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Advances in chromone-based copper(II) Schiff base complexes: synthesis, characterization, and versatile applications in pharmacology and biomimetic catalysis†

Sumit Kumar,^{‡,ab} Aditi Arora,^{‡,a} Vipin K. Maikhuri,^{ID,‡,a} Ankita Chaudhary,^c Rajesh Kumar,^{ad} Virinder S. Parmar,^{abe} Brajendra K. Singh^{ID,*,a} and Divya Mathur^{ID,*,af}

Chromones are well known as fundamental structural elements found in numerous natural compounds and medicinal substances. The Schiff bases of chromones have a much wider range of pharmacological applications such as antitumor, antioxidant, anti-HIV, antifungal, anti-inflammatory, and antimicrobial properties. A lot of research has been carried out on chromone-based copper(II) Schiff-base complexes owing to their role in the organometallic domain and promise as potential bioactive cores. This review article is centered on copper(II) Schiff-base complexes derived from chromones, highlighting their diverse range of pharmacological applications documented in the past decade, as well as the future research opportunities they offer.

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1. Introduction

Chromone, also known as 4H-chromen-4-one or 4H-1-benzopyran-4-one, is an *O*-heterocyclic compound characterized by its benzo- γ -pyrone structure. It is part of a category of naturally occurring molecules that are prevalent in the plant kingdom, ranging from algae to coniferous plants.¹ It forms the central backbone in flavonoids such as flavone, isoflavone, and flavonol (Fig. 1).^{2,3} Chromones have become a compelling presence in the field of medicinal chemistry and drug development initiatives due to their potential to offer a wide range of chemical variations along with a diverse array of pharmacological effects.^{4–6} Both naturally existing and artificially synthesized compounds containing a chromone component have

exhibited a diverse range of biological characteristics including anti-inflammatory^{7,8} anticancer^{9,10} antitumor,¹¹ antimicrobial^{12,13} anti-HIV,¹⁴ antioxidant,¹⁵ and many other activities.^{16,17} Khellin, a furanochromone obtained from the seeds of the *Ammi visnaga* plant, serves as an illustrative instance of a natural chromone employed as a bronchodilator for asthma. Presently, it is also utilized in the management of vitiligo and psoriasis (Fig. 2).¹⁸ Exploration rooted in khellin has given rise to novel chromone derivatives like cromolyn and nedocromil, which are employed as mast cell stabilizers for the treatment of allergic asthma and allergic conjunctivitis.¹⁹ Diosmin, a flavone glycoside is known for its antioxidative, anti-inflammatory, antihyperglycemic, and anti-ulcer properties.²⁰ Another bioflavonoid, apigenin has been reported to modulate histamine release, is a bronchodilator, and also possesses anti-cancer properties.²¹ Flavoxate is known as an anticholinergic agent because of its antimuscarinic effects (Fig. 2).²²

Apart from the development of new chromone derivatives with promising properties, combining chromones with other pharmacophores to achieve hybrid structures with enhanced and multitarget therapeutic applications is an interesting proposition. One such pharmacophore is Schiff bases with varied applications across various domains, including pharmaceutical chemistry,²³ catalysis,²⁴ dye industry,²⁵ corrosion inhibitors^{26,27} and chemo-sensors.^{28,29} The conventional method for creating Schiff bases, which bears the name of Hugo Schiff, encompasses the combination of a carbonyl compound with an appropriate amine, employing suitable solvents and catalysts.³⁰ Schiff bases, a subclass of imines, feature a carbon–nitrogen double bond bound exclusively to alkyl or aryl groups, without

^aDepartment of Chemistry, Bioorganic Research Laboratory, University of Delhi, Delhi, India. E-mail: singhbk@chemistry.du.ac.in

^bDepartment of Chemistry and Environmental Science, Medgar Evers College, 1638 Bedford Avenue, Brooklyn, New York 11225, USA

^cDepartment of Chemistry, Maitreyi College, University of Delhi, Delhi, India

^dDepartment of Chemistry, R. D. S College, B. R. A. Bihar University, Muzaffarpur, India

^eAmity Institute of Click Chemistry and Research Studies, Amity University, Sector 125, Noida 201313, Uttar Pradesh, India

^fDepartment of Chemistry, Daulat Ram College, University of Delhi, Delhi, India. E-mail: dmchem05@gmail.com; divyamathur@dr.du.ac.in

† This article is a tribute to the cherished memory of our esteemed (late) Professor Ashok K. Prasad, who served as our enduring friend, collaborator, colleague, and mentor for many years. His enduring influence and guidance throughout the past several decades have profoundly impacted the careers of numerous aspiring scientists.

