

Cite this: *Chem. Sci.*, 2024, 15, 19546

All publication charges for this article have been paid for by the Royal Society of Chemistry

Strain-induced charge delocalization achieves ultralow exciton binding energy toward efficient photocatalysis†

Junyuan Duan,^{‡ab} Yinghe Zhao,^{ID ‡a} Yu Wu,^a Youwen Liu,^{ID *a} Junnian Chen,^e Ruouo Yang,^a Jiazhao Huang,^a Chuanqi Luo,^a Mao Wu,^a Xiaodong Zheng,^d Pengyu Li,^a Xueliang Jiang,^{ID *b} Jianguo Guan^{ID *c} and Tianyou Zhai^{ID a}

The exciton effect is commonly observed in photocatalysts, where substantial exciton binding energy (E_b) significantly hampers the efficient generation of photo-excited electron–hole pairs, thereby severely constraining photocatalysis. Herein, we propose a strategy to reduce E_b through strain-induced charge delocalization. Taking Ta_2O_5 as a prototype, tensile strain was introduced by engineering a crystalline/amorphous interface, weakening the interaction between Ta 5d and O 2p orbitals, thus endowing a delocalized charge transport and significantly lowering E_b . Consequently, the E_b of strained Ta_2O_5 nanorods (s- Ta_2O_5 NRs) was reduced to 24.26 meV, below the ambient thermal energy (26 meV). The ultralow E_b significantly enhanced the yield of free charges, resulting in a two-fold increase in carrier lifetime and surface potential. Remarkably, the hydrogen evolution rate of s- Ta_2O_5 NRs increased 51.5 times compared to that of commercial Ta_2O_5 . This strategy of strain-induced charge delocalization to significantly reduce E_b offers a promising avenue for developing advanced semiconductor photoconversion systems.

Received 1st September 2024
Accepted 3rd November 2024

DOI: 10.1039/d4sc05873a

rsc.li/chemical-science

Introduction

Photocatalysis has emerged as a promising approach to tackle energy and environmental challenges by harnessing solar energy for conversion into chemical energy,^{1–6} exemplified by the production of hydrogen through photocatalytic water splitting.^{7,8} When a semiconductor catalyst absorbs photons with energy exceeding its bandgap energy (E_g), it excites electrons from the valence band (VB) to the conduction band (CB), while simultaneously generating positive holes in the VB (process I depicted in Fig. 1a).^{1,3,9} Furthermore, the coulombic

attraction between these electrons and holes can give rise to the formation of bound state excitons (process II depicted in Fig. 1a),^{10–13} known as the exciton effect, which is frequently overlooked. Given the neutral nature of excitons,^{14,15} their direct involvement in photocatalytic reactions is precluded; hence, it becomes imperative to mitigate the exciton effect for augmenting the abundance of free charges and thereby enhancing photocatalytic efficiency.^{16–18} According to the Saha–Langmuir equation $\frac{x^2}{1-x} = \frac{1}{n} \left(\frac{2\pi\mu k_B T}{h^2} \right)^{3/2} e^{-E_b/k_B}$ (where x , n , μ , T and h represent the free charge fraction, number of energetic particles, reduced mass of the exciton, temperature and Planck constant respectively),¹⁹ the exciton binding energy (E_b), a key parameter for evaluating the strength of the exciton effect, determines the ratio of free charges to excitons at a certain temperature and excitation density.

Ambient thermal energy ($k_B T \approx 26$ meV, where k_B is the Boltzmann constant and T is temperature) serves as a critical benchmark for evaluating exciton stability.²⁰ Excitons with E_b above this threshold exhibit greater stability and resistance to dissociation into free charges, whereas those with lower E_b are more susceptible to dissociation. To promote efficient exciton dissociation, researchers exert considerable efforts,^{9,21–25} such as typically employing charge spatial separation within heterojunctions, creating an internal electric field that reduces E_b and boosts free carrier generation.^{16,18,26–31} However, due to the

^aState Key Laboratory of Materials Processing and Die & Mould Technology, and School of Materials Science and Engineering, Huazhong University of Science and Technology, Wuhan, 430074, China. E-mail: ywliu@hust.edu.cn

^bHubei Key Laboratory of Plasma Chemistry and New Materials, School of Materials Science and Engineering, Key Laboratory of Green Chemical Engineering Process of Ministry of Education, Engineering Research Center of Phosphorus Resources Development and Utilization of Ministry of Education, Wuhan Institute of Technology, Wuhan, 430205, China. E-mail: jiangxl@wit.edu.cn

^cState Key Laboratory of Advanced Technology for Materials Synthesis and Processing, Wuhan University of Technology, Wuhan, 430070, China. E-mail: guanjq@whut.edu.cn

^dDepartment of Applied Physics, The Hong Kong Polytechnic University, Kowloon, Hong Kong, China

^eSchool of Materials Science & Engineering, Hubei University, Wuhan, 430062, China

† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d4sc05873a>

‡ These authors contributed equally to this work.

