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Ball-milling preparation of ZnFe₂O₄/AgI nanocomposite with enhanced photocatalytic activity†

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Semiconductor photocatalytic technology is increasingly being utilized in wastewater treatment due to its high efficiency, low energy consumption and environmental friendliness. However, single photocatalysts often exhibit low catalytic performance. In this study, a ZnFe₂O₄/AgI composite photocatalyst was initially prepared using a high-energy ball-milling method. For the first time, it was applied to the photocatalytic dehydrogenation of diethyl 1,4-dihydro-2,6-dimethylpyridine-3,5-dicarboxylate (1,4-DHP), as well as photocatalytic degradation of harmful substances such as amaranth (AM), methyl orange (MO) and indole present in wastewater. The composite photocatalyst exhibited superior catalytic performance compared to ZnFe₂O₄ and AgI under visible light irradiation ($\lambda \geq 400$ nm). With optimized composition, the pseudo-first-order rate constants of ZnFe₂O₄/AgI-50% were approximately 6, 20, 64 and 38 times higher than that of AgI for the photooxidation of 1,4-DHP, AM, MO and indole, respectively. The enhanced catalytic activity of the composite was attributed to the formation of heterojunction between ZnFe₂O₄ and AgI, which facilitated the separation and transfer of photogenerated charge carriers. Mechanism studies revealed that photogenerated holes (h⁺) and superoxide radical anions ([•]O₂⁻) played pivotal roles in the photocatalytic reaction process.

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Introduction

Semiconductor-based photocatalysis has been extensively studied for their critical role in environmental remediation and renewable clean energy technologies since its initial discovery.^{1–3} For example, it can be used for dyes degradation such as MB and MO,^{4,5} reduction of Cr(VI),^{6,7} reduction of CO₂,^{8,9} and other fields.¹⁰ However, a single photocatalyst inevitably encounters certain challenges, such as weak redox abilities of photoinduced electrons (e⁻) and holes (h⁺),¹¹ poor capacity for visible light adsorption and fast recombination of charge carriers.¹² To address these unfavorable factors and further enhance the photocatalytic performance of a single semiconductor, an effective strategy is to introduce another photocatalyst to construct a heterojunction composite.^{13,14}

Spinel zinc ferrite (ZnFe₂O₄) has attracted significant attention in the field of photocatalysis due to its strong magnetic recyclability,^{15,16} high photochemical stability,¹⁷ narrow band gap (~1.9 eV),^{13,18} visible-light activity,¹⁹ and environmental

friendliness.^{20,21} Despite extensive investigations into the photocatalytic performances of ZnFe₂O₄ with different morphologies and sizes,^{22,23} it has been found that a single component of ZnFe₂O₄ often exhibits unsatisfactory photocatalytic activity, sometimes even showing nearly no catalytic effect.^{13,24} This is attributed to the rapid recombination of photogenerated e⁻-h⁺ pairs^{25,26} under light irradiation, which marvelously hinders its practical applications. Nevertheless, it has been demonstrated that the construction of ZnFe₂O₄ with other semiconductors through the formation of heterojunction structure can effectively improve its activity.^{13,15,27} For example, Zhang *et al.* prepared ZnFe₂O₄/BiOI hybrid photocatalysts using the solvothermal method. While single ZnFe₂O₄ exhibited negligible photocatalytic activity, the 5%-ZnFe₂O₄/BiOI composite showed optimal degradation efficiency (81.2%) within 100 minutes under visible light irradiation.²⁰ Similarly, Dai *et al.* synthesized magnetic ZnFe₂O₄@ZnSe hollow nanospheres using a two-step method and utilized them for the photocatalytic splitting of water. Despite the negligible photocatalytic activity of pure ZnFe₂O₄, the ZnFe₂O₄@ZnSe-20 composite exhibited high hydrogen production under light irradiation.²⁸

Silver-containing materials have been extensively studied and utilized in various fields, including biomedical application,²⁹ antibacterial agents,^{30,31} photocatalysis,^{30,32} water disinfection,³³ and other related fields.^{34,35} Silver iodide (AgI) has garnered significant attention as a photocatalyst due to its exceptional electrical conductivity, strong absorption of visible

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light,³⁶ and a narrow band gap (~ 2.8 eV).^{37,38} However, pure AgI is unstable under light irradiation and can easily be reduced to metallic silver (Ag), which weakens its catalytic activity.³⁹ To improve the stability of AgI, it can be composited with other materials.^{36,37,39} For example, Xu and coworkers have firstly prepared ZnFe₂O₄/AgI composites successfully with a hydrothermal method, and found that ZnFe₂O₄ nanoparticles were uniformly decorated on the surface of AgI particles. The obtained composite catalysts exhibited excellent photocatalytic disinfection of *E. coli* and degradation of rhodamine B under visible light irradiation due to enhanced the light harvesting ability and improved the separation efficiency of the photo-generated charge carriers.³⁷

Ball milling method is a mechanical technique utilized for the purpose of grinding powders into fine particles.⁴⁰ Additionally, it can cause changes in the structure, oxygen vacancy and even trigger chemical reactions.^{32,41} In recent time, ball milling method has been widely employed for the large-scale preparation of various catalytic materials due to its simplicity, scalability, low cost and environmental friendliness.^{42,43} In this study, ZnFe₂O₄/AgI composites with varying mass proportions of AgI and ZnFe₂O₄ were prepared using the ball-milling method. The photocatalytic performance of the ball-milling prepared ZnFe₂O₄/AgI composite was assessed by the dehydrogenation 1,4-DHP and photocatalytic degradation of AM, MO and indole. Furthermore, we compared the photocatalytic performance of the ball-milling prepared ZnFe₂O₄/AgI composite with that of counterparts prepared *via* deposition-precipitation and a simple mixture of ZnFe₂O₄/AgI. The primary reactive species involved in the photocatalytic process were examined through the application of various chemical scavengers and the electron spin resonance (ESR) technique.

Experimental methods

Materials

The materials information is provided in the ESI (Section 1).†

Synthesis of ZnFe₂O₄/AgI composites

Detailed preparation of ZnFe₂O₄ and AgI can be found in the ESI (Section 2).†

ZnFe₂O₄/AgI composites were prepared by a ball milling method. Typically, the mass ratio of ball to material is 25:1. Specifically, 0.2 g ZnFe₂O₄ and 0.2 g AgI were mixed and placed in a stainless-steel jar, which was fixed on a planetary ball mill (XGB04 type). The ZnFe₂O₄ and AgI mixture was continuously milled with stainless steel balls in the jar for 2 h at 400 rpm. This sample was labeled as ZnFe₂O₄/AgI-50% (ZA-50%), for the mass percentage of AgI was 50%. Other ZnFe₂O₄/AgI composites were prepared using the same procedure as described above, with varying mass proportion of AgI while maintaining the total mass at 0.4 g. The resulted composites were denoted as ZnFe₂O₄/AgI-x% or ZA-x% (where x represents the mass percentage of AgI).

In order to further confirm the high efficiency of the ball milling method, the best proportion ZA-50% was used as comparison. ZA-50% composite prepared by deposition-

precipitation method (DP-ZA-50%) was carried out as following: 80 mg of previously synthesized ZnFe₂O₄ was weighed and added to 30 mL of DI for ultrasonic dispersion. Then, 3.4 mL of AgNO₃ (0.1 mol L⁻¹) and 1.7 mL of KI (0.2 mol L⁻¹) were added dropwise, followed by stirred for 1 h. The precipitate was collected by centrifugation, washed three times using deionized (DI) water and dried at 60 °C. For simple mixing, 25 mg each of ZnFe₂O₄ and AgI were weighed and thoroughly mixed using a glass rod. The resulting sample was designated as SM-ZA-50%.

