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Effect of surface modification on silica supported Ti catalysts for cyclohexene oxidation with vapor-phase hydrogen peroxide†

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Surface modification *via* grafting of organic moieties on a Lewis acid catalyst (silica supported Ti catalyst, Ti-SiO₂) alters the activation of H₂O₂ in vapor-phase cyclohexene epoxidation. Grafting a fluororous group (1*H*,1*H*-perfluoro-octyl) suppresses activity of Ti-SiO₂. Conversely, grafting either a nonpolar group (octyl) or a polar aprotic group (triethylene glycol monomethyl ether) enhances rates and shifts selectivity toward *trans*-1,2-cyclohexanediol.

Post-synthetic modification provides an opportunity to tune the surface properties (*e.g.* hydrophilicity/hydrophobicity) of previously-synthesized, supported catalysts. For example, surface modification to remove surface hydroxyls and to increase hydrophobicity of a surface is an effective way to reduce the negative effects of water in liquid-phase selective oxidation chemistry.^{1,2} This increase in hydrophobicity is particularly useful when water adsorbs onto the active metal site competing with a reactant. Beyond altering surface hydrophilicity/hydrophobicity, other post-synthetic modifications include overcoating of metal oxide layers,^{3,4} grafting of functional organic groups,^{5,6} and depositing additional active metal oxide sites.^{7,8} For liquid phase reactions, any surface modifications must compete against the solvent for any changes to the local environment around the active site. Here, we report grafting of organic molecules to change surface properties for vapor phase cyclohexene epoxidation with vaporized H₂O₂, where we hypothesized that surface modification might have a more direct impact on the elementary steps of catalysis. By condensing the corresponding terminal alcohols with surface silanols, we grafted (Scheme 1) three different types of functional groups on a pre-synthesized Ti-SiO₂ Lewis acid catalyst: octyl groups (Ti-SiO₂-o, nonpolar), triethylene glycol monomethyl ether (Ti-SiO₂-tg, polar aprotic), and 1*H*,1*H*-perfluoro-1-octyl (Ti-SiO₂-F, fluororous).

We prepared a highly dispersed silica supported Ti catalyst (Ti-SiO₂) *via* liquid-phase grafting of trichloro(pentamethylcyclopentadienyl)titanium(IV) onto

a mesoporous silica support at 0.2 Ti atoms per nm², followed by calcination, which is known to give high specific activity in H₂O₂ activation.⁹ We modified the parent Ti-SiO₂ by grafting the corresponding terminal alcohol in refluxing toluene for 24 h, Soxhlet extraction for 24 h in toluene to remove any ungrafted species,² and drying at 100 °C under vacuum.

Successful grafting is indicated by slight decreases in BET surface area (Fig. 1(a) and Table 1) without changes to the shape of physisorption isotherm and by mass losses in thermogravimetric analysis (Fig. 1(b)). Mass losses beyond the low-temperature desorption of water are due to combustion or decomposition of the grafted species. The water desorption temperatures of modified catalysts are similar, which agree with the previous study.² Mass losses beyond the shaded regime in Fig. 1(b) correspond to loadings of 0.75, 0.42, and 0.30 groups per nm², for Ti-SiO₂-o, Ti-SiO₂-tg, and Ti-SiO₂-F, respectively. Grafting 1-octanol on a Ta-SiO₂ catalyst was previously reported to give 0.39–0.64 groups per nm².² Most importantly, all these values are higher than the surface Ti loading of parent supported catalyst, which is 0.2 Ti atoms per nm², so that they should be sufficient to affect catalytic behavior of Ti-SiO₂.

