

Cite this: *J. Mater. Chem. C*,
2024, 12, 3185

Condensed π -molecular arrangement for $-\text{C}_2\text{H}_4\text{SO}_3^-$ armed naphthalenediimide[†]

Ryo Ide,^a Ayumi Kawasaki,^a Takashi Takeda,^{id abc} Shun Dekura,^{id ab}
Nori-hisa Hoshino,^{id d} Wakana Matsuda,^e Shu Seki^e and Tomoyuki Akutagawa^{id *ab}

A highly condensed packing structure with π -molecules is important to achieve high carrier transport properties. To realize condensed π -molecular arrangements, we focused on electrostatic crystal lattices consisting of cation–anion pairs and further designed an anionic *N,N'*-bis(ethyl sulfonate)-naphthalenediimide (**ESNDI**²⁻), which introduces structural flexibility at two $-\text{C}_2\text{H}_4\text{SO}_3^-$ side arms. The size of the counter cation (M^+) was varied from Na^+ , K^+ , Rb^+ , and Cs^+ to design a precise crystal lattice of $(\text{M}^+)_2(\text{ESNDI}^{2-})\cdot(\text{H}_2\text{O})_n$ salts, which is divided into hydrated salts ($\text{M}^+ = \text{Na}^+$ and K^+) and anhydrous salts ($\text{M}^+ = \text{K}^+$, Rb^+ , and Cs^+). Salt for $\text{M}^+ = \text{K}^+$ is a boundary between hydrated and anhydrous crystals, and exhibited reversible H_2O adsorption–desorption behavior at 298 K. In addition, the K^+ salts showed transient conductivity ($\phi\Sigma\mu$) switching behavior associated with the H_2O adsorption–desorption cycle. The anhydrous salts for $\text{M}^+ = \text{K}^+$, Rb^+ , and Cs^+ had isomorphous crystal structures, with the appearance of a two-dimensional (2D) herringbone arrangement in **ESNDI**²⁻. The condensed packing arrangement and highest $\phi\Sigma\mu$ value were observed in $(\text{Rb}^+)_2(\text{ESNDI}^{2-})$ salt. In these anhydrous salts, crystalline domains were easily formed using the simple casting method on the substrate surface. The temperature- and frequency-dependence of the dielectric constant indicated the presence of thermally induced structural fluctuations in these ionic crystal lattices. In hydrated crystals, thermal motions of Na^+ , K^+ , and H_2O were dominant, whereas thermal fluctuations in the rigid $\text{M}^+\cdots\text{O}_3\text{S}^-$ Coulomb lattice were observed in anhydrous crystals as Debye-type relaxation. Precise structural control of the electrostatic crystal lattice determines the optimization conditions for achieving condensed π -molecular arrangement.

Received 16th November 2023,
Accepted 9th January 2024

DOI: 10.1039/d3tc04231f

rsc.li/materials-c

Introduction

The crystal structures formed by organic molecules are governed by diverse intermolecular interactions compared to inorganic materials.^{1–3} Even simple aromatic hydrocarbons composed of carbon and hydrogen atoms, such as naphthalene,⁴ pyrene,⁵ perylene,⁶ and coronene,⁷ undergo π – π and C–H interactions, reflecting their molecular shapes, and various packing patterns such as π -stacking and herringbone arrangements are observed.^{8–10} Intermolecular interactions acting on organic π -molecules have a wide range of energy scales (1–100 kJ mol⁻¹) and directionality, including electrostatic, hydrogen-bonding,

charge-transfer, and van der Waals interactions.^{11–14} The essential feature of organic materials is the high degree of chemically design freedom for the crystal lattice by focusing on intermolecular interactions.

Higher dimensionality of intermolecular interactions is advantageous for high carrier transport properties, which improve in the order of three-dimension (3D) > two-dimension (2D) > one-dimension (1D).^{15–18} While most inorganic materials are formed by 3D chemical bonds, planar π -electronic organic molecules tend to form 1D π -stacking structures, which exhibit instability to external factors such as impurities and lowering of temperature.^{19,20} Many organic materials composed of π -planes need to have a specific molecular design to form a 2D molecular arrangement. This is an important design strategy for achieving high carrier transport properties. One chemical solution to form a 2D electronic structure with high carrier transport properties is to introduce heavy atoms such as S and Se with large polarization into the π -plane composed of carbon and hydrogen atoms.^{21–28} This is a molecular design that has been used with great success in molecular systems such as bis(ethylenedithiolo)-tetrathiafulvalene (BEDT-TTF) and [1]benzothieno[3,2-*b*][1]benzothiophene (BTBT).^{27–30} There are also diverse molecular systems such as

^a Graduate School of Engineering, Tohoku University, Sendai 980-8579, Japan^b Institute of Multidisciplinary Research for Advanced Materials (IMRAM), Tohoku University, 2-1-1 Katahira, Aoba-ku, Sendai 980-8577, Japan.
E-mail: akutagawa@tohoku.ac.jp^c Faculty of Science, Shinshu University, 3-1-1 Asahi, Matsumoto 390-8621, Japan^d Faculty of Engineering, Niigata University, Niigata 950-2102, Japan^e Graduate School of Engineering, Kyoto University, Kyoto, 615-8510, Japan[†] Electronic supplementary information (ESI) available. CCDC 2308337–2308340. For ESI and crystallographic data in CIF or other electronic format see DOI: <https://doi.org/10.1039/d3tc04231f>

phthalocyanine with a wide π -plane and C_{60} with 3D intermolecular interactions can be realized, and various approaches have been proposed to control the strength and dimensionality of the intermolecular interactions.^{31–36}

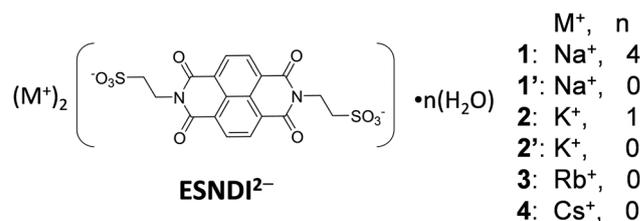
Electrostatic interactions with energies above 100 kJ mol^{-1} can be used to control the arrangement of organic π -molecules and increase intermolecular interactions.^{37–41} Recently, an approach for controlling the arrangement of π -planar cations and anions was reported by Maeda *et al.* to develop organic materials in which the Coulomb interaction can be controlled.^{39–41} We have designed supramolecular cations encapsulated with crown ethers to reduce the Coulomb interactions on anions by shielding positive charges, and have designed a variety of anion arrangements for π -electronic molecules.^{42–46} By screening the positive and negative charges, various types of molecular arrangements similar to those of van der Waals crystals can be realized.

π -Planar naphthalene diimide (**NDI**) molecules exhibit n-type semiconducting properties and are used to fabricate a variety of functional organic materials due to their electronic functions such as carrier transport, molecular adsorption, magnetic and chromic properties.^{47–58} Various molecular designs for the nitrogen sites at the end of the **NDI** molecule are easily possible, and a variety of molecules can be prepared. Therefore, the **NDI** system provides interesting research targets in terms of electronic function design and dense packing of π -molecules. Notably, *N,N'*-bis(cyclohexyl)-**NDI** crystals exhibit high electron mobility due to their 2D molecular arrangement.⁵⁸ For similar n-type organic semiconducting perylene diimide (**PDI**) derivatives, various substituents can be also introduced to the terminal nitrogen atom, and the molecular arrangements from nanoscale to mesoscale, including supramolecular polymers, have been reported. Since the size of the π -electron system differs between **NDI** and **PDI** π -molecules, the substituent introduced to the nitrogen sites has a different effect on the molecular arrangement and packing.^{59–63} The simple **NDI** π -core facilitates detailed studies on the design of the molecular arrangement.

Focusing on effective electrostatic interactions in crystals, we have been studying the control of molecular arrangement of n-type organic semiconducting anionic **NDI** derivatives, and the fabrication of functional materials to design crystal lattices with high chemical stability.^{64–68} An **NDI** derivative with benzenesulfonate groups (**BSNDI**²⁻) showed high thermal stability with the crystal reaching a decomposition temperature around 850 K with alkali metal cations such as Na^+ and K^+ .⁶⁴ However, the structurally rigid benzenesulfonate group allowed only rotational freedom around the $\text{N}-\text{C}_6\text{H}_4\text{SO}_3^-$ bond, resulting in a low degree of freedom in molecular conformation, which did not lead to a suitable electron transport 2D layer with a condensed π -molecular arrangement. To solve the above problem, the **NDI** π -core with a terminal flexible $-\text{C}_2\text{H}_4\text{COO}^-$ group, which improves the molecular conformational freedom, was combined with $\text{M}^+ = \text{Li}^+, \text{Na}^+, \text{K}^+, \text{Rb}^+, \text{and } \text{Cs}^+$ to form cationic-anion salts.^{67,68} As a result, the high structural stability of the Coulomb crystal lattice and control of the electronic

transport properties in conjunction with the H_2O adsorption-desorption behavior were confirmed. It is important to investigate in detail what kind of anionic structural units can realize the condensed **NDI** array with the 2D properties, high thin film formation ability, and high electron mobility from the viewpoints of molecular structure, crystal structure, and control of physical properties.

