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Water-dispersed bismuth-doped strontium pyrovanadate phosphor particles with sub-degree celsius thermal resolution†

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Particle-liquid dispersion technique entails a principle analogous to that of an optically thin arrangement of particles, which is applicable in lab-on-a-chip technology, thin coatings for lighting, solar cell sensitization, medical bioimaging, and temperature measurements in fluids. $\text{Sr}_2\text{V}_2\text{O}_7\text{:Bi}^{3+}$ phosphor particles demonstrated the ability to luminesce when dispersed in distilled water. X-ray powder diffraction patterns confirmed that the phosphor crystallised in the triclinic phase. Its morphologies and chemical compositions were analysed using field emission scanning electron microscopy and energy-dispersive X-ray spectroscopy. The luminescence intensity of the phosphor was optimised via annealing at different temperatures and doping with different Bi^{3+} concentrations. The luminescence intensity of water-dispersed phosphor particles exhibited strong thermal quenching in the temperature range of 22 to 60 °C, and the band maximum showed a significant blue-shift. Relative temperature sensitivity determined from the temperature dependence of luminescence lifetime exhibited a high value around ambient temperatures, with a maximum value of 2.56%/°C at 20 °C and 2.13%/°C at 25 °C. These values were highly comparable to lifetime-based relative temperature sensitivities reported in the literature under similar temperature ranges, demonstrating this material's potential for sensitive thermometry applications around ambient temperatures.

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1. Introduction

Luminescence thermometry (LT) is a non-intrusive optical technique that applies the temperature dependence of luminescence to remotely measure the temperature of a system.^{1–7} By introducing luminescent tracers into transparent fluids or solids or coating them onto surfaces and exciting them with a laser, LT exploits spectral and temporal/spatial approaches for remote sensing with high resolution and can operate in highly luminous environments, which are crucial advantages over infrared thermography and conventional intrusive probes.^{8–11} For these reasons, LT has numerous potential applications in cooling and heating processes, such as in automotive, chemical, and industrial energy processes, e.g., for the design of reactors or heat exchangers;¹² rapidly varying processes, e.g., in fluid flows,¹³ or rotating parts;¹⁴ nanothermometry, e.g., biomedicine,^{10,15} lab-on-a-chip,¹⁶ microelectronics;¹⁷ and temperature measurements in harsh environments with

robust luminosity, e.g., on combustible materials or in internal combustion engines.¹⁸ LT is based on optically active tracers, including quantum dots, metal–ligand complexes, organic compounds (e.g., dyes), and phosphors.¹⁹ Phosphor thermometry is based on inorganic compounds doped with lanthanides (e.g., Eu^{3+} ,²⁰ Ce^{3+} ,²¹ Pr^{3+} ,^{22,23} Nd^{3+} ,^{24,25} and Dy^{3+} ,²⁶ transition metals (e.g., Mn^{3+} , Mn^{4+} ,²⁷ Cr^{3+} ,²⁸ or post-transition metals (e.g., Sn^{2+} ,²⁹ Bi^{3+}).^{30,31} In most cases, the temperature-dependence of luminescence intensity or lifetime of a phosphor is thermally insensitive over an extensive temperature range and shows a steep decrease over a narrow temperature range. It is said to be thermally sensitive in this steep region.

Each of these applications operates in different temperature ranges and therefore requires probes that are thermally sensitive in these ranges. For example, in a high-temperature application such as in combustion engines, for which there are numerous phosphors,³² a hot jet engine blade spins at ~10 000 rpm at ~1000 K (727 °C);¹⁸ this implies that a temperature probe required to measure the temperature of this system should be thermally sensitive in this range. For low-temperature applications, e.g., in live cell imaging,^{10,15,33} in which temperatures lie in the ambient range, the suitable physiological temperature, only few phosphors possess decent thermal sensitivity.³² Thus, this

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study aims to improve measurement precision in the sub-°C temperature range. This can be achieved by developing a phosphor with inherently high thermal sensitivity in the ambient temperature regions, and using changes in the lifetime and luminescence intensity driven by thermal quenching.

In fluid thermometry, phosphor particles are seeded into a liquid or gas as a tracer, with the reasoning that they absolutely assume the temperature (and velocity) of the surrounding fluid. The temperature of the particles is determined by measuring the temperature dependence of the luminescence intensity or the lifetime of the flow *via* laser excitation. For a quality result, low concentrations of phosphor particles should be used to avoid altering the properties of the fluid and multiple scattering. A low particle concentration is also essential for applications such as the production of miniature devices,³⁴ thin coatings for lighting, solar cell sensitization, temperature measurements in gases and fluids, and medical imaging.^{29,31} Among the post-transition metals, Bi³⁺-doped vanadates have shown exceptional thermal sensitivity around ambient temperatures.^{30,31} Bi³⁺ also possesses an intrinsic ability to emit in different ranges of the electromagnetic spectrum, depending on the crystal field strength of the host compound³⁵ (e.g., Y₃Al₅O₁₂:Bi³⁺ emits at 303 nm,³⁶ while ScVO₄:Bi³⁺ emits at 635 nm).³⁷ In general, Bi³⁺ can emit from the ultraviolet to near-infrared spectrum in different hosts.^{31,35–46}

In Bi³⁺-doped vanadates, the onset of thermal quenching and therefore the temperature-sensitive range can also be adjusted by choosing a different host compound. For example, Bi³⁺-doped Y, Gd, or Lu vanadates, quench at higher temperatures (~100 °C) compared to ScVO₄, which quenches around room temperature.^{31,47,48} Another vanadate that has exhibited modest luminescence quenching around room temperature is strontium pyrovanadate (Sr₂V₂O₇).⁴⁹ Sr₂V₂O₇ is a stable compound with interesting crystal chemistry and optical properties.⁵⁰ The decay time of this phosphor is in the microsecond range, which is ideal for fast-flow imaging and for time-gated imaging in applications where fluorescence is not suitable.⁵¹ The luminescence emission spectrum of the phosphor covers almost the entire visible spectrum and the near-infrared region (the first biological window), making it suitable for biological applications. Therefore, this research intends to produce a Sr₂V₂O₇:Bi³⁺ phosphor using the modified Pechini method. To maintain the charge balance in the Sr₂V₂O₇:Bi³⁺ matrix (Sr is divalent, and Bi is a trivalent ion), three Sr²⁺ were replaced by two Bi³⁺ to create Sr_{2–1.5x}Bi_{2x}V₂O₇. The luminescence intensity of the phosphor was enhanced by doping with different concentrations of Bi³⁺ and annealing at different temperatures. In our previous report,³¹ we demonstrated that ScVO₄:Bi³⁺ has exceptional thermometric properties at ambient temperatures, so the luminescence intensity and the quantum efficiency of this phosphor were compared to those of Sr₂V₂O₇:Bi³⁺. The temperature dependence of the luminescence intensity and lifetime were measured by dispersing 40 mg L^{–1} of the powder in distilled water, and exploring the potential application of this phosphor for remote thermometry using the temperature dependence of the luminescence lifetime.

2. Experimental

2.1. Synthesis

Sr_{2–1.5x}Bi_{2x}V₂O₇ ($x = 0, 0.0005, 0.001, 0.003, 0.005, \text{ and } 0.01$) phosphors were prepared by the modified Pechini method. The starting precursors are strontium nitrate (Sr(NO₃)₂, 99.9%), ammonium metavanadate (NH₄VO₃, 99.8%), bismuth(III) oxide (Bi₂O₃, 99.9%), citric acid (C₆H₈O₇, 99.5%), polyethylene glycol (H(OCH₂CH₂)_nOH, (PEG-4000)) and nitric acid. Stoichiometric amounts of Sr(NO₃)₂, NH₄VO₃, and Bi₂O₃ were dissolved in 10 ml of nitric acid in one glass beaker. C₆H₈O₇ and PEG-4000 were also dissolved in 40 ml of deionized water in another beaker. Both mixtures were then stirred in a magnetic hot plate set at 150 and 70 °C, respectively, for 30 min. The two solutions were mixed and stirred for about 4 h. Afterwards, the solution was dried in an oven at 100 °C for 24 h until a resin was formed. The molar ratio of C₆H₈O₇ to all metals is 3 : 1, and that of PEG-4000 to C₆H₈O₇ is 10 : 1. The resin was ground into powder and annealed at 1000 °C for 4 h. Subsequently, the sample doped with 0.1 mol%Bi³⁺ was annealed at 600, 800, 1100 and 1200 °C. For ScVO₄ doped with 1 mol%Bi³⁺, the same procedure was followed using Sc₂O₃ (99.9%) as the source of Sc, and the sample was annealed at 600, 800, and 1000 °C.

