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# pH-Dependent selective extraction of gold(III) from synthetic solution and computer motherboard leachate using a hybrid nanocomposite†

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Recycling gold from electronic waste offers significant benefits for both environmental protection and resource sustainability. However, this process presents considerable challenges due to high costs, prolonged processing times, and interference from coexisting metals. In this study, we synthesized a hybrid mesoporous nanocomposite comprising platelets-like  $\text{CoNi}_2\text{S}_4$  incorporated with  $\text{g-C}_3\text{N}_4$  nanosheets ( $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$ ) for the selective recovery of gold ( $\text{Au(III)}$ ) ions from spent computer motherboards. Comprehensive characterization of the  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  nanocomposite was conducted, including its physicochemical properties, textural and structural characteristics, morphology, and elemental composition. The  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  extractor demonstrated an exceptional adsorption capacity of  $200.6 \text{ mg g}^{-1}$ , with high selectivity at pH 2, rapid equilibrium time of 60 minutes, and satisfactory reusability for over ten cycles. Adsorption isotherm and kinetic studies revealed that the  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  nanocomposite adheres to the Langmuir adsorption model and the pseudo-second-order kinetic model for  $\text{Au(III)}$  ion adsorption. Overall, this study introduces a viable adsorbent that shows considerable promise for industrial-scale  $\text{Au(III)}$  extraction from e-waste.

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## 1. Introduction

Electronic waste (e-waste) represents a rapidly expanding solid waste stream, projected to reach 61.3 million tons annually by 2023.<sup>1,2</sup> Unlike conventional solid wastes, e-waste contains a unique combination of hazardous and resource-rich materials, with only 17.4% currently being effectively recovered due to underdeveloped technologies.<sup>3–5</sup> Spent motherboards (SMBs) from obsolete computers are a significant category of e-waste, containing valuable base metals (*e.g.*, copper, lead, tin) and precious metals (*e.g.*, gold, silver, palladium).<sup>6</sup> Among these, gold (Au) is particularly valuable due to its extensive use in electronics, telecommunications, aerospace, and other high-tech industries.<sup>7–10</sup> Efficient recovery of  $\text{Au(III)}$  ions from SMBs is crucial for sustainable resource utilization and environmental protection.<sup>11,12</sup> Therefore, efficient and sustainable technologies for  $\text{Au(III)}$  recovery from SMBs must be designed.

Methods for  $\text{Au(III)}$  recovery from e-waste involve pyrometallurgy and hydrometallurgy. A pyrometallurgical method often involves directly burning e-waste, which can produce by-pollutants and harmful materials. In contrast, the

hydrometallurgical method provides a more environmentally friendly approach by utilizing digestive liquids to dissolve e-waste components before recovering the aimed species. The obtained solution undergoes subsequent extraction and purification processes.<sup>13,14</sup> Various gold extraction technologies, such as ion exchange, precipitation, flotation, solvent extraction, and adsorption, have been developed.<sup>15</sup> However, many of these approaches have noteworthy drawbacks, including high reagent requirements, the use of poisonous chemicals, and the production of toxic by-waste that necessitates discarding.<sup>16,17</sup> Among these methods, adsorption stands out as the most reliable technique for gold extraction due to its low cost, simplicity, efficiency, and reusability.<sup>18,19</sup> Recent developments have led to the creation of various adsorbents for trapping  $\text{Au(III)}$  ions, including biomass materials,<sup>20</sup> ion imprinting polymers,<sup>21,22</sup> carbon-based materials,<sup>23</sup> and metal-organic frameworks.<sup>21</sup> However, many of these adsorbents suffer from limitations such as lower adsorption capacities, slower kinetics, and reduced selectivity in the presence of competing ions. In this manuscript, we aim to design a stable adsorbent with a high adsorption capacity and selectivity feature for  $\text{Au(III)}$ -recovery from e-waste, even in the existence of other competing ions.

In recent years, graphitic carbon nitride ( $\text{g-C}_3\text{N}_4$ ) has gained significant attention due to its unique physicochemical properties, including high chemical stability, higher electronic characteristics, plentiful functional groups, efficient visible light

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absorption for photocatalytic reactions, and acceptable adsorption capacity.<sup>22–24</sup> In addition,  $g\text{-C}_3\text{N}_4$  has plenty of  $\pi$ -conjugated electrons and nitrogen-containing active sites, enabling the formation of covalent bonds with cations such as  $\text{Au(III)}$  ions. Despite these advantages,  $g\text{-C}_3\text{N}_4$ 's adsorption performance is often limited by its relatively low surface area and restricted pore structure. Various strategies have been employed to enhance  $g\text{-C}_3\text{N}_4$ 's adsorption properties, including texture modification, elemental manipulation, and incorporation with other materials.<sup>25</sup> Introducing metal sulfides, particularly binary metal sulfides such as  $\text{CoNi}_2\text{S}_4$ , offers a promising approach to enhance the adsorption performance of  $g\text{-C}_3\text{N}_4$ . Metal sulfides exhibit improved surface morphology, electrical conductivity, thermal stability, and diverse redox properties, leading to efficient  $\text{Au(III)}$  trapping.<sup>26,27</sup> Binary metal sulfide compounds exhibit superior performance compared to single metal sulfides.<sup>28,29</sup> However, to the best of our knowledge, the direct incorporation of  $\text{CoNi}_2\text{S}_4$  into the  $g\text{-C}_3\text{N}_4$  matrix to trap and extract  $\text{Au(III)}$  ions from e-waste has not been previously reported.

This study introduces a selective recovery method for  $\text{Au(III)}$  ions from e-waste-derived SMBs utilizing a mesoporous  $\text{CoNi}_2\text{S}_4@g\text{-C}_3\text{N}_4$  nanocomposite. We comprehensively evaluated the  $\text{Au(III)}$  adsorption properties of the  $\text{CoNi}_2\text{S}_4@g\text{-C}_3\text{N}_4$  nanocomposite compared to the original  $g\text{-C}_3\text{N}_4$ . The report encompasses four main aspects: (i) synthesis and detailed characterization of  $g\text{-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@g\text{-C}_3\text{N}_4$  nanocomposites, (ii) investigation of the  $\text{Au(III)}$  adsorption properties of  $g\text{-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@g\text{-C}_3\text{N}_4$  adsorbents, considering variables such as solution pH, contact time, sorbent dosage,  $\text{Au(III)}$  concentration, and the presence of interfering ions typically found in e-waste leachates, (iii) assessment of the materials for  $\text{Au(III)}$  recovery from actual SMB leachate, demonstrating their practical utility, and (iv) discussion on the  $\text{Au(III)}$  adsorption mechanisms using isotherm and kinetic models to understand the adsorbents- $\text{Au(III)}$  interaction. Our findings indicate that the mesoporous  $\text{CoNi}_2\text{S}_4@g\text{-C}_3\text{N}_4$  extractor shows considerable promise for selectively extracting gold from SMBs. This innovative approach not only enhances the efficiency of the recovery process but also contributes to the development of sustainable technologies for resource utilization and environmental protection.

## 2. Experimental work

### 2.1. $g\text{-C}_3\text{N}_4$ synthesis

$g\text{-C}_3\text{N}_4$  was synthesized through a method involving the equal proportion mixing of urea and thiourea, followed by thorough grinding. A mixture of 10 g of urea and 10 g of thiourea was thoroughly ground and placed in a covered porcelain crucible. The crucible was then calcined at 600 °C for 3 hours in a muffle furnace under a nitrogen atmosphere. The resulting yellow  $g\text{-C}_3\text{N}_4$  product was washed with deionized water and dried at 90 °C for 12 hours. The yield was approximately 35% (Fig. 1A).

### 2.2. Synthesis of $\text{CoNi}_2\text{S}_4@g\text{-C}_3\text{N}_4$

The  $\text{CoNi}_2\text{S}_4@g\text{-C}_3\text{N}_4$  composite was fabricated through a multi-step process (Fig. 1B). Initially, polyvinylpyrrolidone (PVP, M wt 55 000  $\text{g mol}^{-1}$ ) weighing 0.5 grams was dissolved in 40 mL of deionized water under continuous stirring. To this solution, 2.18 grams of  $\text{Co}(\text{C}_2\text{H}_3\text{O}_2)_2 \cdot 4\text{H}_2\text{O}$  was added, followed by a stirring period of 30 minutes. Subsequently, 1.52 grams of thiourea was incorporated into the mixture, which was then followed by the addition of 2.48 grams of  $\text{Ni}(\text{C}_2\text{H}_3\text{O}_2)_2 \cdot 4\text{H}_2\text{O}$ . The obtained solution was stirred for a further 30 minutes. Thereafter, 0.5 grams of pre-synthesized  $g\text{-C}_3\text{N}_4$  was introduced to the solution and homogenized using ultrasonic agitation for a duration of 60 minutes. The homogenized mixture was then moved to a sealed autoclave for hydrothermal treatment at a temperature of 180 °C for 24 hours. The resulting black precipitate was isolated and thoroughly cleaned with deionized water three times to eliminate impurities. The final  $\text{CoNi}_2\text{S}_4@g\text{-C}_3\text{N}_4$  product was obtained after drying at 75 °C overnight and was subsequently utilized for  $\text{Au(III)}$  adsorption experiments. The yield of  $\text{CoNi}_2\text{S}_4@g\text{-C}_3\text{N}_4$  was approximately 82% based on the total weight of  $g\text{-C}_3\text{N}_4$  and metal precursors used.