Characterization

The characterization information is provided in the ESI (Section 3).†

Photocatalytic evaluation

The photocatalytic performance of ZnFe₂O₄/AgI composites was evaluated by the photooxidative dehydrogenation of 1,4-DHP (0.1 mmol L⁻¹) in a solution prepared with EtOH and DI water (v/v = 1:1).^{44,45} The dehydrogenation efficiency was determined by measuring the absorbance at 374 nm.⁴⁶ In a beaker, 50 mg of ZnFe₂O₄/AgI sample and 50 mL of 1,4-DHP solution were mixed and stirred for 30 minutes under dark condition to achieve an absorption-desorption equilibrium. After equilibrium, about 2.5 mL of the reaction suspension was taken and filtered through a 0.22 μm filter to remove the photocatalyst. The beaker was then placed under a metal halide lamp (70 W) with a UV cut-off filter ($\lambda \geq 400$ nm) to obtain visible light. Samples were taken every minute and tested the absorbance using a Shimadzu UV-2550 spectrophotometer.

The degradation process of amaranth (AM, 30 mg L⁻¹), methyl orange (MO, 10 mg L⁻¹), and indole (20 mg L⁻¹) followed a similar procedure to that of 1,4-DHP. However, the absorbance measurements were taken at different wavelengths: 520 nm⁴⁷ for AM, 464 nm⁵ for MO, and 270 nm⁴⁸ for indole.

Photoelectrochemical measurements

The photoelectrochemical measurements were performed on a CHI 760E electrochemical workstation with a typical three-

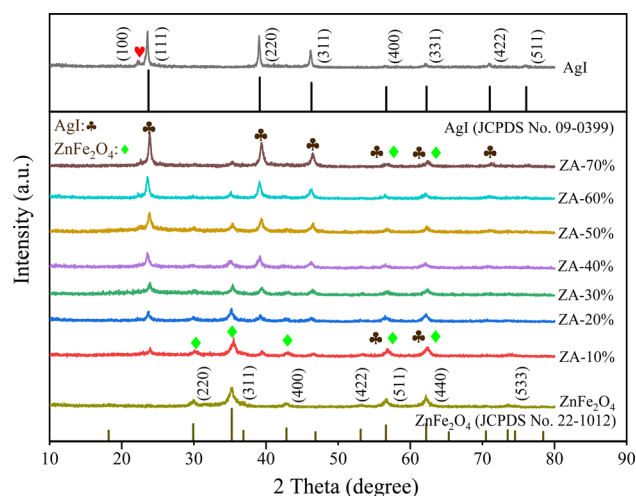


Fig. 1 XRD patterns of AgI, ZnFe₂O₄ and ZnFe₂O₄/AgI composites.



electrode cell. Visible light was generated by a 300 W xenon lamp equipped with a 400 nm cut-off filter. The working electrode was prepared by drop-casting the photocatalyst onto ITO conductive glass. A platinum (Pt) wire and a silver/silver chloride (Ag/AgCl) electrode were used as the counter electrode and reference electrode, respectively. A Na_2SO_4 solution (0.2 mol L^{-1}) was utilized as the electrolyte.

The photocurrent time ($i-t$) curves were obtained under a typical switching cycle of intermittent visible light

illumination with a bias voltage of 0.8 V. The photocurrent was recorded by turning the visible light on or off for a duration of 20 s. Additionally, the electrical impedance spectroscopy diagram (EIS) was performed with an amplitude of 5 mV at a potential of 1.38 V in the frequency range of 0.1 to 100 kHz. Mott-Schottky plots were carried at 1 kHz, and the measured potentials *versus* Ag/AgCl were converted to the normal hydrogen electrode (NHE) scale by the equation of $E_{\text{NHE}} = E_{\text{Ag/AgCl}} + 0.197 \text{ V}$.

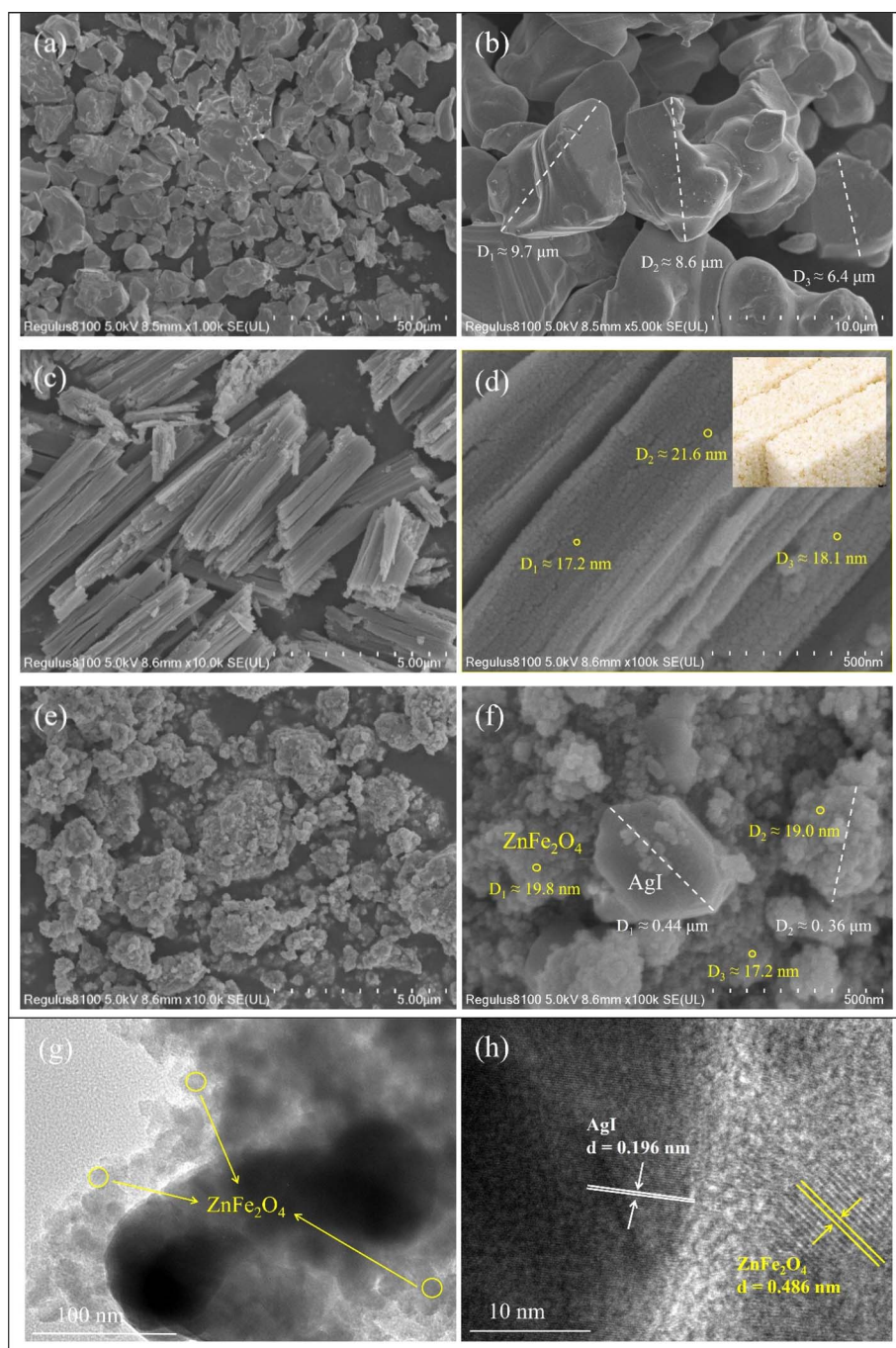


Fig. 2 SEM images with different magnifications: AgI (a and b), ZnFe₂O₄ (c and d) and ZA-50% (e and f); TEM (g) and HRTEM (h) images with lattice fringes of AgI and ZnFe₂O₄ highlighted in the figure.



