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Graphene oxide–cerium oxide nanocomposite modified gold electrode for ultrasensitive detection of chlorpyrifos pesticide

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This research presents a novel approach for the detection of the pesticide chlorpyrifos (CLP) using a gold working electrode immobilized with a graphene oxide–cerium oxide (GO–CeO$_2$) nanocomposite in a phosphate buffer (PBS) solution with a pH of 7.0. Graphene oxide (GO) was synthesized *via* a modified Hummer's method, while cerium oxide (CeO$_2$) nanoparticles were prepared using a coprecipitation technique. The GO–CeO$_2$ nanocomposite was synthesized *via* sonochemical methods. Structural and morphological characterization of the prepared material was conducted using X-ray diffraction (XRD) and scanning electron microscopy (SEM) coupled with energy dispersive spectroscopy (EDX). Fourier transform infrared (FTIR) spectroscopy has been conducted for the confirmation of functional group presence in the prepared materials. Cyclic voltammetry (CV) was employed to investigate the interaction between the prepared material and the analyte. Further investigations using varying scan rates (5 mV s^{-1} to 300 mV s^{-1}) revealed a diffusion-controlled process at the electrode–electrolyte interface. Linear sweep voltammetry (LSV) experiments were conducted across a pH range of 5 to 9, with pH 7.0 showing enhanced response for the target pesticides in the presence of the buffer solution. Subsequent electrochemical measurements were performed at pH 7.0. Chronocoulometry was utilized to measure the effective electrode area for electrochemical interactions. Ultrasensitive square wave voltammetry (SWV) was employed for investigating the sensitivity over a concentration range of 1 fM to 100 μM and yielded the limit of detection (LOD) and limit of quantification (LOQ) as 47.7 fM and 159 fM respectively. Interference studies confirmed the selectivity of the prepared sensor, while stability and reproducibility were assessed through controlled experiments. Electrochemical impedance spectroscopy (EIS) was performed to investigate the interactions at the interface. This study provides insights into the development of selective electrochemical sensors for pesticide detection, with potential applications in environmental monitoring.

Introduction

Pesticides serve as essential neurotoxic agents in agricultural pesticides, offering significant advantages in managing insect pests, diseases, and enhancing plant growth for improved yield and quality. However, the inefficiency of the pesticide application, where less than 0.1% reaches the intended target,[1] necessitates excessive use for economic gains.[2] This trend has led to widespread environmental contamination, adversely impacting soil, water, air, and non-target organisms.[3,4] Consequently, pesticide residue contamination in agricultural produce has emerged as a global concern. Despite the associated risks, pesticides remain indispensable for pest and disease management, leading to overreliance on these chemicals.[5,6] The classification of chlorpyrifos (CLP) as a class II moderately

hazardous pesticide underscores its widespread use as a broad-spectrum organophosphate insecticide, nematicide and acaricide.[7,8] Numerous studies have detected CLP residue in various crops,[9–12] highlighting the urgency for rapid and effective detection methods in both plants and drinking water to mitigate health risks. The World Health Organization's classification and prevalence of CLP residue emphasizes the critical need for urgent research in this area.

Organophosphate (OP) insecticides have become increasingly common as a means of raising agricultural yields.[13] They are extensively employed in pre- and post-harvest interventions to manage several ailments affecting a variety of fruits and vegetables.[14,15] The environment and public health are seriously at risk due to their widespread use. Even though they are categorized as persistent pesticides with harmful effects on the environment, they are nevertheless widely utilized in many nations, which has resulted in significant contamination of food samples and ground water. OP pesticides are strong inhibitors of acetyl cholinesterase (AChE) that can result in

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severe cholinergic toxicity when applied topically, inhaled, or consumed.[16] The European Union (EU) states that in drinking water and food originating from plants, the maximum allowable quantities of pesticides and the compounds linked to them are 0.1 μg L^{-1} and 0.05 mg kg^{-1}, respectively (EEC, 1980).

Gas chromatography (GC), gas chromatography-mass spectrometry (GC-MS), high performance liquid chromatography (HPLC), ultra-high performance liquid chromatography-mass spectrometry (UPLCMS), and immunosensors are widely utilized for pesticide residue detection.[17–20] However, these methods are hindered by intricate pretreatment procedures, time-consuming analyses, and high operational costs, especially when handling large sample volumes in agricultural processing and monitoring industries. Additionally, the use of various chemical reagents during analysis poses environmental risks.[21] In contrast electrochemical techniques have gained prominence in recent decades due to their user-friendly operation, rapid on-site detection capabilities, cost effectiveness and reproducibility. These methods offer a practical alternative for pesticide screening, enabling timely and efficient analysis without the need of extensive sample preparation or hazardous chemicals.[22,23] Electrochemical sensors with their simplicity and sensitivity, present a promising avenue for enhancing pesticide detection efficiency while minimizing environmental impact.

W. S. T. Tun et al.[24] investigated the use of cellulose nanofibers and graphene oxide on screen printed electrode for the sensitive and selective detection of OPs, with chlorpyrifos as the model compound using acetylcholinesterase. The sensitivity of the sensor was assessed using square wave voltammetry, yielding a detection limit (LOD) of 2.2 nM and a quantification limit LOQ of 73 nM. H. Wang et al.[25] developed 2D/2D Z-scheme β-Bi$_2$O$_3$/g-C$_3$N$_4$ heterojunctions for chlorpyrifos detection where the composite improved the electron detection as confirmed by photoluminescence and electrochemical impedance spectroscopy. This method achieved and LOD of 0.03 ng mL^{-1} for chlorpyrifos. Several other nonenzymatic sensors for pesticides have been reported.[26–29] For instance, A. Kumaravel et al.[30] created an electrochemical sensor by immobilizing nano TiO$_2$ and cellulose acetate on a glassy carbon electrode (GCE) to detect chlorpyrifos. Using cyclic voltammetry, differential pulse voltammetry and amperometry they observed the reduction peak at −1.55 V linked to the electroreduction of the pyridine ring's C=N bond. This sensor achieved an LOD of 4.4 μM and LOQ of 14.7 μM. J. N. Nirmala et al.[31] modified a GCE with stearic acid for detecting parathion and methyl parathion. Cyclic voltammetry indicated a diffusion-controlled mechanism, with a 20% current increase compared to the bare electrode. Differential pulse voltammetry showed an LOD of 0.02 μM for parathion and 0.6583 μM for methyl parathion.

