


 Cite this: *RSC Adv.*, 2024, 14, 15706

3D-engineered WO₃ microspheres assembled by 2D nanosheets with superior sodium storage capacity†

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Because of the inadequate sodium storage capacity of graphite, the exploration of high-performance SIB anodes is a crucial step forward. Herein, we report the hydrothermally synthesized self-assembled interconnected nanosheets of WO₃ microspheres possessing admirable sodium storage in terms of cycling stability and acceptable rate capability. Benefitting from the interconnected nature of the nanosheets with a hollow interior, the WO₃ microspheres exhibited a high sodiation capacity of 431 mA h g⁻¹ at 100 mA g⁻¹ and an excellent rate performance of 60 mA h g⁻¹ at 500 mA g⁻¹ with an impressive coulombic efficiency of around 99%. Importantly, even after continuous cycling with increasing current densities, a specific capacity as high as 220 mA h g⁻¹ could be recovered at a current density of 50 mA g⁻¹, suggesting excellent sodium storage reversibility.

 Received 8th March 2024
 Accepted 2nd May 2024

DOI: 10.1039/d4ra01800a

rsc.li/rsc-advances

1. Introduction

Concerns about the depletion of fossil-based resources worldwide can only be effectively addressed with the help of renewable, clean energy technologies.^{1–4} Energy storage technologies play an important role in the success of renewable energy integration.^{5,6} Lithium-ion batteries (LIBs), which have been used to power portable electronics for decades, are currently considered the best option for powering the next generation of electric vehicles and stationary applications. However, the limited reserves (0.006 wt% abundance on the earth) and the rising cost of lithium present significant obstacles to the widespread commercial production of LIBs.⁷ In contrast, sodium (the 4th most abundant element with 2.64 wt% abundance on the earth) resources are abundant and accessible everywhere in the world. For large-scale and stationary-storage applications, SIBs with similar electrochemical storage mechanisms have emerged as potential alternative energy storage systems.⁸ Similar to LIBs, the final performance (energy and power density) of SIBs will be dependent on their electrode materials. Therefore, a step forward for SIBs would be the development of new anode materials with higher capacities and better cycling stability.^{5,9} In recent years, the scientific community has shifted its focus to other materials, particularly conversion-type metal-ion batteries with transition metal oxide electrodes, where the different oxidation states of a compound

may offer a high specific capacity for advanced applications.^{10,11} Cobalt oxides, nickel oxides, manganese oxides, ruthenium oxides, and tungsten oxides have been explored extensively owing to the variable oxidation states of metal ions, which facilitate redox reactions and charge storage.^{12–16} WO₃ has recently been investigated as an anode material because of its electrochemical stability, diversity of crystalline phases, non-toxic nature, and high conductivity. From a fundamental standpoint, WO₃ could be an ideal model system to study intercalation-initiated conversion reactions.^{17,18} Moreover, the morphology of the electrode is another key parameter in enhancing the capacity of the material by shortening the ion diffusion length. Various research groups have explored tungsten oxide at the nanometer scale with different morphologies, including nanowires,¹⁸ nanorods,¹⁹ and nanosized irregular particles.²⁰ Santhosha *et al.* synthesized GO–WO₃ *via* the sol–gel method, and for the first time, reported a sodiation capacity of 927 mA h g⁻¹ at 30 mA g⁻¹.²¹ Fanyan Zeng *et al.* synthesized WO_{3–x} nanorods with *in situ*-hybridized oxygen vacancies in nitrogen-doped carbon nanosheets, which delivered specific capacities of 230.8 mA h g⁻¹ and 184.6 mA h g⁻¹ at 20 A g⁻¹ and 30 A g⁻¹, respectively.²² Because of only limited reports with WO₃, the sodium storage mechanism and the effects of microstructures are not yet clearly understood.

In this work, we have synthesized hierarchical 3D WO₃ microspheres composed of very thin nanosheets using a facile

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† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d4ra01800a>

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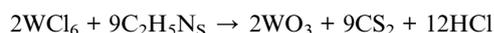


hydrothermal method without any surfactant. The average thickness of the WO₃ nanosheets is ~5 nm. The staggered arrangement of interconnected ultrathin nanosheets forming hollow microspheres can provide more active sites to ease the electrolyte infiltration and enhance the contact between the active materials and electrolytes. Since microspheres have larger surface areas, they can provide more sites for Na⁺ storage, reduce the effects of volume changes, make Na⁺ intercalation into active materials easier, and decrease the diffusion length. While assessing sodium storage behaviour, the as-prepared WO₃ exhibited a high sodiation capacity of 431 mA h g⁻¹ at 100 mA g⁻¹, and excellent rate performance of around 60 mA h g⁻¹ at even 500 mA g⁻¹ with an impressive coulombic efficiency of around 99%.

