


 Cite this: *RSC Adv.*, 2024, 14, 10004

Development of enduring interstitial defects in Mg-doped CuO thin films†

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Doping is one of the easiest methods to enhance the properties of semiconducting thin films. This work presents a systematic study of the influence of Mg doping on the optoelectronic properties of copper oxide (CuO) thin films prepared by spray pyrolysis. The pristine and doped films were prepared on glass substrates at an optimized temperature of 350 °C and a precursor solution concentration of 0.15 M while the Mg doping concentration was varied between 2 and 10 at%. The properties of the prepared films were studied in detail by using various characterization techniques such as X-ray diffraction, field emission scanning electron microscopy, Raman and photoluminescence, X-ray photoelectron spectroscopy and Hall measurement. The results suggest that oxygen interstitial (O_i) defects were enhanced upon doping which has influenced the properties of the film. The size of the crystallites reduced, and the band gap widened after doping. In addition, comprehensive density functional theory analyses have been conducted to gain a thorough understanding of the impact of impurity (*i.e.*, doping/interstitial) atoms on the optical and electrical characteristics of CuO. This study shows that O_i defects improved the conductivity and enhanced the carrier concentration leading to an improvement in the electrical properties of the films.

Received 23rd February 2024

Accepted 21st March 2024

DOI: 10.1039/d4ra01414f

rsc.li/rsc-advances

Introduction

Technological and industrial development has resulted in a huge rise in energy demands for mankind such that the existing non-renewable sources of energy are no longer sufficient for the long run. This has resulted in a high rise in energy-related research, both in production and storage with a special focus on renewable sources. Being the most abundant source, the hottest niche is that of solar energy. The initial solar energy-related research was mostly concentrated on Si technology, which has lost its glory, mainly due to the high fabrication cost and scant availability of Si.¹ Thus, the search for new compounds started off with various thin films, to perovskites, organic compounds and has reached quantum dots and many more. Today we have fabricated solar cells with an efficiency of over 25% using many of these materials. However, most of these faced a setback when it came to cost effectiveness, non-toxicity and sustainability. This has forced the scientific world to search for alternative materials that are environmentally friendly and economical in the long run.

Among the different materials, oxides of copper gained considerable attention not only because of its abundance and low cost but also for its non-toxicity. There are two major stable

oxides of copper: cuprite (Cu_2O) and tenorite (CuO). Both cuprite and tenorite have been widely used for solar energy applications. In spite of the ideal band and high absorption coefficient of CuO thin films, the efficiency of CuO based solar cells has always been lower. The major reason for this is the high resistivity of CuO thin films. A low level of doping was observed to improve the electrical properties by reducing the resistivity of CuO thin films. Higher levels of doping however worsened the efficiency because of the formation of Cu_2O phase and the band gap widening.² Dopants like nitrogen, titanium *etc.* were observed to improve the photovoltaic performance of CuO-based solar cells indicating the positive impact of doping.^{3,4}

In this work, we investigated the effect of Mg doping on the properties of spray deposited CuO oxide films. Isovalent doping can significantly influence the intrinsic defects in the film. In CuO thin films, the intrinsic defects are mainly due to the Cu vacancies (V_{Cu}), Cu interstitials (Cu_i), oxygen vacancies (V_o) or oxygen interstitials (O_i). Out of these defects, V_{Cu} and O_i are acceptor defects that can improve the electrical conductivity of the films. So, if we can enhance these acceptor defects, the electrical properties of the films can be improved significantly. As Mg^{2+} ions have an ionic radius comparable to Cu^{2+} ions, Mg^{2+} can readily substitute Cu^{2+} without causing significant strain or distortion. In addition to these, the strong oxygen affinity of Mg may improve the O_i defects in the films. Hence, Mg was chosen as the suitable dopant. Even though defects can enhance the properties, it is important to control these defects

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† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d4ra01414f>



because higher defects can act as scattering centres for charge carriers affecting their electrical properties. Hence, we are trying to tune the intrinsic defects in CuO thin films especially, O_i defects by a low level of Mg doping (2–10 at%).

Spray pyrolysis is a cost-effective method that is suitable for large area deposition. It's a chemical method in which the precursor solution of suitable metal salts is sprayed onto a heated substrate. As the aerosol falls on the heated substrate, pyrolysis occurs which results in the formation of the required compound on the substrate. The by-products are usually in gaseous form, which are removed from the chamber through the exhaust. Spray pyrolysis is one of the suitable methods for the fabrication of solar cells as it allows large area deposition. Also, the preparation of thin films and even their modification can be done easily without any hassle using this method. There have been several reports on the preparation of pristine and doped CuO thin films by spray pyrolysis in the recent past. However, only a few reports are available on Mg-doped CuO thin films prepared by spray pyrolysis. Also, the influence of Mg doping on the intrinsic defects and properties of CuO thin films has not been investigated. We have explored the influence of Mg doping on the intrinsic defects of spray deposited CuO thin films. The experimental results were verified using Density Functional Theory (DFT) calculations. The results from this study are expected to provide insights into the effect of Mg doping on CuO thin films and also explore its suitability for solar cell application.

Methodology

Experimental details

The Mg-doped CuO thin films were prepared *via* simple spray pyrolysis technique. All the prepared thin films were fabricated on ultrasonically cleaned, acid etched soda-lime glass slides using a series of distilled water, acetone, and isopropyl alcohol baths. Magnesium chloride hexahydrate ($MgCl_2 \cdot 6H_2O$) and copper chloride dihydrate ($CuCl_2 \cdot 2H_2O$) were used as source salts for the preparation of 0.15 M of the spray solution in distilled water. The precursor solution with different Mg doping concentrations (0–10 at%) was prepared by varying the volume of Mg salt solution added to copper chloride solution by keeping the total volume fixed at 10 mL. The samples were named CMO 0, CMO 2, CMO 4, CMO 6, CMO 8 and CMO 10 with the number indicating the atomic percentage of Mg doping. The resulting mixture of salt solutions was magnetically stirred to ensure uniform mixing. The obtained stock solution was then sprayed onto glass substrates preheated to 350 °C on the hot plate inside the spray chamber. The distance between the nozzle of the spray syringe and the substrates was maintained at 14 cm. The air compressed at a pressure of 2 mbar produces a mist of the solution as it is pushed out of the syringe which when comes in contact with the heated substrates undergoes thermolysis resulting in the formation of thin films. The solution was sprayed at regular intervals for about 20 minutes to obtain quality films. The obtained films were deep black in colour and the thickness of the films, measured using Bruker profilometer, was in the range of 550 ± 50 nm. A gradual

increase in the thickness of the films was observed with the increase in Mg doping concentration.

The properties of the films were studied using suitable characterization techniques. The information regarding the structural properties of the films was obtained from X-ray Diffraction (XRD) using Rigaku X-ray diffractometer. The diffractogram of the films was obtained using Cu-K α X-ray source in the 2θ range of 20 to 80° at a scanning rate of 1° min⁻¹. The phase purity of the films was further confirmed by studying the vibrational modes using Raman spectrometer (Horiba LABRAM) with a laser source of 532 nm. The composition of elements and the morphology of the prepared nanoparticles were obtained from Field Emission Scanning Electron Microscopy (FESEM) with Energy Dispersive X-ray Spectroscopy (EDS) (Zeiss SEM EV01S). The optical absorbance of the synthesized films was recorded using SHIMADZU UV 1800 double beam spectrophotometer. Photoluminescence spectra of the films were taken at an excitation wavelength of 532 nm using Horiba LABRAM to understand the defects in the system. The electrical properties of the films were measured using Keithley 2450 source meter employing the van der Pauw technique. Further, the chemical composition of the samples was analysed by X-ray Photoelectron Spectroscopy (XPS) using PHI VersaProbe III having Al K α source of photon energy 1486.6 eV.