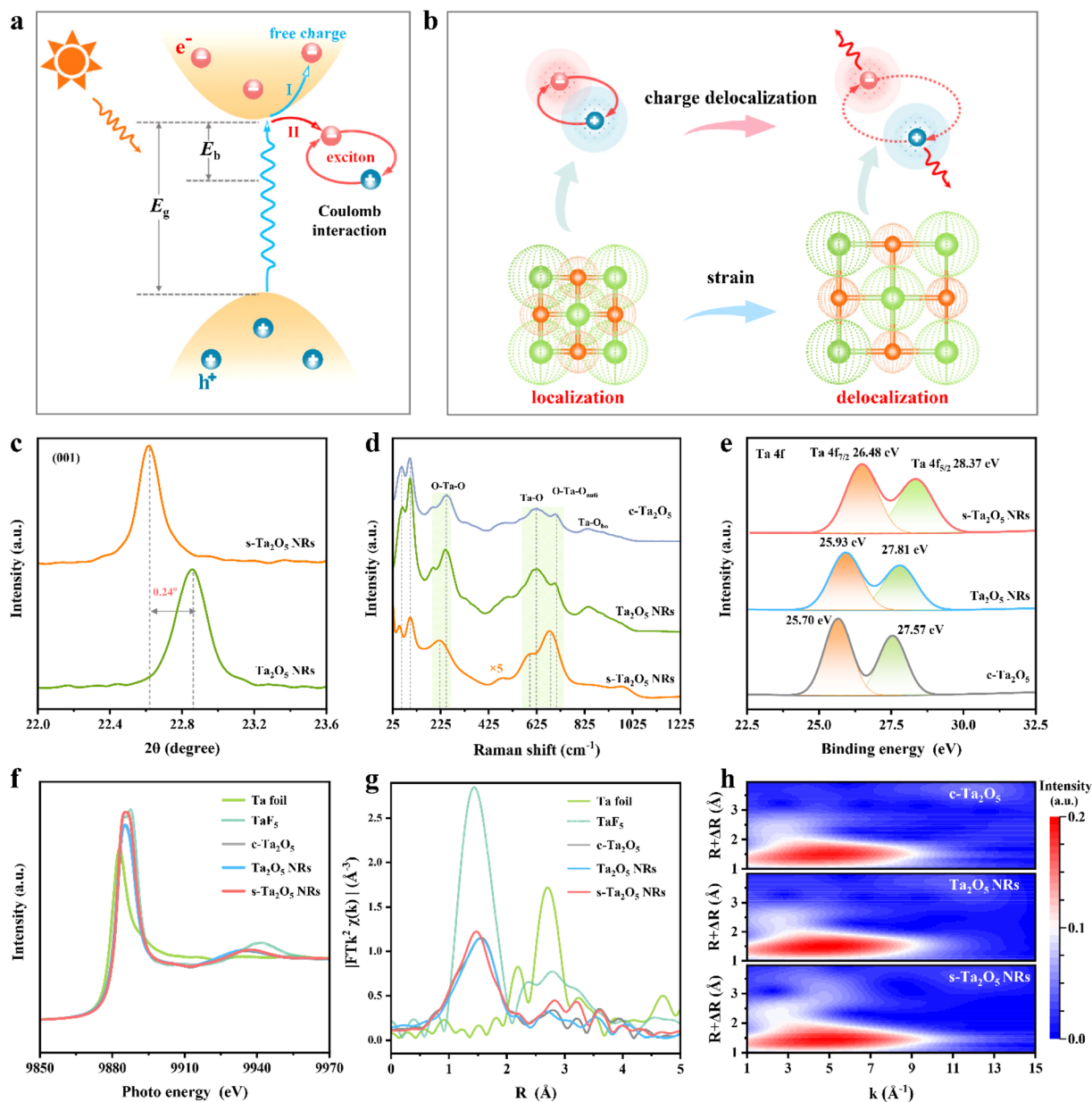


Fig. 1 Design scheme and structural characterization of s-Ta₂O₅ NRs. (a) Schematic of typical charge migration and the exciton effect in photocatalysis. (b) Strain-induced charge delocalization endows a delocalized charge transport pathway, minimizing the charge overlap between the electron and hole of the exciton and thus reducing E_b . (c) XRD patterns with an enlarged view of the (001) peak. (d) Raman spectra. (e) Ta 4f orbital XPS spectra. (f) Ta L3-edge XANES spectra. (g) k^2 -weighted Fourier transform of the EXAFS spectra. (h) Wavelet transform of the k^2 -weighted EXAFS.

significant charge localization inherent in the typical semiconductor catalyst lattice, which often leads to a strong interaction between the electron and hole in the resulting exciton, there are challenges in reducing the E_b below ambient thermal energy. Therefore, developing new strategies to reduce the E_b holds promise for achieving breakthroughs in photocatalytic performance.

Herein, by applying tensile strain to finely adjust charge localization and establish a delocalized transport pathway, we effectively reduce the E_b without relying on external fields or interfaces (Fig. 1b). Taking Ta₂O₅ as a prototype with

exceptional water splitting capability,^{32–34} we achieved charge delocalization by constructing a tensile strain in Ta₂O₅ nanorods (s-Ta₂O₅ NRs). Density functional theory (DFT) calculations further validate that this strain weakens the interaction between Ta 5d and O 2p orbitals, fostering a more delocalized charge transfer channel and significantly increasing the dielectric constant (ϵ_r), thereby achieving a reduced E_b . A tensile strain markedly reduces the E_b of s-Ta₂O₅ NRs to ~ 24.26 meV, below the thermal energy at room temperature ($k_B T$, 26 meV), achieving a reduction of 71.50% in comparison to Ta₂O₅ NRs. This reduction in E_b doubles the carrier lifetime and surface



potential, enhancing the separation of photogenerated charges, increasing the yield and separation rate of free charges, and lowering the hydrogen evolution reaction barrier. Consequently, this facilitates direct electron transfer from Ta atoms to protons, markedly boosting photocatalytic hydrogen production, which is 51.5 times higher than that of *c*-Ta₂O₅. This study demonstrates that by reducing the exciton binding energy through tensile strain, we can optimize charge transfer and promote photogenerated charge-driven photocatalytic reactions, offering a new avenue for the design and preparation of efficient semiconductor photocatalytic materials.

Results and discussion

Structural characterization of *s*-Ta₂O₅ NRs

Herein, the *s*-Ta₂O₅ NRs with a tensile strain were successfully fabricated by ingeniously engineering a crystalline/amorphous interface within them, achieved through the strategic insertion of F atoms into the spaces between the vertices of TaO₆ octahedra and TaO₇ polyhedra. X-ray diffraction (XRD) analysis, as depicted in Fig. 1c and S1,† clearly demonstrates that these *s*-Ta₂O₅ NRs (Fig. S2†) exhibit the characteristic orthorhombic Ta₂O₅, with an evident tensile strain along the [001] axis and an amorphous nature in their orthogonal directions. This was further supported by Raman spectra, which not only confirmed the orthorhombic Ta₂O₅ phase but also evidenced the presence of tensile strain and a level of crystalline integrity reduction in the *s*-Ta₂O₅ NRs (Fig. 1d and Note S1 in the ESI†). Furthermore, X-ray photoelectron spectra (XPS) revealed an increase in the binding energies of the Ta 4f_{5/2} (28.37 eV) and Ta 4f_{7/2} orbitals (26.37 eV) in *s*-Ta₂O₅ NRs, suggesting an oxidation state beyond the typical +5 state,³⁵ attributed to the F atoms' integration into the matrix (Fig. S3a and c†). Energy-dispersive X-ray spectroscopy (EDS) mapping (Fig. S4†) revealed a uniform distribution of Ta, O, and F elements throughout the nanorods, with a relatively weaker signal for F, indicating its minor presence. XPS quantified the atomic proportions (Fig. S3†) as 20.99% Ta, 52.7% O, and 11.85% F, demonstrating that the *s*-Ta₂O₅ NRs not only contain the typical Ta–O bonds but also Ta–F bonds, suggesting a unique interstitial incorporation of F within the lattice, rather than simple substitutional doping. O 1s XPS and electron spin resonance (ESR) measurements (Fig. S3b and S5†) further confirmed the widespread occurrence of oxygen defects.

To further elucidate the structural coordination of Ta atoms within *s*-Ta₂O₅ NRs, the synchrotron radiation X-ray absorption fine structure (XAFS) spectra were recorded, and compared with those of Ta₂O₅ NRs (Fig. S6†) and *c*-Ta₂O₅ (Fig. S7†). Notably, the X-ray absorption near edge structure (XANES) spectra at the Ta L3 edge show that the *s*-Ta₂O₅ NRs display an absorption edge closely aligned with that of the other Ta₂O₅ samples but with a pronounced white line peak (Fig. 1f), suggesting a slightly higher oxidation state of Ta, potentially beyond +5, consistent with that by Ta 4f XPS spectra. Further, Fourier transform analysis of the extended X-ray absorption fine structure (EXAFS) reveals a dominant peak at approximately 1.48 Å for *s*-Ta₂O₅ NRs, indicating a shorter bond length compared to the 1.56 Å Ta–O bond in Ta₂O₅ NRs and *c*-Ta₂O₅, yet longer than the typical

Ta–F bond of approximately 1.44 Å, attributed to the presence of Ta–F alongside Ta–O bonds (Fig. 1g). Additionally, this bond length is slightly longer than that observed in typical Ta₂O₅ samples, yet significantly shorter than the Ta–F bond, indicating a slight increase in coordination number and the incorporation of some fluorine into the *s*-Ta₂O₅ NRs lattice (Fig. 1g and S8†). Wavelet transform analysis further distinguishes the *s*-Ta₂O₅ NRs by showing similar Ta–O coordination within a 1–3 Å range to other samples, but with enhanced contours at 3–4 Å, indicative of a more distinct separation of Ta atoms, suggesting a degree of disordered structure induced by fluorine incorporation (Fig. 1h). According to the above XAFS results, the insertion of F atoms is inferred to have induced a significant degree of disordered structure within the *s*-Ta₂O₅ NRs lattice.