2.2. Characterization

A Bruker D8 Advance X-ray diffractometer was used for the structure analysis. A Jeol JSM-7800F field emission scanning electron microscope (FE-SEM) fitted with an Oxford X-MaxN 80 energy-dispersive X-ray spectroscopy (EDS) apparatus was used for the morphology and the elemental analysis. An FLS980 fluorescence spectrometer (Edinburgh Instruments) equipped with a 450 W xenon lamp as a steady-state excitation source was used to measure the photoluminescence, and a 340 nm pulsed light-emitting diode incorporated into an FLS980 fluorescence spectrometer was used for the lifetime measurements at room temperature. An integrating sphere incorporated in the FLS980 spectrometer was used for the quantum efficiency (QE) measurement. The temperature dependence of the luminescence intensity was measured with a system consisting of a 325 nm He–Cd gas laser as excitation source, a spectrometer, a photomultiplier tube detector, and a lock-in amplifier. The temperature dependence of the lifetime was measured by exciting the sample with the third harmonic (355 nm) of an Nd–YAG laser and using a Thorlabs – PDA8A2 Si fixed gain detector, 320–1000 nm, amplified photodiode connected to and oscilloscope to measure the lifetime.

3. Results and discussion

3.1. Structure

The X-ray powder diffraction (XRPD) patterns of Sr_{2–1.5x}Bi_{2x}V₂O₇ ($x = 0, 0.0005, 0.001, 0.003, 0.005, \text{ and } 0.01$) annealed at 1000 °C are shown in Fig. 1a and b shows the enlarged XRPD patterns in the diffraction angle range of 25° to 35°. The phosphors crystallized in the triclinic phase belonging to the $P\bar{1}$ space group, and matched with the Joint Committee on



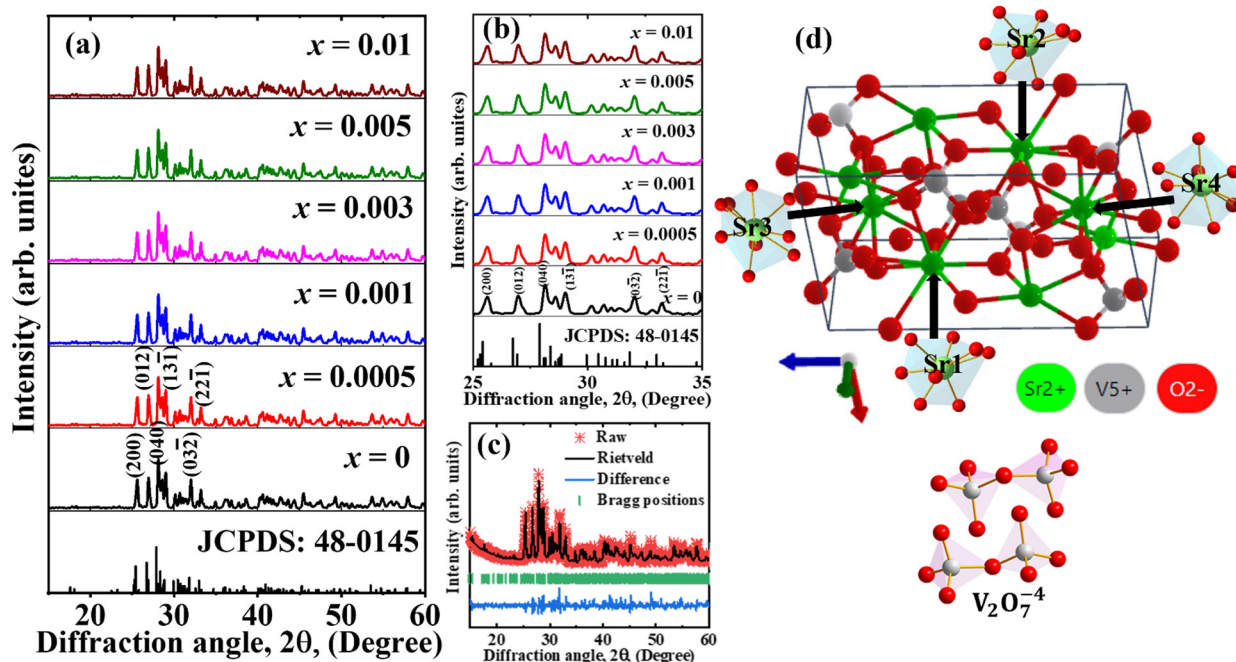


Fig. 1 (a) The XRPD patterns of $\text{Sr}_{2-1.5x}\text{Bi}_{2x}\text{V}_2\text{O}_7$ ($x = 0, 0.0005, 0.001, 0.003, 0.005$, and 0.01) annealed at 1000°C . (b) Enlarged XRPD patterns in the 2θ range of 25° to 35° . (c) The Rietveld refinement of the sample doped with $x = 0.001\text{Bi}^{3+}$. (d) The unit cell of $\text{Sr}_2\text{V}_2\text{O}_7$, showing the various Sr^{2+} and V^{5+} sites.

Powder Diffraction Standards (JCPDS) file no. 48-0145. The Rietveld refined results of the sample doped with $x = 0.001\text{Bi}^{3+}$ is shown in Fig. 1c. The Rietveld refined cell parameters ($a = 7.0571 \text{ \AA}$, $b = 7.1033$, $c = 12.9932 \text{ \AA}$, and $V = 641.25 \text{ \AA}^3$) agreed well with the standard cell parameters ($a = 7.0574$, $b = 7.1032$, $c = 13.0070 \text{ \AA}$, and $V = 641.93 \text{ \AA}^3$). Generally viewed as the “stacking of four layers of heavy atoms,” each layer in the structure of $\text{Sr}_2\text{V}_2\text{O}_7$ contains four inequivalent strontium (Sr1 , Sr2 , Sr3 and Sr4) and four inequivalent vanadium (V1 , V2 , V3 and V4) sites, giving rise to eight Sr^{2+} and eight V^{5+} sites within the unit cell (Fig. 1d).^{50,52} Eight atoms of oxygen coordinated the Sr1 and Sr2 sites with dissimilar Sr–O bond lengths, whereas 9-atoms of oxygen coordinated the Sr3 and Sr4 sites with differing Sr–O bond lengths. The four V^{5+} sites are contained in two different V_2O_7 clusters (Fig. 1d), consisting of two VO_4 tetrahedra clusters linked by an oxygen apex.⁵³ The ionic radii of Sr^{2+} in 8 and 9 coordination are 1.26 and 1.31 \AA , respectively, while that of Bi^{3+} in 8 coordination is 1.17 \AA . Considering the ionic radii, Bi^{3+} can easily substitute the Sr^{3+} sites in the host.

3.2. Morphology and elemental composition

The FE-SEM image (Fig. 2a) of the sample doped with $x = 0.1 \text{ mol\%Bi}^{3+}$ (henceforth, denoted as $x = 0.001\text{Bi}^{3+}$) contains agglomerates of spherical particles of different sizes with an average size of $0.4 \mu\text{m}$, as shown in the particle size distribution plot in Fig. 2b. The inset of Fig. 2a shows the $2.5\times$ enlarged image. The EDS spectrum taken from $x = 0.001\text{Bi}^{3+}$ (Fig. 2c) shows peaks from Sr, V, O, and Bi as expected, as well as Ir and C peaks, which emerge from the iridium used to coat the

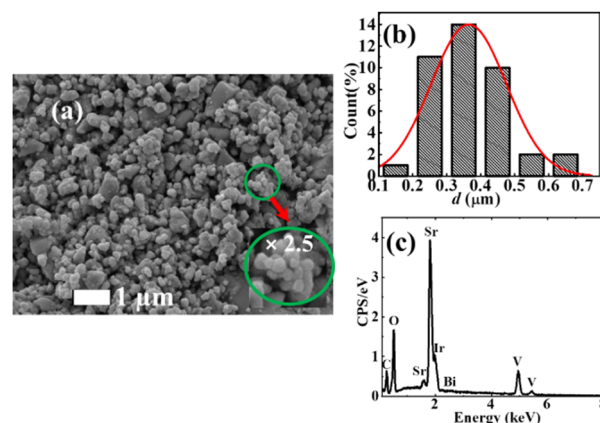


Fig. 2 (a) The FE-SEM image of $x = 0.001\text{Bi}^{3+}$ (inset of (a) is the $2.5\times$ enlargement of the FE-SEM image). (b) The particle size distribution taken from the FE-SEM image, and (c) the EDS spectrum.

powder to prevent charging and carbon tape used to mount the sample and the environmental hydrocarbon.