1.1 Experimental section

Phase pure WO₃ has been synthesized by a facile process using the hydrothermal method. Tungsten hexachloride (0.3 M WCl₆) and thioacetamide (0.6 M C₂H₅NS) were dissolved in 40 ml of deionized water and stirred for around 1 h to get a clear solution. The resultant mixture was transferred into a 50 ml Teflon-lined stainless-steel autoclave and maintained in a preheated oven (180 °C) for 48 h. After naturally cooling to room temperature, the sample was collected, centrifuged, and washed, followed by annealing at 700 °C for 1 hour at a ramp rate of 5°.

The balanced chemical equation is as follows:



This equation indicates that 2 moles of WCl₆ react with 9 moles of C₂H₅NS to produce 2 moles of WO₃, 9 moles of CS₂, and 12 moles of HCl.

The process commences with the nucleation of minute WO₃ nanoparticles (Fig. S1†). This nucleation process initiates sheet formation, wherein individual atoms or molecules congregate to create small clusters of WO₃ nanosheets. Subsequently, microspheres are generated through Ostwald ripening, a phenomenon elucidating the growth of larger particles at the expense of smaller ones within a solution. This growth mechanism arises from the differing solubilities between smaller and larger particles. In the context of tungsten trioxide, larger microspheres are augmented by assimilating material from adjacent smaller ones, resulting in a reduction in the number of smaller particles and a concurrent enlargement of larger ones over time.

Continuing through Ostwald ripening, some microspheres may amalgamate and coalesce, giving rise to larger structures. Under specific conditions, these structures organize into thin, sheet-like formations. The determinants of temperature, concentration, and solvent properties exert a pivotal influence on shaping the final morphology of tungsten trioxide.

1.2 Characterization

The morphological studies of the sample were performed using field-emission scanning electron microscopy (FESEM, Thermo

Scientific Apreo S). The phase and crystal structures were analyzed using the X-ray diffraction pattern obtained using a diffractometer (XRD, BRUKER USA D8 Advance, Davinci). The Raman spectra were obtained using a HORIBA (France) LABRAM HR Evolution instrument and measurements were done using a laser source with an excitation wavelength of 514 nm. The surface chemical states were analyzed by X-ray photoelectron spectroscopy (XPS, physical electronics). High-resolution transmission electron microscopy (HR-TEM, JEOL Japan, JEM-2100 Plus) studies were also performed to discern the morphology and crystal structure at the nanoscale using high-resolution and the selected area electron diffraction pattern (SAED).

1.3 Electrochemical characterization

Cyclic voltammetry (CV) and galvanostatic charge–discharge (GCD) measurements were performed using a biologic electrochemical workstation in the potential range of 0.01–3.0 V. The same workstation was also used to do the electrochemical impedance spectroscopic (EIS) measurements. CR 2032-coin cells were assembled using glass fiber separators. Sodium was used as a counter electrode and 1 mol L⁻¹ NaClO₄ in ethylene carbonate (EC), ethyl methyl carbonate (EMC), and dimethyl carbonate (DMC) (1 : 1 : 1 by volume) was used as the electrolyte. The electrodes were fabricated using a slurry-coating method with the SP carbon and polyvinylidene fluoride (PVDF) in *N*-methyl-2-pyrrolidinone solvent in the weight ratio of 7.5 : 1.5 : 1.

2. Results and discussion

Fig. 1(a) depicts a standard Rietveld refinement plot for the X-ray diffraction pattern of the monoclinic phase of the WO₃ sample. The experimental diffraction pattern is represented by symbols, while the refined one is depicted by a continuous line. The clear sharpness observed in the diffraction peaks not only indicates the high-quality crystallization of the synthesized WO₃ but also the absence of any impurity peaks in the pattern. All the peaks in this spectrum can be perfectly indexed to a monoclinic phase of WO₃. Fig. 2(b) depicts the crystal structure of monoclinic WO₃ in a 3D view. The estimated lattice constants of *a* = 7.3008 Å, *b* = 7.5389 Å, and *c* = 7.6896 Å (JCPDS card # 75-2187) from the diffraction pattern match well with the reported structure of WO₃.²³