Photocatalytic mechanism

In order to elucidate the principle of dehydrogenation of 1,4-DHP and decolorization of AM by this composite, active species capture experiments were performed. EDTA-2Na (10 mmol L⁻¹) and IPA (10 mmol L⁻¹) were used as a hole (h⁺) trapping agent and hydroxyl radical ([•]OH) trapping agent, respectively. N₂ was introduced to remove dissolved oxygen and to investigate the role of [•]O₂⁻.

Results and discussion

Structure and morphology characterization of photocatalysts

Fig. 1 shows the XRD patterns of ZnFe₂O₄, AgI and ZnFe₂O₄/AgI composites. The diffraction peaks of ZnFe₂O₄ correspond to the cubic crystal phase, with distinct diffraction peaks at 2θ values of 29.92° (220), 35.26° (311), 42.84° (400), 53.11° (422), 56.63° (511), 62.21° (440), and 73.51° (533). These values are in good agreement with the standard data of ZnFe₂O₄ (JCPDS No. 22-1012).²⁰ Similarly, AgI exhibits a cubic crystal phase with characteristic peaks at 23.71° (111), 39.13° (220), 46.31° (311), 56.67° (400), 62.26° (331), 71.03° (422), and 76.08° (511) corresponding to AgI (JCPDS No. 09-0399).²⁸ However, the synthesized AgI shows an additional peak (marked by ♥) when compared to the cubic crystal phase (Fig. 1), which corresponds to a characteristic feature of the hexagonal crystal phase of AgI (JCPDS No. 09-0374);^{21,28} 22.32° (100). For the ZnFe₂O₄/AgI composites, as the proportion of AgI increased, the characteristic peaks of AgI (marked by ♣) gradually appeared and strengthened while those for ZnFe₂O₄ (marked by ♦) showed an opposite trend. Apart from these characteristic diffraction peaks, no impurity

peaks were observed. It is worth noting that the diffraction peaks corresponding to specific proportions, such as ZA-70%, exhibited a slight shift towards higher angle regions for the crystal planes of (111), (220), and (311), indicating a reduction in interatomic distance resulting from compressive stress.⁴⁹ This phenomenon indicated that ZnFe₂O₄ and AgI could cause slight lattice shrinkage during ball milling mixing,⁵⁰ but the crystal phase remained unchanged, and pure composite materials were successfully prepared.

To visually observe the microstructure of the samples, SEM test was conducted. As shown in Fig. 2a and b, AgI exhibited a stone-like morphology with a diameter of approximately 5–10 μm. ZnFe₂O₄ displayed a distinct rod-shaped structure (Fig. 2c and d), consisting of many small holes with diameters of about 20 nm, arranged in a configuration that resembles rice candy (Fig. 2d, inset). The microstructure of the ZA-50% composite obtained by ball milling method was depicted in Fig. 2e and f, revealing a large number of small particles with diameters of about 20 nm and an obvious stone-like structure. Based on the previously described morphology of AgI and ZnFe₂O₄, it was further hypothesized that the small particles were the components of ZnFe₂O₄, and the rod-like structure was destroyed after ball milling, leaving only the small particles, and no significant difference in particle size was found.⁵¹ After ball milling, some small particles attached to the stone-like structure of AgI, it can also prove that ZnFe₂O₄ and AgI are fully mixed to form composites, and the size of AgI was decreased from 5–10 μm to about 0.4 μm.

In order to further observe the microscopic morphology of the ZA-50% composite, it was further characterized by TEM and

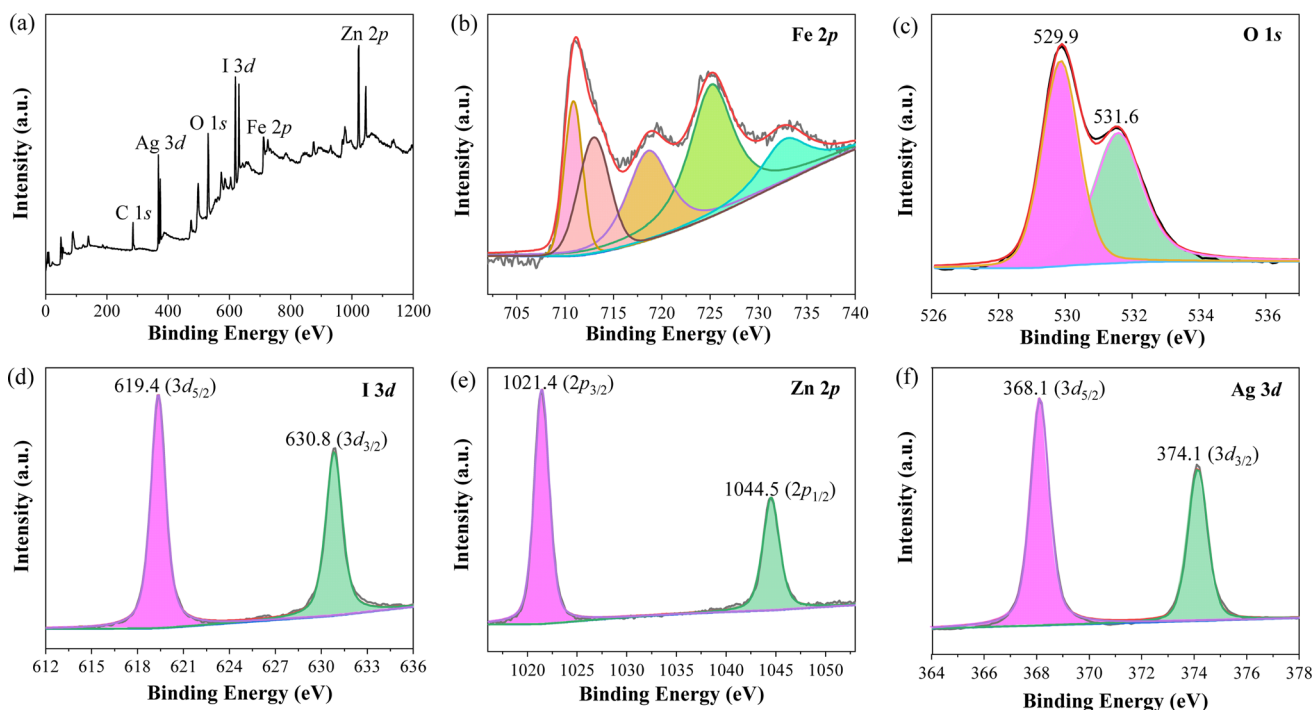


Fig. 3 High-resolution XPS spectra of ZA-50% composite: the complete XPS survey spectrum (a), Fe 2p (b), O 1s (c), I 3d (d), Zn 2p (e) and Ag 3d (f).



