


 Cite this: *RSC Adv.*, 2024, 14, 24130

2H-SnS₂ assembled with petaloid 1T@2H-MoS₂ nanosheet heterostructures for room temperature NO₂ gas sensing†

 Shraddha Hambir,^{ae} Shashikant Shinde,^b H. M. Pathan,^a Som Datta Kaushik,^c Chandra Sekhar Rout^{id}^d and Shweta Jagtap^{id}^{*e}

In this study, we explored the gas-sensing capabilities of MoS₂ petaloid nanosheets in the metallic 1T phase with the commonly investigated semiconducting 2H phase. By synthesizing SnS₂ nanoparticles and MoS₂ petaloid nanosheets through a hydrothermal method, we achieve notable sensing performance for NO₂ gas at room temperature (27 °C). This investigation represents a novel study, and to the best of our knowledge no, prior similar investigations have been reported in the literature for 1T@2HMoS₂/SnS₂ heterostructures for room temperature NO₂ gas sensing. The formed heterostructure between SnS₂ nanoparticles and petaloid MoS₂ nanosheets exhibits synergistic effects, offering highly active sites for NO₂ gas adsorption, consequently enhancing sensor response. Our sensor demonstrated a remarkable sensing response ($R_a/R_g = 7.49$) towards 1 ppm of NO₂, rapid response time of 54 s, baseline recovery in 345 s, good selectivity and long-term stability, underscoring its potential for practical gas-sensing applications.

Received 30th April 2024

Accepted 23rd July 2024

DOI: 10.1039/d4ra03194f

rsc.li/rsc-advances

1 Introduction

Emission of harmful gases from various sectors such as transportation, agriculture manufacturing, construction and power generation affect both human health and the environment. To address this issue, severe regulations are being implemented globally to restrict the emission of air pollutants.^{1,2} Among these gases, NO₂ is well known for being the most harmful gas to the entire ecosystem. If NO₂ concentration exceeds 53 ppb, it can lead to serious health issues such as, lung disease, pulmonary irritations, chronic diseases *etc.* Hence, it is important to keep track of a number of gases in our surroundings on a regular basis^{3,4} and it is highly important to develop a cost-effective gas detection device that is efficient, precise and highly sensitive and selective.

Several methods are available for detection of toxic gases, including chemiresistive, calorimetric, optical, and electrochemical techniques.⁵ However, few methods encounter challenges like limited accessibility, higher costs, and sensitivity constraints *etc.* Resistive gas sensors have garnered

considerable interest due to their compact size, ease of fabrication, straightforward operation, and cost-effectiveness in manufacturing.^{6,7} Metal oxide semiconductors (MOS) have served as effective gas sensors in the past, but their reliance on high operating temperatures led to increased power consumption. Lately, two-dimensional (2D) layer-structured materials have gained significant interest across several domains.⁸ Notably, two-dimensional semiconductor materials distinguish themselves through their extensive band gap coverage, interface free of dangling bonds, high mobility, and rapid carrier transport. A high specific surface area and rich active sites make it possible to adsorb a large number of gas molecules, which is one of the best properties that make 2D structures advantageous in gas sensing applications.⁹ In addition, the electrical characteristics of 2D materials can be modulated by altering the number of layers.¹⁰ These materials undoubtedly provide a reasonably good foundation for developing high-performance gas sensors.⁹ In particular, tin disulfide (SnS₂), a characteristic 2D layered material with weak van der Waals interactions between its layers has received significant interest.¹¹ The non-solubility and non-toxicity of SnS₂ in aqueous solutions make it a very promising material. SnS₂ is characterised by its unique structure at the atomic layer level, which results in an abundance of available functional active sites.¹² Being a more electronegative nanomaterial, SnS₂ accumulates chemically adsorbed molecules on its surface, generating an electrostatic potential. The depletion region of this surface potential corresponds to the size of the entire nanostructure. This unique attribute of n-type SnS₂ nanosheets enhances gas adsorption

^aDepartment of Physics, Savitribai Phule Pune University, India

^bMES's Department of Physics, Nowrosjee Wadia College, Pune 411001, India

^cUGC-DAE Consortium for Scientific Research Mumbai Centre, BARC, Mumbai, India

^dCentre for Nano and Material Sciences, Jain (Deemed-to-be University), Jain Global Campus, Ramnagar, Bangalore, India

^eDepartment of Electronic and Instrumentation Science, Savitribai Phule Pune University, India. E-mail: shweta.jagtap@gmail.com

 † Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d4ra03194f>


sites, making it a strong option for designing and making high-performance NO₂ gas monitors.^{12,13} Moreover, the two-dimensional surface of SnS₂ is favourable for incorporation other semiconductors, which results in the formation of well-contacted heterojunctions that improve carrier conduction.^{14,15} It has been found that the formation of heterojunctions is a highly effective method for manipulating the electronic state of the SnS₂ surface, thereby substantially enhancing its gas sensing characteristics.¹⁶ Comparing heterojunction materials to individual materials, they often perform better due to their diverse morphologies and band alignments. This improvement may be linked to the excellent heterointerface, which promotes fast charge transfer. Furthermore, the curved shape of the conductive band and the valence band in a typical heterojunction structure often causes the Fermi level to attempt equilibrium, which ultimately leads in the formation of a depletion layer. This effect reduces response and recovery times by directly contributing to high conductivity.^{16,17} The construction of two-dimensional heterojunction nanomaterials has been a popular and advanced method for designing gas sensors in recent years. Similar to SnS₂, MoS₂ possesses a traditional layer arrangement with weak van der Waals interactions.¹⁵ Therefore, the development of improved heterointerfaces by employing 2D SnS₂ and MoS₂ nanosheets stands as a promising approach for further improving electrical capabilities. Furthermore, the metastable 1T phase exhibits superior electrical conductivity for charge transfer and higher adsorption energy to NO_x compared to the semiconducting 2H phase, suggesting its suitability for gas sensing applications. For example, MoS₂ nanosheets coated with SnS₂ nanoparticles (MoS₂/SnS₂) were developed by Jia-Bei Liu and colleagues by using the simple hydrothermal process and mechanical exfoliation method. The results of this study indicated that the SnS₂ nanoparticles, which serve as an efficient antioxidative decoration, may increase the stability of MoS₂ nanosheets. This provides a potential way to create high-stability NO₂ gas sensors at ambient temperature.¹⁸

Considering the synergistic effects of heterostructures to boost gas sensing performance. This novel investigation explored the 1T@2H MoS₂/SnS₂ heterostructure by a simple hydrothermal approach to construct a two-dimensional layered 2H-SnS₂ decorated on petaloid 1T-MoS₂ nanosheets for room temperature NO₂ gas sensing, marking the first instance of such research in the literature without prior similar studies. Utilizing a range of characterization techniques, comprehensive structural and functional investigations were conducted on the synthesized material. The hierarchical structure of the SnS₂/MoS₂ sensor proved superior to the pristine SnS₂ sensor, exhibiting good response ($R_g/R_a = 7.49$) to 1 ppm NO₂ at ambient temperature. The sensor is particularly noteworthy for its excellent selectivity and consistent repeatability.

2 Experimental

2.1 Synthesis of SnS₂ nanoparticles

SnS₂ nanoparticles were synthesized using a one-pot hydrothermal process. The procedure involved 0.7 gm of SnCl₄·5H₂O,

added in 30 ml of deionized water with stirring until a clear solution was formed. Subsequently, 1.2 g of CH₃CSNH₂ was added to the above solution, and the mixture was stirred for ten minutes. The resulting solution was then transferred to a 50 ml stainless steel autoclave lined with Teflon. Then the autoclave was maintained at 180 °C for 24 hours to facilitate the formation of SnS₂ nanomaterials, followed by centrifugation, the precipitates subsequently subjected to overnight drying at 60 °C to obtain the final product.