Recently, an **NDI** derivative with the $-\text{C}_2\text{H}_4\text{PO}_3\text{H}^-$ group (**EPNDI**²⁻) was used for the proton conducting crystal *via* a hydrogen-bonding network between $-\text{PO}_3\text{H}^-$ and NH_4^+ in $(\text{NH}_4^+)_2(\text{EPNDI}^{2-})$ salt.^{69,70} In contrast, the formation of pyrophosphate bonds ($-\text{PO}_2-\text{O}-\text{PO}_2-$) through the thermal polymerization reaction between the $-\text{PO}_3\text{H}^-$ units of the adjacent **EPNDI**²⁻ was observed in the $(\text{Na}^+)_2(\text{EPNDI}^{2-})$ crystals. The unit introduced at the terminal nitrogen site of the **NDI** skeleton plays an important role in the physical properties and the condensed arrangement of **NDI**. In this study, we focused on **ESNDI**²⁻ with a flexible ethyl sulfonate ($-\text{C}_2\text{H}_4\text{SO}_3^-$) group for structural flexibility and prepared $(\text{M}^+)_2(\text{ESNDI}^{2-})$ salts with alkali metal ions of $\text{M}^+ = \text{Na}^+, \text{K}^+, \text{Rb}^+, \text{and } \text{Cs}^+$ (Scheme 1). In the previous **BSNDI**²⁻ salts, only rotational degree of freedom around the $\text{N}-\text{C}_6\text{H}_4\text{SO}_3^-$ bond is allowed, whereas in **ESNDI**²⁻, the conformational freedom of the ethyl sulfonate group is enhanced and the various molecular arrangement of the **NDI** π -core is observed. Although the crystal structure and carrier generation by photoirradiation of the **ESNDI**²⁻ molecule combined with $\text{M}^+ = \text{Na}^+, \text{K}^+, \text{Rb}^+, \text{Mg}^{2+}$, and Al^{3+} salts have been reported,^{71–75} there has been no comparative study on the molecular arrangements, formation of condensed π -packing structures, and electron transport properties when the size of M^+ is systematically changed. We focused on $(\text{M}^+)_2(\text{ESNDI}^{2-})$ salts with $\text{M}^+ = \text{Na}^+, \text{K}^+, \text{Rb}^+, \text{and } \text{Cs}^+$ and attempted to investigate changes in the molecular arrangement and electron transport properties in a series of crystals in terms of the appearance of condensed π -molecular arrangements. Hydrated crystal $(\text{Na}^+)_2(\text{ESNDI}^{2-}) \cdot (\text{H}_2\text{O})_{4.0}$ (**1**) and anhydrous crystal $(\text{Na}^+)_2(\text{ESNDI}^{2-})$ (**1'**) were obtained, whereas hydrous $(\text{K}^+)_2(\text{ESNDI}^{2-}) \cdot (\text{H}_2\text{O})_{1.0}$ (**2**) and anhydrous crystal $(\text{K}^+)_2(\text{ESNDI}^{2-})$ (**2'**) were also obtained. The reversible adsorption-desorption cycle of H_2O was confirmed through the sorption isotherms and the reversible conversion between hydrated and anhydrous crystals was possible in a single crystal state for $\text{M}^+ = \text{K}^+$. For Rb^+ and Cs^+ salts, only the anhydrous crystals $(\text{Rb}^+)_2(\text{ESNDI}^{2-})$ (**3**) and $(\text{Cs}^+)_2(\text{ESNDI}^{2-})$ (**4**) were obtained. We examine the molecular arrangement, thin film formation



Scheme 1 Molecular structures of **ESNDI**²⁻ and its salts with $\text{M}^+ = \text{Na}^+, \text{K}^+, \text{Rb}^+, \text{and } \text{Cs}^+$ in this study.



ability, dielectric response, and electron transport properties of the series of crystals 1–4.

Results and discussion

Thermal stability

The thermal stability of crystals 1–4 was evaluated using TG measurements (Fig. S1, ESI[†]). Salts 1 and 2 were stably obtained by hydrated single crystals containing four H₂O molecules and one H₂O molecule as crystallization solvents, whereas salts 3 and 4 were isolated as anhydrous crystals. Crystal 1 exhibits a weight loss of 12.1% upon crystal heating from 353 to 393 K, and changes to anhydrous crystal 1' due to the desorption of four H₂O molecules (calculated weight of 12.0%). The hydrated crystal 2 also forms a dehydrated crystal 2' with one H₂O molecule desorbed, with a weight loss of about 2.71% around 370 K. The decomposition temperatures of dehydrated salts 1', 2', 3, and 4 were 713, 749, 749, and 752 K, respectively, indicating high thermal stability. Salts 1', 2', 3, and 4 showed high thermal stability, whereas the van der Waals crystal, *N,N'*-bis(phenyl)-NDI, showed a melting point at 593–595 K,⁸¹ indicating that M⁺...⁻O₃S⁻ electrostatic Coulomb interactions between the cation and anion were effective in improving thermal stability. The phase transition behavior of salts 1', 2', 3, and 4 after dehydration was also evaluated using DSC measurements (Fig. S2, ESI[†]), and there was no phase transition peak and a single crystalline phase in a wide temperature range of 200–600 K.

Crystal structures

The crystal structures of salts 1–4.0(H₂O) and 3 have been reported previously.^{71–75} However, there is no evaluation of the dimensionality of the electronic structure and electron transport properties. No systematic changes in the molecular arrangement have been reported when cations are increased in size from Na⁺, K⁺, Rb⁺, to Cs⁺. Herein, we attempted to focus on the dimensionality change of the electronic structure of crystals 1–4.^{76,77} In crystal 2, the desorption of H₂O molecules was observed in the single crystalline state upon temperature increase, and the molecular arrangements before and after the H₂O desorption were evaluated in single crystals 1 at 100 K and anhydrous 1' at 300 K.

The molecular arrangements of crystals 1–4 were similar to each other, although the number of crystallographically independent species was different. In particular, anhydrous crystals 2', 3, and 4 had similar lattice constants and similar molecular arrangements (Fig. S1, ESI[†]). For crystals 1 and 2, the presence of H₂O molecules was necessary to form a closest-packing structure. Fig. 1 shows the crystal structure of salt 4 at 100 K. The crystal structure of salt 4 is identical to that of salt 3 at 300 K as previously reported.⁷¹ One ESNDI²⁻ and two Cs⁺ are the crystallographically independent structural units, and the electronic conducting layer consisting of the NDI backbone of ESNDI²⁻ was observed in the *bc* plane. The cation–anion electrostatic inorganic layer in the *bc* plane composed of

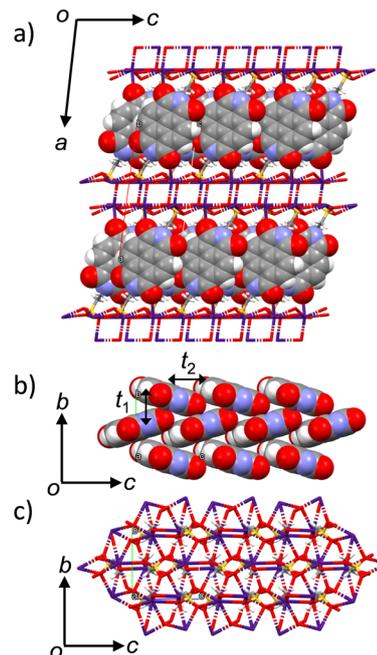


Fig. 1 Crystal structure of salt 4 at 100 K. (a) Unit cell viewed along the *b*-axis using a CPK representation of NDI units. (b) The 2D herringbone NDI arrangement in the *bc* plane with *t*₁ and *t*₂ interactions. (c) The 2D Cs⁺...⁻O₃S⁻ electrostatic networks in the *bc* plane.

Cs⁺...⁻O₃S⁻ and NDI layers was alternatively elongated along the *a*-axis. The electrical conducting NDI layer in the *bc* plane is in a herringbone arrangement, which is expected to form a 2D electronic structure.

Electronic structures

Using the extended Hückel calculation, the transfer integral (*t*, meV) of the LUMOs between the NDI layers of salt 4 was calculated to be *t*₁ = −21.0 meV along the *b*-axis and *t*₂ = −18.6 meV along the *c*-axis, suggesting the formation of a 2D electronic structure (Fig. S3, ESI[†]).⁷⁸ Calculation of the LUMO band of salt 4 using the tight-binding approximation yields a closed 2D Fermi surface assuming a partially occupied state of Fermi energy *E*_F = 0.15, confirming the formation of a 2D electronic structure (Fig. S4, ESI[†]). In the inorganic layer in the *bc* plane, two −SO₃⁻ groups and a C=O oxygen of ESNDI²⁻ formed a hepta-coordinated structure around Cs⁺, forming a cation–anion electrostatic network layer. The average distance (*d*_{Cs–O}) between the oxygen atoms of the −SO₃⁻ group and Cs⁺ was 3.154 Å, whereas the *d*_{Cs–O} between the oxygen atom of C=O and Cs⁺ was 3.279 Å.

