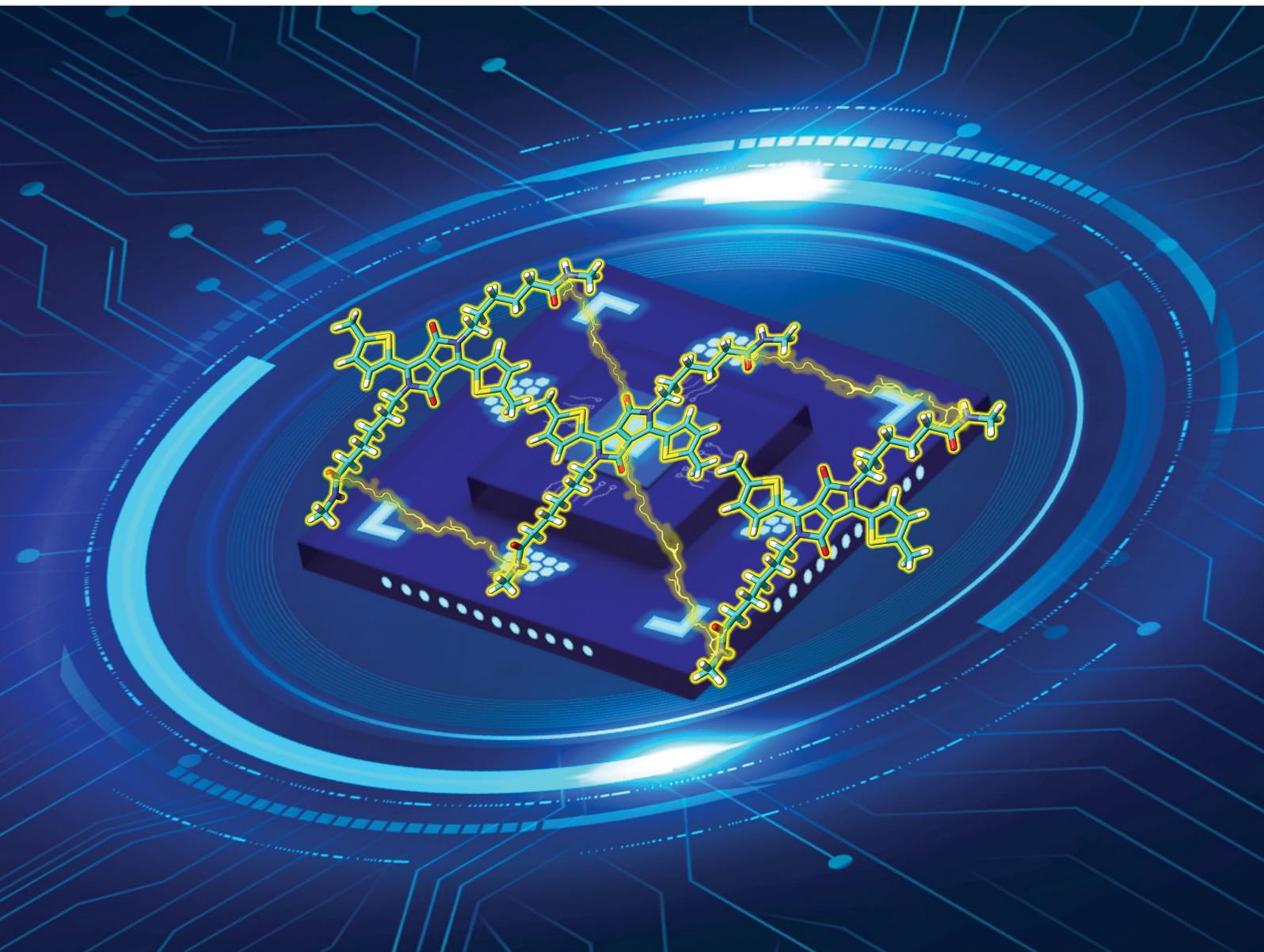


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**PAPER**

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# The effect of hydrogen bond interactions on the electronic nature of DPP-based organic semiconductors: implications on charge transport†

Raúl González-Núñez,<sup>‡</sup> Gabriel Martínez,<sup>‡</sup> Nelson Ricardo Avila-Rovelo,<sup>§</sup> Kyeog-Im Hong,<sup>¶</sup> Amparo Ruiz-Carretero,<sup>\*</sup> and Rocío Ponce Ortiz,<sup>\*a</sup>