To examine the internal structure and strain of *s*-Ta₂O₅ NRs, high-resolution transmission electron microscopy (HRTEM) was performed. The HRTEM image of an individual *s*-Ta₂O₅ nanorod (Fig. 2a) clearly shows that it preferentially grows along the [001] axis, with an observed interplanar spacing of the (001) planes slightly increased to about 0.393 nm (Fig. S9†), a subtle yet notable expansion from the 0.388 nm spacing characteristic of the normal Ta₂O₅ crystal (PDF# 25-0922). Corresponding geometric phase analysis (GPA) further corroborates the presence of a uniform tensile strain of approximately 1.5% along this [001] axis (Fig. 2b). Only a set of bright diffraction spots manifested in the selected area electron diffraction (SAED) pattern (Fig. 2c), which affirms the single-crystal nature of the *s*-Ta₂O₅ NRs. Nevertheless, the occurrence of the ‘tail-dragging’ effect in the diffraction signals perpendicular to the [001] direction points to a significant structural disorder in the orthogonal plane. This is evidenced by high-resolution aberration-corrected high-angle annular dark-field (HAADF) scanning TEM images (STEM, Fig. 2d and e), which reveal a multitude of layer faults, including extra and missing atomic layers in the direction orthogonal to the [001]. Further HRTEM investigations (Fig. S10†) on additional *s*-Ta₂O₅ nanorods consistently demonstrate the aforementioned tensile strain along the [001] axis and a pronounced disordered structure in the orthogonal plane, consistent with the XRD and Raman spectra analyses.

To explore the atomic arrangement within the (001) plane of these special structured *s*-Ta₂O₅ NRs, focused ion beam (FIB) slices were made and then observed by aberration-corrected transmission electron microscopy. The resultant high-resolution HAADF-STEM image taken on the cross-section of an individual *s*-Ta₂O₅ nanorod (Fig. 2f and S11a†) clearly delineates a juxtaposition of both ordered and disordered atomic arrangement, alongside a multitude of discernible cavities. A closer inspection of the region labelled as ‘g’ in Fig. 2f clearly exhibits a highly ordered orthorhombic phase Ta₂O₅ atomic arrangement (Fig. 2g, S11b and S12†), with a slightly larger spacing between (200) planes (~0.317 nm) compared to the normal Ta₂O₅ crystal ($d(200) \approx 0.31$ nm). The corresponding GPA strain tensor mapping (Fig. 2h) further confirms the presence of a tensile strain of approximately 2% in the [200] direction. In stark contrast, the examination of region ‘i’ in



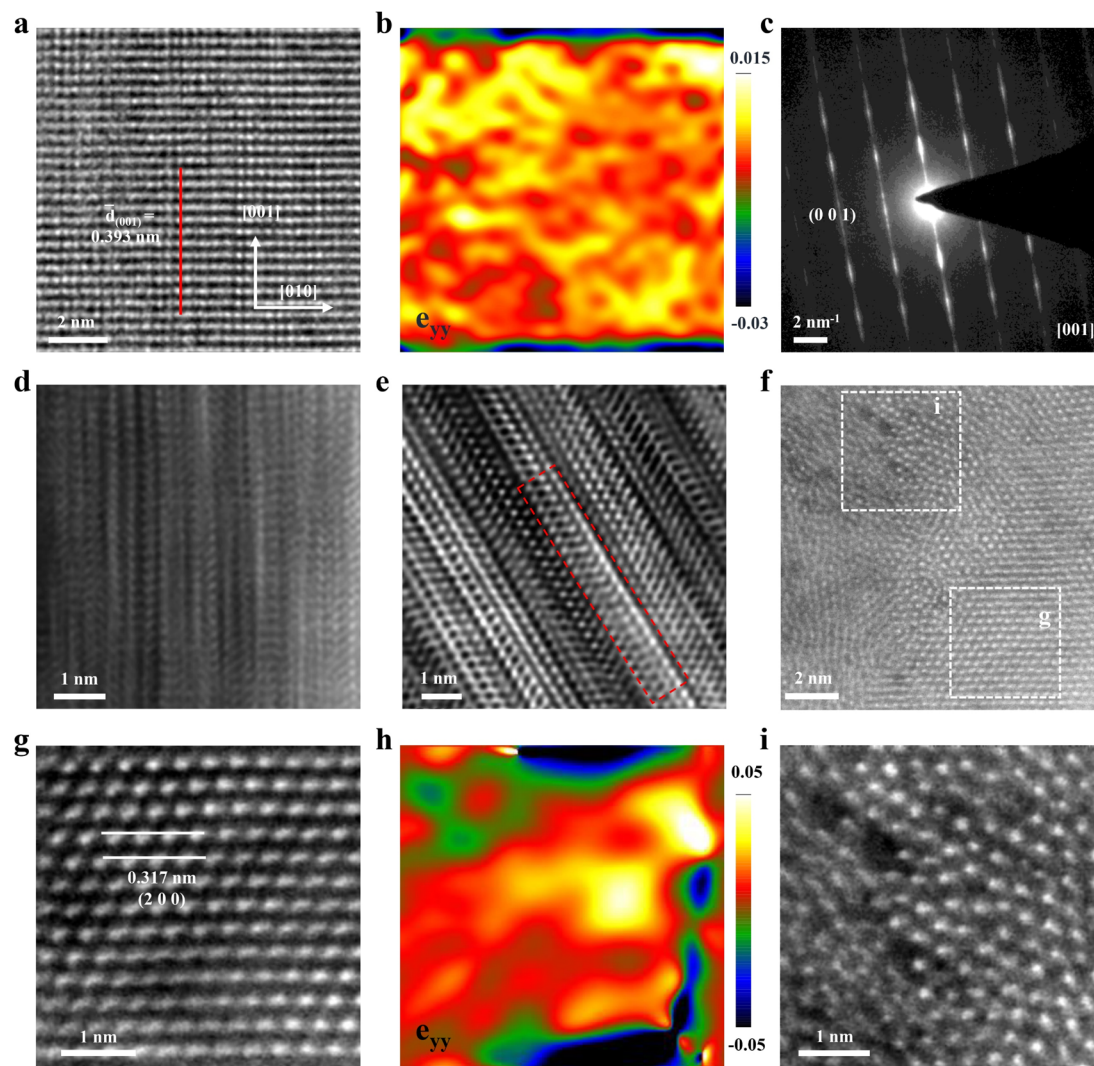


Fig. 2 Structural characterization of s-Ta₂O₅ NRs. (a) HRTEM image of the surface. (b) Corresponding e_{yy} strain tensor derived from (a). (c) SAED pattern on the surface. (d) and (e) Aberration-corrected HAADF-STEM images of the surface. (f) Aberration-corrected HAADF-STEM image of the cross-section. (g) and (i) Enlarged views of the areas g and i marked by rectangles in (f), respectively. (h) Corresponding e_{yy} strain tensor of (g).

Fig. 2f through an enlarged HAADF-STEM image (Fig. 2i) displays a starkly disordered atomic arrangement, replete with many noticeable cavities, indicative of an amorphous structure. More HAADF-STEM images (Fig. S13[†]) randomly sampled from the cross-section of another s-Ta₂O₅ nanorod evidence the repetitive nature of this intricate structural motif, comprising both crystalline and amorphous domains, as well as the consistent occurrence of tensile strain and cavities. Collectively, these observations substantiate that the s-Ta₂O₅ NRs possess a distinct single-crystal growth pattern along the [001] axis, seamlessly incorporating a blend of ordered and amorphous structures in the perpendicular orientation, and are characterized by the tensile strain within the crystalline regions of Ta₂O₅, alongside the pervasive presence of cavities.

