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Significant hydrogen generation *via* photo-mechanical coupling in flexible methylammonium lead iodide nanowires†

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The flexoelectric effect, which refers to the mechanical-electric coupling between strain gradient and charge polarization, should be considered for use in charge production for catalytically driving chemical reactions. We have previously revealed that halide perovskites can generate orders of higher magnitude flexoelectricity under the illumination of light than in the dark. In this study, we report the catalytic hydrogen production by photo-mechanical coupling involving the photoflexoelectric effect of flexible methylammonium lead iodide (MAPbI₃) nanowires (NWs) in hydrogen iodide solution. Upon concurrent light illumination and mechanical vibration, large strain gradients were introduced in flexible MAPbI₃ NWs, which subsequently induced significant hydrogen generation (at a rate of 756.5 μmol g⁻¹ h⁻¹, surpassing those values from either photo- or piezocatalysis of MAPbI₃ nanoparticles). This photo-mechanical coupling strategy of mechanocatalysis, which enables the simultaneous utilization of multiple energy sources, provides a potentially new mechanism in mechanochemistry for highly efficient hydrogen production.

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Introduction

Mechanocatalysis, which triggers chemical reactions by mechanical deformation-induced charge energetics and separation, is an emerging field of advanced catalysis due to its unique potential to achieve the direct conversion of mechanical energy into chemical energy.¹ For instance, piezocatalysis, a typical mechanocatalysis, promotes redox reactions directly by mechanical stimuli-induced polarization charges (piezoelectric effect, *i.e.*, the coupling between strain and polarization charges) and the subsequent physical or chemical electron transfer.² Principally, piezocatalysts with high piezoelectric coefficients and substantial stresses are advantageous for achieving effective piezoelectric charge polarization. Therefore, the low-dimensional piezocatalysts^{3–5} integrated with the

piezoelectric effect and semiconductor properties (where the small size or dimensionality-reduced shape of nanomaterials allows large lattice strain and hence enhanced piezoelectricity)^{6–8} are highly appealing. These materials are not only attractive for exploring new physical phenomena at the nanoscale, but also potentially capable of driving redox reactions with the assistance of environmental vibrations (waves, wind, *etc.*).^{9,10} In this context, there has been a remarkable increase in the studies of piezocatalysis for nanoscale semiconductors in a range of applications, such as dye degradation,¹¹ hydrogen production, CO₂ reduction,¹² tooth whitening,¹³ tumor therapy,¹⁴ and organic synthesis.¹⁵

Piezoelectric-induced charge separation occurs when non-centrosymmetric materials are deformed by external stress.^{16,17} However, upon exposure to random external stresses, internal strain gradients are inevitably introduced in materials, besides the existence of strain.^{18,19} This has drawn attention to another ubiquitous electromechanical coupling effect, which is the flexoelectric effect.²⁰ Flexoelectricity refers to the electromechanical coupling between strain gradient and polarization,²¹ which could induce piezoelectric-like polarization charges; this phenomenon widely exists in insulators, liquid crystals, biological samples, and semiconductors²² (whereas piezoelectric materials must be in non-centrosymmetric crystals). In bulk materials, the flexoelectric effect is sometimes negligible due to the small strain gradient within the fracture limitation, but it becomes more significant when the crystal size is shrunk down

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to the nanoscale.^{23–25} It is noted that a few studies on flexocatalysis of nanomaterials have been reported, mainly focusing on the applications of dye degradation. Zhonglin Wang *et al.* suggested that under ultrasonic excitation, reactive species were created due to strain-gradient-induced flexoelectric polarization in MnO₂ nanosheets composed of nanoflowers, which could be used to degrade organic pollutants, such as methylene blue.²⁶ Yaojin Wang *et al.* also demonstrated the flexocatalysis effect in centrosymmetric SrTiO₃ nanoparticles for degrading organic dyes (rhodamine B).²⁷