The electrical conducting NDI arrangements in crystals 1–4 exhibited a herringbone pattern that could be represented by the ratio of *t*₁ and *t*₂ interactions, wherein the LUMO band changed depending on the cation size (Fig. S3, ESI[†]). In crystal 1, the intermolecular interactions of *t*₁ = −0.69 meV and *t*₂ = +5.24 meV were observed along the *b*- and the *c*-axis, respectively, wherein these values were smaller than those found in salt 4. The negligible magnitude of *t*₁ interaction along the *b*-axis of salt 1 forms a 1D electronic structure, in which the



LUMO band showed a narrow bandwidth of less than 0.005 eV at the Γ - S energy trace (Fig. S5, ESI[†]). Assuming the Fermi surface of salt **1** at $E_F = 0.15$, the formation of a 1D electronic structure is confirmed. In salt **2**, the intermolecular interactions of $t_1 = -2.20$ meV and $t_2 = 18.9$ meV were observed along the a - and the b -axis, respectively, by changing cations from Na^+ to K^+ , which also changed the number of crystallization H_2O molecules from four to one. The intermolecular interaction between the LUMOs increased 3 times due to the replacement the cation from Na^+ to K^+ in salt **2**. Additional $t_3 = 3.20$ meV is also observed along the $a + b$ axis. The structure of the LUMO band indicates the presence of a 1D Fermi surface (Fig. S6b, ESI[†]), assuming $E_F = 0.15$. When crystal **2** changed to anhydrous **2'**, significant enhancement of intermolecular interactions is observed with $t_1 = -27.2$ meV along the b -axis and $t_2 = -19.4$ meV along the c -axis. The resulting LUMO band showed the presence of a circular 2D Fermi surface assuming $E_F = 0.15$ (Fig. S6a, ESI[†]). The crystal change from hydrated salt **2** to anhydrous salt **2'** modulated the closest-packing structure and transformed the molecular arrangement and electronic structure of the **NDI** π -core from a 1D to 2D pattern. In salt **3**, an almost equivalent intermolecular interaction of $t_1 = -25.1$ meV along the b -axis and $t_2 = -21.7$ meV along the b -axis was observed, where the LUMO band assuming $E_F = 0.15$ indicated the presence of a 2D electronic structure (Fig. S7, ESI[†]). The t values in salt **4** were smaller than those in salt **3**, suggesting the best-fit cation size of Rb^+ in the condensed **NDI** π -arrangement.

Cation dependence for packing structures

In the inorganic electrostatic layer of salt **1**, oxygen atoms of two $-\text{SO}_3^-$, four H_2O , and one $=\text{O}$ of ESNDI^{2-} were hepta-coordinated to Na^+ in a pentagonal bipyramidal coordination geometry, forming an electrostatic network in the ab plane, where one Na^+ was coordinated in an equatorial plane of two $-\text{SO}_3^-$ groups and one $=\text{O}$ in a pentagonal configuration with the average distance of $d_{\text{Na-O}} = 2.395$ Å. In addition, the two H_2O molecules coordinated up and down concerning Na^+ , and the average $\text{Na}^+\cdots\text{O}$ distance of $d_{\text{Na-O}} = 2.460$ Å was almost the same as pentagonal coordination. In salt **2**, the two $-\text{SO}_3^-$ groups, one H_2O , and one $=\text{O}$ of ESNDI^{2-} formed a pseudo-dodecahedral-shaped octa-coordinated structure around K^+ , forming an electrostatic network in the ab plane. The interaction between the $=\text{O}$ group and K^+ was observed at $d_{\text{K-O}} = 2.967$ Å, whereas that between the $-\text{SO}_3^-$ oxygen atom and K^+ was observed at an average distance of $d_{\text{K-O}} = 2.883$ Å. In the anhydrous salt **2'**, a clear change in the molecular arrangement of the **NDI** π -core was observed, whereas no significant change in the coordination structure around K^+ was observed. The oxygen atoms of $-\text{SO}_3^-$ and the $=\text{O}$ of the ESNDI^{2-} formed seven-coordinated geometry to the Rb^+ salt, forming a multi-point electrostatic interaction network, with $d_{\text{Rb-O}} = 3.180$ Å for $=\text{O}$ and $d_{\text{Rb-O}} = 2.970$ Å for $-\text{SO}_3^-$. In all salts, two $-\text{SO}_3^-$ groups were coordinated to M^+ , connecting the adjacent **NDI** layers and forming a Coulomb crystal lattice.

Fig. 2 summarizes the $\text{M}^+\cdots\text{O}_3\text{S}^-$ arrangements in the inorganic network layers and the t values of **NDI** for salts **2'**,

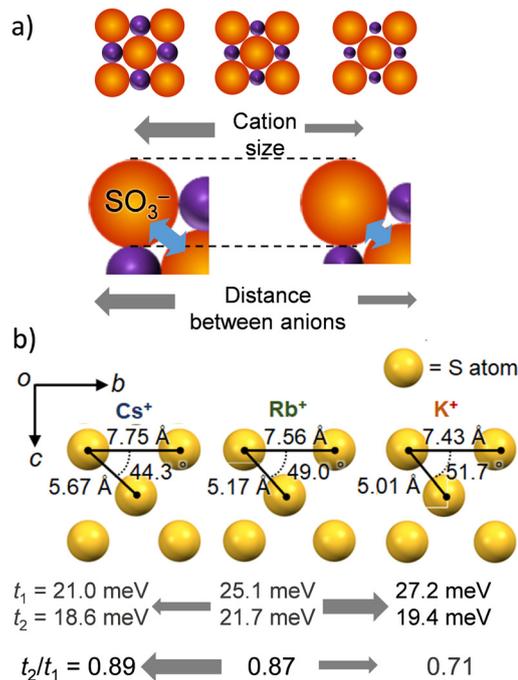


Fig. 2 Change in the inorganic $\text{M}^+\cdots\text{O}_3\text{S}^-$ electrostatic layer from $\text{M}^+ = \text{K}^+, \text{Rb}^+$ to Cs^+ . (a) Modulation of the $\text{M}^+\cdots\text{O}_3\text{S}^-$ electrostatic inorganic layer with increasing M^+ size. The orange and purple spheres represent $-\text{SO}_3^-$ and M^+ , respectively. (b) $\text{S}\cdots\text{S}$ interatomic configuration and t_1 and t_2 values for $\text{M}^+ = \text{K}^+, \text{Rb}^+, \text{Cs}^+$. The M^+ dependent t_1 , t_2 , and their ratio t_2/t_1 are also shown in the bottom.

and **4**. The cation M^+ forms a coordination structure through electrostatic interactions with two $-\text{SO}_3^-$ groups (Fig. 2a). Therefore, a much larger size of M^+ increases the average distance between $-\text{SO}_3^-$ groups. Flexible $-\text{C}_2\text{H}_4\text{SO}_3^-$ groups of the ESNDI^{2-} molecule have a high degree of conformational freedom. An increase in M^+ size alters the intermolecular interactions of the **NDI** π -core responsible for carrier transport property, leading to conformational changes that form a closest-packing structure. Fig. 2b summarizes the $\text{S}\cdots\text{S}$ distances and angles between the three S atoms of the neighboring $-\text{SO}_3^-$ groups of ESNDI^{2-} molecules for $\text{M}^+ = \text{K}^+, \text{Rb}^+, \text{Cs}^+$, where the $\text{S}\cdots\text{S}$ distances $d_{\text{S-S}}$ increase in the order of Cs^+ ($d_{\text{S-S}} = 7.75$ Å) $>$ Rb^+ ($d_{\text{S-S}} = 7.56$ Å) $>$ K^+ ($d_{\text{S-S}} = 7.43$ Å) with increasing M^+ size, whereas the $\text{S}\cdots\text{S}\cdots\text{S}$ ($f_{\text{S-S-S}}$) angle between the three S atoms decreases in the order $f_{\text{S-S-S}} = 51.7^\circ$ (K^+) $>$ 49.0° (Rb^+) $>$ 44.3° (Cs^+). The t values reached a maximum at $t_1 = 25.1$ and $t_2 = 21.7$ meV for $\text{M}^+ = \text{Rb}^+$, and the change of M^+ size from Rb^+ to Cs^+ reduced to $t_1 = 21.0$ and $t_2 = 18.6$ meV, where the intermolecular interactions between the **NDI** π -cores were suppressed. Regarding the dimensionality of the intermolecular interactions between **NDIs**, the dimensionality tended to improve with increasing M^+ size, with t_2/t_1 values of 0.89 (Cs^+) $>$ 0.87 (Rb^+) $>$ 0.71 (K^+). However, the increase in dimensionality and the magnitude of the t value tend to be in conflict, and it can be concluded that the Rb^+ salt was the best condition for obtaining the 2D electronic structure.



