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## Pyrolysis mechanism and pyrolysis kinetics of yellow wine lees

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Yellow wine lees, a by-product produced while brewing yellow wine, can be a helpful biomass resource through pyrolysis. However, there have been very few studies on the pyrolysis of yellow wine lees. The kinetics and mechanism of pyrolysis in yellow wine lees were explored through an extensive study of their chemical and elemental composition. The pyrolysis mechanism of yellow wine lees was further studied using thermogravimetric analysis (TGA) from 30 °C to 900 °C. The TG/DTG analysis showed that yellow wine lees thermally decomposed mainly between 121 °C and 500 °C. The maximum decomposition was observed between 218 °C and 326 °C, with a clear peak at 298 °C. Upon analyzing the 3D-FTIR results, the gas phase products at this stage primarily included inorganic molecules like CO<sub>2</sub>, H<sub>2</sub>O, and CH<sub>4</sub>, along with organic compounds such as esters, alcohols, phenols, amines, ethers, aldehydes, ketones, and acids. The Maillard reaction and ketosis decarboxylation primarily occurred in proteins (amino acids) and carbohydrates. The pyrolysis kinetics of yellow wine lees were analyzed utilizing the distributed activation energy model (DAEM). The results of DAEM were simultaneously verified using the Flynn–Wall–Ozawa (FWO) method. The findings indicated that the pyrolysis of yellow wine lees conforms to the assumptions of infinite parallel reactions and activation energy distribution. As the conversion rate increased during pyrolysis, the activation energy of yellow wine lees initially increased to 210–220 kJ mol<sup>-1</sup>, then stabilized at 190–200 kJ mol<sup>-1</sup> and rapidly decreased to approximately 100 kJ mol<sup>-1</sup>. This study offers a theoretical basis for the application of yellow wine lees using pyrolysis.

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### Introduction

Fossil energy is essential for human existence, but it faces challenges such as limited availability, non-renewability, and environmental contamination resulting from its use. Therefore, it is crucial to thoroughly examine and develop renewable biomass resources as an alternative to fossil fuels.

Yellow wine is one of the three significant brewed wines in China. According to the “2020 National Employment Economic Indicators” issued by the China Liquor Industry Association in 2021, China’s yearly production of yellow wine could reach 2.83 million tons in 2020.<sup>1</sup> The amount of grain needed to make 2 tons of wine from 1 kg of rice is 1.415 million tons. The yield rate of rice lees is determined by the consumption of grains, which ranges from 20 to 30%.<sup>2</sup> Consequently, the production of rice lees can range from 283 000 to 424 500 tons. Yellow wine lees are a by-product generated while brewing yellow wine and can also be considered a useful biomass resource. Yellow wine

lees have a high concentration of protein, amino acids, carbohydrates, fat, cellulose, and other substances, making it a valuable resource with extensive applications.<sup>3</sup> The applications of yellow wine lees include animal feed,<sup>4,5</sup> condiments,<sup>6</sup> lees grain biscuits,<sup>7–9</sup> cosmetics,<sup>10</sup> bacterial cellulose,<sup>11</sup> adsorption functional materials,<sup>12</sup> and energy materials.<sup>13</sup> Energy materials can simultaneously address the strain on fossil energy resources and the ecological harm caused by yellow wine lees waste. Yellow wine lees show potential for efficient utilization, making them a valuable subject of investigation.

Pyrolysis technology is a crucial approach to the regeneration and use of biomass.<sup>14</sup> Biomass pyrolysis is a thermal process that breaks down biomass into combustible gas, liquid bio-oil, and solid biomass charcoal under completely anaerobic or anoxic conditions.<sup>15</sup> Investigating the thermal decomposition mechanism and thermal analysis kinetics of biomass holds substantial importance in biomass utilization and regeneration. However, there have been very few studies on the pyrolysis of yellow wine lees.

Considering the conditions listed above, examining the pyrolysis process and thermal analysis kinetics of yellow wine lees is crucial. A thermogravimetric analyzer (TGA) is frequently used to quantify the reduction in mass of materials at varying temperatures. The resulting mass loss curve, produced from

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TGA, is used to investigate the apparent kinetics of the pyrolysis process. Branca<sup>16</sup> *et al.* studied the pyrolysis kinetics of dried liquor lees through TGA. Lin<sup>17</sup> *et al.* employed TGA to investigate the co-pyrolysis kinetics of sewage sludge and oil shale. Sait<sup>18</sup> *et al.* compared the apparent kinetics of date palm pyrolysis and combustion using TGA.

