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Application of a distinctly bent, trinuclear, end-to-end azide bridged, mixed valence cobalt(III/II/III) complex in the fabrication of photosensitive Schottky barrier diodes†

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A mixed-valence trinuclear cobalt(III)-cobalt(II)-cobalt(III) complex, $[(\mu-1,3-N_3)Co_3L(N_3)_3] \cdot MeOH$ has been synthesized using a tetradentate N_2O_2 donor 'reduced Schiff base' ligand, H_2L {1,3-bis(2-hydroxybenzylamino)2,2-dimethylpropane} and azide as anionic co-ligand. The complex has been characterised by elemental analysis, IR, UV-vis spectroscopy and single-crystal X-ray diffraction studies etc. The cobalt(III)-cobalt(II)-cobalt(III) skeleton in the complex is non-linear and non-centrosymmetric. The redox behavior of the complex was studied by using Cyclic Voltammetry (CV). The complex is found to be a semiconductor material as confirmed by determining the band gap of this complex by experimental as well as theoretical studies. The band gap in the solid state has been determined experimentally. The conductivity of the synthesized complex based device improves considerably in illumination conditions from the non-illuminated conditions. The complex has also been used to fabricate Schottky barrier diodes.

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Introduction

Many di- and poly-nuclear complexes of transition and non-transition metals have interesting photochemical and photo-physical properties, with potential applications in light-emitting diodes, luminescent probes, photovoltaic devices, solar cells, memory devices, field effect transistors *etc.*^{1–8} If a compound of a transition or a non-transition metal shows good semiconducting properties, along with a rectifying nature in a metal–semiconductor junction, it may be used to fabricate a Schottky diode, which is a well-known opto-electronic device. In the last few years, our group has synthesized and characterized a lot of compounds of transition and non-transition metals

showing notable semiconducting properties.^{9–21} These semiconducting materials have been utilized for the fabrication of different opto-electronic devices.^{9–21} Different N,O-donor Schiff bases and their reduced analogues have been used as ligands for the synthesis of these compounds.

Focusing on cobalt, carboxylate-bridged trinuclear mixed valence complexes are very common.^{22–38} In these complexes, the central cobalt is in the +2 state and is present in the O_6 donor environment, whereas terminal cobalt centers are in the +3 state and are present in N_3O_3 or N_2O_4 donor environments. The Co(III)–Co(II)–Co(III) angle is not less than 150° (if not more) in most of the cases and the complexes may be considered as linear trinuclear Co(III)–Co(II)–Co(III) complexes.^{22–38} The bridging carboxylates are *trans* to one another in most of the cases,^{17,28–32,36–38} although *cis* orientations are also found in some cases.^{23,24,28,39} Few such complexes were also used to fabricate opto-electronic devices.^{13,14,40,41} In the present work, we have synthesized a distinctly bent trinuclear mixed valence cobalt complex, where the terminal cobalt centers are bridged by an end-to-end azide group. The structure of the complex has been confirmed by single crystal X-ray diffraction analysis. The band gap of the synthesized material in the solid state has been determined by experimental measurements and compared with the theoretical value obtained from DFT calculations. The band gap indicates that the complex belongs to the semiconductor family. The conductivities of such complexes have not been explored widely. We therefore concentrated on the electric

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conductivity of the complex and aimed to correlate the difference in the conductivity values of the complex with the other reported complexes. This observation leads us to explore the potential of this Co(III)–Co(II)–Co(III) complex in fabricating a photosensitive Schottky barrier diode.

Herein, we report the synthesis and characterization of a mixed valence cobalt complex, $[(\mu-1,3-N_3)Co_3L(N_3)_3] \cdot MeOH$ and its application in opto-electronics.

Experimental section

Materials

All starting materials were commercially available, reagent grade and used as purchased from Sigma-Aldrich. The reactions and all manipulations were carried out under aerobic conditions.

Caution!!! Even though no troubles were experienced in this work, azide salts and perchlorate salts containing organic ligands are potentially explosive. Only a small quantity of the material should be prepared and they should be handled with care.

Synthesis of Schiff base ligand, H_2L^a {2,2'-[(2,2-dimethyl-1,3-propanediyl)bis(nitrilomethylidene)]bis[phenol]}. The N_2O_2 donor Schiff base ligand, H_2L^a , was prepared by refluxing 2,2-dimethylpropane-1,3-diamine (0.2 mL, ~2 mmol) with salicylaldehyde (0.4 mL, ~4 mmol) in methanol (10 mL) for *ca.* 1 h. To obtain the appropriate reduced counterparts of these Schiff base, the ligand was reduced using sodium borohydride (*vide infra*).

Synthesis of the reduced Schiff base ligand, H_2L {1,3-bis(2-hydroxybenzylamino)2,2-dimethylpropane}. The excess solid sodium borohydride was added while stirring continuously to the methanol solution of the Schiff base, H_2L^a , which had been cooled to 0 °C. The resulting solution was acidified with glacial acetic acid and placed under reduced pressure in a rotary evaporator at ~65 °C. The residue was dissolved in 5 mL of water, and then 5 mL of dichloromethane was used to extract it. If there was any extra acid, a small amount of sodium hydrogen carbonate was added to neutralize it. Finally, the ligand solution in dichloromethane was dried using anhydrous sodium acetate and the reduced Schiff base ligand, H_2L was collected in methanol. The ligand was not isolated, but were used *in situ* for the preparation of the complex.

Synthesis of the complex

Synthesis of $[(\mu-1,3-N_3)Co_3L(N_3)_3] \cdot MeOH$. A solution of cobalt(II) perchlorate hexahydrate (1.1 g, ~3 mmol) in methanol

(10 mL) was added to the methanol solution of the reduced Schiff base ligand, H_2L with constant stirring to obtain a dark brown solution. A solution of sodium azide (0.26 g, ~4 mmol) in 10 mL of methanol/water (in a 4 : 1 ratio) was then added to the dark brown solution, with constant stirring. The stirring was continued for *ca.* 3 h. After 8–9 days of slowly evaporating the solution in an open atmosphere, the complex's black crystals were produced. From the product, single crystals that may be used for X-ray diffraction were recovered.

Yield: 0.601 g (~61%, based on cobalt), anal. calc. for $C_{38}H_{50}Co_3N_{16}O_5 \cdot CH_3OH$ (FW = 1001.75): C, 45.93; H, 5.34; N, 21.98%; found: C, 46.19; H, 5.11; N, 22.21%. FT-IR (KBr, cm^{-1}): 3394 (ν_{O-H}), 3236 (ν_{N-H}), 2874–2922 (ν_{C-H}), 2012 (bridging $\nu_{N_3^-}$)–2081 (terminal $\nu_{N_3^-}$); UV-vis, $\lambda_{max}(nm)$, [ϵ_{max} (lit $mol^{-1} cm^{-1}$)] (acetonitrile): 640 (5.4×10^2), 357 (1.4×10^3), 265 (1.8×10^3) and 232 (3.0×10^3). Magnetic moment; $\mu = 5.1$ B.M.; crystal system: orthorhombic.