### 2.3. $\text{Au(III)}$ adsorption methodology

A series of batch trials were performed to investigate the adsorption capacity of  $\text{Au(III)}$  onto  $g\text{-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@g\text{-C}_3\text{N}_4$ . These experiments involved stirring 25 mL of a 30 ppm  $\text{Au(III)}$  solution with varying amounts of the adsorbent materials at



Fig. 1 Schematic design for  $g\text{-C}_3\text{N}_4$  (A) and  $\text{CoNi}_2\text{S}_4@g\text{-C}_3\text{N}_4$  synthesis (B).



room temperature, with the pH adjusted to a specific value. Small amounts of 0.1 M HCl or NaOH solutions were added as micro-drops to alter the pH of solutions. Following the stirring process, the samples were filtered, and the remaining Au(III) concentrations were measured using Atomic Absorption Spectroscopy. To construct adsorption isotherms, 20 mg of the adsorbents were added to 25 mL of an Au(III) solution of known concentration and stirred for 60 minutes at a pH of 2.0. Additionally, the influence of contact time on adsorption efficiency was assessed by introducing 20 mg of the adsorbent into 25 mL of a 30 ppm Au(III) solution, maintaining the temperature and pH at room conditions and 2.0, respectively, and conducting periodic analyses. The adsorption performance of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> and the Au(III)-adsorbed amount ( $q_e$ , mg g<sup>-1</sup>) were calculated by determining the difference in concentration before and after the adsorption process, using the following equations:<sup>30,31</sup>

$$\text{Ads. \%} = \frac{C_o - C_e}{C_o} \times 100$$

$$q_e = (C_o - C_e) \times V/m$$

$C_o$  and  $C_e$  (ppm) represent the original and final concentrations of Au(III), respectively. The dose of the g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents is denoted as  $m$  (g), and  $V$  (L) refers to the Au(III) solution volume. The adsorbed Au(III) ions could be released/eluted using an eluent agent, such as NaOH combined with thiourea, through a batch protocol. The elution proficiency was calculated using the next equation:<sup>32,33</sup>

$$\text{Elution \%} = C_r/C_a \times 100$$

$C_r$  and  $C_a$  (ppm) represent the concentrations of released and adsorbed Au(III) ions, respectively. The recycled g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents were reused for up to 10 adsorption–elution cycles using a batch design. The practical applicability of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents was tested by extracting Au(III) ions from the leach liquor of e-waste. Each experiment was performed in triplicate and repeated thrice to guarantee reproducibility.

### 3. Outcomes and discussion

#### 3.1. Adsorbents characterization

To investigate the morphology and microstructure of the g-C<sub>3</sub>N<sub>4</sub> sample, scanning electron microscopy (SEM) images (Fig. 2A) reveal a nanosheet-like morphology, with these sheets assembling into large blocks approximately 30 μm in size. Transmission electron microscopy (TEM) images (Fig. 2B) show g-C<sub>3</sub>N<sub>4</sub> nanosheets with widths and lengths ranging from 50 to 100 nm and ultrathin thicknesses. Darker regions in the TEM images indicate the overlap of several few-layer g-C<sub>3</sub>N<sub>4</sub> sheets, enhancing the diffusion and trapping performance of Au(III) ions. High-resolution TEM (HRTEM) images (Fig. 2C) of g-C<sub>3</sub>N<sub>4</sub> reveal uniformly distributed lattice fringe distances of 0.21 nm

corresponding to the (002) plane. Selected area electron diffraction (SAED) patterns (Fig. 2D) show two concentric rings, confirming the crystal structure and corresponding to the (100) and (002) planes of g-C<sub>3</sub>N<sub>4</sub>. The morphology of the CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> sample was examined using SEM and TEM (Fig. 2E–H). SEM images (Fig. 2E and F) display platelets-like CoNi<sub>2</sub>S<sub>4</sub> incorporated with g-C<sub>3</sub>N<sub>4</sub> nanosheets. TEM images (Fig. 2G and H) illustrate the grafting of CoNi<sub>2</sub>S<sub>4</sub> platelets among the sheet-like surfaces of g-C<sub>3</sub>N<sub>4</sub>. HRTEM analysis (Fig. 2I) indicates cubic CoNi<sub>2</sub>S<sub>4</sub> planes and g-C<sub>3</sub>N<sub>4</sub>, with inter-planar spacings of 0.3 nm, 0.25 nm, and 0.28 nm corresponding to the (311), (400), and (440) crystal planes of CoNi<sub>2</sub>S<sub>4</sub> nanoparticles, respectively. SAED patterns (Fig. 2J) further confirm the successful formation of the CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> nanocomposite by showing several diffraction rings that correspond to the lattice planes of both CoNi<sub>2</sub>S<sub>4</sub> crystal and g-C<sub>3</sub>N<sub>4</sub>. Mapping of ingredients and energy-dispersive X-ray spectroscopy (EDS) profiles of g-C<sub>3</sub>N<sub>4</sub> (Fig. 2K and S1†) demonstrate the homogenous distribution of C, N, and S elements throughout the surface. The atomic percentages of C, N, and S are 39.58%, 56.41%, and 4.01%, respectively, with the presence of sulfur attributed to the use of thiourea during synthesis. Elemental mapping images (Fig. 2L) of CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> show a homogeneous distribution of C, N, Co, Ni, and S elements, with atomic percentages of 24.68%, 38.6%, 9.45%, 6.83%, and 20.44%, respectively. In summary, these results confirm the successful formation of the CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> nanocomposite.

The X-ray diffraction (XRD) patterns depicted in Fig. 3A were utilized for phase analysis of g-C<sub>3</sub>N<sub>4</sub>, CoNi<sub>2</sub>S<sub>4</sub>, and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub>. The observed XRD peaks are consistent with the g-C<sub>3</sub>N<sub>4</sub> phase (JCPDS 87–1526). Specifically, the notable peak at 27.45° typically assigned to the (002) plane of graphitic materials, while a minor diffraction peak at 12.54° corresponds to an interplanar (100) plane.<sup>34</sup> Distinct peaks attributed to CoNi<sub>2</sub>S<sub>4</sub> are observed at 16.2°, 26.36°, 31.5°, 37.9°, 49.8°, and 54.67°, which align well with the (111), (220), (311), (400), (511), and (440) facets of the cubic phase CoNi<sub>2</sub>S<sub>4</sub> structure, as per standard pattern JCPDS no. 24-0334.<sup>26,35</sup> Furthermore, the XRD pattern of the CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> composite exhibits integrated peaks from both g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>, confirming the successful creation of the CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> nanocomposite.

X-ray photoelectron spectroscopy (XPS) was utilized to investigate the chemical bonding, chemical composition, and valence states of elemental components. The XPS survey spectra indicated the presence of C, N, O, and trace amounts of S in the g-C<sub>3</sub>N<sub>4</sub> sample (Fig. S2A†). The origin of the oxygen and sulfur is likely from the urea and thiourea used in the synthesis process. High-resolution XPS spectra for the C 1s and N 1s regions of g-C<sub>3</sub>N<sub>4</sub> showed distinct peaks at binding energies (BE) of 284.77 eV and 288.25 eV, which correspond to sp<sup>2</sup> C–C and N–C=N bonding, respectively (Fig. S2B†). The N 1s spectrum exhibited peaks at BEs of 398.79 eV and 401.1 eV, characteristic of C–N–H and C–N=N bonds, respectively (Fig. S2C†). These spectral features are consistent with those of pure g-C<sub>3</sub>N<sub>4</sub>, thereby confirming the successful synthesis of g-C<sub>3</sub>N<sub>4</sub>.<sup>36,37</sup> For the CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> sample, the C 1s spectrum exhibits peaks at BE of 289.25 eV and 284.75 eV, while the N 1s spectrum





