

inorganic and organic chemistry faced limited development during that period, likely attributed to challenges in reproducing its synthesis on a large scale and the sensitivity of the anion to moisture and air. A noteworthy resurgence in the chemistry of the phosphoethynolate anion has occurred in the last decade, driven by seminal contributions of Grützmacher and Goicoechea groups.¹⁰ These reports independently presented efficient procedures for synthesizing stable phosphoethynolate anion, overcoming previous limitations. Consequently, $[\text{OCP}]^-$ has emerged as an appealing phosphorus transfer agent, demonstrating efficiency in the synthesis of numerous organophosphorus molecules and phosphorus-based transition metal complexes.¹¹

It is obvious that the rational design of new reactions involving $[\text{OCP}]^-$ as a nucleophile necessitates a comprehensive understanding of its ambident reactivity. In 2014, Grützmacher and Benkő elucidated this ambident reactivity by reacting $[\text{OCP}]^-$ with $\text{iPr}_3\text{Si}-\text{OTf}$, yielding the phosphoalkyne and phosphaketene, which were fully characterized (Scheme 1a).¹² It was concluded that while the oxygen attack is kinetically controlled, the phosphorus attack is thermodynamically controlled. Very recently, the Benkő's group employed Marcus theory to rationalize the ambident reactivity of $[\text{OCP}]^-$ towards carbon electrophiles. Their results indicated a lower intrinsic barrier for the oxygen attack, highlighting the kinetic preference of oxygen attacks.¹³

Building upon these insights and drawing inspiration from precedent contributions by Mayr and coworkers,¹⁴ which revealed that the ambident reactivity of various nucleophiles does not follow the well-established Hard and Soft Acids and Bases theory (HSAB),¹⁵ we embarked on investigating the reactivity of $[\text{OCP}]^-$ towards Mayr's reference electrophiles. The objective is to gain deeper insights into the factors controlling the reactivity of this anion, which is crucial for unlocking its synthetic potential (Table 1).

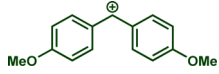

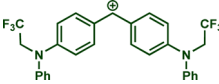
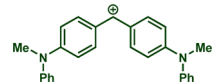
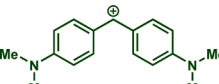
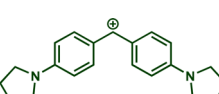
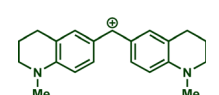
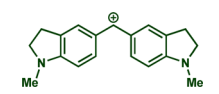
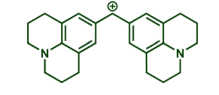
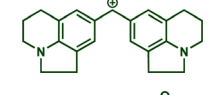
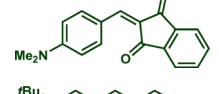
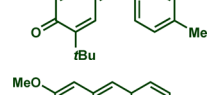
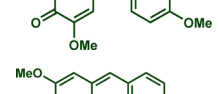
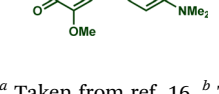
In previous investigations, Mayr *et al.* have shown that numerous nucleophile–electrophile combinations can simply be described by the three parameters eqn (1), where k_2 measures the second-order rate constant of the reaction of an electrophile with a nucleophile, N and s_N are nucleophilicity parameters, and E is the electrophilicity parameter.¹⁸

$$\log k_2 (20^\circ\text{C}) = s_N(E + N) \quad (1)$$

Our investigation commences with the synthesis of sodium phosphoethynolate using a protocol outlined in Scheme 2. This method, previously described by Grützmacher,¹² utilizes inexpensive precursors (sodium, red phosphorus, *t*BuOH, ethylene carbonate).¹² The target compound, $\text{Na}(\text{dioxane})_{2.5}(\text{OCP})$ was successfully obtained on a gram scale. As ion pairing should not be important under the low concentrations used in the kinetic studies, $[\text{Na}(\text{OCP})](\text{dioxane})_{2.5}$ will be referred to as $\text{Na}(\text{OCP})$ throughout this article.

The reaction of $\text{Na}(\text{OCP})$ with different reference electrophiles were carried out in acetonitrile at 20 °C, under pseudo-first-order conditions, by using at least 10 equivalents of the nucleophile with respect to the electrophile. Rates of those

Table 1 Reference electrophiles employed for the determination of the nucleophilicity parameters of $\text{Na}(\text{OCP})$, their electrophilicity

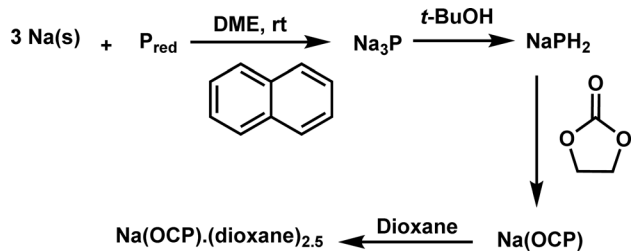
Electrophile		Electrophilicity, E
	1a	0 ^a
	1b	−3.72 ^b
	1c	−3.85 ^a
	1d	−5.89 ^a
	1e	−7.02 ^a
	1f	−7.69 ^a
	1g	−8.22 ^a
	1h	−8.76 ^a
	1i	−9.45 ^a
	1j	−10.04 ^a
	1k	−13.56 ^a
	1l	−15.83 ^a
	1m	−16.38 ^a
	1n	−17.18 ^a

^a Taken from ref. 16. ^b Taken from ref. 17.

reactions were determined either by studying the kinetics of laser-flash photolytically generated benzhydrylium ions or conventional UV-visible spectrophotometry using stable quinone methides or benzhydrylium ions.