CuO–TiO$_2$ hybrid composite, synthesized by Tian et al.,[32] serves as effective electrode modification for methyl parathion (MP) detection without enzymatic assistance. The interaction between MP and CuO nanostructures suppressed the anodic current peak, facilitating MP detection, as evidenced by the differential pulse voltammetry (DPV) analysis yielding a low detection limit of 1.21 ppb with minimal interference. Inspired by these findings, researchers explored the combination of CuO

nanoparticles and two-dimensional materials to enhance sensor performance and stability. Soomro et al.[33] introduced and electrochemical sensor for malathion detection, using hydrothermally synthesized CuO nanostructures and amino acids as bio-templates. The strong affinity between CuO and malathion inhibits the anodic peak current, resulting in efficient detection of organophosphates. This sensor exhibited remarkable stability, and ultra-low limit of detection (LOD) of 0.1 nM. Fu et al.[34] pioneered an electrochemical sensor employing silica/exfoliated graphene composite films for MP determination. The synergistic interaction between nano-silica and graphene significantly improved the sensor's current response compared to bare glassy carbon electrode (GCE). Similarly, Song et al.[35] developed a sensor based on TiO$_2$ nanoparticles and graphene composite, demonstrating excellent electrocatalytic activity and sensitivity towards MP detection, attributed to graphene's rapid electron transport and TiO$_2$ nanoparticles' large surface area. Additionally, Karimian et al.[36] proposed a metal–organic framework (MOF)-based sensing layer functionalized with TiO$_2$–graphene oxide to address challenges associated with MOF-based sensors. This sensor exhibited a low LOD of 0.2 and 1.0 nM for propoxur (PXN) and chlorpyrifos (CLP), respectively, effectively monitoring residues in vegetables and water samples. In the above presented literature, different strategies have been opted to detect the chlorpyrifos in the lower concentrations. But the limit of detection achieved for the organophosphorus pesticide ranges from μM to nM.

In this study, we modified the gold (Au) electrode with GO, CeO$_2$ and GO–CeO$_2$ nanocomposites individually. Structural and morphological analyses were conducted using XRD, SEM and EDX respectively. The modified electrode served as a sensing platform in a conventional electrochemical cell with Ag/AgCl reference electrode and platinum counter electrode. Cyclic voltammetry investigated the incorporation of 1 mM CLP in PBS buffer (pH 7.0), while pH optimization was achieved through LSV (0–1 V). Chronocoulometry assessed the effective area, crucial for detection. SWV evaluated sensor's sensitivity. Controlled experiments examined interference, repeatability, reproducibility, and stability demonstrating the sensor's excellent performance. EIS probed electrode–analyte interactions, contributing to comprehensive understanding. In present work, the strategy has been adopted to use the synergistic effect of both the cerium oxide and that of graphene oxide in order to achieve the ultra-lower limit of detection because of the enhanced electrocatalytic activity of the cerium oxide and high surface area, exceptional conductivity and narrow band gap which facilitates the flow of electrons from graphene oxide, the combine effect of both materials in the form of composite achieved the limit of detection up to femto molar range.

Materials and methods

The chemicals and reagents utilized in this study were of high purity and sourced from reputable suppliers. Carbendazim (97% purity), carbaryl Pestanal (analytical standard), and chlorpyrifos Pestanal™ (analytical standard) were procured

from Sigma-Aldrich. Cerium nitrate hexahydrate Ce(NO$_3$)$_3$·6H$_2$O with a purity of 99%, graphite powder (99% purity), and potassium permanganate (KMnO$_4$) with a purity of 99.5% were also obtained. Additionally, hydrogen peroxide (H$_2$O$_2$) with a concentration of 30%, sodium hydroxide (NaOH) with a purity of 98.5% and sulfuric acid (H$_2$SO$_4$) with a concentration of 98.5% were sourced. Sodium chloride (NaCl) and potassium chloride (KCl) with purities of 99% and 99.5% respectively, were acquired. Furthermore, potassium hydrogen phosphate (KH$_2$PO$_4$) and sodium hydrogen phosphate (Na$_2$HPO$_4$) with purities of 98% were purchased from Sigma-Aldrich. No further modifications were necessary for their use. The pH of the solutions was adjusted by the addition of hydrochloric acid (HCl) and sodium hydroxide (NaOH) to deionized water.

The synthesis of materials was conducted according to established protocols with some modifications. Graphene oxide (GO) was synthesized using a modified Hummer's method.[37] Initially 1 g of graphite was combined with 25 mL of H$_2$SO$_4$ in an ice bath. Subsequently, 3 g of potassium permanganate (KMnO$_4$) was slowly added with continuous stirring at temperatures below 20 °C for three hours. Following this, 50 mL of deionized water was added dropwise, ensuring the temperature remains below 50 °C. The resulting solution transitioned to a deep brown color, indicating GO formation. To complete oxidation, 100 mL of water and 5 mL of hydrogen peroxide were added successively. The product was rinsed with deionized water and ethanol, followed by drying at 60 °C.

