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Iodine/DMSO-catalyzed oxidative deprotection of *N*-tosylhydrazone for benzoic acid synthesis†

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An oxidative deprotection of tosylhydrazones has been demonstrated to afford benzoic acids using iodine and DMSO system. This efficient oxidative deprotection protocol offers exceptional functional group toleration under mild reaction conditions without any initiators or bases. Notably, the tosylhydrazone with the heteroaryl ring or with the aryl ring having base-sensitive hydroxyl and ester functional groups smoothly afforded the corresponding benzoic acid analogues under developed conditions. Moreover, this method features short reaction times, high product yields and easy purification by avoiding column-chromatographic purification.

Introduction

N-tosylhydrazones are popular synthons in synthetic organic chemistry for constructing potent cyclic scaffolds.^{1,2} The synthesis of tosylhydrazones can be straightforwardly achieved in solid form *via* condensation of carbonyl compounds with tosylhydrazine in high yields.^{3,4} However, ascribed to their appealing features and potent diazo and 1,3-dipolar precursors, tosylhydrazones have played a crucial role in the synthesis or modification of various bioactive heterocycles.^{5–10} Additionally, due to their crystalline and stable nature, tosylhydrazones and its derivatives have been used as protection and purification agents for carbonyl compounds.^{11–14} The deprotection of hydrazones into corresponding carbonyl or carboxylic acid derivatives under mild conditions is a vital process in organic synthesis. The development of mild and effective methods for the deprotection of procarbonyl compounds has been of long-standing interest of organic chemists. To date, numerous conditions or catalysts such as copper(i) chloride,¹⁵ clayfen,¹⁶ potassium bromate,¹⁷ quinolinium dichromate (QDC),¹⁸ alumina-supported ammonium chlorochromate,¹⁹ 6-benzyl-4-aza-1-azoniabicyclo[2.2.2]octane dichromate,²⁰ and Amberlyst 15 supported nitrosonium ion²¹ have been reported to accomplish the deprotection of hydrazones into corresponding carbonyl compounds. Although multiple approaches for the regeneration of carbonyl compounds from oximes^{11,22,23} are

reported in scientific literature, only a few reports are available from tosylhydrazones, especially with mild reaction conditions.

In 2000, Bandgar and his team reported the regeneration of carbonyl compounds using hexamethylenetetramine–bromine (HMTAB) and *N*-bromosuccinimide (NBS) (Scheme 1a).²⁴ Similarly, in the same year, Chandrasekhar and his team disclosed the selective cleavage of tosylhydrazone using 2,3-dichloro-5,6-dicyano-1,4-benzoquinone (DDQ) (Scheme 1b).²⁵ In 2006, Movassagh and co-workers utilized K-catalyst and calcium hypochlorite for the deprotection of tosylhydrazones (Scheme 1c).²⁶ Likewise, Jia and his group regenerated the carbonyl compounds using *meta*-chloroperbenzoic acid (*m*CPBA) (Scheme 1d).²⁷ Despite many attempts thus far, existing methods or conditions often suffer from various disadvantages *e.g.* expensive syntheses, limited substrate scope, extended reaction time, and the use of hazardous oxidants and metal ions in certain cases. Noticeably, the usage of a mild oxidation approach for the direct conversion of tosylhydrazones to corresponding carboxylic acids is still desirable. It is therefore necessary to develop new methods based on readily available oxidants and safer chemicals. In this context, iodine and DMSO system have recently received attention as mild and selective oxidizing agent. The I₂/DMSO duo is very suitable for chemoselective oxidation, dehydrogenation, oxidative aromatization, protection/deprotection of various functional groups, and regioselective and stereoselective transformations.^{28–30} On the other hand, the selective oxidation of carbonyl groups is a vital step in organic synthesis since it enables the production of acids, amides, acetals, and esters.^{31–38} Carboxylic acids are the most extensively used oxidized feedstocks and are needed in large amounts as bulk chemicals in various industries, including polymers, fine chemicals, and commercial products.³⁹ Benzoic acids^{32,39,40} and its analogues are biologically important scaffold as they are featured in many natural products and medically potent drugs (Fig. 1).^{41–44}