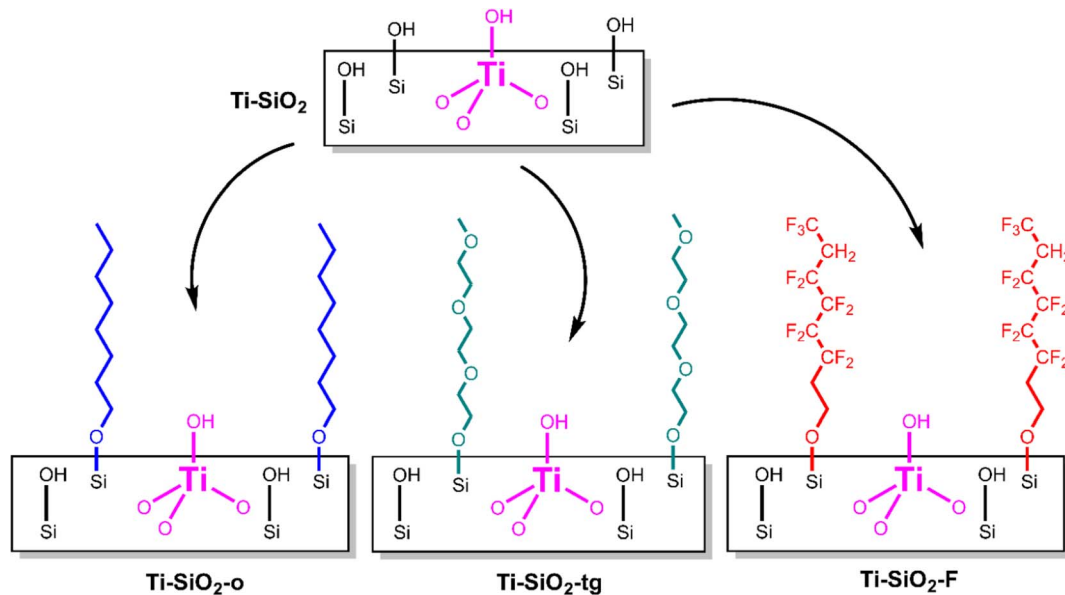
We performed vapor-phase cyclohexene epoxidation at 120 °C, 3 kPa of cyclohexene, and 3 kPa of vaporized H₂O₂ employing our custom built reactor.¹⁰ Here, we used H₂O₂ in acetonitrile, dried over MgSO₄, to minimize initial water content.¹¹ In this study, products were detected with online GC-FID and an in-jet methanizer. We do not observe any C₆ derived products other than cyclohexene epoxide (epoxide) and *trans*-1,2-cyclohexanediol (diol), consistent with our previous work with Ti-SiO₂ at similar conditions.¹⁰ In these systems, cyclohexene first converts to epoxide, and then hydrolyzes to the *trans*-diol (Scheme 2). This stepwise conversion of cyclohexene is consistent with our previous studies,^{7,10} as we do not observe any *cis*-diol that is the product of direct *cis*-dihydroxylation of cyclohexene. Radical oxidation to cyclohexenone or cyclohexenol is

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Scheme 1 Schematic diagram of grafting of 1-octanol (nonpolar), triethylene glycol monomethyl ether (polar aprotic), and 1*H*,1*H*-perfluoro-1-octanol (polar protic) on Ti-SiO₂.