CeO$_2$ nanoparticles were synthesized *via* co-precipitation. A 0.1 M precursor solution was prepared by dissolving cerium nitrate hexahydrate (Ce(NO$_3$)$_3$·6H$_2$O) in 100 mL of deionized water, then heated to 75–80 °C. Sodium hydroxide was dissolved in deionized water to create a 5 M solution. After 4 hours at 80 °C, the sample was washed, dried, and annealed at 450 °C.

Subsequent experiments involved dispersing 25 mg of GO in 50 mL of deionized water, followed by sonication. A 0.5 M Ce(NO$_3$)$_3$·6H$_2$O solution and a 5 M NaOH solution were prepared and added subsequently. After sonication and washing steps, drying occurred at 100 °C for 8 hours and then annealing was done at 400 °C for 3 hours.

In this investigation, a 1 cm × 0.6 cm Au electrode was employed as the working electrode in electrochemical cell for CLP detection. The bare Au electrode underwent ultrasonic washing in sequential steps: first in deionized water, then in ethanol, and finally in acetone. Subsequently, it was air dried at room temperature. The active material was combined with PVDF at a ratio 4:1 in the presence of NMP (1-methyl-2-pyrrolidinone) and homogenized for 30 minutes to produce a slurry. The amount of slurry (15 μL) was then drop-cast onto the surface of the cleaned Au electrode.[38] Following deposition, the Au electrode was dried under a standard bulb lamp for one hour. The resulting electrode configuration is denoted as GO–CeO$_2$/Au.

The electrochemical experiments utilized the Gamry Reference 3000 potentiostat/galvanostat/ZRA, operating with a conventional three-electrode setup. The reference electrode employed was Ag/AgCl, while a platinum wire served as a counter electrode to complete the electrical circuit. X-ray diffraction (XRD) analysis was conducted using Shimadzu XRD-6000 instrument. For imaging and elemental analysis, a field emission scanning electron microscope (FESEM) equipped with energy-dispersive X-ray spectroscopy (EDX) capabilities was utilized, specifically the Tescan Maia 3 model. Fourier transform infrared spectroscopy (FTIR) was performed by ATR Alpha Bruker.

Results and discussion

The X-ray diffraction (XRD) analysis was conducted to characterize the structural properties of the materials under study, including graphene oxide (GO), cerium oxide (CeO$_2$) and graphene oxide–cerium oxide (GO–CeO$_2$) nanocomposite. As illustrated in Fig. 1(a), GO exhibited a diffraction peak centered at $2\theta = 11.6°$, indicative of its layered structure with a d-spacing of 7.5 Å. The presence of oxygen functional groups and water molecules within the carbon layer structure contributed to the increased interlayer spacing.[39] The XRD of CeO$_2$ (red curve) closely matched the literature values (JCPDS no. 81-0792), with distinct peaks corresponding to the face-centered cubic phase of CeO$_2$ crystal at various θ values. In the XRD pattern of GO–CeO$_2$ nanocomposite (blue curve), the suppression of the graphene oxide peak indicated the successful incorporation of CeO$_2$ in the GO matrix. The distortion in the layered structure of
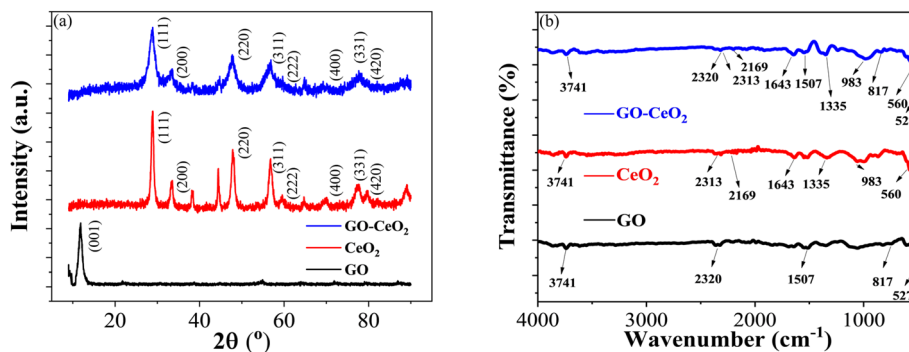

Fig. 1 (a) XRD spectra of GO (black curve), CeO$_2$ (red curve) and GO–CeO$_2$ (blue curve). (b) FTIR spectra of GO (black curve), CeO$_2$ (red curve) and GO–CeO$_2$ (blue curve).

GO due to higher order crystallinity of CeO_2 confirmed the effective formation of the nanocomposite. The mean crystallite size of the nanocomposite was determined to be 12 nm, while CeO_2 nanoparticles exhibited a crystallite size of 15.9 nm, calculated using the Scherrer formula.

Fig. 1(b) displayed the FTIR spectra of GO, CeO_2 and GO–CeO_2 nanocomposites. The IR spectrum of GO is displayed by the black spectral line. The broad absorption band, located at 3741 cm^{-1}, represents the hydroxyl moieties' –OH stretching vibrational mode on the GO matrix's surface. Bending vibrations of carbo–hydroxyl (–C–OH) of carboxylic groups and bending vibration of –C=C– cause the peak that is located at wavenumbers of 1507 cm^{-1} and 817 cm^{-1}, respectively. The red line illustrates the infrared spectrum of CeO_2, which has absorption edges at wavenumbers 983 cm^{-1} and 560 cm^{-1}, respectively. These correspond to the symmetric stretching vibration of the Ce–O–Ce and the stretching of the Ce–O bands in the chains. However, the IR spectrum of GO–CeO_2 contains absorption bands at wavenumbers 3741 cm^{-1}, 1643 cm^{-1}, 1335 cm^{-1}, 983 cm^{-1}, and 817 cm^{-1}. These correspond to the stretching vibration of carbon and epoxide groups and the –OH stretching, –C=C– stretching, –C–H bending, and –OH bending of carboxylic groups, respectively. Conversely, the peaks seen at lower wavenumbers, 560 cm^{-1} and 527 cm^{-1}, are associated with the stretching vibrations of Ce–O and Ce–O–Ce. It therefore validates the effective insertion of CeO_2 nanoparticles onto the GO matrix.