Inspiringly, in our recent work, we have demonstrated that the flexoelectricity in halide perovskites can be significantly enhanced upon the illumination of light, where the materials deliver orders of higher magnitude flexoelectric responses than normal bending-induced flexoelectricity.²⁸ This giant photo-flexoelectric effect provides a whole new paradigm for harvesting multiple forms of energy at the same time. Although these halide perovskite materials are well known to suffer severe decomposition in aqueous solutions,^{29,30} their dynamic equilibrium of dissolution and precipitation in solution can still enable photocatalytic hydrogen iodide (HI) decomposition together with hydrogen (H₂) generation.³¹ For example, Yang Chai *et al.* have recently evidenced that methylammonium lead iodide (MAPbI₃) powder exhibited excellent piezoelectric photocatalytic hydrogen production efficiency in HI aqueous solution.³²

Herein, we report the first case of catalytic hydrogen production by photo-mechanical coupling involving the photoflexoelectric effect from flexible MAPbI₃ nanowires (NWs) in HI solution. It is worth noting that NWs with a large aspect ratio were used as catalysts in this study because they can generate significant flexoelectricity by experiencing a large strain gradient (see discussions below). The photo-enhanced electro-mechanical coupling in MAPbI₃ NWs achieved a significant hydrogen production rate of 756.5 μmol g⁻¹ h⁻¹ by catalytic decomposition of HI, which is much higher than the reported values from similar systems *via* photo- or photopiezocatalysis. The development of this photo-mechanical coupling strategy in mechanocatalysis provides a potential “shortcut” in simultaneous utilization of multiple energy sources, which may shed light on new mechanisms for efficient mechanocatalysis.

Results and discussion

We synthesized MAPbI₃ NWs according to the widely reported wet-chemical method.³³ MAPbI₃ nanoparticles (NPs) and single crystals (SCs) were also synthesized for comparison purposes (see Methods for details). Systematical characterizations were first performed to check the structural information of the as-synthesized samples. Fig. 1a shows the scanning electron microscopy (SEM) image of a crowd of NWs, revealing their typical one-dimensional structures as well as the smooth surfaces (with average diameters of 360 nm and lengths up to 37 μm; see statistical results of their diameters and lengths in Fig. S1†). The transmission electron microscopy (TEM) image in Fig. 1b further demonstrates the smooth surface of a NW. Fig. 1c–e present the energy dispersive spectroscopy (EDS)

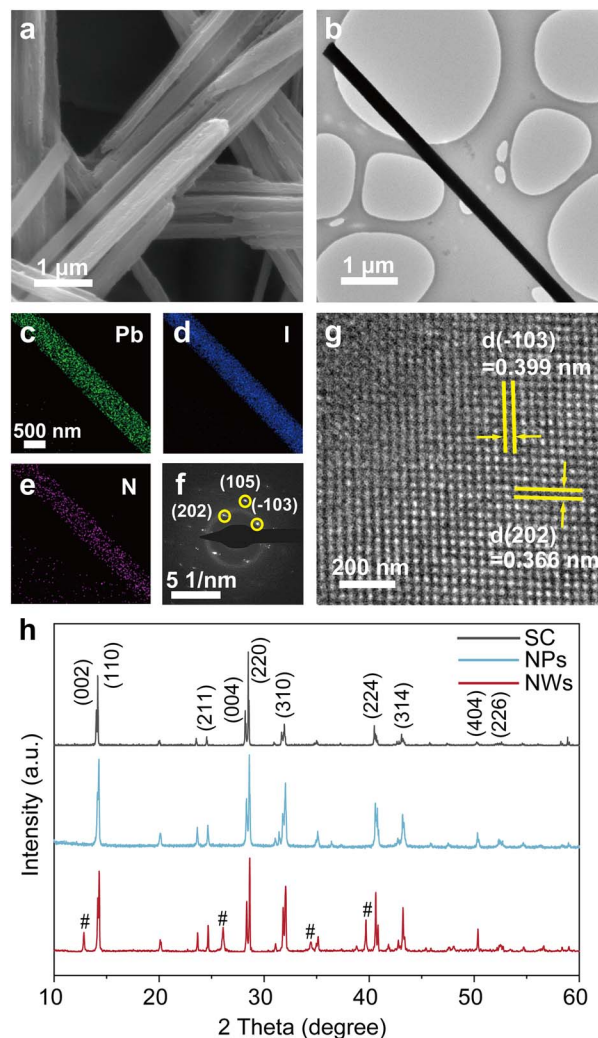


Fig. 1 Structural characterizations for MAPbI₃ NWs. (a) SEM image. (b) Bright-field TEM image. (c–e) EDS elemental mappings for Pb, I, and N on a NW. (f) SAED pattern. (g) Lattice-resolved HRTEM image. (h) XRD patterns for three MAPbI₃ samples.

