98.4% after 2500 cycles at a 2C charge/discharge rate Coulombic efficiency (CE) of 96.6%, indicating a negligible loss of capacity after long-term cycling. The EIS spectroscopy is a more realistic representation of the behavior of the electrolyte with the electrodes in a working battery. The bulk resistance and the charge transfer resistance of the Na||NaQSSE<sub>C</sub>||NVP cell have been measured before and after the long cycle as shown in Fig. 3(b). After cycling, the cell resistance increases from 651 Ohms to 918 Ohms. It is noted that the resistance after cycling is about 1100 Ohms, while before cycling, it is 1120 Ohms, so there is no significant change in the charge transfer resistance.

The Nyquist plot fitting parameters are listed in Table S1 (ESI<sup>†</sup>). These indicate that a stable interphase has been established between the electrodes/electrolyte.<sup>33</sup> There are no adverse reactions that could increase the interfacial resistance. The cell charge/discharge profile is shown in Fig. 3(c)-(i), (ii) for C/10 and 2C rates, respectively, in the 4.0–2.6 V range. Interestingly, the cell voltage difference is almost constant even at a 2C rate after 2500 cycles.

The V 2p and O 1s spectra of the cathode electrolyte interphase (CEI) layer are shown in Fig. 3(d). The two bands at 517.5 eV and 524.8 eV in the V 2p spectra are assigned to the V 2p<sub>3/2</sub> and V 2p<sub>1/2</sub> for the pristine NVP, respectively. The lower ratio of V 2p<sub>3/2</sub> to V 2p<sub>1/2</sub> for the NVP cathode with GF-soaked commercial electrolyte compared to the NaQSSE<sub>C</sub>-based electrolyte indicates the severe destruction of vanadium in the NVP cathode.<sup>34</sup> Thus, a stable CEI layer on the NVP cathode is formed with the NaQSSE<sub>C</sub>-based electrolyte. For the deconvoluted O 1s spectra, it is noted that C–O, C=O, and additional V–O–P/NaClO<sub>4</sub><sup>27</sup> signals are present for the commercial electrolyte-based NVP. In contrast, only C–O and C=O signals are present for the NaQSSE<sub>C</sub>-based electrolyte-based NVP cathode.

The battery's safety is paramount, especially for large-scale energy storage devices. Mainly, the polycarbonate-based electrolyte does not provide safety as it bursts into flames upon ignition. In contrast, our developed NaQSSE<sub>C</sub> is intrinsically safe as it is non-flammable upon ignition, as shown in Fig. S2 (ESI<sup>†</sup>).

In summary, a nonwoven supported quasi-solid-state electrolyte was prepared through solution casting using PVDF-HFP as the polymer matrix. The electrolyte is non-flammable with an optimized PC/TMP/FEC ratio and tested with an open flame test. The electrolyte is highly flexible and provides a conductivity of  $1.16 \times 10^{-4} \text{ S cm}^{-1}$  at RT with a high transport number of 0.68. In addition, the electrolyte is compatible with Na metal and secures an electrochemical voltage window of 4.7 V. The prepared Na||NaQSSE<sub>C</sub>||NVP cell shows exceptional cycle stability and a capacity retention of 98.4% at a high charge/discharge rate of 2C with an initial discharge capacity of  $73 \text{ mA h g}^{-1}$ . This study shows a cost-effective approach using a low-cost, nonwoven-supported polymer electrolyte. It is prepared with an optimal liquid electrolyte ratio to make it a non-flammable sodium metal battery operation.

## Author contributions

Sayan Das: conceptualization, methodology, research, data curation, formal analysis, visualization, writing – original draft preparation, writing – reviewing and editing; Venimadhav Adyam: conceptualization, resources, supervision, draft reviewing and preparation, funding acquisition; Vilas G. Pol: draft reviewing, proof reading and preparation. All authors approved the final version of the manuscript.

## Conflicts of interest

The authors declare no competing financial interest.

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