Using XPS, the chemical composition and oxidation states of the elements were determined. The XPS survey scan spectrum (Fig. 1(c)) shows the presence of W, O, and C. The deconvoluted O 1s core-level spectra are shown in Fig. 1(d). The W 4f (Fig. 1(e)) spectrum is proficiently divided into the tungsten orbitals W 4f_{5/2} and W 4f_{7/2}, which have binding energies of 37.31 eV and 35.21 eV, respectively, confirming the presence of the W⁶⁺ oxidation state of WO₃.^{24,25} The split between the W 4f_{7/2} and W 4f_{5/2} core levels is 2.1 eV, indicating a +6 oxidation state of W in WO₃. The Raman spectrum of the as-synthesized WO₃ structure is depicted in Fig. 1(f), showing the four major characteristic peaks in the 100 cm⁻¹ to 900 cm⁻¹ region. Two strong characteristic bands at 804 and 716 cm⁻¹ represent the W–O–W stretching vibrations of the bridging oxygen from the



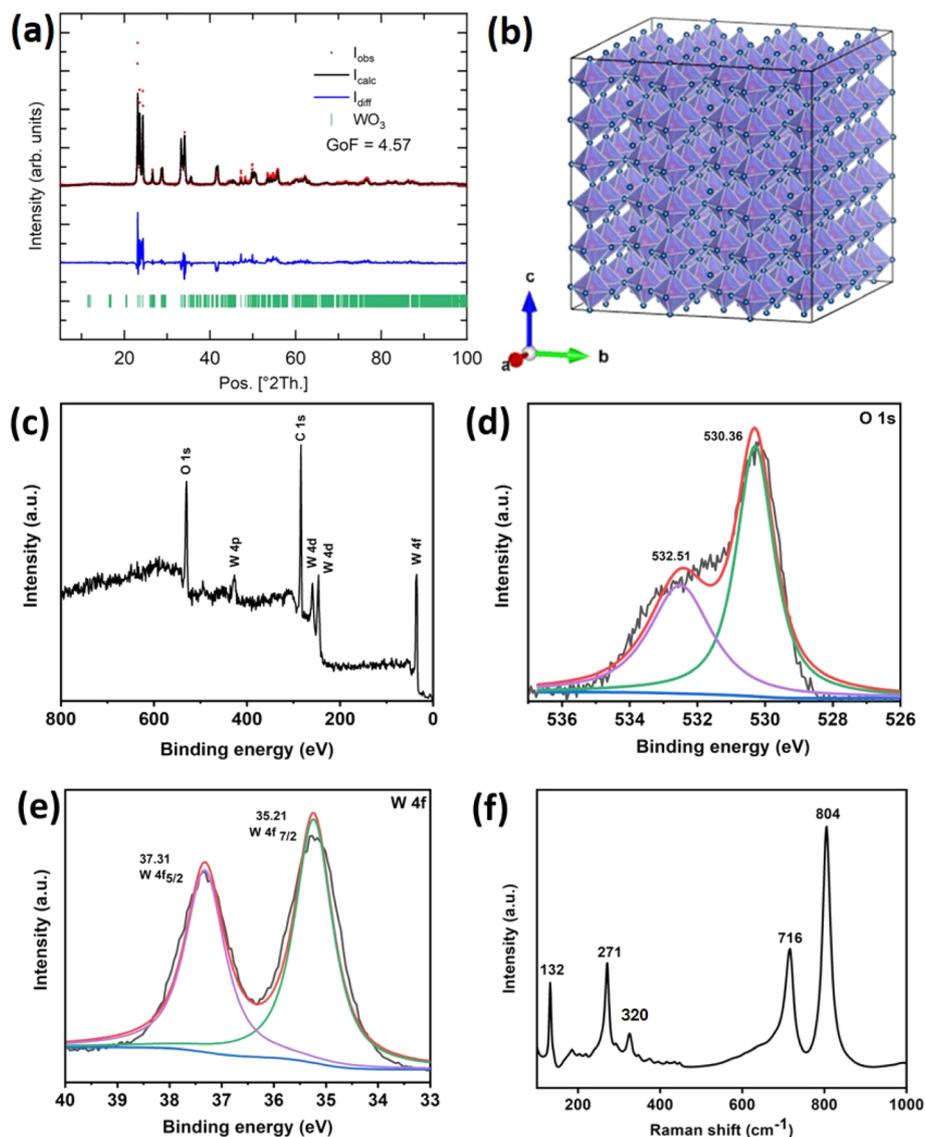


Fig. 1 (a) A standard refinement plot illustrating the X-ray diffraction pattern of WO_3 . (b) A polyhedral representation of monoclinic WO_3 . (c) XPS survey spectrum of tungsten oxide (WO_3). High-resolution XPS spectra of (d) oxygen (O 1s) and (e) tungsten (W 4f). (f) Raman spectrum of as-synthesized WO_3 .