The scanning electron microscopy (SEM) analysis, along with energy dispersive X-ray spectroscopy (EDX), provided insights into the morphology and composition of the materials under investigation. SEM images depicted in Fig. 2(a)–(c) revealed key features of GO, CeO_2 particles and their composite system. Fig. 2(a) exhibited a uniform GO sheet with distinct layers and noticeable wrinkles, suggesting a structured morphology. CeO_2 particles, as observed in Fig. 2(b), indicate a narrow and seemingly homogeneous size distribution. CeO_2 particle concentration increases on the GO sheet as displayed in Fig. 2(c).

EDX analysis shown in Fig. 2(d)–(f), confirmed the elemental composition of the materials. GO demonstrated an oxygen content of 43 atom%, with a carbon to oxygen atomic ratio of 1.24, suggesting a higher oxygen content than the empirical formula proposed by Boehm et al.[40] CeO_2 nanoparticles revealed the presence of oxygen and cerium atoms without any contaminants or surfactants, as indicated by Fig. 2(e). Furthermore, EDX analyses confirmed the elemental components (cerium, oxygen and carbon) in the nanocomposite validating its composition.

Electrochemical measurements were conducted using various techniques to assess the performance of the modified electrode GO–CeO_2/Au and its interaction with the analyte. Cyclic voltammetry (CV) explored the electrochemical behavior within a potential range of 0–1 V. Chronocoulometry (CC) determined the effective area with a potential range of −0.2 to +0.6 V. Square wave voltammetry (SWV) was employed to assess the sensitivity and limit of detection, operating within the 0–1 V potential range, with a pulse amplitude of 25 mV and integration period of 0.02 s. Electrochemical impedance spectroscopy (EIS) investigated the interfacial properties over a frequency range from 0.1 to 10^5 Hz, applying an AC voltage of 500 mV.
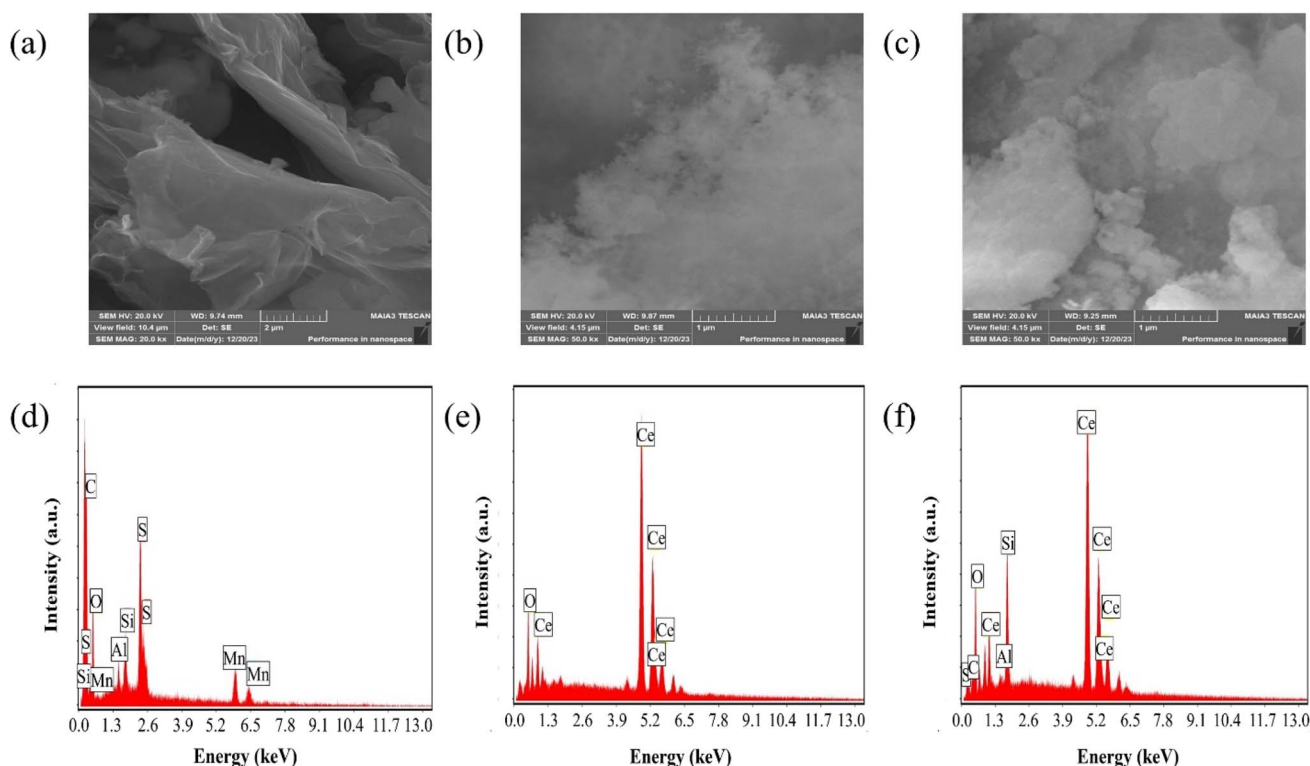