elemental mappings for Pb, I, and N within a selected NW; the synchronous distributions of these constituent elements imply its uniform composition (one more set of data can be found in Fig. S2†). The selected area electron diffraction (SAED) pattern from an isolated NW in Fig. 1f confirms its single crystalline nature, and the pattern was assigned to a tetragonal structure as reported in previous syntheses.^{34,35} Furthermore, the high-resolution TEM (HRTEM) image (Fig. 1g) exhibits distinct lattice fringes with a pair of interplanar spacings of 0.399 and 0.366 nm, corresponding to the (−103) and (202) planes of tetragonal MAPbI₃, respectively. In addition, the X-ray diffraction (XRD) patterns of MAPbI₃ SC, NPs, and NWs (Fig. 1h) show that the diffraction peaks at 14.29° and 28.62° correspond well to the (110) and (220) planes of tetragonal MAPbI₃, indicating their preferential orientation along [110] directions.³⁶ The minor existence of PbI₂ (refer to the low-intensity peaks at 2θ = 12.84°, 26.09°, 34.46°, and 39.69°) is common in these nanomaterials.³⁷



The generation and separation of charge carriers, as well as their subsequent migration into the reaction systems, play a crucial role in the catalytic performance of catalysts. Ultraviolet photoelectron spectroscopy (UPS) and ultraviolet-visible (UV-Vis) absorption spectroscopy were undertaken to construct the band diagram of MAPbI₃ NWs. As shown in Fig. 2a, the as-synthesized MAPbI₃ NWs exhibit a maximum binding energy of 1.21 eV. The respective valence band positions of the NPs and SC samples were also assessed (Fig. S3†) and are consistent with the recent reports.^{38,39} The UV-Vis spectra (Fig. 2b) reveal pronounced light absorption within the three samples. The samples present their sharp absorption edges in the infrared region, which signifies that electrons within the samples undergo efficient transitions from the valence band to the conduction band under the irradiation of visible light. The superior light absorption of NWs as compared to NPs and SC, which was possibly caused by the large surface area and the low stacking density of NWs,^{40,41} implies the higher efficiency of the generation of the energetic electron-hole pairs. Fig. 2c presents the Tauc plot for the optical bandgap calculation of MAPbI₃ NWs, which is 1.54 eV (the corresponding values for the SC and NP samples are 1.51 and 1.54 eV, respectively; see

Fig. S4,† as well as the comparison of the structures for optical bandgaps of MAPbI₃ SC, NWs, and NPs in Fig. S5†). The band diagram of the MAPbI₃ NWs revealed that its conduction band minimum (CBM) and valence band maximum (VBM) corresponded to -0.63 and 0.91 V *versus* the normal hydrogen electrode (NHE), respectively. These energy positions are well-suited for hydrogen reduction (0.046 V *versus* NHE) and iodine oxidation (0.376 V *versus* NHE) within our saturated HI aqueous system, as depicted in the inset of Fig. 2c. Therefore, during the catalytic process, the holes from the VBM exhibit stronger oxidation capability compared to I⁻ ions, leading to the oxidation of I⁻ ions to I₃⁻ ions, while the electrons from the CBM are more reducible than H⁺, resulting in the reduction of H⁺ to form H₂ (*i.e.*, 3I⁻ + 2h⁺ = I₃⁻, 2H⁺ + 2e⁻ = H₂).