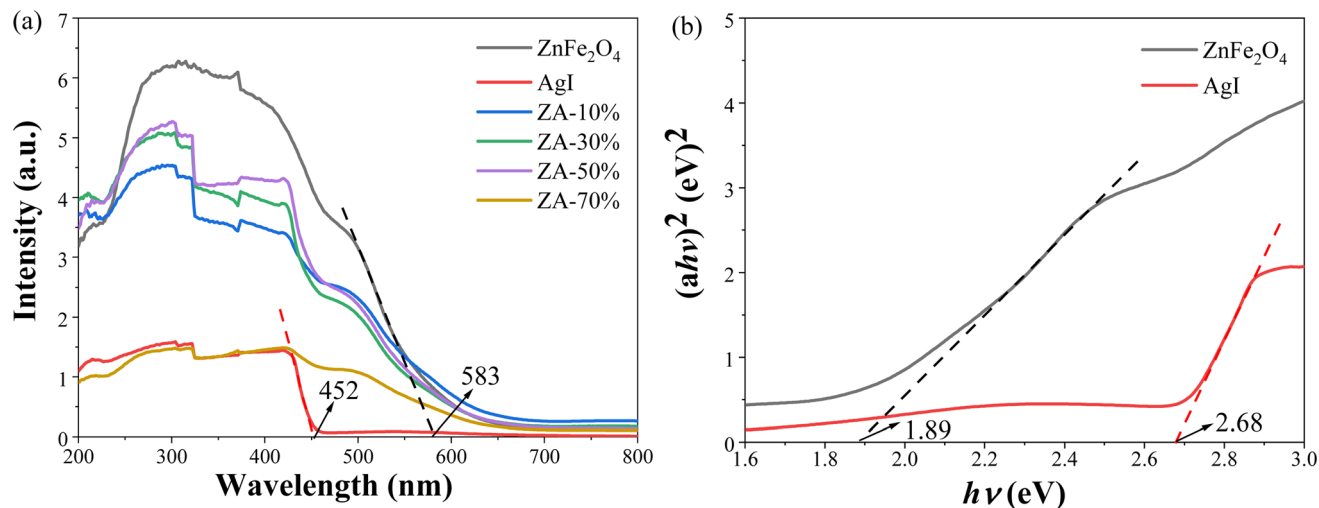


Fig. 4 UV-Vis diffuse reflectance spectra of ZnFe_2O_4 , AgI and ZA-x% (a), calculated band gap of ZnFe_2O_4 and AgI with $(\alpha h\nu)^2$ graph of relationship with energy ($h\nu$) (b).

HRTEM. The images in Fig. 2g and h show obvious bright and dark interlacing phenomena indicating that AgI was tightly bound to ZnFe_2O_4 . In Fig. 2g, the bright area consists of many small particles which were identified as components of ZnFe_2O_4 through SEM analysis. Additionally, in Fig. 2h, the (111) crystal surface belonging to ZnFe_2O_4 (JCPDS No. 22-1012) was observed in the bright area with a lattice spacing of 0.486 nm.^{52,53} On the other hand, the dark area was confirmed as AgI, the (311) crystal

surface of AgI was also observed (JCPDS No. 09-0399), where the lattice spacing was measured as 0.196 nm (Fig. 2h). These results collectively indicated that AgI and ZnFe_2O_4 were tightly combined to form a heterojunction composite after high-energy mechanical ball milling.

XPS survey spectrum was performed to analyze the chemical compositions of the ZA-50% nanocomposite. As shown in Fig. 3a, It was revealed that the composite comprises Zn, Fe, O,

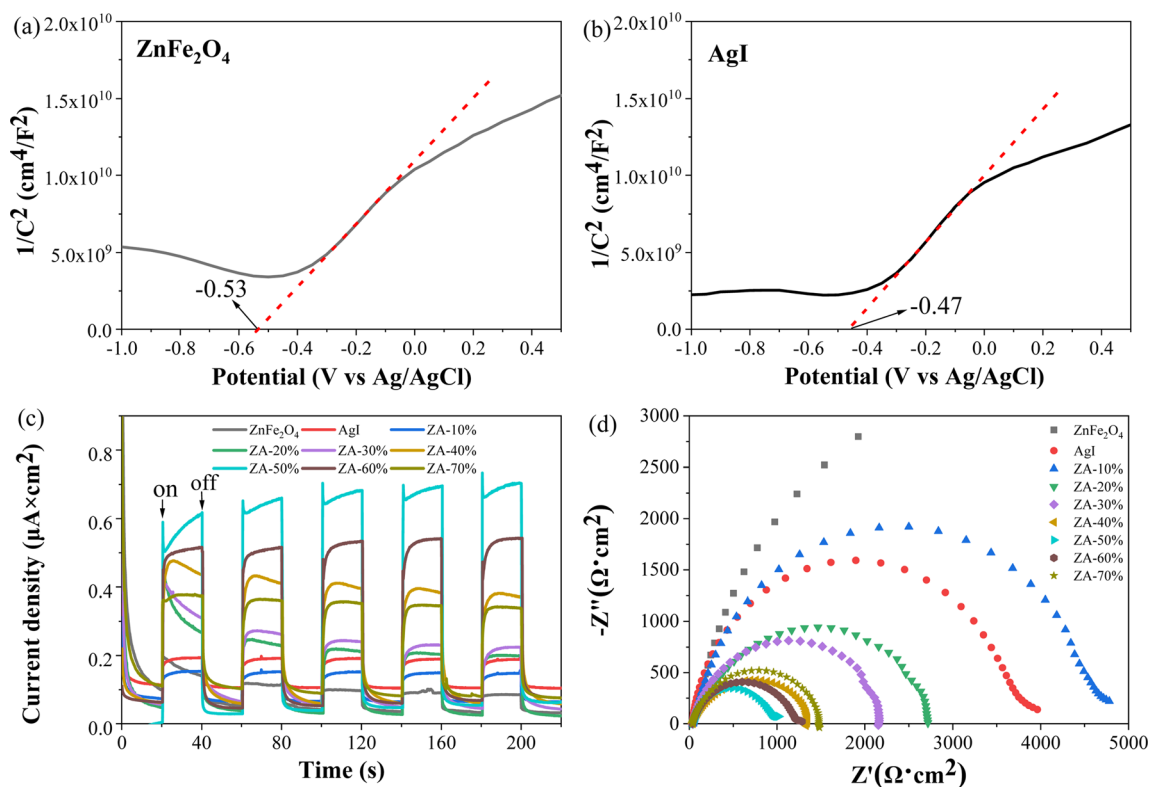


Fig. 5 Mott-Schottky curves of ZnFe_2O_4 (a) and AgI (b), transient photocurrent responses (c), EIS Nyquist plots of the ZnFe_2O_4 , AgI and ZA-x% (d).