Fig. 2 (a–c) SEM micrographs of GO, CeO_2 and GO–CeO_2 respectively. (d–f) Elemental analyses of GO, CeO_2 and GO–CeO_2 respectively.
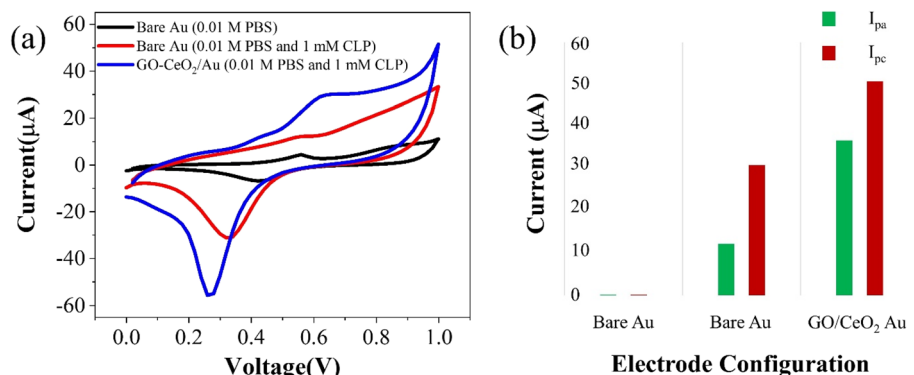
Fig. 3 (a) Cyclic voltammetry response of Au and GO−CeO₂/Au electrode with and without CLP. (b) Peak cathodic and peak anodic current responses for Au and GO−CeO₂/Au electrodes.
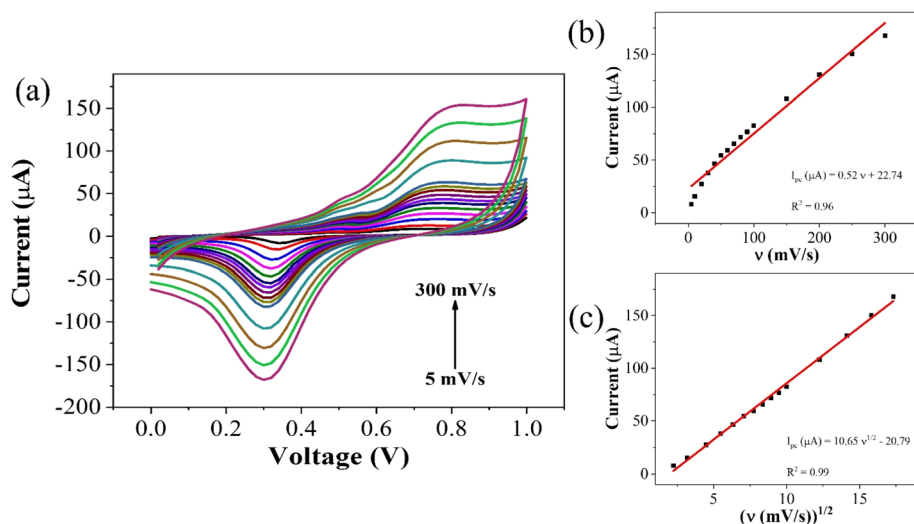

Fig. 4 (a) Scan rate dependent response for CLP at GO−CeO₂/Au electrode (b) plot of peak current *vs. v* (c) plot of peak current *vs. v$^{1/2}$*.

Additionally, linear sweep voltammetry (LSV) optimized pH conditions within the 0–1 V potential range against Ag/AgCl, utilizing a scan rate of 50 mV s^{-1}.

The electrochemical behavior of the developed sensor is depicted in Fig. 3(a). The black trace represents the response on the bare Au electrode in a supporting buffer solution (PBS, pH = 7.0), showing baseline behavior. An enhanced signal (red trace) exhibits pronounced oxidation behavior and a distinct reduction peak at 320 mV upon the addition of 1 mM chlorpyrifos (CLP) analyte. Upon introduction of the modified sensor surface into the electrochemical cell with the analyte, an increased oxidative signal (blue trace) at 620 mV and prominent reduction signal at 230 mV are observed. This indicates a significant catalytic response of GO–CeO₂ nanocomposite with CLP. Previous studies have attributed this peak to electroreduction of the pyridine's ring C=N *via* a 2e$^-$ transfer process in chlorpyrifos. The presence of three chlorine atoms causes electrons delocalization at the −C=N− bond, leading to increased reduction at lower negative potentials.[41,42]

When employing only GO as the immobilizing material onto the surface of the working electrode, no discernible redox behavior is observed. Instead, only capacitive response is evident. Conversely, when CeO₂ was utilized, clear redox behavior is observed, however the reduction peak in the case of GO–CeO₂ nanocomposite is still more pronounced owing to the synergistic effect of the composite.

To elucidate the reaction kinetics at the interface, a scan rate study was conducted over a range of 5–300 mV s^{-1} within the potential range of 0–1 V. Analysis of the voltammograms revealed a rightward shift in the anodic peak, while the cathodic peak initially shifted leftward before moving rightward. The observed behavior suggests a diffusion-controlled reaction occurring at or near the electric double layer (EDL). This interference is supported by the high correlation coefficient (R^2) of 0.99 for plots of peak cathodic current (i_{pc}) against the square root of scan rate, and 0.97 for the plots of i_{pc} against the scan rate itself (Fig. 4). The linear fit equation for the cathodic peak current is expressed as follows:

