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# Hydrogen peroxide assisted synthesis of fluorescent carbon nanoparticles from teak leaves for dye-sensitized solar cells†

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Fluorescent carbon nanoparticles (FCNs) have emerged as promising sensitizers for dye-sensitized solar cells (DSSCs) owing to their unique optical properties and low-cost fabrication. In this study, we synthesized three different types of FCNs, such as hydrogen peroxide (H<sub>2</sub>O<sub>2</sub>)-assisted FCNs (FCN<sub>H</sub>), hydrothermal-assisted FCNs (FCN<sub>A</sub>), and pyrolysis-assisted (FCN<sub>P</sub>). Among them, FCN<sub>H</sub>-based DSSCs exhibited superior fluorescence properties, indicating effective surface oxidation by H<sub>2</sub>O<sub>2</sub>. Then we fabricated DSSCs using FCN<sub>P</sub>, FCN<sub>A</sub>, and FCN<sub>H</sub> as sensitizers and evaluated their photovoltaic activity. The results revealed that FCN<sub>H</sub>-based DSSCs exhibited superior performance compared to FCN<sub>P</sub>-based and FCN<sub>A</sub>-based DSSCs with an open circuit voltage (*V*<sub>OC</sub>) of 0.43 V, a short circuit current (*I*<sub>SC</sub>) of 0.52 mA, and a fill factor (FF) of 0.45. The superior performance of FCN<sub>H</sub>-based DSSCs can be attributed to the improved charge transfer and fluorescence properties of the FCNs after surface oxidation with H<sub>2</sub>O<sub>2</sub>. These findings emphasize the significance of surface oxidation in optimizing the performance of FCNs in DSSCs and suggest the potentiality of FCN<sub>H</sub> as an efficient sensitizer in renewable energy as well as energy harvesting applications for wearable platforms.

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## Sustainability spotlight

The unique optical properties and low fabrication cost of fluorescent carbon nanoparticles (FCNs) have made them a promising sensitizer for dye-sensitized solar cells (DSSCs). Typically, concentrated oxidizing acidic solutions, including HNO<sub>3</sub>, H<sub>2</sub>SO<sub>4</sub>, H<sub>2</sub>SO<sub>4</sub>/HNO<sub>3</sub> mixture and H<sub>3</sub>PO<sub>4</sub>, have been used to induce the oxidation of carbon sources, resulting in the formation of carbon nanodots with surface functional moieties, such as carboxylic acids, hydroxyl groups, or carbonyl groups. Teak leaves (*Tectona grandis* L. f.) were utilized in this study as a natural resource to create affordable, environmentally beneficial H<sub>2</sub>O<sub>2</sub>-assisted functionalized carbon nanomaterials (FCN<sub>H</sub>). The fluorescent carbon nanoparticles (FCNs) were applied in dye sensitized solar cells. The study's findings highlight the importance of surface oxidation in maximizing the performance of functionalized carbon nanomaterials (FCNs) in dye-sensitized solar cells (DSSCs) and point to the possibility of using H<sub>2</sub>O<sub>2</sub>-assisted FCNs as an effective sensitizer in energy harvesting and renewable energy applications for wearable platforms. The Sustainable Development Goals (SDGs) of the UN are met by our work: responsible consumption and production (SDG 12), sustainable cities and communities (SDG 11), and industry, innovation, and infrastructure (SDG 9).

## 1 Introduction

In recent years, fluorescent carbon nanomaterials (FCNs) have gained considerable attention from researchers due to their unique properties, such as tunability in size, shape, excellent

electron conductivity, fluorescence properties, surface chemistry, and optical properties.<sup>1–9</sup> These properties make them attractive for a variety of applications, including optoelectronic devices,<sup>9,10</sup> wearable sensors,<sup>11</sup> drug delivery,<sup>12</sup> supercapacitors,<sup>13</sup> photocatalysis,<sup>14</sup> solar cells,<sup>2,5,15,16</sup> bioimaging and biosensing.<sup>17–19</sup>

Various synthesis strategies have been developed for preparing FCNs, including top-down and bottom-up approaches.<sup>20,21</sup> Top-down approaches involve the fragmentation of a larger carbon structure using physical or chemical methods, followed by size selection and surface modification. Bottom-up approaches involve the formation of carbon nanoparticles from smaller molecular precursors through chemical reactions. Other methods, such as electrochemical and photochemical synthesis, have also been reported.<sup>22</sup> Typically, concentrated oxidizing acidic solutions, including HNO<sub>3</sub>,<sup>23</sup>

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$\text{H}_2\text{SO}_4$ ,<sup>24</sup>  $\text{H}_2\text{SO}_4/\text{HNO}_3$  mixture,<sup>25,26</sup> and  $\text{H}_3\text{PO}_4$ ,<sup>27</sup> have been used to induce the oxidation of carbon sources, resulting in the formation of carbon nanodots with surface functional moieties, such as carboxylic acids, hydroxyl groups, or carbonyl groups.

In the last few years,  $\text{H}_2\text{O}_2$ -assisted quantum dots from different sources, such as graphene,<sup>28</sup> sulfur,<sup>29</sup> and nitrogen-doped carbon,<sup>30</sup> have been synthesized by hydrothermal treatment. In this study, we used teak leaves (*Tectona grandis* L. f.) as a natural resource to synthesize eco-friendly and cost-effective  $\text{H}_2\text{O}_2$ -assisted functionalized carbon nanomaterials ( $\text{FCN}_\text{H}$ ). Teak leaves are a promising source of carbon due to their abundance of various organic compounds.<sup>31</sup>

Hydrothermal treatment is a cost-effective and convenient one-step approach for producing water-soluble FCNs with a narrow size distribution on a large scale.<sup>32</sup> The use of  $\text{H}_2\text{O}_2$  as an oxidizer during the synthesis process can incorporate oxygen-containing functional groups onto the surface of the carbon skeleton and improve the aqueous solubility and biocompatibility of the FCNs.<sup>33</sup> However, in this study, we also fabricated hydrothermal-assisted FCNs using only deionized (DI) water (referred to as  $\text{FCN}_\text{A}$ ), and additional FCNs were synthesized by pyrolyzing teak leaves (referred to as  $\text{FCN}_\text{P}$ ) by using  $\text{H}_2\text{O}_2$ . It was observed that  $\text{FCN}_\text{A}$  exhibited a blue emission color and showed good stability in an aqueous solution under 365 nm UV irradiation, despite the absence of  $\text{H}_2\text{O}_2$ .

