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## Enhancing carbon dioxide capture under humid conditions by optimizing the pore surface structure†

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Metal–organic frameworks (MOFs) exhibit significant potential for mitigating carbon emissions due to their high porosity and tunability. Despite numerous reports on CO<sub>2</sub> capture by MOF sorbents, a common challenge is their poor selectivity for CO<sub>2</sub> over water. Moreover, in-depth studies are much needed to elucidate the relationships among the pore surface structure, hydrophobicity, and CO<sub>2</sub> uptake capacity/selectivity. In this work, we investigate the factors influencing CO<sub>2</sub> adsorption capacity and selectivity under humidity in a series of isorecticular pillar-layer structures, Ni<sub>2</sub>(L)<sub>2</sub>(dabco) (L = bdc, ndc, adc). Our study shows that increasing ligand conjugation not only results in increased hydrophobicity, decreased pore size and BET surface area, but also leads to the change of primary binding sites of water molecules and higher binding energy of CO<sub>2</sub>, all of which contribute to largely increased CO<sub>2</sub> uptake capacity under humid conditions. Additionally, increasing ligand conjugation and consequently hydrophobicity slow down and reduce competitive water adsorption drastically. Notably, the MOF made of ligand with the highest conjugation, Ni<sub>2</sub>(adc)<sub>2</sub>(dabco), exhibits significantly enhanced CO<sub>2</sub> adsorption in N<sub>2</sub>/CO<sub>2</sub> binary mixtures under relatively high humidity (50% RH), with an increase of ~31% and ~36% for the composition of 15/85 and 50/50, respectively, compared to dry conditions. An experimental FTIR study and DFT theoretical calculations confirm that H<sub>2</sub>O occupies different primary binding site in Ni<sub>2</sub>(bdc)<sub>2</sub>(dabco) and Ni<sub>2</sub>(adc)<sub>2</sub>(dabco), and under humid conditions a higher binding energy of CO<sub>2</sub> is achieved with preferential H<sub>2</sub>O/CO<sub>2</sub> co-adsorption in Ni<sub>2</sub>(adc)<sub>2</sub>(dabco), potentially creating additional adsorption sites for CO<sub>2</sub>.

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## Introduction

Global warming has pushed us to the edge of an epoch of frequent unpredictable natural disaster occurrences. In order to limit global warming to 1.5 °C, it is essential for greenhouse gas (GHG) emissions to reach their peak no later than 2024 and decrease by 43% by 2030.<sup>1,2</sup> Managing emissions and cutting down the existing GHG *via* negative emissions technologies

(NETs) are of vital importance besides developing renewable energy.<sup>3,4</sup> Conventional NETs predominantly rely on aqueous phase adsorption of CO<sub>2</sub> *via* amine or inorganic basic solutions, which lack long-term sustainability due to their highly corrosive nature, bulky volume occupancy, and intensive regeneration energy consumption. A more efficient and low-cost alternative NET is currently being extensively explored by the entire scientific community.<sup>5</sup>

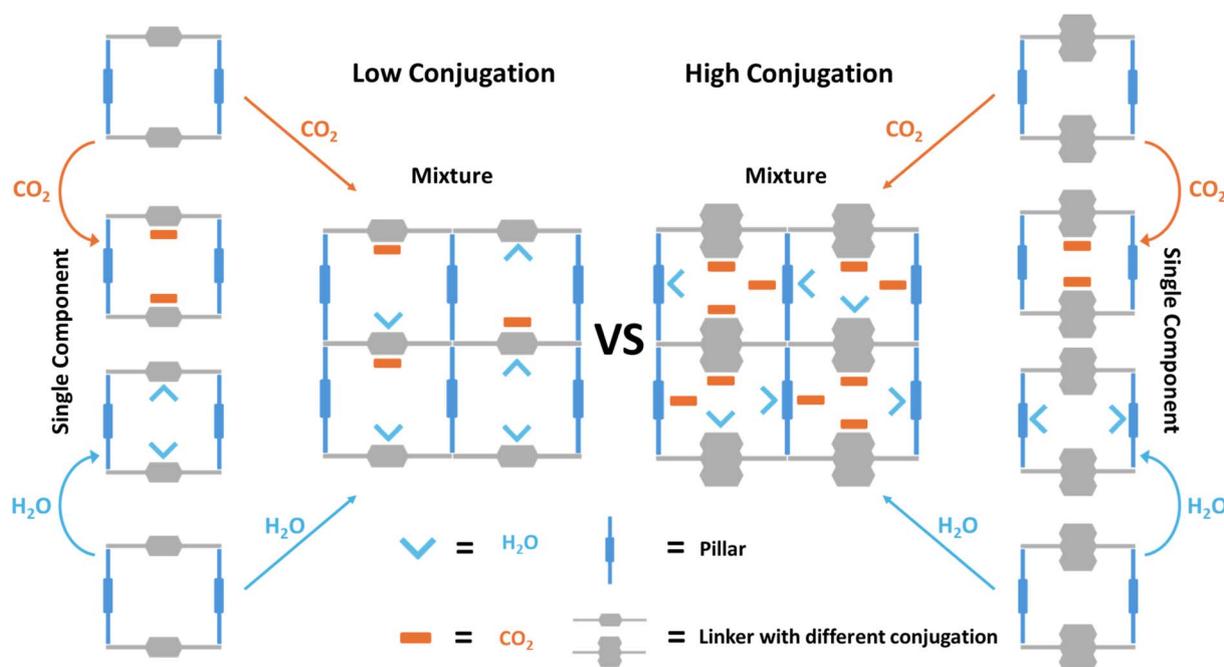
Acclaimed as “magic sponges” for cleaning air, porous solid sorbents have been considered as the solution for the next generation of adsorption-based NETs.<sup>6</sup> Metal–organic frameworks (MOFs), composed of metal ions or clusters and various organic ligands, represent solid sorbents with immense potential for practical applications to slow down global warming.<sup>5,7–13</sup> Having pore sizes reaching down to the angstrom level, these crystalline microporous materials exhibit expansive inner surface areas, high selectivities, systematic structural tunability, and lower energy consumption for regeneration.<sup>14,15</sup> Although numerous MOFs have achieved high CO<sub>2</sub> uptake capacity under certain conditions, further improvement in the selective adsorption and capture of CO<sub>2</sub> from gas mixtures is much needed, especially under moisture and/or dilution