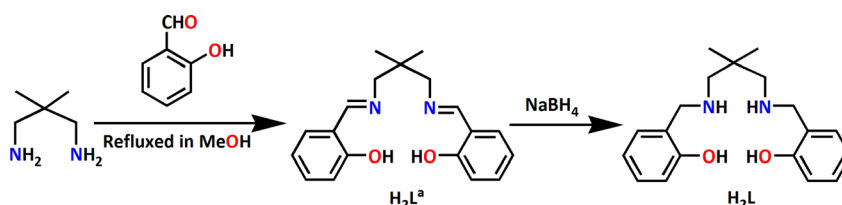
Instrumentation

The details of different instruments are given in the ESI.† Details of the device fabrication and electrical characterization are also added the ESI.†

Results and discussion

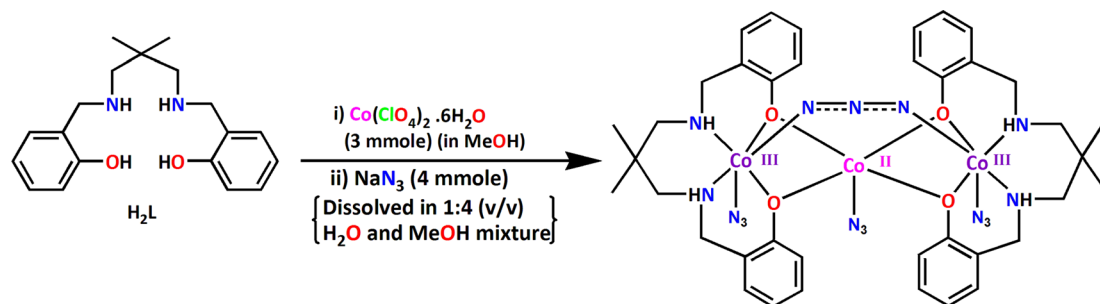
Synthesis

In the present work, the Schiff base ligand, H_2L^a {2,2'-[(2,2-dimethyl-1,3-propanediyl)bis(nitrilomethylidene)]bis[phenol]} has been synthesized by the 1:2 condensation of 2,2-dimethylpropane-1,3-diamine with salicylaldehyde following the literature procedure.¹⁷ The Schiff base has been reduced with excess sodium borohydride to get corresponding reduced Schiff base ligands, H_2L {2,2'-[(2,2-dimethyl-1,3-propanediyl)bis(iminomethylene)]bis[phenol]} (Scheme 1). However, the ligand is not separated or purified; rather, this is utilized in the synthesis of the cobalt complex. The reaction of cobalt(II) perchlorate hexahydrate, sodium azide with H_2L in 3 : 4 : 2 molar ratio in methanol solvent produce $[(\mu-1,3-N_3)Co_3L(N_3)_3] \cdot MeOH$. One of the important observation is that the reaction of cobalt(II) perchlorate hexahydrate, sodium azide and H_2L in 3 : 2 : 2 molar ratio in methanol solvent gave similar results. The complex is a tri-nuclear mixed valence cobalt(III,II,III) complex. The precursor cobalt(II) is partially oxidized by atmospheric oxygen during the course of the reaction. The cobalt(II) is stabilized in weaker NO_4 donor ligand field environment, whereas the cobalt(III) is stabilized in relatively strong field N_4O_2 donor environment. The synthetic route to the complex is shown in Scheme 2.



Scheme 1 Synthesis of the reduced Schiff base ligand, H_2L .





Scheme 2 Synthetic route to the complex. Lattice methanol molecule is not shown for clarity.

Description of crystal structure

The complex belongs to the class of mixed valence trinuclear cobalt(III)-cobalt(II)-cobalt(III) complexes, where the cobalt(II) and the two cobalt(III) are present in the central and the terminal positions, respectively. Charge balance considerations, inter-atomic cobalt-nitrogen, and cobalt oxygen distances were used to confirm the oxidation states of cobalt(III) and cobalt(II). For the central and two terminal cobalt ions, respectively, BVS calculations also support +2 and +3 charges (see Table S1 in ESI†). The detailed structural description of the complex is provided below.

$[(\mu-1,3-N_3)Co_3L(N_3)_3] \cdot MeOH$. X-ray single crystal structure determination reveals that the complex crystallizes in orthorhombic space group, $Pca2_1$ and is non-centrosymmetric. The ORTEP presentation of the complex including the selective atom labeling scheme is shown in Fig. 1. The complex is trinuclear, mixed valence, cobalt(III/II/III) complex. The central cobalt centre, Co(1), is in +II oxidation state, whereas the two terminal cobalt centres, Co(2) and Co(3), are in +III oxidation states. Central cobalt(II) centre is linked with two cobalt(III) centers by four phenoxo oxygen atoms of two $(L)^{2-}$, forming a bis{di(μ -phenoxo)} bridged cobalt(III)-cobalt(II)-cobalt(III) motif. The cobalt(III)-cobalt(II)-cobalt(III) skeleton is imposed to be non-linear. A summary of the crystallographic data and refinement details of the complex is given in Table 1. Important bond lengths and angles are listed in Tables 2 and 3, respectively.

Each of the two terminal cobalt(III) centers Co(2) and Co(3), has a six-coordinate pseudo-octahedral geometry. For Co(2) centre, oxygen atoms, O(1), O(2) and nitrogen atoms, N(1) and N(2) of the deprotonated ligand $(L)^{2-}$ constitute the equatorial plane, and a nitrogen atom, N(5) of the bridging azide and another nitrogen atom, N(11) of a terminal azide, define the axial sites. Similarly, for Co(3) centre, oxygen atoms, O(3), O(4) and nitrogen atoms, N(3) and N(4) of the deprotonated ligand $(L)^{2-}$ constitute the equatorial plane, and a nitrogen atom, N(7) of the end-to-end bridging azide and another nitrogen atom, N(14) of a terminal azide, define the axial positions. On the other hand, the central cobalt(II) center, Co(1), is five coordinated being bonded to four phenoxo oxygen atoms {O(1), O(2), O(3) and O(4)} and one terminal azide nitrogen atom {N(8)}. The geometry of this cobalt(II) center is best described as an intermediate of trigonal bipyramidal and square pyramidal geometry, as confirmed by calculating Addison parameter, $\tau = 0.49$, $\beta = \angle O(2)-Co(1)-O(3) = 151.5^\circ$ and $\alpha = \angle O(1)-Co(1)-N(8) = 122.3^\circ$, where α and β are the two largest L-L-L angles in the coordination sphere of the metal.⁴² The perspective view of CoO_5N moiety with the largest L-M-L angles is shown in Fig. 2. Only one similar complex is reported in literature, which has been synthesized using 1,3-diaminopropane instead of 2,2-dimethyl-1,3-diaminopropane.⁴³

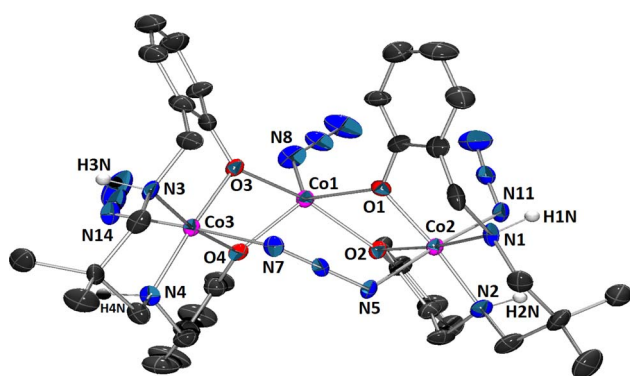


Fig. 1 ORTEP presentation of the complex (ellipsoids are drawn at the 30% probability level) with selective atom-numbering scheme. Hydrogen atoms (except those attached to nitrogen atoms) have been omitted for clarity.