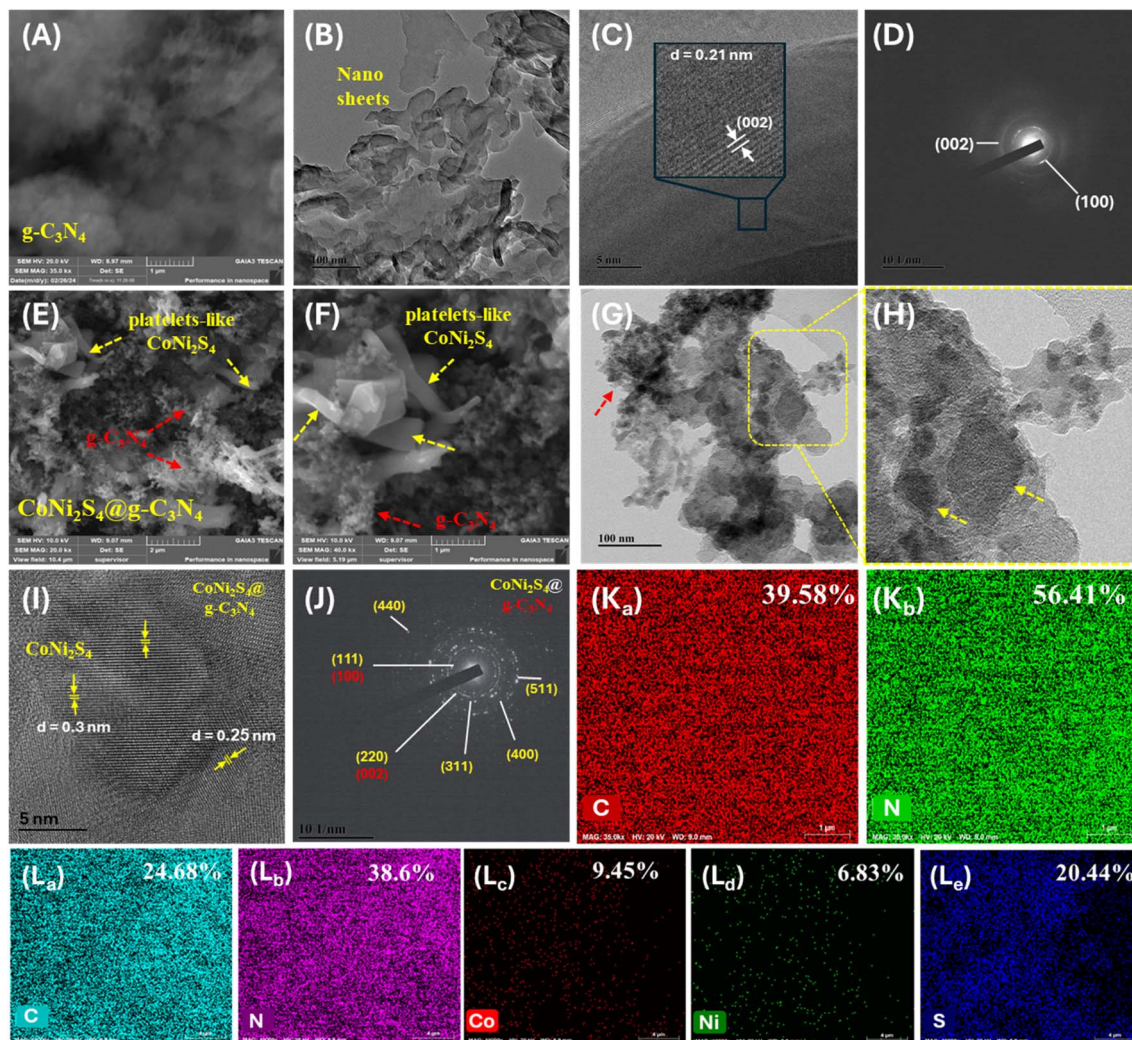


Fig. 2 SEM (A), TEM (B), HRTEM (C), and SAED images (D) of the  $g\text{-C}_3\text{N}_4$  sample; SEM (E and F), TEM (G and H), HRTEM (I), SAED images (J) of  $\text{CoNi}_2\text{S}_4@g\text{-C}_3\text{N}_4$ ; elemental mapping of  $g\text{-C}_3\text{N}_4$  ( $K_{a-b}$ ) and  $\text{FeNi}_2\text{S}_4@g\text{-C}_3\text{N}_4$  ( $L_{a-e}$ ) samples.

displays peaks at BE of 400.3 eV (Fig. 3B and C). Some shifts in BE values are observed, likely due to the grafting of  $\text{CoNi}_2\text{S}_4$  onto the  $g\text{-C}_3\text{N}_4$  surface. In the Co 2p XPS spectrum (Fig. 3D), the peaks at 778.8 and 793.8 eV for Co  $2p_{3/2}$  and Co  $2p_{1/2}$  are characteristic peaks of  $\text{Co}^{3+}$ . While the remaining peaks at 781.5 and 797.9 eV for Co  $2p_{3/2}$  and Co  $2p_{1/2}$  are attributed to the  $\text{Co}^{2+}$  spin-orbits. The corresponding satellite peaks can be observed at 803.5 eV (Co  $2p_{1/2}$ ) and 787.1 eV (Co  $2p_{3/2}$ ).<sup>38</sup> In the Ni 2p XPS spectrum (Fig. 3E), strong peaks at 856.02 and 874.06 eV for Ni  $2p_{3/2}$  and Ni  $2p_{1/2}$  are detected, signifying the presence of  $\text{Ni}^{3+}$ . The corresponding satellite heights can be observed at 862.24 eV (Ni  $2p_{3/2}$ ) and 880.23 eV (Ni  $2p_{1/2}$ ).<sup>35</sup> In the S 2p XPS spectrum (Fig. 3F), the BEs at 162 and 164.6 eV are attributed to S  $2p_{3/2}$  and S  $2p_{1/2}$  core levels, respectively. The peak at 162.3 eV is typically assigned to the  $\text{M}^{\text{II}}\text{-S}$  bond in  $\text{CoNi}_2\text{S}_4$  material. Additionally, the other peak at 164.6 eV is assigned to the sulfur ions in low coordination, which is commonly reported to S-vacancies in the material matrix. The highest detected at 168.87 eV is attributed to the shakeup satellite.<sup>39,40</sup>

$\text{N}_2$  ads-des isotherm analysis was done to evaluate pore features and surface area of the  $g\text{-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@g\text{-C}_3\text{N}_4$  adsorbents. The resulting curves displayed distinct ads/des inflexions and hysteresis loops at  $P/P_0$  ranging from  $\sim 0.4$  to  $\sim 0.9$ , indicating the porosity of the studied adsorbents (Fig. S3A†). The application of the Brunauer–Emmett–Teller (BET) model revealed a specific surface area of  $63.18 \text{ m}^2 \text{ g}^{-1}$  for  $g\text{-C}_3\text{N}_4$ , which decreased to  $8.95 \text{ m}^2 \text{ g}^{-1}$  for the  $\text{CoNi}_2\text{S}_4@g\text{-C}_3\text{N}_4$  sample due to the incorporation of  $\text{CoNi}_2\text{S}_4$  within and around the pores. This incorporation process also resulted in a reduction of pore volume from  $0.21 \text{ cm}^3 \text{ g}^{-1}$  (for  $g\text{-C}_3\text{N}_4$ ) to  $0.033 \text{ cm}^3 \text{ g}^{-1}$  (for  $\text{CoNi}_2\text{S}_4@g\text{-C}_3\text{N}_4$ ). Further analysis using the Barrett–Joyner–Halenda (BJH) method elucidated the distribution of pore sizes in the adsorbents, with narrow peaks observed at 3.02 nm and 2.89 nm, confirming the mesoporous nature of both  $g\text{-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@g\text{-C}_3\text{N}_4$  adsorbents (Fig. S3B†). These mesoporous characteristics are advantageous for efficiently trapping  $\text{Au(III)}$  ions.