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concentrations.<sup>16–18</sup> In contrast to its smaller size, the higher polarizability and quadrupole moment of CO<sub>2</sub> (3.3 Å) enable its effective separation from competitive components such as N<sub>2</sub> (3.64 Å) and O<sub>2</sub> (3.46 Å) from their gas mixtures *via* pore size and sorbent-gas interaction control.<sup>19–21</sup> The real challenge lies in achieving selective adsorption of CO<sub>2</sub> over H<sub>2</sub>O (2.64 Å) under humid conditions while keeping the MOF structure intact; especially in such cases H<sub>2</sub>O and CO<sub>2</sub> usually compete at the same binding sites. Owing to their much smaller size, polar water molecules can readily penetrate into microporous sorbents and form a hydrogen-bonded network inside the pore space, which further accelerates their accumulation.<sup>18</sup> Moreover, water vapor can competitively break the original coordination bonds between ligands and metals and bind to the open metal sites under a high humidity environment,<sup>22,23</sup> which often causes serious degradation of the MOF structure and destroy the porosity.<sup>19</sup> Water molecules with higher binding energy often pre-occupy the competing binding sites in the MOF structure during the adsorption of a gas mixture and suppress the uptake of CO<sub>2</sub>, which refers to “water poisoning”.<sup>24</sup> Notwithstanding a comparatively weaker host-guest interaction, physisorption can be one of the possible solutions to surpass water interference to attain a notable level of adsorption capacity through adeptly tailoring the pore size to an appropriate range and modifying their pore surface structures.<sup>17,18</sup> Furthermore, the regeneration energy can be significantly lower. An excellent example is the recently reported MOF adsorbent, CALF-20,<sup>18</sup> featuring ultra-small pores and physisorption of CO<sub>2</sub>. Due to its exceptional performance in CO<sub>2</sub> capture under both dry and humid conditions, CALF-20 has already entered the initial phase for commercialization. However, achieving a small pore size alone may not be sufficient

to provide a complete solution as it does not necessarily address the competitive adsorption of water molecules. Therefore, an adsorption pathway that favors CO<sub>2</sub> adsorption over H<sub>2</sub>O under humid conditions and a comprehensive understanding of such processes is yet to be developed.

One possible approach to diminish water interference while enhancing CO<sub>2</sub> uptake is to modify the pore surface structure by changing the properties of the binding sites, which may alter the competitiveness of the competing molecules. In 2016, Ding *et al.* demonstrated a strategy for boosting CO<sub>2</sub> uptake by partitioning the channels of MOF-5 with *in situ* polymerization of aromatic acetylenes to simultaneously decrease pore size and increase hydrophobicity.<sup>25</sup> Inspired by this work, we target to systematically study a series of isoreticular pillar-layer type MOFs, Ni<sub>2</sub>(L)<sub>2</sub>(dabco). By alternating the ligand conjugation, we show that the pore size, surface area, hydrophobicity, and pore surface structure can all be optimized concurrently to enhance CO<sub>2</sub> uptake under competitive water co-adsorption. The use of nonpolar aromatic groups is a straightforward and effective method to tune the hydrophobicity and pore surface structure of MOFs. As depicted in Scheme 1, changing the ligand conjugation can not only efficiently optimize the pore surface structure but also lead to the decrease of competitive adsorption of CO<sub>2</sub> and H<sub>2</sub>O, as well as giving rise to new binding sites for CO<sub>2</sub>. The pillar-layer type MOF presented in Scheme 1 is made of two ligands, a linker and a pillar. In the low-conjugation pore environment (Scheme 1, left), CO<sub>2</sub> and H<sub>2</sub>O each preferentially bind to the linker site. However, in the binary mixture, they compete for the same site, resulting in the suppression of CO<sub>2</sub> uptake since H<sub>2</sub>O binds more strongly with the linker. On the other hand, in the case of a high-conjugation pore environment (Scheme 1, right), the preferential binding site of water is



Scheme 1 Schematic showing the ligand conjugation effect on the pore surface structure and competitive binding sites of CO<sub>2</sub> and H<sub>2</sub>O.



shifted to the more hydrophilic pillar ligand due to significantly reduced interaction between H<sub>2</sub>O and the linker ligand, thereby freeing this site for CO<sub>2</sub>. In the binary mixture, the competition between H<sub>2</sub>O and CO<sub>2</sub> will be significantly reduced owing to the formation of energy-favorable H<sub>2</sub>O/CO<sub>2</sub> co-adsorption, which also potentially creates new binding sites for CO<sub>2</sub> to enhance its uptake capacity under humid conditions.

We have successfully synthesized three members of the Ni<sub>2</sub>(L)<sub>2</sub>(dabco) MOFs. All of them are made of the same pillar ligand (1,4-diazabicyclo[2,2,2]octane, dabco) but different linker ligands within the layer with paddle-wheel Ni<sub>2</sub> dimers. The dicarboxylate linker ligands vary from one benzene ring (1,4-benzenedicarboxylic acid, H<sub>2</sub>bdc), to two benzene rings (1,4-naphthalenedicarboxylic acid, H<sub>2</sub>ndc) and three benzene rings (9,10-anthracenedicarboxylic acid, H<sub>2</sub>adc). All three isorecticular MOFs, namely Ni<sub>2</sub>(adc)<sub>2</sub>(dabco) (1), Ni<sub>2</sub>(ndc)<sub>2</sub>(dabco) (2), and Ni<sub>2</sub>(bdc)<sub>2</sub>(dabco) (3), are structurally characterized. Their pore size and surface area decrease with the increasing size and conjugation of the ligands, and their corresponding hydrophobicity is in the order of 3 < 2 < 1. Yet their CO<sub>2</sub> uptake demonstrates a reverse trend of 1 > 2 > 3. Compound 1 reaches a CO<sub>2</sub> uptake of 3.03 mmol g<sup>-1</sup> at 298 K and 101.3 kPa, which is 1.75 times that of compound 3. Moreover, the decreased adsorption amount and rate of water from 3 to 1 show the effectiveness of changing the ligand conjugation to increase the hydrophobicity of the framework. Compound 1 exhibits a very interesting and rare adsorption behavior, with its CO<sub>2</sub> uptake significantly increased under humid conditions, as confirmed by both dynamic adsorption and column breakthrough experiments. *In situ* FTIR measurements and *ab initio* calculations

have verified the change of water binding sites and H<sub>2</sub>O/CO<sub>2</sub> co-adsorption, leading to partial replacement of H<sub>2</sub>O by CO<sub>2</sub> in Ni<sub>2</sub>(adc)<sub>2</sub>(dabco) under humid conditions. The current study is a neat example of tuning the pore surface structure to minimize competitive adsorption and to enhance CO<sub>2</sub> uptake under humidity.

## Results and discussion

### Characterization of the structure and physical properties

The isorecticular Ni<sub>2</sub>(L)<sub>2</sub>(dabco) structural series is formed by a paddle-wheel Ni<sub>2</sub>(COO)<sub>4</sub> node, a neutral N-donor pillar ligand dabco, and a dicarboxylate ligand L with different hydrophobicity (L = bdc, ndc, adc) (Fig. 1a). The pillar-layer structures feature a three-dimensional (3D) network with one-dimensional (1D) channels through *a*- and *c*-axes (Fig. 1b and c). As the micrometer-sized crystals of compound 1 were too small to obtain a satisfactory structure solution by the single crystal XRD method, we synthesized isostructural Zn<sub>2</sub>(adc)<sub>2</sub>(dabco) single crystals under similar conditions and solved its crystal structure (Fig. S2 and Table S1†). The well matched PXRD patterns of the simulated Zn<sub>2</sub>(adc)<sub>2</sub>(dabco) and Ni<sub>2</sub>(L)<sub>2</sub>(dabco) series confirmed their isostructural nature (Fig. 2a). In addition to Ni and Zn, other isostructural M<sub>2</sub>(adc)<sub>2</sub>(dabco) (M = Mg, Ca, Cu, Co) were also successfully synthesized (Fig. S3†). The Ni<sub>2</sub>(L)<sub>2</sub>(dabco) series was chosen for further study since Ni<sub>2</sub>(adc)<sub>2</sub>(dabco) outperforms the other isostructural compounds in CO<sub>2</sub> uptake (Fig. S5a†). Structure analysis revealed that compounds 1–3 crystallize in the tetragonal crystal system, with the space group *I4/mcm*. Each Ni<sup>2+</sup> is 5-coordinated by four oxygen atoms