Table 1 Crystal data and refinement details of the complex

Formula	$C_{38}H_{50}Co_3N_{16}O_5 \cdot CH_3OH$
Formula weight	1001.75
Temperature (K)	273
Crystal system	Orthorhombic
Space group	$Pca2_1$
a (Å)	21.066(8)
b (Å)	12.519(7)
c (Å)	17.232(8)
Z	4
d_{calc} (g cm ⁻³)	1.464
μ (mm ⁻¹)	1.143
$F(000)$	2082
Total reflections	32 229
Unique reflections	6971
Observed data [$I > 2 \sigma(I)$]	5047
No. of parameters	550
$R(int)$	0.143
R_1, wR_2 (all data)	0.0976, 0.1059
R_1, wR_2 [$I > 2 \sigma(I)$]	0.0604, 0.0913

Table 2 Selected bond lengths (Å) of the complex

Co(1)–O(1)	2.012(5)
Co(1)–O(2)	2.092(5)
Co(1)–O(3)	2.085(5)
Co(1)–O(4)	2.012(5)
Co(1)–N(8)	1.973(7)
Co(2)–O(1)	1.904(5)
Co(2)–O(2)	1.956(5)
Co(2)–N(1)	1.981(6)
Co(2)–N(2)	1.968(6)
Co(2)–N(5)	2.004(6)
Co(2)–N(11)	1.916(7)
Co(3)–O(3)	1.905(5)
Co(3)–O(4)	1.930(5)
Co(3)–N(3)	1.979(7)
Co(3)–N(4)	1.970(6)
Co(3)–N(7)	1.985(7)
Co(3)–N(14)	1.918(7)

Table 3 Selected bond angles (°) of the complex

O(1)–Co(1)–O(2)	73.62(19)	N(1)–Co(2)–N(5)	90.3(2)
O(1)–Co(1)–O(3)	91.88(19)	N(1)–Co(2)–N(11)	88.4(3)
O(1)–Co(1)–O(4)	117.5(2)	N(2)–Co(2)–N(5)	88.6(3)
O(1)–Co(1)–N(8)	122.3(3)	N(2)–Co(2)–N(11)	86.8(3)
O(2)–Co(1)–O(3)	152.0(2)	N(5)–Co(2)–N(11)	175.1(3)
O(2)–Co(1)–O(4)	92.41(19)	O(3)–Co(3)–O(4)	78.7(2)
O(2)–Co(1)–N(8)	108.3(2)	O(3)–Co(3)–N(3)	92.4(2)
O(3)–Co(1)–O(4)	72.79(19)	O(3)–Co(3)–N(4)	171.6(2)
O(3)–Co(1)–N(8)	99.8(2)	O(3)–Co(3)–N(7)	91.9(2)
O(4)–Co(1)–N(8)	120.0(3)	O(3)–Co(3)–N(14)	93.0(2)
O(1)–Co(2)–O(2)	79.2(2)	O(4)–Co(3)–N(3)	170.7(2)
O(1)–Co(2)–N(1)	93.2(2)	O(4)–Co(3)–N(4)	93.1(2)
O(1)–Co(2)–N(2)	171.2(2)	O(4)–Co(3)–N(7)	89.5(2)
O(1)–Co(2)–N(5)	90.5(2)	O(4)–Co(3)–N(14)	94.9(2)
O(1)–Co(2)–N(11)	94.3(3)	N(3)–Co(3)–N(4)	95.9(2)
O(2)–Co(2)–N(1)	172.4(2)	N(3)–Co(3)–N(7)	88.3(3)
O(2)–Co(2)–N(2)	92.1(2)	N(3)–Co(3)–N(14)	88.0(3)
O(2)–Co(2)–N(5)	90.2(2)	N(4)–Co(3)–N(7)	90.0(2)
O(2)–Co(2)–N(11)	91.8(2)	N(4)–Co(3)–N(14)	85.7(2)
N(1)–Co(2)–N(2)	95.6(2)	N(7)–Co(3)–N(14)	174.0(3)

The Co(2)⋯Co(1) distances are not sufficiently short enough {3.093(3) Å} to imply any cobalt–cobalt bonding.^{23,39} The bridging angles, \angle Co(1)–O(1)–Co(2) = 104.3(4)°, \angle Co(1)–O(2)–Co(2) = 99.8(4)°, \angle Co(1)–O(3)–Co(3) = 100.9(3)° and \angle Co(1)–O(4)–Co(3) = 102.9(4)° indicate that the cobalt(III)–O–cobalt(II)

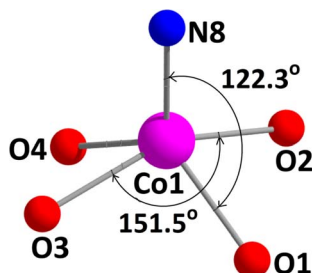
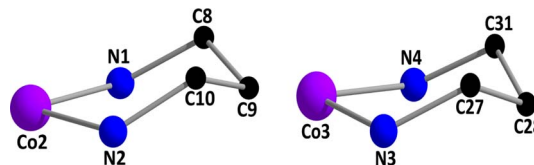
Fig. 2 The perspective view of CoO₅N moiety showing the largest L–M–L angles in the coordination sphere.

Fig. 3 Chair conformations of two saturated six-membered chelate ring in the complex. Only selected atoms have been shown.

fragments are non-collinear. Cobalt(II)–O distances vary from 2.013(9) to 2.099(8) Å (Table 2), values equivalent to those observed in other high-spin cobalt(II) complexes.²² The cobalt(III)–oxygen and cobalt(III)–nitrogen distances (Table 2) are in the range 1.897(8)–1.943(10) and 1.907(11)–1.989(12) Å, respectively, which are typical for low-spin cobalt(III).²² Both the saturated six membered chelate rings, Co(2)–N(1)–C(8)–C(9)–C(12)–N(2) and Co(3)–N(3)–C(27)–C(28)–C(31)–N(4), represent chair conformations (Fig. 3) with puckering parameters,⁴⁴ $Q = 0.540(15)$ Å, $\theta = 18.6(14)^\circ$, $\varphi(2) = 182(4)^\circ$ and $Q = 0.551(13)$ Å, $\theta = 161.7(12)^\circ$, $\varphi(2) = 359(4)^\circ$, respectively.

IR and electronic spectra

The IR and electronic spectra of the complex are in excellent concurrence with its crystal structure. A distinct band is observed in the IR spectrum of the complex at 3394 cm^{−1} due to O–H stretching vibrations of lattice methanol molecule.⁴³ A moderately sharp band around 3236 cm^{−1} corresponding to the presence of N–H stretching vibrations and absence of any peak around ~1600 cm^{−1} (corresponding to imine vibration) indicates that the azomethine groups of the Schiff base has been reduced prior to the complex formation.^{17,36} The band corresponding to the alkyl C–H stretching can be observed in the range of 2874–2922 cm^{−1} for the complex.^{17,23,36,39} A very strong double band observed at 2012 and 2081 cm^{−1} confirms the presence of two types of azide groups (bridging and terminal, respectively) in the complex.⁴⁵ The IR spectrum of the complex is given in Fig. 4.

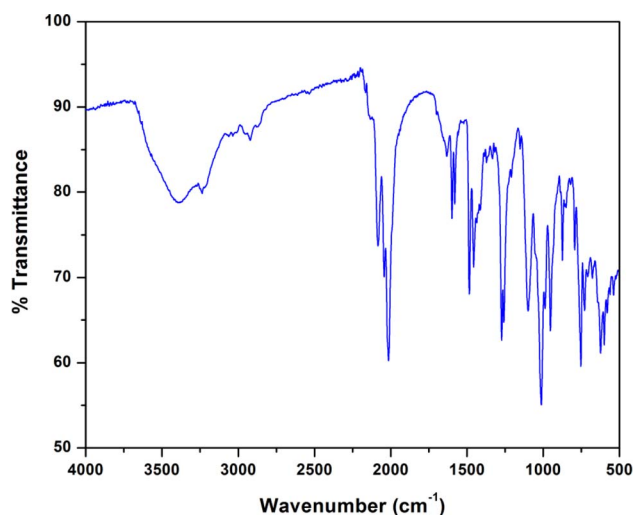


Fig. 4 IR spectrum of the complex.