Following previous investigations by Mayr *et al.*,¹⁹ the carbocation **1j** was generated photolytically upon irradiation of the





Scheme 2 Synthesis of sodium phosphoethynolate.

corresponding phosphonium salt with a 7 ns laser pulse at 266 nm (see ESI†). In the presence of a large excess of Na(OCP), one can follow the monoexponential decays of the absorbance of **1j** ($\lambda_{\text{max}} = 635$ nm), from which the rate constants k_{obs} (s^{-1}) are obtained (Fig. 1a). Interestingly, plots of k_{obs} versus the Na(OCP) concentrations gave linear correlations (Fig. 1b), and the resulting slopes yielded the second-order rate constants k_2 ($\text{L mol}^{-1} \text{s}^{-1}$) which are listed in Table 2. To investigate the role of the counterion on the reactivity of Na(OCP) with benzhydrylium ions, reactions of the former with the carbocation **1n** were studied in the presence of the crown ether 15-crown-5. However, under these conditions, only a very small change in the second-order rate constant was noticed, indicating that the

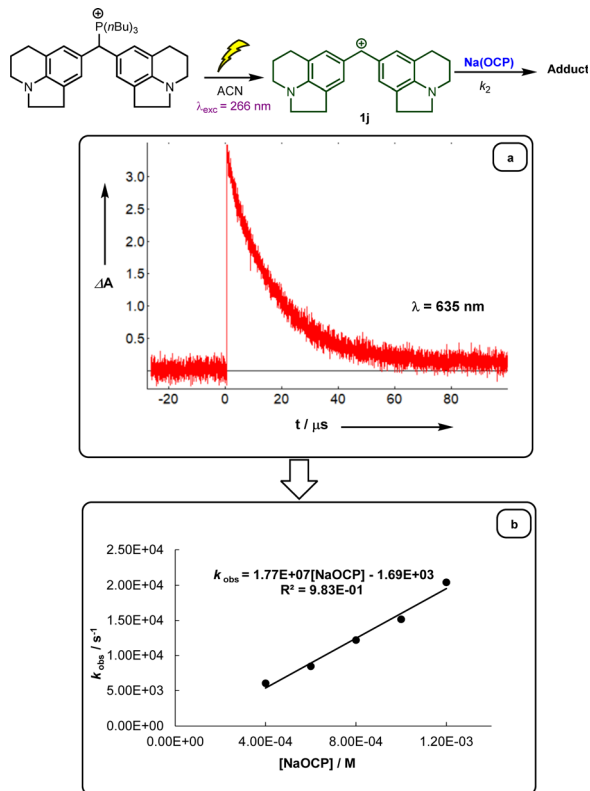


Fig. 1 (a) Decay of the absorbance of carbocation **1j** obtained after irradiation of a 3.07×10^{-5} mol L^{-1} solution of the phosphonium salt **1j**-P(*n*Bu)₃ in acetonitrile in the presence of Na(OCP). (b) Plot of the pseudo-first-order rate constants k_{obs} (s^{-1}) versus the concentration of Na(OCP).

Table 2 Second-order rate constants for the reaction of [OCP⁻] with reference electrophile in acetonitrile at 20 °C

Electrophile	k_2 ($\text{M}^{-1} \text{s}^{-1}$)
1a	— ^a
1b	— ^a
1c	2.52×10^9
1d	2.18×10^9
1e	—
1f	1.04×10^8
1g	9.67×10^7
1h	2.23×10^8
1i	1.08×10^8
1j	1.77×10^7
1k	1.74×10^4
1l	2.98×10^2
1m	2.17×10^2
1n	3.71×10^1

^a Very fast reaction.

counterion does not play a crucial role in the reactivity of [OCP⁻] towards carbocations (see ESI†). The effect of dioxane on the reactivity has also been addressed, and we found that the reaction of dioxane-free Na(OCP) with **1n** react similarly as [Na(OCP)](dioxane)_{2.5}.

In accordance with eqn (1), Fig. 2 shows that the second-order rate constants k_2 correlate linearly with the electrophilicity parameters E of the reference electrophiles **1**. The flattening of the curve at $k_2 \approx 2.5 \times 10^9$ L mol⁻¹ s⁻¹ is obviously due to diffusion control, which is in agreement with previous observations by Mayr *et al.* for the reactions of other benzhydrylium ions with other nucleophiles.^{19b} The nucleophilicity parameters ($N = 19.02$ and $s_N = 0.82$) of [OCP⁻] were derived from the linear part of the curve (*i.e.*, reactions with **1h–1n** in Fig. 2).

It should be noted that the measured second-order rate constants (Table 2) could not be attributed to a single electron transfer mechanism, as the oxidation of Na(OCP) is known to yield the heterobicyclic dianion, (P₄C₄O₄)²⁻. This intermediate was not observed by ³¹P NMR when we investigated the reaction of Na(OCP) with various carbocations **1**.

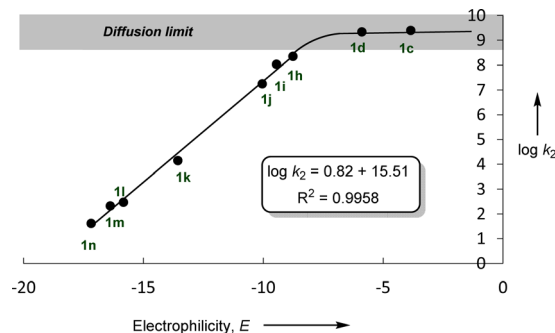
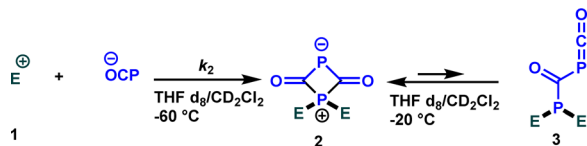


Fig. 2 Plot of $\log k_2$ for the reactions of sodium phosphoethynolate Na(OCP) with reference electrophiles **1**, in acetonitrile at 20 °C versus their electrophilicity parameters E .





Scheme 3 Reactions of Na(OCP) with carbocations **1b** and **1e**.

It is important to emphasize the excellent linearity observed in the correlation ($\log k_2$ vs. E), indicating that the rate-determining step does not change throughout this reaction series. This implies that the same nucleophile terminus center attacks all electrophiles.

The next critical step involves determining whether the identified nucleophilicity parameters align with the attack on oxygen or phosphorus. To elucidate this aspect, we investigated the reaction outcomes of Na(OCP) with both the highly reactive electrophile **1b** as well as the stabilized carbocation **1e** (Scheme 3).