Furthermore, it is essential to comprehend the changes in gas phase products during the pyrolysis process to investigate the pyrolysis mechanism. TGA combined with *in situ* Fourier transform infrared spectroscopy (TG-FTIR) can be used to study the reaction mechanism of materials during pyrolysis and combustion. This technique allows for online identification of the evolution of pyrolysis gas phase products. Mohammed<sup>19</sup> *et al.* utilized TG-FTIR to examine the pyrolysis gas products of heavy fuel oil and vacuum residue. Wei<sup>20</sup> *et al.* investigated the impact of non-lignocellulosic components on the pyrolysis behavior of lees utilizing TG-FTIR. The effect of zinc-containing nanopowder on the catalytic pyrolysis of buckwheat straw was investigated by Tahir<sup>21</sup> *et al.* via TG-FTIR.

This study employed waste yellow wine lees as the raw material. It utilized TGA and TG-FTIR techniques to examine the pyrolysis properties of yellow wine lees at various heating rates. After collecting mass loss data, the activation energy of yellow wine lees at various conversion rates was determined using the corresponding distributed activation energy model (DAEM). The Flynn-Wall-Ozawa (FWO) method was utilized to verify the results, and a pyrolysis kinetic model of yellow wine lees was established to provide a theoretical framework for the successful application of yellow wine lees by pyrolysis.

## Raw material

The experiment employed fresh lees obtained from a yellow wine winery in Shaoxing and solid lees dehydrated using distillation and a belt drier. First, the solid lees were dried in an oven for 8 h at 110 °C, then pulverized with a pulverizer, sieved through a 200-mesh sieve, and sealed in a desiccator for storage, as illustrated in Fig. 1.

Yellow wine lees were analyzed for their ash, protein, crude fiber, starch, and amino acid contents following the national standards GB 5009.4-2016, GB 5009.5-2016, GB/T 5009.10-2013, GB 5009.9-2016, and GB 5009.124-2016. The samples' C, H, N, and S contents were determined using a Flash 2000 organic elemental analyzer manufactured in the US by Thermo Fisher



Fig. 1 Yellow wine lees samples before and after crushing.

Table 1 Chemical composition and ultimate analysis of yellow wine lees

Items	Dry basis, %	
Chemical composition	Ash	1.0
	Crude fibre	8.8
	Protein	27.7
	Starch	24.5
	Fat	6.9
	Amino acids	23.3
Ultimate analysis, %	C	49.7
	H	7.1
	O	36.5 <sup>a</sup>
	N	6.2
	S	0.5

<sup>a</sup> O content calculation formula = 100-C-H-N-S-ash.

Scientific (Table 1). In contrast to Baijiu lees and general biomass ethanol residue lees,<sup>22,23</sup> yellow wine lees contain a comparatively low ash content (1.0%) and crude fiber content (8.8%), according to the findings. The most notable change was in the protein and amino acid content of yellow wine lees, which reached 51.0% with a starch content of 24.5%.

## Experimental methods

### TGA of yellow winelees

The Al<sub>2</sub>O<sub>3</sub> test crucible was utilized to determine the TGA curve with the NETZSCH TG 209F1 Libra. The crushed sample had a mass of approximately 4.2–5.2 mg and was heated at rates of 10, 20, 30, and 40 °C min<sup>-1</sup>. Eventually, the temperature reached 900 °C, with a nitrogen flow rate of 50 mL min<sup>-1</sup>.

### TG-FTIR analysis of yellow wine lees

Yellow wine lees were subjected to a TG-FTIR simultaneous analysis experiment employing a TGA and an FTIR spectrometer (NETZSCH Perseus STA 449 F3 Jupiter). Sample parameters were temperature (25 °C to 900 °C), heating rate (10 °C min<sup>-1</sup>), protective gas (nitrogen) addition, and carrier gas flow rate (50 mL min<sup>-1</sup>).

### Kinetic analysis

(1) **Distributed activation energy model (DAEM).** The process of yellow wine lees pyrolysis involves a series of concurrent reactions. The DAEM approach relies on supposing limitless parallel reactions and the activation energy distribution. Moreover, the DAEM is commonly used in the kinetic investigation of biomass pyrolysis processes. Consequently, the activation energy for pyrolysis of yellow wine lees was determined using the DAEM method. Eqn (1) expresses the correlation between the activation energy and conversion rate when the DAEM model is employed to illustrate the weight loss law of the non-isothermal pyrolysis process<sup>24</sup>

$$\alpha(t) = \int_0^{\infty} 1 - \exp \left\{ \left| -\frac{A}{\beta} \int_0^t \left( \frac{-E}{RT} \right) dt \right| \right\} g(E) dE \quad (1)$$