In this work, we show that hydrogen bond interactions, in addition to directing supramolecular order, can have an intriguing effect on the molecular and electronic properties of DPP derivatives, facilitating redox processes in the solid state. As a result, charge carrier formation can be enhanced in devices, contributing to the improvement of electrical performances of two exceptionally small molecular semiconductors based on a thiophene-capped diketopyrrolopyrrole (DPP) as the electroactive component and amide groups as hydrogen-bonding units. Two different topologies of the amide groups were explored, C-centered and N-centered (**HDPPBA-C** and **HDPPBA-N**, respectively), and the results were compared to the **HDPPH** molecule with no amide groups. Spectroelectrochemical experiments, in conjunction with vibrational spectroscopies and DFT calculations, demonstrate that hydrogen bond formation promotes modifications on the effective conjugation length of the studied semiconductors, which may facilitate the generation of free charge carriers in organic field effect transistors (OFETs). These results open a new strategy towards the simple design of organic semiconductors and control over their morphological and electrical properties by simple hydrogen bond formation.

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## Introduction

Organic materials have appeared in the last few decades as great alternatives for the fabrication of electronic devices, such as organic field effect transistors (OFETs)<sup>1</sup> or organic solar cells.<sup>2</sup> The versatility of organic synthesis allows us to provide small molecules or polymers with the desired optoelectronic properties.<sup>3</sup> Interestingly, the final efficiency of a given device does not only depend on the molecular design but also, very importantly, on the way the materials assemble.<sup>4</sup> It is here that

noncovalent interactions play a crucial role, since they guide the assembly process in solution and when depositing molecules on thin films. In the case of organic semiconductors, since they are based on conjugated segments and aromatic moieties,  $\pi$  stacking is the main non-covalent interaction guiding their molecular packing. Nevertheless, the interplay of  $\pi$  stacking with other non-covalent interactions, such as hydrogen-bonding, has been proven a promising strategy to improve the morphology of organic electronic devices by creating well-connected domains of semiconductors able to transport charges efficiently.<sup>5–12</sup>

Despite the exciting results, where the beneficial role of hydrogen bonding is highlighted, the incorporation of such moieties in organic semiconductors for the fabrication of different organic electronic devices is still underexplored. In addition, in the search of efficient organic semiconductors reaching the performance required for device commercialization, a great effort has been mostly devoted over the years to the optimization of structure–property relationships of pristine organic semiconducting materials, looking for a proper balance between molecular and intermolecular characteristics for efficient charge transport. However, in the last few years, enormous interest has emerged in the use of molecular dopants, which allows tuning the energy gap and the optical and electrical properties of semiconducting materials.<sup>13</sup> Thus, this

<sup>a</sup> Department of Physical Chemistry, University of Málaga, Campus de Teatinos s/n, Málaga 29071, Spain. E-mail: rocioponce@uma.es

<sup>b</sup> University of Strasbourg, Institute Charles Sadron, CNRS, UPR22, 23 Rue du Loess, 67034, Strasbourg Cedex 2, France. E-mail: amparo.ruiz@ics-cnrs.unistra.fr

<sup>c</sup> Instituto de Ciencia de Materiales de Madrid, Consejo Superior de Investigaciones Científicas CSIC, Sor Juana Inés de la Cruz 3, 28049, Madrid, Spain

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‡ R.G.-N and G.M. contributed equally to the manuscript.

§ Current affiliation: Institute for Chemistry and Processes for Energy, Environment and Health, ICPEES, UMR 7515, 25 Rue Becquerel, 67087, Strasbourg Cedex 2, France

¶ Current affiliation: Instituto Madrileño de Estudios Avanzados, IMDEA Nanoscience, Calle Faraday 9, 28049, Madrid, Spain





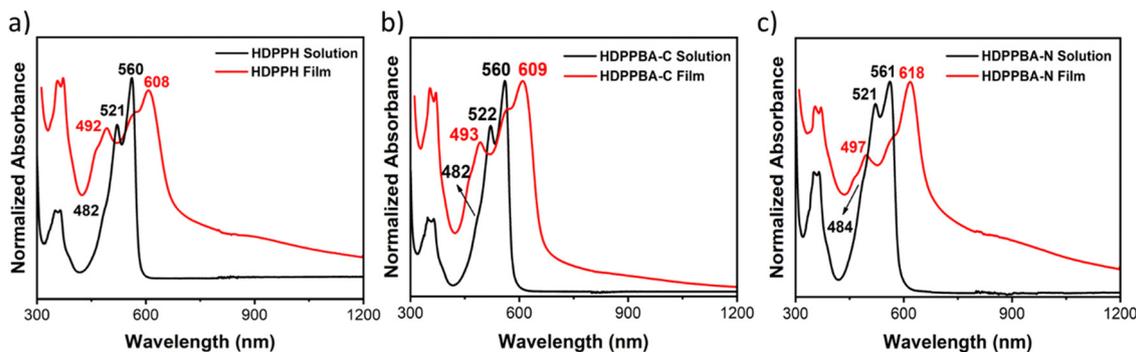


Fig. 1 UV-Vis absorption spectra in chloroform solutions (black) and thin films (red) of (a) HDPPH, (b) HDPPBA-C and (c) HDPPBA-N.