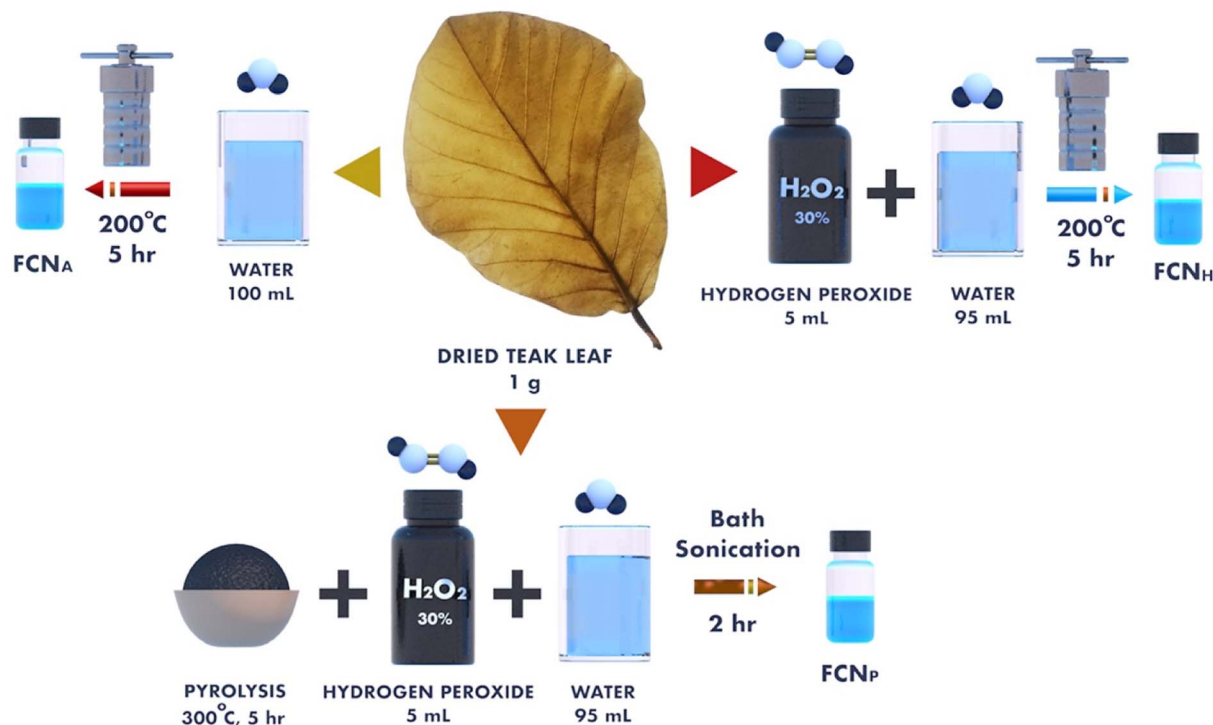
To date, in the field of dye-sensitized solar cells (DSSCs), a limited number of applications of FCNs have been reported.<sup>34</sup> In DSSCs, FCNs act as sensitizers, enhancing light absorption in the visible region and thereby improving overall solar cell efficiency.<sup>34</sup> Nevertheless, organic dyes, particularly ruthenium (Ru) dyes,<sup>35,36</sup>

and D- $\pi$ -A structured porphyrin dyes,<sup>34,37-39</sup> have also been used as photosensitizers in DSSCs. Generally, organic-based sensitizers have demonstrated superior performance compared to FCNs. However, due to certain drawbacks associated with organic dyes, such as low stability, time-consuming synthesis process and toxicity, the scientific community has emphasized the need for the development of simple, cost-effective and eco-friendly DSSCs to overcome these limitations.<sup>34</sup> From this perspective, the use of non-toxic and environmentally friendly FCNs can potentially contribute towards fabricating DSSCs. In this study, blue emitting FCNs are synthesized from teak leaves to construct DSSCs. To the best of our knowledge, this study represents the first attempt to employ  $\text{H}_2\text{O}_2$ -assisted FCNs in DSSCs, with results indicating excellent activity compared to traditional FCNs.

## 2 Experimental section

### 2.1 Chemicals and apparatus

Analytical grade hydrogen peroxide (30%) was purchased from Sigma-Aldrich. Dried teak tree (*Tectona grandis* L. f.) leaves were collected near the Chittagong University of Engineering & Technology area in the Chattogram region of Bangladesh. The titanium dioxide ( $\text{TiO}_2$ ) powder (Degussa P-25) for the photoanode preparation was acquired from the Institute of Chemical Education, University of Wisconsin–Madison, Department of Chemistry, Madison, WI, USA. Fluorine tin oxide (FTO) conducting glass slides were purchased from Hartford Glass Company, Hartford City, Indiana, USA. Acetone, ethanol, and acetic acid were obtained from Sigma-Aldrich (St. Louis, MS, USA) and used without additional purification. DI water was used throughout all segments in this experiment.



Scheme 1 Different synthesis routes of FCNs from dried teak leaves.



## 2.2 Experimental

**2.2.1 Preparation of FCN<sub>H</sub>.** FCNs were synthesized using dried teak leaves as a carbon source through hydrothermal and pyrolysis methods (Scheme 1). For the preparation of FCN<sub>H</sub>, the teak leaves were washed with deionized water, dried, and sieved (0.3 mm) to remove the veins and midrib. Next, 1 g fine teak leaf powder was mixed with 5 mL H<sub>2</sub>O<sub>2</sub> and 95 mL DI water, sonicated and transferred into a PPL-lined autoclave for heating at 200 °C for 5 hours. After cooling down at room temperature, the resulting light-yellow solution was purified and dialyzed against DI water to remove the excess amount of H<sub>2</sub>O<sub>2</sub>. The synthesized FCNs are referred to as FCN<sub>H</sub>. Moreover, FCN<sub>A</sub> and FCN<sub>P</sub> were also synthesized for the purpose of comparison in this study. All samples were collected in powder form through a freeze-drying process.