Ag and I elements. The presence of C 1s (284.8 eV) is attributed to the adventitious hydrocarbon from the XPS instrument itself.⁵² The Fe 2p spectrum (Fig. 3b) displayed binding energies at 710.8 and 713.1 eV, which corresponded well with the Fe 2p_{3/2},^{54,55} Additionally, a peak at binding energy of 725.3 eV aligned closely with the Fe 2p_{1/2},⁵⁴ while the other peaks at 718.8 and 733.2 eV originated from the shake-up satellite of Fe³⁺.^{56,57} From Fig. 3c, two deconvoluted peaks at 529.9 and 531.6 eV were the indicative of lattice oxygen and the defect sites associated with the absorbed water on the sample surface.^{56,58} From Fig. 3d, the high resolution I 3d spectrum of ZA-50% composite exhibited two distinct peaks at 619.4 and 630.8 eV for I 3d_{5/2} and I 3d_{3/2}, respectively.⁵⁹ In the Zn 2p spectra (Fig. 3e), peaks at 1021.4 and

1044.5 eV were assigned to Zn 2p_{3/2} and Zn 2p_{1/2} respectively, which is characteristic of Zn²⁺.^{56,60} The Ag 3d spectrum of ZA-50% displayed peaks for Ag 3d_{5/2} and Ag 3d_{3/2} at binding energies of 368.1 and 374.1 eV, respectively (Fig. 3f). This is characteristic of Ag⁺ according to the literature.⁵⁶

Optical properties analysis

Fig. 4a shows the UV-Vis diffuse reflectance spectra of ZnFe₂O₄, AgI and ZA-10%, 30%, 50%, and 70%. All samples showed strong absorption in the visible region, especially ZnFe₂O₄, and the absorption edges of ZnFe₂O₄ and AgI were about 583 and 452 nm, respectively. For the composites, the light absorption initial increased gradually and then decreased with respect to

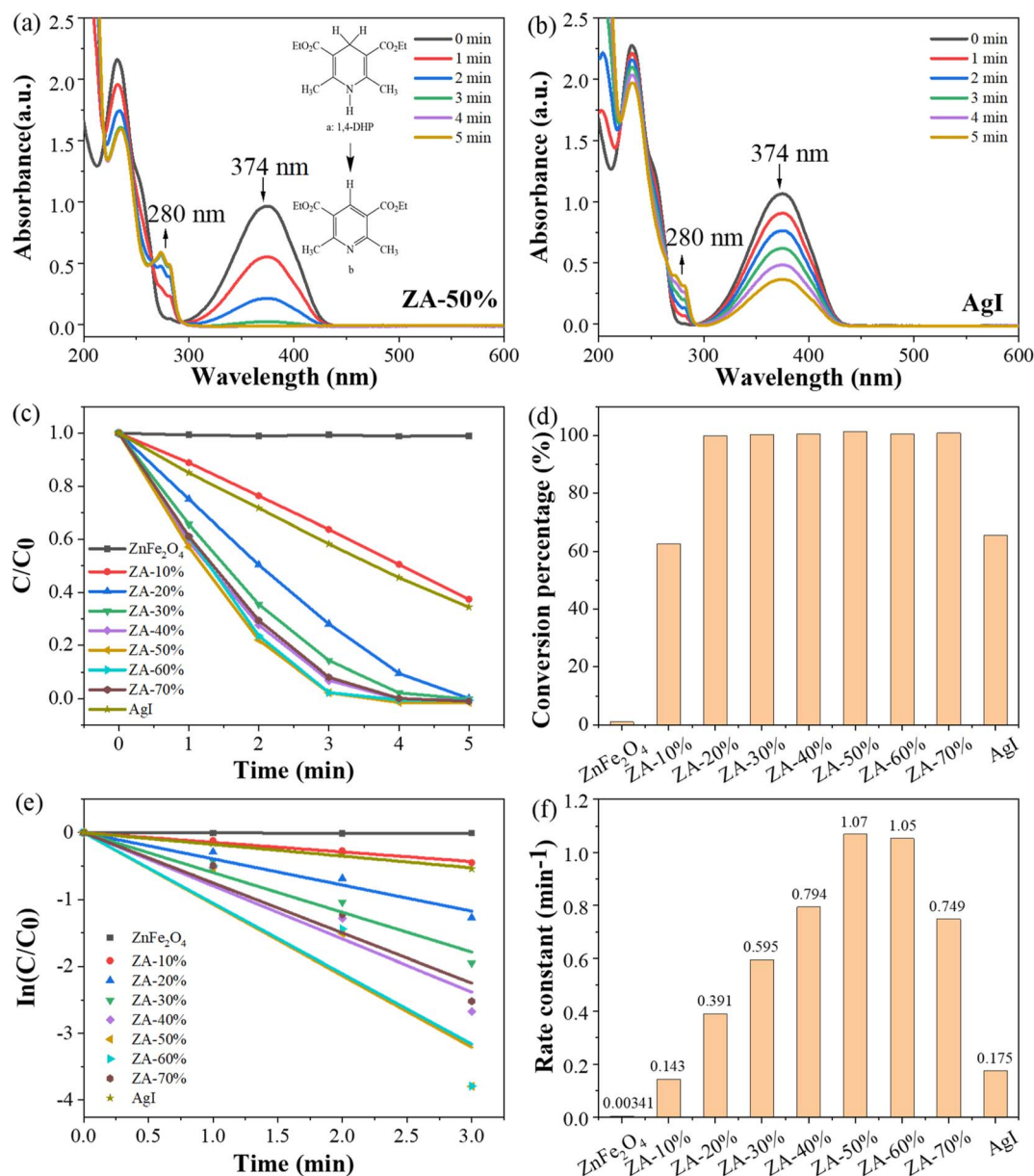


Fig. 6 UV-Vis spectral changes of the 1,4-DHP solution in the presence of ZA-50% (a), AgI (b); ZnFe₂O₄, AgI and ZA-x% photocatalytic oxidation of 1,4-DHP: kinetic diagram (c), dehydrogenation efficiency diagram (d), linear kinetic fitting graph (e), Pseudo-first-order rate constants (f) under visible light irradiation ($\lambda \geq 400$ nm).



the mass ratio of AgI. The band-gap energy (E_g) of ZnFe_2O_4 and AgI were calculated from the DRS data and Kubelka–Munk formula:⁴⁴

$$(\alpha hv) = A(hv - E_g)^{n/2}$$

where α is the absorption coefficient, h is the Planck constant, ν is the incident light frequency, A is a constant, E_g is the band gap energy and n is an integer. The n depends on where the semiconductor is direct ($n = 1$) or indirect ($n = 4$). Both ZnFe_2O_4 (ref. 21 and 37) and AgI ^{14,61} are direct semiconductors ($n = 1$), and the E_g values ZnFe_2O_4 (1.89 eV) and AgI (2.68 eV) were obtained by plotting $(\alpha hv)^2$ versus $h\nu$ and extrapolating the straight line to the horizontal axis (Fig. 4b).

Electrochemistry analysis

Mott–Schottky plots were measured within the voltage from -1.0 to 1.0 V to derive the type and energy band position of the semiconductors. The positive slope of the curves (Fig. 5) indicated that both ZnFe_2O_4 and AgI are n-type semiconductors. The flat-band potentials of ZnFe_2O_4 and AgI were -0.53 V (Fig. 5a) and -0.47 V (Fig. 5b) (vs. Ag/AgCl), respectively. According to the equation of $E_{\text{NHE}}(\text{V}) = E_{\text{Ag/AgCl}}(\text{V}) + 0.197$ V, the flat-band potentials of ZnFe_2O_4 and AgI were -0.33 V and -0.27 V (vs. NHE). In generally, the flat-band potential of an n-type semiconductor is 0.1 V lower than the conduction band

(CB) potential (E_{CB}).^{59,62} Therefore, the CB potentials of ZnFe_2O_4 and AgI were about -0.43 and -0.37 V (vs. NHE), respectively.

The valence band (VB) potential (E_{VB}) were calculated using the following formula:⁴⁵

$$E_{\text{VB}} = E_{\text{CB}} + E_g$$

The VB potentials of ZnFe_2O_4 and AgI were calculated to be 1.46 and 2.31 V (vs. NHE), respectively.