Structural transformation from 2 to 2'

Single crystals **2** can undergo a structural change from monohydrate **2** to anhydrous crystal **2'** while retaining its single crystalline state. Single crystal X-ray structural analyses of both states reveal how the molecular arrangement changes with the H₂O adsorption–desorption cycle. H₂O molecules incorporated in the inorganic K⁺⋯O₃S⁻ network are desorbed from the crystal lattice by thermal annealing, changing the rearrangement of K⁺ and –SO₃⁻ and modulating the molecular arrangement in the **NDI** layer responsible for electron transport property (Fig. 3a). Fig. 3b shows the change in intermolecular interactions in the 2D layer of **NDIs** using *t* values. Intermolecular interactions of *t*₁ = +18.9, *t*₂ = –3.2, and *t*₃ = +2.20 meV are observed in salt **2**, and the small magnitude of the latter two *t* values forms a 1D electronic structure in which the *t*₁ value along the *a*-axis is essentially dominant (Fig. 3b). After H₂O desorption by thermal annealing, the π-stacked **NDI** arrangement tilts toward the molecular long axis and changes to a herringbone arrangement. The desorption of H₂O molecules from single crystal **2** exhibits a drastic effect on the electronic structure.

H₂O adsorption–desorption of salts 1 to 2

The stability of anhydrous salts **1'** and **2'** under atmospheric conditions was evaluated. Salt **1'** was re-adsorbed using H₂O by leaving it in H₂O vapor overnight, and its H₂O adsorption amount was evaluated using TG measurements. The salt **1'** exposed to distilled H₂O vapor by allowing it to stand overnight in a vial at 298 K showed a weight loss of 12.1% equivalent to 4.0 molecules of H₂O in the TG chart (iii in Fig. S8, ESI[†]), which

was in agreement with the weight loss (12.1%) of hydrated crystal **1** (i in Fig. S8, ESI[†]). In contrast, the TG chart for the **1'** salt with H₂O adsorbed in air showed no weight loss, indicating that salt **1'** does not re-adsorb H₂O in the air (ii in Fig. S8, ESI[†]). Similar measurements were examined for salt **2'**. Sample **2'** exposed to distilled H₂O vapor showed a weight loss of 2.71% corresponding to 0.92 amount of H₂O molecules (ii in Fig. S9, ESI[†]), and H₂O re-adsorbed powder **2'** showed a weight loss of 2.11% corresponding to 0.71 molecules of H₂O. Powder **2'** was confirmed to exhibit H₂O re-adsorption behavior even under atmospheric conditions (i in Fig. S9, ESI[†]), which is different from the behavior of salt **1'**.

The adsorption and desorption process of H₂O on crystal **2'** was investigated through the measurement of sorption isotherms at 273 and 298 K (Fig. 3c). In anhydrous crystal **2'**, a rapid increase in adsorption amount (*n*_a, mol mol⁻¹) was observed around relative pressure *P*/*P*₀ = 0.3 with increasing H₂O pressure, confirming a gate-opening H₂O adsorption mechanism. When *P*/*P*₀ is decreased, a stable state of *n*_a = 1.0 is observed around *P*/*P*₀ = 0.1, and a hysteresis behavior for the H₂O adsorption–desorption process is confirmed. Thus, crystal **2'** can be reversibly converted to a monohydrated crystal by changing the relative pressure of H₂O. A similar change is observed for salt **1'** at a relative pressure of *P*/*P*₀ = 0.7, whereas the adsorption of H₂O at *P*/*P*₀ = 1.0 for crystal **1'** at 298 and 273 K is *n*_a = 3.5 and 3.0, respectively, whose amount is lower than that of the four H₂O molecules found in salt **1**. When the relative pressure was decreased from *P*/*P*₀ = 1.0, the hydrated state remained stable and no reversible change to *n*_a = 0 was observed.

Dielectric property

To evaluate the dynamics of polar structural units in salts **1–4**, the temperature (*T*) and frequency (*f*) dependent dielectric responses were evaluated in the AC impedance method using a compressed pellet under nitrogen flow. Fig. 4a and b show the *T*- and *f*-dependent real part dielectric constant ϵ_1 for salts **2** and **2'**.

The ϵ_1 and imaginary part dielectric constant (ϵ_2) of salt **2** are observed at $\epsilon_1 \sim 10$ and $\epsilon_2 \sim 20$ at 220 K, respectively, and these values increase significantly to $\epsilon_1 \sim 400$ and $\epsilon_2 \sim 4000$ at 350 K (Fig. S10a, ESI[†]). Since neutral **NDI** derivatives without ionic pairs or a crystallization solvent do not exhibit a dielectric response,⁶⁴ the dielectric response found in salt **2** is presumably due to the dynamic behavior of the polar H₂O molecule or the ionic component such as K⁺ and K⁺⋯SO₃⁻. A monotonic increase in ϵ_1 and ϵ_2 with increasing *T* is observed for salt **2**. In contrast, ϵ_1 in anhydrous salt **2'** shows a broad peak that depends on *T*- and *f*-values spanning the range of 200 < *T* < 350 K and 1.0 < *f* < 10 kHz (Fig. 4b). The *T*-dependence of the Cole–Cole plots for hydrated salt **2** and anhydrous salt **2'** (Fig. S10 and 11, ESI[†]) showed a semicircular trace in both *T*-ranges, indicating the presence of dielectric relaxation of a Debye-type single-component. For salt **2**, the K⁺ conductivity (σ_{K^+}) obtained from the Cole–Cole plot (Fig. S10, ESI[†]) showed $\sigma_{K^+} = 1.6 \times 10^{-7}$ S cm⁻¹ at 350 K and *E*_a = 466 meV. Next, an

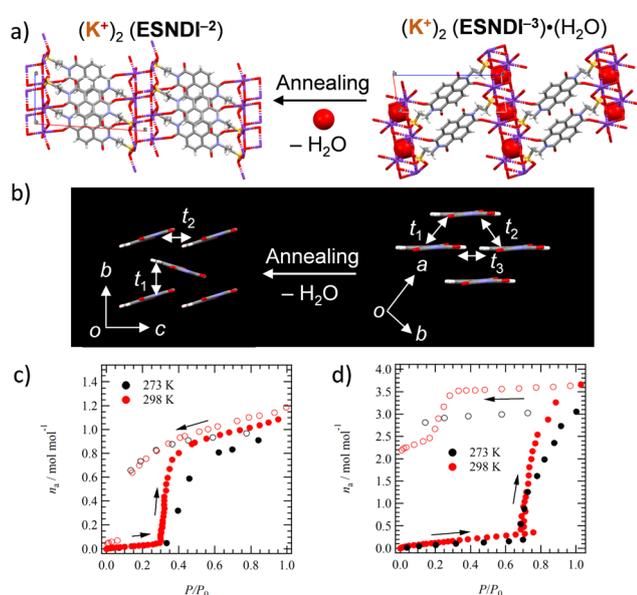


Fig. 3 Structural changes associated with the H₂O adsorption–desorption cycle. (a) Change in the molecular arrangement in salts **2** and **2'**. (b) Change in intermolecular interactions between **NDI** π-cores in salts **2** and **2'**. Adsorption–desorption isotherms of (c) salt **2'** at 273 and 298 K for H₂O, and (d) salt **1'** at 273 and 298 K for H₂O.



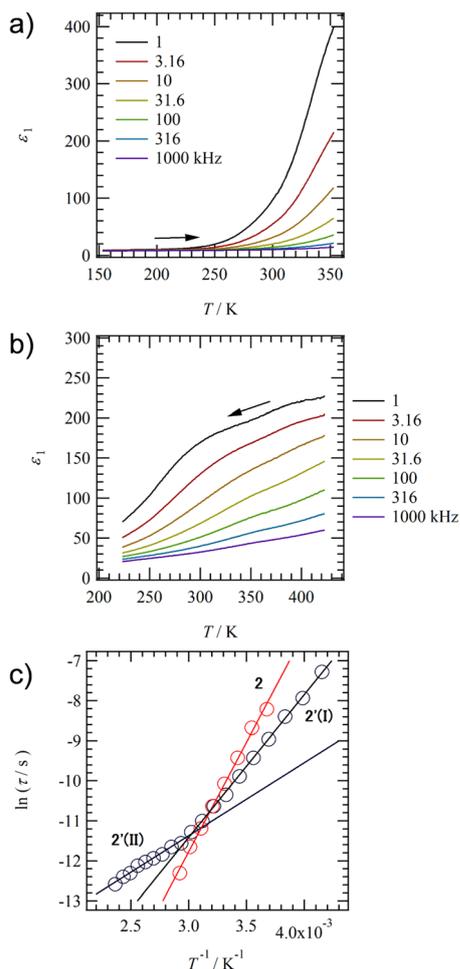


Fig. 4 Dielectric response of hydrated salt **2** and anhydrous salt **2'** using the compressed pellets. T - and f -dependent ϵ_1 of (a) salt **2** and (b) salt **2'**. (c) $\ln(\tau) - T^{-1}$ plots of salts **2** and **2'**. A crossover between two types of relaxation processes I at $T > 330$ K and II at $T < 330$ K is observed in anhydrous salt **2'** around 330 K.