Theoretical details

The DFT calculations were performed using the Quantum Espresso (QE) software, specifically within the framework of a plane-wave self-consistent field (PWscf) method.⁵ The computations employed the generalised gradient approximation (GGA + U) for accuracy⁶ Monoclinic space group $C2/c$ with the lattice constants $a = 4.6837$ Å, $b = 3.4226$ Å, and $c = 5.1288$ Å of pure and Mg-doped CuO were employed for the calculation. Further, employed $U = 3.5$ eV at Cu-3d orbitals with $3 \times 3 \times 3$ supercells. The optimized structures of these supercells are shown in Fig. 1. We have used nearly ~4 at% and ~8 at% doping of Mg on pure and defective CuO structures. A k -mesh of size $4 \times 4 \times 4$ was employed within the first Brillouin zone.⁷ The self-consistent computations reach convergence when the total energy of the system remains constant within a range of 10^{-3} mRy.⁸ The relaxation process for both pure and defective structures carried till the energy convergence reaches 4×10^{-5} eV and continues until the forces per atom decrease to 0.04 eV.⁹

Results and discussions

Structural characterization

The crystalline structure and phase of CMO samples were identified using XRD. The obtained diffraction pattern is depicted in Fig. 2(a). All the samples exhibited similar diffraction patterns corresponding to the monoclinic phase of CuO thin films. The diffractogram of the films consists of two dominant peaks and a few other minor peaks. The major peaks at 36° and 39° correspond to the reflection from (111) and (111) diffraction planes respectively. All the minor peaks were also identified and indexed comparing with the standard JCPDS card



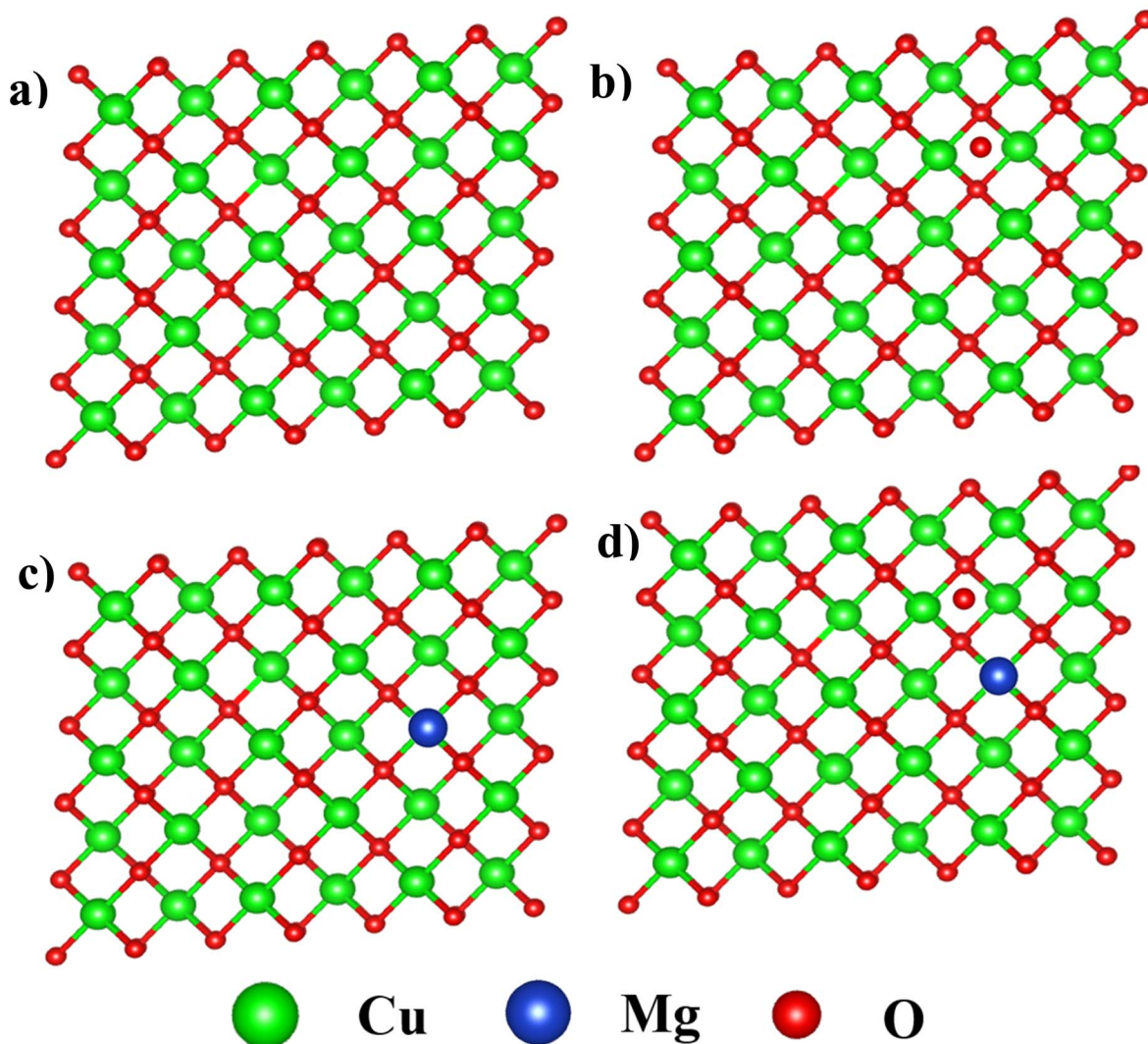


Fig. 1 Crystal structure of (a) pure CuO (b) O_i on CuO, (c) Mg doped CuO and (d) O_i on Mg doped CuO.

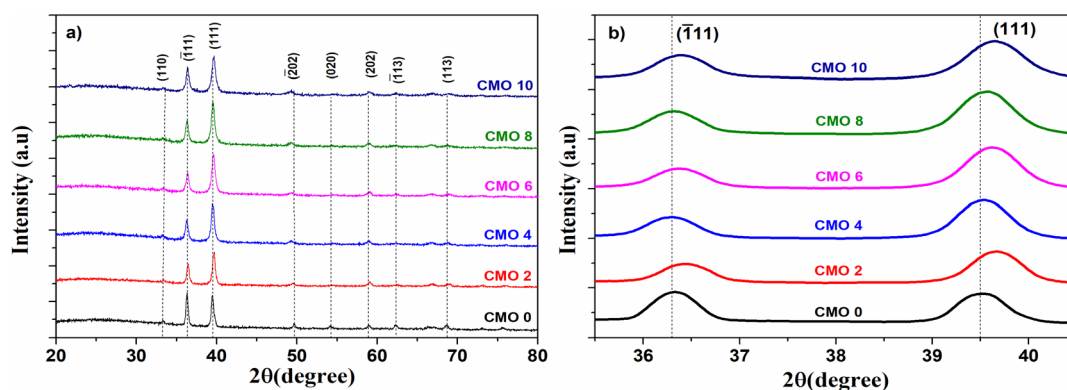


Fig. 2 (a) X-ray diffractogram of prepared CMO films, (b) shift in the peak position of (111) and $(\bar{1}\bar{1}\bar{1})$ reflection planes.

