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Iridium-catalyzed reductive sulfonamidation of alkoxy aryl alkynes†

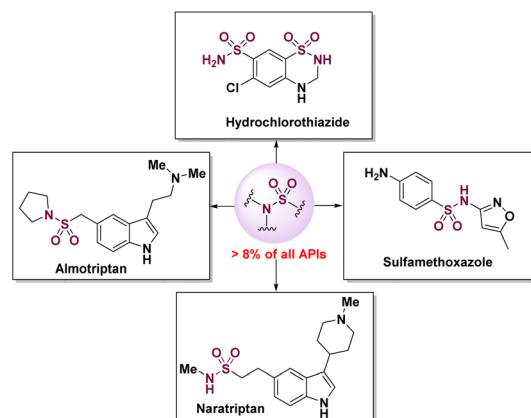
 Yuqiu Liang,^a Chengxiu Liu,^a Penghao Wei,^a Lu Ouyang[†]^{ab} and Youchun Li^{*b}

Sulfonamides are valuable structural building blocks, bioactives, and pharmaceuticals. While there have been great achievements in the sulfonamidation of alkyl and alkenyl carbon, the sulfonamidation of alkynyl carbon has not been studied. Herein, we report the synthesis of *N*-benzylated sulfonamides from alkoxy aryl alkynes and sulfonamides enabled by Ir-catalyzed reductive sulfonamidation using HCO₂H as a hydrogen donor. This process was performed under mild conditions, resulting in the transformation of a variety of substituted benzene, heteroaromatic, and aliphatic sulfonamides. Particularly, the structural diversification of valdecoxib and zonisamide showcased the utility of this protocol.

Introduction

Sulfonamides are not only valuable structural building blocks in synthetic intermediates, but are also commonly found in biological and pharmaceutical fields.¹ For instance, sulfonamides of almotriptan,² sulfamethoxazol,³ hydrochlorothiazide,⁴ and naratriptan⁵ have found applications in the treatments of heavy migraine headache, urinary tract infections, and high blood pressure, respectively (Scheme 1). The latest statistics show that over 8% active pharmaceutical ingredients (APIs) contain sulfonamide skeletons, which have special physicochemical properties of metabolic stability.⁶ Therefore, the extensive application of sulfonamides in medicinal chemistry has attracted the attention of chemists in the synthesis and functionalization of sulfonamides in recent decades.⁷

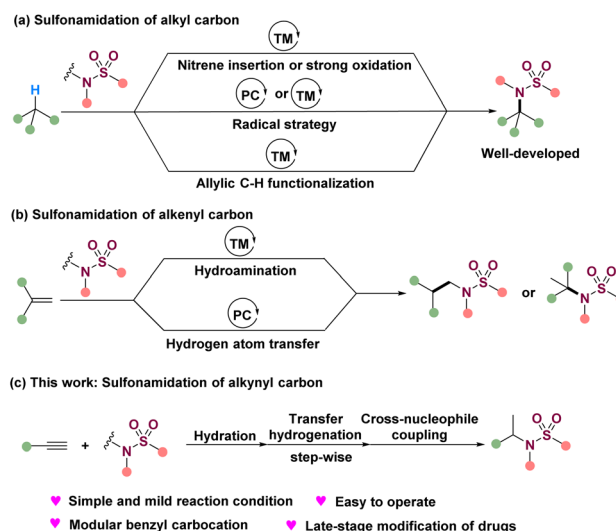
Reactions of primary sulfonamides with aliphatic halides,⁸ alcohols,⁹ and carbonyls¹⁰ present classical strategies for the synthesis of sulfonamides, in which the organic,¹¹ inorganic,¹² Ir,¹³ Ru,¹⁴ Rh,¹⁵ and other metal¹⁶ catalysts are employed. Coupling of primary sulfonamides with aryl halides,¹⁷ boronic acids,¹⁸ and diaryliodonium triflate¹⁹ constitutes another efficient approach to sulfonamide synthesis, where Cu,²⁰ Pd,²¹ and Ni²² metals are commonly utilized as catalysts. Direct sulfonamidation of alkyl carbon provides an atom- and step economy strategy for sulfonamide synthesis, with commendable substrate scope and efficiency (Scheme 2a).²³ However, the inevitable use of hypervalent iodine reagents or strong oxidation,²⁴ excessive equivalents of oxidants,²⁵ and poor



Scheme 1 Drugs containing sulfonamide motifs.