While small Stokes shifts, around 11–15 nm, are recorded in solution, large Stokes shifts are found in the solid state, particularly for the molecules having hydrogen bond interactions (141 nm and 134 nm for HDPPBA-C and HDPPBA-N, respectively *versus* 110 nm for HDPPH). These large Stokes shifts upon hydrogen bond interactions have been previously reported.<sup>41</sup>

The redox properties of HDPPH, HDPPBA-C and HDPPBA-N in solution were experimentally evaluated by cyclic voltammetry (CV) using 1 mM dichloromethane ( $\text{CH}_2\text{Cl}_2$ ) for the three derivatives, which undergo similar reduction and oxidation processes. The compounds exhibit two one-electron oxidation peaks reversible in nature at  $E_{1/2}$  0.36, 0.38 V and 0.41 V and  $E_{1/2}$  0.72, 0.71 and 0.77 V for HDPPH, HDPPBA-C and HDPPBA-N, respectively (Fig. S6, ESI<sup>†</sup>). This observation suggests that the three DPP-based compounds are electron-rich, as they demonstrate the ability to donate two electrons. The oxidation processes are accompanied by a one-electron reduction at  $E_{1/2}$  -1.73 V (C-centered) or  $E_{pa}$  -1.82 and -1.74 V (control and N-centered), which presents reversibility shown by the returning wave in the case of HDPPBA-C, but irreversible for HDPPH and HDPPBA-N. The HOMO and LUMO energy levels were

experimentally calculated from the potential at the onset of oxidation and reduction, respectively, using the frontier orbitals of ferrocene (-4.8 eV below the vacuum level) as the reference. The HOMO levels of HDPPH, HDPPBA-C and HDPPBA-N were calculated to be -5.08, -5.08 and -5.16 eV, respectively. On the other hand, the obtained LUMO values were -3.50, -3.49 and -3.50 eV, respectively. Accordingly, all three derivatives have similar band gaps ranging from 1.59 to 1.64 eV. The minor distinction in the electronic nature of the materials is reasonable, considering that the electron richness of the alkyl (in thiophenes) and amide chains (in lactams) is similar.

CV experiments on thin films were carried out in acetonitrile containing 0.2 M TBAPF6 as a supporting electrolyte (Fig. 2). Similar oxidation and reduction processes are recorded for the three molecules. However, it should be noted that, for both HDPPBA-C and HDPPBA-N compounds, a shoulder is detected from which the HOMO energy levels are calculated (explanation of the origin of the shoulder will be shown below). Slightly destabilized energy values are found for HDPPBA-C (-5.13 eV) and HDPPBA-N (-5.16 eV) compared to HDPPH (-5.21 eV). On the other hand, for LUMO energy levels, we found the same

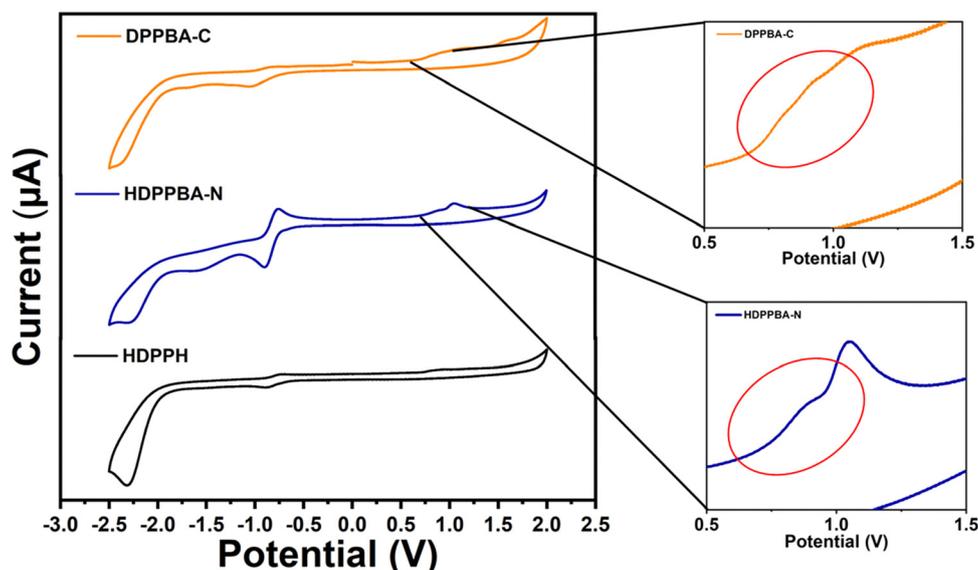


Fig. 2 Thin film cyclic voltammetry spectra recorded in acetonitrile/TBAPF6 (0.2 M) for HDPPH, HDPPBA-C and HDPPBA-N.



tendency, slightly destabilized energy values for **HDPPBA-C** (−3.63 eV) and **HDPPBA-N** (−3.59) compared to **HDPPH** (−3.68 eV).