$$i_{pc}\ (\mu A) = -(4.95 \pm 0.04)v^{1/2} + 6.72 \pm 0.44,\ (R^2 = 0.99)$$
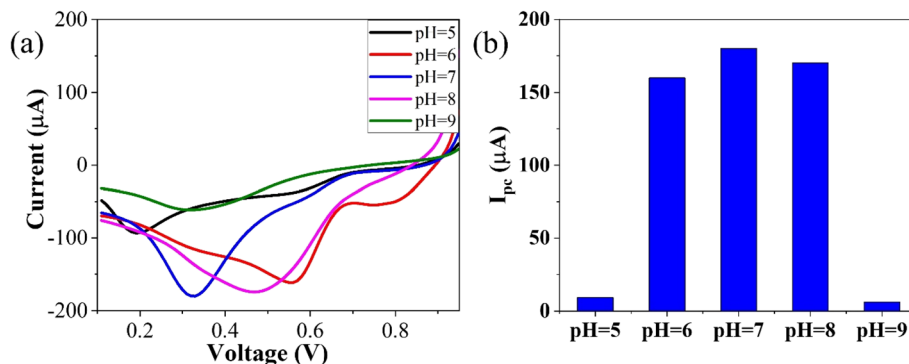
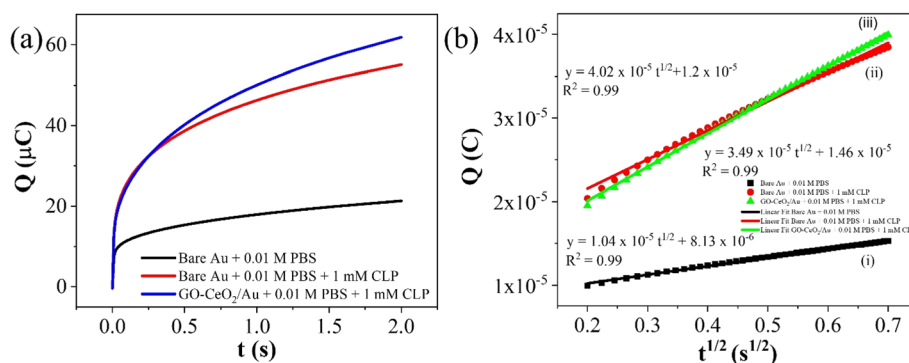Fig. 5 Optimization of reduction peak signals of 1 mM CLP in the presence of 0.01 M PBS.

Fig. 6 (a) Plots of Q–t curves (b) (i) Q–$t^{1/2}$ curves for bare Au in 0.01 M PBS (ii) Q–$t^{1/2}$ curves for bare Au in 0.01 M PBS + 1 mM CLP (iii) Q–$t^{1/2}$ curves for GO–CeO$_2$/Au in 0.01 M PBS + 1 mM CLP.

To optimize the pH, linear sweep voltammetry (LSV) was conducted using the GO–CeO$_2$/Au modified electrode within a potential range of 0–1 V in the presence of phosphate buffer saline (PBS) with a concentration of 0.01 M. The prepared solution containing the analyte CLP at a concentration of 1 mM was tested at pH values ranging from 5 to 9. The voltammograms obtained at different pH values were analyzed, and it was observed that at pH 7.0, the signal exhibited optimal characteristics, as depicted in Fig. 5. The bar graph accompanying the voltammograms illustrated the signal amplitude at each pH level. Based on these findings, a pH of 7.0 was selected as the optimal condition for the subsequent electrochemical investigations, providing consistent and reliable results for further analysis and experimentation, facilitating accurate assessment of GO–CeO$_2$/Au modified electrode's performance in detecting CLP.

The electrochemical method of chronocoulometry (CC) was employed to measure both capacitive and coulombic charge (Q_{ad}) over time, developed at the electric double layer interface between the electrode and analyte. This technique is instrumental in determining both the effective area of the electrode and the surface charge (Γ) of absorbed molecules. The comparison of electrode effective areas is illustrated in Fig. 7(a). In the study, the calculated surface area of bare Au in 0.01 M PBS was found to be 0.00964 cm^2, as shown in Fig. 7(b) (black trace). Conversely, in the presence of 0.01 M PBS (pH = 7.0)

along with 1 mM CLP, the calculated area increased to 0.01658 cm^2, depicted in Fig. 6(b) (red trace). Furthermore, when Au was immobilized with GO–CeO$_2$, the effective area was further augmented to 0.02155 cm^2 as illustrated in Fig. 6(c) (green trace). These measurements were obtained by analyzing the slope of the linear fitted curve correlating charge (Q) and time ($t^{1/2}$), utilizing the Anson equation for the diffusion-controlled process.

$$Q(t) = \frac{2nFA_{eff}C\sqrt{Dt}}{\sqrt{\pi}} + Q_{dl}$$

where $Q(t)$ is the charge as function of time, n is the electron transfer, F is the Faraday's constant, C is the concentration of $[Fe(CN)_6]^{3-/4-}$, and D is the diffusion coefficient and Q_{dl} is the double layer capacitance respectively. This method provides valuable insight into the electrochemical behavior of the system and aids in understanding the kinetics of charge transfer process at the electrode surface, essential for optimizing the performance of electrochemical devices.

To assess the operational range of the prepared sensor, square wave voltammetry (SWV) was conducted across a concentration spectrum spanning from 1 fM to 1 mM. SWV is a highly effective technique utilized for both for qualitative and quantitative determination of the analyte, even when present at trace level in the solution. The methodology involves the imposition of a symmetric square wave potential onto a base
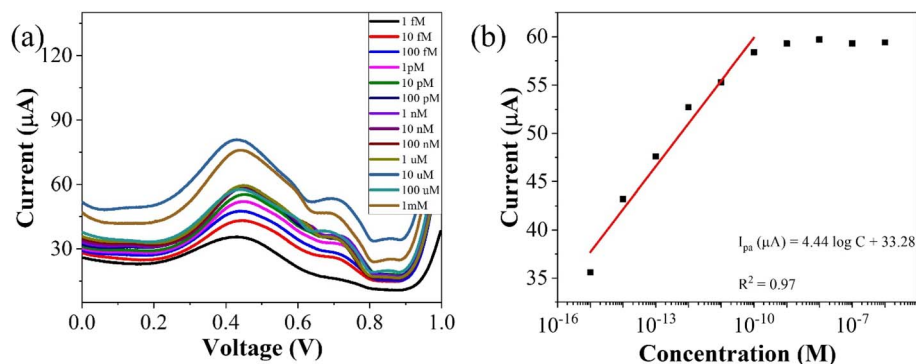
Fig. 7 (a) Concentration profiling using SWV (b) corresponding plot of peak currents *vs.* concentration of the analyte.
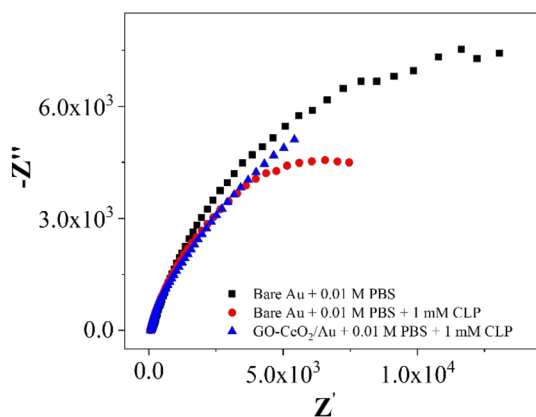