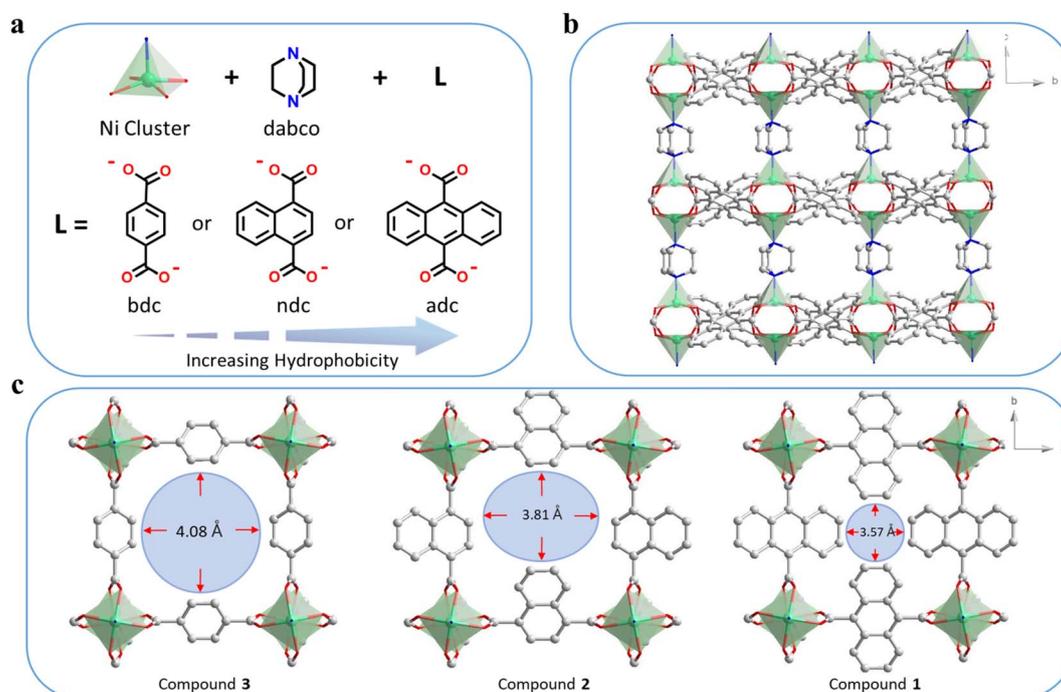


Fig. 1 (a) The chemical components of the Ni<sub>2</sub>(L)<sub>2</sub>(dabco) series; (b) 3D pillar-layer structure of compound 1 viewed along the *a*-axis; (c) 1D channels along the *c*-axis with decreased pore diameters from compounds 3 to 1. Color code: C – grey; O – red; N – blue; Ni – green.



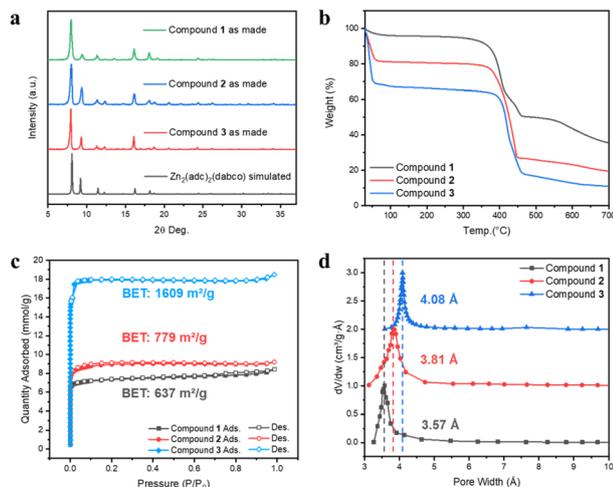


Fig. 2 (a) PXRD patterns of isoreticular Ni<sub>2</sub>(L)<sub>2</sub>(dabco) MOFs and simulated pattern of Zn<sub>2</sub>(adc)<sub>2</sub>(dabco); (b) the TG plots of the Ni<sub>2</sub>(L)<sub>2</sub>(dabco) series; (c) N<sub>2</sub> adsorption-desorption isotherms at 77 K; and (d) pore size distribution (H-K function) of the Ni<sub>2</sub>(L)<sub>2</sub>(dabco) series.

from four different carboxylate groups and one N-donor which forms a square pyramidal coordination environment. Thermogravimetric (TG) analysis of compounds **1**, **2** and **3** after solvent exchange in methanol for two days showed that the initial weight losses were 4.1%, 19.2%, and 32.7%, respectively, attributed to the solvent removal prior to 100 °C. All three samples remained stable up to ~400 °C (Fig. 2b). The permanent porosity of the Ni<sub>2</sub>(L)<sub>2</sub>(dabco) series was estimated using N<sub>2</sub> adsorption data collected at 77 K (Fig. 2c). The Brunauer-Emmett-Teller (BET) surface areas are 637 m<sup>2</sup> g<sup>-1</sup>, 779 m<sup>2</sup> g<sup>-1</sup>,

and 1609 m<sup>2</sup> g<sup>-1</sup>, for compounds **1**, **2** and **3**, respectively. The pore size distributions based on the Horvath-Kawazoe (H-K) function are 3.57, 3.81 and 4.08 Å for compounds **1**, **2** and **3**, respectively, demonstrating a monotonic increase (Fig. 2d). This is well expected, as the surface area and pore size increase/decrease as the size of ligand decreases/increases.

### Single component CO<sub>2</sub> adsorption analysis

Single-component CO<sub>2</sub> adsorption isotherms were obtained at different temperatures to evaluate the adsorption performance of compounds **1**–**3**. Before measurements, solvent exchange was carried out on each sample in MeOH for two days, with the supernatant replaced by fresh MeOH every 8 hours. Activation was subsequently done under dynamic vacuum at 393 K for 3 hours. The results showed a reverse trend between the uptake amount and BET surface area: compound **1** has the lowest BET surface area but the highest CO<sub>2</sub> uptake over the entire pressure range from 0 to 101 kPa. It reaches an adsorption capacity of 3.03 mmol g<sup>-1</sup> (67.9 cm<sup>3</sup> g<sup>-1</sup> STP) at 298 K and 101.3 kPa, while compound **2** and compound **3** can only take up 2.09 mmol g<sup>-1</sup> and 1.73 mmol g<sup>-1</sup>, respectively, under the same conditions (Fig. 3a). Despite its smaller BET surface area (39.6% compound **3**), the CO<sub>2</sub> loading of compound **1** is 1.75 times that of compound **3** at 101 kPa. Moreover, the CO<sub>2</sub> adsorption profile of compound **1** shows a much steeper slope in the low-pressure area (<15 kPa), and the uptake capacity is 1.12 mmol g<sup>-1</sup> at 15 kPa, 4.7 times that of compound **3** (0.24 mmol g<sup>-1</sup>). In addition, the uptake capacity of compound **1** remains significantly higher than those of others at different temperatures (Fig. 3b–d). It also outperformed all other members of the isostructural M<sub>2</sub>(adc)<sub>2</sub>(dabco) series (M = Zn, Cu, Co, Mg, Ca) (Fig. S5a†). Based