We then utilized the piezoelectric response force microscopy (PFM) mode of a scanning probe microscope (SPM) to assess the electromechanical coupling behavior of a single MAPbI₃ NW.⁴² As shown in Fig. 2d, the applied voltage through the PFM tip indeed induced the mechanical response of the underlying NW, which confirms the electromechanical coupling in the MAPbI₃ NW. Furthermore, as the Curie temperature (*i.e.*, the piezo-response should become zero above this temperature) of MAPbI₃ material is around 60 °C,⁴³ we also tested *in situ* the PFM response of NWs at a temperature range across its Curie temperature to verify the origin of the electromechanical coupling (Fig. S6 and S7†). The disappearance of hysteresis in the phase-voltage curves above the Curie temperature confirms the ferroelectric to non-ferroelectric transition for the NW (Fig. S7†). However, as it can be seen in Fig. 2e, the electromechanical coupling coefficient was maintained within 76 to 54 pm V⁻¹ (the system has been calibrated by a standard LiNbO₃ sample with its d₃₃ of 17.3 pm V⁻¹; see Fig. S8†), which indicated that the NW retained more than 70% of the electromechanical coupling response despite the absence of piezoelectric effect at the temperature beyond Curie point; in brief, it is confirmed that the as-measured electromechanical coupling of MAPbI₃ NWs was mainly attributed to the flexoelectric effect.

Furthermore, perovskite bulk materials are well known to be extremely fragile, which would be a significant issue for their applications in mechanocatalysis (in which large deformations of materials are essential for charge generation). However, when the dimensions of a sample are reduced toward the nanoscale, they may manifest exceptional flexibility, enabling them to withstand intense mechanical forces such as ultrasonication and stirring without being damaged. In the SEM/TEM observations, we observed a few bending NWs (see Fig. S9† for a representative example), which inspired the high flexibility of these MAPbI₃ NWs. Consequently, *in situ* experiments were conducted to observe the bending behaviors of MAPbI₃ NWs (the setup is schematically shown in Fig. 3a; see Methods for details). Upon forced bending by a rigid needle at one end, a single NW endured a large curvature (defined as how much the curve deviates from a straight line) before it was broken; yet under controlled deflection, the NW can be cycled between bending and straight states (as shown in Fig. 3b and c; the K values denoted on panels show the corresponding curvatures). During multiple bending processes, it was observed that the

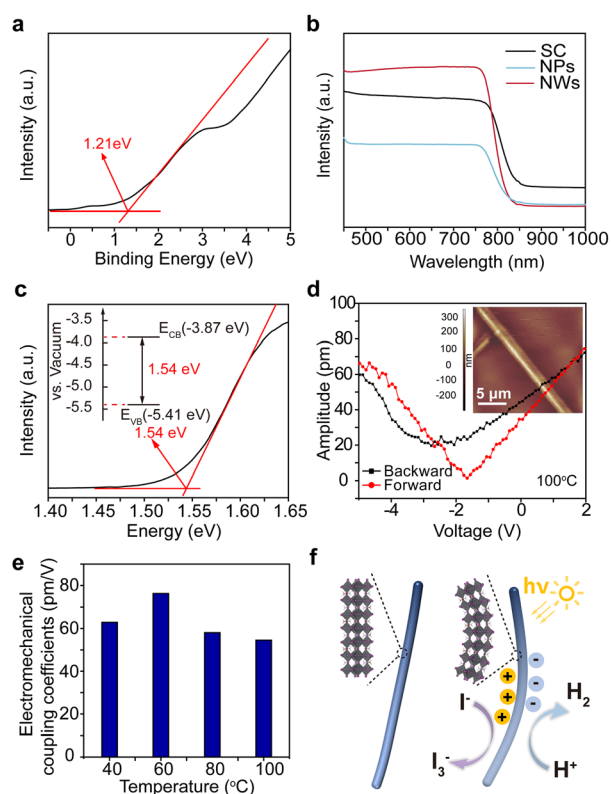


Fig. 2 Optical and electrical characterizations for MAPbI₃ NWs. (a) UPS spectrum for MAPbI₃ NWs. (b) UV-Vis absorption spectra for MAPbI₃ SC, NPs, and NWs. (c) Tauc plot for MAPbI₃ NWs; inset is the schematic diagram for the optical bandgap of MAPbI₃ NWs. (d) PFM response for an MAPbI₃ NW at 100 °C (inset: the corresponding PFM image). (e) Temperature-dependent electromechanical coupling coefficients for NWs. (f) Schematic illustration for the charge production induced by photo-mechanical coupling in MAPbI₃ NWs.