The distribution function of activation energy, denoted as  $g(E)$ , is presumed to follow Gaussian distribution. Eqn (2) represents the correlation between the standard deviation  $\sigma$  and the average activation energy  $E_0$ .<sup>25</sup>

$$g(E) = \frac{1}{\sqrt{2\pi}} \exp \left[ -\frac{(E - E_0)^2}{2\sigma^2} \right] \quad (2)$$

Eqn (3) represents the dynamic form obtained by substituting eqn (2) for eqn (1) and simplifying the result *via* the Miura integration approach:<sup>26</sup>

$$\ln \left( \frac{\beta}{T^2} \right) = \ln \left( \frac{A}{RE} \right) + 0.6075 - \frac{E}{RT} \quad (3)$$

The present investigation employed the Arrhenius equation to calculate the activation energy ( $E$ ), with heating rates ( $\beta$ ) of 10, 20, 30, and 40  $\text{K min}^{-1}$ . The value of  $E$  was determined by dividing the slope of the fitted curve by the universal gas constant ( $R$ ) after plotting  $\ln(\beta/T^2)$  versus  $1/T$  at the same conversion rate.

(2) **Flynn-Wall-Ozawa (FWO).** The FWO is commonly employed as a standard method for verifying the precision of other models. Solid degradation is converted to char and volatile gas during the pyrolysis of yellow wine lees; therefore, solid reaction kinetics are more applicable. The Arrhenius equation, eqn (4), can be used to express the reaction rate equation:

$$\frac{d\alpha}{dt} = A \exp \left( \frac{-E}{RT} \right) f(\alpha) \quad (4)$$

where  $t$  represents the reaction time (s),  $R$  denotes the universal gas constant ( $8.314 \text{ J K}^{-1} \text{ mol}^{-1}$ ),  $T$  signifies the absolute temperature (K),  $A$  represents the pre-exponential factor,  $E$  signifies the activation energy,  $f(\alpha)$  represents the mechanism function of the reaction. The conversion rate, denoted as  $\alpha$ , is defined in eqn (5) as follows:

$$\alpha = \frac{m_0 - m_t}{m_0 - m_\infty} \quad (5)$$

where  $m_0$  is the initial mass, mg;  $m_t$  is the mass of lees corresponding to time  $t$ , mg;  $m_\infty$  is the remaining mass at the final pyrolysis temperature, mg. The pyrolysis of yellow wine lees is a non-isothermal linear heating experiment. The heating rate  $\beta$  is expressed as  $dT/dt$  and is used in eqn (5) to obtain:

$$\frac{d\alpha}{dt} = \frac{A}{\beta} \exp \left( \frac{-E}{RT} \right) f(\alpha) \quad (6)$$

Eqn (7) is obtained by integrating eqn (6):

$$G(\alpha) = \frac{A}{\beta} \int_0^T \exp \left( \frac{-E}{RT} \right) dT \quad (7)$$

After the FWO approach simplifies eqn (7), the kinetic equation is given as follows:<sup>27,28</sup>

$$\ln(\beta) = \lg \frac{AE}{RG(\alpha)} - 2.315 - 0.4567 \frac{E}{RT} \quad (8)$$

Similarly,  $\ln(\beta)$  was plotted *versus*  $1/T$  at the identical conversion rate with heating rates ( $\beta$ ) of 10, 20, 30, and 40  $^\circ\text{C min}^{-1}$ . The activation energy ( $E$ ) was calculated by determining the slope of the fitted curve as  $0.4567E/R$ .

## Results and analyses

### TGA of yellow wine lees

The TGA and DTG curves of the yellow wine lees pyrolysis were used as an example for analysis. The temperature range was 30 to 900  $^\circ\text{C}$ , and the heating rate was 10  $^\circ\text{C min}^{-1}$ . The findings are illustrated in Fig. 1. The overall mass loss during the pyrolysis of yellow wine lees was around 84.65%, as displayed in Fig. 2. The pyrolysis process is categorized into three stages based on the rate at which mass is lost: 30–121  $^\circ\text{C}$ , 121–500  $^\circ\text{C}$ , and 500–900  $^\circ\text{C}$ . As depicted in Fig. 2 of the DTG curve, the initial stage is drying, followed by the rapid degradation stage, which consists of three sub-stages, and finally, the gradual degradation stage.

The initial weight loss peak of the DTG curve corresponded to the temperature range of 30–121  $^\circ\text{C}$  during the first stage. This stage primarily involves the evaporation of excess moisture in yellow wine lees.<sup>29</sup> Due to the initial drying pretreatment of the sample, there was a minimal weight loss of only 3.95%.