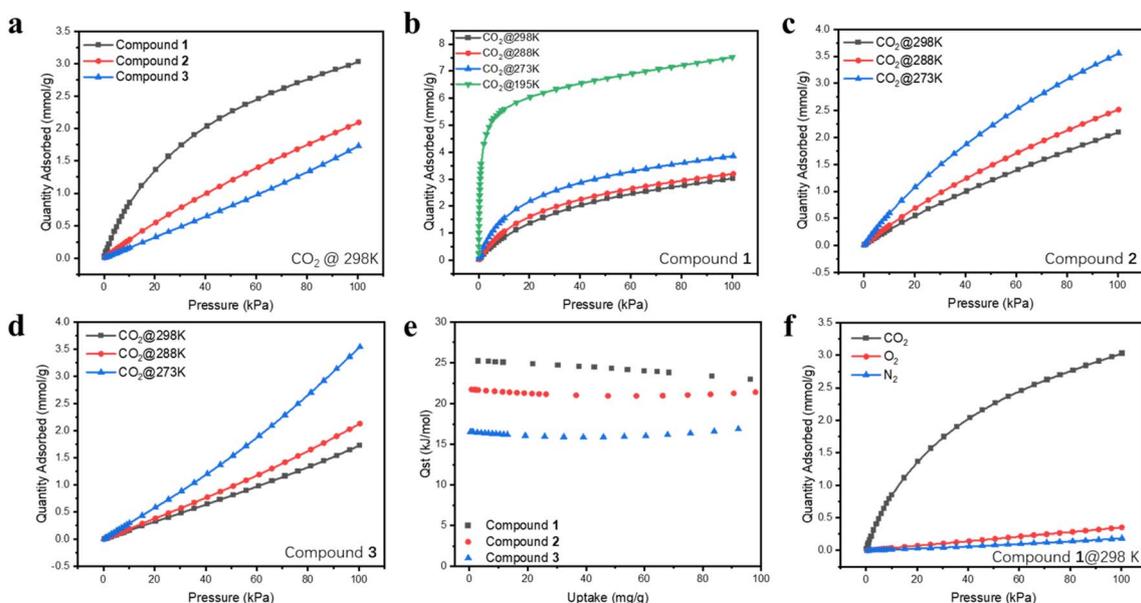


Fig. 3 The single-component CO<sub>2</sub> adsorption isotherms (up to 1 bar) of (a) compounds **1**–**3** at 298 K; (b) compound **1** at 195 K, 273 K, 288 K, and 298 K; (c) compound **2** and (d) compound **3** at 273 K, 288 K, and 298 K; (e) the isosteric heats ( $Q_{st}$ ) of CO<sub>2</sub> adsorption of compounds **1**–**3** calculated from isotherms collected at 273, 288, and 298 K; (f) single component adsorption isotherms of N<sub>2</sub> and O<sub>2</sub> at 298 K and up to 1 bar on compound **1** in comparison with CO<sub>2</sub>.



on these data, it is clear that smaller pore size and an increased degree of conjugation within the pore environment give rise to stronger interactions between CO<sub>2</sub> and the MOF host, as reflected by the  $Q_{st}$  values (Fig. 3e). This phenomenon is particularly notable in the low partial pressure area. At zero loading, compound 1 has the highest isosteric heat (25.2 kJ mol<sup>-1</sup>), compared to 21.7 kJ mol<sup>-1</sup> and 16.5 kJ mol<sup>-1</sup> for compounds 2 and 3, respectively (Fig. 3e). Furthermore, the CO<sub>2</sub> binding energies from the theoretical calculations, 42.61 and 30.27 kJ mol<sup>-1</sup> for compounds 1 and 3, are fully consistent with the experimental data, as discussed below. Considering other gas components in the air mixture, compound 1 adsorbs a very small amount of N<sub>2</sub> and O<sub>2</sub> at 298 K (Fig. 3f). The calculated IAST selectivity of CO<sub>2</sub>/N<sub>2</sub> (15 : 85) for compound 1 gives a very high value at zero loading (Fig. S14†).

### CO<sub>2</sub> adsorption under humid conditions

To assess the effect of water vapor on the MOF sorbents and the CO<sub>2</sub> adsorption performance under humid conditions, dynamic adsorption isotherms of water vapor and CO<sub>2</sub>/N<sub>2</sub> binary mixtures were collected in a gravimetric adsorption analyzer by exposing the samples to variable relative humidity (RH) and monitoring their weight changes. The results show that after optimizing the pore surface structure by increasing the conjugation of ligands, water molecules were significantly impeded from entering the pores. Adsorption kinetics profiles clearly show that the adsorption was slow in all three cases, and the adsorbed amount of water follows the trend: 3 > 2 > 1 throughout the entire RH range being tested (Fig. 4a). The static vapor adsorption isotherms show that the water adsorption capacity in compounds 1–3 remains relatively low under 15 mbar (~50% RH). Above this pressure both compounds 2 and 3 illustrate a sharp surge far surpassing

compound 1 (Fig. S5b†). Compound 1 has a very low equilibrium adsorption capacity of water vapor compared to some top-performing CO<sub>2</sub> sorbents (Fig. 4b and d).<sup>17,18,26–33</sup> A more detailed comparison of its performance with other prototype MOFs is provided in Table S2.† Compound 1 takes up the highest amount of CO<sub>2</sub> under similar RH from CO<sub>2</sub>/N<sub>2</sub> mixtures, demonstrating its outstanding performance under humidity. Note that the rate of CO<sub>2</sub> adsorption was much faster than that of H<sub>2</sub>O, as shown in Fig. 4c for binary CO<sub>2</sub>/N<sub>2</sub> mixtures (15/85 and 50/50) at 303 K (30 °C). For CO<sub>2</sub>, the adsorption equilibrium was reached within several minutes, while for H<sub>2</sub>O, only ~12% of its maximum capacity was reached at 30 minutes (Fig. 4c). More interestingly and strikingly, the uptake of CO<sub>2</sub> under humid conditions (50% RH) increased by a large amount, 30.6% and 36.1% for both CO<sub>2</sub>/N<sub>2</sub> mixtures (15/85 and 50/50) compared to dry conditions (Fig. 4c), which was further confirmed by the increased retention time from column breakthrough experiments. In addition, the sample exhibits high stability. No obvious capacity loss was observed after six-consecutive adsorption cycles (Fig. S15†), and its PXRD pattern remained nearly identical to those of the as-made and simulated patterns upon exposure to open air for 1.5 months (Fig. S15 and S16†), further confirming the robustness and recyclability of this compound.