^aSchool of Pharmacy, Gannan Medical University, Ganzhou 341000, Jiangxi Province, P. R. China. E-mail: oyl3074@163.com

^bThe Affiliated Ganzhou Hospital, Jiangxi Medical College, Nanchang University, Ganzhou 341000, Jiangxi Province, P. R. China. E-mail: liyouchun2007@163.com

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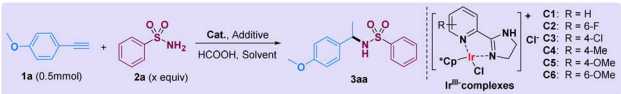
Scheme 2 Sulfonamidation of differential hybridized carbons.



regioselectivity²⁶ limit the application of this strategy. Approaches to sulfonamide synthesis based on the sulfonamidation of alkenyl carbon *via* classical hydroamination or hydrogen atom transfer of alkenes are attractive alternatives (Scheme 2b).²⁷ Notably, asymmetric sulfonamidation of alkenyl carbon for the synthesis of enantioenriched sulfonamides *via* hydrogen atom transfer has also been established.²⁸ While there have been great achievements with regards to sulfonamidation of alkyl and alkenyl carbon, the sulfonamidation of alkynyl carbon has not been studied.

With our continuous research on transfer hydrogenation with Cp*Ir complexes,²⁹ *N*-alkylation³⁰ or *para*-Friedel–Crafts alkylation³¹ were achieved from alkynes *via* hydration and transfer hydrogenation. In a previous work, a relatively stable benzyl carbocation was generated from alkynes *via* a hydration, transfer hydrogenation, and successive dehydroxylation process, which might be captured by primary or secondary sulfonamides to deliver a variety of *N*-benzylated sulfonamides. Using alkynes as substrates comes with various challenges: (a) although hydration of alkynes had been developed in our

Table 1 Optimization of the synthesis of sulfonamide **3aa**^a



Entry	Cat.	Additive (equiv.)	2a (x equiv.)	Solvent	Yield 3aa ^b (%)
1	C3	TsOH (0.6)	1.0	H ₂ O	38
2	C3	TsOH (0.6)	1.0	DMF	N.D.
3	C3	TsOH (0.6)	1.0	TFEA	N.D.
4	C3	TsOH (0.6)	1.0	MeCN	N.D.
5	C3	TsOH (0.6)	1.0	<i>p</i> -Xylene	N.D.
6	C3	TsOH (0.6)	1.0	Dioxane	N.D.
7	C3	TsOH (0.6)	1.0	TFEA/H ₂ O ^c	36
8	C3	TsOH (0.6)	1.0	TFEA/H ₂ O ^d	42
9	C3	TsOH (0.6)	1.0	TFEA/H ₂ O ^e	22
10	C3	TsOH (0.6)	1.0	TFEA/H ₂ O ^f	Trace
11	C3	TsOH (0.6)	1.0	<i>p</i> -Xylene/H ₂ O ^d	30
12	—	TsOH (0.6)	1.0	TFEA/H ₂ O ^d	N.D.
13	C1	TsOH (0.6)	1.0	TFEA/H ₂ O ^d	40
14	C2	TsOH (0.6)	1.0	TFEA/H ₂ O ^d	37
15	C4	TsOH (0.6)	1.0	TFEA/H ₂ O ^d	37
16	C5	TsOH (0.6)	1.0	TFEA/H ₂ O ^d	45
17	C6	TsOH (0.6)	1.0	TFEA/H ₂ O ^d	41
18	C5	TsOH (0.6)	1.2	TFEA/H ₂ O ^d	45
19	C5	TsOH (0.6)	1.4	TFEA/H ₂ O ^d	46
20	C5	TsOH (0.6)	1.6	TFEA/H ₂ O ^d	46
21	C5	TsOH (0.6)	1.8	TFEA/H ₂ O ^d	48
22	C5	TsOH (0.6)	2.0	TFEA/H ₂ O ^d	77
23	C5	TsOH (0.6)	2.5	TFEA/H ₂ O ^d	70
24	C5	TsOH (0.6)	3.0	TFEA/H ₂ O ^d	76
25	C5	—	2.0	TFEA/H ₂ O ^d	N.D.
26	C5	PhSO ₃ H (0.6)	2.0	TFEA/H ₂ O ^d	52
27	C5	R ¹ SO ₃ H (0.6) ^g	2.0	TFEA/H ₂ O ^d	44
28	C5	MSA (0.6)	2.0	TFEA/H ₂ O ^d	59
29	C5	R ² SO ₃ H (0.6) ^g	2.0	TFEA/H ₂ O ^d	35
30	C5	TsOH (0.2)	2.0	TFEA/H₂O^d	81 (79)
31	C5	TsOH (0.4)	2.0	TFEA/H ₂ O ^d	80
32	C5	TsOH (0.8)	2.0	TFEA/H ₂ O ^d	46
33	C5	TsOH (1.0)	2.0	TFEA/H ₂ O ^d	43
34	C5	TsOH (1.5)	2.0	TFEA/H ₂ O ^d	29
35	C5	TsOH (2.0)	2.0	TFEA/H ₂ O ^d	18
36 ^h	C5	TsOH (0.2)	2.0	TFEA/H ₂ O ^d	63
37 ⁱ	C5	TsOH (0.2)	2.0	TFEA/H ₂ O ^d	42
38 ^j	C5	TsOH (0.2)	2.0	TFEA/H ₂ O ^d	N.D.