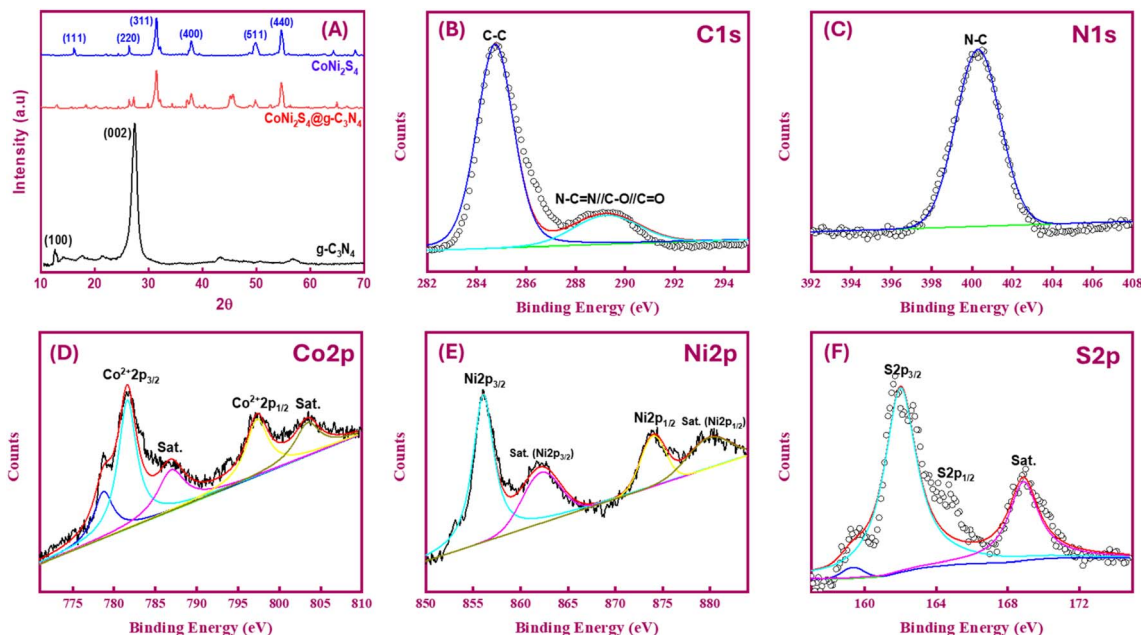


Fig. 3 (A) XRD profiles of g-C<sub>3</sub>N<sub>4</sub>, CoNi<sub>2</sub>S<sub>4</sub>, and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub>; XPS spectra of C 1s (B), N 1s (C), Co 2p (D), Ni 2p (E), and S 2p (F) of CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> sample.

### 3.2. Au(III) adsorption optimization

The pH significantly influenced the Au(III) trapping. Moreover, the selectivity of the adsorbents for Au(III) ions could be controlled by adjusting the pH value. Variations in pH affected the adsorbent active sites' charge, leading to either improved or demoted Au(III)-trapping efficiency. In the bench-top trials, 20 mg of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents were stirred with 25 mL of 30 ppm Au(III) solution for 60 minutes over a pH range of 1 to 8 at room temperature. Data in Fig. 4A revealed that 75% and 98.5% of Au(III) could be trapped at pH 2.0 using g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub>, respectively. The trapping efficiencies of both adsorbents improved faintly as the pH rose from 1 to 2. In an acidic medium (pH 2), Au(III) was present as a negative AuCl<sub>4</sub><sup>−</sup> species. The attraction forces between the active sites of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> and AuCl<sub>4</sub><sup>−</sup> species were mainly induced by the pH environment. At pH 2, the highest positive charge of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents was observed (+22 mV to +32 mV), indicating surface protonation (Fig. 4B). The high protonation (H<sup>+</sup>) at pH 2 boosted the attraction between negatively AuCl<sub>4</sub><sup>−</sup> ions and outer positively charged positions, thereby improving adsorption efficiency. As the pH increased, the surfaces of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> began to acquire a negative charge, growing the repulsion between the negatively AuCl<sub>4</sub><sup>−</sup> species and the negative-charged sites. Additionally, Au(III) may exist as [AuCl<sub>3</sub>OH]<sup>−</sup> and [AuCl<sub>2</sub>(OH)<sub>2</sub>]<sup>−</sup> species, creating an electrostatic repulsion behaviour between the anionic Au(III) species and the negative-charged centres.<sup>41,42</sup>

Fig. 4B indicated that the zero-charge points (ZCP) of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents were found to be at pH 3.5 and 5.1, respectively. At pH < pH<sub>ZCP</sub>, the active sites on the surfaces of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> become positive-charged sites

due to the creation of protonated functional groups. Conversely, at pH > pH<sub>ZCP</sub>, active sites acquire negative charges because of the creation of negatively charged groups. The zeta-potential data proved the highest Au(III)-trapping efficiency obtained at pH 2. The strong binding interactions at this pH ensured maximum adsorption and selectivity for Au(III) ions. Subsequent studies on the trapping of Au(III) ions were conducted using g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> under optimal pH conditions (*i.e.*, pH 2).

Different doses of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents (ranging from 5 to 50 mg) were employed to assess the effect of adsorbent dose on the adsorbent performance toward Au(III)-trapping. As presented in Fig. 4C, the adsorption efficiency of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents increased with the growth of adsorbents quantity. This increase suggests that the adsorption efficiencies of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> are dependent on the availability of active surface sites, and larger quantities of adsorbents provide additional internal and external active sites along the adsorbent surfaces, thereby enhancing efficiency. When the adsorbent quantity was insufficient, the number of interior and exterior surface-active sites was less than what was required for whole trapping. Consequently, 20 mg of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents were found to be enough and suitable for following adsorption experiments.

### 3.3. Adsorption isotherm study

A series of bench-top tests were conducted to evaluate the impact of initial Au(III) concentrations (ranging from 0.1 to 250 ppm) on the adsorption capacities of g-C<sub>3</sub>N<sub>4</sub>, CoNi<sub>2</sub>S<sub>4</sub>, and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents under optimal conditions. As shown in Fig. 5A, the adsorption capacities ( $q_e$ , mg g<sup>−1</sup>) of g-





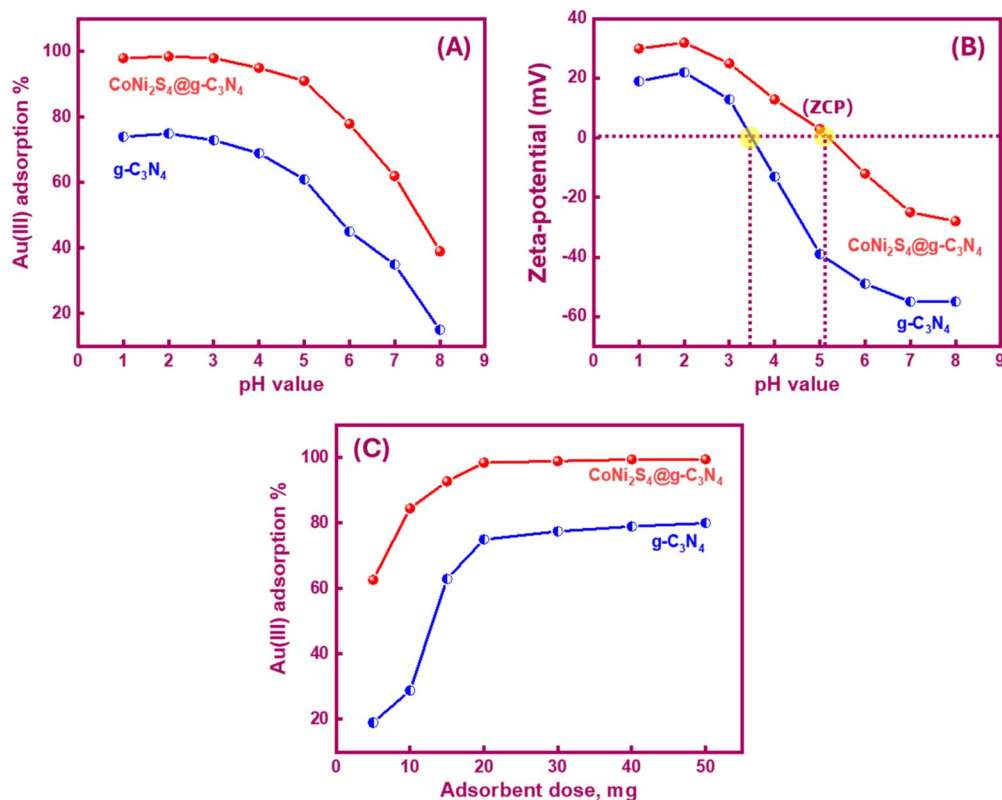


Fig. 4 (A) Impact of pH value on the trapping performance of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents; (B) zeta-potential profile of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents based on the change of pH value; (C) effect of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> dose (mg) on their adsorption performance.

C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> increased sharply at low Au(III) concentrations before reaching equilibrium at higher concentrations. This rapid increase at low concentrations is indicative of type I sorption on porous adsorbents, according to the Langmuir classification. The maximum experimental adsorption capacities were found to be 111.25 mg g<sup>-1</sup> for g-C<sub>3</sub>N<sub>4</sub> and 200.6 mg g<sup>-1</sup> for CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub>. Fig. S4† presents the Au(III) adsorption capacity of the CoNi<sub>2</sub>S<sub>4</sub> sample, which was determined to be 132.86 mg g<sup>-1</sup>. Although this is a significant capacity, it is lower than the 200.6 mg g<sup>-1</sup> achieved by the CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> nanocomposite. This comparison underscores the enhancement in adsorption performance when CoNi<sub>2</sub>S<sub>4</sub> is incorporated into the g-C<sub>3</sub>N<sub>4</sub> matrix. The g-C<sub>3</sub>N<sub>4</sub> support increases the surface area, improves the dispersion of CoNi<sub>2</sub>S<sub>4</sub> particles, and provides additional active sites for adsorption, collectively boosting the overall adsorption capacity and efficiency. The synergistic interaction between CoNi<sub>2</sub>S<sub>4</sub> and g-C<sub>3</sub>N<sub>4</sub> in the nanocomposite results in superior adsorption kinetics and higher selectivity for Au(III) ions, highlighting the crucial role of the g-C<sub>3</sub>N<sub>4</sub> support in the composite material.