‡ All three authors have equal contribution.



any attached hydrogen atoms. Imines, organic compounds characterized by a C=N double bond, are pivotal entities in organic chemistry. They play essential roles in various fields such as synthesis, medicinal chemistry, coordination chemistry, and materials science. With their distinct reactivity and diverse structures, the imine functional group serves as a versatile and crucial component, facilitating the construction of intricate molecular architectures and functional materials.^{31,32} Owing to the electron-rich nature of the azomethine nitrogen atom through which they coordinate with metal ions, Schiff bases are known as one of the strongest chelators. Schiff bases possess remarkable pharmacological activities such as antibacterial, antifungal, and anticancer activities due to the presence of C=N moiety.^{33–35} A noteworthy number of Schiff



Sumit Kumar

motifs of therapeutic importance. He is currently a visiting Research Scholar at Medgar Evers College, City University of New York (CUNY) in the Chemistry and Environmental Science Department.



Vipin K. Maikhuri

of Chemistry, University of Delhi. His research interests include catalysis, heterocyclic chemistry, and carbohydrate chemistry.

base complexes are analogs of naturally occurring molecules and thus are of critical importance.³⁶ In this regard, Schiff bases can also be combined with chromones as they are ideal organic fluorescent probes possessing structural flexibility and pharmacological activity. 3-Formyl chromone serves as a commonly employed starting molecule in the construction of chromone Schiff bases, and it is typically generated through the Vilsmeier–Haack reaction.³⁷

Chromone Schiff bases are organic compounds resulting from the condensation reaction between chromone derivatives and primary amines. Chromone-based Schiff bases have been the focal point of organic, medicinal, and organometallic chemistry. Known for their pharmacological potential, these compounds exhibit diverse activities such as anti-inflammatory,



Aditi Arora

the design and synthesis of carbohydrate and coumarin-modified molecules.



Ankita Chaudhary

Dr Ankita Chaudhary, is currently working as an Assistant Professor at Department of Chemistry, Maitreyi College, University of Delhi, India. She received her PhD degree from the Department of Chemistry, University of Delhi in 2013. She has been a recipient of Junior/Senior Research Fellowship sponsored by CSIR, Science Meritorious Award, Jean and Ashit Ganguly Education Scholarship as well as UGC-Start Up Research Grant. She has over 11 years of experience in teaching postgraduate and undergraduate students. She has published 34 research articles in journals of international and national repute, 05 book chapters, 02 books and has presented her work in numerous international/national conferences. Her research interest encompasses design of novel green methodologies for the synthesis of biologically relevant heterocyclic compounds.



antioxidant, anticancer, antibacterial, and antifungal properties.³⁸ Owing to their established reputation as chelating ligands, these compounds are recognized for their enhanced bioavailability, decreased toxicity, ability to form stable metal complexes, and robust binding affinity for metal ions.^{39–41} The synthesis of chromone Schiff bases typically involves mild conditions and allows for structural modifications to enhance their biological activities. Research has shown that these compounds possess promising therapeutic effects, making them valuable candidates for drug discovery and



Rajesh Kumar

chemistry, biotransformations, and heterocyclic chemistry.

Dr Rajesh Kumar received his MSc degree in Organic chemistry in 2010 and PhD in 2017 from the University of Delhi. He also worked as a Research Assistant in the University of Southern Denmark. After completing his PhD he joined the chemistry department at B. R. A. Bihar University, India. He has published 25 research papers in reputed national and international journals. His research interests lie in nucleic acid



Brajendra K. Singh

for the industrial production of N- and O-arylated compounds using microreactor technology via copper(II) catalyzed cross-coupling reactions. His key interest areas are the study and investigation of bio-catalysis and metal-catalysis, specifically in the design, synthesis and evaluation of bioactive compounds.