(01-077-7718). No other peaks corresponding to other impurity phases like Cu_2O or MgO were observed in the diffraction pattern confirming the phase purity of the prepared samples. This indicates that Mg^{2+} ions can easily penetrate into the CuO

lattice without affecting the crystal structure. However, peaks corresponding to (020) and (113) diminished or disappeared from the diffraction pattern with Mg doping which indicates that incorporation of Mg ions has influenced the grain growth



in CuO thin films. A higher degree of preferential orientation of CuO planes was therefore observed in doped samples.

Incorporation of dopant atoms can cause lattice distortions, mainly due to the size difference between the host and dopant atoms. The lattice distortion can be observed from the change in lattice parameters extracted from the XRD data using unit cell software. For monoclinic CuO, lattice parameters a , b , c is related to interfacial angle β and Miller indices h , k , l by,¹⁰

$$\frac{1}{d_{hkl}^2} = \frac{1}{\sin^2 \beta} \left(\frac{h^2}{a^2} + \frac{k^2 \sin^2 \beta}{b^2} + \frac{l^2}{c^2} - \frac{2hl \cos \beta}{ac} \right) \quad (1)$$

The obtained values of lattice parameters are shown in Table 1. There is an expansion observed along both a and c whereas, the b showed a decrease in its value upon doping. The positions of the two major peaks in the diffractogram are observed to be shifted with Mg doping which is depicted in Fig. 2(b). The position of (111) and $\bar{1}\bar{1}\bar{1}$ showed a shift towards a slightly larger angle. This right shift in the peak position is attributed to the smaller size of Mg²⁺ ions compared to the Cu²⁺ ions. When Cu²⁺ cations are replaced by smaller Mg²⁺ ions, lattice compression occurs which results in the shift in the peak position towards higher angles.¹¹ A deviation from this trend is shown by (111) peak of CMO 4 and CMO 8 samples in which they show a left shift in the peak position. This can be accounted for by the residual stress developed during the film growth. Residual stress causes anisotropic change in lattice parameters leading to lattice distortion. Tensile stress causes the peaks to shift to lower angles whereas compressive stress shifts the peak towards higher angles.¹²

The intensity and full width half maximum (β_{hkl}) of the major peaks in XRD were determined and it was found that the intensity of the peaks decreased, and the peaks broadened after doping. The intensity of diffraction peaks depends on several factors such as the structure factor and atomic scattering factor. When Mg atoms are doped into the CuO lattice, it changes the mean atomic scattering factor and structural factor and there by cause a change in intensity. The broadening of the diffraction peaks is associated with the lattice distortions introduced by the dopant atoms. The orientation of grain growth was determined by calculating the texture coefficient (TC_{hkl})¹³ given by,

$$TC_{hkl} = \frac{I_{hkl}/I_{0hkl}}{\frac{1}{N} \left(\sum_N I_{hkl}/I_{0hkl} \right)} \quad (2)$$

Table 1 Lattice parameters of the prepared CMO thin films obtained from XRD

Sample	a (Å)	b (Å)	c (Å)	β (degrees)	V (Å ³)
CMO 0	4.574	3.370	5.070	97.78	77.45
CMO 2	4.598	3.353	5.082	99.39	77.29
CMO 4	4.625	3.333	5.095	99.48	77.47
CMO 6	4.621	3.319	5.092	99.56	77.02
CMO 8	4.624	3.325	5.103	99.57	77.36
CMO 10	4.630	3.315	5.097	99.69	77.11

where N is the total number of diffraction peaks, I_{hkl} represents the observed intensity of hkl plane in XRD and I_{0hkl} is the intensity in the standard JCPDS card. The value of TC depends on the orientation of crystallites and is represented by considering 1 as the reference value. When the crystallites are dispersed randomly, $TC_{hkl} = 1$. If the value of TC_{hkl} is more than 1, it denotes crystallite abundance, whereas a value of TC_{hkl} less than 1 denotes poor crystallite orientation in the given direction.¹⁴ The TC_{hkl} values of the two prominent peaks ($\bar{1}\bar{1}\bar{1}$) and (111) were calculated and is given in Table 2. It is observed that, before doping, $TC_{\bar{1}\bar{1}\bar{1}}$ has a value greater than 1 and after doping it decreased and for doped samples, TC_{111} is greater than 1. This indicates a clear change in orientation of the crystallites from ($\bar{1}\bar{1}\bar{1}$) plane in the undoped film to (111) plane in the doped films.

The size of the crystallites (D) in the films was calculated from size-strain plot¹⁵ given by,

$$(d_{hkl}\beta_{hkl}\cos\theta)^2 = \frac{K\lambda}{D} (d_{hkl}^2\beta_{hkl}\cos\theta) + \frac{\varepsilon^2}{4} \quad (3)$$

in which, λ is the wavelength of X-ray, d_{hkl} is the lattice distance, K is the shape factor, ε is micro-strain and θ is the diffraction angle. Crystallite size of the films decreased from 16 to 14 nm at 2 at% doping which further decreased to the range of 7–8 nm for higher doping concentration. The decrease in the crystallite size upon doping can be attributed to the substitution of Cu²⁺ by Mg²⁺ ions. With the size of Mg²⁺ ions (0.65 Å) being comparatively smaller than Cu²⁺ ions (0.72 Å), Mg doping produces distortions in the CuO lattice. Other structural parameters like number of crystallites per unit area (N) and dislocation density (δ) were calculated from the following equations¹⁶ and the obtained values are given in Table 2.

$$N = \frac{t}{D^3} \quad (4)$$

$$\delta = \frac{1}{D^2} \quad (5)$$

Fig. 1 depicts the optimized structure of $3 \times 3 \times 3$ CuO supercell containing 108 atoms of Cu and O each. Optimized k -mesh for the CMO structures were obtained to be $4 \times 4 \times 4$ for k_x , k_y and k_z Brillouin zone as presented in Fig. S1(a).† Structural optimization also been done and the same is depicted in Fig. S1(b).† The minimized energy of these structures signifies the stability of the crystal lattice. To explore the impact of Mg doping, two different concentrations: ~4 at% and ~8 at% were theoretically investigated through DFT calculations. Additionally, we computed the formation energy of the CuO crystal structure with intrinsic interstitial oxygen defects. The results revealed a reduction in formation energy from -1.021 eV in pristine CuO to -1.062 eV in 8 at% doped CuO, indicating the stability of CuO post-Mg doping and affirming the suitability of the dopant.