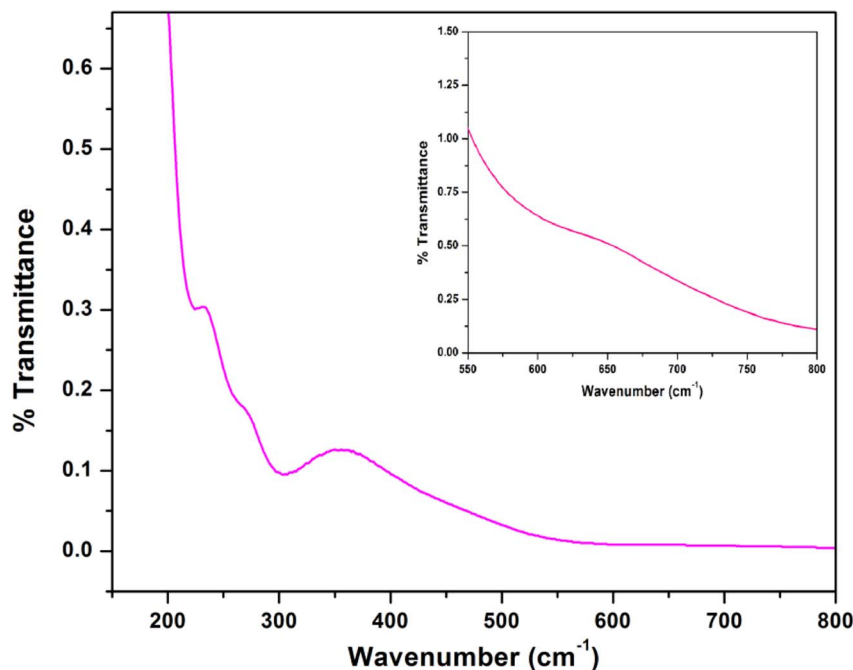


Fig. 5 Electronic spectrum of the complex in acetonitrile. Inset shows the absorption band for low spin cobalt(III) center in the complex.

The electronic spectrum of the complex in acetonitrile medium consists of three most fundamental bands for the complex. The high-energy, intense bands around 265 nm with hump (232 nm) for the complex, may be due to intra-ligand π - π^* transition.^{23,39} Moreover, the intense absorption band at 357 nm in the complex is assigned to charge transfer transitions from the coordinated ligands to the cobalt(II) centers (LMCT).^{23,39} A weak band around 640 nm for the complex may be assigned as one of the two expected transitions for cobalt(II) complex in octahedral geometry.³¹ The transitions from

cobalt(III) may get obscured by other bands. The electronic spectrum of the complex is given in Fig. 5.

Magnetic properties

The room temperature magnetic moment of the complex is 5.1 BM, which suggests that the complex contains just one high-spin cobalt(II) centre. This data, therefore, agree well with the representation of the complex as $\text{Co(III)}(S = 0)\text{-Co(II)}(S = 3/2)\text{-Co(III)}(S = 0)$ trimer.^{23,39} The calculation of the spin density on

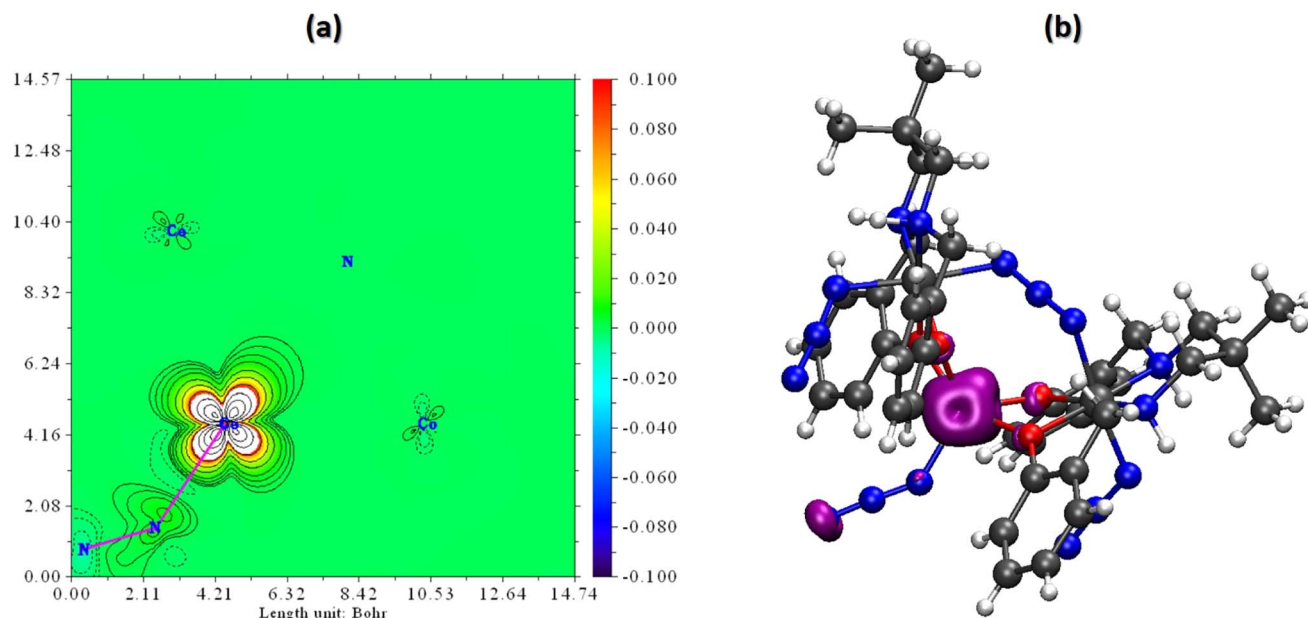


Fig. 6 (a) Spin density plot and (b) spin density fill-contour map plot in a plane of the complex. The red zone clearly indicates the presence of unpaired electrons.

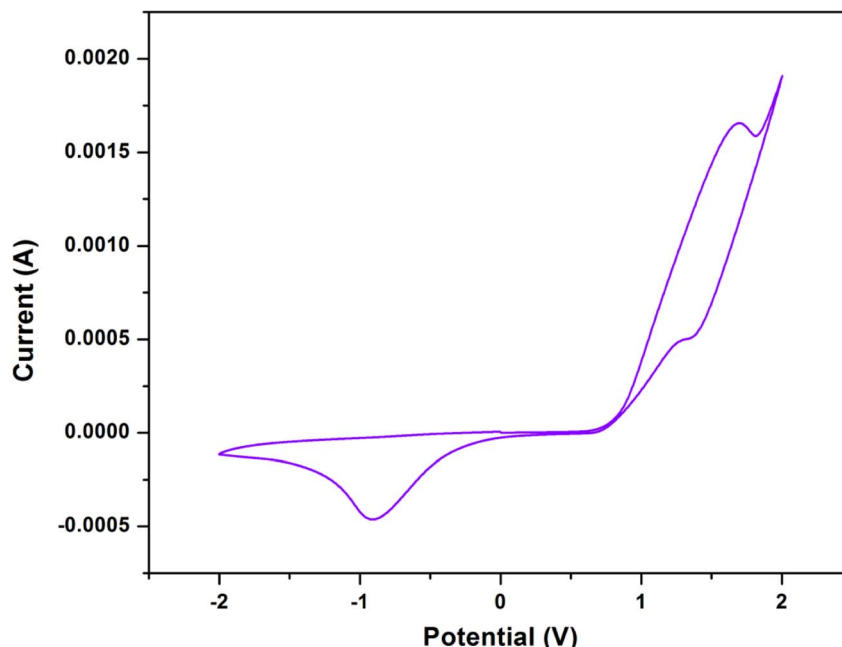


Fig. 7 Cyclic voltammogram of the complex in DMF (using Ag/AgCl reference electrode).