Ultraviolet absorption and valence band spectra (Fig. S14[†]) have established s-Ta₂O₅ NRs as a prototypical semiconductor, with a similar band structure to both Ta₂O₅ NRs and c-Ta₂O₅ samples. It manifests a bandgap (E_g) of 4.10 eV, with the valence

band maximum (VBM) and the conduction band minimum (CBM) situated at 3.21 eV and −0.89 eV respectively. Theoretical simulations show that F-doping in the Ta₂O₅ crystal introduces the density of states near the Fermi level that cross the Fermi surface, yielding a metallic band structure (Fig. S15a–c[†]). However, this contradicts the experimental semiconductor band structure of s-Ta₂O₅ NRs (Fig. S14[†]), indicating that F does not contribute to the band structure of s-Ta₂O₅ NRs. Instead, applying a tensile strain along the [001] axis of the Ta₂O₅ crystal maintains the semiconductor band structure (Fig. S15d[†]), with only slight shifts in the band edges, consistent with experimental observations. This suggests that the subtle alterations in the band structure of s-Ta₂O₅ NRs are mainly attributed to strain effects rather than fluorine doping. These insights, along with the previous TEM, XAS, XPS, XRD, Raman, and EDS analyses, sufficiently verify that the crystal regions under tensile strain consist of pure Ta₂O₅, while fluorine is mainly located in the amorphous areas. Notably, the orthorhombic Ta₂O₅ crystal



structure is formed by corner-sharing or edge-sharing links between TaO_6 octahedra and TaO_7 polyhedra, featuring a distinctive interspace at the junction where a TaO_6 octahedron and a TaO_7 polyhedron share corners (Fig. S16†).³⁶ This interspace, characterized by a lower packing density of oxygen atoms, offers a potential site for F insertion. Given the smaller size and higher electronegativity of F atoms relative to oxygen, they readily occupy these interspaces and form Ta–F bonds, as confirmed by XAS and XPS analyses (Fig. 1e–h and S3†). The insertion of F into the Ta_2O_5 lattice would push away surrounding atoms, culminating in lattice expansion. This lattice expansion effect instigated by F insertion impacts neighboring Ta_2O_5 crystalline regions, engendering tensile strain. Consequently, the synthesized s- Ta_2O_5 NRs exhibit an ordered growth along the [001] direction, alternating between pure Ta_2O_5 crystalline regions and F-doped amorphous areas, containing tensile strain and associated cavities.

Notably, although Ta_2O_5 NRs maintain the same size and morphology as the s- Ta_2O_5 NRs after annealing at 750 °C (Fig. S6†), their XRD pattern and Raman spectrum (Fig. 1c and

d) indicate that the tensile strain disappears. Furthermore, HRTEM images, captured with the electron beam directed along various axes of the Ta_2O_5 NR (Fig. S17†), display well-defined lattice fringes consistent with the orthorhombic Ta_2O_5 crystal structure. This is supported by the presence of sharp diffraction spots, free of the tail-dragging effect, in the corresponding FFT images, suggesting a well-crystallized single-crystal structure. GPA strain tensor mapping further verifies the absence of strain in the Ta_2O_5 NRs (Fig. S18†). After annealing, fluorine is eliminated (Fig. 1g and S3†), and the previously fluorine-doped Ta_2O_5 amorphous regions are converted into an ordered crystalline pure Ta_2O_5 phase (Fig. 1c, d, S3, S6, and S17†), resulting in the removal of tensile strain. Evidently, the creation of localized amorphous regions is the primary cause of the tensile strain in s- Ta_2O_5 NRs.

Effect of tensile strain on E_b in s- Ta_2O_5 NRs

To evaluate the effect of tensile strain on E_b , temperature-dependent photoluminescence (PL) spectroscopy was

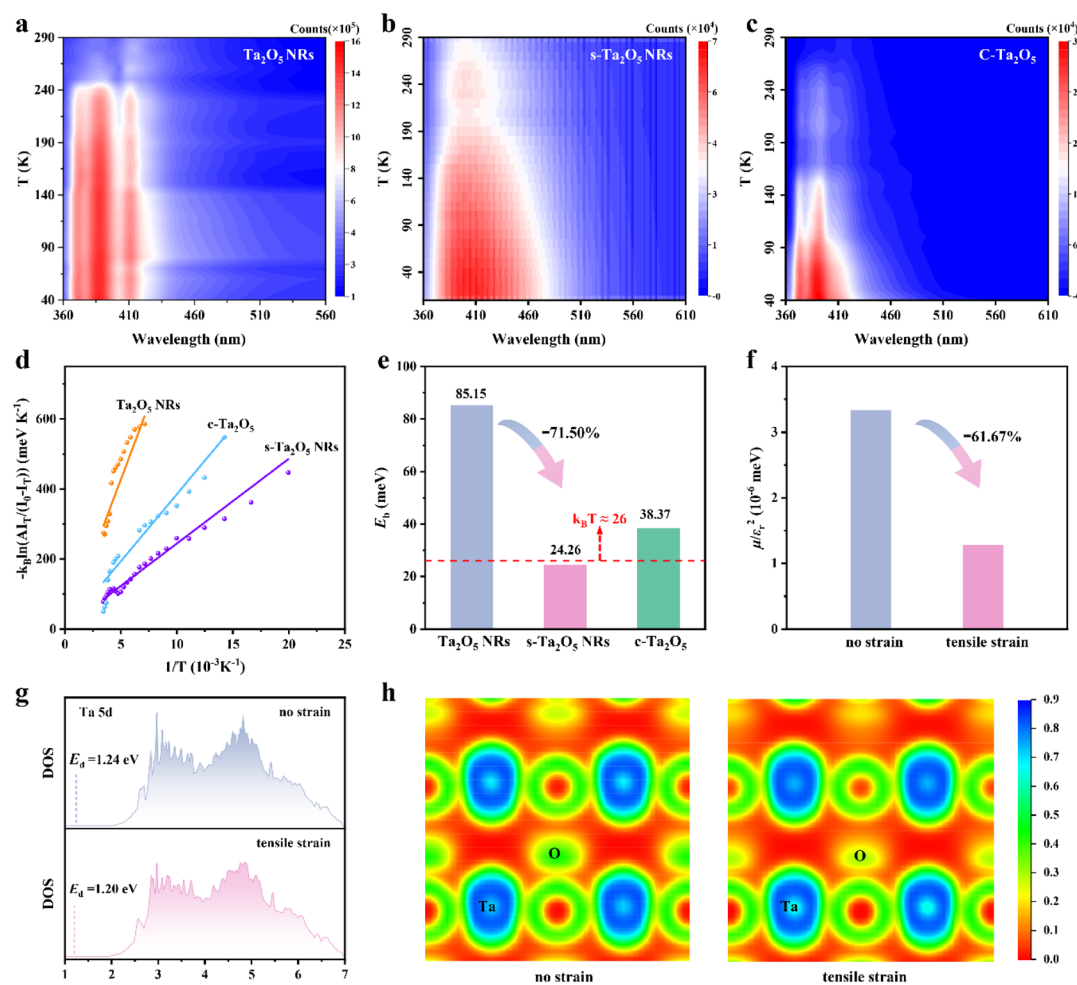


Fig. 3 The effect of tensile strain on the E_b of s- Ta_2O_5 NRs. Temperature-dependent PL spectra of (a) Ta_2O_5 NRs, (b) s- Ta_2O_5 NRs and (c) c- Ta_2O_5 . (d) Integrated PL emission intensity as a function of temperature. (e) Comparative E_b values for s- Ta_2O_5 NRs, Ta_2O_5 NRs, and c- Ta_2O_5 . (f) Calculated μ/ϵ_r^2 values. (g) Ta 5d PDOS and corresponding d-band centers. (h) ELF plots for Ta_2O_5 without (left) and with (right) a tensile strain along [001].



meticulously performed. The PL spectra from all samples (Fig. 3a–c) exhibit a consistent increase in PL peak intensity as the temperature decreases, attributable to phonon-assisted relaxation. Accordingly, the E_b values for various Ta_2O_5 samples were determined by employing the Arrhenius equation $I_T = I_0/(1 + A \exp(-E_b/k_B T))$ to fit the acquired PL data (Fig. 3d and S19†).¹⁸ As stated within semiconductor physics, a reduction in size would weaken the dielectric screening effect, thereby increasing the exciton effect.^{15,37,38} Indeed, Ta_2O_5 NRs exhibit a larger E_b of 85.15 meV compared to c- Ta_2O_5 , which has an E_b of 38.37 meV (Fig. 3e).