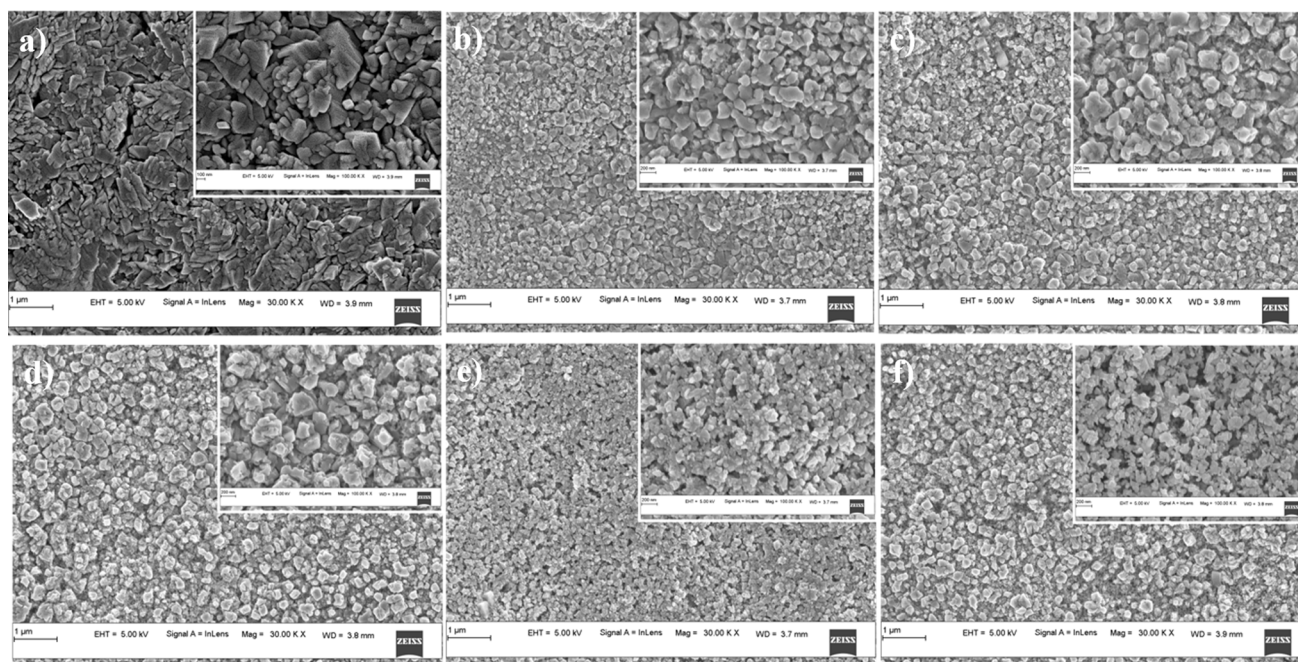
Morphological characterization and elemental analysis

The morphology of the surface of synthesized films were observed using FESEM. Fig. 3 shows the FESEM micrographs of



Table 2 Structural parameters of CMO thin films for different levels of Mg doping

Sample	Crystallite size (nm)	Dislocation density ($\times 10^{16} \text{ m}^{-2}$)	No. of crystallites ($\times 10^{18} \text{ m}^{-2}$)	TC ₁₁₁	TC ₁₁₁
CMO 0	16 ± 1	0.40	0.88	1.15	0.85
CMO 2	14 ± 1	0.49	0.15	0.84	1.16
CMO 4	7 ± 1	2.25	1.55	0.79	1.21
CMO 6	9 ± 1	1.19	0.60	0.71	1.29
CMO 8	8 ± 1	1.73	1.04	0.74	1.25
CMO 10	9 ± 1	1.27	0.66	0.81	1.19

Fig. 3 FESEM micrographs (magnification: 30k \times) of: (a) CMO 0, (b) CMO 2, (c) CMO 4, (d) CMO 6, (e) CMO 8, (f) CMO 10 thin films (insets show images at 100k \times magnification).

spray deposited pristine and Mg doped CuO thin films. All the films appear to be uniform and crack free with small grains. No pinholes were observed in the FESEM images indicating uniform deposition and good quality of the films. The grain size of the doped films appears to be smaller than that of the pristine samples, indicating that Mg doping has influenced the growth of the grains in the films. The calculated grain size of the films is given in Table 3. The decrease in the grain size of the film upon doping is due to the strain introduced by the dopant atom.

Table 3 Compositional analysis of different elements present in CMO films from EDS

Sample	Cu (at%)	O (at%)	Mg (at%)	$\frac{\text{Mg}}{\text{Mg} + \text{Cu}}$	Grain size (nm)
CMO 0	46.21	53.79	0	0	203 ± 8
CMO 2	40.28	58.58	1.14	0.03 (0.04)	142 ± 4
CMO 4	38.94	58.79	2.28	0.06 (0.08)	106 ± 10
CMO 6	38.1	58.73	3.17	0.08 (0.12)	155 ± 3
CMO 8	36.87	58.84	4.28	0.10 (0.16)	93 ± 4
CMO 10	36.10	58.84	5.06	0.12 (0.20)	102 ± 10

The morphology of the Mg doped films appears to be slightly different from that of the undoped film. The doped films have smaller, more or less cubical shaped nanoparticles compared to the large irregularly shaped nanoparticles as observed in the undoped film. This change in the morphology of the films may be associated with the difference in electronegativity of dopant and host atoms. Cu ions have a higher electronegativity value of 1.90 compared to Mg whose electronegativity is 1.31.¹⁷ As the electronegativity of Mg is less than Cu, the growth rate of CuO is reduced by Mg doping which causes the morphological variation.¹¹

The composition of elements in the films was studied using EDS. As a representative, the EDS spectrum of CMO 8 is given in Fig. S2(a).† The peaks corresponding to Cu, O and Mg is observed in the spectrum confirming the presence of these elements in the deposited films. The estimated elemental composition of the films is given in Table 3. It is observed that the films have become oxygen rich in nature upon Mg doping. The Mg dopant concentration is found to be less than that of the source values. This deviation may be due to the difference in the reaction mechanism of both Mg and Cu salts.¹⁸ The



distribution of the elements within the films was obtained using elemental mapping. Fig. S2(b)–(d)† shows the elemental mapping of O, Mg and Cu in CMO 8 sample. The distribution of the elements is observed to be uniform throughout the samples.

UV-Vis spectroscopic analysis

The optical properties of the prepared films were studied using UV-Vis spectroscopy. Fig. 4(a) illustrates the absorption spectra of pristine and Mg doped CuO thin films deposited on glass substrate. All the films exhibited high absorbance in the spectral range of 550–600 nm which makes them suitable for photovoltaic applications. Doping has resulted in an increase in the absorbance of the films, with CMO 10 showing the maximum absorbance at 580 nm. The absorption edge of pristine CuO thin films lies around 748 nm, which showed a nominal shift towards the lower wavelength region indicating a blue shift in the absorbance edge. The blue shift is associated with the smaller crystallite size of the samples after doping.