$\ln(\tau) - T^{-1}$ plot was constructed from the Cole–Cole plot of the anhydrous salt **2'**, and E_a was calculated for dielectric relaxation. The $\ln(\tau) - T^{-1}$ plot shows a slope change around 330 K, and a crossover phenomenon from $E_a(\text{I}) = 310$ meV at $T < 330$ K to $E_a(\text{II}) = 160$ meV at $T > 330$ K appears. The dielectric relaxation of anhydrous salt **2'** is considered to originate from the thermal fluctuation of the ionic crystal lattice because the E_a value at $T > 330$ K is about half of the value at $T < 330$ K and the E_a is more than 200 meV lower than that of the hydrated salt **2**. In general, cations such as Na^+ and K^+ in the ionic crystals are more strongly bound to the crystal lattice than Li^+ .^{79,80} However, K^+ coordinated with H_2O suppressed the anion-cation electrostatic interaction and showed ionic conductivity at 300 K.⁸⁰ In salt **2**, the coordination of H_2O to K^+ shielded the positive charge of K^+ , resulting in K^+ conduction. In contrast, ionic conduction in salt **2'** was suppressed by the strong electrostatic $\text{K}^+ \cdots \text{O}_3\text{S}^-$ interaction, and a dielectric response due to thermal fluctuations at $\text{K}^+ \cdots \text{O}_3\text{S}^-$ bond was observed. The crossover phenomenon is considered to be caused by the

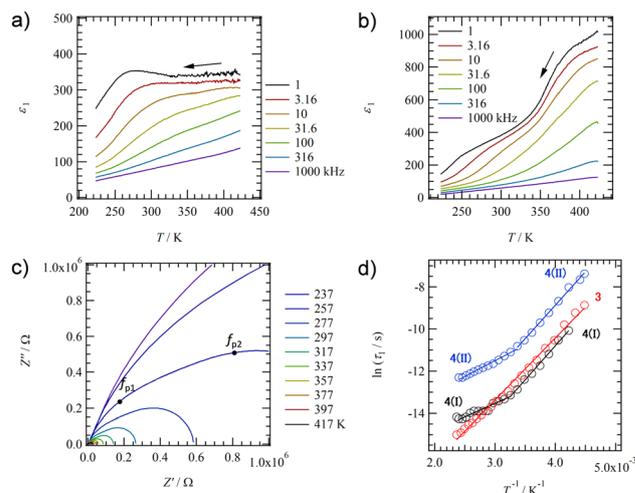


Fig. 5 Dielectric response of salts **3** and **4** using the compressed pellets. T - and f -dependent ϵ_1 of (a) salt **3** and (b) salt **4**. (c) T -dependent Cole–Cole plots of salt **4** with two f -peaks at f_{p1} and f_{p2} . (d) $\ln(\tau) - T^{-1}$ plots of salts **3** and **4**. A crossover between two types of relaxation process I at $T > 310$ K and process II at $T < 310$ K is observed in salt **4**.

T -dependent change in the coordination environment of $-\text{SO}_3^-$ and $=\text{O}$ concerning K^+ .

Fig. 5a shows the T - and f -dependent ϵ_1 behaviour of salt **3**. The ϵ_1 value of salt **3** exhibited a broad peak over a wide measurement T - and f -ranges of $223 < T < 350$ K and $1.0 < f < 31.6$ kHz, respectively, showing similar behaviour to those of salt **2'**. The crystal structure of anhydrous salt **2'** was similar to that of salt **3**, suggesting the same dielectric relaxation mechanism of a Coulomb crystal lattice composed of $\text{M}^+ \cdots \text{O}_3\text{S}^-$ for $\text{M}^+ = \text{K}^+$ and Rb^+ . The T -dependent Cole–Cole plot for salt **3** showed a semicircular trace and used the peak top f_p values to deduce an $\ln(\tau) - T^{-1}$ plot to obtain the E_a value (Fig. 5d and Fig. S14, ESI†). A crossover of E_a values obtained from dielectric relaxation processes I to II with temperature change was already observed for anhydrous salt **2'**, while $E_a = 256$ meV explained the dipole relaxation process of salt **3** in all measured T -ranges. The E_a value of salt **3** has the same origin as the dielectric relaxation process I at the low- T region of salt **2** ($E_a = 160$ meV), which is considered to undergo the thermal fluctuation of the $\text{Rb}^+ \cdots \text{O}_3\text{S}^-$ Coulomb lattice. The larger E_a value for salt **3** than for anhydrous salt **2'** is the result of the much heavier $\text{Rb}^+ \cdots \text{O}_3\text{S}^-$ lattice than the $\text{K}^+ \cdots \text{O}_3\text{S}^-$ lattice, which needs more thermal energy.

Fig. 5b shows the T - and f -dependent ϵ_1 behaviour for salt **4** (Fig. S15, ESI†). Two dielectric relaxation processes were observed in salt **4** as a broad peak at $223 < T < 350$ K and $1.0 < f < 31.6$ kHz on the low- T side (relaxation I) and a broad peak at $350 < T < 423$ K and $1.0 < f < 316$ kHz on the high- T side (relaxation II). This was similar to the behaviour of anhydrous salts **2'** and **3**, and more clearly two different relaxation processes were observed in salt **4**. These three anhydrous salts of **2'**, **3**, and **4** exhibited similar dielectric responses because they have the same molecular arrangement. The T -dependent Cole–Cole plot for salt **4** shows two overlapping semicircular traces, and the presence of two types of



dielectric relaxation (I and II) can be also observed (Fig. 5c). Peak top frequencies of f_{p1} and f_{p2} were determined for the two overlapping semicircles in the Cole–Cole plot at each T , forming two $\ln(\tau) - T^{-1}$ plots (Fig. 5d). For Relaxation I on the low- T side, a crossover from $E_a = 121$ meV at $T > 310$ K to 269 meV at $T < 310$ K was observed around 310 K. For relaxation II, a crossover from $E_a = 110$ meV at $T > 310$ K to 279 meV at $T < 310$ K is also observed around 310 K. Since the activation energies of relaxations I and II are similar to each other and the T -range of the crossover from a low to high- T region is also similar to each other, the motional freedom of dipole moments is interconnected to each other. A similar dielectric relaxation process has been observed for $(M^+)_2(\text{PCNDI}^{2-})$ salts for $M^+ = \text{Rb}^+$ ($E_{a1} = 46.8$ meV and $E_{a2} = 62.8$ meV) and Cs^+ ($E_{a1} = 89.6$ meV and $E_{a2} = 90.9$ meV), which arise from two relaxation processes involving the thermally activated $\text{Rb}^+ \cdots \text{OOC}^-$ Coulomb lattice.⁶⁸ The dielectric relaxation of salt **4** is also attributed to the similar thermal fluctuations of the electrostatic bonding $\text{Cs}^+ \cdots \text{O}_3\text{S}^-$ lattice, and the crossover process is due to the presence of $-\text{SO}_3^-$ and $\text{O}=\text{C}$ groups in different coordination environments around Cs^+ . The E_a value at the low- T side of salt **4** ($E_a = 279$ meV) is larger than that of salt **3** ($E_a = 256$ meV), which is due to the heavy crystal lattice of $\text{Cs}^+ \cdots \text{O}_3\text{S}^-$.

Thin film formation

To evaluate the film formation properties of salt **1–4**, thin films were prepared on a Si substrate with a SiO_2 layer of 200 nm using the drop-casting method, and the surface structure was observed *via* polarized optical microscopy (POM) performed in a cross-Nicole optical arrangement and *via* atomic force microscopy (AFM). POM images of the salt **1** thin film showed the formation of fibrous domain structures, whereas those of salts **2–4** thin films indicated the formation of polycrystalline thin films consisting of dense plate-like crystalline domains (Fig. S16, ESI[†]). Fig. 6 shows an AFM image of the salts **2–4** thin films. The salt **1** thin film consists of fibrous domains with heights ranging from several tens to several hundreds of nanometers, and the formation of a uniform thin film domain was not observed (Fig. S17, ESI[†]). In contrast, the formation of a flat crystalline domain was observed in the salt **2–4** thin film, with domain sizes ranging from 200 to 400 nm. In these thin films, a striped surface structure with overlapping crystalline domains was observed, and the step-terrace between the

crystalline terraces was estimated to be about 1.4–1.8 nm in height from the cross-sectional profile analysis (lower in Fig. 6). The step height corresponds to the length of the a -axis obtained in the single crystal X-ray structure, and the domains had the step-terrace structure characteristic of single crystalline domains. Salts **2–4** can be used to fabricate polycrystalline thin films consisting of single crystalline domains on Si substrates using a simple drop-casting method and exhibit excellent film-forming properties.