Table 4 Cyclic voltammetry data (V) for the complex in DMF medium^a

Scan rate (mV s ⁻¹)	E_{pa} (V) (Co ^{II} → Co ^{III})	E_{pc} (V) (Co ^{III} → Co ^{II})	$E_{1/2}$ (V) (Co ^{II} ↔ Co ^{III})	ΔE_p (Co ^{II} ↔ Co ^{III})	E_{pc} (V) (Co ^{III} → Co ^{II})
200	+1.69	+1.34	+1.51	+0.35	−0.92

^a $E_{1/2}$ denotes the half-wave potential. $E_{1/2} = (E_{pa} + E_{pc})/2$ and $\Delta E_p = (E_{pa} - E_{pc})$.

the cobalt centres in each complex also supports this assignment of oxidation states (*vide infra*).

Spin density

Spin density serves as a metric for gauging the spatial distribution of the disparity between the number of alpha (spin-up) and beta (spin-down) electrons within a quantum system. Essentially, it quantifies the degree to which a specific region in a molecule or solid is more inclined to host electrons with one spin orientation over the other. By examining the spin density of the scrutinized complex, we can deduce the localization of unpaired electrons.

The findings reveal that within this complex, the unpaired alpha electrons are predominantly situated around the central cobalt center, linked to four oxygen atoms and the N₃[−] ligand, exhibiting a spin population of 2.68. Meanwhile, the remaining cobalt centers exhibit fully paired electrons, resulting in a spin population of approximately zero. To facilitate a clear visualization of spin density, two plots are presented in Fig. 6 to depict the results. Mulliken Spin population and atomic charge of the investigated complex are given in Table S3 in ESI.†

Electrochemical studies

Using cyclic voltammetry (CV) in DMF with 10^{−3} M of each complex having tetrabutylammonium bromide (TBAB) as

supporting electrolyte, the redox behaviour of the synthesized complex was examined.

Cyclic voltammetric studies of the complex display two one-electron cyclic voltammetric responses. From them, one is quasi-reversible redox signal and another is irreversible redox signal. The quasi-reversible redox signal consists of an oxidation peak, E_{pa} at +1.69 V and a reduction peak, E_{pc} at +1.34 V, whereas the irreversible signal consists only a reduction peak, E_{pc} at −0.92 V. The quasi-reversible redox signal could be attributed to the Co^{II} → Co^{III} oxidation and the Co^{III} → Co^{II} reduction, whereas the irreversible signal could be attributed to the Co^{III} → Co^{II} reduction only. The cyclic voltammograms of the complex is shown in Fig. 7. The cyclic voltammetric results are consistent with those of other trinuclear mixed valence complexes of cobalt that have been reported in the literature.^{17,38,46,47} The detailed electrochemical data for the complex is given in Table 4.

Optical properties

The optical band gap (E_g) of the complex is determined with the help of Tauc's equation, which corresponds to the electronic transition from the valence band to the conduction band.⁴⁸ From Tauc's equation, we used,

$$(\alpha h\nu)^n = A(h\nu - E_g) \quad (1)$$



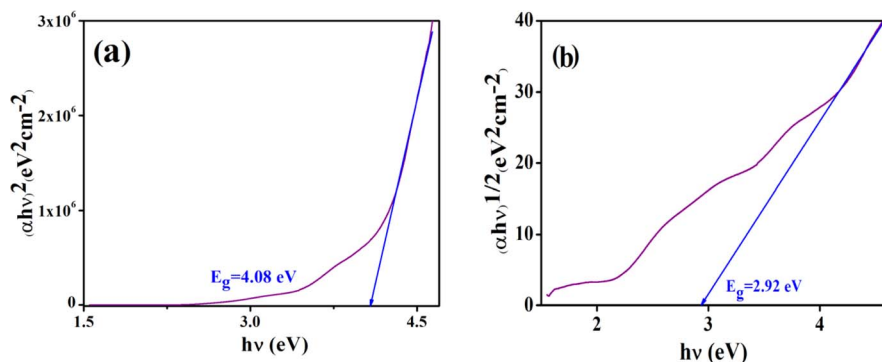


Fig. 8 Tauc plot to evaluate the direct (a) and indirect (b) allowed band gap energy of the complex.

where α is the absorption coefficient, E_g is the optical band gap energy, h is Planck's constant, ν is frequency, A is an arbitrary constant and n is the order of energy transition. The values of n were taken as 2 and 1/2 corresponding to the allowed direct and indirect electronic transitions, respectively. The $(\alpha h\nu)^2$ vs. $h\nu$ plot of the synthesized complex is shown in Fig. 8.⁴⁹ The calculated band gaps of the synthesized complex for direct and indirect electronic transitions are 4.08 eV and 2.92 eV, respectively.

Electrical characterization

The calculated optical band gap offers that material is in the family of a semiconducting in nature. It propelled us to study the semiconducting properties of our material. Hence, we have developed our as-synthesized cobalt based metal (Al)–semiconductor (MS) junction thin film device. The device has been developed following the method⁵⁰ (see ESI†). The current–voltage (I – V) relationship of fabricated devices have been studied in dark as well as illumination (Intensity ~ 100 mW cm^{-2}) conditions at room temperature (300 K) at corresponding applied bias voltage sequentially within the limit ± 1 V to analyze the charge transport behaviour. All the study was carried out under N_2 filled glove box and laboratory temperature of 27 °C.

The I – V characteristic of the cobalt based devices has been recorded under dark and illumination condition (Fig. 9). Under

dark condition, the electrical conductivity of the cobalt based device has been calculated as $1.05 \times 10^{-6} \text{ S m}^{-1}$. However, after exposed under illumination, the conductivity of the same device has been calculated as $2.12 \times 10^{-6} \text{ S m}^{-1}$. It is clear that the conductivity of the synthesized complex based device improves considerably in illumination conditions from the non-illuminated condition.

Besides, the illustrative I – V characteristics of the Al/cobalt based interface under both dark and a photo illumination condition (Fig. 9) represents a nonlinear rectifying behaviour, like Schottky diode (SD). Under photo illumination condition, our material based device exhibits larger current which is the significance of photosensitivity. Hence, we have measured the photosensitivity of our device which has been found to be 0.93.

Here, we have estimated the thermionic emission theory to analyze the obtained I – V characteristics of cobalt based SDs. In this regard Cheung's method is also been employed to extract important diode parameters^{51–54} (see ESI† for detail studies).

The ideality factor (η) and the series resistance (R_s) of our fabricated device have been calculated from the intercept and the slope of $dV/d\ln I$ vs. I plot (Fig. 10) under both conditions. The obtained value of ideality factor of the device both under dark and irradiation conditions has been listed below in Table 5. The value of ideality factor (η) of the fabricated SD has been estimated as 1.17 and 1.11 under dark and illumination condition respectively. The calculated values of ideality factor of the device under both conditions shows a deviation from its ideal value (~ 1). The existence of interface states and series resistance at the junction and presence of inhomogeneities of Schottky barrier height may be the main factor for this non-ideal behaviour.^{54,55}

However, under photo illumination condition the ideality factor of the fabricated SD approaches more ideal (near to the value 1), which is a significant observation. This incident represents the less recombination of interfacial charge carriers and generation of better homogeneity at the barrier of Schottky junctions.⁵¹ From this analysis, it may be concluded that the synthesized cobalt complex based SD possesses fewer number of carrier recombination at the junction *i.e.* better barrier homogeneity even under illumination condition.