**2.2.2 Synthesis of FCN<sub>A</sub>.** To prepare FCN<sub>A</sub>, 1 g of fine dried-teak leaf powder was sonicated in 100 mL of DI water for 1 hour, without the addition of H<sub>2</sub>O<sub>2</sub>. The resulting solution was then transferred to a PPL-lined hydrothermal autoclave and heated at 200 °C for 5 hours. After allowing it to cool at room temperature, the light-yellow solution was filtered through a 0.22 μm membrane filter to remove larger particles and the resulting powder was obtained by freeze-drying.

**2.2.3 Synthesis of FCN<sub>P</sub>.** To synthesize FCN<sub>P</sub> through the pyrolysis process, 1 g of teak leaf powder was heated at 300 °C for 5 hours in a muffle furnace. After cooling to room temperature, the pyrolyzed sample was dispersed in 5 mL H<sub>2</sub>O<sub>2</sub> and 95 mL of DI water with the assistance of a bath-sonicator (Powersonic) for 2 hours. FCN<sub>P</sub> was filtered through a 0.22 μm membrane filter and then collected in dried form through freeze-drying.

## 2.3 Fabrication of dye sensitized solar cells

The titanium dioxide (TiO<sub>2</sub>) films were prepared utilizing a previously documented spin-coating technique.<sup>40–42</sup> To prepare the working electrode, the fluorine-doped transparent tin oxide (FTO) substrate (Hartford Glass Co., Inc. TEC 7, 1" × 1" × 2.2 mm thickness, Solar) was cleaned with detergent, deionized (DI) water and ethanol using an ultrasonic bath for 30 min. To prepare the titanium dioxide (TiO<sub>2</sub>) paste, TiO<sub>2</sub> powder (Degussa P-25) was initially mixed with glacial acetic acid. Then, the prepared TiO<sub>2</sub> paste was applied to the conductive side of FTO using a spin coater and annealed at 450 °C for 30 min.<sup>43–45</sup> The substrate was then immersed in FCN solutions (1 mg of FCNs in 1 mL pure water) overnight for sensitization. Loosely bound particles were removed by rinsing with water and ethanol and the photoanode was dried at room temperature. The counter electrode (cathode) was prepared by painting colloidal graphite on the cleaned FTO-coated glass slide and allowed it to dry at room temperature. Finally, to fabricate the DSSCs, the FCN-sensitized slide and the carbon electrodes were put on top of each other and the redox (I<sup>-</sup>/I<sub>3</sub><sup>-</sup>) electrolyte solution was placed between them.

## 2.4 Characterization and measurements

The fluorescent nanoparticles and the sensitized film's morphology were assessed using field emission scanning

electron microscopy (JSM-7100FA JEOL USA, Inc.). Absorption spectroscopy was conducted using a UV-3600 Plus from Shimadzu, MD, USA, while emission spectroscopy measurements were performed with an RF-5301PC from Shimadzu, MD, USA. The zeta potentials of the samples were determined using an SZ-100 series Dynamic Light Scattering Particle Size Distribution Analyzer, HORIBA Instruments Inc., Irvine, CA, USA. Transmission Electron Microscopy (TEM) images were obtained using a JEM-1400 PLUS (JEOL USA, Peabody, Massachusetts, USA). The TiO<sub>2</sub> paste was deposited onto FTO glass using a WS-650 Series Spin Processor from Laurell Technologies Corporation, PA, USA. Carbon paint obtained from TED PELLA, INC, USA was used for making cathode slides.

## 2.5 Photovoltaic measurements of DSSCs

The cell performance was evaluated using a 150 W fully reflective solar simulator with standard illumination of air-mass 1.5 global (AM 1.5 G) and an irradiance of 100 mW cm<sup>-2</sup> (Sciencetech Inc., London, Ontario, Canada). A GAMRY Instruments Reference 600 Potentiostat/Galvanostat/ZRA (Warminster, PA) was employed for the measurements. Cells were externally biased and the photocurrent generated was measured to determine the current–voltage characteristics of the cells under these conditions. The electrocatalytic performance of the samples was further evaluated by electrochemical impedance spectroscopy (EIS).

## 3 Results and discussion

Surface oxidation is a recognized method for enhancing fluorescence in carbon-based materials by introducing oxygen-containing functional groups *via* the generation of surface defects.<sup>46</sup> H<sub>2</sub>O<sub>2</sub> is known to be effective in promoting the generation of these defects, resulting in an abundance of carbonyl and carboxylic groups on the carbon structure, thereby improving the solubility in aqueous solutions.<sup>47</sup> The prepared H<sub>2</sub>O<sub>2</sub>-supported aqueous solution of FCN<sub>H</sub> exhibited blue emission under 365 nm UV light and formed stable-colored suspensions in DI water, which is likely due to the larger number of functional groups, such as hydroxyl and carboxylic groups.

The UV-vis spectra of all FCNs have been investigated and are presented in Fig. 1a. The spectra showed a weak absorption in the UV region at around 220 nm attributed to the π–π\* electronic transition of the sp<sup>2</sup> conjugated carbon (C=C) bond<sup>48</sup> and a weak absorption peak at around 285 nm due to the functional groups (C=O) present on the surface of all FCNs.<sup>49</sup> Notably, the absorbance peak position of FCN<sub>H</sub> shifted towards the blue region compared to the other two FCNs, signifying a larger number of hydroxyl and carboxylic groups incorporated into the carbon structure through oxidation with H<sub>2</sub>O<sub>2</sub>.