The correspondence between thin films fabricated on quartz substrates and crystal structures obtained from single crystal structure analysis was examined from the XRD measurements of thin films **1–4**. Fig. 7a summarizes the XRD patterns of crystals and thin films for salt **2**. The XRD pattern calculated using the single crystal X-ray structural analysis of **2** at 100 K (pattern i in Fig. 7a) shows good agreement with the XRD pattern of thin film **2**, confirming high crystallinity even in the thin film state. Diffraction peaks at $d_{001} = 15.8$ Å, $d_{002} = 7.90$ Å, and $d_{003} = 5.25$ Å were observed at $2\theta = 5.59$, 11.2, and 16.9°, respectively, corresponding to reflections corresponding to the layered structures, which are consistent with the layer period of $c = 15.7629(3)$ Å for the single crystal at 100 K. Anhydrous thin film **2'** was prepared by heating hydrated thin film **2** to 393 K. The XRD patterns of thin film **2'** were observed around $2\theta = 5.08$, 10.2, 20.5, and 25.7° at $d_{100} = 17.4$, $d_{200} = 8.67$, $d_{400} = 4.33$, and $d_{500} = 3.47$ Å. These are in good agreement with the XRD pattern obtained from the simulation of single crystal **2'** with $a = 17.381(15)$ Å. Therefore, **ESNDI**²⁻ molecules in salts **2** and **2'** are aligned perpendicular to the molecular long axis on the substrate while maintaining crystallinity during the reversible H_2O adsorption-desorption cycle.

For hydrated single crystal **1**, good agreement was also observed between the calculated XRD pattern and that of thin film **1**, where diffraction peaks around $2\theta = 6.37$ and 12.7° corresponded to $d_{002} = 13.9$ Å and $d_{004} = 6.97$ Å, respectively, (Fig. S18, ESI[†]). Therefore, the molecular long axis of **ESNDI**²⁻ was perpendicularly oriented on the substrate surface while

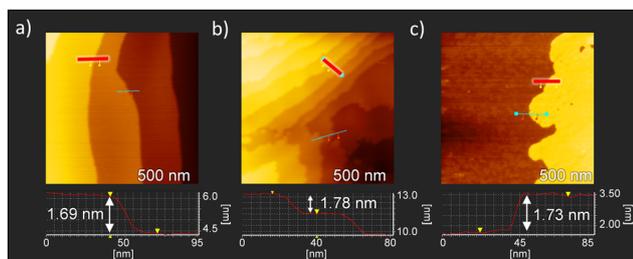


Fig. 6 AFM images of thin films for salts (a) **2**, (b) **3**, and (c) **4** with a scan area of 500×500 nm². The lower part includes the cross-sectional image along the upper red line.

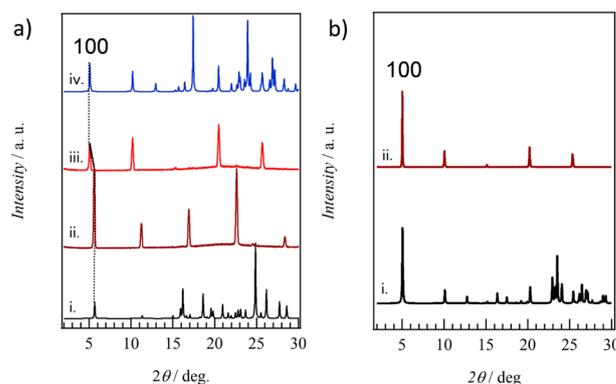


Fig. 7 XRD patterns of thin films. (a) Salts **2** and **2'** of the simulation based on the single crystal X-ray structural analysis of **2** at 100 K (i), a thin film of **2** (ii), single crystal X-ray structural analysis of **2'** at 100 K (iii), and the thin film of **2'**. (b) Salt **3** of the simulation based on single crystal X-ray structural analysis at 100 K (i) and thin film (ii).



retaining the layered structure of $c = 27.7283(9)$ Å confirmed by the single crystal X-ray structural analysis of salt **1**. The XRD pattern of anhydrous thin film **1'** dehydrated by heating at 393 K showed sharp reflections around $2\theta = 6.77$ and 13.6° , in good agreement with that of anhydrous powder **1'** dehydrated by heating. Both thin films **1** and **1'** showed clear Bragg reflections, confirming high crystallinity due to the H₂O adsorption-desorption cycle.

Fig. 7b shows the simulated XRD pattern based on the single crystal X-ray structural analysis of **3** at 100 K and thin film state **3**. Both are in good agreement, and reflections around $2\theta = 5.03, 10.1, 15.1, 20.2,$ and 25.3° correspond to $d_{100} = 17.6, d_{200} = 8.76, d_{300} = 5.87, d_{400} = 4.40,$ and $d_{500} = 3.52$ Å, respectively, for thin film **3**. This periodicity is in good agreement with the lattice constant of $a = 17.6025(4)$ Å in the crystal structural analysis of **3** at 100 K. Similar XRD patterns were also observed in thin film **4** (Fig. S19, ESI[†]), in good agreement with the XRD pattern calculated from single crystal **4** at 100 K. Periodicities of $d_{300} = 5.95, d_{400} = 4.46,$ and $d_{500} = 3.56$ Å were observed, consistent with the lattice constant of $a = 17.9643(4)$ Å based on the single crystal structural analysis of **4** at 100 K. Thin films **3** and **4** were confirmed to retain high crystallinity, consistent with the presence of crystalline domains as found *via* AFM, and were aligned perpendicular to the molecular long axis of ESNDI²⁻ on the substrate surface.

Carrier transport property of thin films

The carrier transport properties of thin films **1–4** were evaluated using FP-TRMC measurements at 298 K in the air (Fig. 8).^{81,82} The same quartz substrate employed for the XRD measurements was used to confirm the crystallinity and

structure of the thin films in the context of FP-TRMC measurements. $\phi\Sigma\mu$ values for thin films **1, 1', 2, 2', 3,** and **4** at 298 K obtained from $\phi\Sigma\mu - t$ plots were $3.7 \times 10^{-9}, 4.4 \times 10^{-9}, 7.5 \times 10^{-9}, 4.4 \times 10^{-8}, 1.5 \times 10^{-7},$ and $8.3 \times 10^{-8} \text{ m}^2 \text{ V}^{-1} \text{ s}^{-1}$, respectively. While no difference in $\phi\Sigma\mu$ values with and without H₂O is observed in thin films **1** and **1'**, a change in $\phi\Sigma\mu$ values according to the existence of H₂O is observed in thin films **2** and **2'** (Fig. 8b). $\phi\Sigma\mu$ in anhydrous thin film **2'** is about 6 times larger than in hydrated **2**, corresponding to the change in the electronic structure from 1D to 2D structure upon H₂O desorption as shown by single crystal X-ray crystal structural analyses (Fig. 3b). The change in electronic transport properties due to the H₂O adsorption-desorption cycle in thin film **2** was confirmed repeatedly, and a reversible H₂O adsorption-desorption cycle and switching behavior of electronic transport properties at the adsorption isotherm were observed. Thin film **3** exhibits the highest $\phi\Sigma\mu$ value in the series, which results from the 2D intermolecular interactions at $t_1 = 25.1$ meV and $t_2 = 21.7$ meV. The $\phi\Sigma\mu$ value of **4** is lower than that of **3** for $t_1 = 21.0$ meV and $t_2 = 18.6$ meV, which represents the 2D electronic structure but results from low absolute t values of intermolecular interactions.

Fig. 9 summarizes the $\phi\Sigma\mu$ values of thin films **1–4** and (M⁺)₂(BSNDI²⁻) at 298 K from the dimensionality of the electronic structure. The $\phi\Sigma\mu$ values of thin films increase in proportion to the intermolecular interactions of the NDI π -core and the dimensionality of the electronic structure. Salt **3** ($t_2/t_1 = 0.86$) and salt **4** ($t_2/t_1 = 0.89$) have a 2D electronic structure, showing prominently higher $\phi\Sigma\mu$ values. In particular, the $\phi\Sigma\mu$ value of thin film **3** was 2.5 times higher than that of (Na⁺)₂(BSNDI²⁻). The (M⁺)₂(BSNDI²⁻) salt was linked through a phenylene spacer unit, which is rigid and has low structural degrees of freedom, whereas (M⁺)₂(ESNDI²⁻) was linked through ethyl groups, which have high structural degrees of freedom. The latter unit enabled the closest-packing structure and the dense π -arrangement in molecular assembly. The high structural freedom for the H₂O adsorption-desorption process also changes the dimensionality of the electronic structure while maintaining a high crystalline state.

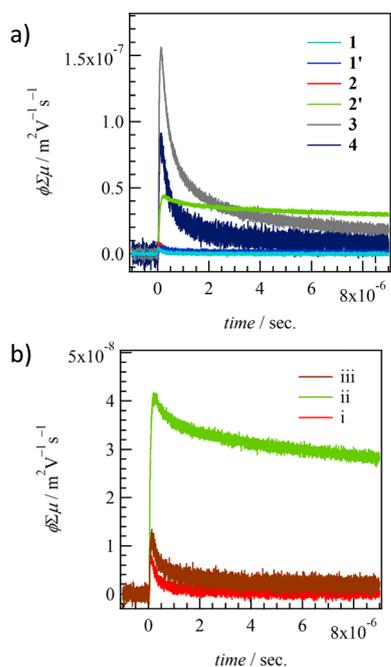


Fig. 8 TRMC kinetics of (a) thin films for **1–4** and (b) H₂O-dependent change of hydrated thin film **2** (i), anhydrous **2'** (ii), to H₂O re-hydrated **2** (iii).