When Na(OCP) (1.2 equivalents) reacted with one equivalent of carbocation **1e** in a dichloromethane/THF (1 : 1) mixture at room temperature, a complex ^{31}P NMR spectrum was obtained. However, when the same reaction was carried out at low temperature ($-60\text{ }^\circ\text{C}$), an immediate disappearance of the Na(OCP) ($\delta^{31}\text{P}\{^1\text{H}\} = -394.1$ ppm) was observed within five minutes and a new species bearing two phosphorus atoms appeared ($\delta^{31}\text{P}\{^1\text{H}\} = 341.8$ ppm (d, $^2J_{\text{PP}} = 36.7$ Hz) and 121.2 ppm (d, $^2J_{\text{PP}} = 36.7$ Hz)). ^1H , ^{13}C and 2D NMR experiments indicated the exclusive formation of the zwitterion **2**. The same intermediate was also detected for other electrophiles **1** (for more information, see ESI ‡). When the temperature was raised to $-20\text{ }^\circ\text{C}$, the zwitterion **2** coexists with the phosphaketene adduct **3** ($\delta^{31}\text{P}\{^1\text{H}\} = -276.2$ ppm (d, $^2J_{\text{PP}} = 167.7$ Hz) and 21.1 ppm (d, $^2J_{\text{PP}} = 167.8$ Hz)) in 5 to 1 ratio (2/3).

Grützmaier, Stephan, and co-authors reported the formation of structurally analogous complexes when they studied the reactions of the phosphoethynolate anion with a variety of boranes as Lewis acids. 20

To understand how these intermediates **2** and **3** were formed and to support the analysis of the kinetic data, we analyzed putative reaction mechanisms for the combination of free $[\text{OCP}]^-$ with the tropylium cation **1b** using density functional theory (RI-DSD-PBEP86-D3(BJ)/def2-QZVPP/SMD(THF)//M06-2X/6-31+G(d,p)/SMD(THF)). These results are summarized together with selected transition state structures in Scheme 4. Our calculations predict that $[\text{OCP}]^-$ and **1b** initially form a reactant complex $[\text{OCP}/\mathbf{1b}]$ that is more stable than the separated reactants. This complex is most likely held together *via* coulombic interactions. However, due to the large number of potential conformers of the reactant complex, the energetic value for this complex in Scheme 4 corresponds to the optimized structure at the end of the corresponding IRC calculation. Therefore, it is not unlikely that there are more stable bimolecular complexes. Within this complex, the OCP anion undergoes a rapid reaction with the tropylium ion (**TS1**, $\Delta G^\ddagger = 10\text{ kJ mol}^{-1}$) and forms the P-alkylated intermediate **4b**. This is

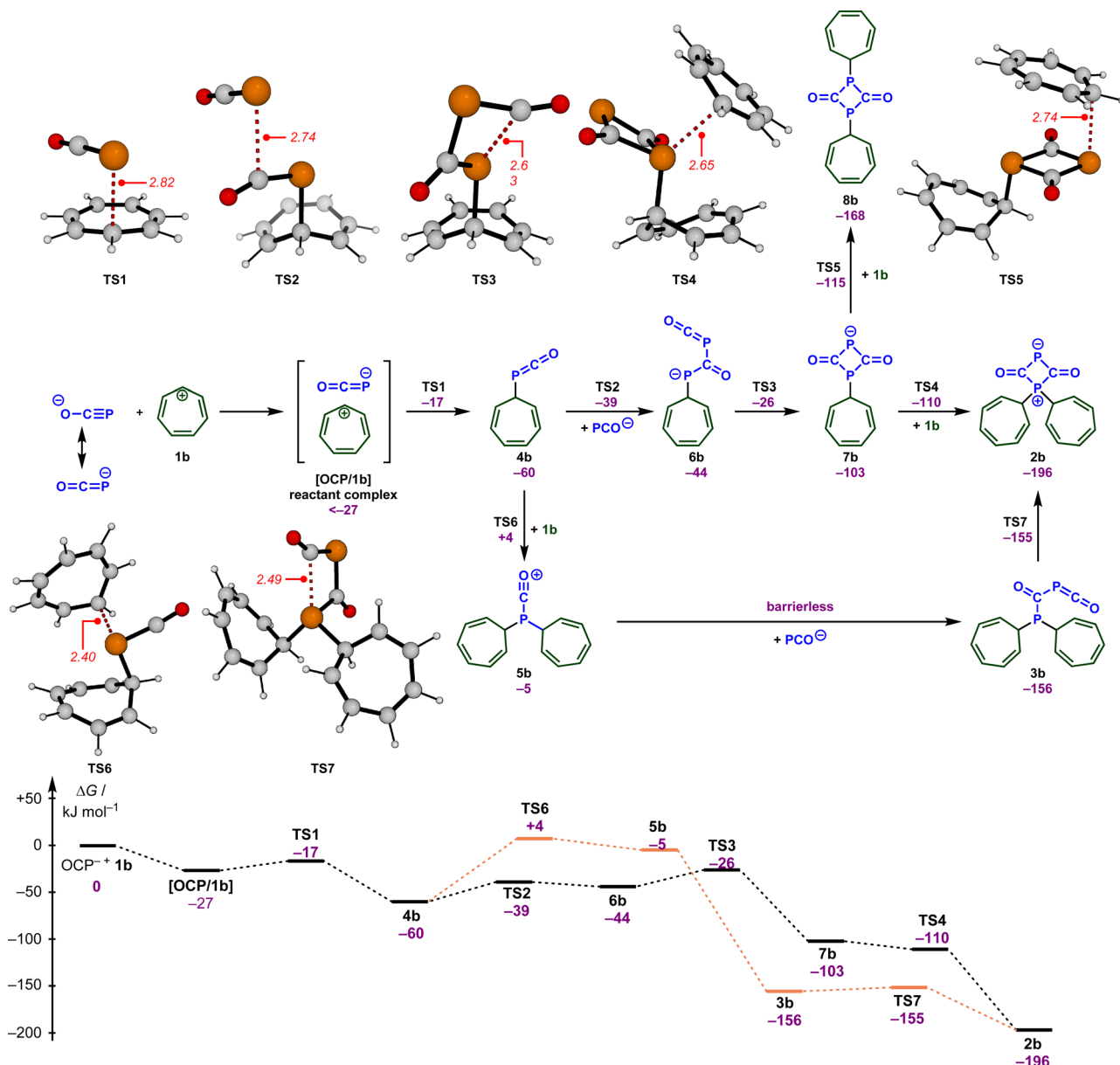
in perfect agreement with the high nucleophilicity (see above) and electrophilicity ($E = -3.72$) 17 of the tropylium ion. The alternate reactivity at the oxygen atom of $[\text{OCP}]^-$ (not shown in Scheme 4) leads to an O-alkylated compound that lies 101 kJ mol^{-1} above intermediate **5b**.