The surface functional groups of the three different types of the synthesized FCNs were investigated using Fourier transform infrared (FTIR) spectroscopy, as presented in Fig. 1b. The FCN<sub>H</sub> sample exhibited a higher overall peak intensity compared to FCN<sub>A</sub> and FCN<sub>P</sub>. All FCNs exhibited identical C=C stretching



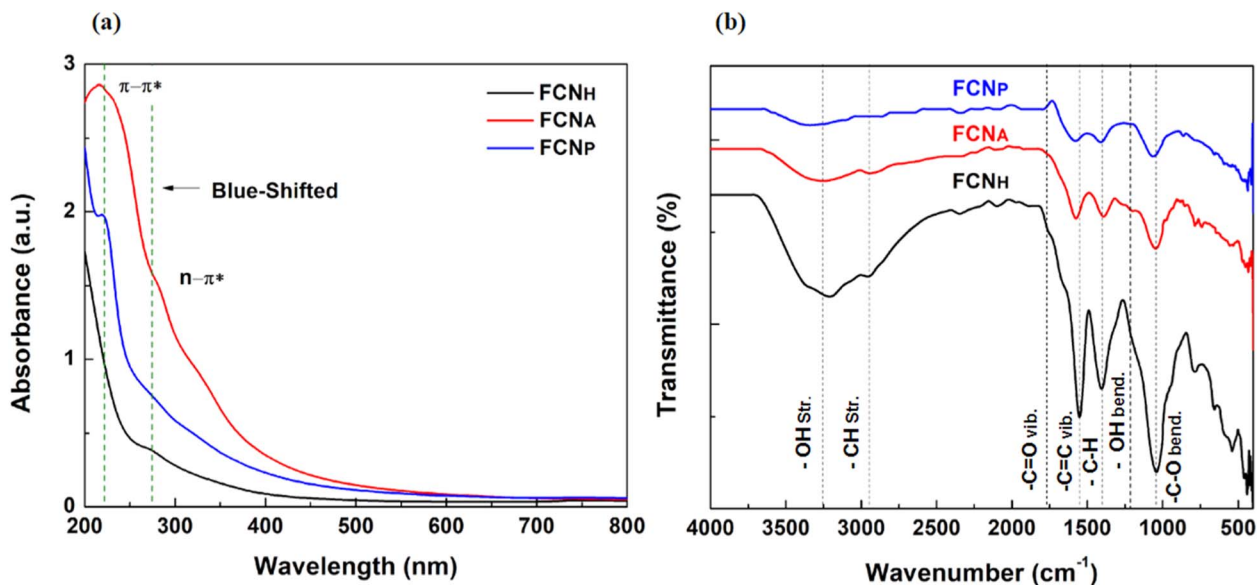


Fig. 1 (a) UV-vis spectra and (b) FT-IR of the as-synthesized FCNs.

and C–H bending peaks at around 1600 and 1027  $\text{cm}^{-1}$ , respectively, attributed to the alkyl hydrocarbon from the teak leaf. The broad signature band peak of –OH stretching at 3200  $\text{cm}^{-1}$  to 3400  $\text{cm}^{-1}$  was detected in all FCNs, confirming the adsorption of water molecules onto sample surfaces, along with an –OH bending peak at around 1390  $\text{cm}^{-1}$ . Additionally, a distinguished C=O stretching peak was detected at around 1740  $\text{cm}^{-1}$  with high intensity in FCN<sub>H</sub>, confirming the successful incorporation of a considerable number of oxygen functional groups into the carbon moieties after treatment with H<sub>2</sub>O<sub>2</sub>. A superimposed C–H stretching peak of each sample was also observed in the region 2400–2950  $\text{cm}^{-1}$ .<sup>50</sup>

The photoluminescence (PL) spectra were recorded with an excitation wavelength of 340 nm. All synthesized carbon nanoparticles showed fluorescence emission peaks in the blue region, ranging from 430 nm to 435 nm. These results indicated the maximum wavelength at which the emitted fluorescence was observed for all types of FCNs (Fig. 2).

The zeta potential, which quantifies the surface charge of colloidal particles in a solution, was measured and is presented

in Fig. 3. The zeta potential values of the three types of FCNs were assessed using electrophoretic light scattering in an aqueous solution. All samples exhibited negative zeta potential values, confirming negatively charged surfaces. This negative charge is attributed to the presence of oxygenated functional groups on the particle surfaces, which dissociate in the solution and generate negatively charged ions. Among the three samples, FCN<sub>H</sub> displayed the most negative zeta potential value of –27.7 mV, demonstrating a higher surface negative charge density. The formation of a larger number of functional groups on the particle surface can be attributed to H<sub>2</sub>O<sub>2</sub> treatment. The H<sub>2</sub>O<sub>2</sub> treatment likely facilitated the introduction of more oxygen functionalities, leading to an increased negative charge on the surface. On the other hand, FCN<sub>A</sub> exhibited a zeta potential value of –11.2 mV, while FCN<sub>P</sub> showed a more negative value of –23.6 mV, suggesting a notable increase in the negative surface charge compared to FCN<sub>A</sub>. Notably, FCN<sub>H</sub> exposed an even stronger negative charge than FCN<sub>P</sub>, demonstrating a certain level of passivation or resistance toward chemical reactions.<sup>51,52</sup> This passivation effect could limit or

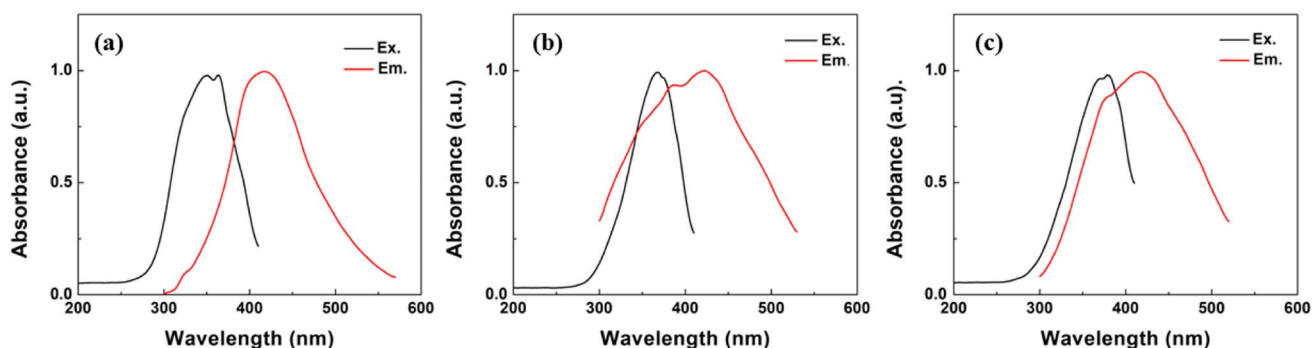


Fig. 2 PL spectra of (a) FCN<sub>H</sub>, (b) FCN<sub>A</sub> and (c) FCN<sub>P</sub>.