2.2 Synthesis of 1T@2H-MoS₂/SnS₂ heterostructures

SnS₂/MoS₂ heterostructures were also formed by a hydrothermal process. To produce a uniform suspension of SnS₂, 0.2 g of SnS₂ was introduced to 90 ml of deionized water and ultrasonicated for 5 minutes. Later SnS₂ solution was mixed vigorously for two hours with addition of 1.79 g of ammonium molybdate ((NH₄)₆Mo₇O₂₄·2H₂O), 0.68 g of CH₃CSNH₂, and 0.885 g of C₁₉H₄₂BrN. Following that, the prepared mixture was transferred into a Teflon-lined stainless-steel autoclave with a volume of 180 ml which was then subjected to heating at 180 °C for a duration of 24 hours. After cooling to room temperature, the final product was thoroughly washed multiple times with deionized water, ethanol and dried at 60 °C. For synthesis of MoS₂ similar procedure were carried out without addition of SnS₂.

2.3 Material characterization

X-ray powder diffraction (XRD) analysis was performed using a Bruker D8 Advance instrument with Cu-K α radiation ($\lambda = 0.15406$ nm). The investigation covered a range from 5° to 80° (2 θ) to identify the crystal phase structures of the samples. The optical spectra of the material were obtained using a UV-vis spectrometer (Jasco V750). Additionally, the morphology of the synthesized material was examined through field emission scanning electron microscopy (FESEM) using a FET Nova Nano SEM 450. The current–voltage (I – V) characteristics of the semiconductor were assessed using a semiconductor parameter analysers system, (Keithley 4200A). X-ray photoelectron spectroscopy (K alpha + X-ray photo spectrometer with X-ray source – Al K). The spectra were analyzed using the XPSpeak41 software. The core peaks were deconvoluted and identified with the help of literature. The analysis of the specific surface area and pore size distribution of the synthesized material was done using the Brunauer–Emmett–Teller (BET) model and the Barrett–Joyner–Halenda (BJH) technique respectively.

2.4 Gas sensor fabrication and gas sensing measurement

For the gas sensing test, a gold electrode was deposited on a 1 cm by 1 cm alumina substrate through the thermal evaporation method. The device features interdigitated electrodes (IDES) designed to facilitate gas detection. It includes two main electrodes, each 0.6 cm in length, separated by a distance of 0.4 cm. The interdigitated pattern comprises parallel fingers from each electrode, interleaved with one another. Each finger in the IDE is 0.04 cm wide, with a gap of 0.02 cm between adjacent fingers. The dimensions of the sensing film are



illustrated in the ESI (Fig. S1).† Film fabrication was carried out using the drop-casting method. A suitable quantity of the synthesized material was dispersed in 5 ml of ethanol by ultrasonication for about 10 minutes. Subsequently, 10 μ l of the obtained dispersion was placed at the centre of the alumina substrate, which had pre-printed interdigital gold electrodes. Gas sensing measurements were conducted using a tabletop gas sensing unit, as reported in the literature.^{19,20} We utilized a piston cylinder with a capacity of 500 ml containing a gas at a concentration of 1000 ppm. To calculate the volume required to achieve a concentration of 1 ppm in a 250 ml chamber, we applied the dilution formula:

$$C_1 V_1 = C_2 V_2$$

where, C_1 = concentration of gas in cylinder = 1000 ppm, V_1 = volume required for 1 ppm, C_2 = concentration that we need *i.e.* 1 ppm, V_2 = volume of chamber *i.e.* 250 ml and using this V_1 = 0.2 ml.

Therefore, to achieve a 1 ppm concentration in a 250 ml chamber, we need 0.25 ml of gas from the cylinder. We then used syringe to collect 0.25 ml of the gas from the 1000 ppm cylinder.

3 Result and discussion

3.1 Structure and morphology

Using X-ray diffraction (XRD), the crystal structure and phase purity of the synthesized materials were evaluated. Fig. 1 illustrates the XRD pattern of the as-prepared MoS₂, and 1T@2H-MoS₂/SnS₂ heterostructures. Fig. 1a shows diffraction peaks at $2\theta = 8.67^\circ, 17.72^\circ, 32.28^\circ, 34.55^\circ,$ and 56.51° , which correspond to the crystal planes (002), (004), (100), (103), and (110) of 1T@2H-MoS₂ (JCPDS, No. 37-1492) further it was also observed that the peak appeared at $2\theta = 8.67^\circ$ *i.e.* (002) plane have significant displacement from 14° to 8.67° , suggesting that the

interlayer spacing in 1T-MoS₂ was greater than in 2H-MoS₂. This shift is likely attributed to the intercalation of ammonium ions.²⁴ Fig. 1b shows an XRD pattern with two separate sets of peaks, one for MoS₂ petaloid nanosheets and the other for SnS₂ nanoparticles. XRD pattern of the as-prepared SnS₂ nanoparticles is illustrated in Fig. S2.† Diffraction peaks at $8.89^\circ, 17.95^\circ,$ and 56.66° , which correspond to the (002), (100), and (110) planes of 1T@2H-MoS₂ in the SnS₂/MoS₂ composite. In addition to this peak at $15.08^\circ, 33.30^\circ$ and 52.50° corresponds to the (001), (101) and (111) crystallographic planes of SnS₂. Findings suggested the integration of both materials into a mixed lattice. Slight shifts in the SnS₂ peaks towards lower angles and the MoS₂ peaks towards higher angles indicated the presence of heterostructures qualities between MoS₂ and SnS₂.²²

X-ray Photoelectron Spectroscopy (XPS) measurements were performed to determine the presence of the 1T@2H-MoS₂ mixed phase and to examine the bonding arrangement, chemical composition, and electronic structure. The entire XPS spectrum of a SnS₂-decorated MoS₂ heterostructures in Fig. 2a shows Sn, Mo, S, and tiny quantities of carbon and oxygen. The presence of oxygen is a result of ambient oxygen adsorption on the composite surface. In Fig. 2b two prominent peaks, corresponding to Mo 3d_{5/2} and Mo 3d_{3/2} at around 228.9 and 231.84 eV, respectively, in the XPS spectra of the Mo 3d region for the SnS₂/1T@2H-MoS₂ composite show the existence of the 1T phase of MoS₂. Two smaller 2H phase peaks with binding energies at 229.9 and 233.2 eV shift by about 1 eV to higher binding energies than 1T-MoS₂.²³⁻²⁶ This change indicates the presence of a trace quantity of semiconducting MoS₂ in the 2H phase. The +6-oxidation state of Mo (Mo⁶⁺ 3d_{3/2}) is responsible for another weak peak at 236.9 eV which shows that pure MoS₂ is partially oxidized. High-resolution S 2p spectra (Fig. 2c) show two peaks at 160.82 and 162.51 eV for S 2p_{1/2} and S 2p_{3/2} of S²⁻ from MoS₂ one unidentified peak were observed at 160.03. The SnS₂/MoS₂ heterostructures is further confirmed by two

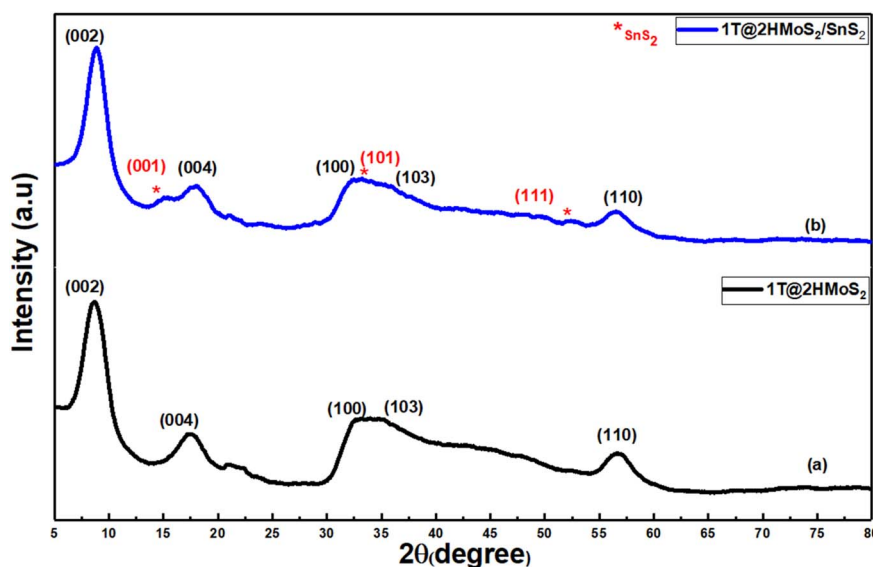


Fig. 1 XRD diffraction patterns of (a) 1T@2H-MoS₂, (b) 1T@2H-MoS₂/SnS₂ heterostructures.