In the next steps, **4b** is transformed into the experimentally observed intermediate **2b**. This requires the reaction with another $[\text{OCP}]^-$ and a second tropylium cation. According to our DFT calculations, the formal $[2 + 2]$ cycloaddition between $[\text{OCP}]^-$ and **4b** proceeds in a stepwise fashion and occurs very quickly (**TS2**, **TS3**, $\Delta G^\ddagger = 21$ and 34 kJ mol^{-1}) and forms the diphosphatedione **6b**. In this sequence, another $[\text{OCP}]^-$ will initially attack the electrophilic carbon atom within phosphaketene **4b**. This is followed by a ring-closing reaction to yield the diphosphatedione anion **7b**. Alkylation of **7b** with a second equivalent of the tropylium cation again proceeds very rapidly and can occur at both phosphorus atoms through **TS4** or **TS5**. While negative barriers were calculated on the RI-DSD-PBEP86 potential energy surface in both cases, small barriers were determined on the M06-2X surface. Eventually, the zwitterion **2b** is formed with a high thermodynamic driving force ($\Delta G = -196\text{ kJ mol}^{-1}$) as the more stable species.

Alternatively, intermediate **4b** can first react with the tropylium ion through **TS6**, which requires a slightly larger barrier of 64 kJ mol^{-1} . No transition states could be located on the potential energy surface for the reaction of **5b** with the OCP anion and all potential energy surface scans resulted in a barrierless addition. This indicates that this step will proceed very rapidly under the reaction conditions. Cyclization of phosphaketene **3b** finally results in the zwitterion **2b** again without a significant barrier. In agreement with the experimental observations (Scheme 3), the computational investigations also indicate that zwitterion **2b** is more stable than the phosphaketene **3b**, however, the thermodynamic difference seems to be substantially overestimated in the calculations. To better understand this deviation, we first calculated the thermodynamic differences for other substituents on the phosphorus atom (see the ESI for details ‡). Regardless of the substituent, a comparable strong preference for the zwitterions was observed in all cases. Similarly, different computational methods (*e.g.*, DLPNO-CCSD(T), B2GP-PLYP, M06-2X, ω 97X-V) also resulted in almost identical energy differences in favor of the zwitterion. Finally, we realized that solvation seems to be an important aspect. In the gas phase, both structures **2b** and **3b** are almost isoenergetic, and with increasing polarity, the zwitterion **2b** substantially benefits from solvation, which then leads to an overestimation. Thus, the overestimation can be traced back to issues arising from the charge separation within the zwitterions.

These results seem to contradict earlier observations by Slootweg and coworkers, who reported a different product for the reaction of sodium phosphoethynolate with 1,2,3-tris-*tert*-butylcyclopropenium tetrafluoroborate **10** (Scheme 5). 14 Indeed, while the first step of the reaction leads to a phosphaketene, similar to the reaction with our reference electrophiles **1**, a fast $[2 + 2]$ cycloaddition yields bis(cyclo-propenyl) diphosphatedione **80**. The different reactivity of these





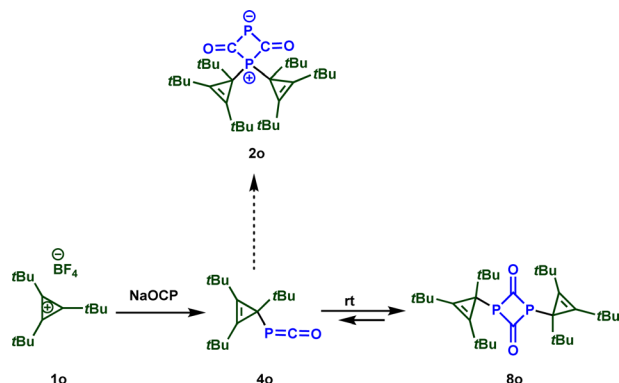
Scheme 4 Calculated reaction free energies (in kJ mol⁻¹) and selected transition states (bond lengths in Å) for the proposed reaction mechanism for the formation of the zwitterion **2b** and the phosphaketene **3b**.

electrophiles is probably caused by the steric demand within **1o**. While the 1,3-disubstituted diphosphetandione **8b** is thermodynamically less stable than the zwitterion **2b** (see Scheme 4), the situation reverts for the electrophile **1o**. According to our calculations, **8o** is substantially more stable than **2o** ($\Delta\Delta G = -48$ kJ mol⁻¹) (Scheme 5).

Based on the results highlighted above, one can assign the measured nucleophilicity ($N = 19.02$, $s_N = 0.82$) to the phosphorus site of the phosphoethynolate, which is five orders of magnitude more reactive than cyanate anion²¹ and ten times more reactive than the N-terminus of thioisocyanate.²² This high reactivity may explain the capability of this anion to react with a wide variety of electrophiles, including weak ones such as carbodiimide ($E \approx -20$) (Fig. 3).^{23,24}

After determining the phosphorus nucleophilicity of the OCP anion, we now wondered how this knowledge can be used in synthetically useful transformations. The investigation of the mechanism of the reaction of sodium phosphoethynolate with stabilized carbocations, for instance **1b**, revealed the formation of the zwitterion **2b** that was detected at -60 °C. To convert the latter into an isolable intermediate, we treated it with the NHC carbene **9**. Notably, this resulted in the formation of azolium phosphoethynolate **10b** in quantitative yield, which can be described as the NHC adduct of **4b**. The structure of **10b** was confirmed by a single crystal X-ray diffraction experiment (Scheme 6). In the solid state, the C–P and C–O bond lengths are 1.737 and 1.264 Å, respectively. These results are in good agreement with those reported by Stephan, Cummins *et al.*²⁵ for





Scheme 5 Reaction of sodium phosphoethynolate with 1,2,3-tris-tert-butylcyclopropenium **1o**.^{11j}