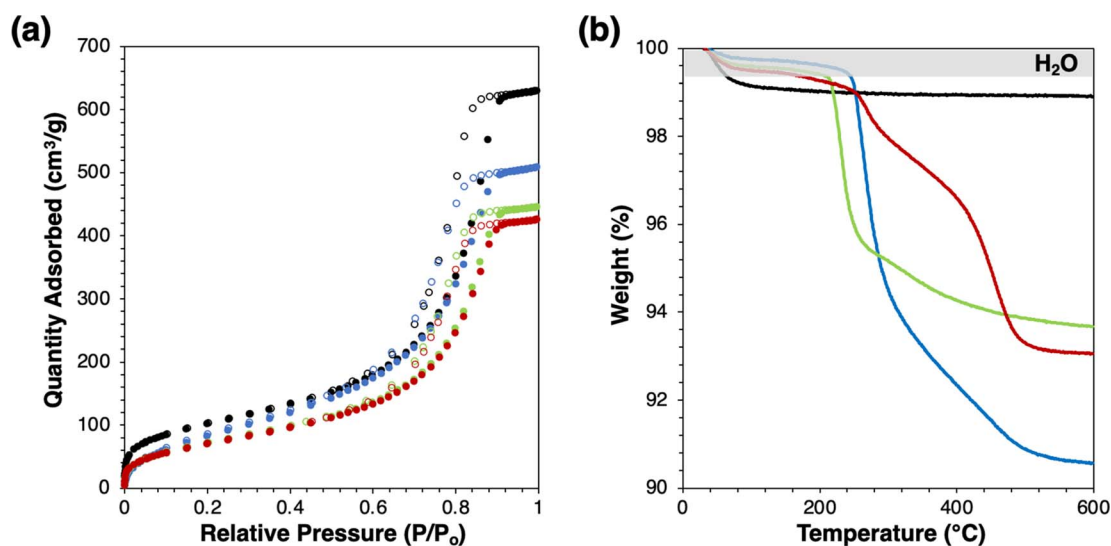


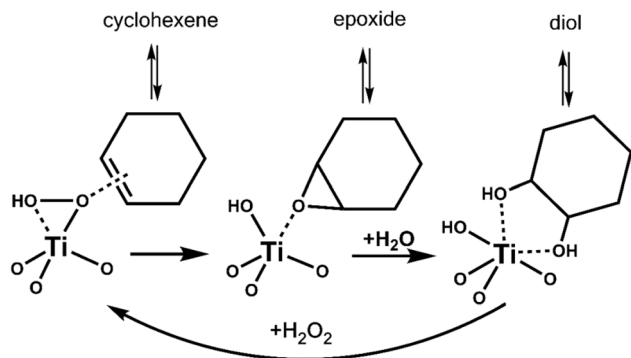
Fig. 1 (a) N₂ physisorption isotherm and (b) mass loss profile by thermogravimetric analysis of catalysts (black: Ti-SiO₂, blue: Ti-SiO₂-o, green: Ti-SiO₂-tg, red: Ti-SiO₂-F).

Table 1 Summary of catalyst properties and activities

Catalyst	BET surface area [m ² g ⁻¹]	Organic surface density [# nm ⁻²]	Steady state TOF ^a [mol _{C₆} mol _{Ti} ⁻¹ h ⁻¹]	Steady state selectivity ^a [%] (epoxide/diol/CO _x)
Ti-SiO ₂	370	—	6.5	39/42/19
Ti-SiO ₂ -o	340	0.75	15.5	6/86/8
Ti-SiO ₂ -tg	270	0.42	11.0	4/85/11
Ti-SiO ₂ -F	290	0.30	1.6	18/32/50

^a TOF values of production of epoxide and diol values at 600 min. Steady state operation is reached after 200–400 minutes at these conditions.





Scheme 2 Reaction network of cyclohexene epoxidation to cyclohexene oxide (epoxide), and its hydrolysis to yield *trans*-1,2-cyclohexane diol (diol). Surface modifications near the active site can alter the strength of adsorption of reactants and intermediates, altering product selectivity.

not observed. Background over-oxidation to CO and CO₂ occurs at a rate of approximately 1.2 to 4.1 mol_{cyclohexene} mol_{Ti}⁻¹ h⁻¹ or 0.2 to 0.8% conversion at these conditions, regardless of catalyst.

The parent Ti-SiO₂ shows an initial turnover frequency (TOF = mol_(epoxide+diol) mol_{Ti}⁻¹ h⁻¹) of 19.6 h⁻¹ at 50 min time-on-stream (TOS), which decays to a steady-state rate of 6.5 h⁻¹ at 600 min TOS. As seen most clearly in the selectivity plot, steady-state is reached after ~200 minutes, with only slow catalyst deactivation thereafter (Fig. 2). The steady state selectivity is 39%/42% to epoxide and diol respectively, with the remainder going to background overoxidation to CO_x. Grafting of a fluoros group (Ti-SiO₂-F) almost totally suppresses C₆ product formation. The small amount of remaining C₆ formation has a selectivity of 18%/32% to epoxide and diol, relatively similar to the parent catalyst and suggesting the existence of small patches of unfunctionalized surface. Otherwise, the