### Separation of CO<sub>2</sub>/N<sub>2</sub> binary mixtures by column breakthrough experiments

The column breakthrough experiments were carried out to assess the selectivity of CO<sub>2</sub> over N<sub>2</sub> under both dry and wet conditions to mimic practical separation processes and to evaluate the CO<sub>2</sub> adsorption performance under competitive adsorption of water molecules. In all cases, N<sub>2</sub> gas was eluted out quickly after a delay between 50 and 125 s g<sup>-1</sup>, while CO<sub>2</sub> was retained in the column for various time periods after the departure of N<sub>2</sub>. The CO<sub>2</sub> retention times under dry conditions are 500 s g<sup>-1</sup>, 170 s g<sup>-1</sup>, and 110 s g<sup>-1</sup>, for compounds 1, 2 and 3, respectively (Fig. 5a–c). These results are consistent with their single component CO<sub>2</sub> adsorption behaviors. On the other hand, under wet conditions (at 50% RH), the retention time of compound 3 decreases slightly, while it increases significantly for compound 1, about 20% compared to that under dry conditions (Fig. 5a). The decrease in the retention time for compound 3 is likely due to the competitive adsorption of water at the same binding site and the increase in the retention time for compound 1 is fully consistent with the observed ~30.6% increase in the CO<sub>2</sub> uptake amount in Fig. 4c under the same RH. In order to understand this desirable yet rare behavior, we further carried out *in situ* FT-IR spectroscopic experiments and *ab initio* calculations to elucidate the role and the binding site of water molecules during the co-adsorption process and how it facilitated the enhanced CO<sub>2</sub> adsorption in compound 1.

### *In situ* FTIR study

To reveal the competitive and co-adsorption of CO<sub>2</sub> and H<sub>2</sub>O in compound 1 under humid conditions, we further conducted *in situ* infrared (IR) spectroscopic experiments. After sample activation, CO<sub>2</sub> alone was measured first in a dry sample. Given that

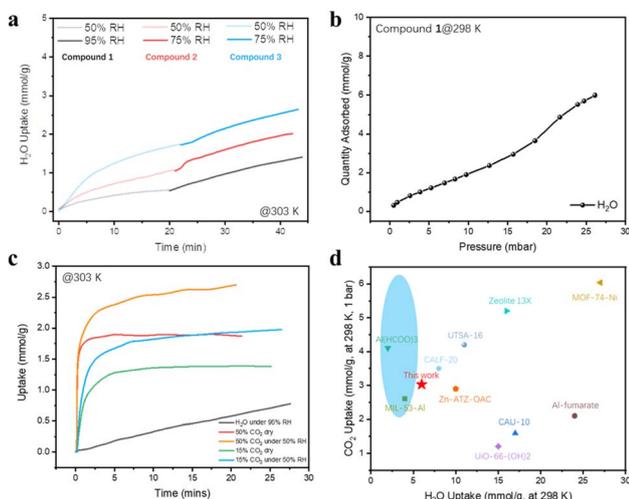


Fig. 4 (a) Dynamic adsorption kinetics of water vapor in compounds 1–3 at 303 K; (b) adsorption isotherm of water vapor at 298 K; (c) adsorption kinetics of CO<sub>2</sub> (in 15/85 and 50/50 binary mixtures of CO<sub>2</sub>/N<sub>2</sub> under dry and 50% RH conditions) and H<sub>2</sub>O (95% RH) in compound 1 at 303 K; (d) comparisons of single component CO<sub>2</sub> uptake (1 bar, 298 K) and H<sub>2</sub>O vapor uptake at 298 K of compound 1 with those of previously reported best-performing MOFs.



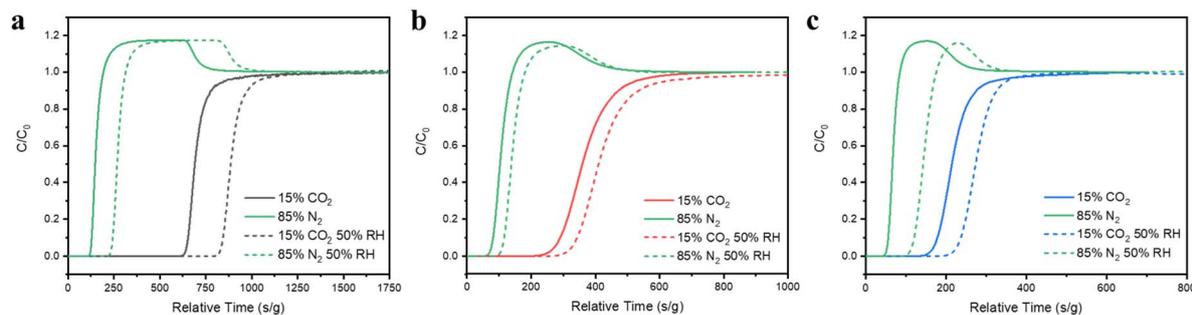


Fig. 5 Experimental column breakthrough curves of compound (a) 1, (b) 2, and (c) 3 for CO<sub>2</sub>/N<sub>2</sub> (15/85) binary mixtures under dry (solid line) and wet (dashed line) conditions at 1.0 bar and 298 K.

the gas phase CO<sub>2</sub> spectrum was prohibitively high (no signal on the detector) above 10 torr and thus it was impossible to directly observe the adsorbed CO<sub>2</sub> in the presence of a gas phase, the gas CO<sub>2</sub> was evacuated by pumping the cell and the spectra were immediately collected within 5 seconds of evacuation once the pressure of CO<sub>2</sub> drops below 500 mTorr (negligible gas-phase IR absorption). The adsorbed CO<sub>2</sub> was detected by its characteristic  $\nu_{\text{as}}(\text{CO}_2)$  band at 2334 cm<sup>-1</sup> as shown in Fig. 6a.<sup>34</sup> After fully evacuating CO<sub>2</sub> under vacuum, the sample was then exposed to water moisture at 11 torr, equivalent to 50% RH at 24 °C. The adsorbed H<sub>2</sub>O was typified by its stretching ( $\nu$ ) and bending ( $\beta$ ) bands, respectively (see Fig. 6b bottom and Fig. S6†). A closer inspection reveals the  $\beta(\text{H}_2\text{O})$  band shows two components at 1637 and 1625 cm<sup>-1</sup>, which correspond to two types of H<sub>2</sub>O in compound 1 as later verified by theoretical calculations. The two blue-shift differently in reference to gas phase H<sub>2</sub>O at 1595 cm<sup>-1</sup> due to their different strength of H bonding interaction with the sample.<sup>35–37</sup> The larger shift points to a stronger interaction, whereas the smaller one indicates a weaker interaction with the structure.<sup>35,36</sup> Along with water adsorption, we see the perturbation of phonon

modes of the MOF, *e.g.*, stretching bands of the COO<sup>-</sup> group in the region 1600–1400 cm<sup>-1</sup>,<sup>10,38</sup> as indicated by the derivative-like feature, which is caused by water inclusion.<sup>10</sup> Keeping water moisture inside, we further introduced CO<sub>2</sub>. The time dependent spectra were recorded to monitor the impact of CO<sub>2</sub> on pre-adsorbed H<sub>2</sub>O. As shown in Fig. 6b, the  $\beta(\text{H}_2\text{O})$  band diminishes significantly by around 50% after exposure to CO<sub>2</sub> for 15 min, indicating that half of pre-adsorbed water was displaced by CO<sub>2</sub> which indicates an energy-favorable process. The loss mainly occurred on the 1625 cm<sup>-1</sup> component, suggesting the higher susceptibility of weakly bound H<sub>2</sub>O to be displaced by incoming CO<sub>2</sub>. From this displacement phenomenon, we can conclude that the thermodynamically favorable co-adsorption configurations of H<sub>2</sub>O and CO<sub>2</sub> play a dominant role when the low concentration of water moisture exists. The final adsorption of CO<sub>2</sub> was further measured through monitoring the  $\nu_{\text{as}}(\text{CO}_2)$  band after evacuating gas phase CO<sub>2</sub>. Interestingly, the  $\nu_{\text{as}}(\text{CO}_2)$  band shows a slight blue shift (+2 cm<sup>-1</sup>) compared with the value observed in the dry sample (Fig. 6a). Given the sensitivity of the  $\nu_{\text{as}}(\text{CO}_2)$  position to its chemical environment, we infer that the binding/interaction of CO<sub>2</sub> within compound 1 is enhanced due to the presence of co-adsorbed H<sub>2</sub>O, which further assists CO<sub>2</sub> to bind onto the sites that are not accessible under dry conditions, as further verified by the subsequent *ab initio* calculation.