Fig. 8 Nyquist plots of Au in presence of 0.01 M PBS and AU and GO–CeO$_2$/Au in the presence of 0.01 M PBS and 1 mM CLP.

staircase potential applied to the working electrode. Each complete square wave cycle corresponds temporally to one step within the staircase waveform. Within this cycle, the current is sampled twice: once at the concentration of the forward pulse (I_f), and once at the termination of reverse pulse (I_r). This process engenders square wave modulation, wherein reverse pulses prompt the inverse reaction of any products engendered by the preceding forward pulse. Plotting the net current ($I_f - I_r$) against the base staircase potential yields a voltammogram, typically exhibiting a peak-shaped morphology symmetrically aligned about the half wave potential.

The benefits of employing the SWV over the differential pulse voltammetry (DPV) are manifold: firstly, the SWV rapid scan rate facilitates expedited analysis, minimizes the consumption of the electroactive species *vis-à-vis* a DPV, and mitigates issues associated with electrode surface blocking. Secondly, SWV boasts enhanced sensitivity compared to DPV, owing to its incorporation of reverse current considerations. Finally, SWV exhibits superior capacity for discerning background signals, particularly those stemming from oxygen reduction process. The results of the square wave voltammogram illustrates the anodic peak response of the targeted analyte, as depicted in Fig. 8(a), while the ensuing linear correlation is discernible in Fig. 8(b).

The linear fit equation is represented as follows,

$$i_{pa} \text{ (µA)} = (3.96 \pm 0.21)\log x + 97.41 \pm 2.22, \ R^2 = 0.97$$

The slope of this curve (m) for the blank solution and the standard deviation (s) can be used for the estimation of limit of detection as well as for limit of quantification as given below,

$$\text{LOD} = \frac{3s}{m}$$

$$\text{LOQ} = \frac{10s}{m}$$

Table 1 Comparison of different metal-oxide frameworks with GO for the detection of pesticides with current study

| Material	Electrochemical technique used	Analyte	LOD	Reference
CuO-NPs/3DGR	DPV	Malathion	0.01×10^{-9} M	43
TiO$_2$/graphene/GCE	CV, LSV	Methyl parathion	1.00×10^{-9} M	44
CoTCPP/Co$_3$O$_4$/GO	DPV	Methyl parathion	1.10×10^{-8} M	45
Porphyrin/Co$_3$O$_4$/graphene	DPV	Methyl parathion	1.10×10^{-8} M	46
CuO-NPs/3DGR	DPV	Malathion	0.01×10^{-9} M	47
Au–ZrO$_2$–GNs/GCE	SWV	Methyl parathion	3.79×10^{-9} M	48
Au–Pd/rGO/CPE	SWASV	Parathion	8.00×10^{-9} M	49
TiO$_2$/CA/GCE	DPV	Chlorpyrifos	3.50×10^{-6} M	50
CNFs/GO/CS–GO/AChE/SPCE	SWV	Chlorpyrifos	3.79×10^{-9} M	51
GO–CeO$_2$/Au	SWV	Chlorpyrifos	47.7×10^{-15} M	This work
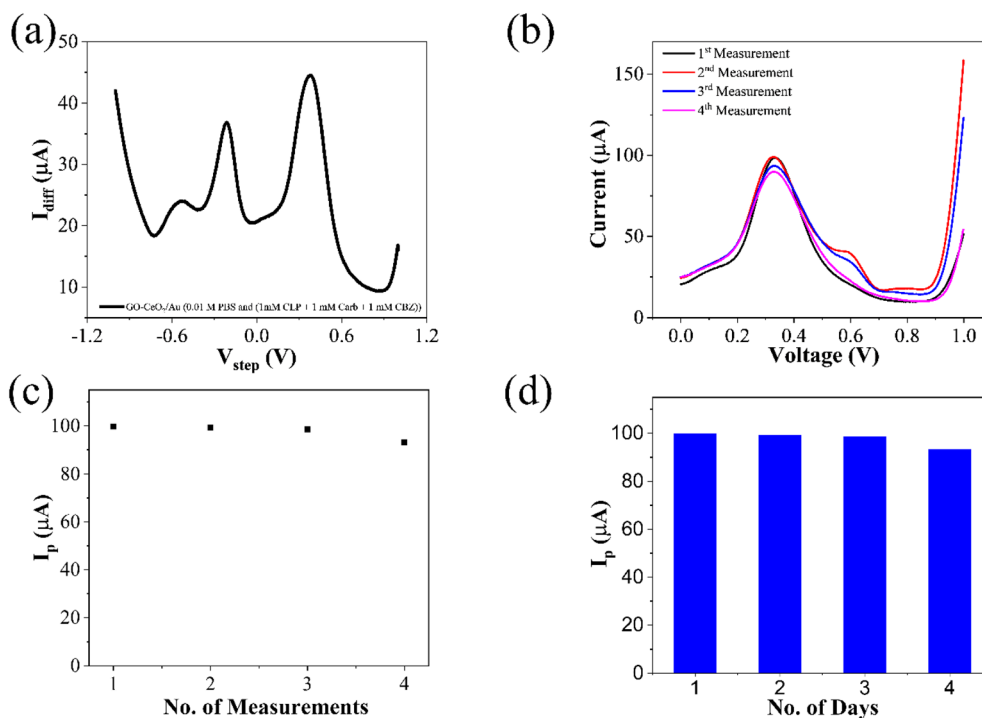
Fig. 9 (a) Interference study of CLP in the presence of pesticides: carbaryl (CARB) and carbendazim (CBZ) (b) repeatability of the prepared sensor (c) reproducibility over several measurements and (d) stability of the prepared sensor over a period of four days.