monoclinic crystalline phase of WO_3 . The bands at 320 cm^{-1} , 271 cm^{-1} and 132 cm^{-1} represent the bending vibrations of the (O–W–O) mode.^{26,27}

The FESEM images (Fig. 2(a) and (b)) reveal that each microsphere consists of many interconnected nanosheets, which form a spherulitic microspherical hierarchical structure with a diameter ranging between 1.0 and 1.5 μm . The sheet-like internal microstructure remarkably increases the surface area, thus enhancing ionic diffusion. The nanosheets show hyperbranched flower structures due to subsequent nucleation. The interconnected nanosheets provide sufficient ionic channels for the electrolyte to establish large accessible internal surfaces for an electrochemical redox reaction, resulting in a high specific capacity.²⁸

The very thin and transparent nanosheets were observed at the surface of the micro-sphere by transmission electron microscopy (TEM, Fig. 2(c)). The hollow nature of the

microspheres was visible through a fracture that may have occurred during sample preparation by ultrasonication. Fig. 2(d) and (e) presents high-resolution TEM (HRTEM) displaying a lattice spacing of 0.36 nm corresponding to the (200) plane of WO_3 , which is in line with the XRD. The EDX spectra (Fig. 2(f)) also confirmed the presence of tungsten and oxygen in the sample.

N_2 adsorption–desorption measurements were utilized to assess the specific surface areas and pore structures of WO_3 , as depicted in Fig. S2(a and b).† According to the IUPAC classification, the N_2 adsorption–desorption isotherms exhibited pseudo-type IV hysteresis. The BET specific surface area of WO_3 was determined to be $12.240\text{ m}^2\text{ g}^{-1}$. Fig. S2(b)† illustrates pore-size distribution plots obtained using the BJH method for pristine WO_3 . The average BJH pore volume of pristine WO_3 was measured at $0.045\text{ cm}^3\text{ g}^{-1}$, indicating a broad pore-size