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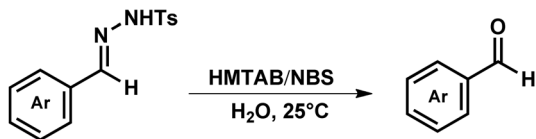
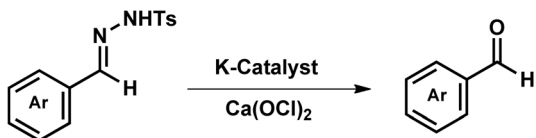
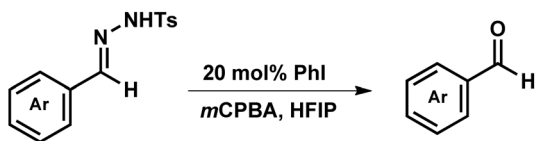
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Deprotections of Tosylhydrazones (Previous Works)

1a) Bandgar *et al.*1b) Chandrasekhar *et al.*1c) Movassagh *et al.*1d) Jia *et al.*

1e) Present work

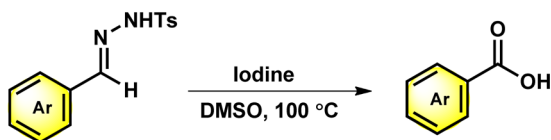
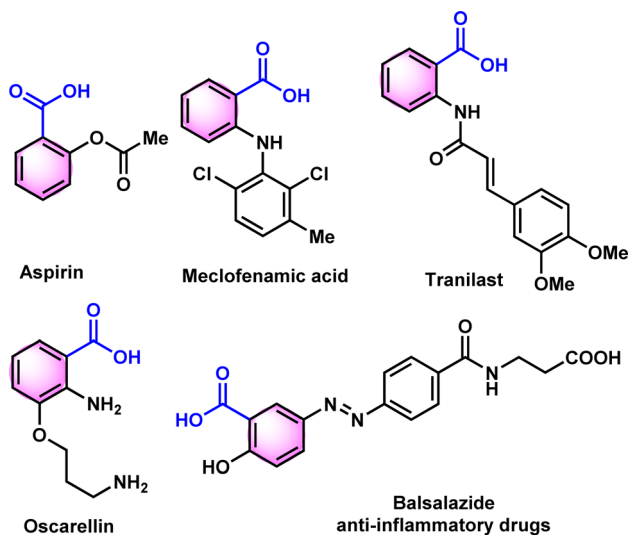
Scheme 1 Deprotection of *N*-tosylhydrazones.

Fig. 1 Representative biologically active benzoic acids.

They have a wide range of applications, including antimicrobial preservatives, antifungals, tablet and capsule lubricants, and UV protection agents.^{45–47} Utilizing a blend of medicinal plants that consist of Benzoic acid derivatives, amino acids, antioxidant chemicals, and certain minerals may be the optimal approach in formulating a phytomedicine for dealing with anemia or sickle cell disease (SCD).^{43,48–50} In 1971, Pierre and colleagues found that benzoic acid derivatives have anti-sickling properties *in vitro* using the root extract of *Fagara xanthoxyloides*.⁴⁸ Similarly, Elekwa and his colleagues have revealed the anti-sickling benefits of *p*-fluorobenzoic acid.⁵¹

Hence, by deliberately avoiding hazardous substances, we have discovered that tosylhydrazones can undergo direct oxidative deprotection to produce benzoic acid derivatives (Scheme 1e). This developed method can be used in chemical industries as a cost-effective, efficient and ecologically benign synthesis.^{52–58} Herein, we report the metal and base-free deprotection of tosylhydrazones to the oxidative product of its parent aldehyde group using iodine and DMSO system with optimal conditions.

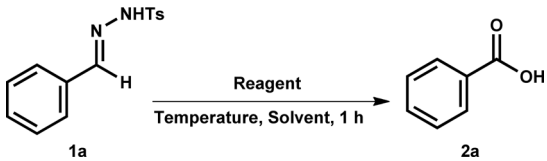
Results and discussion

The investigation started by examining the reaction conditions using tosylhydrazone **1a** as the model substrate.