For direct band gap semiconductors, the optical band gap can be determined from Tauc relation given by,

$$\alpha h\nu = B(h\nu - E_g)^{1/2} \quad (6)$$

where α is the absorption coefficient, B is the energy independent constant, $h\nu$ is the incident photon energy.¹⁹ The Tauc plot for all the films is shown in Fig. 4(b). The band gap of undoped film was 1.5 eV, which increased to 1.69 eV at 6% doping. Further increase in the doping caused a decrease in the band gap. The optical properties of thin films are largely depended on the structural parameters such as micro strain and crystallite size which is known as quantum confinement effect.²⁰ The inverse relationship between E_g and R is given by,²¹

$$E(R) = E_g(\infty) + E_b \left(\frac{\pi a_B}{R} \right)^2 \quad (7)$$

where R is radius of spherical quantum dots, a_B and E_b are the exciton Bohr radius and binding energy respectively, $E_g(\infty)$ is

the band gap of bulk material and $E(R)$ is the lowest transition energy. However, this relation is valid only if $R < a_B$. The value of a_B lies between 6.6 to 28.7 nm.²² The values of a_B and E_b are constant for CuO. As the crystallite size of the prepared CMO films are less than a_B , the widening of optical band gap occurs with decrease in the crystallite size. The decrease in the band gap at higher level of doping is associated with band filling.²³ At higher level of doping (above 6 at%), large number of O_i defects are introduced leading to an increase in the number of charge carriers, *i.e.*, holes. Thus, O_i defect levels created within the energy band gap may overlap with the valence band and thereby effectively reduced the band gap.

We have also simulated the optical spectra using first principle calculations to compare experimental bandgap as well as finding excitonic binding energy for pure and Mg doped CuO (see Fig. S3†). Observed simulated optical absorption spectra are well matched with our experimental findings. Using optical absorption spectra, we have also calculated excitonic binding energy²³ and found that in case of Mg doped CuO, excitonic binding energy is higher than pristine samples which is due to doping and interstitial effect in CuO (see Fig. S3†).

Raman analysis

CuO having monoclinic structure belongs to the space group of $C2/c$ ($C_2^6_{2h}$). The primitive unit cell of CuO has two molecules, *i.e.*, four atoms, resulting in three acoustic and nine phonon branches represented as,²⁴

$$\Gamma = 4A_u + A_g + 5B_u + 2B_g \quad (8)$$

Out of the nine optical modes, six are IR active and three are Raman active modes.²⁵ The Raman modes include the A_g and $2B_g$ symmetries. The acoustic modes are represented by A_u and $2B_u$ symmetries and the remaining $3A_u$ and $3B_u$ are IR active modes.²⁶ All the Raman active modes of CuO are because of the displacement of oxygen atoms and in all the cases, Cu atoms remain stationary. This is due to the site symmetry of oxygen atom.²⁷ A_g mode represents the in-phase or out-phase rotation

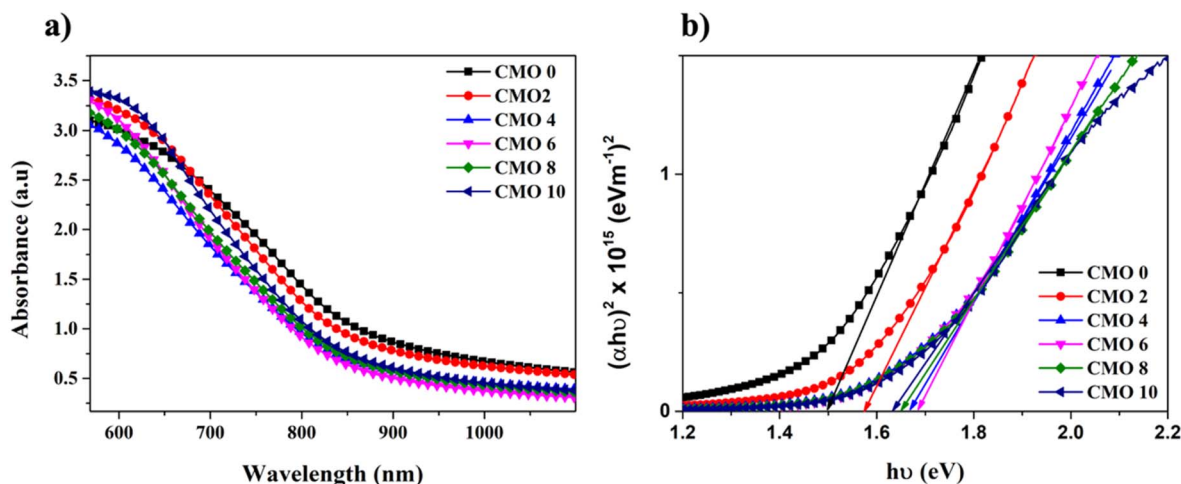


Fig. 4 (a) Absorbance spectra and (b) Tauc plot of CMO films.



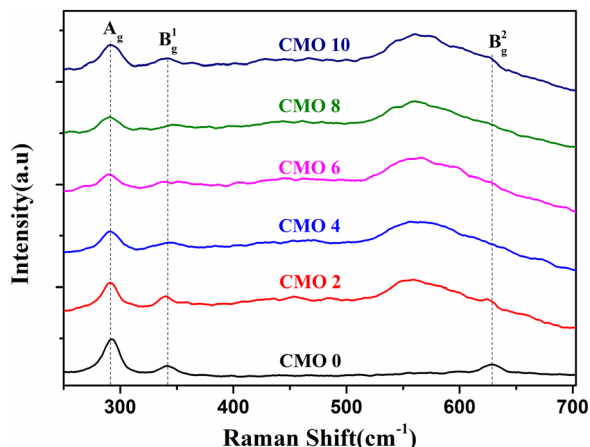


Fig. 5 Raman spectra of spray deposited CMO films.

of CuO and B_g^1 , B_g^2 modes are due to the bending of CuO and symmetric stretching of oxygen respectively.^{28,29}

Fig. 5 shows the Raman spectra of Mg doped CuO thin films deposited by spray pyrolysis on glass substrate. Spectra shows all the three Raman active modes which are centred at 293, 342 and 628 cm^{-1} in the pristine sample. The peak at 293 cm^{-1} corresponds to the A_g mode whereas the peaks at 342 and 628 cm^{-1} are assigned to B_g^1 and B_g^2 modes respectively. These values are in good agreement with those found in the literature.³⁰ While the position of A_g mode remained the same for all the doped samples, there was a shift in the peak positions of B_g modes. This is due to the fact that vibrations of B_g modes are associated with the oxygen atom and the increase in the oxygen content after doping has resulted in the shift in peak position.

The position of B_g^2 mode showed a much higher redshift as compared to B_g^1 . Also, the B_g^2 is much broader compared to other Raman peaks present in the sample. The broadness in Raman peaks is mainly associated with poor crystallinity, the presence of defects and the quantum confinement effect.³¹ As observed from the XRD, the large redshift and increase in the broadness of the peak can be assigned to a decrease in the crystallite size of the nanoparticles and the confinement of vibrations along the radial direction.³¹ The broad Raman peaks can also be due to the presence of defects in the system. After Mg-doping, there has been an enhancement in the amount of oxygen present in the sample as observed from EDS. This increase in the oxygen content increased the amount of interstitial oxygen defects, which is detailed in the PL analysis.

Photoluminescence study

The optical properties and defect levels within the band gap of the prepared nanoparticles were studied using photoluminescence (PL) spectroscopy. The deconvoluted PL emission spectra of the synthesized CuO and Mg doped CuO thin films between the wavelength range of 550 to 950 nm is depicted in Fig. 6. All the films showed similar PL spectra with peaks in the green and IR region and a broad emission in the orange-red region. A total of 4 emission bands were observed in the PL spectrum of each sample. The green emission was observed at 565 nm and IR emission at 884 nm. The broad emission at orange-red region were deconvoluted to two peaks centred at 676 and 726 nm.