3.3. Optimization of the luminescence intensity by doping

Fig. 3a shows the photoluminescence (PL) excitation spectra of $\text{Sr}_{2-1.5x}\text{Bi}_{2x}\text{V}_2\text{O}_7$ ($x = 0, 0.0005, 0.001, 0.003, 0.005$, and 0.01) obtained by monitoring the emission at 550 nm . The spectra exhibited a broad band spanning from the ultraviolet to the near-visible range, which broadened toward blue for the Bi^{3+} -doped samples. The spectra consist of a peak at 265 nm and a maximum peak around 335 nm for the undoped $\text{Sr}_2\text{V}_2\text{O}_7$ ($x = 0$). The second band, observed around 335 nm for the undoped

$\text{Sr}_2\text{V}_2\text{O}_7$, shifted to ~ 340 nm and broadened towards the longer wavelength for the doped samples (Fig. 3c). The bands at 265 and 335 nm are due to the absorption from the $^1\text{A}_1$ ground state of VO_4^{3-} to its $^3\text{T}_2$ and $^3\text{T}_1$ excited states, respectively (see Fig. 4c).^{54,55} The shift in the peak maximum position and broadening observed in the doped sample is associated with two transitions; (i) the absorption of Bi^{3+} ions originating from the $^1\text{S}_0 \rightarrow ^3\text{P}_1$ (around 340 nm) transition, and (ii) the metal-to-metal charge-transfer (MMCT) absorption (the little bump around 362 nm in the doped samples) corresponding to a transition from the Bi^{3+} ground state to the host conduction band states. The MMCT state originates from the interaction between the $\text{Bi}^{3+}(6s^2)-\text{V}^{5+}(3d^0)$ configuration. The values of the MMCT transitions can be calculated using the empirical equations proposed by Philippe Boutinaud⁵⁶ presented in eqn (1) for the 4-coordinated V^{5+} :

$$\text{MMCT}(\text{Bi}^{3+}, \text{cm}^{-1}) = 70\,000 - 52\,000 \frac{\chi_4(\text{M}^{n+})}{d_{\text{coor.}}(\text{Bi}^{3+} - \text{V}^{5+})} \quad (1)$$

and in eqn (2) for the 8-coordinated Sr^{2+}

$$\text{MMCT}(\text{Bi}^{3+}, \text{cm}^{-1}) = 55\,000 - 45\,500 \frac{\chi_{\text{CN} > 4}(\text{M}^{n+})}{d_{\text{coor.}}(\text{Bi}^{3+} - \text{Sr}^{2+})} \quad (2)$$

where

$$d_{\text{coor.}} = d_{\text{host}} + \frac{1}{2}[r(\text{Bi}^{3+}) - r(\text{host})] \quad (3)$$

In eqn (1) and (2), $\chi_4(\text{M}^{n+})$ and $\chi_{\text{CN} > 4}(\text{M}^{n+})$ which in this case are $\chi_4(\text{V}^{5+})$ and $\chi_8(\text{Sr}^{2+})$, refer to the optical electronegativity of the 4-coordinated V^{5+} and 8-coordinated Sr^{2+} , respectively. In eqn (3), $d_{\text{host}} = d(\text{Bi}^{3+}-\text{V}^{5+})$ or $d(\text{Bi}^{3+}-\text{Sr}^{2+})$ are the shortest distances between $\text{Bi}^{3+}-\text{V}^{5+}$ and $\text{Bi}^{3+}-\text{Sr}^{2+}$, respectively; $r(\text{Bi}^{3+})$ and $r(\text{host})$ are the ionic radii of 8-coordinated Bi^{3+} and 8-coordinated Sr^{2+} ions or 4-coordinated V^{5+} , respectively. The values of $d(\text{Bi}^{3+}-\text{Sr}^{2+})$, 3.92 Å, and $d(\text{Bi}^{3+}-\text{V}^{5+})$, 3.28 Å, were

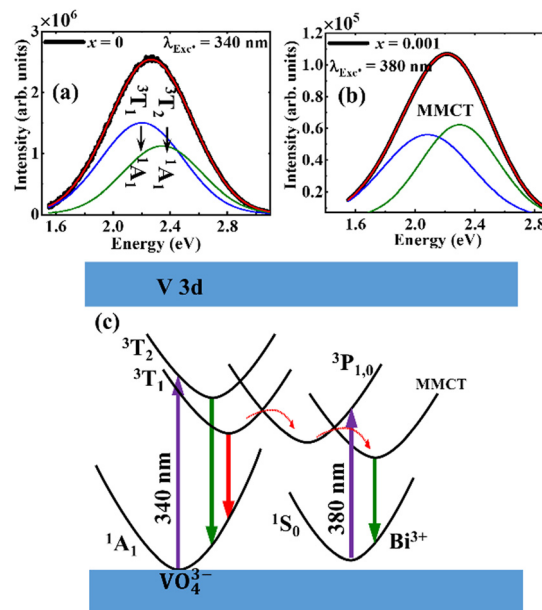


Fig. 4 Fitted PL emission spectra of (a) undoped $\text{Sr}_2\text{V}_2\text{O}_7$ and (b) $x = 0.001\text{Bi}^{3+}$ excited at 340 and 380 nm, respectively; (c) photoluminescence mechanism.

obtained by determining the distances between $\text{Sr}0-\text{Sr}5$ and $\text{Sr}7-\text{V}13$, respectively, from CIF data⁵⁷ using Diamond software.⁵⁸ It should be noted that Bi^{3+} will substitute Sr sites in the host. Therefore, the shortest Sr-Sr and Sr-V bond lengths were used for the calculation. Using the values of the parameters listed in Table 1, the value of the $\text{Bi}^{3+}-\text{Sr}^{2+}$ MMCT transition was calculated to be $41\,358\text{ cm}^{-1}$ (242 nm), and it is not observed in the spectra. This makes sense since Sr^{2+} does not have a d^0 or d^{10} configuration. The value of the $\text{Bi}^{3+}-\text{V}^{5+}$ MMCT was calculated to be $27\,934\text{ cm}^{-1}$ (358 nm), which is very close to the small shoulder observed around 360 nm in Fig. 3a, suggesting that the $\text{Bi}^{3+}-\text{V}^{5+}$ MMCT transition