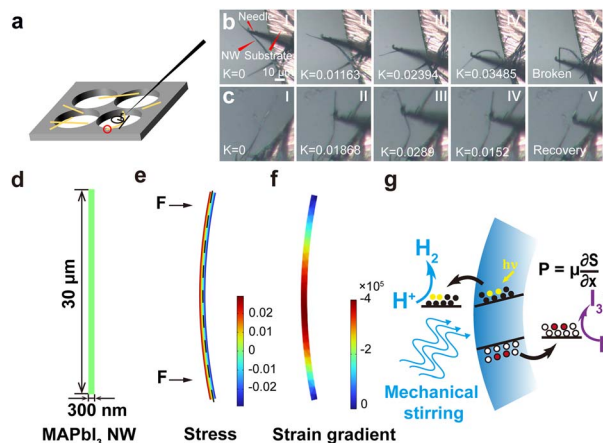


Fig. 3 Bending behaviors of MAPbI₃ NWs. (a) Schematic diagram for the *in situ* bending test of a single MAPbI₃ NW. (b) *In situ* bending of an MAPbI₃ NW until broken. (c) *In situ* bending of an MAPbI₃ NW and its recovery. (d) A single NW model for FEM. (e) FEM results for bending-induced stress distribution within a single NW. (f) FEM distributions of bending-induced strain gradient within the NW. (g) Schematic illustration for the photo-mechanical coupling in a bending MAPbI₃ NW.

NW exhibited excellent flexibility (Fig. S10; also refer to the ESI Movie†), which is suitable for repeated bending-recovery cycles in mechanocatalysis.

Moreover, finite element method (FEM) simulations were employed to survey the bending-induced strain gradient within a single NW. For a model of MAPbI₃ NW, the length is set to 30 μm, and the diameter is set to 300 nm (Fig. 3d). The stress distribution (ϵ_{11}) within the MAPbI₃ NW upon experiencing bending moments at both ends is shown in Fig. 3e, in which one side (the red area) of the NW undergoes tensile stress, while the opposite side (the blue area) experiences compressive stress (the dashed line indicates the neutral plane, where $\epsilon = 0$). Correspondingly, the strain progressively increases from the center towards the surface, while it decreases from the center to both ends of the NW along the axial direction. The strain gradient, however, manifests distinct variations across the NW due to the structural characteristics of this simply supported NW (Fig. 3f). One can see that the strain gradients at both ends of the NW are relatively low, whereas in the central region, there is significant augmentation in strain gradient (up to 10^5). This observation demonstrates that the morphology of NW facilitates the formation of strain gradients, subsequently giving rise to significant mechano-electrical coupling for NWs. The polarization induced by this strain gradient would generate an electric field (E) perpendicular to the radial direction within the MAPbI₃ NW, together with the net charges at both surfaces. In our previous work, we obtained a flexoelectric coefficient of $\sim 2000 \mu\text{C m}^{-1}$ for halide perovskites;²⁸ as a result, a polarization strength of about 200 C m^{-2} in the MAPbI₃ NW can be calculated. This significant internal polarization could lead to a completely tilted band structure for MAPbI₃ material and drive the efficient dissociation of charge carriers, facilitating electrons and holes to be separated in opposite directions. When MAPbI₃ NWs are subjected to a varying field of external mechanical stimulations (*i.e.*, magnetic stirring), the

repeated bending and recovery cycles of NWs would drive the random changes in their internal electrical fields (both strengths and directions), which induce the inclinations of the energy levels of NWS and continuous accumulation of charges (which was initiated by photo-induced separation) for catalysis (*i.e.*, photo-mechanocatalysis).⁴⁴ In this case, H⁺ readily combines with the accumulated charges to produce H₂ (Fig. 3g).