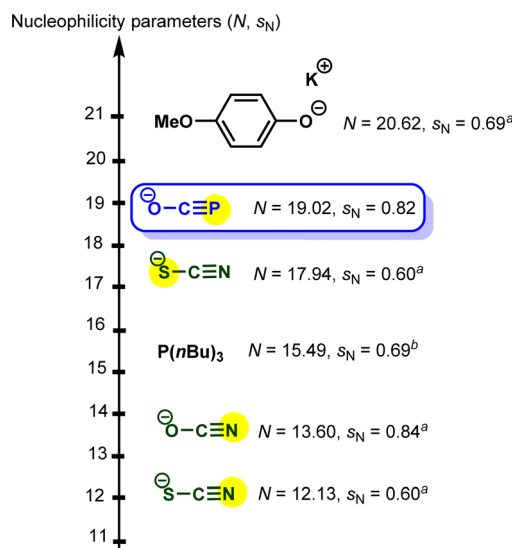
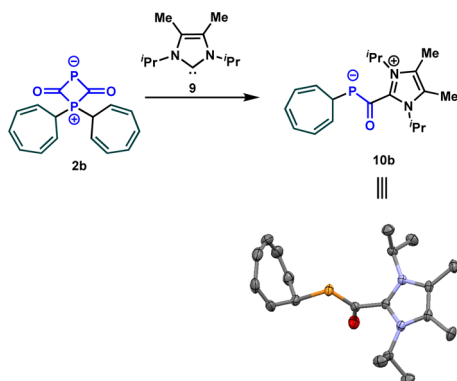


Fig. 3 Embedding sodium phosphoethynolate in the Mayr nucleophilicity scale. ^a Solvent: acetonitrile. ^b Solvent: dichloromethane.^{21–24}



Scheme 6 Reaction of the zwitterion **2b** with the carbene **9**.

acylphosphide anions, where experimental and computational studies revealed that the short C–P and long C–O bond are the consequence of the delocalization of electron density from the phosphide lone pair into $\pi^*(\text{C}=\text{O})$ orbital.

Importantly, when the zwitterion **2e** is dissolved in (dichloromethane/THF) in the presence of 0.5 equiv. of water, the secondary phosphine **11e** is formed, as observed by ³¹P NMR spectroscopy (see ESI[†]). Conducting the reaction in the presence of D₂O resulted in the formation of deuterated phosphine, providing confirmation that the proton originates from water (Scheme 7).

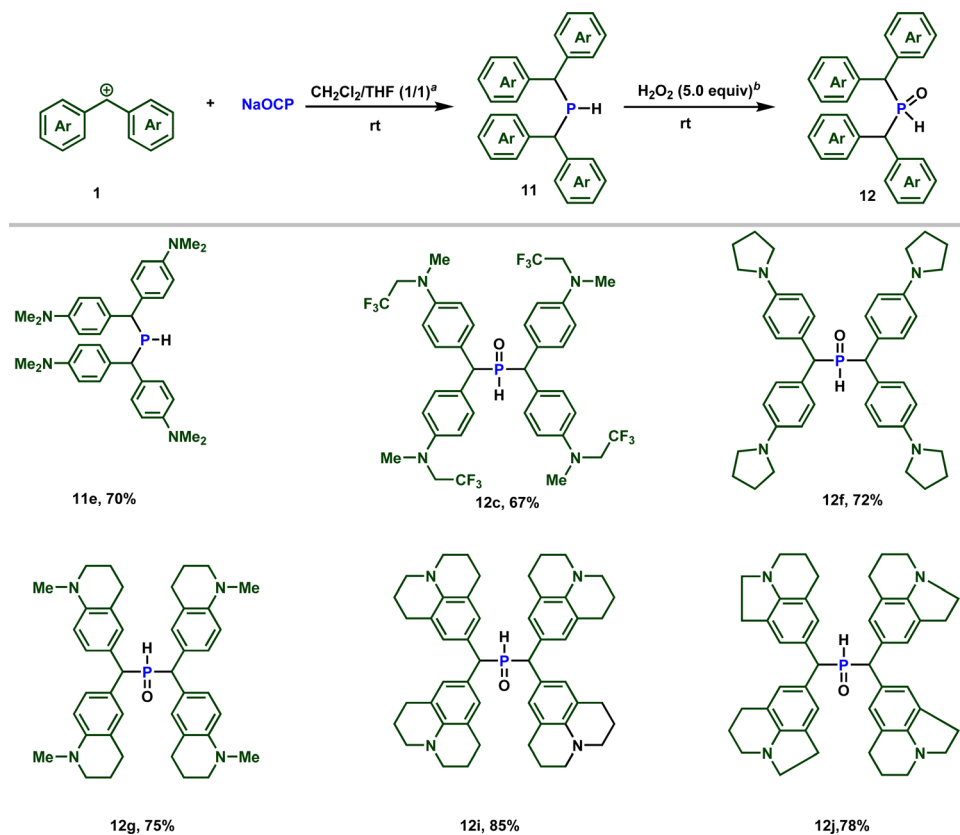
Building upon these observations, we combined different stabilized carbocations **1** with Na(OCP) in THF/dichloromethane (1:1) mixture, containing 0.5 equiv. of water. Notably, employing stabilized carbocations (**1c**, **f**, **g**, **i**, **j**) led to the formation of secondary phosphines **11** and secondary phosphine oxides **12** in good to excellent yields, as confirmed by ³¹P NMR. While the phosphine **11e** is stable to be isolated by column chromatography, the other phosphines oxidize rapidly when exposed to air. Consequently, these were treated with 5 equivalents of H₂O₂ and isolated as secondary phosphine oxides **12** in synthetically useful yields (Scheme 7).

Due to the high electrophilicity and Lewis acidity of **1a** and **1b**,^{19a} their reactions with the OCP anion produced a mixture of secondary and tertiary phosphines. Clearly, the formed secondary phosphines promptly react with **1a** and **1b** to give the tertiary phosphines. Both phosphines are oxidized to the phosphine oxides and isolated separately. The overall isolated yields are consistently good to very good (Scheme 8).