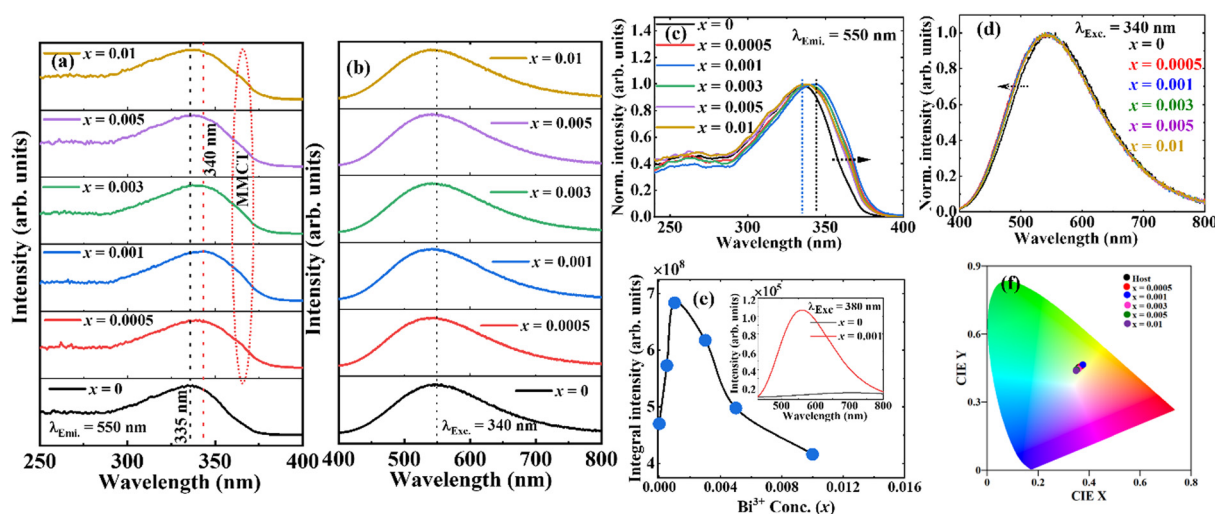


Fig. 3 (a) The PL excitation spectra measured when monitoring the emission at 550 nm, and (b) the emission spectra excited at 340 nm for $\text{Sr}_{2-1.5x}\text{Bi}_{2x}\text{V}_2\text{O}_7$ ($x = 0, 0.0005, 0.001, 0.003, 0.005$, and 0.01) phosphors; (c) normalized excitation spectra; (d) normalized emission spectra; (e) the graph of the integral emission intensity versus Bi^{3+} doping concentrations; the inset of (e) is the emission spectra of the undoped $\text{Sr}_2\text{V}_2\text{O}_7$ and $x = 0.001\text{Bi}^{3+}$ -doped sample excited at 380 nm, and (f) CIE coordinates of the phosphors. The fitted line is only used to help guide the eye.



Table 1 Values of the parameters used to calculate MMCT values

Parameter	Value	Ref.
$\chi_s(\text{Sr}^{2+})$	1.12	62
$\chi_4(\text{V}^{5+})$	2.47	62
$r(\text{Bi}^{3+})$ (Å)	1.17	63
$r(\text{Sr}^{2+})$ (Å)	1.26	63
$r(\text{V}^{5+})$ (Å)	0.36	63
$d(\text{Bi}^{3+}-\text{Sr}^{2+})$ (Å)	3.92	
$d(\text{Bi}^{3+}-\text{V}^{5+})$ (Å)	3.28	
$d_{\text{coor.}}(\text{Bi}^{3+}-\text{V}^{5+})$ (Å)	3.69	
$d_{\text{coor.}}(\text{Bi}^{3+}-\text{Sr}^{2+})$ (Å)	3.87	

could be the origin of the luminescence emission. In addition, Boutinaud and Cavalli,^{59,60} claimed that the ratio of the electro-negativity to the $\text{Bi}^{3+}-\text{V}^{5+}$ distance, $\chi_4(\text{M}^{n+})/d_{\text{coor.}}(\text{Bi}^{3+}-\text{V}^{5+})$, appears to be a decent way to predict the trends in Bi-activated phosphors. Back *et al.*⁶¹ suggested that for a $\chi_4(\text{M}^{n+})/d_{\text{coor.}}(\text{Bi}^{3+}-\text{V}^{5+})$ ratio ≤ 0.52 , the $^3\text{P}_{1,0} \rightarrow ^1\text{S}_0$ transition dominates the PL emission spectrum. For the ratio ≥ 0.53 , the MMCT is the emitting state. The calculated ratio in this material is 0.67, suggesting that the emission is dominated by the MMCT state.

When excited at 340 nm, the PL emission of the phosphors exhibited a broad spectrum covering almost the entire visible wavelength range (Fig. 3b). The band maxima of the emission spectra shifted from 550 nm for the undoped $\text{Sr}_2\text{V}_2\text{O}_7$ to 540 nm for the doped samples. The normalized emission spectra also showed that the emission bands broadened slightly to a shorter wavelength with Bi^{3+} doping (Fig. 3d). The broad band observed in the undoped $\text{Sr}_2\text{V}_2\text{O}_7$ is common in many vanadate hosts, and it comprises two bands from the VO_4 clusters emerging from the electronic transitions from $^3\text{T}_2 \rightarrow ^1\text{A}_1$ and $^3\text{T}_1 \rightarrow ^1\text{A}_1$ (see Fig. 4a).^{55,64,65} Gupta *et al.*⁶⁶ reported two emission bands at 450 and 600 nm from undoped $\text{Sr}_2\text{V}_2\text{O}_7$, which merged to a single broadband with maxima at 540 nm and a reduced bandwidth following Bi^{3+} doping. However, no significant change was observed in the luminescence intensity with Bi^{3+} doping. They assigned these effects to a reduced energy difference between the ^3T and ^1A levels of VO_4^{3-} , resulting from the distortion of the VO_4 tetrahedra structure by the Bi^{3+} ions doping. Matsushima *et al.*⁶⁵ also reported multiple emission bands from an undoped $\text{Sr}_2\text{V}_2\text{O}_7$. Other authors, on the contrary, have reported a single broad emission band with a maxima around 550 nm from undoped $\text{Sr}_2\text{V}_2\text{O}_7$.^{65,67–71} A notable difference from the above reports is the purity of the starting vanadate precursors. Although Gupta *et al.*⁶⁶ did not state the purity of the vanadate precursors, Matsushima *et al.* used 99% purity precursors, while other authors used 99.9% purity. Therefore, the reported double emission bands from $\text{Sr}_2\text{V}_2\text{O}_7$ by Gupta *et al.* and Matsushima *et al.* could originate from luminescence centers from other impurities in the material, and not VO_4 . Gupta *et al.*⁶⁶ also concluded that Bi^{3+} doping into $\text{Sr}_2\text{V}_2\text{O}_7$ caused quenching of the luminescence intensity, and attributed the emission band in the Bi^{3+} doped sample to the host transitions. However, their lowest Bi^{3+} doping concentration was 0.5 mol%, at which the luminescence intensity has already started quenching,

according to the findings herein (Fig. 3e). According to Fig. 3e, the luminescence intensity of the phosphors increased with a Bi^{3+} doping concentration up to 0.1 mol%, and started to quench upon further doping due to concentration quenching.⁵⁴ Therefore, we believe that the change in the shape of the emission spectra observed by Gupta *et al.*⁶⁶ could be due to the bands taking the shape of the Bi^{3+} emission spectrum with increasing Bi^{3+} doping concentration, which is rational considering that the band maximum (540 nm) agrees with what is reported herein.

The shift in emission wavelength (from 550 to 540 nm) observed in Fig. 3b after Bi^{3+} doping suggests that the emission from the Bi^{3+} transitions overlaps with the host emission. To further confirm this, the emission of the undoped $\text{Sr}_2\text{V}_2\text{O}_7$ and the $x = 0.001\text{Bi}^{3+}$ -doped samples were measured using 380 nm excitation (in this region, the excitation of the undoped sample is flat). From these spectra shown in the inset of Fig. 3e, the emission from the undoped $\text{Sr}_2\text{V}_2\text{O}_7$ was flat, but the doped sample exhibited a broad band with features similar to those excited at 340 nm in Fig. 3b. The ground state of Bi^{3+} is $^1\text{S}_0$, having electronic configuration $[\text{Xe}] 4f^{14} 5d^{10} 6s^2$, and the excited states in increasing order of energy are $^3\text{P}_0$, $^3\text{P}_1$, $^3\text{P}_2$ and $^1\text{P}_1$, with $6s6p$ configuration.⁵⁴ Due to the nature of the electronic configuration, the position and width of the luminescence emission band of the Bi^{3+} ions highly depend on the crystal field strength of the host matrix.^{72,73} These shifts in band positions usually observed in the Bi^{3+} -doped matrices are due to the outer 6p shell having lone pair electrons which can be influenced by its environment.⁷⁴ Thus, as the crystal field experienced by the Bi^{3+} ions in the host increases, the position of the emission band maximum shifts to a longer wavelength.^{75,76} For example, in the SrO , LuVO_4 and ScVO_4 hosts, the emission maximum of Bi^{3+} was observed at 444, 576 and 635 nm,^{43,47,77} respectively. Hence, the crystal field strength of these hosts compared to that of $\text{Sr}_2\text{V}_2\text{O}_7$ is presumably in the order: $\text{ScVO}_4 > \text{LuVO}_4 > \text{Sr}_2\text{V}_2\text{O}_7 > \text{SrO}$. The CIE coordinates of the phosphors (Fig. 4f) exhibited coordinates of warm white light without significant variation with the doping concentration.