The emissions in PL spectrum can be intrinsic, due to the transition between conduction band and the valence band which are usually referred to as the band edge emission. Along with this there are several defect related emissions known as the

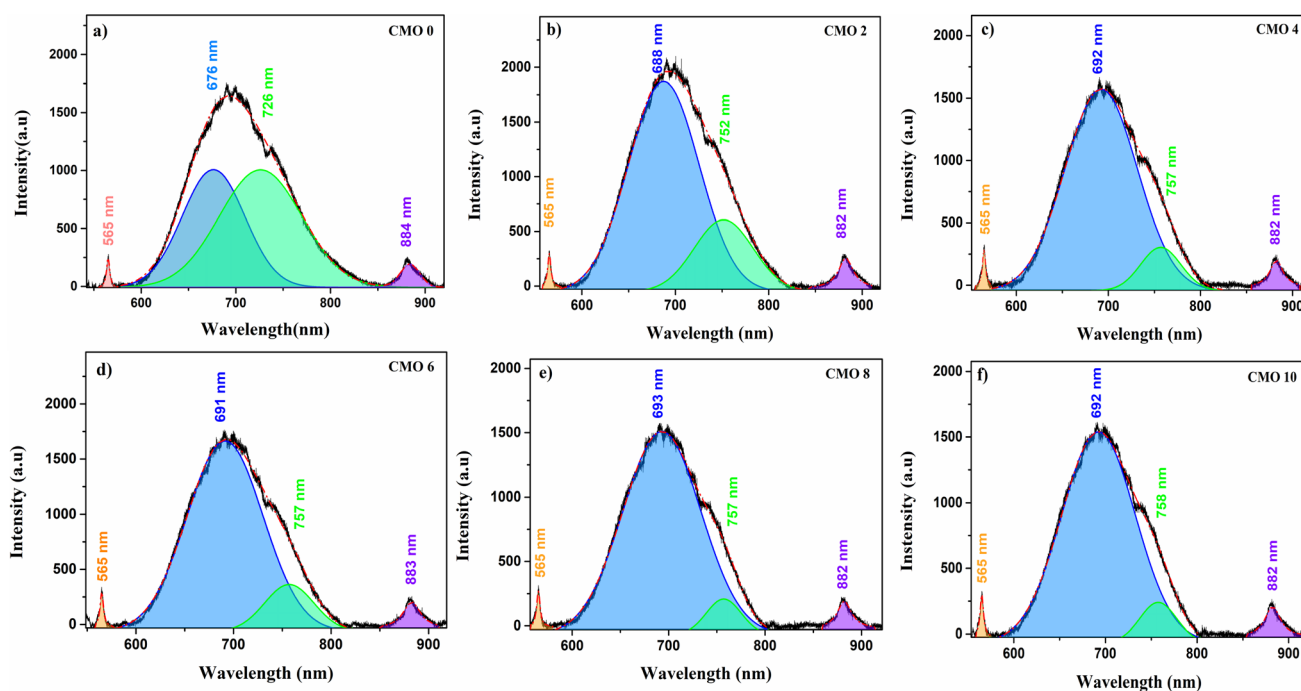


Fig. 6 Photoluminescence spectra of (a) CMO 0, (b) CMO 2, (c) CMO 4, (d) CMO 6, (e) CMO 8 and (f) CMO 10 films.



extrinsic emission. The peak at 884 nm correspond to the band edge emission. The remaining three peaks are defect related emissions that corresponds to the various defects present in the samples. The intrinsic defects in CuO are mainly due to the Cu vacancies (V_{Cu}), Cu interstitials (Cu_i), oxygen vacancies (V_o) or oxygen interstitials (O_i).³² The peak at 565 nm is from the transition between the excited energy level to the V_{Cu} levels.

The red emission at 676 nm is associated with O_i .^{33,34} As generation of V_{Cu} is energetically favoured in the host lattice, a higher concentration of such defects could be expected in CuO nanoparticles. However, V_{Cu} being the most stable defect in CuO, doesn't create changes in the electronic structure and doesn't usually contribute to emissions.³⁵ An enrichment of V_{Cu} thereby enhances the oxygen ions resulting in O_i defects. Both V_{Cu} and O_i act as acceptor levels and enhance the p-type conductivity of CuO.³⁶ The emission at 726 nm corresponds to recombination of electrons from the donor level with the free holes.³⁷ V_o usually acts as the donor level in CuO and hence it can be assumed that this particular emission also corresponds to the oxygen vacancies as well.

The influence of Mg doping in the position and intensity of defect peaks can be observed from Fig. 6. Though the position of both band edge and green emission remained the same, there was an increase in the intensity of these emissions. A significant variation in the intensity and position of defect peaks were observed in the doped samples. While the intensity of both 676 and 726 nm peaks were nearly the same in the case of undoped sample, an enhancement in the intensity of 676 nm peak and a decrease in the intensity of 726 nm peak can be observed with doping. A redshift in the position of these two peaks can also be found. The enhancement in the intensity of 676 nm peak and the redshift of these peaks can be assigned to the increase in the amount of interstitial oxygen defect upon doping.³⁴

To confirm the existence of oxygen at interstitial in Mg doped CuO observed in experimental results, we have done theoretical calculations for the pure and Mg-doped CuO along with O at interstitial positions (see Fig. 7). Formation energies for the pure and Mg-doped CuO along with various kinds of neutral native point defects³⁸ were estimated and observed results have been plotted as functions of the Fermi energy. In the calculations, the zero of the energy was set at the top of the valence

band. From the figure, it is clearly observed that interstitial oxygen is the most stable defect with increase in Mg doping as compared to other native defects.

In isovalent doping, most of the changes in the properties of the films are caused due to lattice deformations introduced by the dopant atom. When dopant atom replaces the metal cations, the surface charge density varies causing a change in the energy of the ionic bond (E_i). Usually, E_i increases as the size of cation decreases. This causes a reduction in the oxygen vacancies.³⁹ The presence of O_i causes an expansion in the CuO lattice which can lead to the redshift observed. This may be the reason for the increase in the a and c lattice parameters as found from XRD. Area ratio of O_i to V_o PL emission peaks were calculated and is given in the Table 4. The O_i/V_o is maximum for CMO 8 indicating that CMO 8 has the maximum O_i defects.