However, introducing tensile strain results in a substantial decrease in the E_b for s- Ta_2O_5 NRs (which have the same size and morphology as Ta_2O_5 NRs, Fig. S2 and S6†) to 24.26 meV, which is significantly lower than that of both the unstrained Ta_2O_5 NRs and c- Ta_2O_5 samples, and even below ambient thermal energy ($k_B T \approx 26$ meV), suggesting an easy dissociation of excitons into free charges during photocatalysis. According to the equation $E_b = \mu E_0/m_0 \epsilon_r^2$, E_b is predominantly determined by the variables μ and ϵ_r .^{39,40} To further examine the effect of strain on E_b , the μ and ϵ_r values for Ta_2O_5 crystal models (Fig. S20†), both with and without a tensile strain along the [001] direction, were calculated (Fig. S21†). Fig. 3f illustrates that a tensile strain along the [001] direction markedly causes a 61.67% decrease in the value of μ/ϵ_r^2 , aligning well with the experimentally observed trends in E_b variation. Both experimental and theoretical results robustly support the assertion that tensile strain can significantly lower the E_b in Ta_2O_5 photocatalysts, which in turn weakens the coulombic interaction between photoexcited

electrons and holes, enhancing charge separation efficiency, liberating more free charges, and promoting charge-driven photocatalytic reactions.

To explore the effect of tensile strain on E_b , we performed DFT calculations to assess the density of states (DOS) and electron localization function (ELF) in Ta_2O_5 , both with and without the application of tensile strain. The d-band center theory offers an effective method for evaluating the bonding conditions in transition metals.^{41–43} As illustrated in Fig. 3g, there is a discernible reduction in the partial density of states (PDOS) for Ta 5d above the Fermi level. Concurrently, there is a slight shift in the d-band center (E_d) from 1.24 eV to 1.20 eV. These observations, combined with the O 2p PDOS (Fig. S22†) suggest that tensile strain could weaken the antibonding interactions between Ta and O atoms, consequently decreasing the overlap between the Ta 5d and O 2p orbitals. The ELF offers an effective description of electron delocalization in solids,^{44,45} serving as a valuable tool for the chemical bonds' classification.^{46,47} Further scrutiny of the ELF contour plots on the (001) crystal plane (Fig. 3h and S23†) reveals a reduction in ELF values near oxygen atoms under a tensile strain along the [001] axis, indicating a lower charge density and increased electron delocalization in these areas. These observations collectively imply that tensile strain reduces the interaction between Ta 5d and O 2p orbitals, promoting a more delocalized pathway for efficient charge transport.⁴⁸ This, in turn, lessens the coulombic interaction between photoexcited electrons and holes in excitons, thereby lowering the E_b .

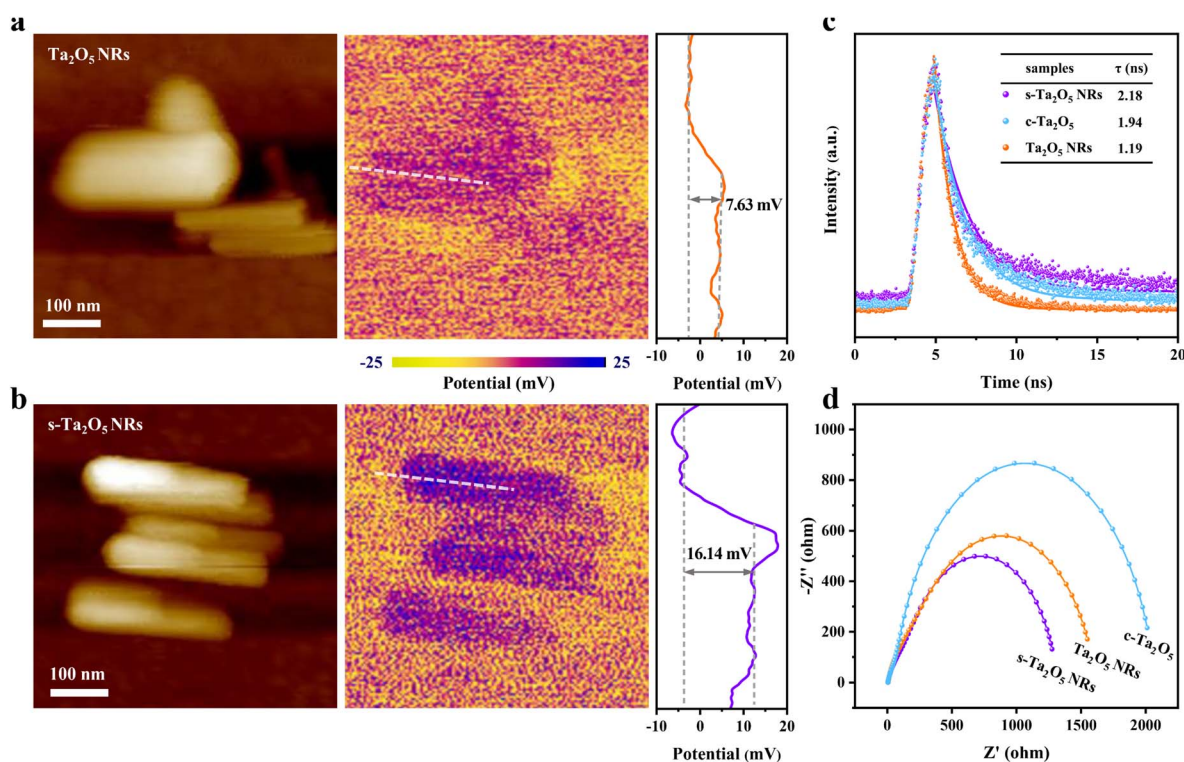


Fig. 4 The effect of tensile strain on charge transport in s- Ta_2O_5 NRs. Atomic force microscopy images (left) and the corresponding surface potentials measured with KPFM (right) of (a) Ta_2O_5 NRs and (b) s- Ta_2O_5 NRs. (c) Time-resolved transient PL decay. (d) EIS spectra.



Effect of tensile strain on charge transport

Reducing E_b facilitates the separation of photoexcited charge carriers, providing more free charge carriers for the photocatalytic reactions. To confirm this point, the influence of tensile strain on the charge transport was studied using Kelvin probe force microscopy (KPFM). Fig. 4a and b reveal that the surface potential of s-Ta₂O₅ NRs is considerably two times higher than that for unstrained Ta₂O₅ NRs, suggesting a lower work function and Fermi level, which are conducive to electron separation. Moreover, the influence of tensile strain on the lifetime of photo-generated charge carriers was further investigated using advanced time-resolved PL spectroscopy. In Fig. 4c, the PL lifetime of the s-Ta₂O₅ NRs reaches up to 2.18 ns, notably surpassing that for both Ta₂O₅ NRs (1.19 ns) and c-Ta₂O₅ (1.94

ns). This is indicative of the tensile strain's efficacy in diminishing electron-hole recombination rates within the s-Ta₂O₅ NRs, in line with the observed trend in the reduction of E_b . In parallel, electrochemical impedance spectra (EIS) provided additional evidence, where the s-Ta₂O₅ NRs exhibit the smallest semicircle in the Nyquist plots (Fig. 4d), signifying exceptional charge transfer capabilities and the most efficient charge migration when benchmarked against Ta₂O₅ NRs and c-Ta₂O₅. Collectively, these experimental results robustly validate that the introduction of tensile strain substantially diminishes exciton binding energy, thereby extending the lifetime of photo-generated charges and promoting effective charge separation and migration, which is highly beneficial for photogenerated charge-involved photocatalytic reactions.