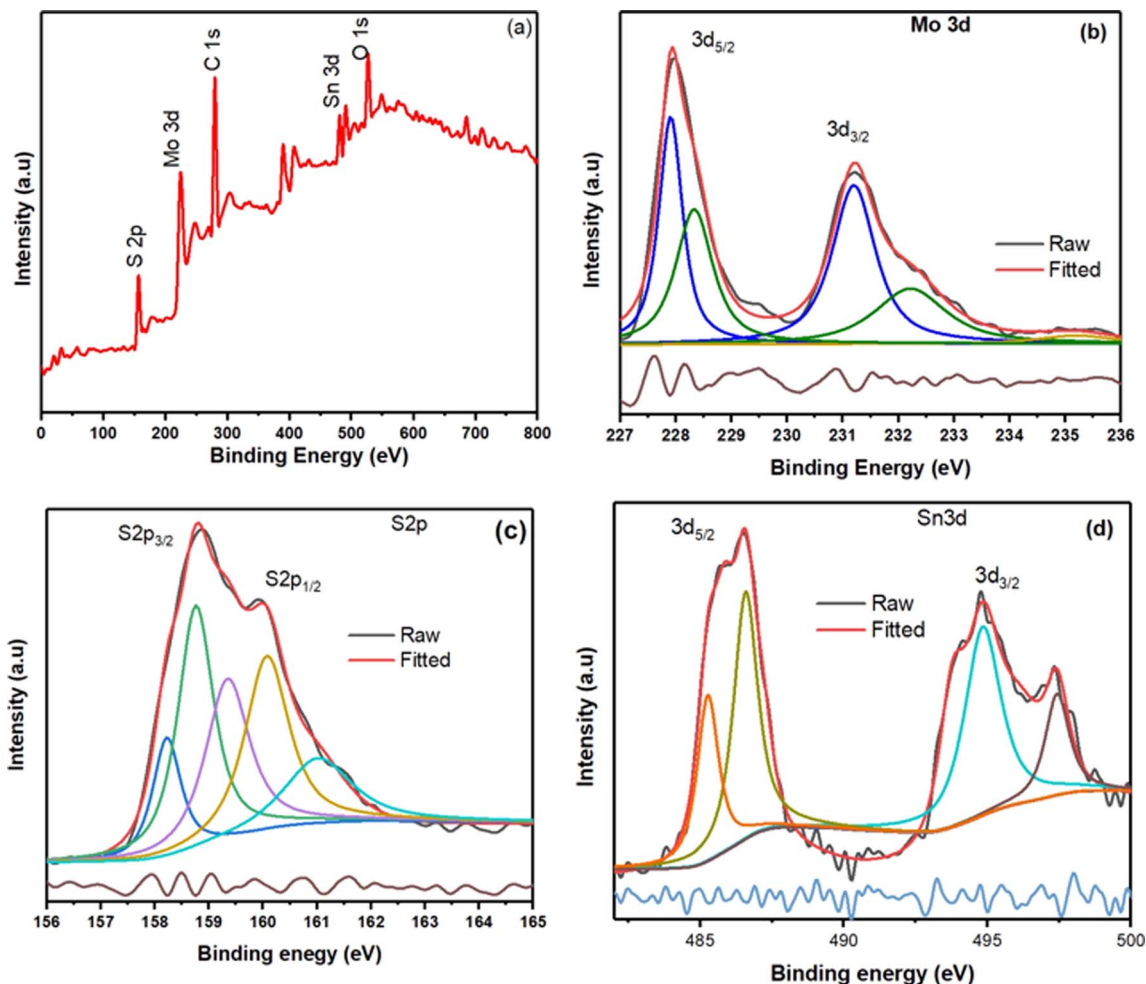


Fig. 2 (a) XPS survey spectra of $\text{SnS}_2/1\text{T}@2\text{H-MoS}_2$ heterostructures, high resolution XPS spectra of (b) Mo 3d. (c) S 2p. (d) Sn 3d.

distinctive peaks in the S 2p spectra of $\text{SnS}_2/\text{MoS}_2$ that are located at 160.40 and 163.47 eV, respectively and correspond to $\text{S}^{2-} 2\text{p}_{1/2}$ and $\text{S}^{2-} 2\text{p}_{3/2}$ from SnS_2 .²⁷ In the Sn 3d region from Fig. 2d, the peaks at 487.5 eV and 496.1 eV are typical Sn(IV) peaks. Interestingly, two more peaks at 485.2 eV and 493.6 eV suggest that Sn has changed in chemical states, which may be due to SnS_2 and MoS_2 interactions.²⁸

The microstructures of $1\text{T}@2\text{H-MoS}_2$ and its heterojunction with SnS_2 were investigated using Field Emission Scanning Electron Microscopy (FESEM). Fig. 3a shows pristine $1\text{T}@2\text{H-MoS}_2$, which features a combination of small nanoparticles and petaloid nanosheets ranging from 1–3 μm in length. High-resolution Fig. 3b reveals small pores on the petaloid nanosheets and width of the nanosheets is around 30–80 nm. Pure SnS_2 , illustrated in Fig. S4(a),[†] consists of closely packed small nanoparticles. The pristine $1\text{T}@2\text{H-MoS}_2$ also displays some nanoparticle morphology, whereas in the composite, as shown in Fig. 3c, spherical SnS_2 particles developed on the surface of the porous $1\text{T}@2\text{H-MoS}_2$ petaloid nanosheets. Fig. 3d shows that the significant aggregation of the SnS_2 nanoparticles, resulting in the formation of a small sheet-like structure. Fig. 3e showed TEM image of $1\text{T}@2\text{H-MoS}_2/\text{SnS}_2$ which consist of SnS_2

and MoS_2 nanoparticles which are around 10 to 20 nm in size, along with MoS_2 nanosheets. Further to confirm the presence of nanoparticles and petaloid nanosheets in $1\text{T}@2\text{H-MoS}_2/\text{SnS}_2$, EDS analysis was also carried out as shown in Fig. 3f which clearly shows presence of sulfur (S), molybdenum (Mo), and tin (Sn) elements. Individual elemental distribution is also shown in Fig. 3g–i for Mo, Sn and S respectively.