X-ray photoelectron spectroscopy

The chemical phase of the prepared film can be further confirmed by identifying the valence state of Cu and other elements present in the sample. The chemical state of the elements present in the film was identified using XPS. The XPS survey spectrum of CMO 8 sample is illustrated in Fig. 8. The survey spectrum included 2p peaks of Cu and Mg along with 1s peaks of O and C. Additional minor peaks corresponding to Cu 3s, 3p and Cu LMM auger peaks were also observed in the spectrum indicating the presence of these elements in the sample. Any shift in the binding energy due to the surface charging is eliminated by carbon correction at 284.8 eV.⁴⁰ The

Table 4 The estimated optical band gap value from Tauc plot and ratio of area of oxygen interstitial (O_i) to oxygen vacancy (V_o) peak in PL for CMO thin films

Sample	E_g (eV)	O_i/V_o
CMO 0	1.50	0.8
CMO 2	1.57	3.5
CMO 4	1.67	8.3
CMO 6	1.69	6.2
CMO 8	1.65	12.8
CMO 10	1.63	10.6

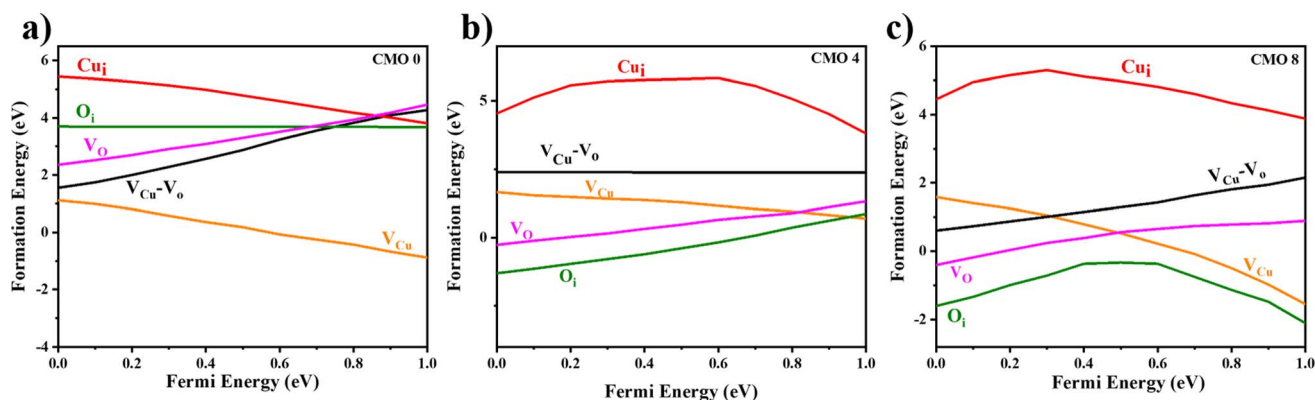


Fig. 7 Calculated formation energy of CuO (a) without doping (b) with ~4 at% Mg doping (c) with ~8 at% Mg doping.



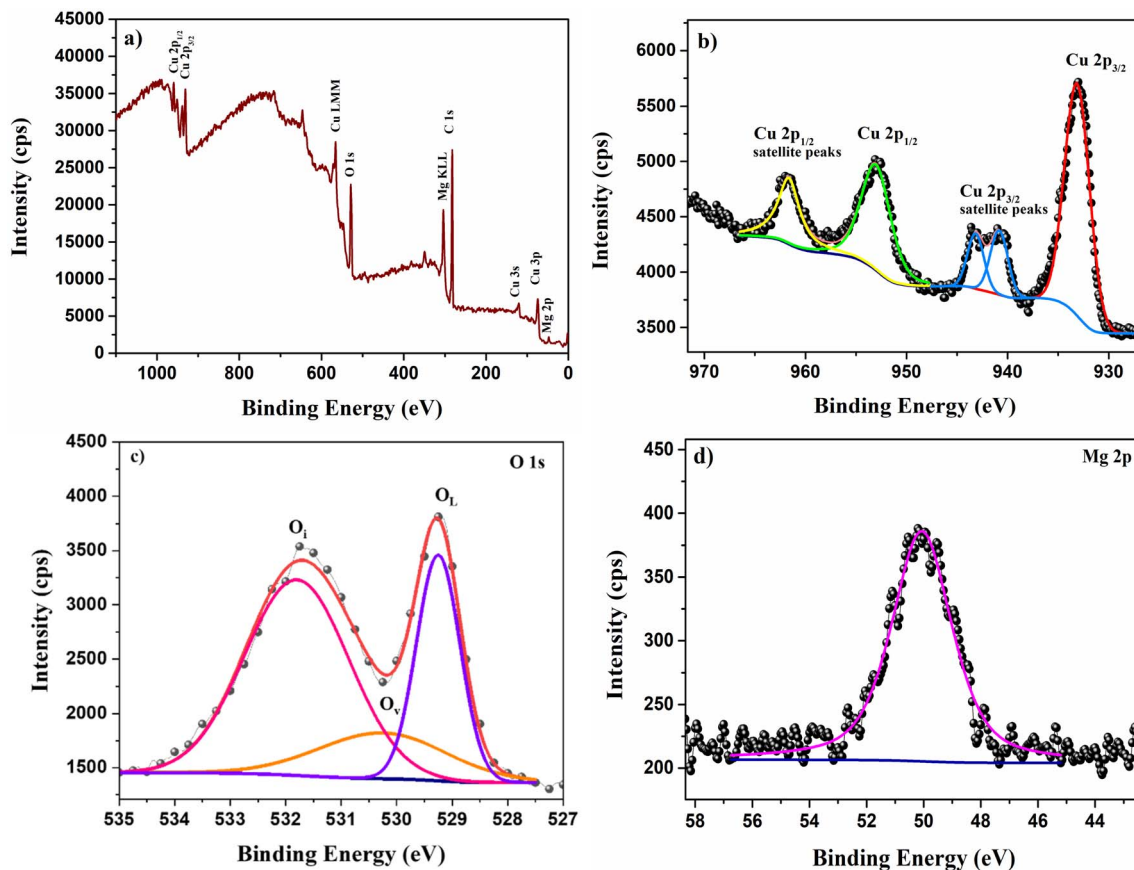


Fig. 8 XPS (a) survey and core spectra of (b) Cu 2p, (c) O 1s, (d) Mg 2p subshells.

peaks present in the core spectrum of each element were deconvoluted and fitted using the Gaussian function to separate different species.

The fitted profile of the high resolution Cu 2p photoemission spectrum of CMO 8 film (Fig. 8(b)) displayed two prominent peaks centred around 953.02 and 933.12 eV. These two peaks with a peak separation of 19.9 eV is attributed to the spin orbit splitting of 2p orbital state of Cu and is in good agreement with the reported values in the literature.⁴¹ The peak at lower binding energy corresponds to Cu 2p_{3/2} and the one at the higher binding energy represents Cu 2p_{1/2}. Furthermore, strong shake up peaks were observed in the higher binding energy region of both Cu 2p_{3/2} and Cu 2p_{1/2}. Copper has two stable oxidation states: Cu⁺ and Cu²⁺ and the presence of shakeup peaks in the higher binding energy region is a clear indication of +2 oxidation state of Cu.⁴² Satellite peaks are characteristics of the materials having d⁹ configuration in the ground state.⁴³ The two shake up peaks at 940.84 and 943.16 eV arise due to the multiplet splitting in 2p⁵ 3d⁹ configuration.⁴⁴ These satellite peaks are not observed in the Cu⁺ configuration as the completely filled d¹⁰ configuration does not allow screening *via* charge transfer to the d states.⁴⁴ However, Mg doping has resulted in a shift in the peak positions towards lower binding energy. When Mg replaces Cu atoms from the CuO lattice, the difference in the electronegativity creates variation in the electron density cloud of Cu²⁺. As the electronegativity of Mg is

lesser than Cu, the electron density cloud around Cu²⁺ increases with doping causing a decrease in the binding energy.¹⁷