Langmuir and Freundlich isotherm models were employed to discover the nature of Au(III)-adsorbents interaction. Additionally, these models were used to determine the theoretical adsorption capacity of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents, based on the following equations:<sup>43,44</sup>

$$\frac{C_e}{q_e} = \frac{1}{K_L Q_0} + \left(\frac{1}{Q_0}\right) C_e$$

$$R_L = \frac{1}{1 + K_L C_0}$$

$$\ln q_e = \ln K_f + \frac{1}{n} \ln C_e$$

$Q_0$ ,  $K_L$  and  $K_f$  are the theoretical adsorption capacities of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents (mg g<sup>-1</sup>), Langmuir constant (L mg<sup>-1</sup>), and Freundlich constant. And  $n$  represents the sorption intensity. The coefficient of determination ( $R^2$ ) for the linear relationships confirmed that the Langmuir model is more appropriate than the Freundlich model (Fig. 5B and C), indicating the creation of chemical bonds between Au(III) ions and active sites of the g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> in a single layer. The  $Q_0$ ,  $K_L$ ,  $K_f$ , and  $n$  values were determined and listed in Table 1. The theoretical  $Q_0$  values for g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents were 121.95 mg g<sup>-1</sup> and 205.34 mg g<sup>-1</sup>, respectively, which aligned well with the experimental results. Additionally,  $1/n < 1$  confirmed the chemical interaction between Au(III) ions and used adsorbents. Findings specified



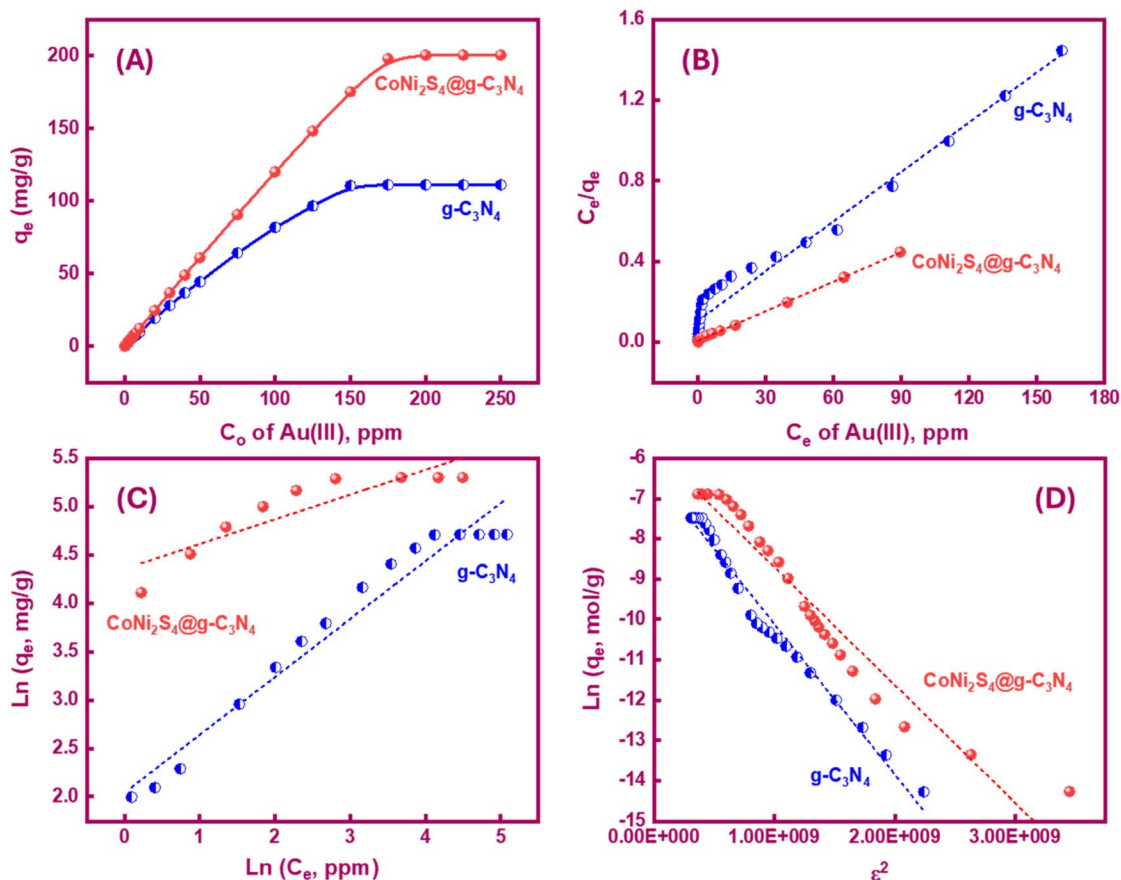


Fig. 5 (A) Effect of preliminary [Au(III)] on the  $q_e$  values of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents, and linear plots of Langmuir (B), Freundlich, and (C) Dubinin–Radushkevich (D–R) (D) isotherm models.

that the g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents are effective at both low and high Au(III) concentration levels.

Furthermore, the Dubinin–Radushkevich (DR) isotherm model characterize of the adsorption process based on the porous structure of g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents:<sup>45,46</sup>

$$\ln q_e = \ln q_o - \delta \epsilon^2$$

$$\epsilon = RT \ln \left( 1 + \frac{1}{C_e} \right)$$

where  $\epsilon$  is the Polanyi potential,  $C_e$  (mol L<sup>-1</sup>) is the equilibrium Au(III)-concentration,  $q_o$  (mol g<sup>-1</sup>) represents the maximum Au(III) adsorption capacity, and  $\delta$  (mol<sup>2</sup> J<sup>-2</sup>) is the constant related to the Au(III)-sorption energy. The Au(III)-sorption energy ( $E$ , J mol<sup>-1</sup>) offers insight into the Au(III)-sorption nature. The value of  $E$  can be concluded *via* the subsequent relationship:<sup>47</sup>

Table 1 Isotherm and kinetic parameters of Au(III)-sorption using g-C<sub>3</sub>N<sub>4</sub> and CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub>

	Langmuir model	Freundlich model			DR	PFO	PSO		IPD			
g-C <sub>3</sub> N <sub>4</sub>	Experiential $q_m$ (mg g <sup>-1</sup> )	111.25			$R^2$	0.97						
	$R^2$	0.97	$R^2$	0.95	$q_o$ mol g <sup>-1</sup>	0.0017	$R^2$	0.65	$R^2$	0.98	$R^2$	0.654
	$Q_o$	121.95	$K_F$ , mg g <sup>-1</sup>	7.69	$\delta$ mol <sup>2</sup> J <sup>-2</sup>	$3.7 \times 10^{-9}$	$q_e$ , mg g <sup>-1</sup>	100.69	$q_e$ , mg g <sup>-1</sup>	32.5	$I$	12.65
	$K_L$ , L mg <sup>-1</sup>	0.076	$n$	1.66	$E$ J mol <sup>-1</sup>	11 624.76	$K_1$ , min <sup>-1</sup>	0.0033	$K_2$ , g mg <sup>-1</sup> min <sup>-1</sup>	0.00313	$K_{id}$ mg g <sup>-1</sup> min <sup>-0.5</sup>	2.02
					$E$ kJ mol <sup>-1</sup>	11.624						
CoNi <sub>2</sub> S <sub>4</sub> @g-C <sub>3</sub> N <sub>4</sub>	Experiential $q_m$ (mg g <sup>-1</sup> )	200.6			$R^2$	0.93						
	$R^2$	0.998	$R^2$	0.77	$q_o$ mol g <sup>-1</sup>	0.0031	$R^2$	0.45	$R^2$	0.999	$R^2$	0.64
	$Q_o$	205.34	$K_F$ , mg g <sup>-1</sup>	77.47	$\delta$ mol <sup>2</sup> J <sup>-2</sup>	$2.92 \times 10^{-9}$	$q_e$ , mg g <sup>-1</sup>	181.97	$q_e$ , mg g <sup>-1</sup>	38.46	$I$	30.6
	$K_L$ , L mg <sup>-1</sup>	0.611	$n$	3.85	$E$ J mol <sup>-1</sup>	13 085.6	$K_1$ , min <sup>-1</sup>	0.0019	$K_2$ , g mg <sup>-1</sup> min <sup>-1</sup>	0.0124	$K_{id}$ mg g <sup>-1</sup> min <sup>-0.5</sup>	0.843
					$E$ kJ mol <sup>-1</sup>	13.085						

$$E = \frac{1}{\sqrt{2\delta}}$$

$$\log(q_e - q_t) = \log q_e - \left(\frac{K_1}{2.303}\right)t$$

$$\frac{t}{q_t} = \frac{1}{K_2 q_e^2} + \left(\frac{1}{q_e}\right)t$$

If the  $E$  value is  $<8 \text{ kJ mol}^{-1}$ , the sorbate-sorbent interaction is classified as physisorption, while  $E$  value ranging from 8 to  $16 \text{ kJ mol}^{-1}$  signifies a chemisorption process. In the current study, the  $E$  values for  $\text{g-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  sorbents were estimated as 11.624 and  $13.085 \text{ kJ mol}^{-1}$ , confirming the chemical nature of the  $\text{Au(III)}$  adsorption process (Fig. 5D and Table 1).