From Fig. 5c, it could be seen that the ZA-50% displayed the highest photocurrent density under light irradiation, indicating that the separation efficiency of e^- and h^+ were the highest. Similarly, the differences in charge carriers can be seen by comparing the arc radius from the electrochemical impedance spectra (EIS) Nyquist and a smaller radius indicates superior electronic conductivity and faster charge transfer.⁶³ The arc radius of ZA-50% was the smallest (Fig. 5d), indicating that its photogenerated carriers had the best transmission speed after photoexcitation,⁶⁴ inhibiting the recombination of photo-generated carriers.

Evaluation of photocatalytic performance

The photocatalytic activity of the $\text{ZnFe}_2\text{O}_4/\text{AgI}$ composite was evaluated by the photocatalytic dehydrogenation of 1,4-DHP under visible light irradiation ($\lambda \geq 400$ nm). As shown in Fig. 6a, the intensity of the characteristic peak of 1,4-DHP (374 nm)

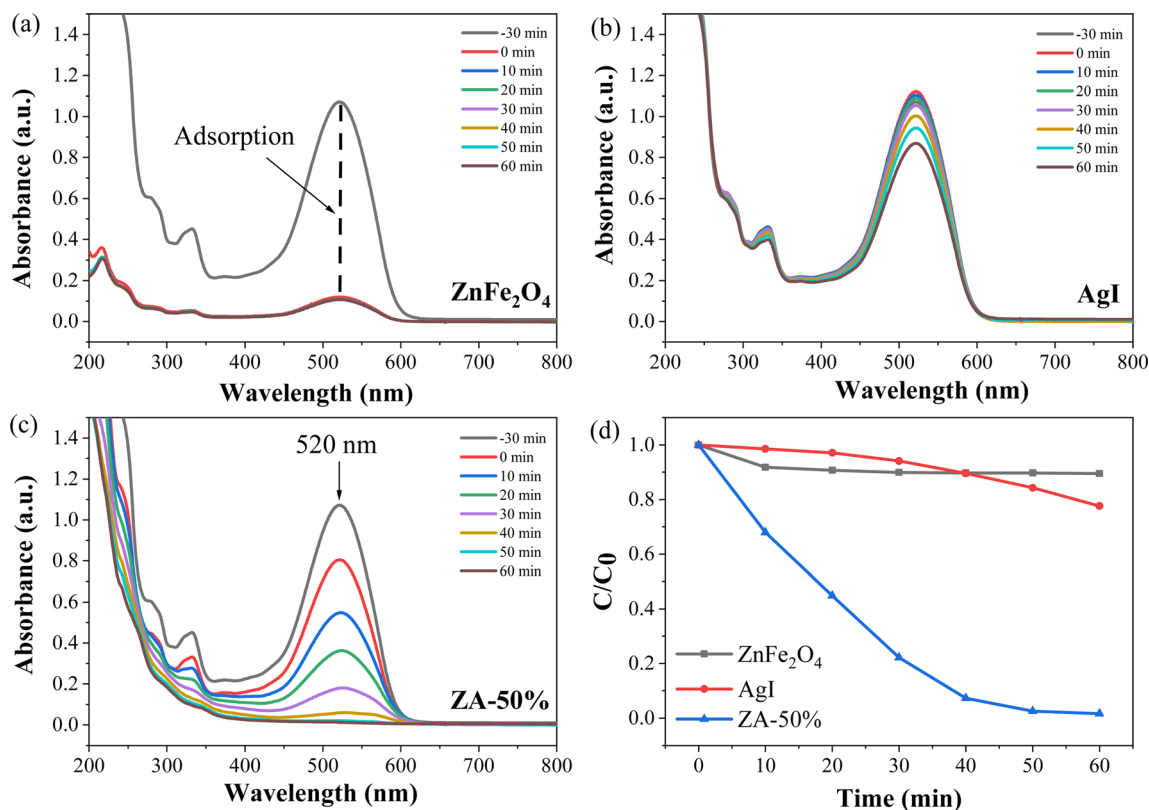


Fig. 7 UV-vis spectral changes of the AM solution under visible light irradiation ($\lambda \geq 400$ nm) in the presence of ZnFe_2O_4 (a), AgI (b), ZA-50% (c), the kinetic diagram of ZnFe_2O_4 , AgI and ZA-50% to degrade AM (30 mg L^{-1}) (d).



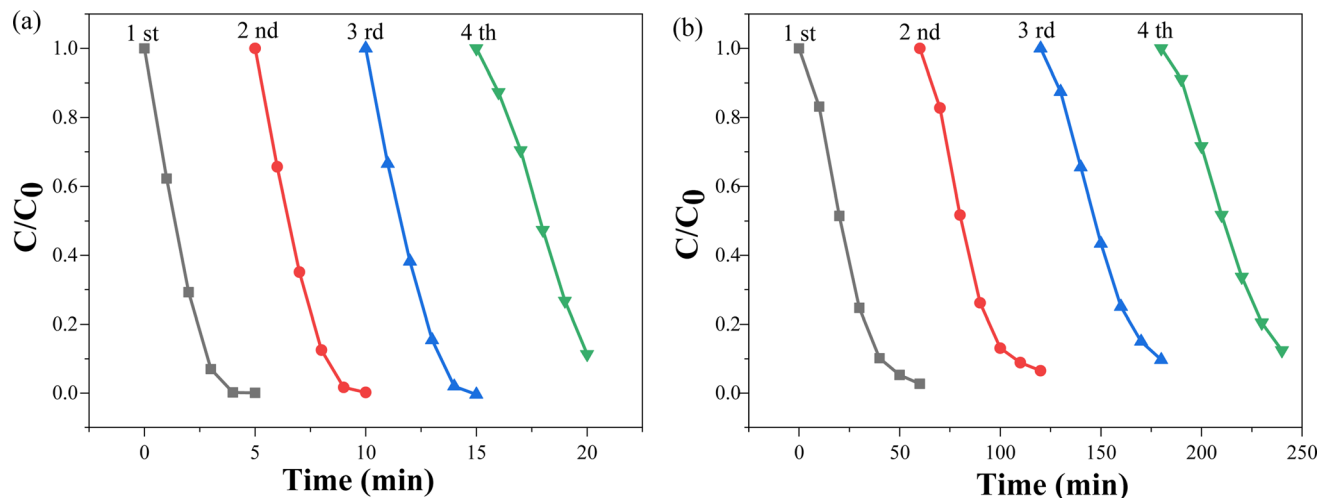


Fig. 8 Cyclic kinetics curve of photooxidation of 1,4-DHP (a) and photodegradation of AM (b) by ZA-50%.

decreased rapidly, accompanied by the emergence of a new peak at 280 nm⁴⁶ in the presence of ZA-50% composite. This indicates the rapid removal of two hydrogen atoms from the six-membered ring of 1,4-DHP, effectively converting it to its pyridine derivatives⁴⁵ (Fig. 6a, inset). The absorption peak of 1,4-DHP at 374 nm also decreased in the presence of AgI (Fig. 6b), but at a much slower rate compared with ZA-50%. However, there was almost no change observed in the UV-Vis spectra of 1,4-DHP in the presence of ZnFe₂O₄ (data not shown).