We have calculated the value of barrier height (ϕ_B) from the intercept of $H(I)$ vs. I plot (Fig. 10). In this study the barrier

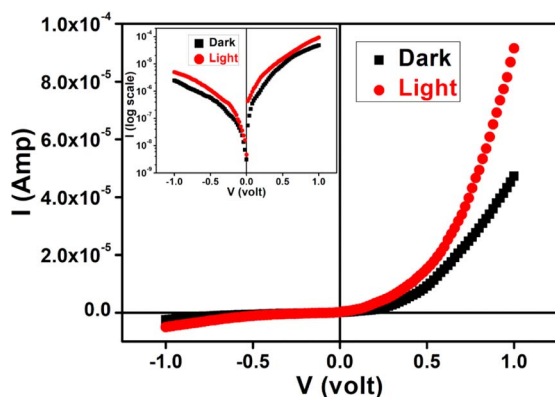


Fig. 9 I – V characteristics curve for cobalt based/Al structured thin film devices under dark and photo illumination condition.

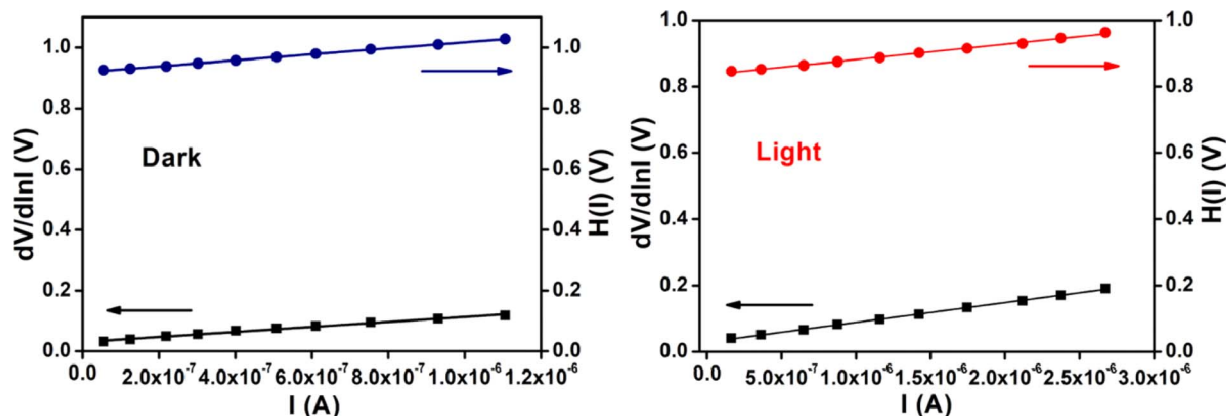


Fig. 10 $dV/d\ln I$ vs. I curve and $H(I)$ vs. I curve for our synthesized cobalt based thin film device under dark, and photo illumination condition.

Table 5 Schottky device parameters of cobalt based SD

Condition	Conductivity ($S\ m^{-1}$)	Photosensitivity	Ideality factor (η)	Barrier height (eV)	Avg. R_s from $dV/d\ln I$ & H ($k\Omega$)
Dark	1.05×10^{-6}	0.93	1.17	0.78	90
Light	2.12×10^{-6}		1.11	0.75	52.5

potential height of the device represents a reduction under irradiation condition, which has the significance. This reduction may be occurred due to the accumulation of the produced photo induced charge carriers near the conduction band. The slope of $H(I)$ vs. I plot also gives the series resistance (R_s). The measured potential height (ϕ_B), ideality factor (η) and series resistance (R_s) under dark and illumination condition for the metal (Al)–semiconductor (cobalt based material) (MS) junctions have been listed in Table 5. The diminution of R_s due to expose under the irradiation condition of our device signifies its applicability in the field of optoelectronics.

For an in-depth analysis of the charge transport phenomena at MS junction, we have further studied the I - V curves. The characteristic I - V curves under both conditions in the logarithmic scale reveal that it can be differentiated in two regions (region I and region II) having different slope values (Fig. 11).

The first region (Region-I), with the slope value is ~ 1 , symbolizes to the ohmic regime. Here the generated current follows the relation $I \propto V$. In which portion the current is proportional to V^2 that is called as second region (Region-II) (Fig. 11). Here the value of slope is about 2, which is the characteristic of a trap free space charge limited current (SCLC) regime.^{51,56} If the injected carriers are more than the background carriers, the injected carriers spread and generate a space charge field. In this region the currents are controlled by this space charge field and are known as SCLC.^{51,56} The device performance was estimated by adopting SCLC theory.

Following this model, the effective carrier mobility has been estimated from higher voltage region of I vs. V^2 plot (Fig. 12) by Mott–Gurney equation.^{51,54,56}

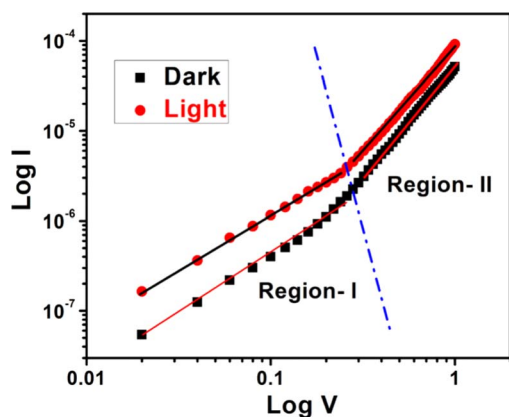


Fig. 11 $\log V$ vs. $\log I$ curves for the as-synthesized cobalt based thin film device under dark and illumination condition.

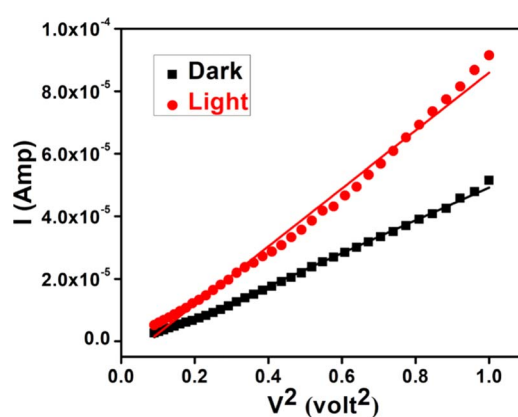


Fig. 12 I vs. V^2 curves for the synthesized cobalt based thin film device under dark and illumination condition.



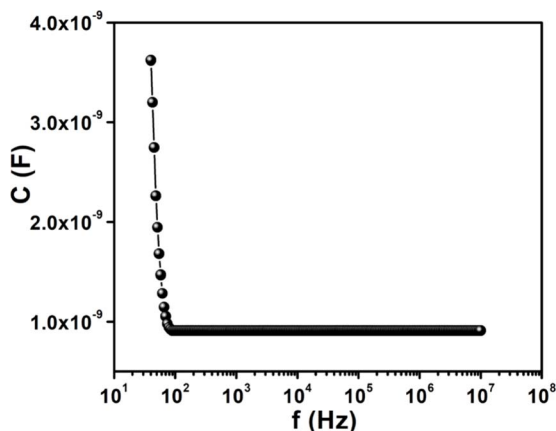


Fig. 13 Capacitance vs. frequency graph for determination of dielectric constant.