Thin films of **HDPPH**, **HDPPBA-C** and **HDPPBA-N** were deposited by drop casting solutions of the semiconducting materials in chloroform on Si/SiO<sub>2</sub> substrates. After thermal annealing at 120° for 3 hours, their morphologies were studied by atomic force microscopy (AFM). Clear differences were found comparing **HDPPH** with the amide derivatives. While the control molecule forms large crystalline structures (Fig. 3a), **HDPPBA-C** and **HDPPBA-N** display a morphology composed of smaller fibers (Fig. 3b and c). This is ascribed to the formation, in **HDPPBA-N** and **HDPPBA-C**, of J-type aggregates in solution and in thin films formed by hydrogen bonding while the control molecule only assembles thanks to  $\pi$ - $\pi$  stacking.<sup>40</sup>

Analysing the AFM data, we found a decrease of the average roughness for **HDPPBA-N** (104 nm) and **HDPPBA-C** (45 nm), in accordance with the formation of small fibres in these compounds, with respect to **HDPPH** (169 nm). It should be noted that, in the area analysed, the maximum peak height is found in **HDPPH** (931 nm), which forms a large crystalline structure, while the values of maximum peak heights, for **HDPPBA-C** and **HDPPBA-N**, are 169 nm and 532 nm, respectively.

XRD measurements were also carried out on the thin films (Fig. S7, ESI<sup>†</sup>), indicating a more crystalline structure for **HDPPH** compared to **HDPPBA-C** and **HDPPBA-N**. Even though crystallinity is desirable to achieve efficient charge transport, **HDPPBA-C** and **HDPPBA-N** provide interconnected amorphous fibrillar networks, while **HDPPH** shows large crystalline structures

with sharp boundaries among crystallites. We believe that these two morphological scenarios are crucial to compare their charge transport properties since the interconnection among amorphous aggregates has been proven beneficial in polymeric semiconductors<sup>42</sup> and other hydrogen-bonded oligomers,<sup>5</sup> but in this case, we are pushing the limits by miniaturizing the conjugated skeleton of the derivatives explored.

### Electrical properties of thin film transistors

The photoconductivity properties of **HDPPBA-N**, **HDPPBA-C** and **HDPPH** measured at the nanoscale by electrodeless techniques based on microwave conductivity have been already reported by us. However, their charge mobility in full devices was not measured, having only a previous screening indicating that the hydrogen-bonded derivatives had superior values of photoconductivity, and charge carrier lifetimes are one order of magnitude larger than the control derivative.<sup>40</sup>

In this work, in order to measure the electrical characteristics of the studied compounds, solution processed OFETs were fabricated in a top-contact/bottom-gate structure (see the detailed procedure in the ESI<sup>†</sup>). The charge carrier mobilities ( $\mu$ ), threshold voltage ( $V_T$ ) and current on/off ratio ( $I_{ON}/I_{OFF}$ ) were extracted from the saturated region in the device transfer curves and the electrical parameters of the best performing devices are summarized in Table 1. Complete device characterization can be found in the ESI<sup>†</sup>. Compounds **HDPPBA-C** and **HDPPBA-N** show similar maximum p-type field effect mobilities, with values of  $2 \times 10^{-2} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$  and  $1 \times 10^{-2} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ , respectively (see output and transfer plots in Fig. 4). In contrast,