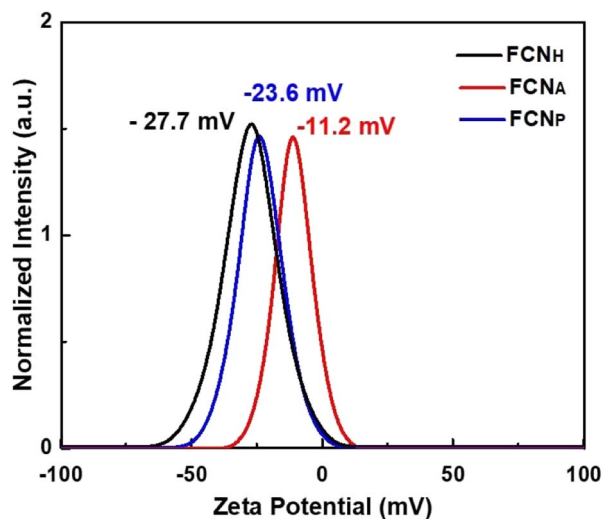


Fig. 3 Zeta Potential measurement of all FCNs.

slow down the functionalization reactions, even with  $\text{H}_2\text{O}_2$  treatment. Even though  $\text{H}_2\text{O}_2$  introduced oxygenated functional groups, the surface properties of pyrolytic carbon nanoparticles might impede the further incorporation of functional groups,

resulting in lower zeta potential values compared to  $\text{FCN}_\text{H}$ .<sup>53</sup> The observed percentages in EDX (Fig. S1 and Table S1†) reflect the relative abundance of carbon (C) and oxygen (O) atoms within each sample.  $\text{FCN}_\text{H}$  exhibited a composition of 79.53% C and 20.47% O atoms, indicating a relatively higher oxygen content. The treatment with  $\text{H}_2\text{O}_2$  mainly introduced oxygen-containing functional groups onto the surface of the carbon dots through oxidation reactions. Similarly,  $\text{FCN}_\text{A}$  displayed a composition of 81.59% C and 16.52% O, which confirmed the slightly lower oxygen content compared to  $\text{FCN}_\text{H}$ . The reduction in oxygen content might be due to the breaking of weaker oxygen-carbon bonds under the high-temperature and pressure conditions of hydrothermal treatment.<sup>54</sup> The composition of 42.71% carbon and 14.35% oxygen in  $\text{FCN}_\text{P}$  suggests a significant reduction in both carbon and oxygen contents compared to the other treatments. The reduction in both carbon and oxygen contents implies that the pyrolysis process led to the removal of various oxygen-functional groups, as well as a partial decomposition of carbon components, resulting in a composition with relatively lower carbon and oxygen concentrations.

In this study, we also investigated the surface morphology of  $\text{FCN}_\text{H}$  using transmission electron microscopy (TEM) and field-emission scanning electron microscopy (FE-SEM). The TEM images displayed individual particles within a size range of 30–

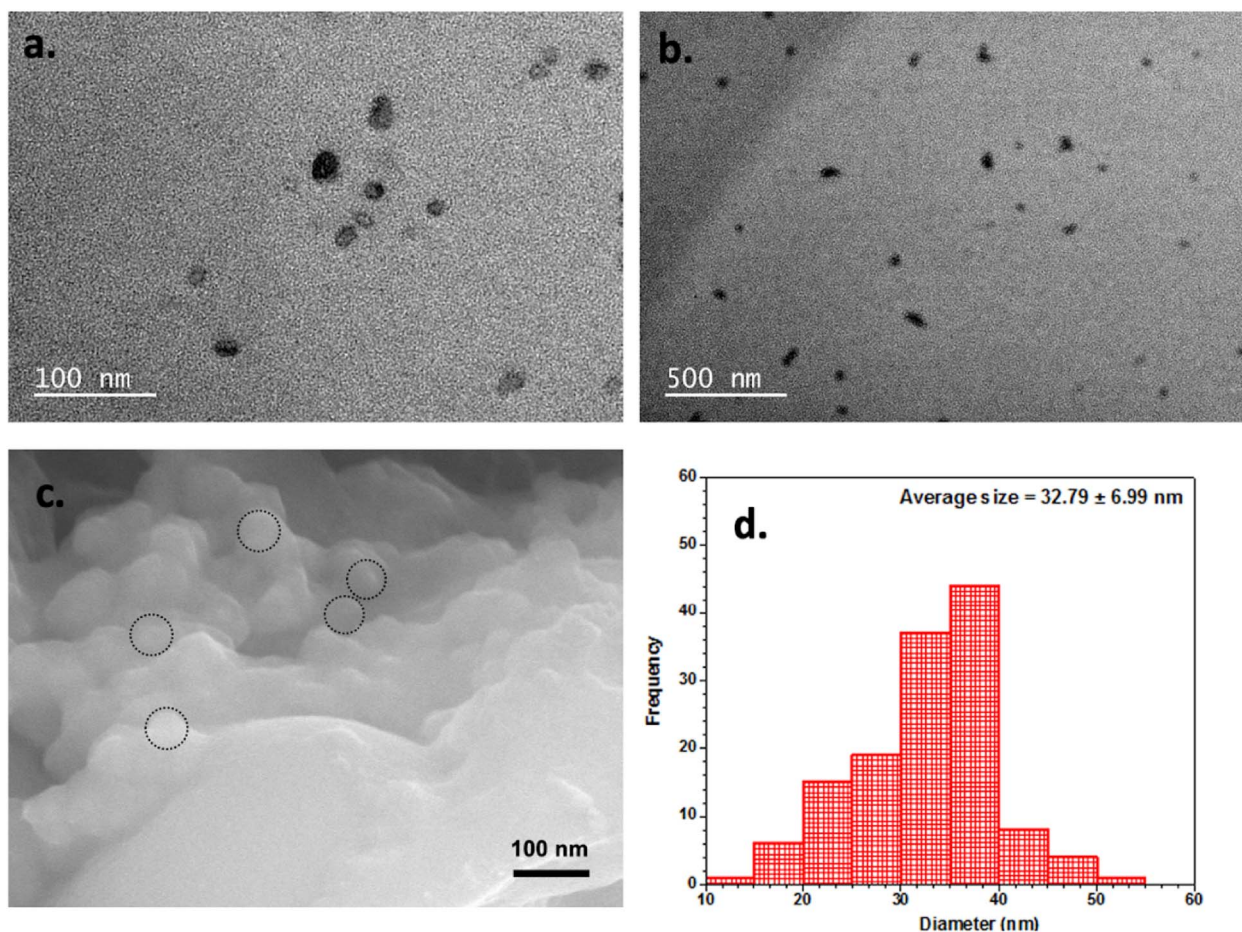


Fig. 4 (a) High-resolution TEM image, (b) TEM image, (c) FE-SEM image and (d) particle size distribution histogram of  $\text{FCN}_\text{H}$ .