The second stage, which involves three sub-stages, had a temperature range of 121–500  $^\circ\text{C}$  and a weight loss of 75.20%. In the first substage, the temperature varied between 121 and 218  $^\circ\text{C}$ , with a small pyrolysis peak at 175  $^\circ\text{C}$ . The weight loss during this stage can be attributed to several factors. The primary source of this occurrence is the evaporation of volatile organic compounds and the crystal water present in cellulose, hemicellulose, and lignin molecules.<sup>30</sup> Additionally, it results from dehydration, decomposition, and the decomposition of unstable short chains.<sup>31–33</sup>

Additionally, it pertains to the decarboxylation reaction between amino acids and proteins.<sup>34</sup> During the second substage, the temperature varied between 218 and 326  $^\circ\text{C}$ , with a maximum pyrolysis peak at 298  $^\circ\text{C}$ . This value is comparable

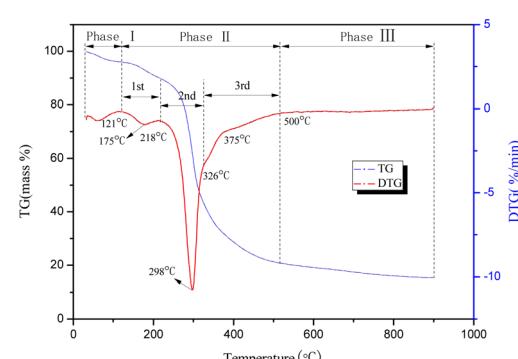


Fig. 2 TG/DTG curve of yellow wine lees at a  $10 \text{ }^\circ\text{C min}^{-1}$  heating rate.



to the maximum pyrolysis peak of protein measured by Sibilla *et al.*<sup>35</sup> the weight loss was about 41.68% due to protein, cellulose, and starch, which have degradation temperatures ranging from 217 to 450 °C, 290 to 368 °C, and 246.6 to 366.5 °C, respectively, according to previous research.<sup>34,36</sup> The temperature range of the third sub-stage was between 326 and 500 °C. This stage primarily involves the thermal decomposition of lignin and fat, which occurs within the temperature range of 356.2 to 423.5 °C.<sup>34</sup> The initial and subsequent reactions occurred at temperatures ranging from 326 to 375 °C and 375 to 500 °C, respectively.

During the third stage, the temperature range was 500–900 °C, and the yellow wine lees experienced a mass loss of 5.5%. The primary reason is that the solid result contains residual organic macromolecules that continue to undergo subsequent pyrolysis processes. At elevated temperatures, the carbon–carbon atoms held in place undergo spatial rearrangements and interact with the gas products produced by pyrolysis. This leads to numerous mesopores and micropores forming within the solid product of pyrolysis, increasing specific surface area.

### TG-FTIR analysis of yellow wine lees

With a temperature range of 30 to 900 °C and a 10 °C min<sup>-1</sup> heating rate, the 3D infrared spectrum of pyrolysis of yellow wine lees is displayed in Fig. 3. The pyrolysis and volatilization of yellow wine lees predominantly took place between the temperature range of 121 to 500 °C (Fig. 3). A close correlation can be seen between the appearance of the absorption peak and the weight loss that was discovered in the DTG curve that is depicted in Fig. 2. Under the conditions described above, every volatile component displays distinct infrared absorption peaks at 298 °C, the highest mass loss rate (Fig. 4). Fig. 4 shows that the volatile components have eight distinct infrared absorption bands at 650–750, 1000–1200, 1600–1900, 2000–2250, 2250–2400, 2700–3100, and 3400–3800 cm<sup>-1</sup>. Table 2 presents the findings of an analysis of the material composition of the volatile components using the functional groups associated with each characteristic absorption band.

The volatile components of yellow wine grains after pyrolysis consist primarily of inorganic compounds (CO, CO<sub>2</sub>, H<sub>2</sub>O, CH<sub>4</sub>) and organic compounds (esters, alcohols, phenols, amines, ethers, aldehydes, ketones, and acids), as indicated in Table 2.

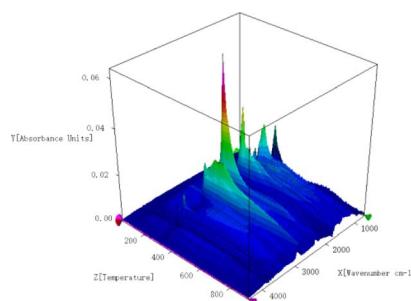


Fig. 3 3D infrared spectrum of pyrolysis of yellow wine lees at a 10 °C min<sup>-1</sup> heating rate.