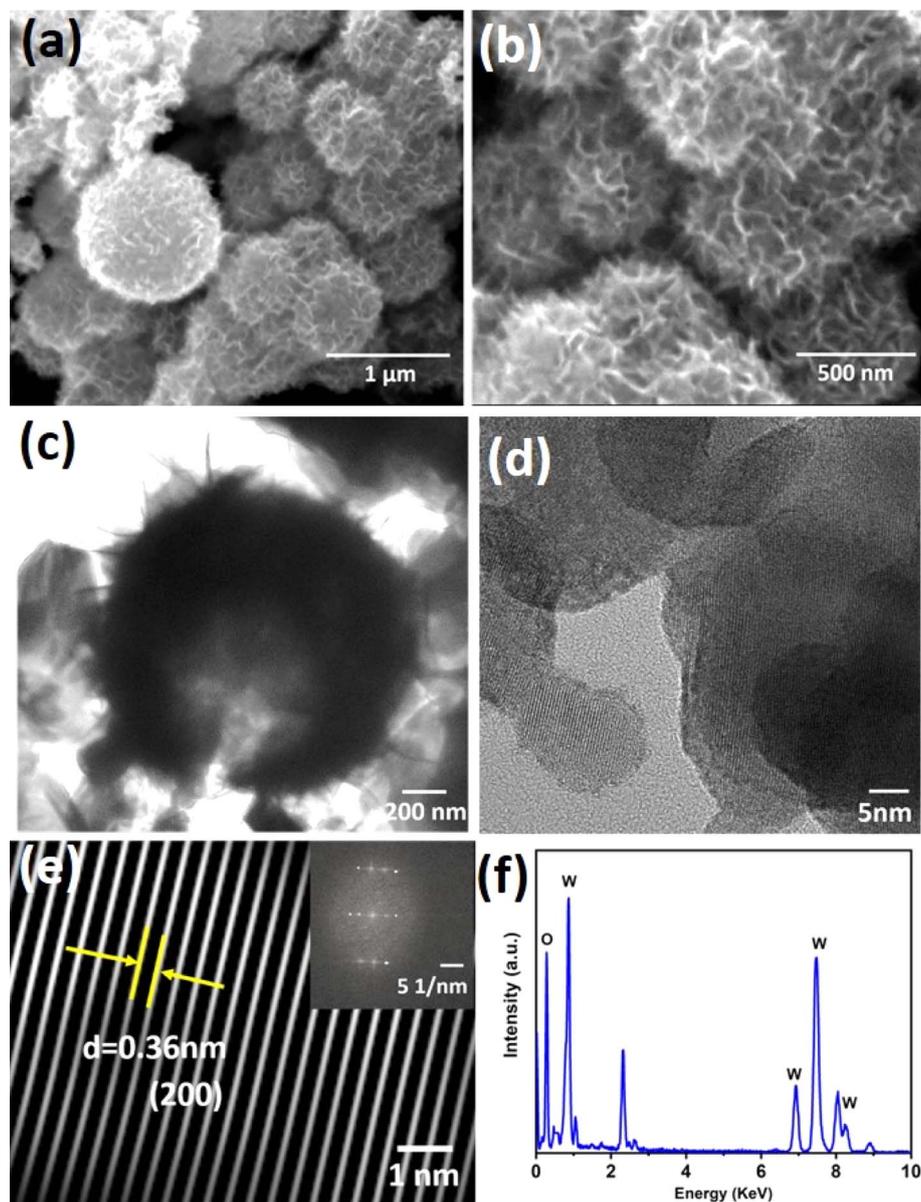


Fig. 2 FESEM image of WO_3 (a) at low magnification and (b) at high magnification. (c) Bright-field (d) high-resolution TEM images of WO_3 . (e) A filtered inverse FFT image showing the lattice plane. The inset shows the FFT (fast Fourier transform) pattern. (f) EDX (energy dispersive X-ray analysis) of the corresponding HRTEM image.

distribution ranging from 19 nm to 30 nm, suggestive of the presence of macropores. The prepared WO_3 microspheres demonstrated a high specific surface area and a favorable pore-size distribution.

The electrochemical studies were conducted *via* CV and GCD in a potential window of 0.01 V to 3 V (Fig. 3). The cyclic voltammogram for the first three consecutive cycles at a scan rate of 0.2 mV s^{-1} is presented in Fig. 3(a). The broad peak from $\sim 0.86 \text{ V}$ to 0.29 V in the first cycle during the reduction process corresponds to the insertion of Na^+ into the interlayers of WO_3 , followed by the peak at 0.62 V associated with the conversion of WO_3 into elemental W and the irreversible SEI layer formation, as well as the insertion of Na^+ (sodium ion) to form metallic W and Na_2O . During the oxidation process, a broad peak was observed at

$\sim 0.91 \text{ V}$ due to the extraction of Na^+ as well as the formation of W to WO_3 . From the 2nd cycle onwards, the reduction and oxidation peaks appeared at 0.8 V and 0.7 V , respectively, denoting the conversion reaction between WO and W.



The charge and discharge profile for a few specific cycles at 100 mA g^{-1} is shown in Fig. 3(b). The first discharge and charge capacities were $431.3 \text{ mA h g}^{-1}$ and $222.31 \text{ mA h g}^{-1}$, respectively. A huge capacity deprivation with an initial coulombic efficiency of 51.8% was observed, due to the formation of a solid electrolyte interphase (SEI) on the surface of the electrode by the