40 nm, as depicted in Fig. 4a and b. The FE-SEM results, depicted in Fig. 4c, along with the size distribution measurement (Fig. 4d), corroborated and strengthened the findings from TEM, confirming the presence of nano-sized particles in the sample. Additionally, the FE-SEM images in Fig. S2† revealed that FCN<sub>P</sub> possessed smaller particle sizes compared to FCN<sub>A</sub>. Notably, the introduction of H<sub>2</sub>O<sub>2</sub> during the synthesis process resulted in a more significant reduction in particle size for FCN<sub>P</sub> compared to FCN<sub>A</sub>. The DLS data (Fig. S3†) obtained in this study were also consistent with the FE-SEM images, supporting the observation that the particle size of FCN<sub>P</sub> is smaller than that of FCN<sub>A</sub>.

The photoluminescence (PL) lifetime values of the carbon dots were measured for FCN<sub>H</sub>, FCN<sub>A</sub>, and FCN<sub>P</sub> (Fig. 5). For FCN<sub>H</sub>, the PL lifetime values were 2.20 ns and 8.69 ns, pointing to two distinct decay processes. Similarly, FCN<sub>A</sub> exhibited two PL lifetime values of 1.28 ns and 7.07 ns, indicating the presence of two decay processes in its emission behavior. On the

Table 1 Fluorescence lifetime measurement of FCNs

Sample name	Lifetime			
	$\tau_1$ (ns)	Std dev.	$\tau_2$ (ns)	Std dev.
FCN <sub>H</sub>	2.20	$1.21 \times 10^{-11}$	8.69	$2.74 \times 10^{-11}$
FCN <sub>A</sub>	1.28	$8.39 \times 10^{-11}$	7.07	$3.03 \times 10^{-11}$
FCN <sub>P</sub>	2.09	$1.32 \times 10^{-11}$	8.57	$3.00 \times 10^{-11}$

other hand, FCN<sub>P</sub> displayed PL lifetime values of 2.09 ns and 8.57 ns when tested with the N719 dye. Therefore, it can be inferred from the PL lifetime values that FCN<sub>H</sub> has a longer PL lifetime compared to FCN<sub>A</sub> and FCN<sub>P</sub>, proposing a lower rate of non-radiative recombination (Table 1). This characteristic potentially contributes to the enhanced solar-to-electric power conversion efficiency in the FCN<sub>H</sub>-based DSSCs. Conversely, FCN<sub>A</sub> and FCN<sub>P</sub>, with their shorter PL lifetimes, are likely to exhibit higher rates of non-radiative recombination, thus resulting in a lower solar-to-electric power conversion efficiency in the DSSCs that utilized these dyes.

### 3.1 Use of FCNs as sensitizers in dye sensitized solar cells

When they are exposed to sunlight, dye molecules (S) absorbed on the TiO<sub>2</sub> film absorb photons and are excited from the highest occupied molecular orbital (HOMO) to the lowest unoccupied molecular orbital (LUMO) state as shown in Scheme 2. Upon photoexcitation, the dye species (S\*) injects an electron into the conduction band of the TiO<sub>2</sub> electrode, leading to its oxidation (S<sup>+</sup>). The oxidized dye species accepts an electron from the electrolyte (I<sup>-</sup>), resulting in the restoration of the ground state of the dye (S). The injected electron is then transported through the nano-porous TiO<sub>2</sub> film to the FTO layer and further conducted through an external circuit to a load, where it is converted into electrical energy. The electron from the external load subsequently diffuses to the cathode and is transferred to the electrolyte (I<sub>3</sub><sup>-</sup>), thus regenerating the electrolyte system.

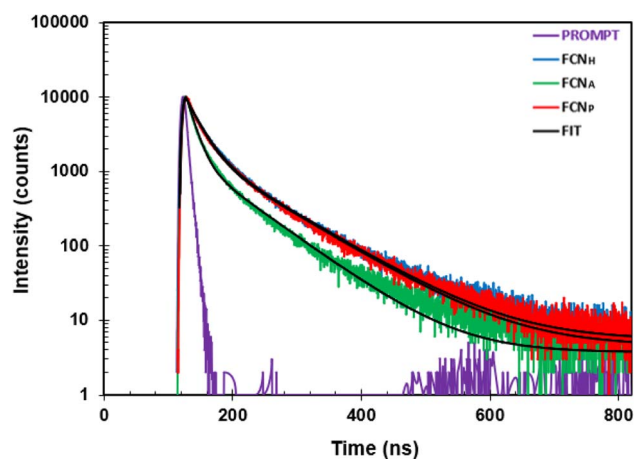
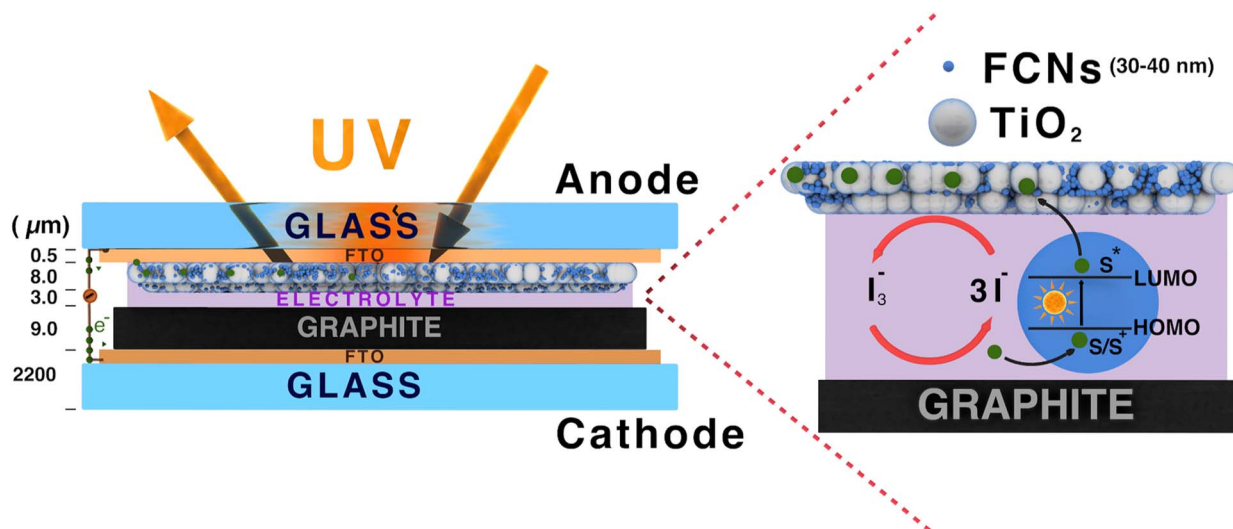