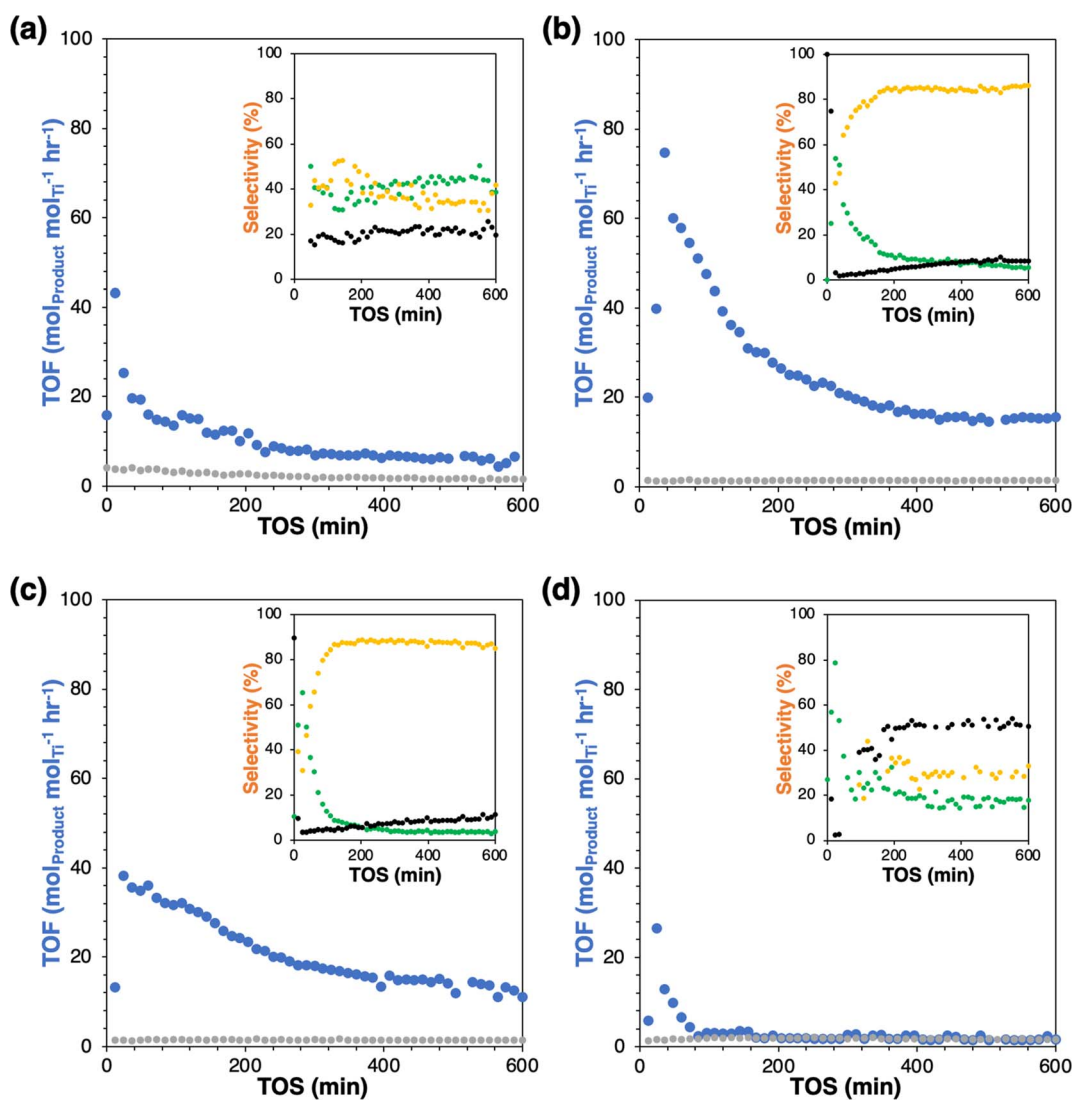


Fig. 2 Time-on-stream turnover frequency (TOF) and selectivity (inset) of (a) Ti-SiO₂, (b) Ti-SiO₂-o, (c) Ti-SiO₂-tg and (d) Ti-SiO₂-F (blue: TOF of C₆ products, grey: TOF of CO_x/6, green: epoxide selectivity, yellow: diol selectivity, black: CO_x/6 selectivity).



conclusion is that the fluorinated surface makes binding and activation of cyclohexene unfavorable by inhibiting the adsorption of cyclohexene on the surface. Conversely, grafting of either nonpolar octyl or polar aprotic tri(ethylene glycol) groups on Ti-SiO₂ increases rates by at least 1.7-fold at steady-state, relative to the parent catalyst. These enhanced catalytic rates have two effects. First, the loss of C₆ to background over-oxidation drops dramatically, from 19% in the parent catalyst to 8–11% in the modified catalysts. Moreover, the C₆ selectivity shifts to substantially favor hydrolysis of the epoxide to the diol, giving approximately 5%/85% for epoxide and diol, respectively, at steady state.