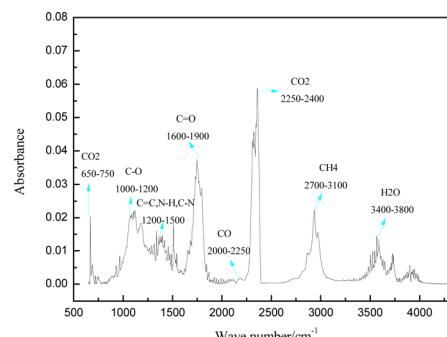


Fig. 4 Characteristic infrared absorption peaks of each volatile component at 298 °C.

The detected frequencies of 670 and 2360 cm<sup>-1</sup> correspond to the bending and stretching vibrations of the C=O bond, respectively. The vibrations are mainly a result of the breaking of the carbonyl group (–C=O) and the removal of the carboxyl group (–COOH) to produce CO<sub>2</sub>. The stretching vibration of the C–O bond in the 1000–1200 cm<sup>-1</sup> range was primarily associated with alcohol, phenol, ether, and ester compounds produced during the pyrolysis of hemicellulose, starch, fat, and lignin. The C=C bond stretching vibrations at 1200–1500 cm<sup>-1</sup> are primarily caused by aromatic compounds in the pyrolysis products.

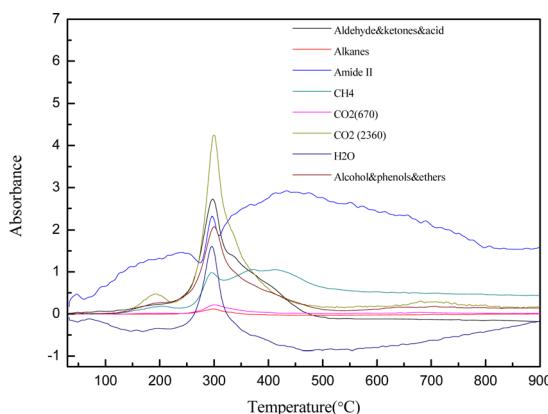
On the other hand, the N–H and C–N bond stretching vibrations in yellow wine lees are primarily caused by reactions between proteins and reducing sugars. The reactions involved include Maillard translation, deamination, dehydration, decarboxylation, and the formation of amides by the pyrolysis of cellulose, hemicellulose, and lignin.<sup>36–39</sup> The C–C and C–H stretching vibrations in this distinctive absorption band are primarily associated with the alkanes produced by the pyrolysis of lignin, hemicellulose, and fat.<sup>44</sup> The stretching vibration at 1745 cm<sup>-1</sup> corresponds to the C=O bond. This vibration is predominantly generated through the pyrolysis of cellulose and lignin in the sample, which yields ketone and aldehyde compounds in addition to organic acid compounds generated during protein pyrolysis.<sup>45</sup> 2175 cm<sup>-1</sup> is the stretching vibration of the C–O bond, primarily generated when the C–O–C and C=O bonds are cleaved and deoxygenated, forming CO.<sup>40,41</sup> The peak at 2937 cm<sup>-1</sup> corresponds to the C–H bond's stretching vibration, indicating that the material composition is primarily CH<sub>4</sub>. This is mainly associated with the random scission of aliphatic side chains, a degradation of C=C and methoxy (–OCH<sub>3</sub>) functional groups.<sup>40</sup> The thermal degradation process predominantly leads to water production through mechanisms such as the cross-linking reaction of cellulose and lignin, the extraction of free water in yellow wine lees, and the dehydration and condensation of proteins.

Fig. 5 depicts the variation in absorption intensity of volatile components as a function of temperature, with a 10 °C min<sup>-1</sup> heating rate and a temperature range between 30 and 900 °C. As per Lambert–Beer's law, a linear relationship existed between relative volatile component concentration and the absorption intensity at a given wave number. Consequently, the absorption



Table 2 FTIR analysis of yellow wine lees pyrolysis

Functional groups	Wavenumbers range (cm <sup>-1</sup> )	Species	Ref.
C=O	650–750	CO <sub>2</sub>	40 and 41
C–O	1000–1200	Alcohols, phenols, ethers, esters	32, 42 and 43
C=C, N–H, C–N, C–C, C–H (benzene skeleton)	1200–1500	Aromatics, amide II, alkanes	44
C=O	1600–1900	Aldehydes, ketones, organic acids	45
C–O	2000–2250	CO	40 and 41
C=O	2250–2400	CO <sub>2</sub>	40 and 41
C–H	2700–3100	CH <sub>4</sub>	46
O–H	3400–3800	H <sub>2</sub> O, phenols, aliphatic alcohol	34

Fig. 5 Infrared absorption intensity of yellow wine lees at various temperatures with a 10 °C min<sup>-1</sup> heating rate.