Dr Brajendra K. Singh currently holds the position of Assistant Professor in the Department of Chemistry, University of Delhi. He visited the Department of Chemistry, University of Leuven, Leuven, Belgium as International Scholar for the period November 2005–March 2007 where he developed new strategies to carry out microwave reactions under simultaneous cooling. He has developed an efficient scalability procedure

development.⁴² A tremendous increase in the biological property of Schiff bases is observed when they are chelated with transition metal ions. Such metal complexes can be efficiently prepared through the addition of a suitable Schiff-base ligand to a metal precursor in an appropriate ratio and experimental conditions (Fig. 3). This is also true for chromone-based Schiff bases as exemplified by many examples. Chromone Schiff-base nano-complexes (**I**) such as those of Cu²⁺, Ni²⁺, Co²⁺, Fe³⁺, Zn²⁺, Cd²⁺, and UO₂⁶⁺, and Zn²⁺ have been successfully prepared and evaluated as antimicrobial and antitumor agents, exhibiting



Virinder S. Parmar

Head of the Department of Chemistry and as Chairman of the Board of Research Studies, and Provost of Gwyer Hall at this university. Professor Parmar's research interests include: green/sustainable Chemistry, nanotechnology, organic synthesis, nucleic acid chemistry, advanced materials, medicinal chemistry, biocatalysis and the chemistry of natural products. He has mentored 85 PhD and postdoctoral scientists.

Prof. V. S. Parmar is a Professor of Chemistry at CUNY-Graduate Center, and Departments of Chemistry, City College & Medgar Evers College, The City University of New York, and Institute of Click Chemistry Research and Studies, Amity University, Noida, India. He has also been a faculty member at St Stephen's College and the University of Delhi for 44 years, he recently retired as Professor of Chemistry and has served as



Divya Mathur

Dr Divya Mathur obtained her PhD in Chemistry from the University of Delhi in 2013. She did her postdoctoral research at Ghent University, Belgium, and worked as a visiting researcher at Acadia University, Canada. She is a recipient of the Erasmus Mundus Postdoctoral Scholarship by the European Union, the Canadian Commonwealth Scholarship by the Government of Canada, and National Postdoctoral Fellowship, Junior and Senior Research fellowships by the Government of India. She has over 8 years of experience in teaching postgraduate and undergraduate students. She has published over 30 research papers in reputed national and international journals. Presently, she is working as an Assistant Professor in Chemistry at Daulat Ram College, University of Delhi. Her research areas include nucleoside chemistry, carbohydrate chemistry, biocatalysis, and heterocyclic chemistry.



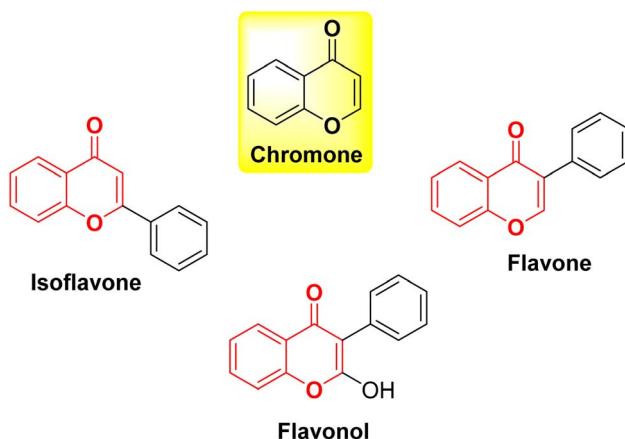


Fig. 1 Chromone skeleton in various flavonoids.

lower toxicity compared to cis-platin.⁴³ Chromone Schiff base (**II**) complexes with Pd^{2+} have also been successfully prepared, characterized, and investigated for circulating tumor DNA (ctDNA) binding, radical scavenging, and antimicrobial potency.⁴⁴ In a recent study, a fluorescent molecule appended with rhodamine and chromone (**III**) was synthesized as an efficient sensor for Fe^{3+} , Al^{3+} , and Cr^{3+} at trace levels. Further, they were assessed for anticancer activity and were found to be effective against MCF-7 cells, glucocorticoid, and progesterone receptors.⁴⁵ Barve *et al.*⁴⁶ synthesized copper ion conjugated chromone Schiff base complexes (**IV**) these compounds demonstrated their effectiveness as inhibitors against BT20, PC-3, as well as both COLO 357 and BxPC-3 cancer cell lines. Thus, it can be affirmatively suggested that the metal complexes exhibit better biological activity. The increase in activity can be ascribed to structural modifications resulting from coordination with the metal, and thus chelation makes metal complexes more potent compared to Schiff bases.⁴⁷ The report by El-Saghier *et al.*⁴⁸ established that Cu^{2+} complexes exhibit

superior inhibition in comparison to the other complexes. This was ascribed to the presence of free mobilized electrons of Cu^{2+} complexes that permit strong oxidative inhibitory activity against microorganisms. This re-emphasizes that chromone-based Schiff bases are capable of forming d-block metal complexes amongst which copper complexes are predominantly effective as they have exhibited potent radical scavenging, anti-cancer, antimicrobial, and chemosensing properties.