### 3.4. Kinetic study of $\text{Au(III)}$ adsorption

Determining the required time for the  $\text{Au(III)}$ -trapping process is economically important to achieve maximum adsorption efficiency within the lowest consumed time as much as possible. So, numerous stirring times (2.5–90 minutes) were tested to explore the impact of stirring time on  $\text{Au(III)}$ -sorption efficiency. Outcomes in Fig. 6A specified that the  $\text{Au(III)}$ -trapping efficiency encouraged with longer stirring times, reaching 75% and 98.5% for  $\text{g-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$ , respectively, within 60 minutes. Thus, equilibrium could be achieved within 60 minutes. Pseudo-first/second-order (PFO & PSO) kinetic models were employed to investigate the sorption interaction mechanism. The PFO and PSO kinetic equations are described as follows:<sup>48</sup>

$K_1 \text{ (min}^{-1}\text{)}$  and  $K_2 \text{ (g mg}^{-1} \text{ min}^{-1}\text{)}$  are the rate constants of PFO and PSO models.  $q_e$  and  $q_t$  represent the quantity of trapped  $\text{Au(III)}$  ions ( $\text{mg g}^{-1}$ ) at equilibrium and at time  $t$ . The constants can be determined from the slope and intercept of the linear plots of  $\log(q_e - q_t)$  and  $t/q_t$  against  $t$  (Fig. 6B and C). The results indicated that the PSO model was more appropriate for illustrating the kinetic pathway of  $\text{Au(III)}$  adsorption onto  $\text{g-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$ , suggesting a chemisorption mechanism. The  $K_2$  value for the  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  adsorbent (0.0124) was greater than that of the  $\text{g-C}_3\text{N}_4$  adsorbent (0.00313) (Table 1), indicating that  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  has a significantly higher adsorption rate for  $\text{Au(III)}$  ions. This implies that the composite material  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  enhances the adsorption kinetics, likely due to improved surface properties, increased active sites, or better interaction between sorbents and  $\text{Au(III)}$  ions.

The  $\text{Au(III)}$ -sorption process typically progresses through multiple stages, including adsorbate transfer, external prevalence, intra-particle diffusion (IPD), and eventually chemical or

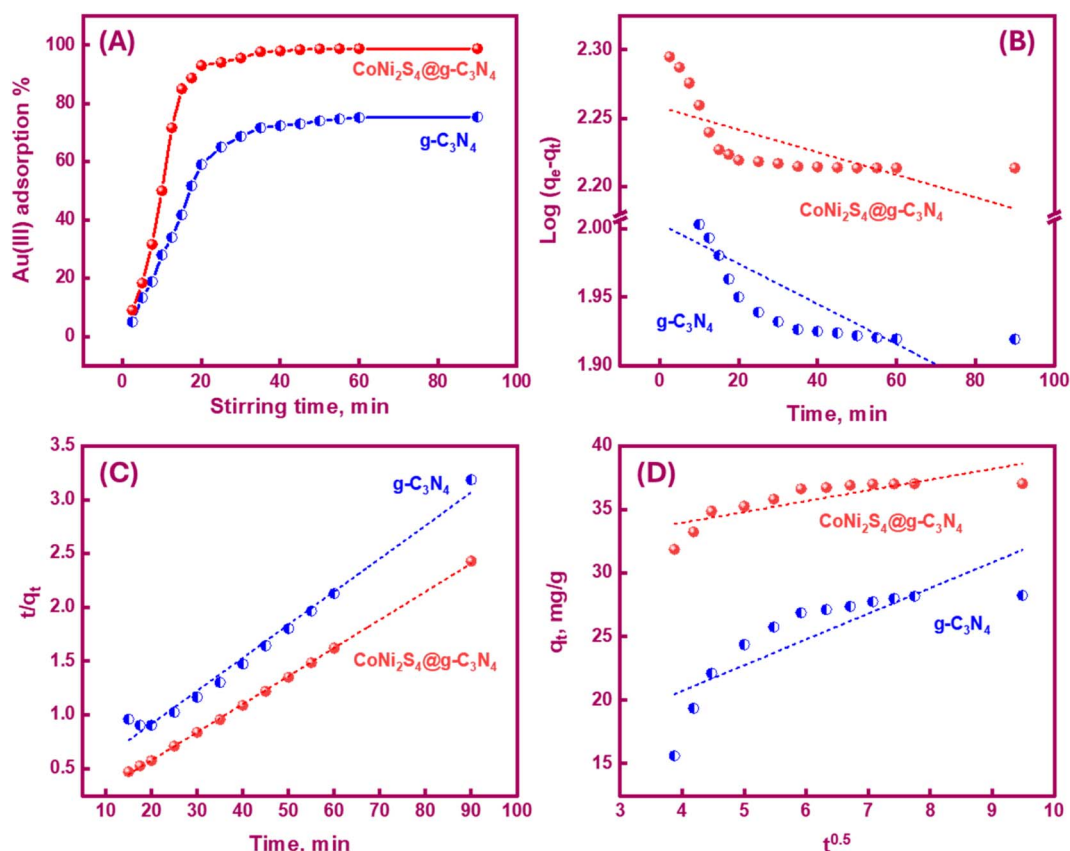


Fig. 6 Effect of stirring time on the trapping efficiency of  $\text{Au(III)}$  adsorption using  $\text{g-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  adsorbents, and linear plots of PFO (B), PSO (C), and intra-particle diffusion (D) kinetic models under the best trapping conditions.





physical interaction between the sorbents' active sites and Au(III) ions. The IPD model assumes that the prevalence of Au(III) is the rate-controlling step during the adsorption procedure, and the diffusion direction can vary. This kinetic model was used to evaluate the IPD rate constant. The IPD equation can be represented as next:<sup>49</sup>

$$q_t = K_{id}t^{0.5} + I$$

$K_{id}$  ( $\text{mg g}^{-1} \text{min}^{-0.5}$ ) is the rate constant of the IPD. The intercept  $I$  of the IPD equation depends on the thickness of the external boundary layer. The plot of  $q_t$  versus  $t^{0.5}$  for  $\text{g-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  adsorbents is shown in Fig. 6D. Table 1 shows the calculated values of  $R^2$ ,  $K_{id}$ , and  $I$ . For  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$ , the  $K_{id}$  value is lower than that for  $\text{g-C}_3\text{N}_4$ , suggesting that the diffusion of Au(III) along  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  matrix is slower than  $\text{g-C}_3\text{N}_4$ . This may be due to the dense grafting of  $\text{CoNi}_2\text{S}_4$  along the  $\text{g-C}_3\text{N}_4$  surface and pores, facilitating quick access to active sites. A higher  $I$  value enhances the influence of the outer surface prevalence, improving the rate of the Au(III)-sorption process. Rapid sorption of Au(III) ions using  $\text{g-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  adsorbents occurred within the first 60

minutes, primarily due to the prevalence of the external layer or mesopores.

### 3.5. Selectivity study of used adsorbents

Selectivity is a critical factor in metal extraction processes, as it aims to isolate specific ions in their pure states from complex mixtures containing various ions. Here, we investigated the impact of competing ions on the Au(III)-trapping efficiency in both single and combined systems. At the optimal pH of 2, 20 mg of  $\text{g-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  were stirred with 25 mL of a 30 ppm Au(III) solution, along with 30 ppm of Na(I), K(I), Ag(I), Mg(II), Ca(II), Pd(II), Mn(II), Fe(III), Cu(II), Ni(II), and Al(III). As shown in Fig. 7A, the adsorbents demonstrated a high selectivity for Au(III), with minimal adsorption of other cations. The solution's pH was a significant determinant of this selectivity. In acidic conditions, the high concentration of  $\text{H}^+$  ions hindered the adsorption of other cations. Furthermore, when Au(III) was presented alongside another metal ion in binary systems and exposed to the adsorbents at pH 2, findings indicated that competing cations had a negligible impact on Au(III) ions sorption (Fig. 7B).