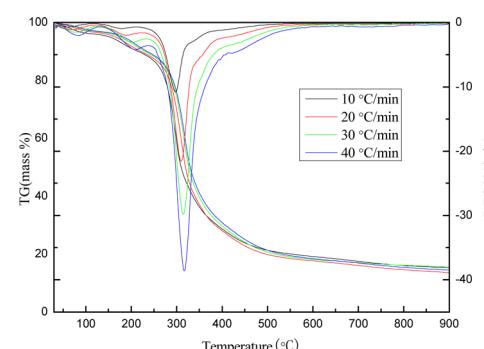
intensity of volatile components throughout the pyrolysis process can be utilized to determine the relative content of volatiles. Fig. 5 demonstrates a clear correlation between the variation of absorption peaks of volatile components and the mass loss observed in the DTG curve. Most volatile components displayed absorption peaks primarily during the second stage, ranging from 121 to 500 °C. The temperature at which its maximum absorption peak appeared corresponds to the position of the maximum pyrolysis peak of DTG, which primarily consists of volatile components, including CO<sub>2</sub>, H<sub>2</sub>O, aldehydes, ketones, acids, alcohols, and aromatic compounds. During the second phase of pyrolysis, CO<sub>2</sub> exhibited two absorption peaks. The initial absorption peak was observed at 121–218 °C, followed by another at 220–500 °C. The initial absorption peak is mainly attributed to the decarboxylation of carbohydrates and proteins.

In contrast, the second peak is primarily ascribed to the transesterification of fats and the deoxygenation of oxygen-containing groups at high temperatures.<sup>47,48</sup> CH<sub>4</sub> showed three absorption peaks in the second stage of pyrolysis, corresponding to the 121–218 °C, 218–325 °C, and 325–500 °C stages of the DTG curve. The absorption peaks observed within the temperature ranges of 121–218 °C and 218–325 °C primarily result from the thermal decomposition of cellulose, lignin, and protein alkyl side chains. The absorption peak observed between 325 and 500 °C is attributed mainly to the profound

degradation and rearrangement reactions of other biomass and the cleavage of the methoxy group (–OCH<sub>3</sub>) on the aromatic ring of lignin.<sup>49–51</sup> As mentioned above, the analysis revealed that the pyrolysis reaction of yellow wine lees involved the Maillard reaction and ketosis decarboxylation reactions of proteins (amino acids) and carbohydrates.

### Influence of heating rate on the pyrolysis process of yellow wine lees

The heating rate substantially affects the pyrolysis behavior of lees, as demonstrated by the TG/DTG curves of yellow wine grains at various heating rates in Fig. 6. Consistent with the findings of numerous researchers, the pyrolysis peak of lees shifted to a higher temperature as the heating rate increased, and the maximal mass loss peak temperature was increased from 298 °C to 318 °C.<sup>52–55</sup> This could be because the internally generated volatile components have not yet evaporated. On the other hand, the externally produced volatile components from pyrolysis reacted with the surface coke to form tar, leading to surface coking, restricted heat transfer, and increased temperature gradient within the lees. As a result, phenomena such as the broadening of the DTG peak shape and the elevation of the pyrolysis temperature associated with the maximal weight loss peak were observed. Moreover, the weight loss rate at a given temperature point and the maximum rate of weight loss through pyrolysis increased in response to an increase in the heating rate; this finding is in agreement with the pyrolysis results of lees reported by Wang<sup>56</sup> *et al.*, this is because the

Fig. 6 TG/DTG curves of yellow wine lees with various heating rates (10, 20, 30, and 40 °C min<sup>-1</sup>).

increased heat conduction between the sample and the reactor, which is induced by the accelerated heating rate, causes the interior of the lees to reach the maximum pyrolysis temperature more quickly when the pyrolysis process begins; volatile components precipitate rapidly, which in turn increases the rate at which weight is lost.<sup>57,58</sup>

### Analysis of pyrolysis kinetics of yellow wine lees

The apparent activation energies at various conversions were found using the DAEM and FWO plots to pyrolyze yellow wine lees. Fig. 7 and 8 depict the Arrhenius linear fitting diagrams for the DAEM kinetic model and the FWO model-free integration approach applied to yellow wine lees. The Fig. 7(a) and 8(a), 7(b) and 8(b) and 7(c) and 8(c) had conversion rates of 0.02–0.18, 0.20–0.55 and 0.60–0.80, respectively. For most conversions, the regression coefficients were high (>0.95) and qualified as a satisfactory fit for assessing apparent activation energy using various techniques. This also results in a rather stable and reliable activation energy value. Fig. 8 shows that both the FWO and the DAEM models produce nearly identical curve-fitting degrees, supporting the idea that the calculated activation energy value is accurate.