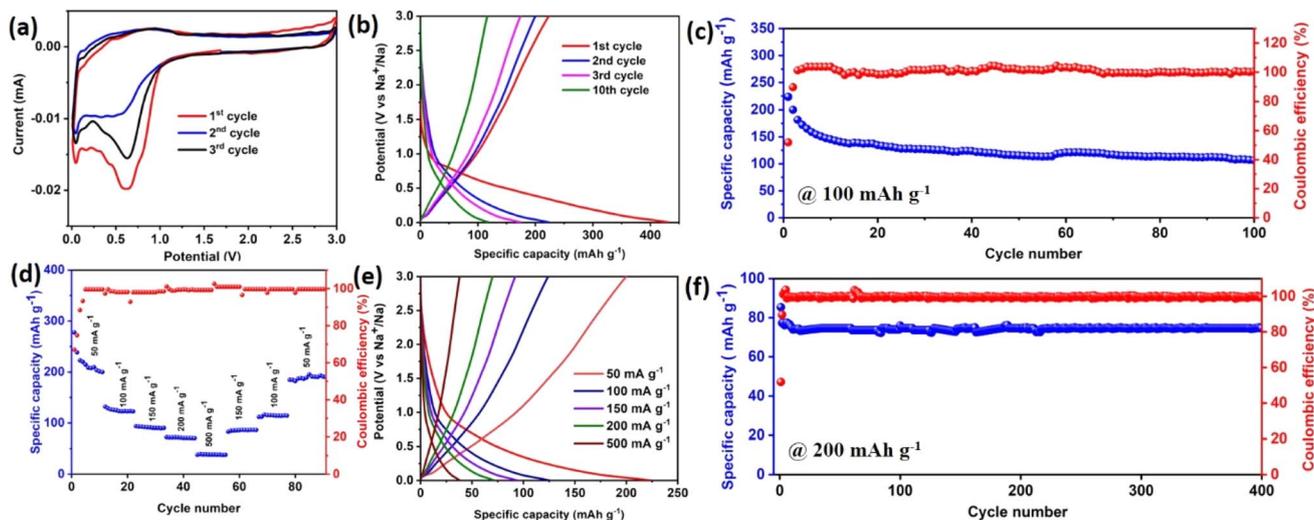


Fig. 3 (a) Cyclic voltammogram (CV) at a sweep rate of 0.2 mV s^{-1} . (b) Galvanostatic charge–discharge profiles of different cycles at a current density of 100 mA g^{-1} . (c) Long cycling and coulombic efficiency at a current density of 100 mA g^{-1} . (d) Rate performance and coulombic efficiency of WO_3 . (e) The corresponding charge–discharge profiles at different current densities are shown. (f) Long-term cycling at a current density of 200 mA h g^{-1} for up to 400 cycles.

decomposition of the electrolyte.²⁹ The theoretical capacities of carbon materials were less than 250 mA h g^{-1} .³⁰ The as-synthesized WO_3 showed a better experimental capacity as compared to carbon materials. After the 10th cycle, the capacity fading was stable up to the 100th cycle. The coulombic efficiency was also enhanced with cycling and stabilized, showing $\sim 100\%$.

Besides the satisfactory specific capacity and superior cycling stability, the WO_3 electrode also showed satisfactory rate capability. As displayed in Fig. 3(c and d), on increasing current densities from 50 to 100, 150, 200, and 500 mA g^{-1} , reversible capacities of 222.5, 124.0, 92.8, 70.0 and 40 mA h g^{-1} were achieved, respectively. Interestingly, even after prolonged cycling at increasing current densities, a specific capacity of up to 220 mA h g^{-1} was recovered at a current density of 50 mA g^{-1} , indicating outstanding sodium storage reversibility. A long cycling test was performed at a current density of 200 mA h g^{-1} for up to 400 cycles as shown in Fig. 3(e). The test showed a negligible capacity loss. The capacity observed after 400 cycles was $\sim 75 \text{ mA h g}^{-1}$ with a coulombic efficiency of 100%.

The kinetic properties were evaluated by obtaining EIS of the sample before and after cycling up to 100 cycles at 100 mA g^{-1} . The Nyquist plots are shown in Fig. S3(a)† and were fitted by a Randles circuit (Fig. S4(a) and (b)†), and the resultant fitting parameters are provided in Table 1. The estimated charge transfer resistance R_{ct} values for the WO_3 electrode before and after 100

cycles were 109Ω and 182Ω , respectively. Using the following equation, we determined the diffusion coefficient of sodium ions:

$$D = \frac{R^2 T^2}{2A^2 n^4 F^4 C^2 \sigma^2} \quad (2)$$

where T , A and n are the absolute temperatures, area of the electrode and the number of electrons during charge/discharge, respectively. R is the gas constant, F is the Faraday constant, C is the sodium ion concentration of sodium ions, and σ is the Warburg factor.

The Warburg factor is denoted by σ and the formula is $\sigma = Z'/\omega^{-1/2}$. It can be understood in terms of the impedance of the diffusion barrier layer and the semi-infinite diffusion impedance in the lower frequency slope that comes after the semi-circle. The value of σ was found to be inversely proportional to the Na^+ diffusion coefficient.³¹ The slope values (from Fig. S1(b)†) before and after 100 cycles were 4.78 and 8.14, respectively. The formula above could be used to determine the diffusion coefficients of Na ions, which were 5.01×10^{-12} and 1.72×10^{-12} , respectively.