Initially, reaction was conducted in the DMSO without the use of iodine at 100 °C for 1 h. Unfortunately, the reaction was failed to afford the target product **2a** in the absence of iodine (Table 1, entry 1). Next, we attempted to manipulate the reaction environment by adding various reagents such as NaI, TBAI, NH₄I and KI (Table 1, entries 2–5). However, using these reagents were futile and resulted in no product formation. Further, the use of iodine reagent in catalytic amount (10 mol% or 0.1 equiv.) in DMSO afforded the needed product **2a** in 30% yield (Table 1, entry 6). It indicates that the iodine reagent is essential for this conversion. Subsequently, increasing the equivalent of iodine from 0.1 equiv. to 0.5, 1.0 and 1.5 equiv. improved the reaction efficiency to provide the **2a** in 56%, 78% and 87% yields, respectively (Table 1, entries 7–9). Further, elevation in temperature from 100 to 120 °C also did not have any impactful effect on the productivity of this transformation (Table 1, entry 10). Reducing reaction temperature from 100 °C to room temperature and 70 °C significantly affected the reaction outcome, no reaction occurred at room temperature while only 60% yield of **2a** was observed at 70 °C (Table 1, entries 11 and 12). Moreover, we noticed the reaction outcome by switching to the other solvents. Changing solvents from DMSO to toluene, acetonitrile (MeCN), 1,4-dioxane, water, MeOH, DMF, THF, chloroform (CHCl₃) and DMA was entirely ineffective and **2a** was either not observed or obtained in poor yields (Table 1, entries 13–22). It suggested that DMSO has a critical role in this reaction transformation. Finally, we found that the use of 1.5 equiv. of iodine in DMSO at 100 °C was the optimum condition for the satisfactory yield of **2a**.

Under optimized reaction conditions, the accessibility of the protocol was observed with the use of a range of benzaldehyde tosylhydrazones **1** (Table 2). The reaction of benzaldehyde tosylhydrazones with electron-donating substituents such as