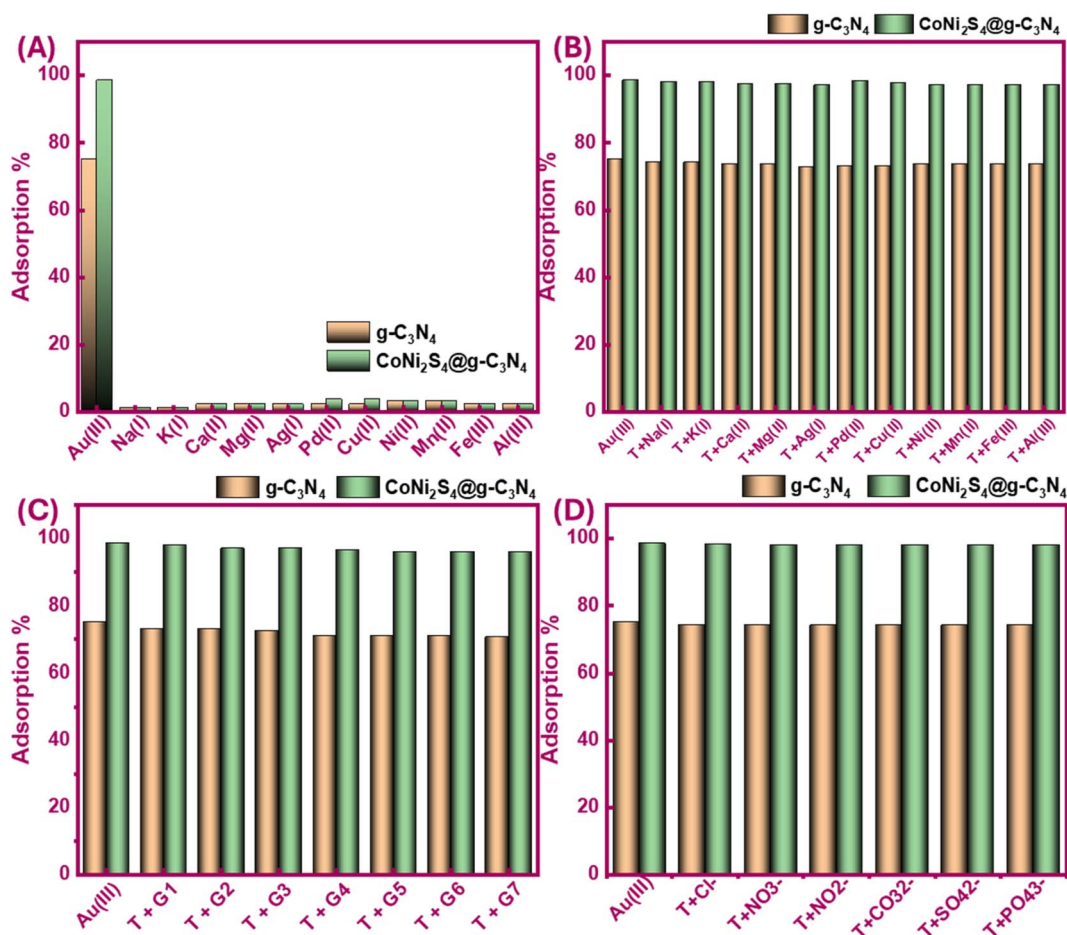


Fig. 7 Selectivity profile of  $\text{g-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  to adsorb Au(III) ions and other co-existing ions in single (A), binary (B), and group (more than one competitive ion) systems (C), and anions (D) systems at optimum trapping conditions. G1{Au(III) + Ag(I)}, G2{Au(III) + Cu(II)}, G3{Au(III) + Mn(II) + Al(III)}, G4{Au(III) + Co(II) + Hg(II) + Pb(II)}, G5{Au(III) + Ni(II) + Ca(II) + Mn(II) + Fe(III)}, G6{Au(III) + Al(III) + Mg(II) + Pb(II) + Hg(II) + Cr(III)}, G7{Au(III) + Hg(II) + Cd(II) + K(I) + Pd(II) + Fe(III)}.

Further batch experiments were conducted in a mixture system using more than one cation with Au(III) (G1 to G7), as explained in Fig. 7C caption. The findings in Fig. 7C demonstrated no noteworthy variation in Au(III) adsorption, maintaining a high adsorption efficiency of 96–98%, despite the presence of increased concentrations of diverse cations. Overall, both  $g\text{-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  exhibited strong anti-cation interference capabilities, effectively capturing Au(III) with high selectivity. We also evaluated the impact of co-existing anions on the adsorption behavior of Au(III). Fig. 7D shows that the Au(III)-trapping efficiency of  $g\text{-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  decreased slightly by 1–2% in the company of  $\text{Cl}^-$ ,  $\text{NO}_3^-$ ,  $\text{NO}_2^-$ ,  $\text{SO}_4^{2-}$ ,  $\text{CO}_3^{2-}$ , and  $\text{PO}_4^{3-}$  ions. This minor reduction in adsorption efficiency suggests a relatively small adverse impact of these anions on Au(III)-sorption performance. Despite this minor interference, the robust binding between negatively Au(III) species and positive-charged sites of  $g\text{-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  at pH 2 remained dominant.

### 3.6. Recycling and reusability of spent extractors

Spent  $g\text{-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  adsorbents can be recycled using a batch elution/desorption protocol, minimizing the cost of the extraction process and the generated further waste. Elution is essential for obtaining target ions in their pure forms. In bench-top trials, Au(III)-loaded adsorbents were treated with NaOH and thiourea (0.1 M) under continuous stirring to investigate the effect of elution time (Fig. 8A). Thiourea forms a stable complex with Au(III) ions  $\{[\text{Au}(\text{CS}(\text{NH}_2)_2)_2]^+\}$  in the presence of NaOH, facilitating the desorption of gold from the adsorbent. NaOH adjusts the pH and provides an alkaline environment that enhances the desorption efficiency of the thiourea complex. Results showed that over 99% of the loaded Au(III) were recovered within 60 minutes. The treated  $g\text{-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  adsorbents were then filtered, dried, and reused for subsequent Au(III) adsorption. Fig. 8B explains that the recycled  $g\text{-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  adsorbents could be reused for ten cycles. After the tenth cycle, the adsorption efficiency of  $g\text{-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  adsorbents decreased to 55% and 85%, respectively, while the elution efficacy stayed high (>99%). Despite the decrease in adsorption efficiency due to the stripping agent's adverse effects during repeated operations, which may destroy some of the surface-active sites. The findings indicate that the  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  adsorbent is effective in e-waste treatment and Au(III) ion recovery. This is attributed to its superior efficacy, cost-efficiency, reusability, and robustness even after multiple reuse cycles.

### 3.7. Au(III) extraction from SMB

Recycling e-waste has garnered global attention as a sustainable strategy for accessing precious metals like gold while mitigating environmental impact. Our study focuses on e-waste recycling using highly selective extractors through a straightforward, rapid, and efficient approach. In this investigation, we utilized mesoporous  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  adsorbent to selectively capture Au(III) ions amidst competitive ions from actual leachate. The extraction of Au(III) from SMB comprised multiple phases: (i)

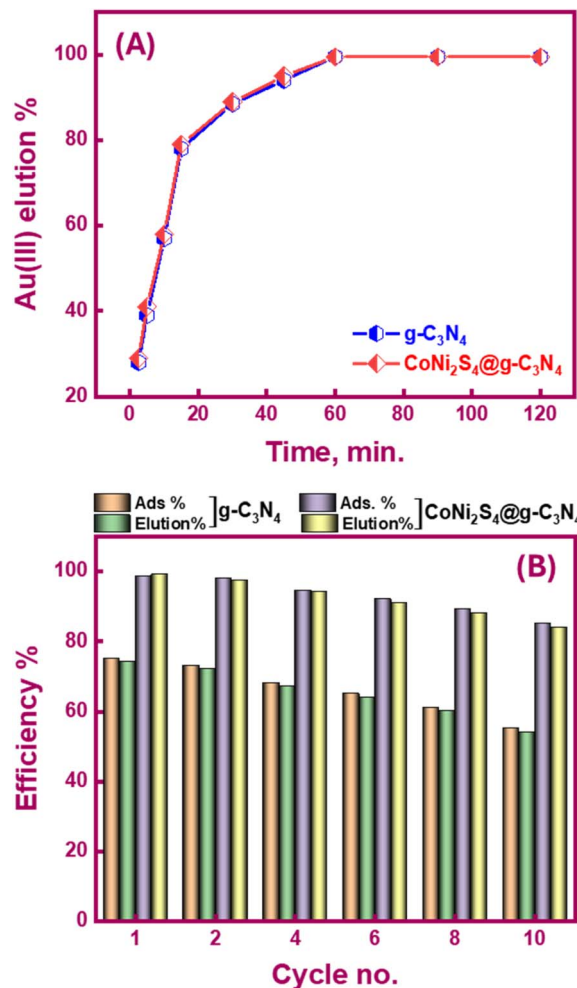


Fig. 8 (A) Influence of stirring time on the elution efficacy of Au(III) ions from used  $g\text{-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  adsorbents; (B) reusability of  $g\text{-C}_3\text{N}_4$  and  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  adsorbents under the optimum adsorption and elution circumstances.

leaching step to release Au(III) ions from SMB to solutions, (ii) adsorption step to adsorb Au(III) ions using  $\text{CoNi}_2\text{S}_4@\text{g-C}_3\text{N}_4$  adsorbent under the optimum adsorption conditions, and (iii) elution step to collect Au(III) ions in a pure form.