Fig. 5 Photoluminescence decay curves at a 420 nm emission wavelength of the FCNs.



Scheme 2 A schematic diagram of the proposed DSSC (not to scale).



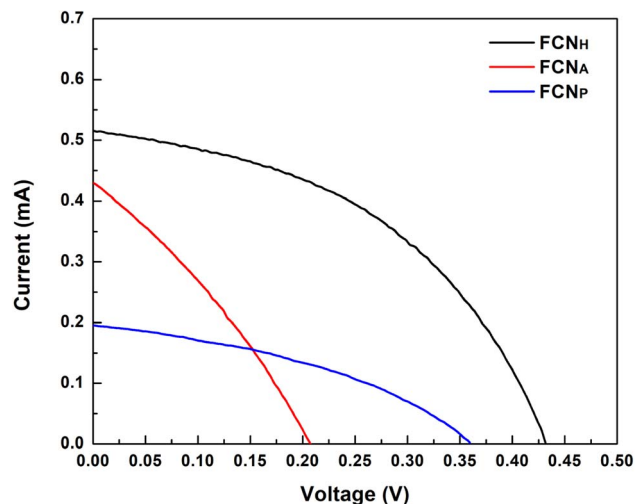


Fig. 6 Current–voltage characteristics of DSSCs fabricated from FCNs.

The current–voltage ( $I$ – $V$ ) measurement of the samples was carried out with a standard illumination of air-mass 1.5 global (AM 1.5 G) having an irradiance of  $100 \text{ mW cm}^{-2}$ . The  $I$ – $V$  values of the samples are used to calculate the solar-to-electric energy efficiency of the solar cells. The  $I$ – $V$  values of the measurements are displayed in Fig. 6. It was observed that  $\text{FCN}_\text{H}$  exhibits promising characteristics as it is evident from its open circuit voltage ( $V_{\text{OC}}$ ) of 0.43 V and short-circuit current ( $I_{\text{SC}}$ ) of 0.52 mA, thus demonstrating the successful functioning of the proposed DSSC (Table 2). Furthermore, the fill factor (FF) of 0.45 obtained from the fabricated DSSC ensures a good electrical contact and low series resistance. However, the power conversion efficiency (PCE) of the fabricated dye sensitized solar cell for  $\text{FCN}_\text{H}$  was 0.10%, offering further improvement in terms of light absorption and charge separation efficiency (Table 2).

### 3.2 Electrochemical impedance spectroscopy (EIS)

Electrochemical impedance spectroscopy (EIS) studies were carried out to understand the transport and recombination of the photoexcited electrons in the fabricated DSSC. The EIS studies expressed in Bode phase and Nyquist plots are displayed in Fig. 7 and 8, respectively.

The charge transfer process at the photoanode/electrolyte interface as well as the electron lifetime ( $\tau$ ) corresponds to the peak frequency in the Bode phase plot (Fig. 7). It is desirable to have a large electron lifetime as in the case of  $\text{FCN}_\text{H}$ . Additionally, the inclusion of  $\text{FCN}_\text{H}$  leads to a notable decrease in

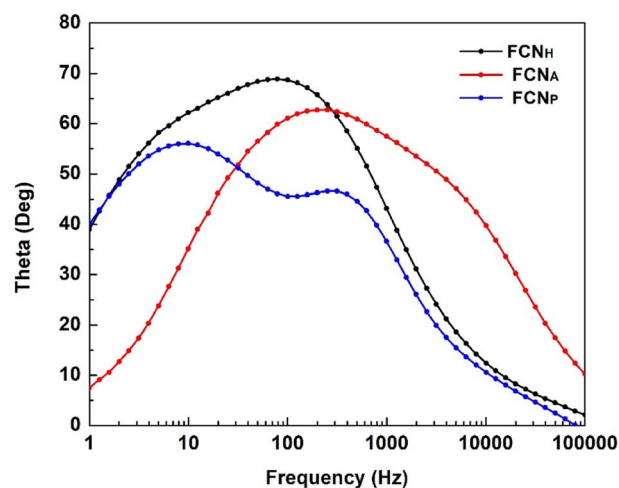


Fig. 7 Bode plot of the FCN DSSCs.

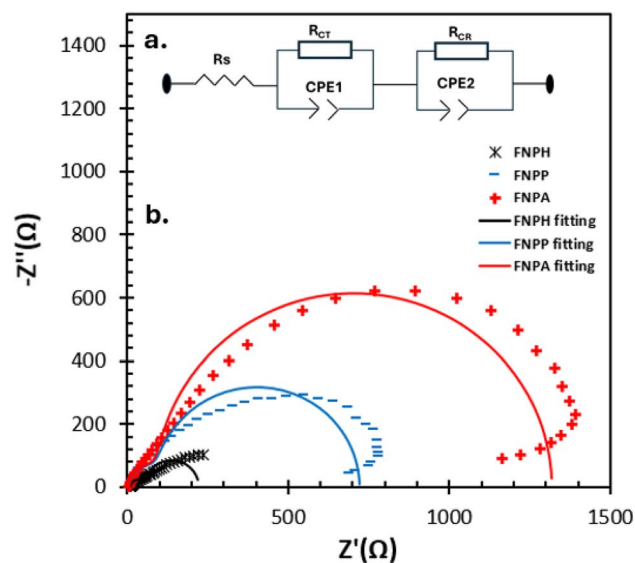


Fig. 8 (a) A full cell's equivalent circuit model and (b) matching Nyquist plots of  $\text{FCN}_\text{H}$ ,  $\text{FCN}_\text{A}$ , and  $\text{FCN}_\text{P}$ -based DSSCs in the dark at a bias of  $V_{\text{OC}}$  between 0.1 Hz and 100 Hz with an amplitude of 10 mV.

charge transfer resistance, thus resulting in the enhancement of the overall performance of the DSSC.