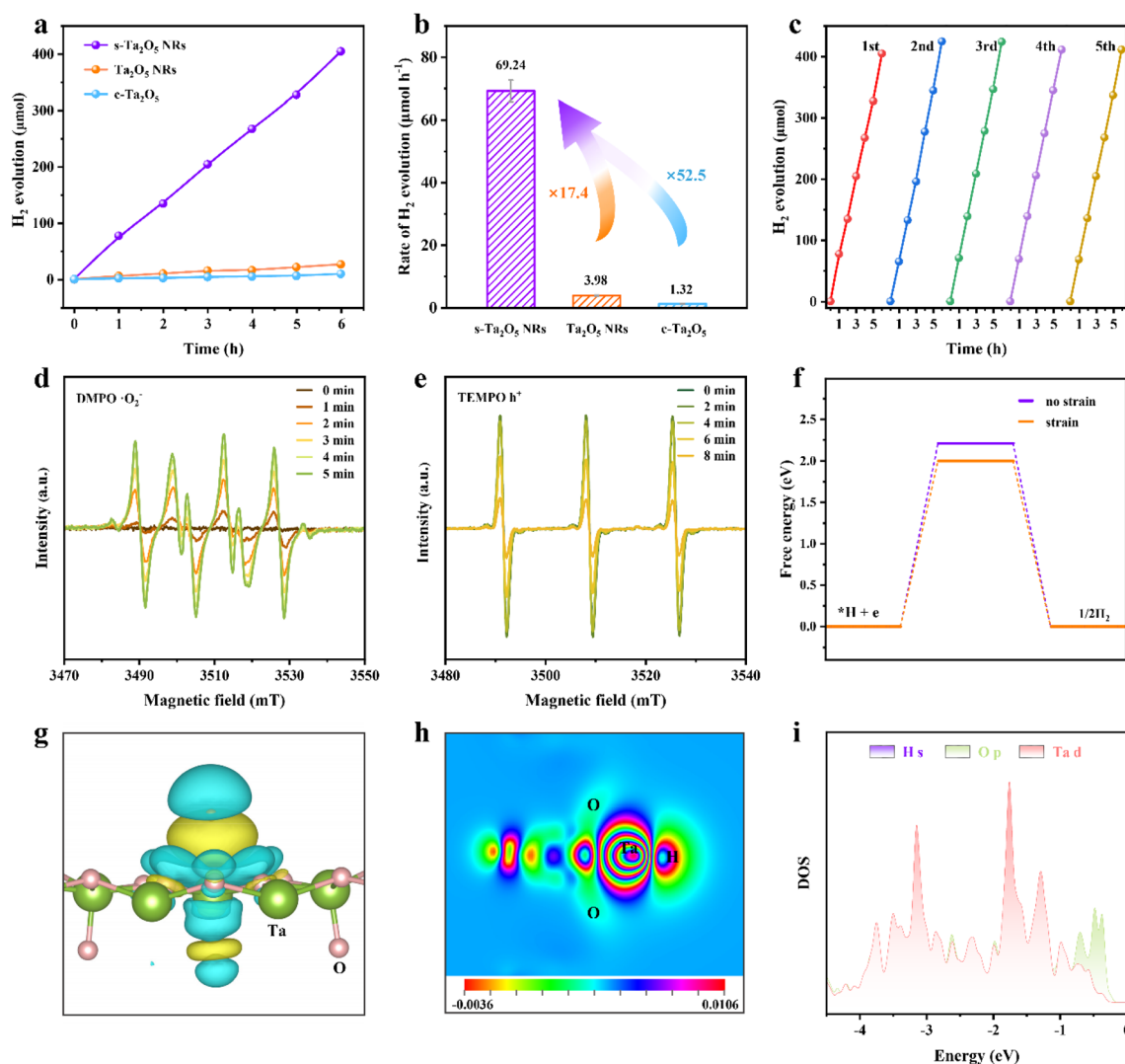


Fig. 5 Photocatalytic water splitting performances and mechanism of s-Ta₂O₅ NRs. (a) Photocatalytic H₂ production from water splitting for s-Ta₂O₅ NRs, Ta₂O₅ NRs, and c-Ta₂O₅ under UV light irradiation. (b) Comparative photocatalytic hydrogen evolution rates for s-Ta₂O₅ NRs, Ta₂O₅ NRs, and c-Ta₂O₅, respectively. (c) Repeated hydrogen generation performance of s-Ta₂O₅ NRs. *In situ* ESR spectra of (d) DMPO-trapped $\cdot\text{O}_2^-$ radicals and (e) TEMPO-trapped h^+ in s-Ta₂O₅ NRs. (f) Gibbs free energy diagram for H₂ production through water splitting. (g) Charge density difference of H on Ta₂O₅ with a tensile strain along the [001] direction. The isosurface level is 0.00025 au. The yellow and blue shadows indicate charge accumulation and depletion, respectively. (h) Corresponding slice image of (g) across the (010) face. (i) PDOS of H 1s, Ta 5d, O 2p orbitals with H adsorbed on Ta₂O₅ with a tensile strain along the [001] direction.



Photocatalytic activity

To validate that tensile strain-induced decrease in E_b can be leveraged to substantially promote charge separation, photocatalytic water splitting was further evaluated. Hydrogen production from photocatalytic water splitting using s-Ta₂O₅ NRs under ultraviolet illumination in a methanol/water solution was conducted and compared with that of Ta₂O₅ NRs and c-Ta₂O₅. Fig. 5a exhibits a direct proportional relationship between the duration of light illumination and the amount of hydrogen evolved, indicating a consistent photocatalytic activity over various Ta₂O₅ samples. Notably, s-Ta₂O₅ NRs achieve a hydrogen production rate of 69.24 $\mu\text{mol h}^{-1}$, significantly surpassing that of Ta₂O₅ NRs (3.98 $\mu\text{mol h}^{-1}$) and c-Ta₂O₅ (1.32 $\mu\text{mol h}^{-1}$), representing an increase by approximately 16.4 and 51.5 times, respectively (Fig. 5b). No hydrogen was detected in the absence of either irradiation or photocatalysts, suggesting H₂ production solely attributable to the photocatalytic reaction on the Ta₂O₅ catalysts. The s-Ta₂O₅ NRs demonstrate a steady linear rise in hydrogen production over five consecutive cycles (Fig. 5c), indicating robust catalytic activity and stability. Subsequent SEM and XRD analyses post-catalysis show no notable changes in the morphology or crystal structure (Fig. S24†), corroborating the structural integrity of s-Ta₂O₅ NRs throughout the photocatalytic reaction. The structural characterization above suggests that s-Ta₂O₅ NRs have comparable structure sizes (the specific surface areas of s-Ta₂O₅ NRs and Ta₂O₅ NRs were measured at approximately 15.8 and 11.9 m² g⁻¹ respectively) and nearly identical energy bands to Ta₂O₅ NRs. Moreover, no significant hydrogen was detected when both F-doped and undoped amorphous Ta₂O₅ samples were used as catalysts. Additionally, the F element is solely found in the amorphous region in s-Ta₂O₅ NRs; these areas will not be involved in photo-excitation and production of photogenerated charges. Taken together, these results emphatically validate that the resultant tensile strain in s-Ta₂O₅ NRs effectively decreases exciton binding energy, thereby enhancing charge separation and ultimately leading to an exceptional enhancement of photocatalytic activity.