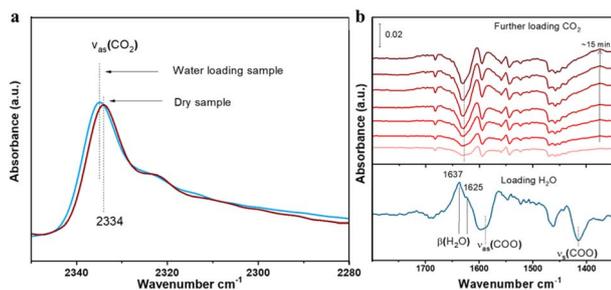


Fig. 6 (a) IR spectra of adsorbed CO<sub>2</sub> into dry (activated) and water-loaded (~50% RH) compound 1. CO<sub>2</sub> was loaded at ~120 torr and 24 °C in both cases. (b) IR spectra of loading H<sub>2</sub>O (bottom line) and CO<sub>2</sub> (top lines) sequentially into the compound 1 sample. The bottom spectrum was referenced to the activated sample to show the adsorbed H<sub>2</sub>O. The top six spectra were referenced to the bottom one to show the changes of pre-adsorbed H<sub>2</sub>O upon further loading CO<sub>2</sub>. The  $\beta(\text{H}_2\text{O})$  band was analyzed and presented here as the  $\nu(\text{OH})$  band of H<sub>2</sub>O is partially masked by the combination bands of gas phase CO<sub>2</sub> (Fig. S7†). Notation and acronym:  $\nu$ , stretch;  $\beta$ , bending; as, asymmetric; and s, symmetric.

### Theoretical calculations of the CO<sub>2</sub> binding energies and binding sites

Here, we chose two MOFs, compounds 1 and 3, to illustrate the trends and mechanisms of adsorption. We made use of an *ab initio* thermal annealing method and GCMC with the help of VASP and RASPA, respectively, for an efficient sampling of the configurational space.<sup>39,40</sup> Several low-energy configurations were identified and further subjected to full *ab initio* geometry optimization. There is a clear distinct binding spot for CO<sub>2</sub> in both 1 and 3 and H<sub>2</sub>O in compound 3 (Fig. 7). However, two low-energy configurations were found for H<sub>2</sub>O in compound 1 (Fig. S8†). The CO<sub>2</sub> binding energies in compounds 1 and 3 at the most favorable site are 42.61 and 30.27 kJ mol<sup>-1</sup>, respectively. Hence, the binding of CO<sub>2</sub> is significantly stronger with compound 1 in comparison with compound 3. These results



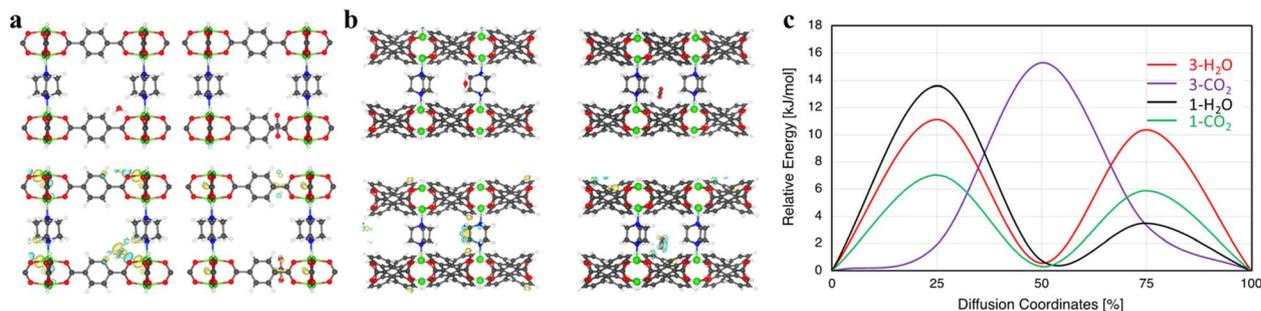


Fig. 7 (a) (Top) H<sub>2</sub>O (left) and CO<sub>2</sub> (right) optimum binding locations in compound **3**. (Bottom) Guest interaction with compound **3**. (b) (Top) H<sub>2</sub>O (left) and CO<sub>2</sub> (right) optimum binding spots in compound **1**. (Bottom) Guest interaction with compound **1**. Yellow and blue represent charge accumulation and depletion, respectively, at an iso-value of  $0.0005 \text{ e } \text{\AA}^{-3}$ . (c) Diffusion energy barrier for H<sub>2</sub>O and CO<sub>2</sub> in compounds **1** and **3**. Note: diffusion coordinate describes the progress of the molecules along their diffusion pathway, 0% means no translation and 50% corresponds to a translation halfway through the unit cell.

explain the significantly higher single-component CO<sub>2</sub> uptake in compound **1** and align with our experimental results.

Similarly, the H<sub>2</sub>O binding energies were calculated to be 45.72 and 33.19 kJ mol<sup>-1</sup> in compounds **1** and **3**, respectively. Upon further inspection, we found that, although CO<sub>2</sub> and H<sub>2</sub>O in compound **3** mainly interact with the bdc linker, there are clear distinctions. The main interaction of H<sub>2</sub>O is with the oxygen of the bdc linker; however, CO<sub>2</sub> strongly interacts with the C–C connected to O of the bdc linker (Fig. 7a). It can also be seen that H<sub>2</sub>O only interacts with one linker while CO<sub>2</sub> interacts with two of the nearby linkers (Fig. S9†). On the other hand, the interaction of guest molecules and preferential binding sites are very different in compound **1**. For example, CO<sub>2</sub> only interacts with the adc linker while due to the hydrophobic nature of adc, H<sub>2</sub>O's main interaction is with the pillar ligand dabco, with a secondary interaction at the adc linker (Fig. S8†). Due to the presence of different binding pockets for water, we find two low-energy structures with two H<sub>2</sub>O molecules, respectively (Fig. S10†). The separations between them are in the range of  $\sim 2.12\text{--}1.94 \text{ \AA}$ , depending on the configuration. Moreover, the water–water interaction is quite visible even at two water molecules per pore (Fig. S11†).

We further study the kinetics of guest molecules in both compounds **1** and **3** (Fig. 7c). It is observed that the guest molecules show different diffusion behavior in these MOFs. The diffusion energy barrier for H<sub>2</sub>O is 11.13 kJ mol<sup>-1</sup> vs. 15.28 kJ mol<sup>-1</sup> for CO<sub>2</sub> in compound **3**. Consequently, a faster diffusion of H<sub>2</sub>O is found in compound **3** in comparison with CO<sub>2</sub>. However, an opposite behavior in compound **1** is noted where we see a much faster diffusion of CO<sub>2</sub> compared to H<sub>2</sub>O. The diffusion barrier for CO<sub>2</sub> and H<sub>2</sub>O in compound **1** is 7.04 and 13.59 kJ mol<sup>-1</sup>, respectively, suggesting almost two times faster kinetics for CO<sub>2</sub>. These outcomes are in perfect agreement with our experimental observations, showing that CO<sub>2</sub> is adsorbed much faster compared with water in dynamic adsorption in compound **1**.