The exploration of copper coordination chemistry is influenced by the crucial role copper assumes in numerous physiological processes, where it serves as an intrinsic component within metalloproteins like tyrosinase, cytochrome C oxidase, Cu/Zn superoxide dismutase, and blue copper proteins.^{49,50} Equally significant is the fact that there are numerous copper(II) complexes with four, five or six-coordinate and different geometries such as square planar, flattened tetrahedral, trigonal bipyramidal, square-pyramidal, distorted octahedral, and penta-coordinated structures.^{51–53} Owing to the pharmacological importance of chromones and Schiff bases, together with the multifaceted role of copper, there is a substantial interest in the development of chromone Schiff base copper complexes.

The primary goal of the current study is to perform an extensive review of published data from the past decade (2013–2023) pertaining to the synthesis and diverse applications of copper complexes derived from chromone-based Schiff bases. The impetus for undertaking this research stemmed from the observation that no comprehensive review had been available until now on Cu^{2+} complexes involving chromone Schiff bases. In this work, we initially delve into the use of chromone-based Schiff base ligands as a selective colorimetric tool for detecting Cu^{2+} ions. Subsequently, we provide a comprehensive examination of ongoing research on chromone Schiff base copper complexes, highlighting their wide-ranging medicinal and pharmaceutical applications, such as antimicrobial, antioxidant, and antitumor properties. Further, we have explained

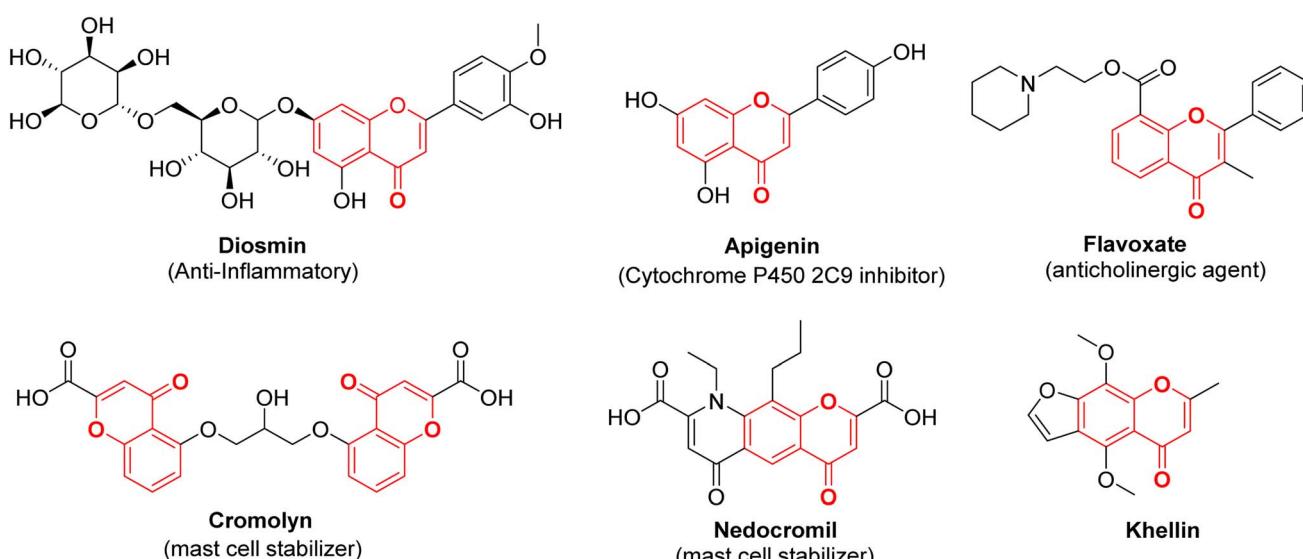


Fig. 2 Representative chromone derivatives as pharmaceutical agents.