The deconvoluted O 1s core level spectrum is shown in Fig. 8(c). It consists of three peaks located at 529.25, 530.25 and 531.80 eV. The peak at lower binding energy (P1) corresponds to the lattice oxygen and the peak at the middle binding energy (P2) is attributed to the oxygen vacancies.⁴⁵ The peak at higher binding energy (P3) is associated with interstitial oxygen.^{45,46} This indicates the co-existence of both O_i and V_o which is also observed in PL spectra. As observed in PL, the area ratio of O_i/V_o is greater than 1 indicating that O_i is the dominant defect in the prepared films which is in line with the PL data. One interesting observation is that the oxygen defect peak corresponding to O_i has a significant area greater than the lattice oxygen itself. Such enhancement in the defect peaks was also observed by M. Nesa *et al.*⁴⁷ with Zn doping. Fig. 8(d) displays the fitted profile of Mg 2p containing only a single peak centred at 50.07 eV. This peak is due to the +2 oxidation state of Mg.⁴⁸ No other peaks corresponding to any other phase were observed in the profile indicating that Mg²⁺ ions have substituted Cu²⁺ ions from its lattice, which is backed by both EDS as well as XRD observations.

The percentage composition of the elements present in the sample was quantified from XPS spectra using the relative sensitive factor approach. In this method, we assume that the intensity of the signals corresponding to a given element is proportional to its concentration.⁴⁹



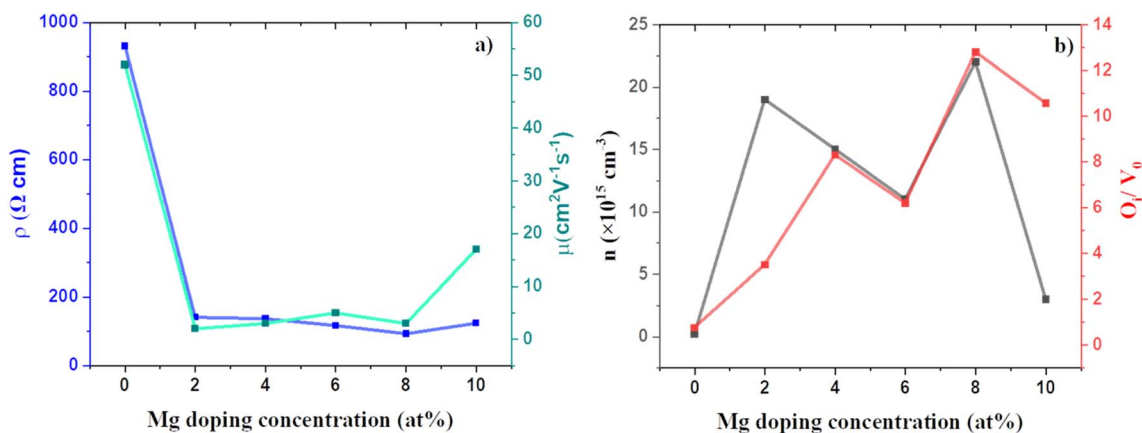


Fig. 9 Variation of (a) electrical resistivity, mobility and (b) carrier concentration, area ratio of O_i to V_o PL peaks of CMO films with doping.

$$I_i = \alpha_i x_i \quad (9)$$

where x_i is the atomic fraction and α_i is the sensitivity factor of the i th element. So, the concentration of i th element is given by,

$$x_i = \frac{\frac{I_i}{\alpha_i}}{\sum_k \left(\frac{I_k}{\alpha_k} \right)} \quad (10)$$

Thus, the calculated amount of Cu, O and Mg present in CMO 8 sample is found to be 36, 60 and 4% respectively. These values are in good agreement with those obtained from EDS results. The higher percentage of O in the samples indicates that our films are oxygen rich in nature.

Electrical properties

The effect of Mg doping on the electrical properties of CuO thin films was studied using Hall measurements. Fig. 9 shows the variation in the obtained electrical properties of doped and undoped films. All the films showed p-type conductivity. The presence of V_{Cu} is the primary reason for the p-type conductivity in CuO thin films.⁵⁰ Mg doping has resulted in a considerable improvement in the conductivity and carrier concentration of CuO films with CMO 8 showing the maximum conductivity and carrier concentration. The carrier concentration of the films was increased by an order of two after doping. The increase in conductivity and carrier concentration can be associated with the defects introduced in the films after doping, especially the oxygen interstitials. Enhancement in the oxygen interstitials increases the p-type conductivity.^{34–36} The increase in oxygen interstitials with doping is substantiated by the enhanced oxygen content from EDS. Further, the enhanced O_i peaks in PL and the XPS also confirm this observation. Fig. 9(b) represents the variation of carrier density and PL peak area ratio of O_i/V_o with the change in Mg doping concentration. The identical trend in the variation of carrier concentration and O_i/V_o elucidates this argument. On the other hand, the mobility of the films showed an inverse trend. The mobility of the films has declined which is due to the smaller crystallite size. With

doping, the carrier concentration has increased, and the grain size has decreased. The change in grain size has influenced the mobility significantly. To see the change in the electronic properties with Mg doping in CuO we have also calculated the total density of states for (a) pure CuO, (b) 4 at% Mg doped CuO and, (c) 8 at% Mg doped CuO (see Fig. S4†). Electronic bandgap of pristine were observed at 1.50 eV and increases to 1.67 eV in case of 4 at% Mg doping while this decreases to 1.65 eV in case of 8 at% Mg doping in CuO. Variation in bandgap is due to structural disorder in case of 8 at% Mg doping and hence it is more favouring O_i and also found in the formation energy calculations.

Conclusion

The influence of Mg doping on the optoelectronic properties of spray deposited CuO thin films has been investigated. CuO and Mg doped CuO thin film were synthesized over a glass substrate by spray pyrolysis method. Mg doping has resulted in an enrichment of oxygen atoms in the films which has affected structural, optical, and electrical properties. The EDS result revealed the presence of oxygen in excess. The detailed analysis of PL and XPS data has shown that this excess oxygen has remained as interstitials which leads to the lattice expansion. Mg^{2+} ions, being smaller than Cu^{2+} , were easily incorporated into the CuO lattice without affecting its monoclinic structure. However, the presence of O_i has resulted in the expansion of lattice parameters as observed from XRD results. The crystallite size of the films was observed to be decreased and the band gap of the films widened with doping. Mg doping has improved the conductivity and carrier concentration of the films with CMO 8 showing the least resistivity and maximum carrier concentration. CMO 8 showed the maximum O_i/V_o ratio in PL which explains the maximum conductivity and carrier concentration of this film. Extensive DFT calculations confirms the stability of O_i defects in Mg doped CuO thin films. Presence of O_i in Mg doped CuO shows significant changes in emission and absorption spectra, as well as it also helps to improve the electrical properties. To conclude, Mg doping can be used to improve the electrical properties of CuO films, by enhancing the



O_i defects, without compromising their good optical properties and thus can be used for improving the performance of CuO based solar cells.

Author contributions

Adithya Prakash: conceptualization, investigation, formal analysis, methodology, visualization, writing – original draft. Vikash Mishra: methodology (DFT calculations). M. G. Mahesha: conceptualization, supervision, writing – review & editing.

Conflicts of interest

Authors have no conflict of interest to declare.

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