To investigate the impact of AgI proportion on the photocatalytic activity of the composites, photocatalytic dehydrogenation of 1,4-DHP was conducted at various AgI ratios (Fig. 6c). It was observed that the conversion of 1,4-DHP to its pyridine derivatives was completely executed with 5 min when the AgI content ranged from 20% to 70% (Fig. 6c and d). The plots of $\ln(C/C_0)$ versus reaction time were analyzed in order to determine the pseudo-first-order rate constants of the dehydrogenation reaction (Fig. 6e). The rate constants calculated from

Fig. 6e were compared in Fig. 6f. All the composites exhibited an increased rate constants compared to AgI and ZnFe₂O₄. The ZA-50% sample displayed the highest rate constant (1.07 min⁻¹), which was approximately 313 times higher than that of ZnFe₂O₄ (0.00341 min⁻¹) and 6 times higher than that of AgI (0.175 min⁻¹), respectively (Fig. 6f).

To further investigate the photocatalytic capability of the composite catalyst, an organic dye, AM, was selected for photodegradation. It was observed ZnFe₂O₄ exhibited a strong adsorption capacity and but almost no catalytic activity (Fig. 7a). On the other hand, AgI showed minimal adsorption capacity and weaker catalytic activity (Fig. 7b) compared to ZA-50% (Fig. 7c). The adsorption capacity of ZA-50% fell between that of ZnFe₂O₄ and AgI, while demonstrating the best photocatalytic ability by degrading 98.4% of AM in just 60 min (Fig. 7c and d). Furthermore, degradation experiments for MO and indole were carried out as detailed in the ESI (Fig. S1 and S2).†

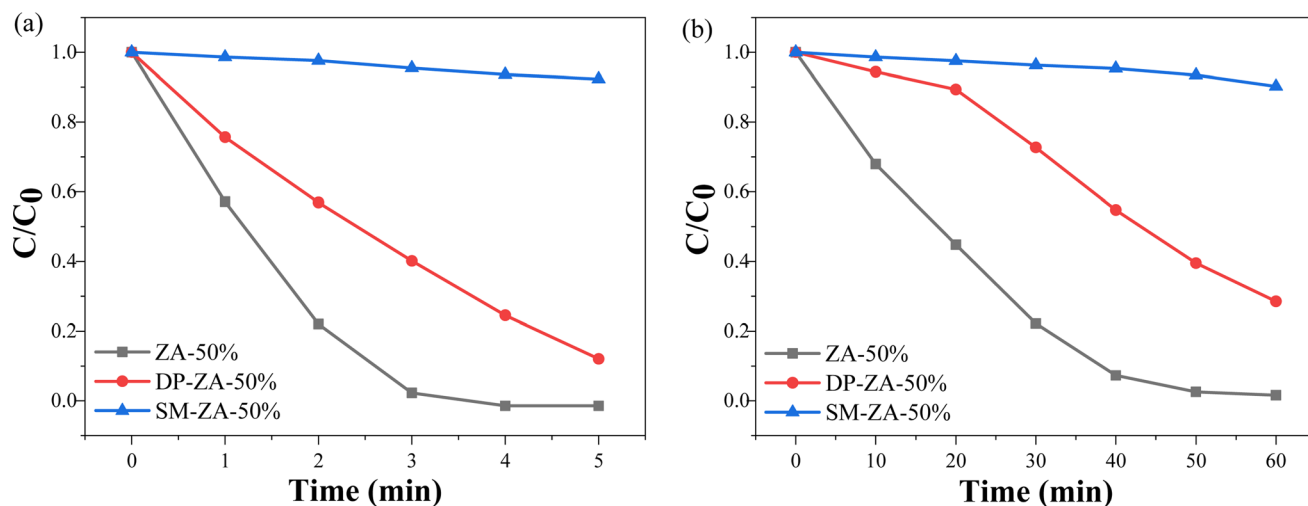


Fig. 9 Kinetics of photo-oxidative dehydrogenation of 1,4-DHP (a), photocatalytic degradation of AM (b) by ZA-50%, DP-ZA-50%, SM-ZA-50% under visible light.



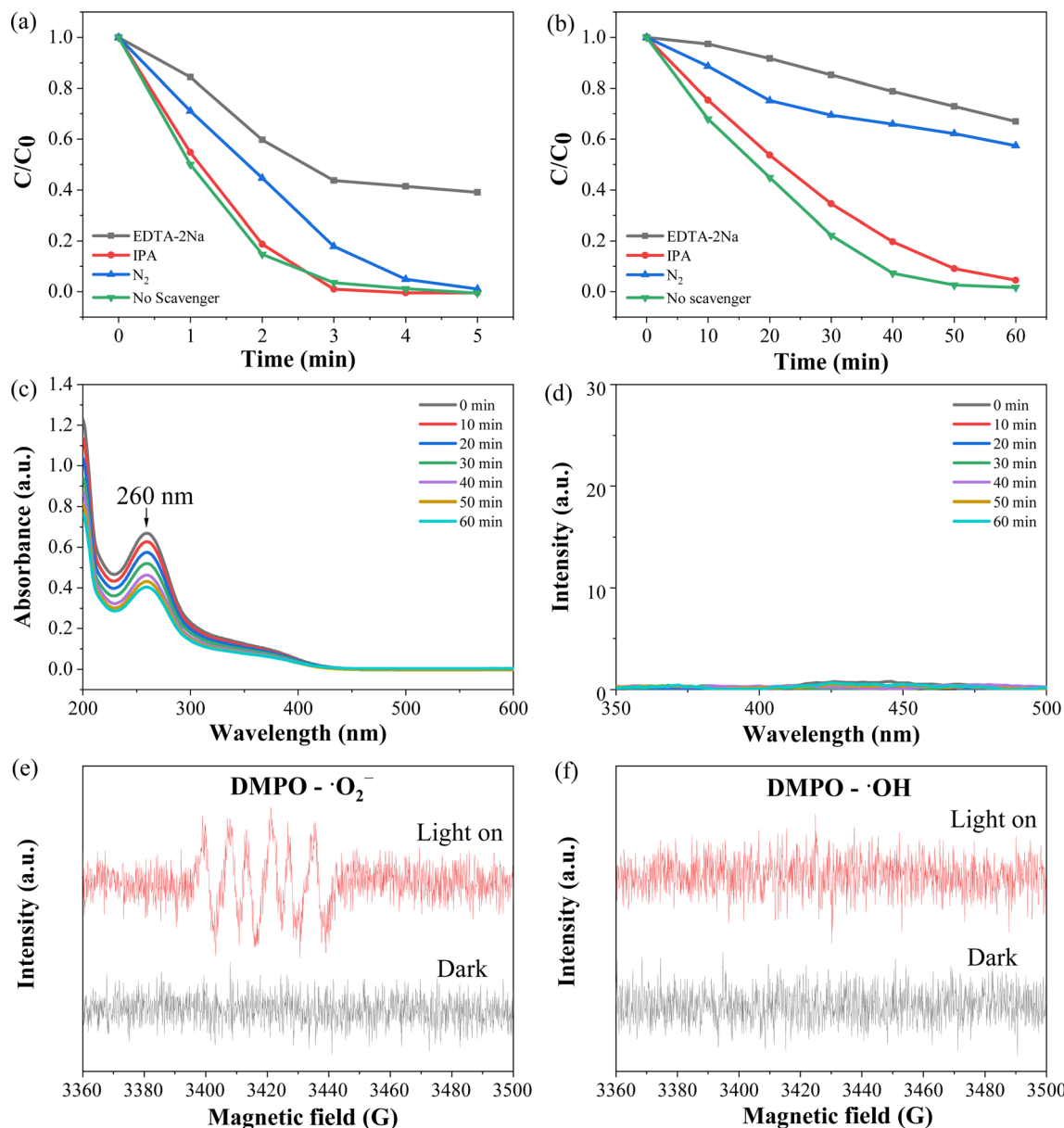


Fig. 10 The effect of certain selected scavengers on the photooxidation of 1,4-DHP and photodegradation of AM by ZA-50% under visible light ($\lambda \geq 400$ nm): kinetic diagrams of 1,4-DHP oxidation (a) and AM degradation (b); UV-Vis absorption spectrum evolution of nitroblue tetrazolium solution (c); fluorescence spectra of TA solution (d); EPR signals of DMPO- $\cdot O_2^-$ adducts (e), and DMPO- $\cdot OH$ adducts (f) in the presence of ZA-50%.