The leaching process for Au(III) ion releasing from SMB can be performed using a hydrometallurgical approach, which could be done through the next steps (see Fig. 9):

(i) Gathering SMBs: SMBs were isolated from obsolete computers and crushed into tiny fragments to ease the leaching process.

(ii) Leaching SMBs: Here, 100 g of crushed SMB was stirred overnight with a mixture of  $\text{H}_2\text{SO}_4$  (5 M) and  $\text{H}_2\text{O}_2$  (30%) to leach metals such as Ni(II), Fe(III), Al(III), Zn(II), and Cu(II) ions. The remaining solid components were then filtered and separated.

(iii) Dissolving gold ions: The resulting solution was stirred with  $\text{HNO}_3/\text{HCl}$  solution (1 : 3 v : v), leading to a soluble solution of Au(III) ions and other ions.

(iv) Precipitation of other ions: Fe(III), Sn(IV), and Al(III) ions were precipitated using NaOH by adjusting the pH to 4–4.5,





Fig. 9 SMB hydrometallurgical processing and extraction process scheme of the Au(III) ions from the SMBs using mesoporous CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbent.

while Ag(I) ions were separated using NaCl as AgCl. Precipitates were then eliminated through centrifuge filtration.

(v) The obtained clear solution was employed for selective trapping of Au(III) ions using the CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbent.

As shown in Table 2, Au(III) ions and other competitive ions concentrations in the resulting liquor were quantified using Atomic Absorption Spectroscopy before interaction with the CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbent. The CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbent achieved an adsorption efficiency of 97.6%, and the elution efficiency for these adsorbents was 98.5%. This study demonstrates that the developed CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbent is highly effective in extracting, recovering, and purifying Au(III) from obsolete SMB. These findings suggest the potential application of CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbents in environmental cleanup and e-waste management, offering a sustainable solution for recovering valuable Au(III) ions from e-waste sources selectively.

### 3.8. Advantages and potential limitations of the proposed method

In this section, we comprehensively compare our proposed method for Au(III) extraction using the CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> nanocomposite with other established techniques, highlighting the specific advantages and potential limitations.

(I) Leaching process: pyrometallurgy involves high-temperature processing, which can efficiently extract gold and other metals from e-waste. However, these methods often require significant energy input, generate toxic by-products, and have lower selectivity, contaminating the recovered metals. In contrast, hydrometallurgy uses aqueous solutions to leach metals, which can be more environmentally friendly than pyrometallurgical methods. Therefore, we followed the hydrometallurgy method in the leaching process.

(II) Adsorption process: the CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> nanocomposite demonstrates a high adsorption capacity of 200.6 mg g<sup>-1</sup>, rapid adsorption kinetics, and excellent reusability, maintaining over 85% efficiency after ten cycles. These features make it a superior adsorbent compared to many conventional materials, as shown in Table 3. Table 3 compares the adsorption capacities of various adsorbents for extracting Au(III) ions from aqueous solutions. The results demonstrate that CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> excels in Au(III) adsorption, surpassing the performance of several other materials. These findings suggest that the CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> nanocomposite holds considerable promise for selectively capturing Au(III) ions from e-waste.

(III) Specific advantages of our proposed method: (i) the CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> nanocomposite shows exceptional selectivity for Au(III) ions, even in the presence of competing metal ions, and a high adsorption capacity, outperforming many traditional adsorbents; (ii) the process operates under mild conditions, reducing energy requirements and minimizing the use of hazardous chemicals. The reusability of the adsorbent further reduces operational costs and environmental impact; (iii) our method has been tested with actual e-waste leachates, demonstrating its effectiveness in real-world scenarios. The results indicate that the method can be easily scaled up for industrial applications, providing a viable solution for sustainable gold recovery.

(IV) Potential limitations: (i) using Co and Ni in the nanocomposite raises concerns about the sustainability and availability of these critical minerals. However, the high reusability and potential for recycling these materials mitigate this concern to some extent; (ii) while the adsorbent can be reused, the regeneration process must be optimized to ensure minimal loss of adsorption capacity and material integrity over multiple cycles.

Table 2 Selective adsorption study of Au(III) ions using CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> adsorbent at specific extraction/elution conditions

[Coexisted metal ions] (ppm)	[Au(III)] (ppm)	[Au(III)], after adsorption (ppm)	[Au(III)], after elution (ppm)	Adsorption %	Elution %
Ag(I) 0.06, Pd(II) 0.12, Cu(II) 5.25, Ni(II) 1.7, Mn(II) 0.07, Fe(III) 9.3, Al(III) 10.8, Zn(II) 3.85, Pb(II) 0.03, Hg(II) 0.02, Cd(II) 0.02	9.8	0.235	9.42	97.6	98.5





Table 3 Comparing adsorption capacity for different materials that adsorb Au(III) ions

Adsorbents	Adsorption capacity mg g <sup>-1</sup>	Ref.
g-C <sub>3</sub> N <sub>4</sub>	111.25	Here
CoNi <sub>2</sub> S <sub>4</sub> @g-C <sub>3</sub> N <sub>4</sub>	200.6	
Zr-MOF functionalized with mercapto-1,3,4-thiodiazole	301.5	50
UiO-66-NH <sub>2</sub> modified by amidinothiourea	227.68	51
Dowex Marathon MSA commercial resin	73.53	52
Cellulose-based bio-adsorbent	5.07 mmol g <sup>-1</sup>	53
Humic acid	182.82	54
N-(2-[bis(2-aminoethyl)amino]ethyl)aminomethyl-polystyrene polymer bead	173.18	55
Imprinted ionic material SiO <sub>2</sub> (BGS/RHA)-TMPDT-Im-Au)	10.44	56
Ureido polymers containing large repeating ring	37.6	57
Clay mineral composite	108.3	58
Functionalized silica coating mercapto on iron sand magnetic material	125	59

## 4. Conclusion

In this study, we successfully synthesized a hybrid mesoporous nanocomposite of CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> for the selective recovery of Au(III) ions from SMB. The CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> nanocomposite achieved a remarkable adsorption capacity of 200.6 mg g<sup>-1</sup> at pH 2, significantly better than that of g-C<sub>3</sub>N<sub>4</sub> alone (111.25 mg g<sup>-1</sup>). The high selectivity for Au(III) ions over other competitive ions in both single and multi-ion systems underscores the potential of CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> as an effective adsorbent for gold recovery from e-waste. The adsorption process reached equilibrium within 60 minutes, fitting well with the PSO kinetic model. This rapid adsorption rate is advantageous for practical applications, where time efficiency is critical. The CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> nanocomposite exhibited excellent reusability, maintaining high adsorption efficiency (>85%) after ten cycles of adsorption and desorption. This recyclability, combined with the ease of synthesis, makes CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> a cost-effective and sustainable solution for industrial-scale gold recovery. In practical applications, the CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> nanocomposite demonstrated high efficiency in extracting Au(III) ions from the leach liquor of SMB. The adsorption efficiency was 97.6%, with an elution efficiency of 98.5%, highlighting the nanocomposite's practical utility in e-waste management and resource recovery. In conclusion, the CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> nanocomposite presents a promising and viable adsorbent for the choosy recovery of Au(III) from e-waste. Its high adsorption capacity, rapid kinetics, excellent selectivity, and reusability make it an attractive candidate for sustainable gold recovery processes. Future studies could focus on scaling up the synthesis process and exploring the application of CoNi<sub>2</sub>S<sub>4</sub>@g-C<sub>3</sub>N<sub>4</sub> for recovering other valuable metals from various types of e-waste.

## Data availability

Relevant data are included in the paper and its ESI files.† Other data and parameters generated or analyzed during the study are available from the corresponding author upon reasonable request.

## Conflicts of interest

There are no conflicts to declare.

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