Table 6 Charge conducting parameters of cobalt based thin film devices

Condition	ϵ_r	$\mu_{\text{eff}} (\text{m}^2 \text{V}^{-1} \text{s}^{-1})$	$\tau (\text{sec})$	$\mu_{\text{eff}} \tau (\text{m}^2 \text{V}^{-1})$
Dark	2.47	3.02×10^{-7}	2.55×10^{-6}	7.701×10^{-13}
Light		5.33×10^{-7}	1.46×10^{-6}	7.782×10^{-13}

$$I = \frac{9\mu_{\text{eff}}\epsilon_0\epsilon_r A}{8} \left(\frac{V^2}{d^3} \right) \quad (2)$$

where, I , ϵ_0 , ϵ_r , and μ_{eff} is the current, permittivity of free space, relative dielectric constant of the synthesized material and effective mobility, respectively.

The relative dielectric constant has been estimated as 2.47, measured from the plot of capacitance against frequency (Fig. 13) of synthesized cobalt based film format at constant bias potential.

Further we have calculated the transit time (τ). These are some major parameters to analyze charge transport across the junction. From the slope of SCLC region (region-II) of Fig. 11, using the following equations we have estimated τ .^{51,57}

$$\tau = \frac{9\epsilon_0\epsilon_r A}{8d} \left(\frac{V}{I} \right) \quad (3)$$

All the parameters estimated in the SCLC region demonstrate that the charge transport properties of the material improve under illumination condition (Table 6). The higher mobility implied higher transport rate under illumination, while the number of charge carriers also increased under the same condition. The diode parameters of the synthesized complex based SD demonstrate superior charge transfer kinetics under irradiation condition.

The absence of region having slope higher than 2 in the IV characteristics can be attributed to the limited voltage range considered in our analysis, specifically in the positive voltage region. In the presented data, we focused on the voltage range up to 1 V, in which the slope of the 0–0.25 V range was found to be approximately 1 (ohmic region) and subsequently the slope as 2 (SCLC region).

Here, we would like to mention that extending the positive voltage range might reveal additional regions with higher slopes, more indicative of exponential behavior. Also, it is to be noted that, the I – V characteristic behavior of any diode not only depends on the applied voltage range, but also on the nature of the constituting materials.

It's crucial that the absence of higher slope regions within the explored voltage range doesn't imply a non-exponential nature. The diode exhibits exponential characteristics in the observed range, having the value of the slope as 1, and subsequently 2. Thus, the limited range of the positive voltage axis does not negate the overall exponential nature of the diode's behavior.

Table 7 Comparison of the electrical parameters of the present complex with other related cobalt complexes reported in literature^a

Cobalt complexes used in Schottky diode	Condition	Photosensitivity	Device conductivity (S m^{-1})	Ideality factor	Barrier height (eV)	Rectification ratio	Reference
$[\text{Co}^{\text{II}}(\text{N}_3)\{\text{Co}^{\text{III}}(\mu\text{-L})(\text{N}_3)_2(\mu\text{-N}_3)\} \cdot \text{MeOH}]$	Dark	0.93	1.05×10^{-6}	1.17	0.78	90	This work
	Light		2.12×10^{-6}	1.11	0.75	52.5	
$[\text{Co}^{\text{II}}\{\mu\text{-L}^1\}(\mu\text{-OAc})\text{Co}^{\text{III}}(\text{OAc})_2] \cdot 2.67\text{H}_2\text{O}$	Dark	—	5.858×10^{-4}	1.826	0.596	44.326	17
$[\text{Co}^{\text{II}}\{\mu\text{-L}^1\}(\mu\text{-OOCPh})\text{Co}^{\text{III}}(\text{DMSO})_{0.8}(\text{OOCPh})_{0.2}\}_2](\text{ClO}_4)_{1.6}$	—	—	6.804×10^{-5}	2.144	0.645	6.164	17
$[\text{Co}^{\text{II}}\{\mu\text{-L}^2\}(\mu\text{-OOCPh})\text{Co}^{\text{III}}(\text{DMSO})_{0.75}(\text{OOCPh})_{0.25}\}_2](\text{ClO}_4)_{1.5}$	—	—	9.358×10^{-5}	1.911	0.644	6.381	17
$[(\text{N}_3)\text{Co}^{\text{III}}\text{L}^3(\mu\text{-C}_6\text{H}_4(\text{NO}_2)\text{CO}_2)\text{Co}^{\text{II}}(\text{N}_3)]$	Dark	—	2.04×10^{-5}	1.13	0.69	14	56
$[(\text{N}_3)\text{Co}^{\text{III}}\text{L}^4(\mu\text{-C}_6\text{H}_4(\text{NO}_2)\text{CO}_2)\text{Co}^{\text{II}}(\text{N}_3)]$	—	—	2.58×10^{-4}	1.12	0.62	36	56
$[(\text{H}_2\text{O})\text{Co}^{\text{II}}\text{L}^3(\mu\text{-O}_2\text{CR}^1)\text{Co}^{\text{II}}(\text{NCS})]$	Dark	—	2.04×10^{-5}	1.82	0.72	—	57
$[(\text{DMSO})\text{Co}^{\text{II}}\text{L}^4(\mu\text{-O}_2\text{CR}^2)\text{Co}^{\text{II}}(\text{NCS})]$	—	—	1.54×10^{-3}	0.76	0.58	—	57
$[\text{Co}(\text{adc})(4\text{-ppy})(\text{H}_2\text{O})_2]_n$	Dark	—	2.61×10^{-5}	3.08	0.75	112.47	38
$[\text{Co}(\text{adc})(4\text{-bppy})(\text{H}_2\text{O})_2]_n$	—	—	2.52×10^{-6}	2.15	0.79	18.59	38

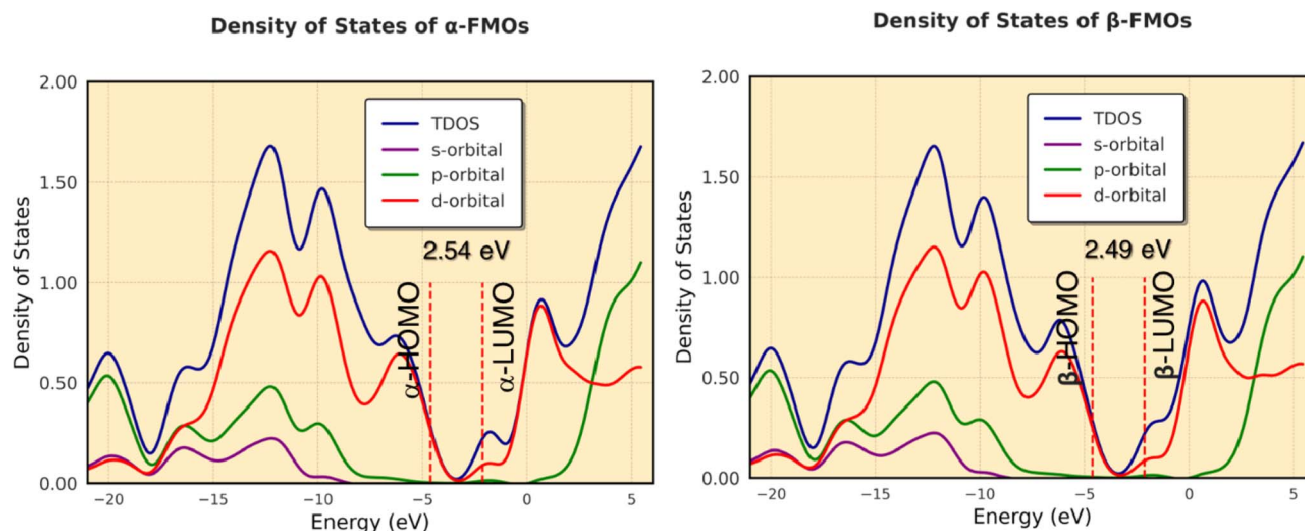
^a $\text{H}_2\text{L}^1 = [1,3\text{-bis}(2\text{-hydroxybenzylamino})2,2\text{-dimethylpropane}]$ and $\text{H}_2\text{L}^2 = [2,2' \text{-}[1,1' \text{-(propane-2,2-diylidimino)diethylidene}]diphenol]$, $\text{H}_2\text{L}^3 = [(2,2\text{-dimethyl-1,3-propanediyl})\text{bis}(\text{iminomethylene})\text{bis}(6\text{-methoxyphenol})]$, $\text{H}_2\text{L}^4 = [(2,2\text{-dimethyl-1,3-propanediyl})\text{bis}(\text{iminomethylene})\text{bis}(6\text{-ethoxyphenol})]$, $\text{HO}_2\text{CR}^1 = 3\text{-methyl-4-nitrobenzoic acid}$, $\text{HO}_2\text{CR}^2 = 4\text{-methyl-3-nitrobenzoic acid}$, $\text{H}_2\text{adc} = \text{acetylenedicarboxylic acid}$, 4-ppy = 4-phenylpyridine and 4-bppy = 4-(4-bromophenyl)-pyridine).