3.1.1 BET analysis. The N_2 adsorption–desorption loops for both the pure material and the heterostructures exhibit type IV adsorption isotherms (shown in Fig. 4a). The hysteresis loop (type H1 hysteresis loop) in the adsorption/desorption curve suggests the presence of mesopores in these materials²⁹ with $1\text{T}@2\text{H-MoS}_2/\text{SnS}_2$ showing lower porosity compared to $1\text{T}@2\text{H-MoS}_2$ (shown inset Fig. 4a). This difference is attributed to the introduction of SnS_2 nanoparticles, which provide additional active sites, thereby enhancing the sensing response of the heterostructure compared to the pristine material. The calculated surface area of $1\text{T}@2\text{H-MoS}_2$ and $1\text{T}@2\text{H-MoS}_2/\text{SnS}_2$ is $16.56 \text{ m}^2 \text{ g}^{-1}$ and $25.45 \text{ m}^2 \text{ g}^{-1}$ respectively. The increased surface area of $1\text{T}@2\text{H-MoS}_2/\text{SnS}_2$ can be attributed to interconnected pores and the presence of small SnS_2 nanospheres scattered on the surface of petaloid nanosheets. In addition, the



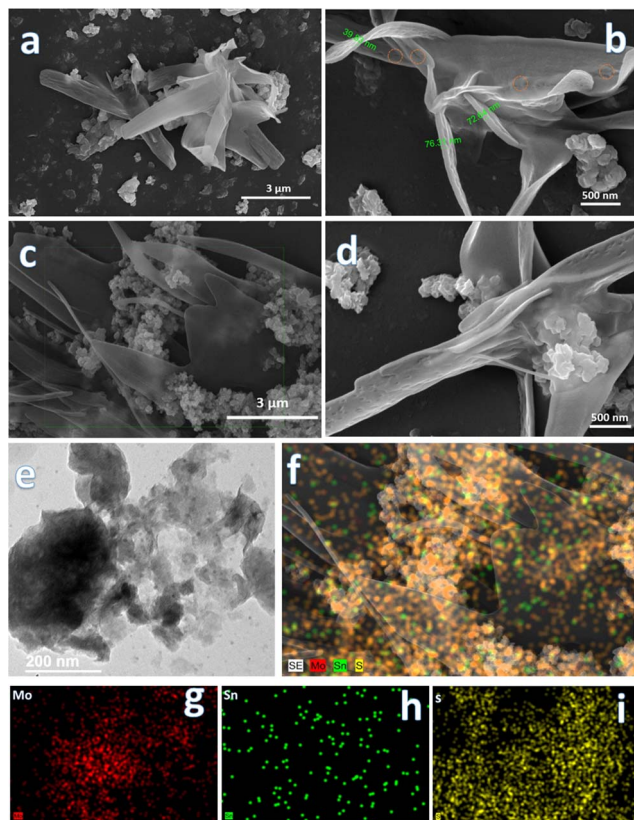


Fig. 3 FESEM images of (a and b) 1T@2H-MoS₂, (c and d) 1T@2H-MoS₂/SnS₂, (e) TEM image 1T@2H-MoS₂/SnS₂, (f–i) EDS mapping of 1T@2H-MoS₂/SnS₂ heterostructures.

heterostructures exhibit larger pore radius and pore volume (7.54 nm and $9.59 \times 10^{-2} \text{ cm}^{-1}$) compared to the pristine 1T@2H-MoS₂ (3 nm and $2.48 \times 10^{-2} \text{ cm}^{-1}$). Thus, adding SnS₂ to the MoS₂ structure improves porosity and surface area, possibly by modifying the materials morphology and pore structure.

3.2 Gas sensing properties

A comprehensive investigation of the detecting capabilities of the synthesized material has been carried out, with pure SnS₂, MoS₂, and 1T@2H-MoS₂/SnS₂ heterostructures in presence NO₂ at room temperature. Fig. 4b–d depicts the sensor response of SnS₂, MoS₂, and 1T@2H-MoS₂/SnS₂ with respect to 1 ppm NO₂ gas over three consecutive cycles. The behaviour of pure SnS₂ is characteristic of n-type semiconductors, exhibiting an increase in resistance as shown in Fig. 4b when NO₂ gas is introduced. Based on the gas sensing properties of 1T@2H-MoS₂, as depicted in Fig. 4c, there is an observed increase in resistance upon exposure to NO₂ gas, indicating that electrons are the majority charge carriers. For instance, in one study conducted by Zong *et al.* a field-effect transistor (FET) gas sensor was developed using MoS₂ that contains a heterophase of the 1T metallic phase and the 2H semiconducting phase. The researchers also observed a decrease in conductivity when this heterophase MoS₂ was exposed to NO₂ gas.³⁰ In the SnS₂/

1T@2H-MoS₂ heterostructures, the interaction with NO₂ gas is affected by the characteristics of both constituent materials. Therefore, heterostructures also exhibit a similar response, with an increase in resistance when exposed to NO₂ as demonstrated in Fig. 4d. The response values for 1 ppm NO₂ gas were observed as 7.49 for SnS₂/1T@2H-MoS₂, 5.4 for MoS₂, and 3.49 for SnS₂. The findings suggest that the addition of SnS₂ nanoparticles considerably boosts the sensitivity of the 1T@2H-MoS₂/SnS₂ at room temperature. Furthermore, after each cycle, the resistance consistently returns to the original baseline without experiencing a significant amount of attenuation. This exceptional repeatability, which ensures constant and consistent output across many sensing cycles, is an essential feature for real-world applications.

The 1T@2H-MoS₂/SnS₂ heterostructures demonstrates remarkable potential for NO₂ gas sensing with its significantly faster response and recovery times compared to individual SnS₂ and MoS₂. In particular, the SnS₂/1T@2H-MoS₂ gas sensor with response times of about 54 seconds and recovery times of about 345 seconds. This performance is higher than MoS₂ (response time 80 s and recovery time 102 s) and the SnS₂ sensor (response and recovery time 75 s and 91 s respectively). The faster response and recovery of the heterostructures may be attributed to a synergistic effect between SnS₂ and MoS₂, leading to enhanced gas adsorption and improved charge transfer dynamics.

Table 1 presents comparison between the 1T@2H-MoS₂/SnS₂ heterostructures and other relevant sensors, including SnS₂-based sensors and MoS₂-based sensors reported in the literature for NO₂ sensing. This comparative analysis aims to provide a comprehensive evaluation of sensing performance, taking into account crucial parameters such as response, response speed, recovery time, and operating temperature. Despite some literature reports indicating sensors with higher responses, our sensor stands out due to its room temperature sensing capability, exceptional selectivity, and stability. This recognition of performance at room temperature, along with its outstanding selectivity and stability, positions our sensor as a promising candidate for widespread use, even in scenarios where other sensors may have demonstrated greater responses according to existing literature.

To get more knowledge of the NO₂ detecting capabilities of the gas sensors, the dynamic response curves were determined at NO₂ concentrations of 1, 5, 10, 25, and 50 ppm, respectively. As shown in Fig. 5a, the response values of the SnS₂/MoS₂ sensors steadily rise with increasing NO₂ concentrations, with the 1T@2H-MoS₂/SnS₂ gas sensor exhibiting the greatest response value over the entire test range. Furthermore, by plotting the logarithm of sensor response on the Y-axis and the logarithm of gas concentration on the X-axis in Fig. 5b linear correlation is observed across the entire concentration range. The SnS₂/1T@2H-MoS₂ sensor demonstrates linear relationship, supported by an *R*-squared value of 0.987.

The selectivity of the 1T@2H-MoS₂/SnS₂ sensor were carried out using 1000 ppm concentration of other interfering gases, such as, ammonia (NH₃), carbon monoxide (CO), ethanol (C₂H₅OH), acetone (CH₃COCH₃), and carbon dioxide (CO₂) and