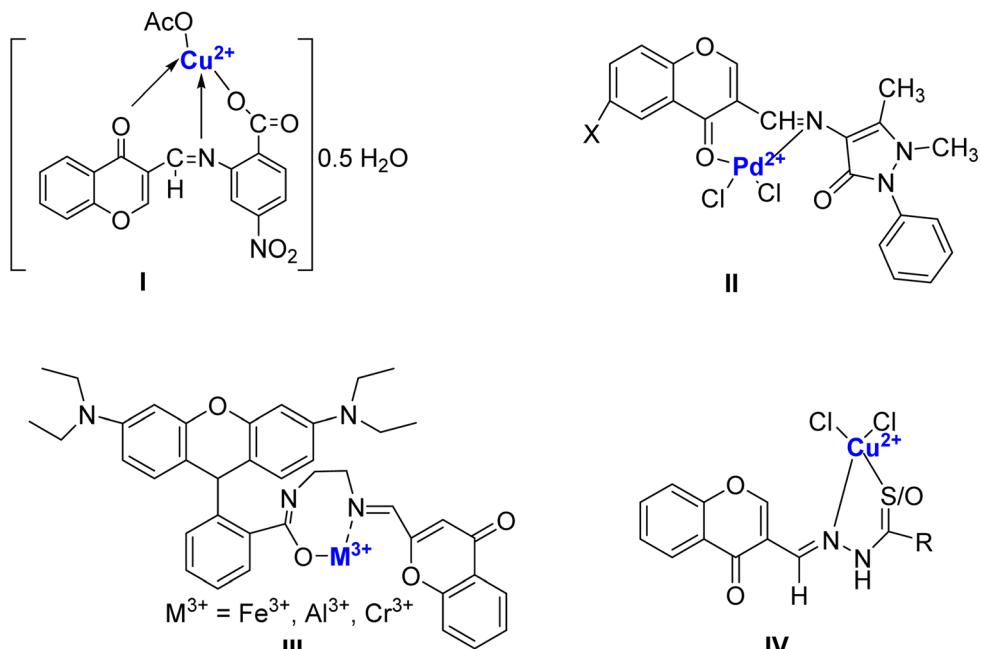


Fig. 3 Pharmacologically significant chromone-based Schiff base complexes reported in the literature.

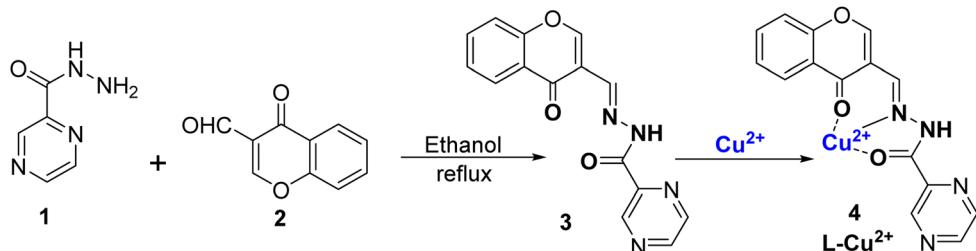
the biomimetic catalytic activity of chromone-based Schiff bases together with some miscellaneous applications. Finally, the challenges and future developments have been deliberated upon to promote the research on copper chromone Schiff base agents.

2. Copper detection by chromone-based Schiff bases

Copper is regarded as an abundant and important metal which play a very useful role in various biological, medicinal, and environmental fields.⁵⁴ On one hand, excessive use of copper metal can cause serious problems in humans such as retardation of growth, Wilson's disease, and liver destruction while on the other, less intake of copper is also harmful, causing anemia, leukopenia, and myelopathy.⁵⁵ Excess Cu²⁺ ions in drinking water sources can cause damage to animals, humans, and plants. According to the U. S. EPA, the highest range of Cu²⁺ ions in water should not be more than 20 μM .^{56,57} According to WHO, the maximum limit of Cu²⁺ ions is 31.4 μM .⁵⁸ Thus, preparing an efficient sensor for tracing Cu²⁺ ions in environmental bodies and organic species becomes all the more

important.⁵⁹⁻⁶¹ Many traditional methods are available, which are otherwise costly or difficult to handle such as electrochemical sensing and atomic absorption spectroscopy (AAS).⁶²⁻⁶⁴ Therefore, there is a need for a simple, cost-effective, accurate method for monitoring Cu²⁺ ions in samples. Colorimetric sensors are the solution for the detection of copper because they not only reduce the cost but also make the process of detection easy and fast. Due to this, the design and use of colorimetric probes is a flourishing area of research.^{65,66} Nowadays, different variety of organic receptors and nanomaterials are used as colorimetric sensors.⁶⁷⁻⁶⁹ In recent times, Schiff bases and chromone-derived colorimetric and optical sensors have also been used for metal ion detection.⁷⁰⁻⁷² This is due to their specific structure, and large binding affinity with metal ions.^{73,74} Herein, we have discussed chromone Schiff base sensors which bind to Cu²⁺ ions making exclusive coloured copper complexes and thus can be used as colorimetric sensors.