In a word, it was clearly that the $ZnFe_2O_4/AgI$ composite photocatalyst displayed better catalytic activity than that of $ZnFe_2O_4$ and AgI.

In order to assess the stability of the composite, the catalysts on the filter membrane and in the reactor were recovered and reused for four cycles of catalytic process. Despite a slight reduction in photocatalytic efficiency, the 1,4-DHP dehydrogenation rate remained high at 88.7% (Fig. 8a) and the AM degradation rate was 87.6% (Fig. 8b) after four cycles.

To demonstrate the superiority of ball milling method, ZA-50% was compared with DP-ZA-50% and SM-ZA-50%. In both dehydrogenation of 1,4-DHP (Fig. 9a) and degradation of AM

(Fig. 9b), ZA-50% showed the highest photocatalytic activity, surpassing that of DP-ZA-50%, while SM-ZA-50% exhibited the weakest photocatalytic ability.

Photocatalytic mechanism

It is widely accepted that the main active species in photocatalysis are photogenerated holes (h^+), hydroxyl radicals ($\cdot OH$) and superoxide radicals ($\cdot O_2^-$), and these primary active species can be detected through trapping experiments using specific scavengers.^{14,23} As shown in Fig. 10a and b, the addition of EDTA-2Na as a hole scavenger⁶⁵ significantly inhibited the photocatalytic reaction, indicating that h^+ play a crucial role in



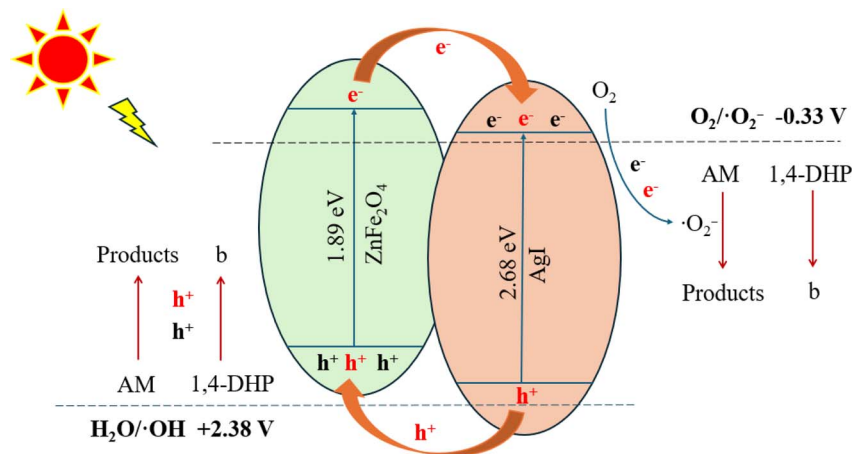


Fig. 11 Hypothetical mechanism diagram of photooxidation of 1,4-DHP and photodegradation of AM by ZnFe₂O₄/AgI composite under visible light ($\lambda \geq 400$ nm).

photocatalysis. Additionally, the removal of dissolved O₂ through N₂ bubbling resulted in a decrease in photocatalytic efficiency, suggesting that $\cdot\text{O}_2^-$ is also involved in the photocatalytic process since it is generated by the reduction of O₂ with photogenerated electrons ($\text{O}_2 + e^- \rightarrow \cdot\text{O}_2^-$).^{44,66,67} However, the introduction of isopropyl alcohol (IPA) as a $\cdot\text{OH}$ scavenger exhibited minimal inhibitory effects, implying that $\cdot\text{OH}$ may not participate in the photocatalytic process.⁴⁴

To further validate the generation of active species during the photocatalysis, NBT was employed to illustrate the generation of $\cdot\text{O}_2^-$ radicals.⁴³ The decreased peak intensity at 260 nm indicated the production of $\cdot\text{O}_2^-$ (Fig. 10c). Moreover, a fluorescence method^{68,69} was utilized to determine the generation of $\cdot\text{OH}$ radicals. Nanosized TiO₂ (Degussa P25) was also employed as a reference, as detailed in the ESI (Fig. S3).[†] As shown in Fig. 10d, no distinct peak at 425 nm was observed, indicating that $\cdot\text{OH}$ radicals were not generated during the photocatalytic process in the presence of ZA-50% composite. Additionally, electron paramagnetic resonance (EPR) measurements revealed the production of $\cdot\text{O}_2^-$ under light irradiation in the presence of ZA-50% composite sample (Fig. 10e). However, it was observed that $\cdot\text{OH}$ was not generated (Fig. 10f).^{32,70}

Based on the above observations, a reasonable mechanism is proposed in Fig. 11. Under visible light, ZnFe₂O₄ and AgI are excited, generating e^- in the CB and h^+ in the VB. Electrons in the CB of ZnFe₂O₄ are transferred to the CB of AgI, which has a more negative potential than the standard redox potential of O₂/ $\cdot\text{O}_2^-$ (-0.33 V vs. NHE).⁴³ This leads to the reduction of O₂ to $\cdot\text{O}_2^-$, resulting in the degradation of AM and dehydrogenation of 1,4-DHP to pyridine derivatives. Similarly, the h^+ in the VB of AgI are transferred to the VB of ZnFe₂O₄ for degradation of AM and dehydrogenation of 1,4-DHP. Both of ZnFe₂O₄ and AgI have a more negative VB potential than H₂O/ $\cdot\text{OH}$ (+2.38 V vs. NHE),⁴⁵ so no $\cdot\text{OH}$ radicals can be generated, consistent with previous trapping experiments and the EPR results. The enhanced photocatalytic activity of ZnFe₂O₄/AgI composite is attributed to the efficient charge separation enabled by the heterojunction formation.

Conclusion

ZnFe₂O₄/AgI composites were prepared using a ball-milling method, exhibiting enhanced photocatalytic activity under visible light compared to ZnFe₂O₄ and AgI. Among these composites, the ZnFe₂O₄/AgI-50% sample exhibited the most effective catalytic activity, completely oxidizing 1,4-DHP in just 5 minutes. The degradation efficiency of AM reached 98.4% within 1 hour, MO reached 89.3% within 3 hours, and indole reached 70.9% within 1 hour. The pseudo-first-order rate constants of ZnFe₂O₄/AgI-50% were found to be significantly higher than that of AgI for the photooxidation of 1,4-DHP (6 times), AM (20 times), MO (64 times), and indole (38 times). Mechanism experiments confirmed that h^+ and $\cdot\text{O}_2^-$ played a primary role in the catalytic process. The enhanced photocatalytic activity of ZnFe₂O₄/AgI composite is attributed to efficient charge separation and transfer enabled by heterojunction formation. This composite shows promise for use in other fields due to its efficient photocatalytic ability.

Data availability

All relevant data are within the manuscript and its additional files.

Conflicts of interest

There are no conflicts of interest to declare.

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