How could compound **1** capture CO<sub>2</sub> with an increased capacity when water molecules are present in the system with competitive co-adsorption, especially when water has a higher binding energy and demonstrates obvious water–water

interactions in the pore? The favorability of co-adsorption configuration over lower water loading has been reported in TIFSIX-3-Ni, and the CO<sub>2</sub> binding energy is notably higher than that of H<sub>2</sub>O in these systems.<sup>41</sup> However, for compound **1**, as the water interaction is stronger than that of CO<sub>2</sub>, why does the CO<sub>2</sub> uptake increase when water co-adsorption is present in the mixture? To unlock this mystery, our next step was to study CO<sub>2</sub>/H<sub>2</sub>O co-adsorption and their preferential binding sites in detail.

For compound **1**, we found two low energy preferential co-adsorption configurations as shown in Fig. 8. The total binding energy of the most favorable co-adsorption configuration (1H<sub>2</sub>O + 1CO<sub>2</sub>) is 94.5 kJ mol<sup>-1</sup>, which is stronger than the binding of 2H<sub>2</sub>O. Additionally, the secondary co-adsorption configuration is just 4.73 kJ mol<sup>-1</sup> higher in energy compared to the most favorable one. Given the small difference in energy, both types of co-adsorption likely occur under humid conditions. In this configuration, the pillar ligand offers a new binding site for CO<sub>2</sub> which is not accessible under dry conditions. Since the H<sub>2</sub>O molecules at the secondary binding site can be partially replaced by CO<sub>2</sub> based on the results from the IR experiments, this explains the  $\sim 30\%$  increase in CO<sub>2</sub> uptake and prolonged retention time under the humid conditions for compound **1**. Such synergistic effects of H<sub>2</sub>O/CO<sub>2</sub> co-adsorption that enhances the CO<sub>2</sub> binding energy and adsorption capacity are quite rare, especially in ultra-microporous materials (Fig. S12†).<sup>41</sup> Note that there is an increase in CO<sub>2</sub> binding

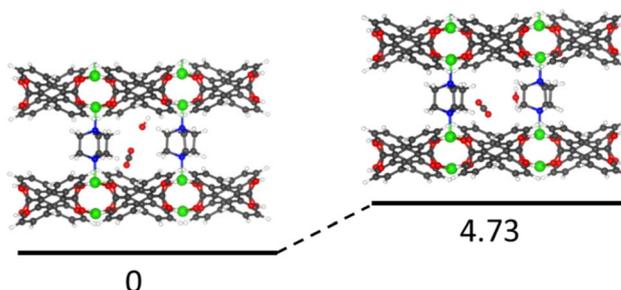


Fig. 8 The two low energy co-adsorption configurations of H<sub>2</sub>O/CO<sub>2</sub> in compound **1**.



energy when water is present in the structure. For example, in two favorable co-adsorption states, the binding energy of CO<sub>2</sub> is 50.62 kJ mol<sup>-1</sup> and 44.05 kJ mol<sup>-1</sup>, respectively, which are both higher than that of CO<sub>2</sub> in a single component adsorption experiment (42.61 kJ mol<sup>-1</sup>).

To summarize, the higher binding energy and faster kinetics of CO<sub>2</sub> upon adsorption in compound **1** can be attributed to its higher conjugation and hydrophobicity. Conversely, much slower water kinetics and changes in the pore surface environment hinder water accumulation within pores. The adsorption enhancement achieved in this study further highlights the importance of optimizing the pore surface structure through ligand functionalization.

## Conclusions

We have successfully synthesized a series of isorecticular Ni<sub>2</sub>(L)<sub>2</sub>(dabco) MOFs using ligands with different degrees of conjugation. The increase in the conjugation level of the ligands correlates with an increase in ligand size and a decrease in pore size and surface area, resulting in a different pore surface structure. Compound **1** with the lowest BET surface area and the smallest pore size reaches 1.75 times higher CO<sub>2</sub> uptake capacity compared to compound **3** with the largest BET surface area and pore size and the lowest hydrophobicity under dry conditions. The calculated CO<sub>2</sub> isosteric heats of adsorption and binding energies are fully consistent with these experimental observations. The enhancement of CO<sub>2</sub> uptake is more prominent in the lower pressure region. More significantly, compound **1** exhibits a substantial increase in CO<sub>2</sub> uptake (>31%) in binary CO<sub>2</sub>/N<sub>2</sub> mixtures under humid conditions, as confirmed by both dynamic adsorption isotherm and column breakthrough measurements. The possible reasons are elucidated by *in situ* FTIR analysis along with theoretical *ab initio* calculations that suggest that the co-adsorption configuration of H<sub>2</sub>O and CO<sub>2</sub> is favorable and generates extra binding sites to accommodate more CO<sub>2</sub>, leading to an enhancement of CO<sub>2</sub> uptake. Optimizing the pore surface structure of MOFs by increasing ligand conjugation can be an effective approach to significantly enhance CO<sub>2</sub> adsorption, both under single component dry conditions and in gas mixtures with competitive water co-adsorption under humid conditions.

## Experimental section

### Synthesis of H<sub>2</sub>adc

The dicarboxylic linker 9,10-anthracenedicarboxylic acid (H<sub>2</sub>adc) was synthesized according to a reported procedure with some modifications.<sup>42</sup> In a 500 mL two-necked round bottom flask, 8 grams of 9,10-dibromoanthracene were dissolved in 300 mL of dry ether and stirred at 0 °C. Meanwhile, 33 mL of 1.6 M *n*-butyllithium was added to the mixture under nitrogen protection and kept stirring for 2 hours. Pure carbon dioxide gas was purged into the resulting yellowish mixture at -78 °C (dry ice and acetone bath) and the entire reaction mixture was stirred overnight. A bright yellow solid was precipitated out once the water and hydrochloric acid (6 M) were added to the

solution. The solid was filtered, washed with a copious amount of fresh water and collected. The yield is 90%. <sup>1</sup>H NMR (500 MHz, DMSO-d<sub>6</sub>, δ): 8.04 (dd, 4H, Ar H), 7.66 (dd, 4H, Ar H). The peaks of H<sub>2</sub>adc were assigned and integrated to confirm the structure (Fig. S1†). The synthesis of this ligand was scaled up to 20 g and can be scaled up further with appropriate reaction vessels.

### Synthesis of Ni<sub>2</sub>(L)<sub>2</sub>(dabco)

The Ni<sub>2</sub>(L)<sub>2</sub>(dabco) MOF series were synthesized *via* similar solvothermal reactions with some modifications and improvements to the previous procedure. Ni(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O (145 mg, 0.5 mmol), 1,4-diazabicyclo[2,2,2]octane (dabco, 56 mg, 0.5 mmol) and 0.5 mmol of ligands (H<sub>2</sub>adc 133 mg, or H<sub>2</sub>ndc 108 mg, or H<sub>2</sub>bdc 83 mg) were dispersed in 10 mL of DMF solution in a 20 mL glass vial. The mixture was then stirred at room temperature for 2 hours followed by filtering using a syringe Nylon filter. The clear filtrate was collected and sealed in a new vial, which was then put into an oven pre-heated to 120 °C and reacted for two days. The crystalline precipitates were filtered and washed with fresh DMF and MeOH several times. The yields were calculated based on the ligand amount and were 53%, 73%, and 71% for compounds **1**, **2** and **3**, respectively. The products were characterized by powder X-ray diffraction (PXRD) and thermogravimetric (TG) analysis. The largest batch we have achieved was ~5 g and the synthesis procedure can also be modified to larger scales.