Table 8 Comparison of the electrical parameters of the complex with other related complexes reported in literature^a

Transition and non-transition metals complexes used in Schottky diode	Condition	Photosensitivity	Device conductivity (S m ⁻¹)	Ideality factor	Barrier height (eV)	Rectification ratio	Reference
{[Cd ₂ (HL ⁵) ₂ (N(CN) ₂) ₂]·H ₂ O} _n	Dark	2.89	1.78 × 10 ⁻⁷	2.75	0.42	15.48	60
	Light		6.15 × 10 ⁻⁷	2.35	0.38	46.23	
[Cd(HL ⁶) ₂ (N(CN) ₂) ₂] _n	Dark	2.54	1.07 × 10 ⁻⁷	2.96	0.44	14.73	60
	Light		2.44 × 10 ⁻⁷	2.62	0.41	37.69	
[(N ₃)CoL ⁷ Na(N ₃)] _n	Dark	1.45	0.88 × 10 ⁻⁸	0.27	0.90	25.57	14
	Light		2.38 × 10 ⁻⁸	0.91	0.87	135.49	
{[Zn(HL ⁸) ₂]·DMF} _n	Dark	—	2.53 × 10 ⁻⁹	2.04	0.39	—	61
	Light		5.97 × 10 ⁻⁵	1.88	0.21	—	
[Zn ₄ (bdc) ₄ (ppmh) ₂ (H ₂ O)] _n	Dark	20.95	6.20 × 10 ⁻⁷	3.41	0.60	93	62
	Light		1.37 × 10 ⁻⁶	2.91	0.58	32	
[{CuL ⁹ Na} ₂ (μ _{1,1,3} -NCS)HgCl(μ-Cl)(μ _{1,3} -NCS)] _n	Dark	57	1.48 × 10 ⁻⁶	0.35	0.71	21.66	16
	Light		8.40 × 10 ⁻⁵	0.58	0.62	30.75	
[Cu(L ¹⁰)(μ _{1,3} -N ₃)] _∞	Dark	—	12.54 × 10 ⁻⁵	2.25	1.16	3	63
	Light		20.94 × 10 ⁻⁵	1.59	0.60	129	
[Cu ₂ (L ¹¹) ₂ (μ _{1,1} -N ₃)(μ _{1,3} -N ₃)] _∞	Dark	—	13.45 × 10 ⁻⁵	3.21	0.63	4.31	63
	Light		28.14 × 10 ⁻⁵	1.27	0.62	104	

^a HL⁵ = [2-methoxy-6-((E)-(quinolin-3-ylimino)methyl)phenol], HL⁶ = [2-methoxy-6-((E)-(quinolin-5-ylimino)methyl)phenol], H₂L⁷ = [N,N'-bis(5-bromosalicylidene)-2,2-dimethyl-1,3-propanediamine], HL⁸ = (E)-N'-(thiophen-2-ylmethylene)isonicotinohydrazide, H₂bdc = 1,4-benzene dicarboxylic acid, ppmh = N-pyridin-2-yl-N'-pyridin-4-ylmethylene-hydrazine, H₂L⁹ = N,N-bis(3-methoxysalicylidene)propane-1,3-diamine, H₂L¹⁰ = 2-[(2-dimethylamino-ethylamino)-methyl]-4-nitrophenol, H₂L¹¹ = 2-[(2-diethylamino-ethylamino)-methyl]-4-nitrophenol.

Fig. 14 The total and partial DOS calculations of the valence and conduction bands for α - and β -FMOs of the complex.

The comparison of electrical parameters of our synthesized complex with some other cobalt complexes^{17,40,58,59} reported in literature is given in Table 7. It demonstrates that our synthesized complex exhibits photosensitivity whereas the other reported complexes do not. The ideality factor value, on the other hand, indicates that our synthesized complex produced a more ideal diode than other reported complexes. So, these kinds of materials can pave the way for a very promising future in optoelectronic device application.

It may be interesting to compare the electrical parameters of the present complex with some other complexes of different transition and non-transition metals with different Schiff bases and their reduced analogues. Table 8 gathers the electrical

properties of these complexes. This demonstrates that our synthesized mixed valence cobalt complex is showing higher conductivity, better rectification ratio and better ideality factor than many of other reported complexes.

Band gap and density of states of the investigated complex

The difference in energy between the highest occupied and lowest unoccupied MOs ($\Delta\epsilon = E_{\text{LUMO}} - E_{\text{HOMO}}$, eV) is sometimes assumed as the band gap. We calculated $\Delta\epsilon$ by DFT calculations as shown in Fig. 14. The calculated band gaps are 2.54 eV (from α -HOMO- α -LUMO) and 2.49 eV (from β -HOMO- β -LUMO) for the complex. The results show that the band gaps resulting from α -FMOs for the complex agree with the ranges



of experimental values as shown in Fig. 8, respectively. Fig. 14 shows the total and partial DOS calculations, which clearly state that the valence band is mainly dominated by the d-orbitals of the cobalt atoms. In this complex, both in the valence band and conduction band, the contribution of the s- and p-orbitals is poor.

Summary and conclusion

The synthesis and characterization of a trinuclear cobalt(III/II/III) complex with N₂O₂ donor reduced Schiff base ligand have been described in the present paper. Single crystal X-ray diffraction has confirmed its structure. More flexibility of the reduced Schiff bases compared to their Schiff base precursor may be the driving force in forming the complex, as the non-reduced analogue of the Schiff bases fails to prepare the complex. Electrochemical measurements were performed in DMF solution to account for the redox behavior of the complex. The band gap of this semiconductor material in the solid state has been determined and it confirms that it is a semiconductor. We have also shown the potential of this complex in fabricating a photosensitive Schottky device. Theoretical studies have been used to rationalize the experimental observations and to provide an explanation to the enhanced conductivity of the complex upon illumination. It is a promising candidate to have applications in semiconducting research and in the fabrication of optoelectronic devices.

Conflicts of interest

There are no conflicts to declare.

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