^a Reaction conditions: **1a** (0.5 mmol), Cat. (1.0 mol%), HCO₂H (10.0 equiv.), and solvent (1.5 mL) for 12 hours at 80 °C (under air). ^b Yield was determined by NMR with dimethyl terephthalate as the internal standard. ^c The ratio of the mixed solvent was 1 : 8 (v/v). ^d The ratio of the mixed solvent was 1 : 4 (v/v). ^e The ratio of the mixed solvent was 1 : 2 (v/v). ^f The ratio of the mixed solvent was 1 : 1 (v/v). ^g R₁ = 4-Cl-Ph; R₂ = 2-naphthyl. ^h 60 °C. ⁱ 100 °C. ^j Without HCO₂H.



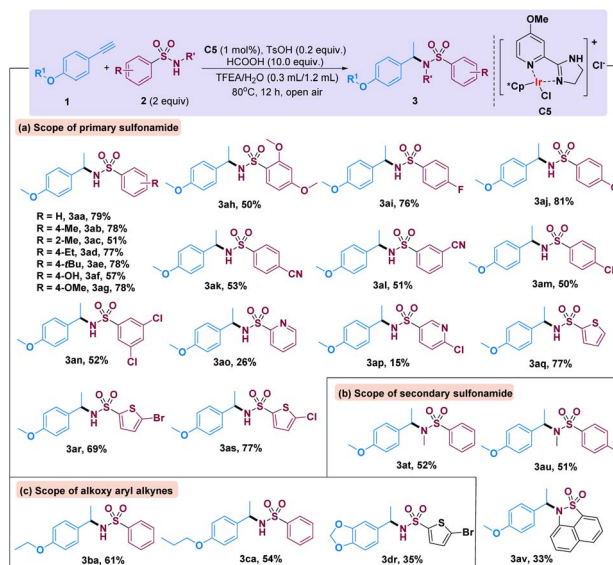
previous work,³¹ there is a risk of deactivation of hydration using primary or secondary sulfonamides as nucleophilic reagents; (b) a similar outcome of poisoning subsequent transfer hydrogenation is possible under these reaction conditions; (c) a weaker nucleophilic property is noted while using nitrogen atoms as a nucleophilic reagent under acidic conditions, and using sulfonamides as nucleophilic agents will likely have similar related issues, hindering the final cross nucleophilic coupling process. Despite these difficulties, through protracted and unremitting efforts, herein, we realized the reductive sulfonamidation of alkynes with primary and secondary sulfonamides using metal catalysis, which provides inspiration for the synthesis of diversified sulfonamides (Scheme 2c).