In 2021, Tomer *et al.*⁷⁵ successfully synthesized a Schiff base ligand denoted as ligand 3, which was derived from chromone. The synthesis of this ligand involved a condensation reaction between pyrazine-2-carbohydrazide 1 and 3-formylchromone 2, as illustrated in Scheme 1. The synthesized ligand 3 underwent



Scheme 1 Schiff base ligand 3 was synthesized, followed by the preparation of its Cu²⁺ complex 4.

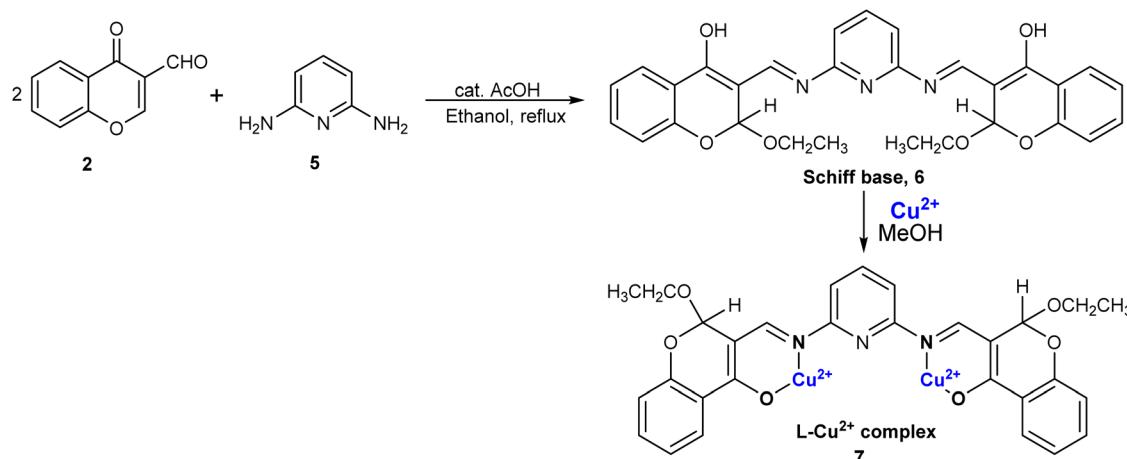


characterization such as $^1\text{H-NMR}$, HRMS, FT-IR, and UV-Vis spectroscopy. Notably, ligand **3** was employed as a highly effective colorimetric probe for the detection of Cu^{2+} ions in different water sources, including canal water, groundwater, and tap water. The addition of Cu^{2+} ions to ligand **3** resulted in an immediate change in color from colorless to yellow, and the color change was because of the complex formation between the ligand and Cu^{2+} ions, referred to as complex **4**. Importantly, there was no color change with other tested metal ions. The absorption spectra of ligand **3** displayed two prominent absorption peaks at 258 nm and 311 nm. Nonetheless, with the introduction of Cu^{2+} ions, a fresh absorption peak became apparent at 428 nm, while the 311 nm band diminished. This shift in absorption, known as a bathochromic shift, was attributed to the complexation of Cu^{2+} ions and ligands through an intramolecular charge transfer process. When Cu^{2+} ions were incrementally introduced into the solution, another band at 428 nm continued to appear, while the 311 nm band gradually decreased. The researchers observed that ligand **3** displayed remarkable sensitivity across a wide pH range, particularly in neutral and basic conditions within the pH range of 6–11, although it was not effective at very low pH values, such as 2–3. The limit of detection for Cu^{2+} ions was determined to be $3.9 \times 10^{-7} \text{ M}$, and the association constant was calculated to be $2.3 \times 10^5 \text{ M}^{-1}$ using the Benesi–Hildebrand equation. The stoichiometry between ligand **3** and Cu^{2+} ions as 1 : 1 was confirmed through Job's plot, which was also substantiated by HRMS and DFT. Moreover, the practical applicability of this developed ligand was assessed by testing its performance in actual water samples for the identification and quantification of Cu^{2+} ions.