The fitted PL emission spectra of the undoped $\text{Sr}_2\text{V}_2\text{O}_7$ (excited at 340 nm) and the $x = 0.001\text{Bi}^{3+}$ doped sample (excited at 380 nm) shown in Fig. 4a and b, respectively, consist of two bands each, belonging to the $^3\text{T}_1 \rightarrow ^1\text{A}_1$ (2.20 eV) and $^3\text{T}_2 \rightarrow ^1\text{A}_1$ (2.35) transitions of VO_4 , and the MMCT (2.09 and 2.30 eV) transitions.^{65,78} Since there are four Sr sites in $\text{Sr}_2\text{V}_2\text{O}_7$, there could be up to four bands for Bi^{3+} MMCT emission.

However, in the present case, the experimental data could be fitted well by only two bands, which could mean that Bi^{3+} either preferentially occupies only certain sites or that their MMCT emission energies are too similar to easily resolve. The MMCT energy levels lies slightly below the $^3\text{P}_{0,1} \rightarrow ^1\text{S}_0$ transitions of Bi^{3+} . The existence of two emission bands in the MMCT states has been reported by Cavalli *et al.*⁷⁸ The evidence of the existence of more than one emission band in the MMCT state can be seen in the temperature-dependent luminescence spectra in Fig. 8a. It is evident from Fig. 8a that the luminescent intensity in the long wavelength is more responsive to the



change in temperature than the short wavelength region. Another evidence of the existence of two emission bands in the MMCT state is the double exponential nature of the decay curve (see Fig. 10a and b). The energy diagram describing the luminescence mechanism of the undoped and Bi^{3+} -doped $\text{Sr}_2\text{V}_2\text{O}_7$ excited at 340 nm and the Bi^{3+} -doped $\text{Sr}_2\text{V}_2\text{O}_7$ excited at 380 nm are shown in Fig. 4c. Soon after electrons are excited from the ground state ($^1\text{A}_1$) of the VO_4^{3-} group to one of the V 3d levels ($^3\text{T}_2$) using a 340 nm lamp (this energy is not enough to promote an electron to the $^1\text{T}_2$ level), the excited electrons exhibit emission from the $^3\text{T}_1 \rightarrow ^1\text{A}_1$ and $^3\text{T}_2 \rightarrow ^1\text{A}_1$ levels of VO_4^{3-} group for the undoped sample. Meanwhile, for the Bi^{3+} -doped samples, the host emission is reabsorbed by the $^3\text{P}_{0,1}$ state of Bi^{3+} . Instead of giving out emission, the electrons from the $^3\text{P}_{0,1}$ state of Bi^{3+} undergo a nonradiative transition and move to the lower-lying $\text{Bi}^{3+}\text{-V}^{5+}$ MMCT level, leading to emission from the MMCT $\rightarrow ^1\text{S}_0$ transition. Furthermore, when the Bi^{3+} -doped sample was excited at 380 nm, electrons moved from the $^1\text{S}_0$ ground state to the $^3\text{P}_{0,1}$ excited state of Bi^{3+} , undergo nonradiative transition to the MMCT state, and emission was observed from the MMCT $\rightarrow ^1\text{S}_0$ transition.

3.4. Optimization of the luminescence intensity by annealing

The effect of annealing temperatures on the luminescence properties of the samples doped with $x = 0.001\text{Bi}^{3+}$ was studied in the temperature range of 600–1200 °C (Fig. 5). The excitation spectra (Fig. 5a) shared common characteristics with those shown in Fig. 3a and c, except for the change in the excitation peak maximum observed with the change in the annealing temperatures. At the lower annealing temperatures (600, 800, and 1000 °C), the excitation maximum shifted to the longer wavelength (328, 334, and 340 nm, respectively) with increasing

temperature. For higher annealing temperatures (1100 and 1200 °C), the excitation maximum shifted to shorter wavelengths (337 and 334 nm, respectively). This effect is due to a change in the covalency between the anion (O^{2-}) and the cations (V^{5+} and Bi^{3+} ions).^{79–82} Conventionally, the electronic interaction decreases with increasing covalency, hence decreasing the energy level of the excited state of the metal ion⁸³ – a phenomenon known as the nephelauxetic effect.⁸⁴ Thus, the initial shift towards a longer wavelength (lower energy) at low annealing temperatures (600–1000 °C) is due to an increase in the covalency, while the subsequent shift to a shorter wavelength (higher energy) at higher annealing temperatures (1100–1200 °C) implies a decrease in the covalency. Also, the crystal field strength increases with covalency and decreases with increasing coordination number.^{85–88} In ref. 79–93, it was shown that oxygen vacancies have the propensity to increase with temperature at low annealing temperatures, and subsequently decrease at higher annealing temperatures in different materials. Hence, the initial increase in oxygen vacancy at lower annealing temperatures (600–1000 °C) will decrease the coordination number, and hence increase the crystal field, leading to a decrease in the energy level of the excited state. Meanwhile, the later decrease in oxygen vacancy at higher annealing temperatures (1100–1200 °C) will increase the coordination number, causing the crystal field to decrease and increasing the energy level of the excited state. The luminescence emission spectra (Fig. 5b) have similar features to those shown in Fig. 3b, but the emission maxima shifted from lower to higher wavelengths (543, 546, and 550 nm) for the samples annealed at 600, 800 and 1000 °C, respectively, and then to lower wavelengths (548 and 545 nm) for the samples annealed at 1100 and 1200 °C, respectively. These shifts in the wavelength maximum are concurrent with those observed in the excitation spectra (Fig. 5a). The plot of the integrated emission intensity *versus* sample annealing temperature (Fig. 5c) shows that the best luminescence intensity was attained from the samples annealed at 1000 °C. The decrease in luminescence intensity at high annealing temperatures is likely to be related to the evaporation of Bi^{3+} ions from the surface of the material.⁹⁴

The luminescence of orthovanadates doped with Bi^{3+} has been widely studied. Among these orthovanadates, YVO_4 , LuVO_4 , ScVO_4 and GdVO_4 doped with Bi^{3+} have shown high quantum efficiencies (QE) with internal quantum efficiencies (IQE) of 75%, 68%, 35%³⁶ and 80%,⁴⁵ respectively. Ref. 77 reported a higher IQE of 56% for $\text{ScVO}_4\text{:Bi}^{3+}$. Generally, the onset of the thermal quenching, and hence the temperature-sensitivity range, is a function of the intrinsic properties of the host material.⁹⁵ In the orthovanadate-doped Bi^{3+} , although YVO_4 , GdVO_4 , and LuVO_4 have better QE, they quenched at higher temperatures (~ 100 °C) compared to ScVO_4 , which quenched at around room temperature.^{31,47,48}

Abram *et al.*³¹ and Kang *et al.*⁷⁷ reported the luminescence of $\text{ScVO}_4\text{:Bi}^{3+}$, and both obtained the optimum luminescence intensity at 1 mol%. To rank the QE of $\text{Sr}_{2-1.5x}\text{V}_2\text{O}_7$ doped with Bi^{3+} among the aforementioned materials, we compared the luminescence intensities of $\text{Sr}_{2-1.5x}\text{V}_2\text{O}_7$ doped with $x = 0.001\text{Bi}^{3+}$ (the

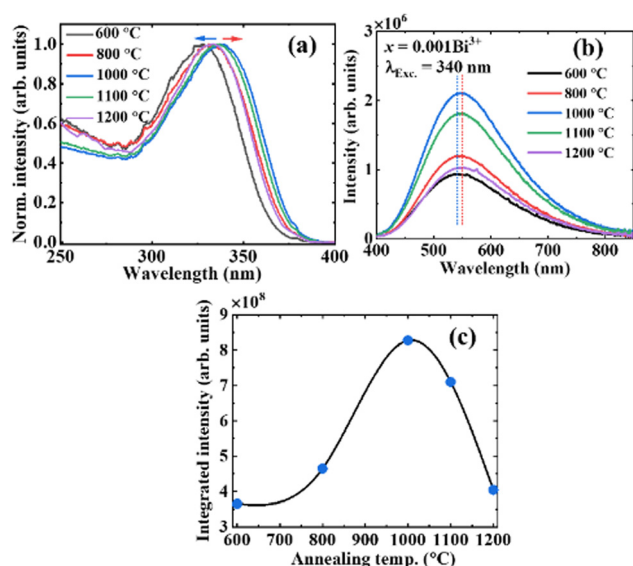


Fig. 5 The luminescence (a) normalized excitation and (b) emission spectra of the sample doped with $x = 0.001\text{Bi}^{3+}$ annealed at 600, 800, 1000, 1100 and 1200 °C, and (c) the plot of the integrated luminescence intensity *versus* annealing temperature. The fitted line is only used to help guide the eye.