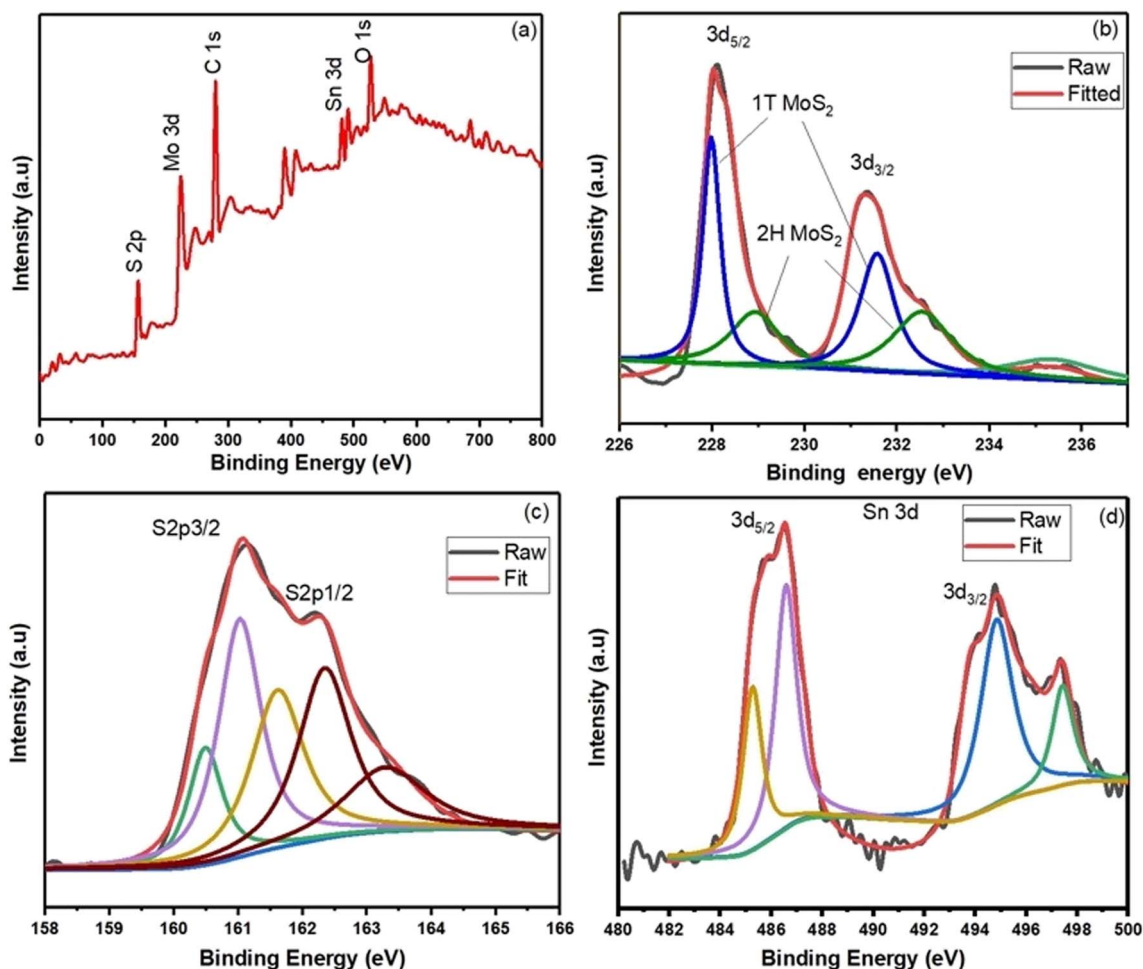


Fig. 4 (a) N_2 adsorption–desorption isotherms of pure 1T@2H-MoS₂ and 1T@2H-MoS₂/SnS₂ heterostructures, sensor response of, (b) SnS₂, (c) 1T@2H-MoS₂, (d) 1T@2H-MoS₂/SnS₂ towards 1 ppm of NO₂ concentration for three consecutive cycles.

100 ppm of NO₂ gas. According to the findings, which are shown in Fig. 5c, resistance increases and decreases with gas exposure are represented by positive and negative response values, respectively. The trend of selectivity for individual gas molecules is determined by their oxidation potential and electrophilicity, which are impacted by partial charge transfer and adsorption energy. Compared to other gases (CO, C₂H₅OH, H₂, C₃H₇OH), the sensor notably shows a greater detecting response to nitrogen-based compounds (*e.g.*, NO₂, NH₃). This conclusion

is consistent with work by Ray *et al.*,³⁸ who used first-principal calculations to demonstrate that the gas adsorption energy of MoS₂ for NO₂ (268.6 MeV) is larger than that for NH₃ (110.1 MeV), supporting the claim that adsorption and selectivity are related. Additionally, because of the physisorbed paramagnetic NO₂ molecules on the surface of SnS₂, which provide a magnetic dipole and greater physical affinity, the device exhibits a higher sensing response to N-based substance.^{39,40} To explore the impact of relative humidity on the sensing

Table 1 Comparison of NO₂ sensing performance between current data and reported data based on MoS₂ and SnS₂

Sensing material	NO ₂ concentration (ppm)	Working temperature (°C)	Response	Response/recovery time (s)	References
Au/SnS ₂ /SnO ₂ heterojunctions	8	80	22.3	174/359.6	31
MoS ₂ /rGO composite	3	160	23%	—	32
3D-MoS ₂ /PbS	100	RT	25%	30/235	33
SnS ₂ /MoS ₂	10	RT	6.2	3.3/25.3	34
SnO ₂ @SnS ₂	0.2	RT	5.5	950/1160	35
SnS ₂ /SnS	0.5	RT	2.5	375/1590	36
SnS ₂ /vertical flakes	50	120	1.64	41/379	37
1T@2H-MoS ₂ /SnS ₂	1	RT	7.49	54/345	This work



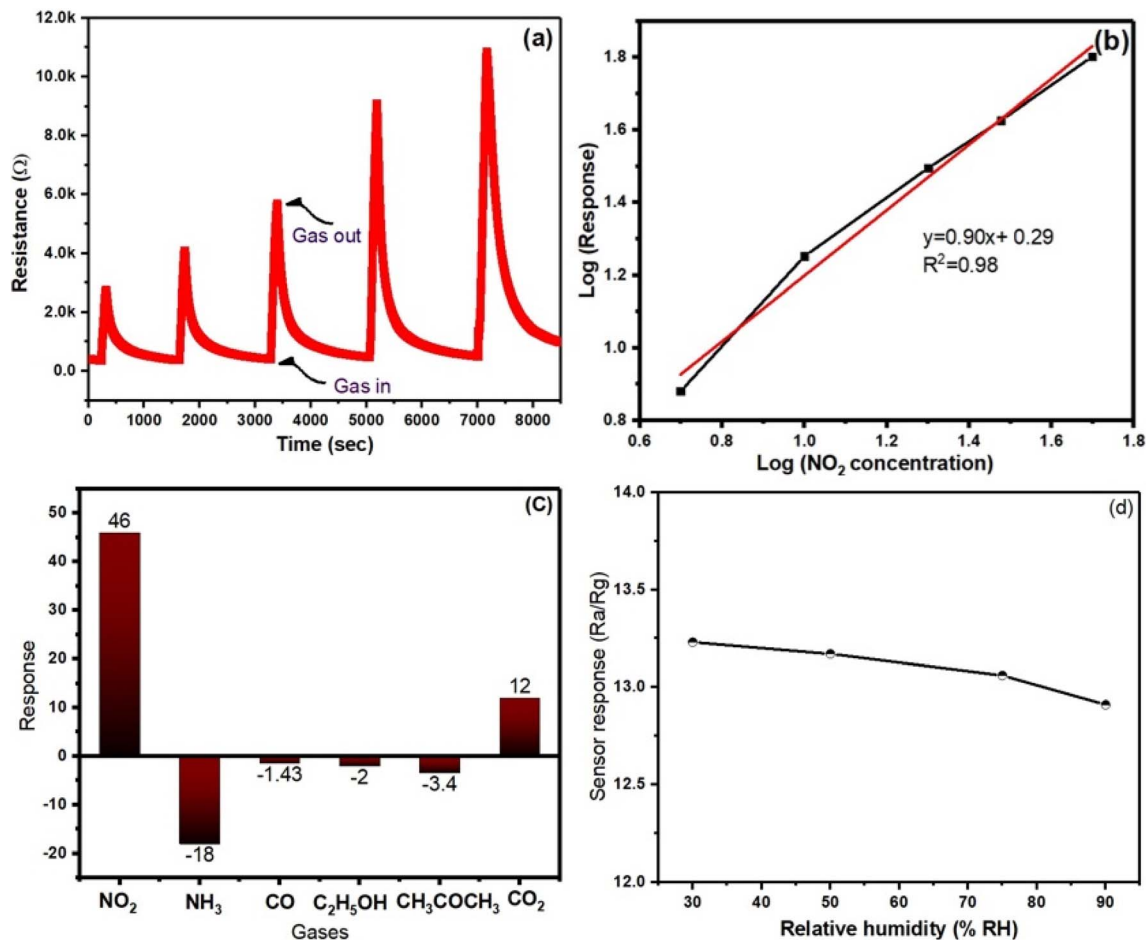


Fig. 5 (a) Response of $\text{SnS}_2/1\text{T@2H-MoS}_2$ towards various concentration of NO_2 . (b) Sensitivity of $\text{SnS}_2/1\text{T@2H-MoS}_2$ towards various concentration of NO_2 . (c) Response of $1\text{T@2H-MoS}_2/\text{SnS}_2$ towards 100 ppm of various gases and 10 ppm of NO_2 at similar condition. (d) Response of $1\text{T@2H-MoS}_2/\text{SnS}_2$ towards 5 ppm NO_2 under different humidity level.