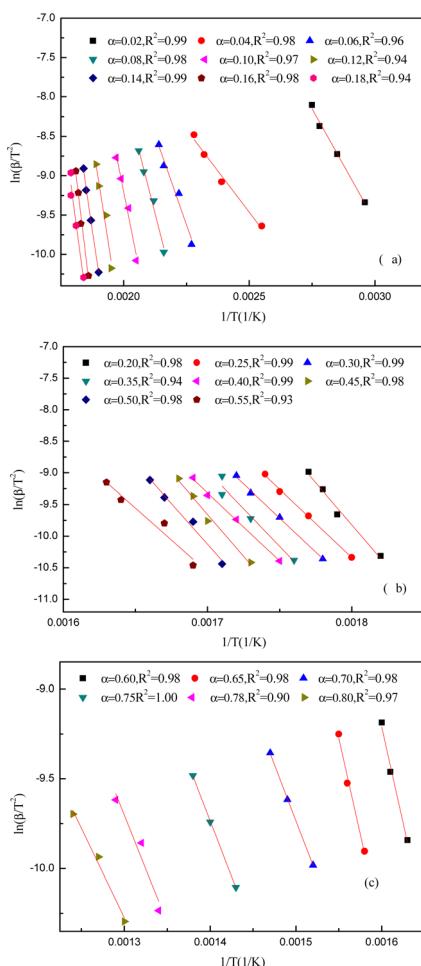


Fig. 7 Linear fitting of DAEM model of yellow wine lees under different conversion rates (a–c).

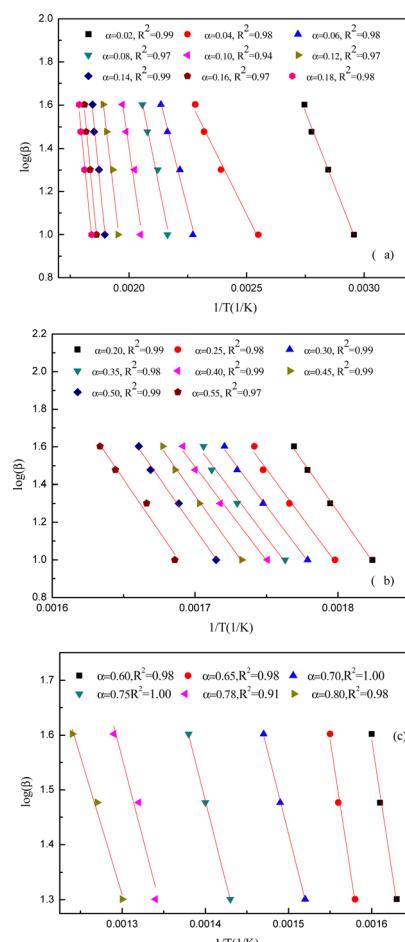


Fig. 8 Linear fitting of FWO model of yellow wine lees under different conversion rates (a–c).

The activation energies were determined for various conversion rates (as indicated in Fig. 9 and Table 3). The DAEM and FWO models calculated the activation energies were 35–220 and 40–210 kJ mol<sup>-1</sup>, respectively. The two strategies yield nearly identical calculation results. Thus, the yellow wine lees pyrolysis law is highly reliable and complies with the DAEM hypothesis mechanism of infinite parallel reactions and activation energy distribution. As depicted in the figure, the

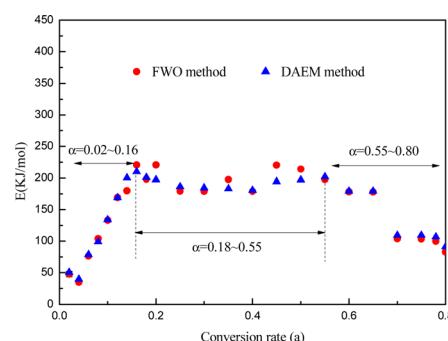


Fig. 9 Activation energy and conversion rate diagram of yellow wine lees.

Table 3 The kinetic parameters of pyrolysis of yellow wine lees

Conversion rate, %	<i>E</i> (kJ mol <sup>-1</sup> )	Ln <i>A</i> (min <sup>-1</sup> )	<i>R</i> <sup>2</sup>
0.02	47.55	7.00	0.99
0.04	34.89	6.68	0.98
0.06	76.04	17.14	0.96
0.08	104.08	23.68	0.98
0.10	133.23	29.22	0.97
0.12	169.26	35.37	0.94
0.14	179.84	36.47	0.99
0.16	220.84	44.37	0.98
0.18	198.08	38.95	0.94
0.20	220.64	41.69	0.98
0.25	179.34	33.87	0.99
0.30	179.23	32.71	0.99
0.35	197.66	35.75	0.94
0.40	179.07	31.24	0.99
0.45	220.15	37.41	0.98
0.50	214.05	35.21	0.98
0.55	197.90	28.15	0.93
0.60	178.44	36.11	0.98
0.65	178.11	33.40	0.98
0.70	103.91	21.49	0.98
0.75	103.20	20.23	1.00
0.78	99.80	19.04	0.90
0.80	82.95	12.71	0.97

activation energy value of yellow wine lees increased initially, stabilized, and then declined as the conversion rate increased.