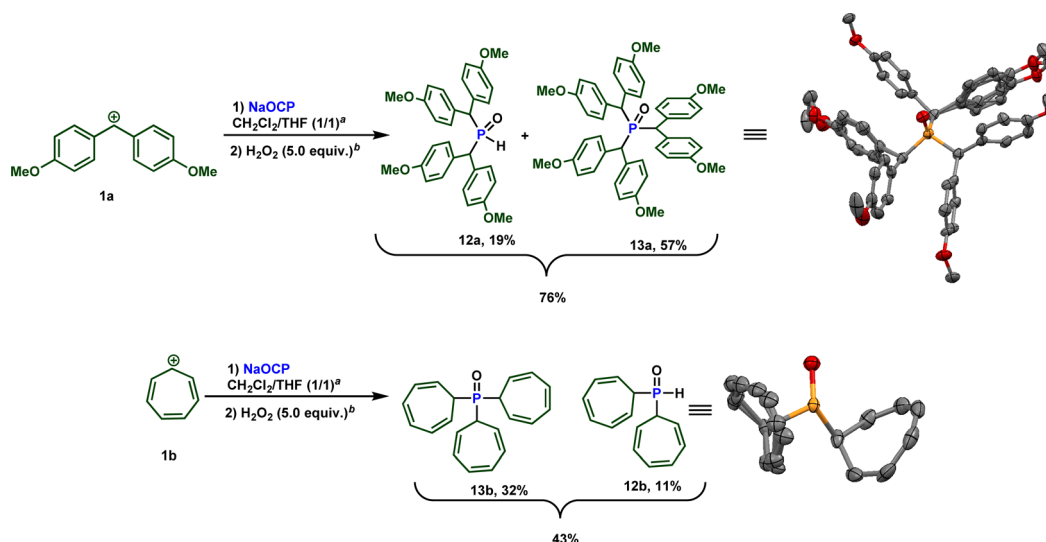
It is important to note that while an elegant synthesis of acyl phosphines and related molecules from the phosphoethynolate anion and related structures from [OCP][−] has been previously documented by Goicoechea,^{11b} Stephan,²⁶ and others, direct synthesis of the sterically hindered phosphines **12** at **13** has not been described to our knowledge. Furthermore, traditional methods for obtaining these types of molecules would necessitate the use of strong Brønsted bases. Therefore, this approach offers a more convenient route as it enables access to these compounds under mild and sustainable conditions.

We again relied on DFT calculations to gain further insights into the mechanism of this useful transformation (Scheme 9). To simplify the computational studies, we continue to use the tropylium cation **1b** as the initial electrophile. Based on our DFT calculations, the formation of the secondary phosphine **11b**, CO₂, and phosphaneylidene methane (OCPH) is thermodynamically feasible ($\Delta G = -40 \text{ kJ mol}^{-1}$). Mechanistically, we propose that the zwitterion **2b** initially reacts with water to yield intermediate **14**. No barrier could be detected for the attack of water and the computational studies indicate that the attack of water and the proton transfer from water to the phosphorus atom occurs in a concerted fashion. This intermediate quickly loses OCPH *via* **TS8** and affords the phosphanecarboxylic acid **15**. A concerted decarboxylation is generally feasible and affords the secondary phosphine **11b** in an overall favorable process, but the barrier for this step is very high (**TS9**, $\Delta G^\ddagger = 155 \text{ kJ mol}^{-1}$). Thus, a concerted process is very unlikely under the reaction conditions. Instead, the OCP anion could deprotonate the carboxylic acid and the decarboxylation takes place through an anionic transition state **TS10** ($\Delta G^\ddagger = 125 \text{ kJ mol}^{-1}$). This is still a substantial barrier, but might be slightly overestimated due to the involved proton transfer. A more likely





Scheme 7 Reactivity of sodium phosphoethynolate towards stabilized carbocations. ^a 0.5 equiv. of water. ^b H₂O₂ 30% (w/w) in water was used.

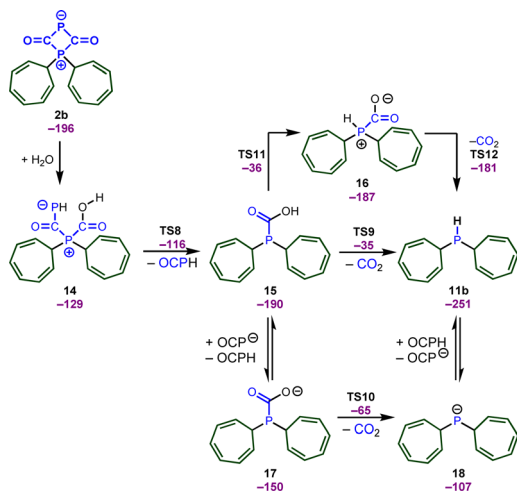


Scheme 8 Reactivity of sodium phosphoethynolate towards highly reactive carbocations **1a** and **1b**. ^a 0.5 equiv. of water. ^b H₂O₂ 30% (w/w) in water was used.

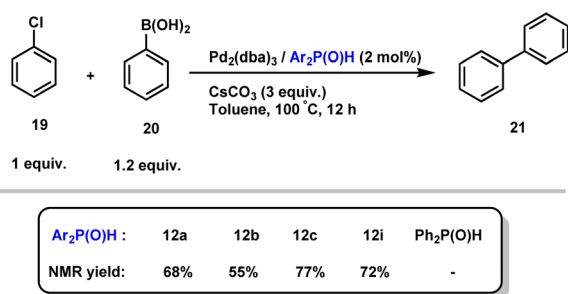
scenario is the formation of a zwitterionic phosphanecarboxylic acid **16**. Although a formal 1,3 proton shift through **TS11** is very unfavorable, another molecule **15** or the OCP anion will facilitate the proton transfer. The zwitterion **16** readily undergoes the decarboxylation (**TS12**) and affords the secondary phosphane in a very rapid reaction.

We should emphasize that consistent with the computational studies, the formation of OCPH was experimentally confirmed, as its dimer was detected when **2b** was reacted with water (0.5 equiv.) at $-60\text{ }^{\circ}\text{C}$ (see ESI,[†] pages S93 and S94) (³¹P {¹H} NMR spectrum at 305.8 and 75.9 ppm with a ²J_{P-P} coupling constant of 16.0 Hz, as well as a doublet of doublets at 7.32 ppm





Scheme 9 Potential pathways for the formation of secondary phosphines.



Scheme 10 Secondary phosphine oxide/ $\text{Pd}_2(\text{dba})_3$ catalysed the Suzuki cross-coupling reaction of **19** with **20**.