### Gas adsorption measurements

Prior to the gas adsorption test, the as-made samples were first washed thoroughly with DMF and then soaked in fresh MeOH for two days to allow solvent exchange to complete. Around 80 mg of exchanged sample was activated at 393 K under dynamic vacuum for three hours prior to adsorption experiments. Gas adsorption measurements were carried out on a Micromeritics 3Flex volumetric adsorption analyzer. The BET surface area and pore size distribution were determined based on the nitrogen adsorption-desorption data at 77 K. Liquid nitrogen, dry ice-isopropyl alcohol bath, and a circulating-bath digital temperature controller were used for measurements at 77 K, 195 K, and temperatures around room temperature, respectively.

The single-component dynamic adsorption isotherms for CO<sub>2</sub> and H<sub>2</sub>O vapor were collected in a gravimetric adsorption analyzer TGA Q50 (TA Instruments). The bubbler contained water-run ultrahigh purity N<sub>2</sub> as the carrier gas. The partial pressure of CO<sub>2</sub> or H<sub>2</sub>O was adjusted by controlling the ratio of pure N<sub>2</sub> gas and CO<sub>2</sub> or saturated water vapor. Approximately 20 mg of exchanged samples was activated under a constant N<sub>2</sub> flow at 393 K for 60 min. Adsorbed amounts were monitored continuously by weight changes in the sample throughout the measurements.

### Isosteric heat of adsorption (Q<sub>st</sub>)

To determine the affinity between the scaffold and the adsorbates, the Clausius-Clapeyron equation<sup>43</sup> was employed to



calculate the adsorption heats for CO<sub>2</sub> on compounds 1–3, which is defined as follows:

$$Q_{\text{st}} = -RT^2 \left( \frac{\partial \ln P}{\partial \ln T} \right) n_{\text{a}}$$

where  $Q_{\text{st}}$  stands for the adsorption heat of CO<sub>2</sub>,  $P$  and  $T$  represent the pressure and temperature.  $R$  is the universal gas constant and  $n_{\text{a}}$  is the adsorbed amount in moles. Before the calculation, the isotherms tested at 273 K, 288 K, and 298 K were fitted using the virial equation:<sup>44</sup>

$$\ln(P) = \ln(N) + \left( \frac{1}{T} \right) \sum_{i=0}^m a_i \times N^i + \sum_{j=0}^n b_j \times N^j$$

$$Q_{\text{st}} = -R \times \sum_{i=0}^m a_i \times N^i$$

where  $P$  is the pressure described in pascals,  $N$  is the adsorbed amount in mmol g<sup>-1</sup>,  $T$  is the temperature in kelvin,  $a_i$  and  $b_j$  are virial coefficients, and  $m$  and  $n$  are the number of coefficients used to describe the isotherms.

### In situ infrared (IR) spectroscopic measurements

*In situ* IR measurements were performed on a Nicolet™ iS50 FTIR spectrometer using a liquid N<sub>2</sub>-cooled mercury cadmium telluride (MCT-A) detector. The spectrometer is equipped with a vacuum cell that is placed in the main compartment with the sample at the focal point of the infrared beam. To avoid the direct pressing of the MOF sample that may cause damage to the crystalline structure, the sample (~5 mg) was made into a slurry form by mixing with a small amount of methanol and pasted onto a KBr pellet. The powder sample was dried and directly attached to the KBr pellet after blowing a nitrogen stream for a minute, and then transferred into the cell that is connected to a vacuum line for evacuation. The sample was activated by evacuation at 150 °C for 3 h and then cooled back to room temperature for CO<sub>2</sub> adsorption measurement. H<sub>2</sub>O was loaded by exposing the sample to ~11 torr water vapor at 24 °C for ~5 min until adsorption reaches saturation. CO<sub>2</sub> loading was followed by introducing 80 torr gas into the cell and keeping for ~15 min.

### Computational and simulation details

To deepen our understanding of molecular-level guest–guest and guest–host interactions, *ab initio* calculations were carried out utilizing the VASP code.<sup>45,46</sup> These calculations captured the necessary van der Waals interactions through the non-local vdW-DF1 functional with a plane-wave basis set and in conjunction with standard PAW pseudopotentials.<sup>47–50</sup>  $\Gamma$ -Point sampling along with a kinetic energy cutoff of 600 eV gave well-converged results. The geometry optimization criteria were: 10<sup>-6</sup> eV for SCF loops and 0.005 eV Å<sup>-1</sup> for Hellmann–Feynman forces. The binding energy ( $E_{\text{b}}$ ) is calculated from the energy difference between the loaded MOF ( $E_{\text{total}}$ ) and the individual fragments (activated empty MOF and isolated gas-phase energy of the molecule):

$$E_{\text{b}} = E_{\text{MOF}} + E_{\text{guest}} - E_{\text{total}}$$

Diffusion energy barriers were determined using the climbing-image nudged elastic band method (cNEB), a transition-state search algorithm,<sup>51</sup> and five transition images were selected.

### Column breakthrough measurements

Breakthrough tests were carried out at 298 K and 1 bar in an auto mixed-gas breakthrough apparatus (BSD-MAB) equipped with a vapor generator. The mass of compounds 1–3 filled into the column (I.D. 6 mm and length 80 mm) was 0.273 g, 0.51 g, and 0.79 g, respectively. The adsorbents were activated at 120 °C for 2 h under purging helium gas (20 mL min<sup>-1</sup>). When the temperature was cooled down to 25 °C, the helium flow was stopped and the feed gas mixtures (CO<sub>2</sub>/N<sub>2</sub>, 15/85; v/v, dry or wet) at a flow rate of 10 mL min<sup>-1</sup> were introduced into the adsorption column. The wet feed gas was pre-mixed with a relative humidity of 50%. The gases at the outlet were analyzed using a mass spectrometer (MKS). After the adsorption reached equilibrium, the column was purged with helium gas (20 mL min<sup>-1</sup>) at 100 °C for 2 h for desorption.

### Data availability

Crystallographic data for Zn-DABCO-ADC have been deposited at the CCDC under 2360909 and can be obtained from <https://doi.org/10.5517/ccdc.csd.cc2k7qcx>. All data supporting this research (like PXRD, TGA, adsorption isotherms, NMR, FT-IR, etc.) are included in the main article and/or ESI.†

### Author contributions

G. Z.: methodology, investigation, formal analysis, visualization, and writing – original draft; F. X.: investigation and formal analysis; S. U., T. T.: computational investigation and formal analysis; H. W., L. M.: investigation and formal analysis (breakthrough); K. T., S. M.: investigation and formal analysis (IR); S. T.: investigation (single-crystal); J. L.: conceptualization, supervision, validation, and writing – review & editing.

### Conflicts of interest

There are no conflicts to declare.

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