To delve into the underlying photocatalytic mechanisms, *in situ* ESR spectroscopy was employed to monitor the reactive species throughout the photocatalytic process. 2,2,6,6-Tetramethyl-1-piperidinyloxy (TEMPO) and 5,5-dimethyl-1-pyrroline-*N*-oxide (DMPO) radicals as scavengers were utilized for the direct observation of photogenerated holes and superoxide radicals ($\cdot\text{O}_2^-$) respectively. The DMPO signal, absent under dark conditions, markedly increases with light exposure and continues to rise over time (Fig. 5d), indicating a steady production of photogenerated electrons. This is further corroborated by the detection of photogenerated $\cdot\text{OH}$ (Fig. S25†). In contrast, the TEMPO signal diminishes when exposed to light and further decreases with prolonged illumination (Fig. 5e), suggesting the sustained formation and accumulation of photogenerated holes. Further, DFT calculations reveal that the tensile strain can lower the Gibbs free energy for hydrogen production, thereby lowering the energy barrier for both water molecule adsorption and hydrogen gas release

(Fig. 5f), facilitating the water-splitting reaction. Charge density difference calculation shows a reduced charge on Ta atoms and a heightened charge around H atoms, indicating a direct electron transfer from Ta to H atoms (Fig. 5g and h). PDOS analysis further confirmed a robust interaction between Ta 5d and H 1s orbitals (Fig. 5i). Integrating ESR analyses with DFT calculations, it is inferred that the introduction of tensile strain significantly lowers E_b , endows a strong interaction between Ta 5d and H 1s orbitals, and reduces the Gibbs free energy of hydrogen production, thereby facilitating efficient electron transfer for water reduction and hydrogen generation, as well as sustaining photogenerated holes for the photocatalytic oxidation reactions.

Conclusions

In summary, we have successfully fabricated s-Ta₂O₅ single crystal nanorods with tensile strain by ingeniously engineering a crystalline/amorphous interface within them through the strategic insertion of F atoms into the interspaces shared by the vertices of TaO₆ octahedra and TaO₇ polyhedra. This strain weakens the interaction between Ta 5d and O 2p orbitals in s-Ta₂O₅ NRs, thereby reducing charge localization and building a more delocalized charge transport channel. This significantly decreases the E_b to below ambient thermal energy (~ 26 meV), marking a 71.50% decrease compared to their unstrained counterparts. A lower E_b contributes to enhanced carrier lifetime and surface potential, improved charge separation efficiency, and a lowered hydrogen evolution reaction barrier, thereby endowing direct electron transfer from Ta atoms to protons. As a result, the s-Ta₂O₅ NRs exhibit a markedly improved photocatalytic hydrogen production activity, with generation rates surpassing those of Ta₂O₅ NRs and c-Ta₂O₅ by 16.4 and 51.5 times, respectively. The study not only offers an effective strategy to reduce exciton binding energy through strain-induced charge delocalization, but also paves a novel avenue for the design of highly efficient photoconversion systems.

Data availability

All experimental procedures, computational sections, and associated data are included in the article and ESI.†

Author contributions

J. D., Y. L., and T. Z. conceived and directed the project. J. D. designed and executed the experiments. J. D. and Y. L. co-wrote the paper. J. D., Y. W., Y. Z., and Y. L. performed and analyzed the DFT calculations. J. D. and Y. L. conducted and interpreted the XAFS spectroscopy. J. D. and C. L. collaboratively conducted PL. J. D., J. H., and Y. Z. jointly performed AFM and KPFM. J. D. and D. Z. carried out and discussed the GPA. J. D., Y. L., X. J., J. G., and T. Z. revised the paper. J. C. participated in the discussion of the catalytic mechanism. All authors discussed the results and contributed to the manuscript preparation.



Conflicts of interest

The authors declare no competing financial interest.

Acknowledgements

This work was financially supported by the National Natural Science Foundation of China (22071069, 22173048, and 22103026), the Project Funded by the China Postdoctoral Science Foundation (2019M662604), and the Scientific Research Fund Project of Wuhan Institute of Technology (no. 23QD06). We thank the Analytical and Testing Center at Huazhong University of Science and Technology for their technical support, especially Dr Jianquan Zhao. We also acknowledge the support of the HPC platform at HUST. Furthermore, the authors greatly appreciate the beam time allocation and assistance provided by beamline BL11B at the Shanghai Synchrotron Radiation Facility.

References

- 1 T. Hisatomi and K. Domen, *Nat. Catal.*, 2019, **2**, 387–399.
- 2 W. H. Lee, C. W. Lee, G. D. Cha, B.-H. Lee, J. H. Jeong, H. Park, J. Heo, M. S. Bootharaju, S.-H. Sunwoo, J. H. Kim, K. H. Ahn, D.-H. Kim and T. Hyeon, *Nat. Nanotechnol.*, 2023, **18**, 754–762.
- 3 X. P. Tao, Y. Zhao, S. Y. Wang, C. Li and R. G. Li, *Chem. Soc. Rev.*, 2022, **51**, 3561–3608.
- 4 J. Y. Y. Loh, N. P. Kherani and G. A. Ozin, *Nat. Sustainability*, 2021, **4**, 466–473.
- 5 S. P. Che, J. H. Zhang, F. Z. Mou, X. Guo, J. E. Kauffman, A. Sen and J. G. Guan, *Research*, 2022, **2022**, 9816562.
- 6 M. Z. Li, J. Z. Wang, Q. Y. Wang, H. L. Lu, G. F. Wang and H. G. Fu, *Chem. Sci.*, 2024, **15**, 1860–1869.
- 7 P. Zhou, I. A. Navid, Y. Ma, Y. Xiao, P. Wang, Z. Ye, B. Zhou, K. Sun and Z. Mi, *Nature*, 2023, **613**, 66–70.
- 8 J. Y. Duan, G. Y. Mou, S. Zhang, S. Wang and J. G. Guan, *J. Mater. Chem. A*, 2015, **3**, 14686–14695.
- 9 Q. Wang, C. Pornrungraj, S. Linley and E. Reisner, *Nat. Energy*, 2022, **7**, 13–24.
- 10 B. Jin, N. Zuo, Z. Y. Hu, W. J. Cui, R. Y. Wang, G. Van Tendeloo, X. Zhou and T. Y. Zhai, *Adv. Funct. Mater.*, 2020, **30**, 2006166.
- 11 H. Wang, W. X. Liu, X. He, P. Zhang, X. D. Zhang and Y. Xie, *J. Am. Chem. Soc.*, 2020, **142**, 14007–14022.
- 12 S. Kang, K. Kim, B. H. Kim, J. Kim, K. I. Sim, J. U. Lee, S. Lee, K. Park, S. Yun, T. Kim, A. Nag, A. Walters, M. Garcia-Fernandez, J. Li, L. Chapon, K. J. Zhou, Y. W. Son, J. H. Kim, H. Cheong and J. G. Park, *Nature*, 2020, **583**, 785–789.
- 13 Q. C. Burlingame, X. Liu, M. L. Ball, B. P. Rand and Y. L. Loo, *Energy Environ. Sci.*, 2023, **16**, 1742–1751.
- 14 J. Pradhan, A. Das, K. Kundu, Chahat and K. Biswas, *Chem. Sci.*, 2022, **13**, 9952–9959.
- 15 P. Chen, T. L. Atallah, Z. Y. Lin, P. Q. Wang, S. J. Lee, J. Q. Xu, Z. H. Huang, X. D. Duan, Y. Ping, Y. Huang, J. R. Caram and X. F. Duan, *Nature*, 2021, **599**, 404–410.
- 16 Y. Qian, Y. Han, X. Zhang, G. Yang, G. Zhang and H.-L. Jiang, *Nat. Commun.*, 2023, **14**, 3083.
- 17 C. Li, J. Liu, H. Li, K. Wu, J. Wang and Q. Yang, *Nat. Commun.*, 2022, **13**, 2357.
- 18 Z.-A. Lan, G. Zhang, X. Chen, Y. Zhang, K. A. I. Zhang and X. Wang, *Angew. Chem., Int. Ed.*, 2019, **58**, 10236–10240.
- 19 M. B. Price, P. A. Hume, A. Ilina, I. Wagner, R. R. Tamming, K. E. Thorn, W. T. Jiao, A. Goldingay, P. J. Conaghan, G. Lakhwani, N. Davis, Y. F. Wang, P. Y. Xue, H. Lu, K. Chen, X. W. Zhan and J. M. Hodgkiss, *Nat. Commun.*, 2022, **13**, 2827.
- 20 M. H. Huang, S. Mao, H. Feick, H. Q. Yan, Y. Y. Wu, H. Kind, E. Weber, R. Russo and P. D. Yang, *Science*, 2001, **292**, 1897–1899.
- 21 A. Raja, L. Waldecker, J. Zipfel, Y. Cho, S. Brem, J. D. Ziegler, M. Kulig, T. Taniguchi, K. Watanabe, E. Malic, T. F. Heinz, T. C. Berkelbach and A. Chernikov, *Nat. Nanotechnol.*, 2019, **14**, 832–837.
- 22 Y. Lei, Y. Li, C. Lu, Q. Yan, Y. Wu, F. Babbe, H. Gong, S. Zhang, J. Zhou, R. Wang, R. Zhang, Y. Chen, H. Tsai, Y. Gu, H. Hu, Y.-H. Lo, W. Nie, T. Lee, J. Luo, K. Yang, K.-I. Jang and S. Xu, *Nature*, 2022, **608**, 317–323.
- 23 A. Raja, A. Chaves, J. Yu, G. Arefe, H. M. Hill, A. F. Rigosi, T. C. Berkelbach, P. Nagler, C. Schueller, T. Korn, C. Nuckolls, J. Hone, L. E. Brus, T. F. Heinz, D. R. Reichman and A. Chernikov, *Nat. Commun.*, 2017, **8**, 15251.
- 24 Y. B. Shi, J. Li, C. L. Mao, S. Liu, X. B. Wang, X. F. Liu, S. X. Zhao, X. Liu, Y. Q. Huang and L. Z. Zhang, *Nat. Commun.*, 2021, **12**, 5923.
- 25 N. Kornienko, J. Z. Zhang, K. K. Sakimoto, P. D. Yang and E. Reisner, *Nat. Nanotechnol.*, 2018, **13**, 890–899.
- 26 B. Wu, J. Z. Fan, J. Y. Han, Y. Su, M. P. Zhuo, J. H. Sun, Y. Gao, S. Chen, J. J. Wu, Z. S. Wang and X. D. Wang, *Adv. Mater.*, 2023, **35**, 2206272.
- 27 J. Kosco, S. Gonzalez-Carrero, C. T. Howells, T. Fei, Y. F. Dong, R. Sougrat, G. T. Harrison, Y. Firdaus, R. Sheelamantula, B. Purushothaman, F. Moruzzi, W. D. Xu, L. Y. Zhao, A. Basu, S. De Wolf, T. D. Anthopoulos, J. R. Durrant and I. McCulloch, *Nat. Energy*, 2022, **7**, 340–351.
- 28 A. Ciarrocchi, F. Tagarelli, A. Avsar and A. Kis, *Nat. Rev. Mater.*, 2022, **7**, 449–464.
- 29 S. Nishioka, K. Hojo, L. Xiao, T. Gao, Y. Maseki, S. Yasuda, T. Yokoi, K. Sayama, T. E. Mallouk and K. Maeda, *Sci. Adv.*, 2022, **8**, eadc9115.
- 30 J. V. Passarelli, C. M. Mauck, S. W. Winslow, C. F. Perkinson, J. C. Bard, H. Sai, K. W. Williams, A. Narayanan, D. J. Fairfield, M. P. Hendricks, W. A. Tisdale and S. I. Stupp, *Nat. Chem.*, 2020, **12**, 672–682.
- 31 F. Li, Z. H. Fang, Z. H. Xu and Q. J. Xiang, *Energy Environ. Sci.*, 2024, **17**, 497–509.
- 32 V. Gurylev, *Mater. Today Sustainability*, 2022, **18**, 100131.
- 33 P. Zhang, J. J. Zhang and J. L. Gong, *Chem. Soc. Rev.*, 2014, **43**, 4395–4422.