properties of the sensor, 10 ppm NO_2 was exposed to $1\text{T@2H-MoS}_2/\text{SnS}_2$ under varying humidity levels (30% RH, 50% RH, 75% RH, 90% RH). The results from Fig. 5d indicate a slight decrease in the response value (13.25 ± 0.06) as the RH increases. This phenomenon can be attributed to the adsorption of water molecules, which reduces the active sites available for the target gas, consequently leading to a decline in sensor response. Therefore, it is reasonable to conclude that the influence of humidity is minimal, ensuring the reliability of the sensor in practical applications at room temperature. Furthermore, the sensor stability was examined over a twelve-week period. Results indicate that throughout this duration, the $1\text{T@2H-MoS}_2/\text{SnS}_2$ sensor consistently shows sensing response ranging from 16.8 to 17.76 (in Fig. 6a). Therefore, development of a heterojunction between MoS_2 and SnS_2 considerably improves the sensors stability in air.

3.3 Gas sensing mechanism

Gas sensing relies on the surface interactions between the sensor material and the target gas, which alter the carrier density or the surface depletion region, consequently affecting the sensors electrical resistance. Initially, in ambient

conditions, oxygen molecules adsorb onto the sensors surface, capturing electrons from the conduction band to form chemisorbed oxygen species (O_2^-). This process creates a wide electron depletion layer, reducing the charge carrier concentration and increasing the sensor resistance. The presence of chemisorbed oxygen enhances gas sensing response and facilitates charge transfer. When the sensor is exposed to NO_2 , a gas with electrophilic properties, it extracts additional electrons from the surface, forming adsorbed nitrite ions (NO_2^-). This interaction with NO_2 modifies the concentration of adsorbed oxygen and leads to the conversion of NO_2 into NO_3^- . Consequently, electrons trapped in oxygen species are released back to the conduction band, widening the depletion layer and increasing the sensors electrical resistance further.⁴¹

The findings demonstrate that $1\text{T@2H-MoS}_2/\text{SnS}_2$ heterostructures exhibit superior gas response compared to their individual counterparts. Enhanced sensor response due to the combined action of the geometric (as shown in Fig. 7) and electronic aspects. The geometric aspect results in SnS_2 nanoparticles on the MoS_2 surface with more exposed active sites, while the electronic aspect creates a heterostructures at the interface. According to the UV-vis spectra, pure SnS_2 shows



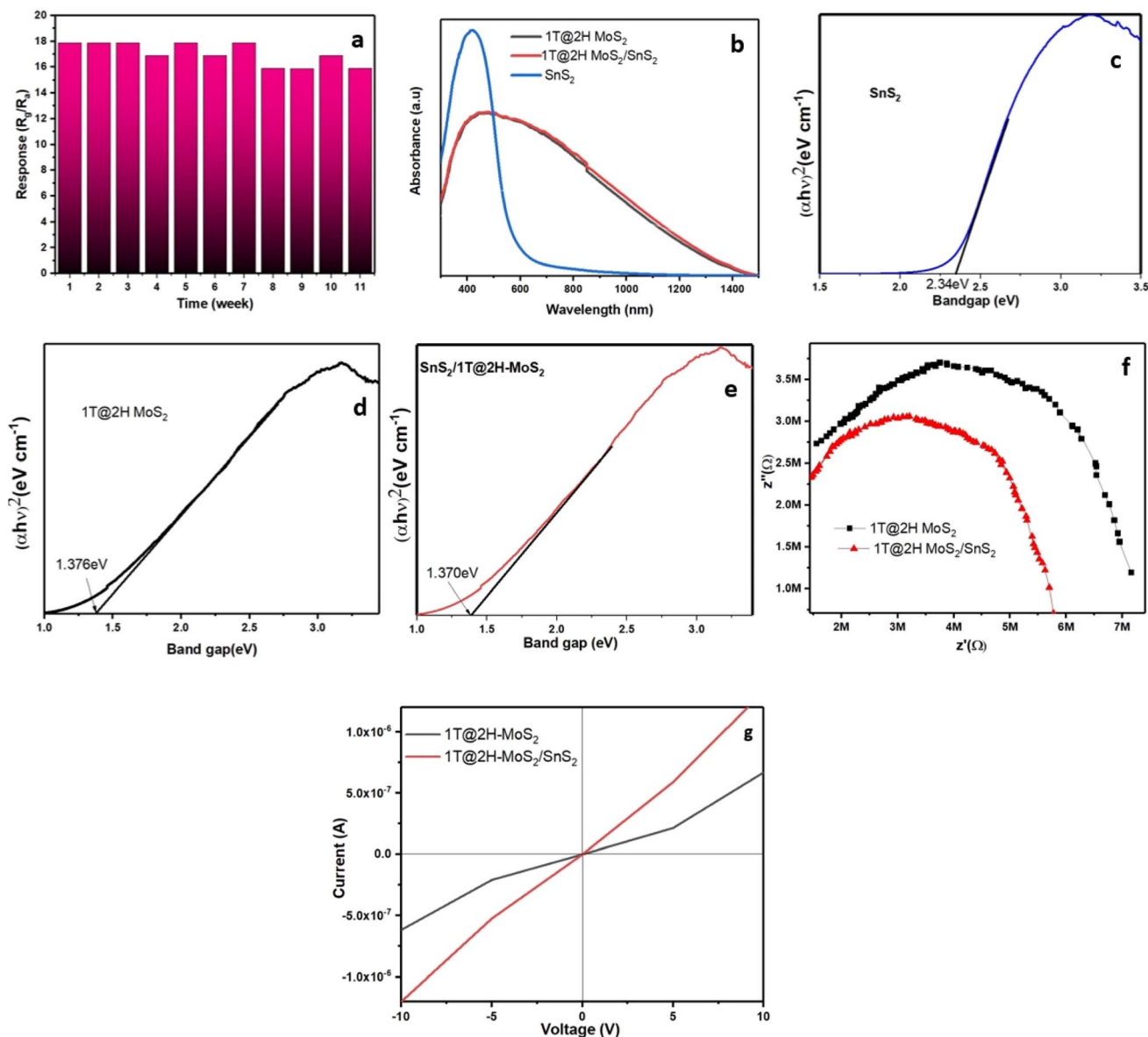


Fig. 6 (a) 1T@2H-MoS₂/SnS₂ sensor stability for 12 weeks at 10 ppm NO₂ in air, (b) UV-vis spectra of SnS₂, 1T@2H-MoS₂ and 1T@2H-MoS₂/SnS₂ Tauc plot of (c) SnS₂, (d) 1T@2H-MoS₂, (e) 1T@2H-MoS₂/SnS₂. (f) EIS plot of 1T@2H-MoS₂ and 1T@2H-MoS₂/SnS₂ in presence of 10 ppm of NO₂. (g) I-V spectra of MoS₂ and SnS₂/1T@2H-MoS₂.