This behavior is quite different from that observed in the condensed phase, where grafting groups to remove surface silanols tends to decrease yields slightly and increase epoxide selectivity relative to diol by suppressing water sorption at the active site.² In the vapor phase and for these wide-pore materials, the lack of a liquid solvent phase means that surface modification can more directly influence the stability of reaction intermediates. As suggested in Scheme 2, the surface modifications appear to be strengthening the adsorption of cyclohexene and the intermediate epoxide, leading to corresponding increases in rate and selectivity to the hydrolysis product. In addition, a recent study by Leonhardt *et al.* proposed computationally that an epoxide molecule can remain adsorbed to one facet of the Ti-OH site while still leaving another coordination site available for oxidation of an incoming cyclohexene.¹² In that mechanism, enhancing epoxide adsorption at the active site increases hydrolysis to the diol without inhibiting – or even enhancing – overall product formation rates, such as we have observed with the octyl- and tg-modified surfaces. Overall, these observations show that surface grafting can play a significant role in modifying the reactivity of catalysts in the nascent field of selective oxidation with vaporized H₂O₂. Also, the results presented here contribute/expand to the current design strategy of post-modification of heterogeneous catalysts with simple method. Additional studies will be carried out to understand the precise mechanistic origins of these changes in rate and selectivity and develop further strategies to tune catalyst surface properties.

Data availability

The data supporting this article have been included solely in the manuscript.

Conflicts of interest

The authors declare no competing financial interest.

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References

- 1 D. A. Ruddy and T. D. Tilley, Highly selective olefin epoxidation with aqueous H₂O₂ over surface-modified TaSBA15 prepared via the TMP method, *Chem. Commun.*, 2007, (32), 3350–3352.
- 2 N. Morlanés and J. M. Notestein, Grafted Ta-calixarenes: Tunable, selective catalysts for direct olefin epoxidation with aqueous hydrogen peroxide, *J. Catal.*, 2010, 275(2), 191–201.
- 3 A. T. Y. Wolek, M. A. Ardagh, H. N. Pham, S. Alayoglu, A. K. Datye and J. M. Notestein, Creating Brønsted acidity at the SiO₂-Nb₂O₅ interface, *J. Catal.*, 2021, 394, 387–396.
- 4 A. T. Y. Wolek, K. E. Hicks and J. M. Notestein, Tuning acidity in silica-overcoated oxides for hydroalkoxylation, *J. Catal.*, 2023, 426, 113–125.
- 5 C. Khoury, S. Holton, D. Shpasser, E. Hallo, A. Kulkarni, F. C. Jentoft and O. M. Gazit, Elucidating Cooperative Interactions between Grafted Amines and Tin or Titanium Sites on Silica, *ACS Catal.*, 2022, 12(16), 9846–9856.
- 6 C. A. Contreras, R. A. Hackler, J. M. Notestein and P. C. Stair, Orientation of 1,1'-Bi-2-naphthol Grafted onto TiO₂, *J. Phys. Chem. C*, 2022, 126(18), 7980–7990.
- 7 S. Ahn, N. E. Thornburg, Z. Li, T. C. Wang, L. C. Gallington, K. W. Chapman, J. M. Notestein, J. T. Hupp and O. K. Farha, Stable Metal–Organic Framework-Supported Niobium Catalysts, *Inorg. Chem.*, 2016, 55(22), 11954–11961.
- 8 I. S. Kim, S. Ahn, N. A. Vermeulen, T. E. Webber, L. C. Gallington, K. W. Chapman, R. L. Penn, J. T. Hupp, O. K. Farha, J. M. Notestein and A. B. F. Martinson, The Synthesis Science of Targeted Vapor-Phase Metal–Organic Framework Postmodification, *J. Am. Chem. Soc.*, 2020, 142(1), 242–250.
- 9 M. A. Ardagh, D. T. Bregante, D. W. Flaherty and J. M. Notestein, Controlled Deposition of Silica on Titania-Silica to Alter the Active Site Surroundings on Epoxidation Catalysts, *ACS Catal.*, 2020, 10(21), 13008–13018.
- 10 S. Ahn, S. L. Nauert, K. E. Hicks, M. A. Ardagh, N. M. Schweitzer, O. K. Farha and J. M. Notestein, Demonstrating the Critical Role of Solvation in Supported Ti and Nb Epoxidation Catalysts via Vapor-Phase Kinetics, *ACS Catal.*, 2020, 10(4), 2817–2825.
- 11 N. E. Thornburg, A. B. Thompson and J. M. Notestein, Periodic Trends in Highly Dispersed Groups IV and V Supported Metal Oxide Catalysts for Alkene Epoxidation with H₂O₂, *ACS Catal.*, 2015, 5(9), 5077–5088.
- 12 B. E. Leonhardt, M. Head-Gordon and A. T. Bell, The Effects of ≡Ti–OH Site Distortion and Product Adsorption on the Mechanism and Kinetics of Cyclohexene Epoxidation over Ti/SiO₂, *ACS Catal.*, 2024, 14(5), 3049–3064.