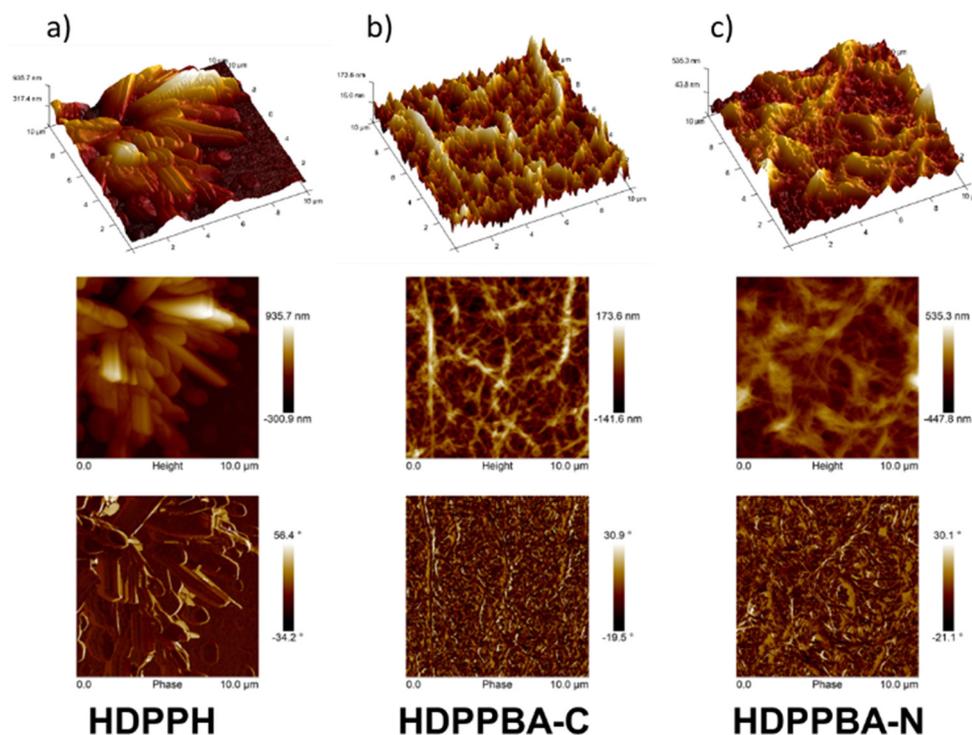


Fig. 3 AFM images of semiconducting thin films deposited on substrates at room temperature with a posterior thermal treatment at 120° for 3 hours: (a) **HDPPH**, (b) **HDPPBA-C** and (c) **HDPPBA-N**.



**Table 1** OFET electrical data for best-performing deposited films of the indicated semiconductors measured in a vacuum

Compound	Subst. treatment <sup>a</sup>	$\mu_{\text{h}}$ (cm <sup>2</sup> V <sup>-1</sup> s <sup>-1</sup> )	$V_{\text{T}}$ (V)	$I_{\text{ON}}/I_{\text{OFF}}$
<b>HDPPH</b>	OTS (120)	NA	NA	NA
<b>HDPPBA-C</b>	OTS (120)	$2 \times 10^{-2}$	-12	$2 \times 10^4$
<b>HDPPBA-N</b>	OTS (120)	$1 \times 10^{-2}$	-3	$4 \times 10^5$

<sup>a</sup> Substrate treated with octadecyltrichlorosilane. The thermal annealing temperature is shown in parentheses.

the **HDPPH** compound is not active under any tested conditions (see details in the ESI<sup>†</sup>), which indicates that the introduction of amide groups in these molecules promotes and favors charge transport.

In order to understand the extent and reasons behind this improvement, which may not be solely related to morphological and crystallinity changes due to hydrogen bond formation, a physical chemistry investigation including vibrational spectroscopy and spectroelectrochemistry was performed, as shown in the following sections.

### Analysis of electronic properties by spectroscopy.

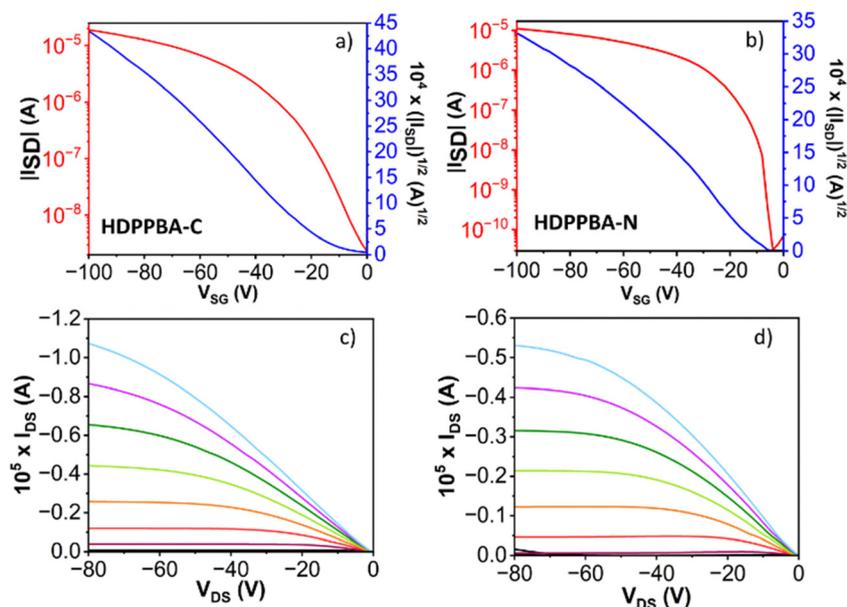
Attenuated total reflectance (ATR) IR spectra of **HDPPBA-N**, **HDPPBA-C** and **HDPPH** semiconductors as powder samples were recorded to account for the hydrogen bond formation and its implications on the electronic structure of the conjugated skeletons (Fig. 5, left). It should be noted that while a single  $\nu(\text{C}=\text{O})$  vibration is found in **HDPPH** at 1656 cm<sup>-1</sup> arising from the cross-conjugated carbonyl group of the DPP lactam group, in the case of **HDPPBA-N** and **HDPPBA-C**, as expected, two distinct bands appear at around 1657 and 1639–1642 cm<sup>-1</sup>. The IR band recorded at 1657 cm<sup>-1</sup> is associated with the cross-conjugated lactam ring while the one recorded at 1639–1642 cm<sup>-1</sup>