The effective electron lifetime ( $\tau$ ) in a DSSC can be determined by analyzing the maximum frequency in the Bode plot using eqn (1) (ref. 55 and 56)

$$\tau = \frac{1}{2\pi f_{\text{max}}} \quad (1)$$

Table 2 Current–voltage characteristics of FCNs in DSSCs

Sample	Electrolyte	$V_{\text{max}}$	$I_{\text{max}}$	$V_{\text{OC}}$ [V]	$I_{\text{SC}}$ [ $\text{mA cm}^{-2}$ ]	FF [%]	PCE [%]
$\text{FCN}_\text{H}$	$\text{I}^-/\text{I}_3^-$	0.29	0.35	0.43	0.52	0.45	0.10
$\text{FCN}_\text{P}$		0.11	0.25	0.21	0.43	0.31	0.03
$\text{FCN}_\text{A}$		0.22	0.12	0.36	0.20	0.39	0.03





Table 3 Calculated charge transfer resistance and lifetime parameters

Samples	$R_S$ ( $\Omega$ )	$R_{CT}$ ( $\Omega$ )	$R_{CR}$ ( $\Omega$ )	$\tau$ (ms)
FCN <sub>H</sub>	47.79	11.84	161.7	2.01
FCN <sub>P</sub>	132.7	6.398	582	0.504
FCN <sub>A</sub>	140.2	6.141	1173	0.632

The Bode plot analysis of FCN<sub>H</sub> reveals a peak at 79 Hz, corresponding to an estimated electron lifetime of 2 ms. These findings are depicted in Fig. 8 and summarized in Table 3. The longer electron lifetime observed in FCN<sub>H</sub> can be associated with a reduced occurrence of recombination within the TiO<sub>2</sub> network, thus resulting in an enhanced electron diffusion through lower trapping states. This phenomenon accelerates electron transfer, increases electron density, and ultimately improves the overall performance of the device.

The Nyquist plots of FCN<sub>H</sub>, FCH<sub>A</sub>, and FCH<sub>P</sub>-based DSSCs measured in the dark at forward bias ( $V_{OC}$ ) are displayed in Fig. 8b. Each spectrum exhibits two distinct semicircles: the high-frequency region on the left side corresponds to the charge transfer at the electrolyte/counter electrode interface, while the intermediate-frequency region on the right side represents the back charge transfer from the photoanode to the electrolyte.<sup>57,58</sup> Fig. 8a (inset) shows the equivalent circuit that was used to analyze the EIS spectra. The two interfaces were represented by a parallel combination of a capacitor and a resistance. The device's total series resistance ( $R_S$ ), charge transfer resistance ( $R_{CT}$ ) at the electrolyte/counter electrode interface, and charge recombination resistance ( $R_{CR}$ ) at the photoanode/electrolyte interface are depicted in the figure. By fitting an equivalent circuit, the values of  $R_S$ ,  $R_{CT}$ , and  $R_{CR}$  are obtained and shown in Table 3. In comparison to FCN<sub>H</sub> (47.79  $\Omega$ ), a large value of  $R_S$  (132.7 and 140.2  $\Omega$ ) in FCH<sub>A</sub> and FCH<sub>P</sub>, respectively, suggests a slower charge transport in these two contexts. Although the  $R_{CT}$  value in FCN<sub>H</sub> is marginally higher, all devices'  $R_{CT}$  values are consistent because the electrolyte and counter electrode were identical. In FCH<sub>A</sub>, a noticeably high charge recombination resistance (1173  $\Omega$ ) was found. This might be due to a gradual recombination in FCH<sub>A</sub>. The study on the  $I$ - $V$  curves of all the DSSCs indicates that FCN<sub>H</sub> has higher power conversion efficiency compared to FCN<sub>A</sub> and FCN<sub>P</sub>. The obtained charge transfer resistance ( $R_{CT}$ ) is much less in the case of FCN<sub>H</sub> than FCN<sub>A</sub> and FCN<sub>P</sub>, and therefore is clearly associated with the reduced charge recombination and consequently the enhanced electron lifetime due to the abundance of oxygen containing functional groups.<sup>59</sup>

## 4 Conclusion

In this study, we have demonstrated an easy H<sub>2</sub>O<sub>2</sub>-based FCN synthesis from an economical and available natural precursor teak leaf. H<sub>2</sub>O<sub>2</sub> not only increases the degree of surface oxidation of the carbon framework, but also improves their stability in aqueous solution. The prepared all FCNs showed blue emission fluorescence with high aqueous stability. Then, DSSCs were fabricated using FCN<sub>H</sub>, FCN<sub>A</sub> and FCN<sub>P</sub> as sensitizers and

their photovoltaic performance was evaluated. The FCN<sub>H</sub>-based DSSCs showed superior performance over FCN<sub>P</sub>-based and FCN<sub>A</sub>-based DSSCs, achieving an open circuit voltage ( $V_{OC}$ ) of 0.43 V, a short circuit current ( $I_{SC}$ ) of 0.52 mA, and a Fill Factor (FF) of 0.45. This enhanced performance of FCN<sub>H</sub>-based DSSCs was achieved through surface oxidation with H<sub>2</sub>O<sub>2</sub>. We anticipate that these bio-based FCNs could provide new insights into the fabrication process of DSSCs. Overall, the development of H<sub>2</sub>O<sub>2</sub>-assisted FCNs has opened new avenues for the application of carbon-based nanomaterials in various fields, and their potential uses are still being explored.

## Conflicts of interest

There are no conflicts to declare.

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