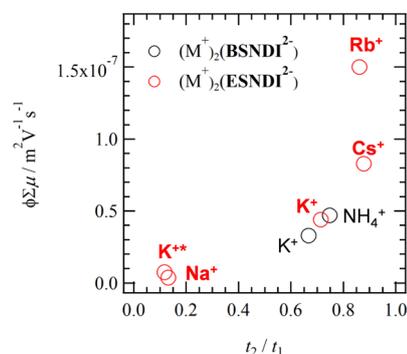


Fig. 9 $\phi\Sigma\mu - t_2/t_1$ plots of thin films **1–4** (red circle) and (M⁺)₂(ESNDI²⁻) (black circle). Points K⁺ and K⁺* correspond to thin films **2'** and **2**, respectively.



Conclusions

In ionic n-type organic semiconductors, the effects of structural flexibility of the spacer unit and the size of the counter cation (M^+) on the packing structure and physical properties were investigated in $(M^+)_2(ESNDI^{2-}) \cdot (H_2O)_n$ salts. Salts **1** ($M^+ = Na^+$) and **2** ($M^+ = K^+$) containing four molecules and one H_2O molecule, respectively, as crystal solvents were obtained, whereas anhydrous crystals were obtained as single crystals for $M^+ = K^+$, Rb^+ , and Cs^+ . Around a boundary, a single-crystal-to-single-crystal transformation with adsorption–desorption of H_2O was observed in salt **2** for $M^+ = K^+$. For the electrostatic crystal lattice consisting of $M^+ \cdots O_3S^-$, the cation sizes of Na^+ and K^+ were small and H_2O molecules were introduced to fill the crystalline space. In contrast, no crystalline space for H_2O to be introduced was confirmed in $M^+ = Rb^+$ and Cs^+ , forming a closest-packing structure with a 2D **NDI** π -arrangement. As a result, the highest transient conductivity $\phi\Sigma\mu$ was observed in the Rb^+ salt. The $\phi\Sigma\mu$ decreased in the order, salts **3** (Rb^+) > **4** (Cs^+) > **2'** (K^+) > **2** (K^+) \approx **1** (Na^+) \approx **1'** (Na^+), and reversible conductivity switching was observed for **2** and **2'**. Salts with 2D intermolecular interactions form crystalline domains on the substrate surface, which leads to a high $\phi\Sigma\mu$ value. Thermal fluctuations in the $M^+ \cdots O_3S^-$ lattice and ionic conductivity for Na^+ were observed *via* dielectric spectroscopy. The anhydrous salts **2'**, **3**, and **4** exhibited isomorphous crystal structures and showed the formation of a 2D electronic structure. Among them, salt **3** exhibited the lowest dipole fluctuations near room temperature, which achieved a condensed molecular arrangement and a high $\phi\Sigma\mu$ value. The designed structural flexibility of the spacer unit of $-C_2H_4SO_3^-$ in **ESNDI**²⁻ compared to **BSNDI**²⁻ is considered to provide more freedom for the realization of the condensed π -arrangement. A high degree of structural freedom is an important design parameter that both improves film formation and increases the packing density of the π -arrangement, which in turn affects the thermal stability and electron transport properties of n-type semiconductor materials.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by a Grant-in-Aid for Scientific Research on KAKENHI (Grant Numbers: JP20H05865, 22K19004, and 20H05862), the Japan Science and Technology Agency, Core Research for Evolutional Science and Technology (Grant Number: JPMJCR18I4), and the “Crossover Alliance to Create the Future with People, Intelligence and Material” project supported by the Ministry of Education, Culture, Sports, Science, and Technology.

Notes and references

- 1 A. I. Kitaigorodsky, *Molecular Crystals and Molecules*, Academic Press, New York, 1973.
- 2 J. D. Dunitz and A. Gavezzotti, *Chem. Soc. Rev.*, 2009, **38**, 2622–2633.

- 3 A. J. Edwards, C. F. Mackenzie, P. R. Spackman, D. Jayatilaka and M. A. Spackman, *Faraday Discuss.*, 2017, **203**, 93–112.
- 4 S. C. Abrahams, J. M. Robertson and J. G. White, *Acta Cryst.*, 1949, **2**, 233–238.
- 5 A. Camerman and J. Trotter, *Acta Cryst.*, 1965, **18**, 636–643.
- 6 D. M. Donaldson, J. M. Robertson and J. White, *Proc. R. Soc. London, Ser. A*, 1953, **220**, 366–376.
- 7 E. Fawcett and J. Trotter, *Proc. R. Soc. London, Ser. A*, 1966, **289**, 311–321.
- 8 M. D. Curtis, J. Cao and J. W. Kampf, *J. Am. Chem. Soc.*, 2004, **126**, 4318–4328.
- 9 A. Guijarro, J. A. Vergés, E. San-Fabián, G. Chiappe and H. Louis, *Chem. Phys. Chem.*, 2016, **17**, 3548–3557.
- 10 C. Wang, D. Hashizume, M. Nakano, T. Ogaki, H. Takenaka, K. Kawabata and K. Takimiya, *Chem. Sci.*, 2020, **11**, 1573–1580.
- 11 T. Akutagawa, *Mater. Chem. Front.*, 2018, **2**, 1064–1073.
- 12 T. Akutagawa, *Bull. Chem. Soc. Jpn.*, 2021, **94**, 1400–1420.
- 13 M. Stein and M. Heimsaat, *Crystals*, 2019, **9**, 665.
- 14 S. P. Thomas, P. R. Spackman, D. Jayatilaka and M. A. Spackman, *J. Chem. Theory Comput.*, 2018, **14**, 1614.
- 15 C. Rolin, E. Kang, J.-H. Lee, G. Borghs, P. Heremans and J. Genoe, *Nat. Commun.*, 2017, **8**, 14975.
- 16 P. J. Skabara, J.-B. Arlin and Y.-H. Geerts, *Adv. Mater.*, 2013, **25**, 1948–1954.
- 17 S. Giannini and J. Blumberger, *Acc. Chem. Res.*, 2022, **55**, 819–830.
- 18 N. Kasuya, J. Tsurumi, T. Okamoto, S. Watanabe and J. Takeya, *Nature Mater.*, 2021, **20**, 1401–1406.
- 19 J. C. Scott, *Semiconductor and Semimetals. Highly Conducting Quasi-One-Dimensional Organic Crystals*, ed. E. Conwell, Academic Press, New York, 1988, p. 385.
- 20 P. A. Lee, T. M. Rice and P. W. Anderson, *Phys. Rev. Lett.*, 1973, **31**, 462–465.
- 21 J. Shinagawa, Y. Kurosaki, F. Zhang, C. Parker, S. E. Brown and D. Jérôme, *Phys. Rev. Lett.*, 2007, **98**, 147002.
- 22 T. Okamoto, M. Mitani, C. P. Yu, C. Mitsui, M. Yamagishi, H. Ishii, G. Watanabe, S. Kumagai, D. Hashizume, S. Tanaka, M. Yano, T. Kushida, H. Sato, K. Sugimoto, T. Kato and J. Takeya, *J. Am. Chem. Soc.*, 2020, **142**, 14974–14984.
- 23 W. Jiang, Y. Li and Z. Wang, *Chem. Soc. Rev.*, 2013, **42**, 6113–6127.
- 24 M. Mas-Torrent and C. Rovira, *J. Mater. Chem.*, 2006, **16**, 433–436.
- 25 F. G. Brunetti, J. L. López, C. Atienza and N. Martín, *J. Mater. Chem.*, 2012, **22**, 4188–4205.
- 26 M. Bendikov and F. Wudl, *Chem. Rev.*, 2004, **104**, 4891–4945.
- 27 T. Yamamoto and K. Takimiya, *J. Am. Chem. Soc.*, 2007, **129**, 2224–2225.
- 28 H. Ebata, T. Izawa, E. Miyazaki, K. Takimiya, M. Ikeda, H. Kuwabara and T. Yui, *J. Am. Chem. Soc.*, 2007, **129**, 15732–15733.
- 29 G. Saito, T. Enoki, K. Toriumi and H. Inokuchi, *Solid State Commun.*, 1982, **42**, 557–560.
- 30 G. Saito and Y. Yukihiro, *Bull. Chem. Soc. Jpn.*, 2007, **80**, 1–137.