is due to the alkyl chain amide groups forming hydrogen bonds. As expected, the value of the latter is significantly lower than the one predicted by theoretical calculations on isolated molecules of both **HDPPBA-N** and **HDPPBA-C** (1706 and 1708 cm<sup>-1</sup>), which indicates a lengthening of the C=O double bond due to intramolecular interactions (see Fig. S10–S21, ESI<sup>†</sup>).

In addition, and unexpectedly, IR spectra also point out to changes on the electronic properties of the conjugated skeletons due to hydrogen bond formation. To analyze this, we now focus our attention on the collective vibration over the thiophene-DPP-thiophene chain. This collective vibration appears at 1553 cm<sup>-1</sup>, being a quite sharp band in **HDPPH**. However, when hydrogen bonds take part in **HDPPBA-N** and **HDPPBA-C**, the aforementioned IR band widens showing two clear contributions at 1555–1554 cm<sup>-1</sup> and 1544 cm<sup>-1</sup>. The latter indicating some participation of the carbonyl group of the lactam ring in the supramolecular hydrogen bond interaction, which promotes enhanced  $\pi$ -conjugation in molecules having such interactions.

This can be understood considering that, when the lactam carbonyl group is involved in the hydrogen bond formation, the cross-conjugated path due to the carbonyl lactam is blocked and thus the molecular  $\pi$ -conjugation improves through the linear conjugated path (see Scheme 2). Nonetheless, it seems reasonable that only a fraction of the molecules may interact through those carbonyl groups, being the hydrogen bonding through lateral alkyl chains with amide groups the most probable one. The same conclusion is gathered from the FT-Raman spectra of the semiconductors as bulk samples, where two distinct contributions are found in the Raman band due to the collective (C=C/C-C) of the linear conjugated path (Fig. 5, right and Fig. S22–S33, ESI<sup>†</sup>).

It is expected that the hydrogen bonding through the imide carbonyl group may influence the electronic properties of the



**Fig. 4** Transfer (a) and (b) and output (c) and (d) of **HDPPBA-C** (a) and (c) and **HDPPBA-N** (b) and (d). The transfer characteristics were measured at a constant source–drain voltage of  $-80$  V. The gate voltage in the output plots varies from 0 to  $-80$  V in steps of 10 V.





charge formation in field effect transistor devices. In fact, while the control molecule **HDPPH** was inactive in OFETs, field-effect mobilities with values of  $2 \times 10^{-2} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$  and  $1 \times 10^{-2} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$  were recorded for **HDPPBA-C** and **HDPPBA-N**, respectively. Although, in thin film devices, a correct supramolecular organization is commonly crucial to boost the final electrical performance, this molecular electronic property modulation by intermolecular interactions may have a positive impact on the device performance.

### Analysis of electronic properties by spectroelectrochemistry.

#### Charge carriers' stabilization

In order to evaluate the ability of the studied semiconductors to stabilize charge carriers and to probe if non-covalent interactions play a role in the formation of these charged species, as indicated by thin film CV, *in situ* spectroelectrochemical studies were first conducted in dilute dichloromethane solution ( $10^{-4} \text{ M}$ ) at room temperature using  $0.1 \text{ M Bu}_4\text{NPF}_6$  as the supporting electrolyte and an optically transparent thin-layer electrochemical (OTTLE) cell. As shown in Fig. 6 and Fig. S34–S36 (ESI<sup>†</sup>), the three compounds behave similarly in solution, as expected, due to the absence or minimization of intermolecular interactions in diluted solutions, and all stabilize one positive charge at identical redox potentials. TD-DFT calculations support the experimental results. Therefore, no discernible changes on the molecular and electronic structures of the isolated molecules are found by the inclusion of amide groups on the lateral groups.