a high absorption, especially in the ultraviolet region, with a notable decrease in absorption beyond 400 nm. SnS₂ coated 1T@-2HMoS₂, on the other hand, exhibits absorption from UV to near-infrared light regions. Which indicates that it is capable of absorbing superior amounts of ultraviolet and visible light. Because of the presence of the 1T-MoS₂ metallic phase in the composite, there is a possibility that the increased absorption in the visible light range is associated with plasmon resonance absorption (illustrate in Fig. 6b). In case of semiconductor band gap is affected by a number of parameters, such as grain size, doping, and composition. A very small change was observed by the introduction of SnS₂ in 1T@2H-MoS₂/SnS₂ than the pristine 1T@2H-MoS₂. The relationships between (αhν)² and photon energy are shown in Fig. 6c-e. The bandgap of SnS₂ and

1T@2H-MoS₂/SnS₂ are 2.34 eV and 1.36 eV respectively. The lower band gap observed in the composite material than the pristine materials. It is often simpler for electrons to go from the valence band to the conduction band when the band gap is narrower. This promotes the growth of chemisorbed oxygen on the surface of SnS₂/1T@2H-MoS₂, ultimately resulting in an elevated reaction rate.⁴²

To further understand the interface charge transfers on the sensor surface, Electrochemical impedance spectroscopy (EIS) with equivalent circuit (illustrate in Fig. S5 and S6†) was employed in presence of 10 ppm of NO₂ gas. Typically, impedance spectra exhibit semi-circles at low-frequency regions (illustrate in Fig. 6e), reflecting the surface charge characteristics of the sensor material. Notably, 1T@2H-MoS₂/SnS₂ displays



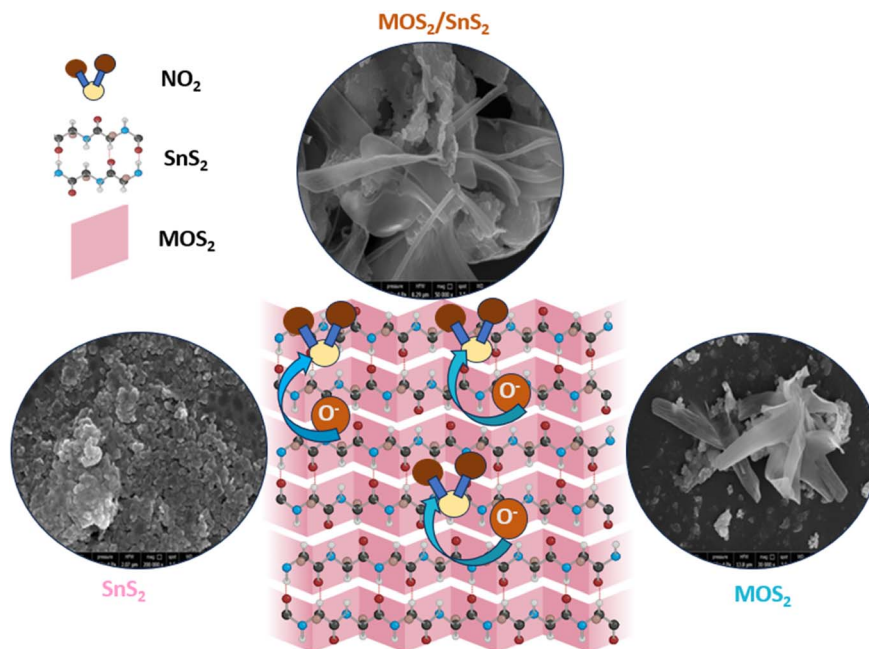


Fig. 7 Sensing mechanism of 1T@2H-MoS₂/SnS₂ heterostructures.

a smaller semicircle indicating lower electron transfer resistance compared to other devices. These results underscore the superior gas sensing properties of 1T@2H-MoS₂/SnS₂ as a sensing material. Consequently, the findings suggest that the resistance and charge transfer resistance of pure 1T@2H-MoS₂ (5.48 MΩ) is higher than those of 1T@2H-MoS₂/SnS₂ (3.72 MΩ). In addition to this *I-V* characteristics of the synthesized materials obtained at room temperature under air are illustrated in Fig. 6g. Both forward and reverse-biased regimes of the corresponding *I-V* curves exhibit linearity, suggesting the ohmic nature of the contact. The average electrical resistances in air are 15 MΩ, and 6 MΩ of 1T@2H-MoS₂ and 1T@2H-MoS₂/SnS₂ respectively. However, a recent study by L. Liu *et al.*³⁴ observed lower conductivity (resulting in higher resistance) in the case of SnS₂/MoS₂-II (SMS-II) compared to the pure SnS₂ and MoS₂ in air at room temperature. Contrary to these findings, we observed higher conductivity in the case of the 1T@2H-MoS₂/SnS₂ heterostructure compared to the pristine materials. This enhanced conductivity in the heterostructure may be attributed to two main factors: firstly, the introduction of SnS₂ into the MoS₂ lattice can act as a dopant, introducing additional charge carriers. Secondly, the presence of SnS₂ can modify the band structure of MoS₂, leading to alterations in the electronic properties of the heterostructure. Furthermore, the addition of SnS₂ to the MoS₂ structure improves porosity and surface area, contributing to the enhanced sensing response. Moreover, SnS₂ and MoS₂ possessing work functions of 4.2–4.5 eV and 5.2–5.4 eV respectively, heterojunctions form between them. Electron flow occurs from SnS₂ to MoS₂ until their Fermi levels are balanced, creating an electron depletion layer on MoS₂ and bending the energy band of SnS₂.³⁴ This leads to a change in electrical resistance in 1T@2H-MoS₂/SnS₂ heterostructures.

Upon exposed to NO₂ gas at the optimal operating temperature, trapped electrons are released back to the conduction band of 1T@2H-MoS₂/SnS₂ heterostructures due to the reaction between adsorbed O₂ species and NO₂ molecules. Consequently, the electrical resistance of 1T@2H-MoS₂/SnS₂ heterostructures significantly decreases, resulting in an enhanced gas sensing response.

3.4 Conclusion

In summary, this study successfully synthesized SnS₂ nanoparticles decorated on petaloid 1T@2H-MoS₂ nanosheets using a simple one-pot hydrothermal method, achieving room temperature NO₂ gas sensing for the first time. This heterostructures demonstrated effective performance as a high-quality NO₂ sensing than the pristine materials. The 1T@2H-MoS₂/SnS₂ sensor exhibited remarkable response characteristics ($R_a/R_g = 46$ at 100 ppm) including a quick response time and efficient recovery. The petaloid nanosheet morphology of 1T@2H-MoS₂ within the heterostructure significantly enhances its performance in gas sensing applications. This structure provides a high surface-to-volume ratio, which is key in increasing the number of active sites available for gas interactions. This feature is important in improving both the sensitivity and selectivity of the heterostructure toward specific gases. In addition to this outstanding gas sensing capabilities can be attributed to the formation of heterojunctions between the 1T@2H-MoS₂ and SnS₂ which provide abundant electron transfer channels between the constituent materials. By exploring ternary composite materials, controlling defects in the surface by functionalization or doping, and tailoring the band gap will give new opportunities for enhancing gas adsorption and reaction efficiency.



Data availability

All data generated or analyzed during this study are included in this article.

Author contributions

Shraddha Hambir: conceptualization of this study, methodology, formal analysis, writing original draft and investigation. Shashikant Shinde: conceptualization and XPS analysis. Shweta Jagtap: conceptualization, supervision, validation, project administration and writing-review and editing and funding acquisition. Chandra Sekhar Rout: supervision, writing-review and editing. H. M. Pathan: supervision, validation, project administration and writing-review and editing. Som Datta Kaushik: supervision, validation, project administration and writing-review and editing.

Conflicts of interest

The authors declare no conflict of interest.

Acknowledgements

Authors would like to thank UGC-DAE-CSR for financial assistance under the project CRS/2021-22/03/546.