It is therefore expected that the proposed photo-mechanical coupling in MAPbI₃ NWs should deliver significant catalytic performance. Consequently, we evaluated the catalytic performance of MAPbI₃ NWs in a MAPbI₃-saturated HI solution (though not in water because the material is soluble in water, with a solubility of 0.645 mol L^{-1} at 20 °C (ref. 31)) under simultaneous light illumination and mechanical stirring (see Methods). The as-synthesized MAPbI₃ NWs were added into a homemade quartz vessel filled with MAPbI₃-saturated HI solution, which was hermetically sealed and equipped with a magnetic agitator to induce mechanical stimulation along with a xenon lamp for illumination. Gas chromatography (GC) was employed to monitor the H₂ production from the reaction system. As anticipated, a substantial quantity of H₂ was produced by the photo-mechanocatalytic HI splitting induced by the flexible MAPbI₃ NWs. The results are illustrated in Fig. 4a, which was juxtaposed with the results obtained solely through photocatalysis and mechanocatalysis of the NWs (which were driven by either light or stirring, respectively). The hydrogen generation rates for photo-mechanocatalysis, mechanocatalysis, and photocatalysis were 756.5 , 192.2 , and $164.2 \mu\text{mol g}^{-1} \text{ h}^{-1}$, respectively; in other words, the amounts of H₂ generated by the photomechanocatalysis, mechanocatalysis, and photocatalysis in our experimental conditions were recorded as 75.65 , 19.22 , and $16.42 \mu\text{mol}$ in 2 hours, respectively (Fig. 4b; the original GC data can be referred to in Fig. S11 and S12†). It is thus evident that the hydrogen production *via* the photomechanocatalytic process is significantly higher than the combined hydrogen production from the mechanocatalytic and photocatalytic processes, which suggests a remarkable synergistic effect of photo-mechanical coupling on this system. Next, to verify the effect of MAPbI₃ morphology on the photo-mechanocatalytic performance, another group of control experiments was conducted by comparing the hydrogen generation rates of MAPbI₃ NWs and NPs, and the time-dependent hydrogen evolution profiles for MAPbI₃ NWs and NPs are shown in Fig. 4c. It is clear that the NWs exhibited a significantly higher hydrogen generation rate compared to the NPs (Fig. 4d), demonstrating the geometrical advantages of NWs in flexoelectric charge production,⁴⁵ as well as in photomechanocatalytic hydrogen production.

To mitigate the interference of hydrogen generation from the self-decomposition of HI under illumination and mechanical agitation, we replicated the previously mentioned experiment in the absence of MAPbI₃ NWs. No noticeable hydrogen evolution was detected in the scenarios involving sole illumination, sole agitation, or the combination of illumination and agitation (Fig. S13†). In this context, the hydrogen generation in previous experiments can be attributed to the photomechanocatalytic cleavage of HI induced by the MAPbI₃ NWs. To the best of our knowledge, the hydrogen production rate of MAPbI₃ NWs as reported in this work is one of the highest values for the single-