The reaction kinetics of WO_3 was investigated *via* scanning cyclic voltammetry at different sweep rates ranging from 0.2– 0.8 mV s^{-1} , shown in Fig. 4(a). As the scan rate increased, these CVs showed the same contours and uninterrupted peak shifts. Therefore, the peak locations were similar to the CV results shown in Fig. 4(a) after the first cycle, indicating low polarization and strong electrochemical behavior. Eqn (3) and (4) were used to evaluate the charge storage contribution (diffusion-controlled contribution and capacitive contribution):³²

$$i = av^b \quad (3)$$

$$\log i = b \log v + \log a \quad (4)$$

Table 1 The results from EIS and diffusion coefficients before and after cycling, obtained from the Nyquist plot

	$R_1 (\Omega)$	$R_2 (\Omega)$	$R_3 (\Omega)$	Slope (σ)	$D (\text{cm}^2 \text{ s}^{-1})$
Before cycling	6.15	109.55	—	4.78	5.01×10^{-12}
After cycling	32	182	91	8.14	1.72×10^{-12}



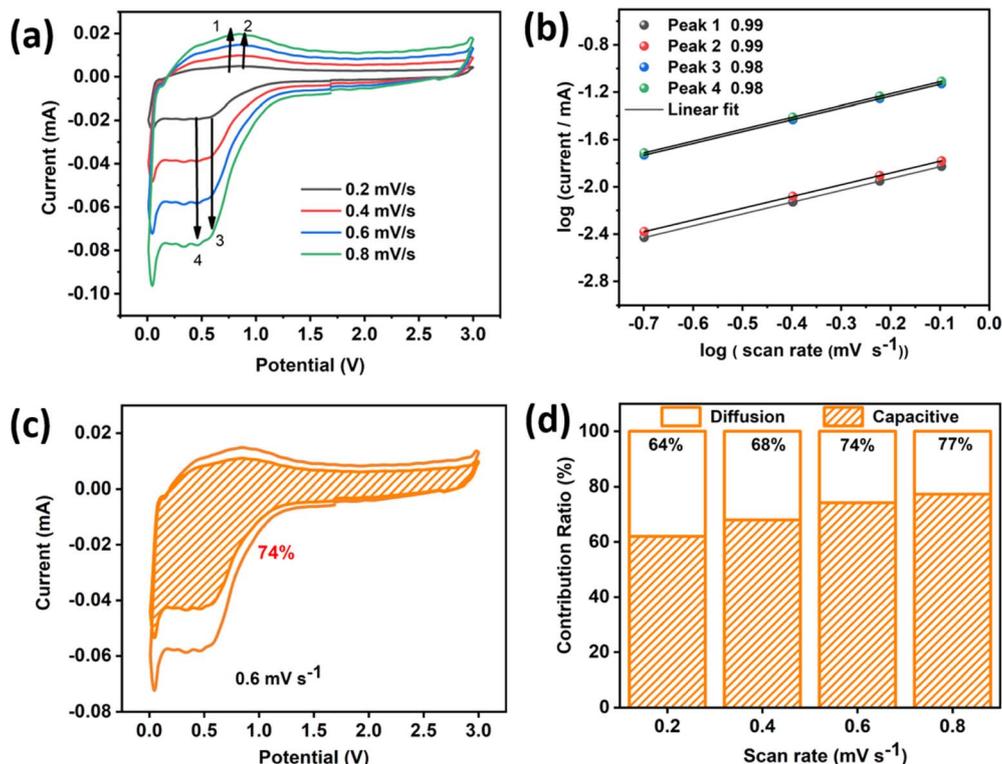


Fig. 4 (a) CV plots at various scan rates ranging from 0.2 to 0.8 mV s⁻¹. (b) Log(peak current, i) vs. log(scan rate, v), (c) the pseudocapacitive contribution (orange) at 0.6 mV s⁻¹. (d) Storage contribution for various scan rates.

where i is the peak current and v is the scan rate, respectively; a and b are the changeable parameters.

By the $\log(v)$ vs. $\log(i)$ plot, the value of b can be calculated, which shows the capacitive and ion-diffusion contribution. In a capacitive-controlled process, $b = 1.0$, whereas in the ion diffusion process governing the electrochemical reaction, $b = 0.5$. When b falls within the range of 0.5 to 1.0, both capacitive and intercalation processes contribute to the storage of charge. For the as-synthesized WO₃, the calculated b values were 0.99, 0.99, 0.98, and 0.98, for peaks 1, 2, 3, and 4, respectively as shown in Fig. 4(b). The charge storage contribution of WO₃ can be further estimated by the following eqn (5):

$$i = k_1v + k_2v^{1/2} \quad (5)$$

where, k_1v corresponds to the capacitive contribution and $k_2v^{1/2}$ corresponds to the diffusive contribution. As illustrated in Fig. 4(c), the capacitive process of WO₃ contributes 74% to the total capacity at the scan rate of 0.6 mV s⁻¹. This result demonstrates that the material provides specific capacitance predominantly through a capacitive contribution. The variation of the capacitive and insertion contribution with respect to scan rates was also calculated and the changes are depicted in Fig. 4(d).