Results and discussion

We initially examined the reductive sulfonamidation of alkynes by employing 4-ethynylanisole **1a** and benzene sulfonamide **2a** as model substrates, Cp*Ir complexes as catalyst,³² and HCO₂H as a hydrogen donor (Table 1). Interestingly, the desired product **3aa** was produced at a 38% yield using H₂O as a solvent and TsOH as an additive (Table 1, entry 1). Screening of further reaction parameters indicated that the H₂O and Cp*Ir catalyst were essential for successful reductive sulfonamidation (Table 1, entries 2–11). Increasing the ratio of TFEA would decrease the yield of **3aa** (Table 1, entries 7–10). For instance, the yield of **3aa** was reduced to trace even though the ratio of TFEA and H₂O was loaded over 1 : 1 (Table 1, entry 10). Additionally, catalyst optimization (Table 1, entries 12–17) showed that the Cp*Ir complex **C5** could slightly enhance the sulfonamidation process leading to a 45% yield of the product **3aa** (Table 1, entry 16). Satisfyingly, increasing the loading of **2a** would sharply improve the yield of **3aa** (Table 1, entries 18–24). Of note, the control experiment demonstrated that a Lewis acid was crucial for this reductive sulfonamidation process (Table 1, entries 25–35) and decreasing the loading of TsOH to 0.2 equiv. resulted in the best yield of **3aa** (Table 1, entry 30). However, decreasing or increasing the reaction temperature was harmful to the production of **3aa** (Table 1, entries 36 and 37). Control experiment showed that HCO₂H was the essential hydrogen donor in this transformation, indicating that H₂O only act as a reaction media (Table 1, entry 38).

With the successfully optimized conditions, the substrate scope with respect to aryl alkynes and aryl sulfonamides was investigated (Table 2). As anticipated, aryl sulfonamides with electron-donating groups at differential positions were well tolerated, including methyl, ethyl, *tert*-butyl, hydroxyl, and methoxy, delivering the corresponding products (**3ab–3ag**) in good to excellent yields. Additionally, di-substituted aryl sulfonamide (**2h**) also performed well in this system. Aryl sulfonamides with electron-withdrawing groups, such as fluorine (**2i**), chlorine (**2j**, **2n**), nitrile (**2k**, **2l**), and trifluoromethyl (**2m**), were also efficient substrates to afford similar yields of the desired products (**3ai–3an**). However, significantly different

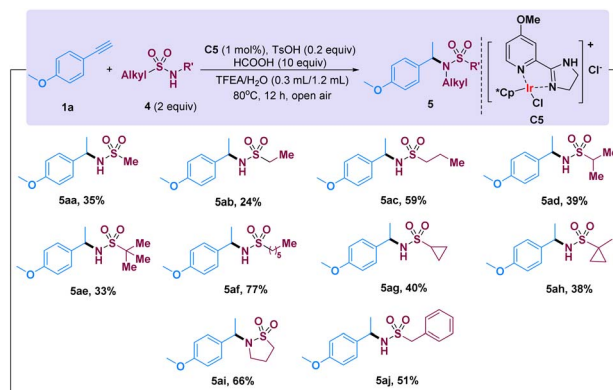
Table 2 Substrate scope of aromatic sulfonamides and alkoxy aryl alkynes^a



^a Reaction conditions: **1** (0.5 mmol), **2** (1.0 mmol), **C5** (1.0 mol%), HCO₂H (10.0 equiv.), and TFEA (0.3 mL), H₂O (1.2 mL) for 12 hours at 80 °C, and isolated yield.

yields were achieved with heteroaromatic sulfonamides (**2o–2s**) as substrates. For instance, low yield (15–26%) of corresponding products **3ao**, **3ap** were obtained using pyridine sulfonamides as substrates, while 2-thiophene-sulfonamides delivered the desired products **3aq–3as** in good yields (69–77%). This difference could be attributed to the difference in the density of π electrons. Interestingly, switching the primary aryl sulfonamides to secondary aryl sulfonamides also allowed the formation of the desired products (**3at–3av**) in moderate yields. Furthermore, other alkoxy-substituted aryl alkynes (**2b–2d**) were

Table 3 Substrate scope of alkyl sulfonamides^a



^a Reaction conditions: **1a** (0.5 mmol), **4** (1.0 mmol), **C5** (1.0 mol%), HCO₂H (10.0 equiv.), and TFEA (0.3 mL), H₂O (1.2 mL) for 12 hours at 80 °C, and isolated yield.



Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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