In contrast, the enhanced effective  $\pi$ -conjugation imparted by hydrogen bond formation, evidenced in the previous section and shown in Scheme 2, is expected to influence charge

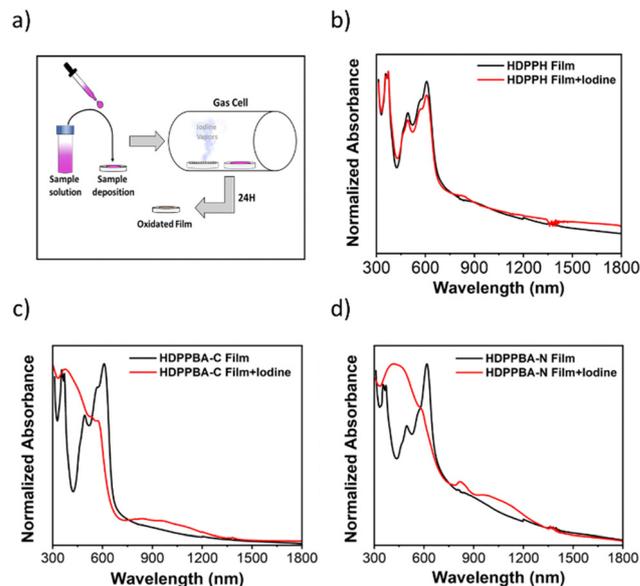


Fig. 7 (a) Schematic representation of the oxidation process with iodine vapour in a gas cell. Solid state oxidation of (b) **HDPPH**, (c) **HDPPBA-C** and (d) **HDPPBA-N**, upon iodine vapour exposition.

injection, stabilization and transport in the solid state, where intermolecular interactions are crucial. Thus, thin film oxidation was also attempted by exposing the films deposited onto the quartz substrate to iodine vapour (a very weak oxidant) for 24 hours (Fig. 7). Remarkably, different results are found for semiconductors having hydrogen-bonding units, **HDPPBA-N** and **HDPPBA-C**, compared to the control material, **HDPPH**. It should be noted that, while the formation of a positive charged

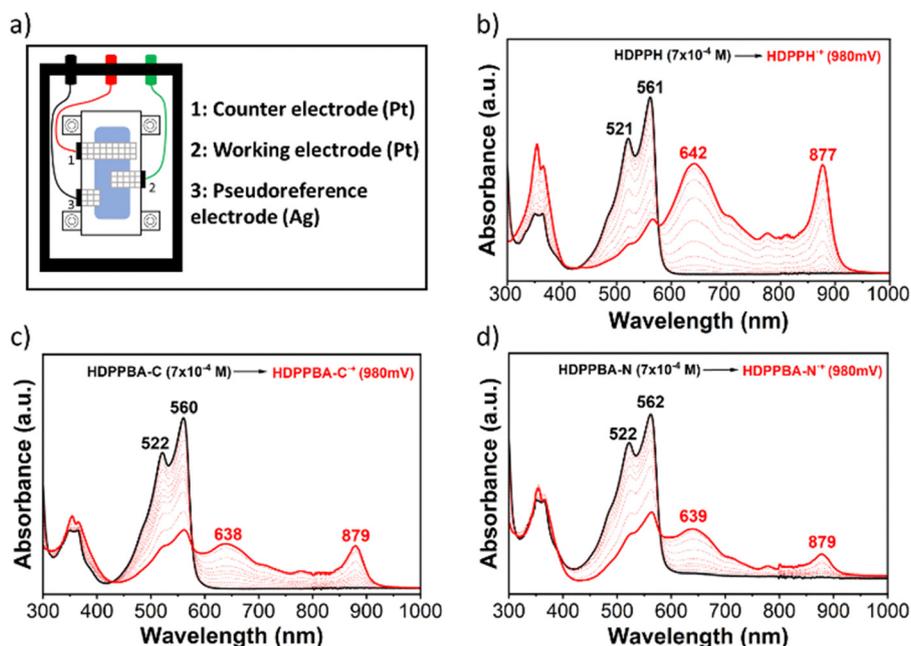


Fig. 6 (a) A scheme of optically transparent thin layer electrochemical cell (OTTLE). UV-Vis-NIR absorption spectra of (b) **HDPPH** (c) **HDPPBA-C**, and (d) **HDPPBA-N** electrochemically oxidized by the progressive increase of the oxidation potential.