spectra shown in Fig. 5) and $\text{ScVO}_4:0.01\text{Bi}^{3+}$ annealed at different temperatures. The QE of the best sample among the two materials was also compared. From the XRPD of both materials annealed at different temperatures, the FE-SEM image, particle size distribution, and EDS of $\text{ScVO}_4:0.01\text{Bi}^{3+}$ annealed at 800 °C are shown in S1a and b, and S2a–c (ESI†), respectively. The PL excitation and emission spectra of $\text{ScVO}_4:0.01\text{Bi}^{3+}$ annealed at different temperatures are shown in S3a and b (ESI†). Fig. 6 compares the luminescence intensities of $\text{Sr}_{2-1.5x}\text{V}_2\text{O}_7$ doped with $x = 0.001\text{Bi}^{3+}$ and $\text{ScVO}_4:0.01\text{Bi}^{3+}$ annealed at different temperatures. The luminescence intensity of $\text{ScVO}_4:0.01\text{Bi}^{3+}$ attained its maximum from the sample annealed at 800 °C, with the intensity slightly lower than that of $\text{Sr}_{2-1.5x}\text{V}_2\text{O}_7$ doped with $x = 0.001\text{Bi}^{3+}$ annealed at 1000 °C. The IQE of $x = 0.001\text{Bi}^{3+}$ and $\text{ScVO}_4:0.01\text{Bi}^{3+}$ annealed at 1000 and 800 °C were measured to be 14% and 12%, respectively, which ranked the two materials in a close bracket of brightness. The difference in the values of the reported QE of $\text{ScVO}_4:0.01\text{Bi}^{3+}$ reported in ref. 36 and 77 and the value measured herein could be due to the difference in the synthesis method, suggesting that solid-state synthesis is a better method in terms of brightness.

3.5. Luminescence by particle dispersion

The particle dispersion method is essential since it has a similar mechanism to the behaviour of phosphor particles in optically thin arrangements of particles, for example, in applications such as thin coatings for lighting, solar cell sensitization, temperature measurements in fluids, and medical imaging.²⁹ This method allows for the temperature control (10–70 °C) of the particle dispersion to extract the temperature dependence of the luminescence and lifetime, which are the critical parameters. For comparison, the normalized luminescence excitation and emission spectra of powder and the phosphor particle–water dispersion (40 mg L^{−1}) of the $x = 0.001\text{Bi}^{3+}$ are shown in Fig. 7a and b. The luminescence of the dispersed material and the powder material have similar features. According to ref. 29, the particle–water dispersion allows for quantitative determination of the signal per mass of phosphor particles, which enables the comparison of the absolute luminescence signal of different phosphors. For $\text{ScVO}_4:0.01\text{Bi}^{3+}$, as high as 7×10^{13} photons per mg was reported

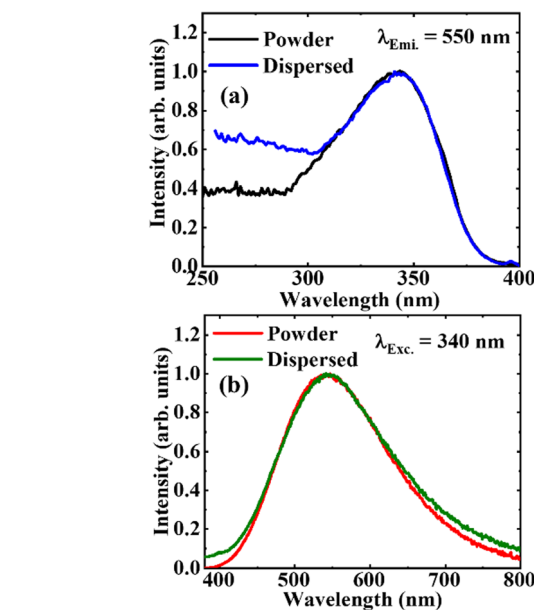


Fig. 7 The luminescence (a) excitation and (b) emission spectra of $x = 0.001\text{Bi}^{3+}$ measured from powder and phosphor particles–water dispersion.

in ref. 31, which is comparable with some of the best phosphors used in flow temperature imaging.²⁹ The small quantity of powder required for the particle–water dispersion suggests that excessively high mass concentrations are not required in experiments.

3.6. Temperature dependence of the luminescence intensity

The temperature dependence of the luminescence intensity of $\text{Sr}_{2-1.5x}\text{V}_2\text{O}_7$ doped with $x = 0.001\text{Bi}^{3+}$ was measured in the temperature range of 22 to 60 °C by dispersing 40 mg of the phosphor particles in a liter of distilled water, and using a 325 nm He–Cd laser as the excitation source (Fig. 8a). The luminescence intensity exhibited strong thermal quenching in the measured temperature range, and the position of the peak maximum displayed a blue-shift from 555 nm at 22 °C to 545 nm at 60 °C. Fig. 8b shows how the luminescence emission intensity and the peak maximum position behave with an increase in temperature. The temperature dependence of the luminescence intensity (I) was fitted using a quadratic equation given by eqn (4).

$$I = A + BT + CT^2 \quad (4)$$

where T is the temperature, and A , B and C are the fitting parameters. The peak maximum showed a linear regression with increasing temperature. The colour responses (Fig. 8c) showed a pronounced colour change, shifting from yellow towards white, confirming the blue-shift of the spectra maximum. The inset of Fig. 8c is the yellow light observed from the dispersed particles measured at 22 °C.

The thermal quenching observed in Fig. 8a resulted from increased electron–phonon interaction as the temperature was increased. The excited electrons could gain adequate phonon energy to overcome the energy barrier, leading to a non-

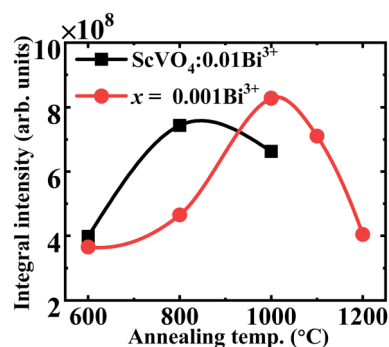


Fig. 6 The integrated luminescence intensities versus annealing temperature for $\text{Sr}_{2-1.5x}\text{V}_2\text{O}_7$ doped with $x = 0.001\text{Bi}^{3+}$ and $\text{ScVO}_4:0.01\text{Bi}^{3+}$. The fitted line is only used to help guide the eye.