($^1J_{\text{P-H}} = 172$, $^3J_{\text{P-H}} = 16.0$ Hz)). This finding aligns well with Goicoechea's observation that OCPH undergoes dimerization.²⁷

Finally, given that hindered secondary phosphine oxides have proven to be effective ligands for transition metal-catalyzed reactions, as initially demonstrated by Li,²⁸ we evaluated their combination with $\text{Pd}_2(\text{dba})_3$ in the Suzuki cross-coupling reaction of aryl chloride **19** with phenylboronic acid **20**.²⁹ Importantly, the desired biphenyl product **21** in good yields, showing the efficiency of our secondary phosphine oxides (Scheme 10). Interestingly, no reaction occurred when diphenylphosphine oxide was used as a ligand, highlighting the crucial role of sterically hindered phosphines.

Conclusions

In conclusion, we herein provide an experimental quantification of the phosphorus nucleophilicity of the sodium phosphoethynolate. This reactivity is both kinetically and thermodynamically favourable. Mechanistic investigations allowed the identification and full characterization of key intermediates such as the zwitterion **2**, and also the optimization of experimental conditions enabling the synthesis of synthetically useful organophosphorus molecules such as

sterically hindered secondary phosphines oxides. The latter have shown to be effective ligands in Suzuki cross coupling reaction.

Data availability

All experimental procedures, details of the calculations, and additional data can be found in the ESI.†

Author contributions

T. H. V. N. conducted synthetic and kinetic experiments, while S. C. analysed the kinetic data and carried out experiments using laser-flash photolysis equipment. S. M.-L. conducted the X-ray experiments. M. B. performed all calculations. S. L. conceived and supervised the project. All authors participated in discussing the results, providing comments, and proof-reading the manuscript.

Conflicts of interest

The authors declare no competing interests.

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Notes and references

- D. E. C. Corbridge, *Phosphorus 2000. Chemistry, Biochemistry & Technology*, Elsevier, Amsterdam, 2002.
- (a) W. Schipper, *Eur. J. Inorg. Chem.*, 2014, 1567–1571; (b) F. H. Westheimer, *Science*, 1987, **235**, 1173–1178; (c) D. M. Karl, *Nature*, 2000, **406**, 31–33; (d) L. V. Kochian, *Nature*, 2012, **488**, 466–467.
- K. H. Büchel, H. H. Moretto and P. Woditsch, *Industrial Inorganic Chemistry*, Wiley VCH, New York, 2nd edn, 2000, pp. 65–101.
- For a comprehensive review on phosphinate chemistry, see: J. L. Montchamp, *Acc. Chem. Res.*, 2014, **47**, 77–87.
- (a) D. J. Scott, *Angew. Chem., Int. Ed.*, 2022, **61**, e202205019; (b) D. J. Scott, J. Cammarata, M. Schimpf and R. Wolf, *Nat. Chem.*, 2021, **13**, 458–464; (c) M. Till, J. Cammarata, R. Wolf and D. J. Scott, *Chem. Commun.*, 2022, **58**, 8986–8989; (d) J. E. Borger, A. W. Ehlers, J. C. Slootweg and K. Lammertsma, *Chem.–Eur. J.*, 2017, **23**, 11738–11746; (e) M. Scheer, G. Balázs and A. Seitz, *Chem. Rev.*, 2010, **110**, 4236–4256; (f) B. M. Cossairt, N. A. Piro and C. C. Cummins, *Chem. Rev.*, 2010, **110**, 4164–4177.