The conversion rate of lees in the initial region was  $0.02 \leq \alpha \leq 0.16$ , while the corresponding temperature ranged from 65 to 264 °C. At  $\alpha = 0.04$ , the yellow wine lees exhibited a low activation energy of 35–40 kJ mol<sup>-1</sup> and a temperature of 119 °C. This is consistent with the first weight loss peak observed in TGA, primarily due to free moisture vaporization in the lees. The activation energy increased substantially, reaching 210–220 kJ mol<sup>-1</sup> at  $\alpha = 0.16$ . The main reason is that lower activation energies are produced when hemicellulose is thermally degraded because weakly connected spots in the linear chains of hemicellulose polymers are easily reached. The linear chain then experiences random rupture following the weak link breaking, raising the activation energy.<sup>59</sup> The substantial variation in apparent activation energy at this point showed the presence of a multi-step decomposition mechanism.<sup>19</sup> The temperature in the second region ranged from 270 to 320 °C, and the conversion rate for the yellow wine lees was between 0.18 and 0.55, inclusive. At this stage, the primary substances undergoing pyrolysis are protein and cellulose. In the present study, the activation energy remains consistently around 190 to 200 kJ mol<sup>-1</sup>, indicating that this process follows a single-step breakdown mechanism.<sup>19</sup> The activation energy gradually dropped to approximately 180 kJ mol<sup>-1</sup> when  $\alpha = 0.60$ –0.65. However, the activation energy decreased to approximately 100 kJ mol<sup>-1</sup> at  $\alpha = 0.70$ –0.80. This might be because macro-molecular compounds, including cellulose, hemicellulose, lignin, starch, and protein, produce active substances, including  $-\text{OCH}_3$ , during the initial stage of pyrolysis, facilitating subsequent stages. These findings are in agreement with the TG-FTIR analysis.

## Conclusions

Yellow wine lees, the primary organic waste generated during the Chinese yellow wine brewing process, are a significant renewable resource capable of substituting fossil fuels for bio-energy production. Pyrolysis is a viable conversion method for yellow wine lees. It can efficiently utilize lees to produce valuable biofuels. This work investigated the pyrolysis process of yellow wine lees using TG/DTG, TG-FTIR, DAEM, and FWO kinetic model approaches. First, the elemental and chemical composition studies showed that the main components of yellow wine grains include lipids, crude fiber, protein, starch, and amino acids.

Further, it possesses a small amount of ash and a significant concentration of protein and amino acids, making it a beneficial primary substance for producing biochar, particularly nitrogen-doped carbon compounds. Subsequently, the process of decomposing yellow wine lees using pyrolysis is divided into three distinct stages. The main pyrolysis stage for yellow wine lees was 121–500 °C, with a maximal pyrolysis rate of 298 °C. The Maillard reaction and ketosis-decarboxylation reaction mainly occurred at this stage. The volatile components after pyrolysis include inorganic molecules such as CO<sub>2</sub>, H<sub>2</sub>O, and CH<sub>4</sub> and organic compounds such as esters, alcohols, phenols, amines, ethers, aldehydes, ketones, and acids. Finally, the pyrolysis kinetics of yellow wine lees were simulated utilizing the DAEM and FWO models. The findings indicated that the pyrolysis of yellow wine lees conforms to the assumptions of infinite parallel reactions and activation energy distribution. The kinetic mechanism identified in this work can serve as a theoretical base for the subsequent design and advancement of the pyrolysis process of yellow wine lees.

However, further studies are needed to explore further aspects of the pyrolysis process of yellow wine lees. The effect of pyrolysis conditions and the concentration of proteins and amino acids in the raw materials on the surface morphology of biochar is currently not understood. Moreover, it is imperative to advance the understanding of the creation mechanism of specific volatile compounds.

## Author contributions

Kong Ling-Niao: investigation, data curation, and writing-original draft. Ge Song-Tao: resources and visualization. Yuan Yang: project administration and funding acquisition. Feng Feng: conceptualization and writing-review and editing.

## Conflicts of interest

There are no conflicts to declare.

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