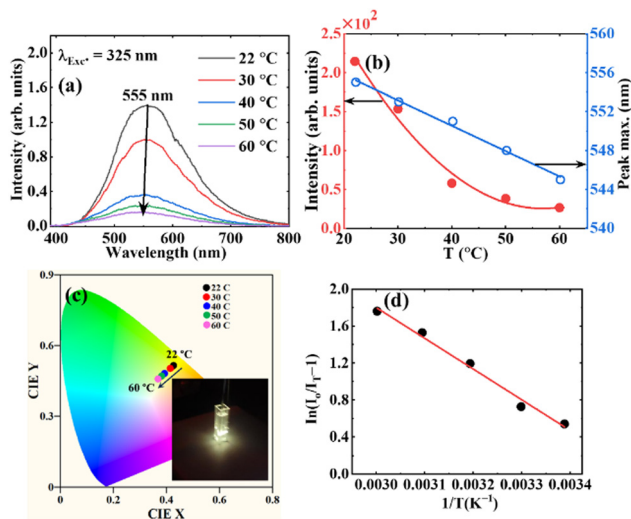


Fig. 8 (a) Temperature dependence of the luminescence intensity of water-dispersed $\text{Sr}_2\text{V}_2\text{O}_7$ doped with $x = 0.001\text{Bi}^{3+}$; (b) luminescence intensity and peak position versus temperature; (c) CIE coordinates showing the colour response of the phosphor; inset of (c) is an image of the dispersed particles under 325 nm He–Cd laser as the excitation, and (d) $\ln(I_0/I_T - 1)$ versus inverse temperature.

radiative transition, and hence thermal quenching. This energy barrier is the activation energy (ΔE). The activation energy can be calculated using the Arrhenius equation,⁹⁶ eqn (5):

$$\frac{-\Delta E}{k_B} \cdot \frac{1}{T} = \ln\left(\frac{I_0}{I_T} - 1\right) \quad (5)$$

where k_B is the Boltzmann constant ($8.629 \times 10^{-5} \text{ eV K}^{-1}$), I_0 is the initial emission intensity (in this case, at $T = 22^\circ\text{C}$) and I_T is the emission intensities at higher temperatures. From eqn (5), ΔE is inversely related to the T ; hence, the bigger the activation energy, the weaker the thermal quenching. The value of the ΔE was determined by plotting the $\ln(I_0/I_T - 1)$ versus $1/T$, taking the best fit (Fig. 8d) and equating $\Delta E/k_B$ to the slope of the graph. The value of the ΔE is about 0.28 eV, which is similar to the value reported by Li *et al.*⁴⁹

3.7. Temperature dependence of the lifetime

The temperature dependence of the lifetime of the $x = 0.001\text{Bi}^{3+}$ sample dispersed in distilled water was measured between 20°C and 70°C using the third harmonic of a pulsed Nd:YAG laser (355 nm, 8 ns, 30 Hz), laser fluence of 10 mJ cm^{-2} , an amplified photodetector, and oscilloscope (Fig. 9). A 550 nm long pass filter was used to block the 355 and 532 nm lines of the laser from entering the photodetector.

Fig. 10a shows the decay curves of the dispersed phosphors. For the sake of comparison, the lifetime of the same sample was measured at room temperature (shown in Fig. 10b), by monitoring the emission at 550 nm, and excited using an EPLED 340 nm incorporated in the FLS980 fluorescence spectrometer. The decay curves were fitted using a double-exponential function given by eqn (6):

$$I(t) = A \exp\left(-\frac{t}{\tau_1}\right) + B \exp\left(-\frac{t}{\tau_2}\right) \quad (6)$$

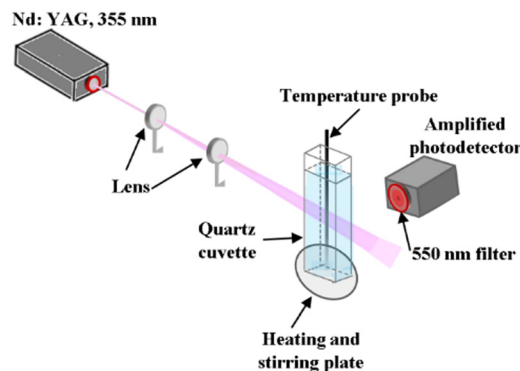


Fig. 9 Experimental set-up for the temperature-dependence of lifetime measurement.

where $I(t)$ is the luminescence intensity at time ' t ', and τ_1 and τ_2 are the first and second-order lifetime components corresponding to decay constants A and B , respectively.⁹⁷ The double-exponential decay curve suggests the presence of at least two types of decay pathways in the phosphor,⁹⁸ which is evident in the presence of two MMCT emission bands, as shown in Fig. 4b and c. The average decay time τ for a double exponential curve, defined by eqn (7),

$$\tau = \frac{A\tau_1^2 + B\tau_2^2}{A\tau_1 + B\tau_2} \quad (7)$$

was calculated to be 4.6 and $5.6 \mu\text{s}$ for the dispersed phosphor measure at 20°C and the powder sample measured using a fluorescence spectrometer, respectively. The lifetime obtained using fluorescence spectroscopy is $1 \mu\text{s}$ longer, presumably because it was measured at a fixed emission wavelength (550 nm). Meanwhile, for the dispersed particles, the lifetime was measured for a long spectral range (551–800 nm) (the spectral region displayed in blush red in the inset of Fig. 10b). The lifetime versus temperature plot was fitted using a cubic polynomial function given by eqn (8):

$$\tau = A + BT + CT^2 + DT^3 \quad (8)$$

where A , B , C and D are fitting parameters of the polynomial, and T is the temperature, and the plot is shown in Fig. 10c. It is apparent from Fig. 10c that the lifetime strongly decreases in the 20 – 40°C temperature range, showing a tenuous decrease with temperature in the 45 – 60°C range and a steep drop at the lower tail of the graph. The lifetime and the temperature maintain a similar relationship with the luminescence intensity, defined by eqn (9), from which the ΔE can also be determined.^{99,100}

$$\frac{-\Delta E}{k_B} \cdot \frac{1}{T} = \ln\left(\frac{\tau_0}{\tau_T} - 1\right) \quad (9)$$

Here, τ_0 is the initial lifetime (at $T = 20^\circ\text{C}$ in this case), and τ_T are the lifetimes at the subsequent temperatures. From the best fit of the graph of $\ln(\tau_0/\tau_T - 1)$ vs. $1/T$ (Fig. 10d), the ΔE was calculated to be 0.24 eV, which is in good agreement with the value obtained from the luminescence data.



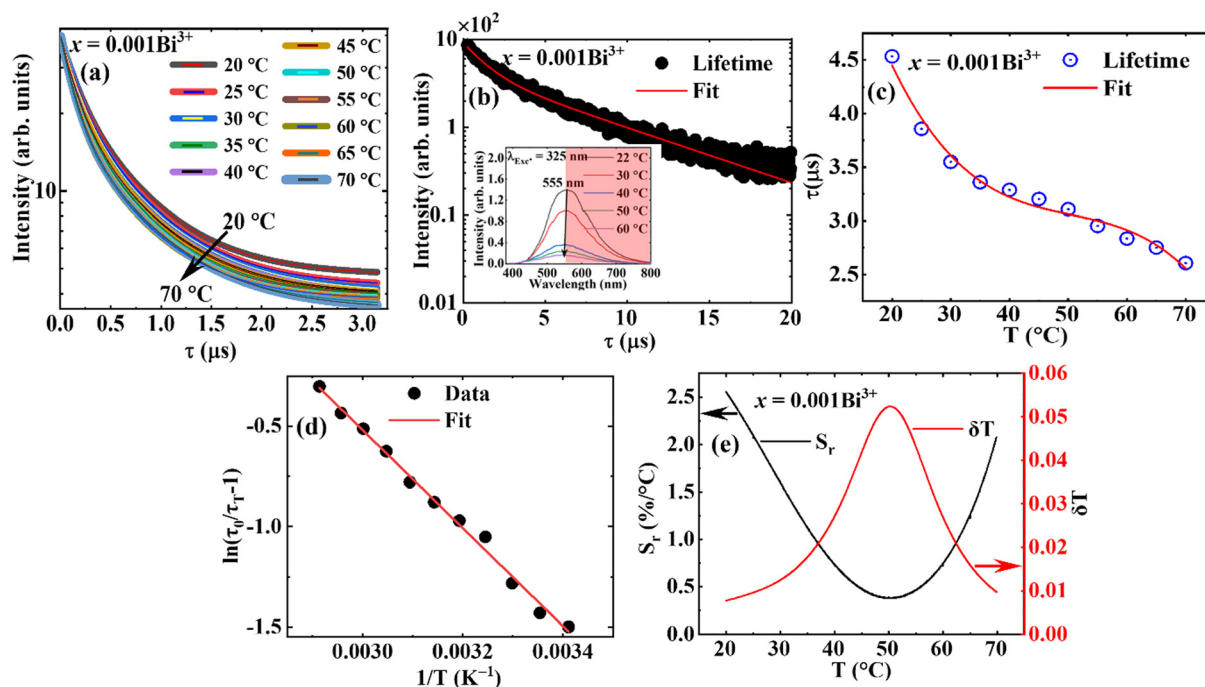


Fig. 10 (a) The temperature-dependence decay curve of the dispersed particles; (b) the decay curve of the powder sample measured at room temperature; the inset of (b) shows the spectral region (displayed in bluish red) the temperature dependence of the lifetime was measured; (c) the plot of the lifetime of the dispersed particles *versus* temperature, (d) the plot of $\ln(\tau_0/\tau_T - 1)$ *versus* $1/T$, and (e) the relative sensitivity and temperature resolution.