- 6 (a) J. M. Goicoechea and H. Grützmacher, *Angew. Chem., Int. Ed.*, 2018, **57**, 16968–16994; (b) Z. J. Quan and X. C. Wang, *Org. Chem. Front.*, 2014, **1**, 1128–1131; (c) Z. Quan, W. Wang and X. Wang, *Chin. J. Org. Chem.*, 2015, **35**, 2301–2312.
- 7 (a) M. Regitz, *Chem. Rev.*, 1990, **90**, 191–213; (b) K. B. Dillon, F. Mathey and J. F. Nixon, *Phosphorus: The Carbon Copy*, Wiley, Chichester, 1998.
- 8 For phosphaketenes obtained from OCP anion, see: (a) S. Alidori, D. Heift, G. Santiso-Quinones, Z. Benkő, H. Grützmacher, M. Caporali, L. Gonsalvi, A. Rossin and M. Peruzzini, *Chem.–Eur. J.*, 2012, **18**, 14805–14811; (b) L. Liu, D. A. Ruiz, F. Dahcheh, G. Bertrand, R. Suter, A. M. Tondreau and H. Grützmacher, *Chem. Sci.*, 2016, **7**, 2335–2341; (c) N. Del Rio, A. Baceiredo, N. Saffon-Merceron, D. Hashizume, D. Lutters, T. Müller and T. Kato, *Angew. Chem., Int. Ed.*, 2016, **55**, 4753–4758; (d) Y. Xiong, S. Yao, T. Szilvási, E. Ballester-Martínez, H. Grützmacher and M. Driess, *Angew. Chem., Int. Ed.*, 2017, **56**, 4333–4336.
- 9 (a) G. Becker, W. Schwarz, N. Seidler and M. Westerhausen, *Z. Anorg. Allg. Chem.*, 1992, **612**, 72–82; (b) G. Becker, G. Heckmann, K. Hübler and W. Schwarz, *Z. Anorg. Allg. Chem.*, 1995, **621**, 34–46.
- 10 (a) F. F. Puschmann, D. Stein, D. Heift, C. Hendriksen, Z. A. Gal, H. F. Grützmacher and H. Grützmacher, *Angew. Chem., Int. Ed.*, 2011, **50**, 8420–8423; (b) D. Heift, Z. Benkő and H. Grützmacher, *Dalton Trans.*, 2014, **43**, 831–840; (c) I. Krummenacher and C. C. Cummins, *Polyhedron*, 2012, **32**, 10–13; (d) A. R. Jupp and J. M. Goicoechea, *Angew. Chem., Int. Ed.*, 2013, **52**, 10064–10067.
- 11 For selected examples on the use of OCP anion in organic and organometallic chemistry, see: (a) A. M. Tondreau, Z. Benkő, J. R. Harmer and H. Grützmacher, *Chem. Sci.*, 2014, **5**, 1545–1554; (b) A. R. Jupp and J. M. Goicoechea, *J. Am. Chem. Soc.*, 2013, **135**, 19131–19134; (c) Y. H. Wu, Z. F. Li, W. P. Wang, X. C. Wang and Z. J. Quan, *Eur. J. Org. Chem.*, 2017, 5546–5553; (d) R. Suter, Y. Mei, M. Baker, Z. Benkő, Z. Li and H. Grützmacher, *Angew. Chem., Int. Ed.*, 2017, **56**, 1356–1360; (e) C. Camp, N. Settineri, J. Lefèvre, A. R. Jupp, J. M. Goicoechea, L. Maron and J. Arnold, *Chem. Sci.*, 2015, **6**, 6379–6384; (f) C. J. Hoerger, F. W. Heinemann, E. Louyriac, L. Maron, H. Grützmacher and K. Meyer, *Organometallics*, 2017, **36**, 4351–4354; (g) L. Liu, D. A. Ruiz, D. Munz and G. Bertrand, *Chem*, 2016, **1**, 147–153; (h) D. W. N. Wilson, A. Hinz and J. M. Goicoechea, *Angew. Chem., Int. Ed.*, 2018, **57**, 2188–2193; (i) L. Liu, D. A. Ruiz, F. Dahcheh, G. Bertrand, R. Suter, A. M. Tondreau and H. Grützmacher, *Chem. Sci.*, 2016, **7**, 2335–2341; (j) T. Krachko, A. W. Ehlers, M. Nieger, M. Lutz and J. C. Sloatweg, *Angew. Chem., Int. Ed.*, 2018, **57**, 1683–1687; (k) M. G. Jafari, Y. Park, B. Pudasaini, T. Kurogi, P. J. Carroll, D. M. Kaphan, J. Kropf, M. Delferro, M. Baik and D. J. Mindiola, *Angew. Chem., Int. Ed.*, 2021, **60**, 24411–24417; (l) M. M. Hansmann, *Chem.–Eur. J.*, 2018, **24**, 11573–11577.
- 12 D. Heift, Z. Benkő and H. Grützmacher, *Dalton Trans.*, 2014, **43**, 5920–5928.
- 13 Á. Horváth, B. D. Lőrincz and Z. Benkő, *Chem. Eur. J.*, 2023, **29**, e2023006.
- 14 H. Mayr, M. Breugst and A. R. Ofial, *Angew. Chem., Int. Ed.*, 2011, **50**, 6470–6505.
- 15 (a) R. G. Pearson, *Science*, 1966, **151**, 172–177; (b) R. G. Pearson, *J. Am. Chem. Soc.*, 1963, **85**, 3533–3539; (c) R. G. Pearson and J. Songstad, *J. Am. Chem. Soc.*, 1967, **89**, 1827–1936.
- 16 H. Mayr and A. R. Ofial, *SAR QSAR Environ. Res.*, 2015, **26**, 619–646.
- 17 H. Mayr, K.-H. Müller, A. R. Ofial and M. Bühl, *J. Am. Chem. Soc.*, 1999, **121**, 2418–2424.
- 18 (a) H. Mayr and M. Patz, *Angew. Chem., Int. Ed.*, 1994, **33**, 938–957; (b) H. Mayr, T. Bug, M. F. Gotta, N. Hering, B. Irrgang, B. Janker, B. Kempf, R. Loos, A. R. Ofial, G. Remennikov and H. Schimmel, *J. Am. Chem. Soc.*, 2001, **123**, 9500–9512; (c) H. Mayr, B. Kempf and A. R. Ofial, *Acc. Chem. Res.*, 2003, **36**, 66–77; (d) *The database of Mayr reactivity parameters (N , s_N , and E)*, <https://www.cup.lmu.de/oc/mayr/reaktionsdatenbank2/>, accessed 27/05/2024.
- 19 (a) H. Mayr, J. Ammer, M. Baidya, B. Maji, T. A. Nigst, A. R. Ofial and T. Singer, *J. Am. Chem. Soc.*, 2015, **137**, 2580–2599; (b) J. Ammer, C. Nolte and H. Mayr, *J. Am. Chem. Soc.*, 2012, **134**, 13902–13911; (c) J. Ammer, C. F. Sailer, E. Riedle and H. Mayr, *J. Am. Chem. Soc.*, 2012, **134**, 11481–11494; (d) R. Loos, S. Kobayashi and H. Mayr, *J. Am. Chem. Soc.*, 2003, **125**, 14126–14132; (e) J. Ammer and H. Mayr, *J. Phys. Org. Chem.*, 2013, **26**, 956–969.
- 20 K. M. Szkop, A. R. Jupp, R. Suter, H. Grützmacher and D. W. Stephan, *Angew. Chem., Int. Ed.*, 2017, **56**, 14174–14177.
- 21 A. A. Tishkov and H. Mayr, *Angew. Chem., Int. Ed.*, 2005, **44**, 142–145.
- 22 R. Loos, S. Kobayashi and H. Mayr, *J. Am. Chem. Soc.*, 2003, **125**, 14126–14132.
- 23 L. Liu, J. Zhu and Y. Zhao, *Chem. Commun.*, 2014, **50**, 11347–11349.
- 24 For the electrophilicity of carbodiimides, see: Z. Li, R. J. Mayer, A. R. Ofial and H. Mayr, *J. Am. Chem. Soc.*, 2020, **142**, 8383–8402.
- 25 K. M. Szkop, M. B. Geeson, D. W. Stephan and C. C. Cummins, *Chem. Sci.*, 2019, **10**, 3627–3631.
- 26 (a) K. M. Szkop, A. R. Jupp, H. Razumkov and D. W. Stephan, *Dalton Trans.*, 2020, **49**, 885–890; (b) D. A. Petrone, K. M. Szkop, L. Miao, P. S. Onge, Z. W. Qu, S. Grimme and D. W. Stephan, *Angew. Chem., Int. Ed.*, 2021, **60**, 18547–18551.
- 27 A. Hinz, R. Labbow, C. Rennick, A. Schulz and J. M. Goicoechea, *Angew. Chem., Int. Ed.*, 2017, **56**, 3911–3915.
- 28 G. Y. Li, *J. Org. Chem.*, 2002, **67**, 3643–3650.
- 29 (a) T. M. Shaikh, Ch-M. Weng and F.-E. Hong, *Coord. Chem. Rev.*, 2012, **256**, 771–803; (b) A. Gallen, A. Riera, X. Verdaguier and A. Grabulosa, *Catal. Sci. Technol.*, 2019, **9**, 5504–5561.