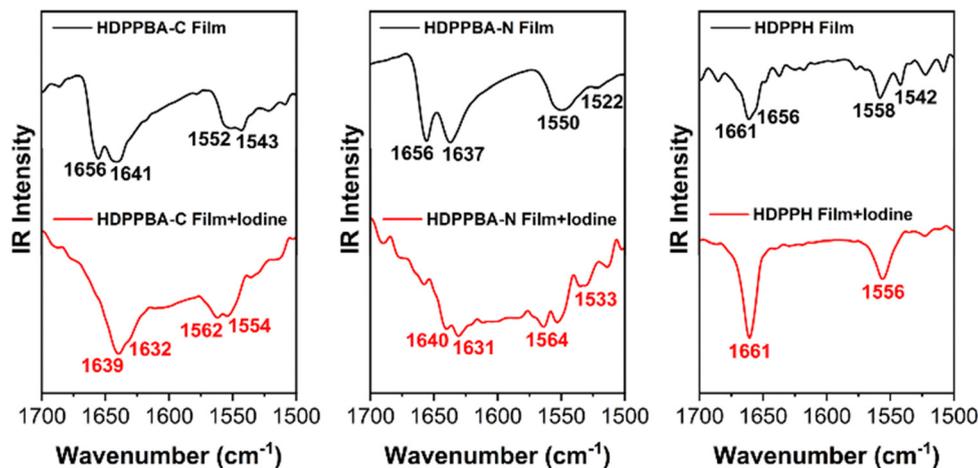


Fig. 8 Film IR spectra of HDPPBA-C (left), HDPPBA-N (center) and HDPPH (right). Recorded spectra before iodine vapour exposition (in black) and after 24 h exposition to iodine vapour (in red).

species is found for derivatives with hydrogen bonding abilities, no oxidation is found in the case of HDPPH, where no hydrogen bond interactions are possible.

To analyze this, the IR spectra of the semiconductor thin films, prior and after exposition to iodine vapour, were recorded. As shown in Fig. 8, no remarkable differences are found on the IR spectra of HDPPH before and after iodine exposition. However, in the case of HDPPBA-N and HDPPBA-C, great changes are recorded according to the radical cation formation. Therefore, the enhanced  $\pi$ -conjugated path, due to hydrogen bond formation through the imide carbonyl group, has indeed an effect on the oxidative potential of the semiconducting thin films, and thus it should facilitate the generation of free charge carriers in OFETs, contributing to the charge carrier mobility improvement shown above.

It should be noted that, while HDPPH was not active in OFETs, field effect mobilities of  $2 \times 10^{-2} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$  and  $1 \times 10^{-2} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$  were recorded for HDPPBA-C and HDPPBA-N, respectively. As far as we know, this is the first time that an enhancement of the effective  $\pi$ -conjugation length of organic semiconductors, by a simple chemical design promoting non-covalent interactions, has been shown to boost the ability of organic semiconductors to stabilize charge carriers and ultimately enhance their electrical performance.

## Conclusions

In conclusion, we present here a molecular engineering strategy that allows simultaneously control over the semiconductors morphological and intrinsic electronic properties by simply promoting hydrogen-bond formation in remarkably small DPP derivatives. We show that the electrical improvement found in semiconductors having C- or N-centered amide groups can not be solely related to supramolecular characteristics in the thin film, but it could be also ascribed to the modulation of the molecular electronic properties by the hydrogen-bond interactions. In fact, a comprehensive spectroelectrochemical and spectroscopic study points out to a complex scenario in which hydrogen-bonds

formation promote the linear conjugated path in DPP semiconductors respect to the cross-conjugated path, when the carbonyl of the DPP lactam group is involved in the hydrogen-bonds formation. We have demonstrated that this enhancement of the semiconductor effective-conjugation facilitates charge carriers formation in the thin film, as shown by cyclic voltammetry and upon iodine oxidation, which may also contribute to the boost of the electrical performance in OFETs. We feel that this synthetic approach could be extrapolated to many already existing efficient DPP-based conjugated cores, boosting their electronic performance by simply modifying their pending alkyl solubilizing units.

## Author contributions

G. M. synthesized and provided the molecules for the study and interpreted the results. R. G.-N. fabricated all the devices, performed the electrochemical oxidation experiments, interpreted the results. N. R. A. R. synthesized and provided the C-centered derivative. K. H. performed the cyclic voltammetry measurements on thin film. A. R. C. and R. P. O. supervised the work, interpreted the results, and wrote the manuscript.

## Data availability

The data supporting this article have been included as part of the ESI.†

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

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Research in Chemistry (FRC) LabEx Emerging Investigators Grant 2018 (CSC-ARC-18). A. R. C. and G. M. thank the Graduate School of Complex Systems Chemistry of Strasbourg for his doctoral fellowship funded by the French National Research Agency (CSC-IGSANR-17-EURE-0016). The authors thank the Agence Nationale de la Recherche (ANR JCJC TOTAL-BOND 2020). R. P. O and A. R. C. acknowledge the MICIIN for the REDES project "RED2022-134503-T". The authors would also like to thank the computer resources, technical expertise and assistance provided by the SCBI (Supercomputing and Bioinformatics) centre of the University of Málaga. The Vibrational spectroscopy (EVI) lab of the Research Central Services (SCAI) of the University of Málaga is also gratefully acknowledged.

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