To investigate the phase change and structural changes that occur in the material upon charge–discharge cycling, the cycled cell was opened and the anode material was recovered. The cell was cycled up to 100 cycles at a current density of 100 mA g⁻¹ before the sample was recovered. This post-electrochemical sample was analyzed using XRD and SEM. The XRD spectrum

of the cycled anode material is shown in Fig. 5(a). The presence of tungsten trioxide and the absence of elemental tungsten in the XRD spectra indicates reversibility and the peaks for sodium oxide may have originated from the conversion reaction after the first cycle.

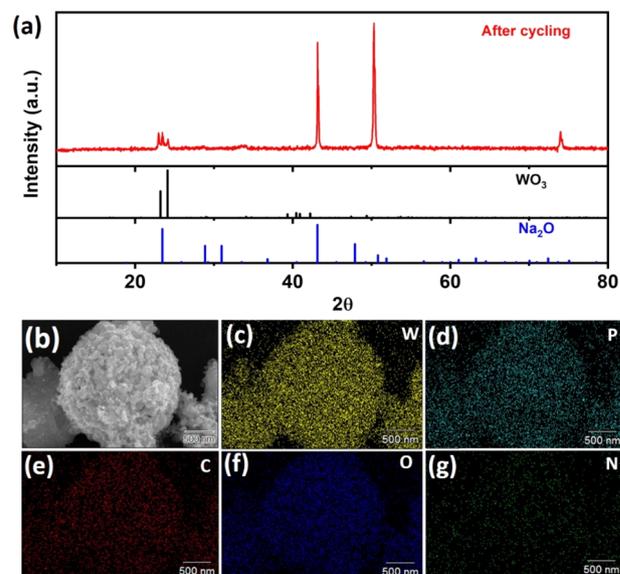


Fig. 5 (a) XRD spectra of the electrode after cycling. (b) HRSEM image of the electrode material after cycling. (c) The corresponding elemental mapping images of W (tungsten), (d) P (phosphorus), (e) C (carbon), (f) O (oxygen), (g) N (nitrogen).



The SEM image and elemental mapping of the cycled sample are shown in Fig. 5(b–g). The SEM image did not reveal the nanosheets, which were likely obscured by PVDF, activated carbon, and dissociated electrolyte products on the surface. Despite this, the resilient spherical shape persisted, underscoring impressive structural stability. For further clarification of the microstructural changes in the post-cycling electrode material, two SEM images are provided in Fig. S5(a and b).† Elemental mapping indicates the presence of carbon, oxygen, phosphorus, and nitrogen, hinting at the accumulation of electrolyte, PVDF, and conductive carbon on the surface. The detailed observation of the transformation of the morphology and structure of the cycled anode proves that the sample's reversibility and structural stability are extremely good, enabling it to withstand continuous sodiation and desodiation.

3. Conclusion

We have successfully synthesized WO_3 microspheres having a complex framework, made up of a nanosheet network. The as-synthesized anode material exhibited excellent electrochemical properties for Na-ion insertion. In response to ramping up current densities from 50 to 100, 150, 200, and 500 mA g^{-1} , reversible capacities of 222.5, 124.0, 92.8, 70.0 and 40 mA h g^{-1} were achieved, respectively. Significantly, after the uninterrupted cycling with increasing current densities, a specific capacity as high as 220 mA h g^{-1} was regained at a current density of 50 mA g^{-1} , claiming exquisite sodium storage reversibility. The reaction kinetics of WO_3 was also investigated, and the capacitive process of WO_3 contributed 74% to the entire capacity at the sweep rate of 0.6 mV s^{-1} . Such a high value implies that the self-assembled WO_3 microspheres offer ample sites for surface storage by the capacitive mechanism, which is reflected in rate capability, and enough space to buffer the volume change, which is correlated with the cycling stability.

Conflicts of interest

The authors declare no potential conflicts of interest with respect to the research, authorship, and/or publication of this article.

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