- 34 Z. Wang, Y. Inoue, T. Hisatomi, R. Ishikawa, Q. Wang, T. Takata, S. Chen, N. Shibata, Y. Ikuhara and K. Domen, *Nat. Catal.*, 2018, **1**, 756–763.
- 35 M. A. Lange, I. Khan, P. Opitz, J. Hartmann, M. Ashraf, A. Qurashi, L. Praedel, M. Panthofer, A. Cossmer, J. Pfeifer, F. Simon, M. von der Au, B. Meermann, M. Mondeshki, M. N. Tahir and W. Tremel, *Adv. Mater.*, 2021, **33**, 2007434.
- 36 Y. Yang and Y. Kawazoe, *Phys. Rev. Mater.*, 2018, **2**, 034602.
- 37 Y. Zhang, F. K. Wang, X. Feng, Z. Zhang, K. L. Liu, F. F. Xia, W. X. Liang, X. Z. Hu, Y. Ma, H. Q. Li, G. C. Xing and T. Y. Zhai, *Adv. Funct. Mater.*, 2022, **32**, 2205757.
- 38 L. K. Li, J. Kim, C. H. Jin, G. J. Ye, D. Y. Qiu, F. H. da Jornada, Z. W. Shi, L. Chen, Z. C. Zhang, F. Y. Yang, K. Watanabe, T. Taniguchi, W. Ren, S. G. Louie, X. H. Chen, Y. B. Zhang and F. Wang, *Nat. Nanotechnol.*, 2017, **12**, 21–25.
- 39 M. Baranowski and P. Plochocka, *Adv. Energy Mater.*, 2020, **10**, 1903659.
- 40 A. Miyata, A. Mitioglu, P. Plochocka, O. Portugall, J. T. W. Wang, S. D. Stranks, H. J. Snaith and R. J. Nicholas, *Nat. Phys.*, 2015, **11**, 582–587.
- 41 N. Hong Nhan, T. Reier, H.-S. Oh, M. Gliech, P. Paciok, V. Thu Ha Thi, D. Teschner, M. Heggen, V. Petkov, R. Schloegl, T. Jones and P. Strasser, *Nat. Catal.*, 2018, **1**, 841–851.
- 42 K. Jia, J. Ma, J. Wang, Z. Liang, G. Ji, Z. Piao, R. Gao, Y. Zhu, Z. Zhuang, G. Zhou and H.-M. Cheng, *Adv. Mater.*, 2023, **35**, 2208034.
- 43 H. Yang, C. L. Dong, H. M. Wang, R. J. Qi, L. Q. Gong, Y. R. Lu, C. H. He, S. H. Chen, B. You, H. F. Liu, J. L. Yao, X. L. Jiang, X. P. Guo and B. Y. Xia, *Proc. Natl. Acad. Sci. U. S. A.*, 2022, **119**, e2202812119.
- 44 A. Savin, R. Nesper, S. Wengert and T. F. Fassler, *Angew. Chem., Int. Ed.*, 1997, **36**, 1809–1832.
- 45 X. Feng, K. J. Bu, T. Liu, S. H. Guo, Z. D. Sun, T. H. Fu, Y. S. Xu, K. L. Liu, S. J. Yang, Y. H. Zhao, H. Q. Li, X. J. Lu and T. Y. Zhai, *Angew. Chem., Int. Ed.*, 2022, **62**, e202217238.
- 46 D. B. Chesnut and A. Savin, *J. Am. Chem. Soc.*, 1999, **121**, 2335–2336.
- 47 B. Silvi and A. Savin, *Nature*, 1994, **371**, 683–686.
- 48 J. Y. Duan, T. Y. Liu, Y. H. Zhao, R. O. Yang, Y. Zhao, W. B. Wang, Y. W. Liu, H. Q. Li, Y. F. Li and T. Y. Zhai, *Nat. Commun.*, 2022, **13**, 2039.