References

- H. Khan, A. Zavabeti, Y. Wang, C. J. Harrison, B. J. Carey, M. Mohiuddin, A. F. Chrimes, I. A. De Castro, B. Y. Zhang, Y. M. Sabri, S. K. Bhargava, J. Z. Ou, T. Daeneke, S. P. Russo, Y. Li and K. Kalantar-Zadeh, *Nanoscale*, 2017, **9**, 19162–19175.
- S. Hambir and S. Jagtap, *R. Soc. Open Sci.*, 2023, **10**(4), 221135.
- M. Ikram, L. Liu, Y. Liu, L. Ma, H. Lv, M. Ullah, L. He, H. Wu, R. Wang and K. Shi, *J. Mater. Chem. A*, 2019, **7**, 14602–14612.
- M. Modak, S. Mahajan, M. Shinde, S. Rane and S. Jagtap, *J. Mater. Sci.: Mater. Electron.*, 2022, **33**, 26205–26224.
- T. Aldhafeeri, M.-K. Tran, R. Vrolyk, M. Pope and M. Fowler, *Inventions*, 2020, **28**.
- J. Lin, M. Kilani and G. Mao, *Adv. Mater. Technol.*, 2023, **8**, 2202038.
- S. Mahajan, S. Shinde, M. Modak and S. Jagtap, *Microelectron. Eng.*, 2022, **266**, 111887.
- N. Yi, Z. Cheng, H. Li, L. Yang, J. Zhu, X. Zheng, Y. Chen, Z. Liu, H. Zhu and H. Cheng, *Mater. Today Phys.*, 2020, **15**, 100265.
- H. Long, A. Harley-Trochimczyk, T. Pham, H. Long, A. H. Trochimczyk, T. Pham, Z. Tang, T. Shi, A. Zettl, C. Carraro and M. A. Worsley, *Adv. Funct. Mater.*, 2016, **26**, 5158–5165.
- L. Liu, M. Ikram, L. Ma, X. Zhang and H. Lv, *J. Hazard. Mater.*, 2020, **393**, 122325.
- R. Zhao, T. Wang, M. Zhao, C. Xia, X. Zhao, Y. An and X. Dai, *Phys. Chem. Chem. Phys.*, 2017, **19**, 10470–10480.
- F. Deng, X. Lu, X. Pei, X. Luo, S. Luo and D. D. Dionysiou, *J. Hazard. Mater.*, 2017, **332**, 149–161.
- Z. Yang, C. Su, S. Wang, Y. Han, X. Chen, S. X. Z. Yang, C. Su, S. Wang, Y. Han, X. Chen, S. Xu, Z. Zhou, N. Hu, Y. Su and M. Zeng, *Nanotechnology*, 2020, **31**, 075501.
- S. Ahmad, M. M. Ali Khan and F. Mohammad, *ACS Omega*, 2018, **3**, 9378–9387.
- M. Ikram, L. Liu, Y. Liu, L. Ma, H. Lv, M. Ikram, L. Liu, Y. Liu, L. Ma, H. Lv, M. Ullah, L. He, H. Wu, R. Wang and K. Shi, *J. Mater. Chem. A*, 2019, **7**, 14602–14612.
- Y. Huang, W. Jiao, Z. Chu, G. Ding, M. Yan, X. Zhong and R. Wang, *J. Mater. Chem. C*, 2019, **7**, 8616–8625.
- L. Li, C. Zhang and W. Chen, *Nanoscale*, 2015, **7**, 12133–12142.
- J. B. Liu, J. Y. Hu, C. Liu, Y. M. Tan, X. Peng and Y. Zhang, *Rare Metals*, 2021, **40**, 1536–1544.
- M. Modak and S. Jagtap, *Ceram. Int.*, 2022, **48**, 19978–19989.
- S. Mahajan, S. Shinde, M. Modak and S. Jagtap, *Microelectron. Eng.*, 2022, **266**, 111887.
- M. Acerce, D. Voiry and M. Chhowalla, *Nat. Nanotechnol.*, 2015, **10**, 313–318.
- K. Das, S. Patnaik, S. Mansingh, A. Behra, A. Mohanty, C. Acharya and K. M. Parida, *J. Colloid Interface Sci.*, 2020, **561**, 551–567.
- J. Wu, J. Liu, J. Cui, S. Yao, M. Ihsan-Ul-Haq, N. Mubarak, E. Quattrocchi, F. Ciucci and J. K. Kim, *J. Mater. Chem. A*, 2020, **8**, 2114–2122.
- Z. Liu, Z. Gao, Y. Liu, M. Xia, R. Wang and N. Li, *ACS Appl. Mater. Interfaces*, 2017, **9**, 25291–25297.
- D. Wang, X. Zhang, S. Bao, Z. Zhang, H. Fei and Z. Wu, *J. Mater. Chem. A*, 2017, **5**, 2681–2688.
- M. Acerce, D. Voiry and M. Chhowalla, *Nat. Nanotechnol.*, 2015, **10**, 313–318.
- Y. Li, Z. Yin, G. Ji, Z. Liang, Y. Xue, Y. Guo, J. Tian, X. Wang and H. Cui, *Appl. Catal., B*, 2019, **246**, 12–20.
- Y. Wang, Y. Deng, L. Fan, Y. Zhao, B. Shen, D. Wu, Y. Zhou, C. Dong, M. Xing and J. Zhang, *RSC Adv.*, 2017, **7**, 24064–24069.
- P. Chopade, V. Kashid, N. Jawale, S. Rane, S. Jagtap, A. Kshirsagar and S. Gosavi, *Phys. Chem. Chem. Phys.*, 2023, **25**, 10567.
- B. Zong, Q. Li, X. Chen, C. Liu, L. Li, J. Ruan and S. Mao, *ACS Appl. Mater. Interfaces*, 2020, **12**, 50610–50618.
- B. Huang, Q. Zhu, H. Xu, X. Li, X. Li and X. Li, *Sens. Actuators, B*, 2023, **380**, 133303.
- Z. Wang, T. Zhang, C. Zhao, T. Han, T. Fei, S. Liu and G. Lu, *Sens. Actuators, B*, 2018, **260**, 508–518.
- X. Xin, Y. Zhang, X. Guan, J. Cao, W. Li, X. Long and X. Tan, *ACS Appl. Mater. Interfaces*, 2019, **11**, 9438–9447.
- L. Liu, M. Ikram, L. Ma, X. Zhang, H. Lv, M. Ullah, M. Khan, H. Yu and K. Shi, *J. Hazard. Mater.*, 2020, **393**, 122325.
- D. Liu, Z. Tang and Z. Zhang, *Sens. Actuators, B*, 2020, **324**, 128754.
- Q. Sun, J. Wang, J. Hao, S. Zheng, P. Wan, T. Wang, H. Fang and Y. Wang, *Nanoscale*, 2019, **11**, 13741–13749.
- A. Kumar, N. Sharma, A. Gutal, D. Kumar, G. Zhang, H. Kim, P. Kumar, M. Paranjothy, M. Kumar and M. S. Strano, *ACS Sens.*, 2023, **8**, 1357–1367.



- 38 S. J. Ray, *Sens. Actuators, B*, 2016, **222**, 492–498.
- 39 P. Bharathi, S. Harish, M. Shimomura, M. K. Mohan, J. Archana and M. Navaneethan, *Chemosphere*, 2024, **346**, 140486.
- 40 J. Z. Ou, W. Ge, B. Carey, T. Daeneke, A. Rotbart, W. Shan, Y. Wang, Z. Fu, A. F. Chrimes, W. Wlodarski, S. P. Russo, Y. X. Li and K. Kalantar-Zadeh, *ACS Nano*, 2015, **9**, 10313–10323.
- 41 M. Chen, Z. Wang, D. Han, F. Gu and G. Guo, *J. Phys. Chem. C*, 2011, **115**, 12763–12773.
- 42 S. Hambir and S. Jagtap, *J. Mater. Sci.: Mater. Electron.*, 2023, **34**, 1716.