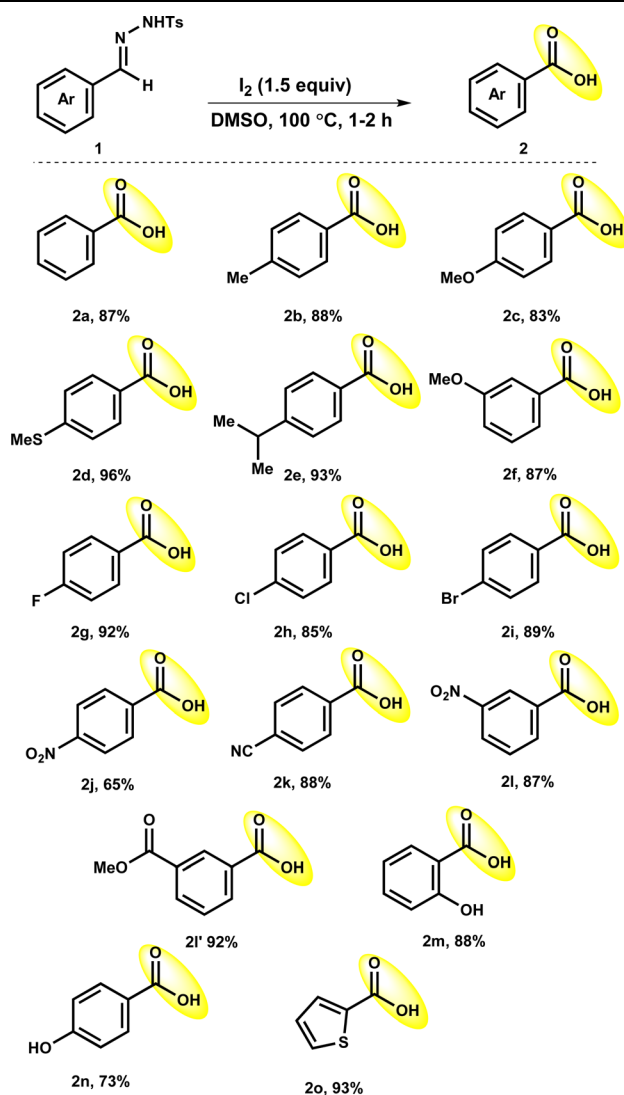


Table 1 Optimization of reaction conditions^a


Entry	Reagent (equiv.)	Solvent	T (°C)	Yield ^b (%)
1	—	DMSO	100 °C	NR
2	NaI (1.0)	DMSO	100 °C	NR
3	TBAI (1.0)	DMSO	100 °C	NR
4	NH ₄ I (1.0)	DMSO	100 °C	NR
5	KI (1.0)	DMSO	100 °C	NR
6	I ₂ (0.1)	DMSO	100 °C	30
7	I ₂ (0.5)	DMSO	100 °C	56
8	I ₂ (1)	DMSO	100 °C	78
9	I ₂ (1.5)	DMSO	100 °C	87
10	I ₂ (1.5)	DMSO	120 °C	86
11	I ₂ (1.5)	DMSO	RT	NR
12	I ₂ (1.5)	DMSO	70 °C	60
13	I ₂ (1.5)	Toluene	100 °C	15
14	I ₂ (1.5)	MeCN	80 °C	Trace
15	I ₂ (1.5)	1,4-Dioxane	100 °C	10
16	I ₂ (1.5)	H ₂ O	100 °C	NR
17	I ₂ (1.5)	MeOH	65 °C	NR
18	I ₂ (1.5)	DMF	150 °C	NR
19	I ₂ (1.5)	THF	60 °C	NR
20	I ₂ (1.5)	CHCl ₃	60 °C	NR
21	I ₂ (1.5)	DMF	100 °C	NR
22	I ₂ (1.5)	DMA	100 °C	NR

^a Reaction conditions: **1a** (0.36 mmol), I₂ (0.54 mmol), solvent (1.5 mL), 100 °C. ^b Isolated yields. NR = no reaction. RT = room temperature.

methyl (**1b**), methoxy (**1c**, **1f**), and isopropyl (**1e**) either at *meta*- or *para*-position of aryl rings proceeded efficiently to deliver the corresponding products **2b** (88%), **2c** (83%), **2e** (93%) and **2f** (87%). It is worth mentioning that the oxidizable thiomethyl (**1d**) group could sustained in the optimized reaction condition to give **2d** in a 96% yield. Tosylhydrazones of benzaldehyde having halogen substituents such as fluoro (**1g**) chloro (**1h**) and bromo (**1i**) also worked well, resulting in the production of the corresponding products **2g**, **2h** and **2i** in high yields of 92%, 85% and 89%, respectively. The strong electron-withdrawing nitro group at the *para*-position of tosylhydrazone **1j** had little impact on the oxidative deprotection process, giving **2j** a reduced yield of 65%. It was interesting to observe that other electron-withdrawing groups *para*-cyano (**1k**), nitro (**1l**) and ester (–COOMe; **1l'**) group at *meta*-position did not have any negative influence on reaction outcome, providing expected products **2k**, **2l** and **2l'** in excellent yields (87–93%). Furthermore, tosylhydrazones **1m** and **1n** with hydroxy substitutions reacted well to afford the **2m** and **2n** in good yields (73–88%). It is worth mentioning that tosylhydrazones with base-sensitive functional groups such as ester (**1l'**) and hydroxy (**1m** and **1n**) groups were well tolerated under the optimized conditions. Delightfully, the reaction of 4-methyl-*N'*-(thiophen-2-ylmethylene)benzene-sulfonohydrazide (**1o**) showed credible reactivity to afford **2o** in 93% yield.

Table 2 Oxidative deprotection of tosylhydrazones to benzoic acid derivatives^{a,b}

^a Reaction conditions: **1a** (0.36 mmol), I₂ (0.54 mmol), solvent (1.5 mL), 100 °C. ^b Isolated yields.

Some of the control studies were designed to understand the mechanistic pathway (Scheme 2). Firstly, when the reaction of benzaldehyde **3** was performed under optimized reaction conditions, benzoic acid (**2a**) was observed only in trace amounts (Scheme 2a). Next, we speculated that *in situ* generated *p*-TSA acid might be involved as a catalyst in benzaldehyde to benzoic acid formation. Therefore, a reaction of **3** was performed with 0.5 equiv. of *p*-toluenesulfonic acid (*p*-TSA), but **2a** formation was not observed. Further, we carried out the reaction mixture's LCMS analysis for the reaction of **1b** under standard conditions. LCMS data suggested the generation of *p*-toluenesulfonic acid (TsH) instead of PTSA (see the ESI†). Hence, these experiments suggested no involvement of benzaldehyde intermediate in the developed oxidative deprotection process of *N*-tosylhydrazone.



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