- 31 O. A. Melville, B. H. Lessard and T. P. Bender, *ACS Appl. Mater. Interfaces*, 2015, **7**, 13105–13118.
- 32 T. Mori, *Bull. Chem. Soc. Jpn.*, 2016, **89**, 973–986.
- 33 H. Hasegawa, T. Naito, T. Inabe, T. Akutagawa and T. Nakamura, *J. Mater. Chem.*, 1998, **8**, 1567–1570.
- 34 T. Inabe and H. Tajima, *Chem. Rev.*, 2004, **104**, 5503–5534.
- 35 C. Gaul, S. Hutsch, M. Schwarze, K. S. Schellhammer, F. Bussolotti, S. Kera, G. Cuniberti, K. Leo and F. Ortman, *Nat. Mater.*, 2018, **17**, 439–444.
- 36 Y. Zhang, I. Murtaza and H. Meng, *J. Mater. Chem. C*, 2018, **6**, 3514–3537.
- 37 C. F. J. Faul, *Acc. Chem. Res.*, 2014, **47**, 3428–3438.
- 38 D. Wu, R. Liu, W. Pisula, X. Feng and K. Müllen, *Angew. Chem., Int. Ed.*, 2011, **50**, 2791–2794.
- 39 Y. Haketa and H. Maeda, *Chem. Commun.*, 2017, **53**, 2894–2909.
- 40 S. Sugiura, T. Kubo, Y. Haketa, Y. Hori, Y. Shigeta, H. Sakai, T. Hasobe and H. Maeda, *J. Am. Chem. Soc.*, 2023, **145**, 8122–8129.
- 41 H. Tanaka, Y. Kobayashi, K. Furukawa, Y. Okayasu, S. Akine, N. Yasuda and H. Maeda, *J. Am. Chem. Soc.*, 2022, **144**, 21710–21718.
- 42 T. Akutagawa, T. Hasegawa, T. Nakamura, S. Takeda, T. Inabe, K. Sugiura, Y. Sakata and A. E. Underhill, *Chem. – Eur. J.*, 2001, **7**, 4902–4912.
- 43 T. Akutagawa, K. Shitagami, S. Nishihara, S. Takeda, T. Hasegawa, T. Nakamura, Y. Hosokoshi, K. Inoue, S. Ikeuchi, Y. Miyazaki and K. Saito, *J. Am. Chem. Soc.*, 2005, **127**, 4397–4402.
- 44 K. Sambe, N. Hoshino, T. Takeda, T. Nakamura and T. Akutagawa, *J. Phys. Chem. C*, 2020, **124**, 13560–13571.
- 45 K. Sambe, N. Hoshino, T. Takeda, T. Nakamura and T. Akutagawa, *Cryst. Growth Des.*, 2020, **20**, 3625–3634.
- 46 K. Sambe, N. Hoshino, T. Takeda, T. Nakamura and T. Akutagawa, *Cryst. Growth Des.*, 2021, **21**, 5928–5942.
- 47 H. E. Katz, A. J. Lovinger, J. Johnson, C. Kloc, T. Siegrist, W. Li, Y.-Y. Lin and A. A. Dodabalapur, *Nature*, 2000, **404**, 478–481.
- 48 J. Penneau, B. J. Stallman, P. H. Kasai and L. L. Miller, *Chem. Mater.*, 1991, **3**, 791–796.
- 49 L. Zhao, D. Zhang, Y. Zhu, S. Peng, H. Meng and W. Huang, *J. Mater. Chem. C*, 2017, **5**, 848–853.
- 50 S.-L. Suraru and F. Würthner, *Angew. Chem., Int. Ed.*, 2014, **53**, 7428–7448.
- 51 N. Sakai, J. Mareda, E. Vauthey and S. Matile, *Chem. Commun.*, 2010, **46**, 4225–4237.
- 52 M. Cui, R. Murase, Y. Shen, T. Sato, S. Koyama, K. Uchida, T. Tanabe, S. Takaishi, M. Yamashita and H. Iguchi, *Chem. Sci.*, 2022, **13**, 4902–4908.
- 53 T. Tanabe, T. Sato, K. Fuku, S. Takaishi and H. Iguchi, *ChemPlusChem*, 2023, e202300140.
- 54 T. Ono, Y. Tsukiyama, S. Hatanaka, Y. Sakatsume, T. Ogoshi and Y. Hisaeda, *J. Mater. Chem. C*, 2019, **7**, 9726–9734.
- 55 Y. Takashima, V. M. Martínez, S. Furukawa, M. Kondo, S. Shimomura, H. Uehara, M. Nakahama, K. Sugimoto and S. Kitagawa, *Nat. Commun.*, 2011, **2**, 168.
- 56 A. Mizuno, Y. Shuku, M. M. Matsushita, M. Tsuchiizu, Y. Hara, N. Wada, Y. Shimizu and K. Awaga, *Phys. Rev. Lett.*, 2017, **119**, 057201.
- 57 A. Mizuno, Y. Shuku, R. Suizu, M. Tsuchiizu and K. Awaga, *CrystEngCom.*, 2021, **23**, 5053–5059.
- 58 D. Shukla, S. Nelson, D. C. Freeman, M. Rajeswaran, W. G. Ahearn, D. M. Meyer and J. T. Carey, *Chem. Mater.*, 2008, **24**, 7486–7491.
- 59 M.-M. Ling, P. E. M. Gomez, M. Koenemann, J. Locklin and Z. Bao, *Adv. Mater.*, 2007, **9**, 1123.
- 60 A. Nowak-Król and F. Würthner, *Org. Chem. Front.*, 2019, **6**, 1272–1318.
- 61 O. Ogi, V. Stepanenko, K. Sugiyasu, M. Takeuchi and F. Würthner, *J. Am. Chem. Soc.*, 2015, **137**, 3300–3307.
- 62 G. S. Perez, S. Dasgupta, W. Żuraw, R. F. Pineda, K. Wojciechowski, L. K. Jagadamma, I. Samuel and N. Robertson, *J. Mater. Chem. A*, 2022, **10**, 11046–11053.
- 63 R. S. Wilson-Kovacs, X. Fang, M. J. L. Hagemann, H. E. Symons and C. F. J. Faul, *Chem. – Eur. J.*, 2022, **28**, e202103443.
- 64 A. Kawasaki, T. Takeda, N. Hoshino, W. Matsuda, S. Seki and T. Akutagawa, *J. Phys. Chem. C*, 2019, **123**, 15451–15457.
- 65 A. Kawasaki, T. Takeda, N. Hoshino, W. Matsuda, S. Seki and T. Akutagawa, *Cryst. Growth Des.*, 2020, **20**, 1276–1284.
- 66 A. Kawasaki, T. Takeda, N. Hoshino, W. Matsuda, S. Seki, G. K. H. Shimizu and T. Akutagawa, *ACS Appl. Mater. Interfaces*, 2023, **15**, 1661–1674.
- 67 H. Abe, A. Kawasaki, T. Takeda, N. Hoshino, W. Matsuda, S. Seki and T. Akutagawa, *ACS Appl. Mater. Interfaces*, 2020, **12**, 37391–37399.
- 68 H. Abe, A. Kawasaki, T. Takeda, N. Hoshino, W. Matsuda, S. Seki and T. Akutagawa, *J. Am. Chem. Soc.*, 2021, **143**, 1046–1060.
- 69 R. Ide, A. Kawasaki, T. Takeda, S. Dekura, N. Hoshino, W. Matsuda, S. Seki and T. Akutagawa, *J. Phys. Chem. C*, 2023, **127**, 16709–16720.
- 70 R. O. Macron, J. G. dos Santos, K. M. Figueiredo and S. Brochsztein, *Langmuir*, 2006, **22**, 1680–1687.
- 71 H. Ke, X.-M. Zhu, S.-M. Xie, P.-X. Ming and J. Z. Liao, *Inorg. Chem. Front.*, 2022, **9**, 2568–2574.
- 72 X.-M. Zhu, Y. Jiang, S.-M. Xie, P.-X. Ming, C.-H. Zhang, H. Ke and J.-Z. Liao, *Dyes Pigm.*, 2022, **207**, 110747.
- 73 K. Ke, J.-Z. Liao and C.-Z. Lu, *Acta Cryst.*, 2019, **C75**, 1128–1133.
- 74 J.-Z. Liao, J.-F. Chang, L. Meng, H.-L. Zhang, S.-S. Wang and C. Z. Lu, *Chem. Commun.*, 2017, **53**, 9701–9704.
- 75 J.-Z. Liao, S.-S. Wang, X.-Y. Wu, R. Yu, C.-Z. Lu and X.-L. Chen, *Dalton Trans.*, 2018, **47**, 1027–1031.
- 76 Crystal Structure: Single Crystal Structure Analysis Software. Ver. 4.30, 2018. Rigaku Corporation and Molecular Structure Corporation.
- 77 G. M. Sheldrick, SHELX2014 Programs for Crystal Structure Analysis; Universität Göttingen: Göttingen, Germany, 2014.
- 78 T. Mori, A. Kobayashi, Y. Sakaki, H. Kobayashi, G. Saito and H. Inokuchi, *Bull. Chem. Soc. Jpn.*, 1984, **57**, 627–633.



- 79 G. Greeuw and J. F. Verwey, *J. Appl. Phys.*, 1984, **56**, 2218–2224.
- 80 N. Yoshinari, S. Yamashita, Y. Fukada, Y. Nakazawa and T. Konno, *Chem. Sci.*, 2018, **10**, 587–593.
- 81 S. Seki, A. Saeki, T. Sakurai and D. Sakamaki, *Phys. Chem. Chem. Phys.*, 2014, **16**, 11093–11113.
- 82 Y. Tsutsui, G. Schweicher, B. Chattopadhyay, T. Sakurai, J.-B. Arlin, C. Ruzié, A. Aliev, A. Ciesielski, S. Colella, A. R. Kennedy, V. Lemaire, Y. Olivier, R. Hadji, L. Sanguinet, F. Castet, D. Beljonne, J. Cornil, P. Samorì, S. Seki and Y. H. Geerts, *Adv. Matter.*, 2016, **28**, 7106–7114.