Table 2 Comparison of different studies for the detection of chlorpyrifos pesticides with current study

Electrode/modifiers	Quantification method	pH, electrolyte	Linear range	Detection limit	Ref.
AChE/Pin5COOH/Fe$_3$O$_4$NP/GCE	AMP	pH 7.0 phosphate buffer	1.5 to 70 nM	9.1 nM	52
AChE/ZrO$_2$/rGO	AMP	pH 7.0 phosphate buffer	10^{-9} to 10^{-4} M	10^{-13} M	53
CuO NFs-SWCNTs/GCE	DPV	pH 7.0 phosphate buffer	0.1 to 150 ng L^{-1}	70 pg mL^{-1}	54
NA/AuNPs/rGO–NH$_2$/AChE/GCE	CV	pH 8.0 phosphate buffer	0.021 to 0.122 mg mL^{-1}	14 ng mL^{-1}	55
FTO–AuNPs–chl–Ab	DPV	pH 7.5 phosphate buffer	1 fM to 1 μM	10 fM	56
GO–CeO$_2$/Au	SWV	pH 7.0 phosphate buffer	1 fM to 100 nM	47.7 fM	Present work

By using the above relations, we have estimated the LOD as 47.7 fM and LOQ as 159 fM respectively.

In the investigations of interfacial interactions between the modified working electrode surface and the target analyte, electrochemical impedance spectroscopy was conducted over a frequency range spanning from 0.1 to 10^5 Hz. Experimental conditions include a DC voltage 0 V *vs.* E_{ref} and an AC voltage of 500 mV RMS, with an initial delay of 10 s. EIS spectra depicted the response of the bare Au electrode in the presence of 0.01 M PBS pH = 7.0 (black trace), the response of bare Au electrode in the presence of 0.01 M PBS pH = 7.0 and 1 mM CLP (red trace), and the response of the modified electrode (GO–CeO$_2$/Au) in the presence buffer solution of 0.01 M PBS pH = 7.0 and analyte concentration 1 mM CLP (blue trace). Comparison of various metal-oxide composites with graphene oxide (GO) for pesticide detection with the current study is presented in Table 1.

The interference experiment was conducted by mixing three pesticides namely carbendazim, carbaryl along with the target

pesticide chlorpyrifos having molarities (1 mM each) under the optimized conditions by using the square wave voltammetry in the potential range from −1 to +1 V *vs.* Ag/AgCl in the presence of 0.01 M PBS (pH = 7.0). Three distinguished peaks at −531 mV for carbendazim, −212 mV for carbaryl and 381 mV for chlorpyrifos showed the interference capability of the prepared sensor. While the enhanced peak showed the selectivity of the prepared sensor towards target pesticide chlorpyrifos in (Fig. 9(a)). Repeatability assessment involved four consecutive SWV measurements within the optimized potential window, demonstrating close alignment of all four reduction current peaks with previous measurements, indicating significant repeatability (Fig. 9(b)). Reproducibility was evaluated using for different electrodes under optimized parameters, showcasing closely matched responses (Fig. 9(c)), affirming the sensor's practical utility. Stability assessment involved monitoring the voltametric response of the same electrode over four days, with deviations from ideal behavior measured as changes in current.

Corresponding reduction peaks are depicted in bar graphs (Fig. 9(d)), showing good agreement with initial measurements.

Table 2 represents the comparison of the developed systems for the detection of chlorpyrifos, which also included the electrochemical techniques employed, optimized value of pH along with buffer. Linear range of the prepared systems and limit of detection for the analyte under consideration.

Conclusion

The successful synthesis of a GO–CeO$_2$ nanocomposite *via* sonochemical method is highlighted, with SEM images confirming the uniform distribution of CeO$_2$ nanoparticles on graphene oxide sheets. The modified electrode exhibited excellent electrocatalytic activity for the detection of chlorpyrifos (CLP) in the presence of carbendazim and carbaryl in 0.01 M PBS (pH = 7.0). Graphene oxide sheets effectively prevented CeO$_2$ nanoparticles agglomeration, enhancing sensitivity and selectivity for CLP detection. Scan rate studies indicate a diffusion-controlled mechanism at the interface, while chronocoulometry probed the actual electrochemical area involved. All measurements were conducted at pH 7.0, with ultra-sensitive square wave voltammetry revealing a limit of detection (LOD) and limit of quantification (LOQ) of 47.7 fM and 159 fM, respectively. Interface studies, repeatability, reproducibility, and stability assessments confirmed the sensor's practical applicability. Electrochemical impedance spectroscopy provided additional insights into interface interactions. The study contributes valuable knowledge to the development of highly sensitive selective electrochemical sensors for pesticide detection, with potential applications in environmental and agricultural sectors.

Data availability

The data that support the findings of this study are available on request from the corresponding author.

Author contributions

Sami Ullah Khan: performed the experiments and prepared initial draft. Muhammad Atif: structural studies, and Zulqurnain Ali: morphological studies. All authors have contributed to the interpretation of the results and written the manuscript. Waqas Khalid: supervision, project administration, manuscript preparation.

Conflicts of interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

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