Table 2 Lifetime-based relative sensitivities of different phosphors and their lifetime at 25 °C

Phosphor	S_r , max%/°C	T (S_r , max) °C	S_r @25 °C	Tau@25 °C	$\lambda_{Exc.}$ (nm)	$\lambda_{Emi.}$ (nm)	Ref.
Ba ₃ LuGa ₂ O _{7.5} :Bi ³⁺	2.21	50	1.64	335.43 (ns)	351	500	100
Ca ₄ (PO ₄) ₂ O:Eu ²⁺ , Eu ³⁺	2.07	20	2.00	0.80 (μs)	355	700	102
InTaO ₄ :Cr ³⁺	2.27	42	1.42	4.58 (μs)	500	880	103
Sr ₄ Al ₁₄ O ₂₅ :Mn ⁴⁺	1.50	147	0.21	1.55 (ms)	400	651	104
Nd _{0.5} Y _{0.4} ³⁺ Yb _{0.1} PO ₄	1.22	82	1.14	0.11 (ms)	808	980	105
Nd _{0.5} Lu _{0.4} ³⁺ Yb _{0.1} PO ₄	0.85	127	0.71	0.12 (ms)	808	980	105
Nd _{0.5} La _{0.4} ³⁺ Yb _{0.1} PO ₄	1.05	147	0.84	0.16 (ms)	808	980	105
Nd _{0.5} Gd _{0.4} ³⁺ Yb _{0.1} PO ₄	0.74	172	0.45	0.09 (ms)	808	980	105
Zn _{0.9} Mn _{0.1} Al ₂ O ₄ :Cr ³⁺	0.80	−13	0.78	0.04 (ms)	427	512	106
ZnGa ₂ O ₄ :Cr _{0.5%} ³⁺ , Bi _{0.5%} ³⁺ @SiO ₂	1.93	200	0.41	2.00 (ms)	430	692	107
SrZrO ₃ :Eu ³⁺	0.30	187	0.01	0.45 (ms)	237	616	108
ZnGa ₂ O ₄ :Cr ³⁺	0.58	147	0.20	4.17 (ms)	406	684	109
LiGa ₅ O ₈ :Cr ³⁺	0.76	22	0.74	2.50 (ms)	406	684	109
Ga ₂ O ₃ :Cr ³⁺	0.46	97	0.35	1.11 (ms)	406	684	109
La ₂ MgGeO ₆ :Bi ³⁺ , Eu ³⁺	4.09	110	1.12	0.33 (μs)	333	402	110
Ba ₂ LaNbO ₆ :Mn ⁴⁺	2.77	134	0.25@27 °C	0.22 (ms)	332	681	111
Ca ₂ SiO ₄ :Mn ²⁺	4.25	102	1.55	16.00 (ms)	365	550	112
Mg ₂ TiO ₄ :Cr ³⁺	0.75	157	0.41	1.38 (ms)	473	712	113
Bi ₂ Ga ₄ O ₉ :Cr ³⁺	1.60	450	0.42	20 (μs)	355	775	114
ScVO ₄ :1%Bi ³⁺	2.83	49	2.52	2.12 (μs)	355	600	115
YVO ₄ :1%Bi ³⁺	2.34	191	0.00	3.92 (μs)	355	545	115
Sr ₂ V ₂ O ₇ :0.001Bi ³⁺	2.56	20	2.13	3.85 (μs)	355	540	This work

The relative sensitivity S_r can be calculated from the temperature dependence of the lifetime data using eqn (10):¹⁰¹

$$S_r = \left| \frac{1}{\tau(T)} \frac{d\tau(T)}{dT} \right| \times 100\% \quad (10)$$

where $\tau(T)$ is the luminescence lifetime of the phosphor at temperature T , and $d\tau(T)/dT$ is the temperature derivative of $\tau(T)$.

From Fig. 10e, the S_r exhibited a parabolic form with a maximum value of 2.56%/°C at 20 °C, and decreased with an increase in temperature to the lowest value of 0.34%/°C at 50 °C, and started to increase. Table 2 compares the maximum value of the lifetime-based S_r (S_r max), the S_r value at 25 °C, and the lifetime at 25 °C of $x = 0.001\text{Bi}^{3+}$ to different phosphors in the literature. The S_r max of Sr_{2−1.5x}V₂O₇ doped $x = 0.001\text{Bi}^{3+}$ compared very well with the best-reported values, except for La₂MgGeO₆:Bi³⁺, Eu³⁺. However, in the temperature region of

interest ($\sim 25^\circ\text{C}$), the S_r value is among the best. Materials with a very long decay time, $>100\ \mu\text{s}$, or a very short decay time, $<10\ \text{ns}$, are either very slow or very fast for imaging applications. The decay time of the $\text{Sr}_{2-1.5x}\text{V}_2\text{O}_7$ doped with $x = 0.001\text{Bi}^{3+}$ is in the microsecond range, which can be detected using cameras with a short interframe time, making it suitable for rapid imaging such as in liquid and gas flows, where phosphors can be used as tracer particles.^{31,51}

Since the relative temperature sensitivity is continuous at a particular temperature, it yields thermometer with a continuous temperature resolution, δT , which is a crucial performance criterion in the performance of a thermometer.¹¹⁶ The temperature resolution or uncertainty defines the modest temperature variation that produces an obvious change in the measured parameter,¹¹⁷ and it is given by eqn (11).¹¹⁶

$$\delta T = \frac{1}{S_r} \frac{\delta \tau}{\tau} \quad (11)$$

$\delta \tau / \tau$ is the relative standard error in determining the lifetime,^{105,118,119} and was determined to be 0.020% in our previous study.¹⁰² The best temperature resolution (0.008 $^\circ\text{C}$, Fig. 10d) was obtained at 20 $^\circ\text{C}$ and became poorer with the increase in temperature up to 50 $^\circ\text{C}$, and started to improve again.

4. Conclusions

$\text{Sr}_2\text{V}_2\text{O}_7$ doped Bi^{3+} phosphors were synthesised by the modified Pechini method. The optimal emission intensity was obtained from the sample doped with 0.1 mol% Bi^{3+} annealed at 1000 $^\circ\text{C}$. Different annealing temperatures did not cause significant changes in the structure, but the luminescence intensity varied with annealing, with the optimal emission intensity obtained from the sample annealed at 1000 $^\circ\text{C}$. The luminescence intensity and quantum efficiency (QE) of this phosphor are slightly better than those of $\text{ScVO}_4\text{:Bi}^{3+}$. However, the value of the QE of $\text{ScVO}_4\text{:Bi}^{3+}$ measured herein is below the reported values of the same materials synthesised by solid-state method,^{36,77} suggesting that solid-state is a better synthesis method.

The phosphor particles dispersed in distilled water demonstrated high relative sensitivity around ambient temperatures. The temperature-dependent luminescence intensity of the dispersed particles showed strong quenching in the 22 to 60 $^\circ\text{C}$ temperature range and blue-shifted from 555 to 545 nm. The temperature-dependence of the lifetime of the dispersed particles demonstrated high relative temperature sensitivity, with a maximum value of 2.56%/ $^\circ\text{C}$ at 20 $^\circ\text{C}$ and a value of 2.13%/ $^\circ\text{C}$ at 25 $^\circ\text{C}$, and corresponding temperature resolutions of 0.008 and 0.009 $^\circ\text{C}$, respectively. This value is among the best values reported for lifetime-based relative temperature sensitivities in these temperature ranges with a lifetime of a few microseconds.

Author contributions

Hendrik C. Swart: funding acquisition, supervision, resources and writing – review & editing. Robin E. Kroon: supervision, resources and writing – review & editing. Lucas J. B. Erasmus:

writing – review & editing. Simon N. Ogugua: writing – original draft, conceptualization, investigation, methodology and validation, visualization.

Data availability

The authors confirm that the data supporting the findings of this study are available within the article and its ESI.†

Conflicts of interest

There are no conflicts to declare.

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