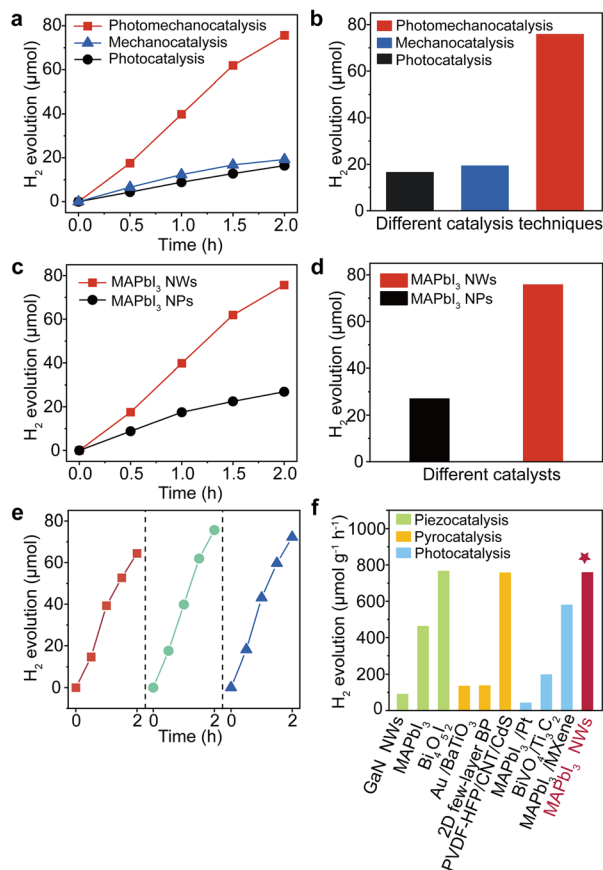


Fig. 4 Hydrogen production performances for MAPbI₃ NWs via photomechanocatalysis. (a) Time-dependent hydrogen evolution profiles for photomechanocatalysis, photocatalysis, and mechanocatalysis for MAPbI₃ NWs. (b) The comparative amounts of hydrogen generated from MAPbI₃ NWs after 2 h using different catalysis techniques. (c) Time-dependent hydrogen evolution profiles for MAPbI₃ NWs and NPs via photomechanocatalysis. (d) The comparative amounts of hydrogen generated from photomechanocatalysis after 2 h using MAPbI₃ NWs and NPs. (e) The test of stability for photomechanocatalytic hydrogen production using MAPbI₃ NWs over three cycles. (f) Comparison of hydrogen production rates among different catalytic modes and catalysts in recent reports. The listed examples include piezocatalysis (GaN NWs,³ MAPbI₃,³² Bi₄O₅I₂ (ref. 46)), pyrocatalysis (Au/BaTiO₃,⁴⁷ 2D few-layer black phosphorene,⁴⁸ PVDf-HFP/CNT/CdS⁴⁹), photocatalysis (MAPbI₃/Pt,⁵⁰ BiVO₄/Ti₃C₂ nanosheets,⁵¹ MAPbI₃/MXene⁵²), and our work.

catalyst systems in HI (Table S1†). Moreover, the solar HI splitting efficiency of our experiment can reach up to 2.67% (the calculation can be found in Methods), which represents a significant improvement compared to the existing reports. Moreover, the cycling stability of MAPbI₃ NWs as catalysts in photomechanocatalytic hydrogen production was also assessed; as shown in Fig. 4e, the hydrogen production rates have been constantly maintained across three catalytic cycles, demonstrating the good repeatability of MAPbI₃ NWs in this catalytic reaction. In addition, a comprehensive analysis was conducted to investigate the microstructural evolution of MAPbI₃ NWs during the reaction process. The XRD spectrum of MAPbI₃ NWs after catalytic reaction is presented in Fig. S14.† Compared to

the XRD pattern of NWs before the reaction (Fig. 1h), the peaks of MAPbI₃ remained essentially unchanged, indicating the NWs were basically stable in the catalytic system, while the peaks of PbI₂ disappeared as it dissolved in HI. Moreover, the NWs after different reaction times (0, 1, and 2 h) were isolated from the reaction system and further observed under SEM (Fig. S15†). In the prolonged reaction, the NWs exhibited slight surface damage due to mechanical stirring and possible photo-degradation, yet the structure of the NWs was largely preserved. We further compared the hydrogen production rate in our experiment with the existing catalytic hydrogen generation approaches, including piezoelectric, pyroelectric, and photocatalytic pathways (as shown in Fig. 4f). The catalytic activity of MAPbI₃ NWs with photo-mechanical coupling exhibits a superior rate than, or at least, is comparable to a multitude of extensively investigated piezoelectric, pyroelectric, and photocatalytic systems.

Conclusions

In summary, an effective strategy for hydrogen production using flexible MAPbI₃ NWs upon concurrent light illumination and mechanical stirring was demonstrated with a hydrogen yield rate as high as 756.5 μmol g⁻¹ h⁻¹ (the corresponding catalytic efficiency is 2.67%), which surpasses the rates achieved by sole photocatalysis and flexocatalysis by nearly five-fold and four-fold, respectively. The significant photomechanocatalysis from MAPbI₃ NWs is attributed to the photo-mechanical coupling, which tilts the band structure of MAPbI₃ under large strain gradients and thus accelerates the charge production at the surfaces of NWs upon light illumination. Our work demonstrates the potential of MAPbI₃ NWs for hydrogen production via simultaneously harvesting mechanical and solar energy from the environment and may shed light on the paradigm shift in mechanocatalysis.

Data availability

The data that support the findings of this study are available from the corresponding authors upon reasonable request.

Author contributions

L.-F. F. conceived and supervised the project. L.-F. F. and Y.-C. Z. designed the experiments. J.-W. H. conducted TEM observations and data analysis. M.-Y. Z. and Z.-Y. Z. conducted the *in situ* bending test. K.-Q. N. and X.-X. L. conducted UPS measurements. Z.-G. W. performed FEM calculations. Y.-C. Z., T.-F. T., Z. W., and L.-F. F. wrote the manuscript with the help and input of all authors. All authors have given their approval to the final version of the manuscript.

Conflicts of interest

There are no conflicts to declare.



Acknowledgements

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