A chromone-functionalized pyridine-based chemosensor was also prepared by the same group for the sensing of Cu^{2+} ions.⁷⁶ In their study, the researchers reported the synthesis of a novel Schiff base labeled as Schiff base **6**. This compound was prepared by condensing **2** with 2,6-diaminopyridine **5** in ethanol as the solvent, and glacial acetic acid as catalyst (Scheme 2). A significant finding was that the binding stoichiometry between Schiff base **6** and Cu^{2+} ions was established as 1 : 2, a conclusion

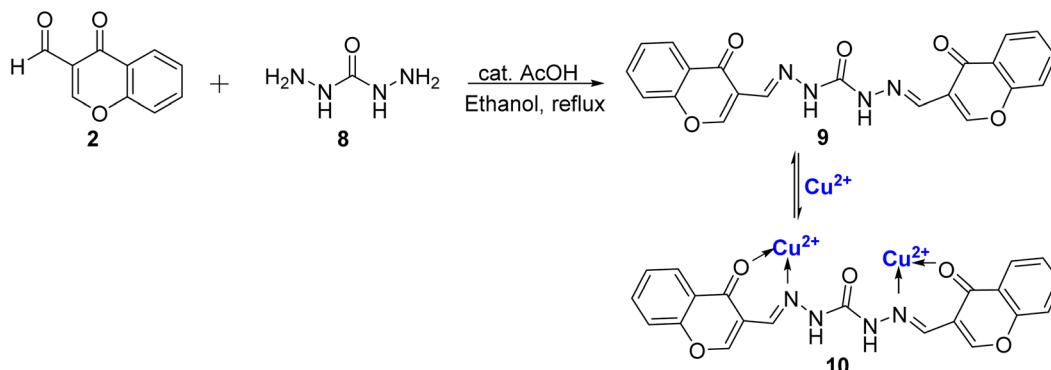
supported by both Job's plot and HRMS spectra data. Schiff base **6** demonstrated exceptional sensitivity and selectivity for cupric ions in comparison to other metal ions, as revealed by absorbance and emission titration studies. The absorption spectrum of Schiff base **6** exhibited three distinct absorption bands at 306 nm, 357 nm, and 401 nm. Upon the introduction of cupric ions, a new absorption band emerged at 347 nm, attributable to ligand-to-metal charge transfer, while a reduction in intensity was observed at 306 nm and 401 nm. Regarding fluorescence emission spectra, Schiff base **6** displayed a peak at 465 nm. The addition of Cu^{2+} ions resulted in a decrease in fluorescence intensity because of the paramagnetic nature of Cu^{2+} ions, accompanied by the appearance of a new peak at 458 nm. This fluorescence quenching was specific to the binding of Schiff base **6** to Cu^{2+} ions, with no similar response observed with other ions. pH studies indicated that the optimal pH range for detecting cupric ions was between 6 and 11. At lower pH values, protonation of the azomethine linkage occurred, leading to reduced sensing capability. Schiff base **6** exhibited an association constant of $3.26 \times 10^4 \text{ M}^{-1}$ and a regression coefficient of 98.7% for Cu^{2+} ions. It displayed a detection value of 1.2 μM for the Schiff base– Cu^{2+} complex **7**, which was below the permissible limit for Cu^{2+} ions, thus confirming its effectiveness as a probe. Schiff base **6** was successfully employed in actual water samples, including canal water, distilled water, groundwater, and tap water for the detection of Cu^{2+} ions.

Tomer *et al.*⁷⁷ presented another Schiff base probe based on chromone **9**, in their study. This probe exhibited remarkable selectivity for Cu^{2+} ions and possessed the added capability of detecting *para*-nitrotoluene (*p*NT) with impressively low detection limits. The synthesis of this probe **9**, was accomplished through a straightforward, one-step condensation reaction involving **2** and **8**, (Scheme 3). Notably, Schiff base probe **9** displayed noticeable naked-eye colorimetric changes, shifting from colorless to yellow upon the introduction of Cu^{2+} ions, setting it apart from the response to other ions tested. In terms of its spectroscopic features, the Schiff base probe **9** exhibited two distinct absorption bands at 315 nm and 278 nm. With the



Scheme 2 Synthesis of Schiff base **6** and its complex with Cu^{2+} ions **7**.



Scheme 3 Synthesis of probe 9 and its complex with Cu^{2+} ions 10.

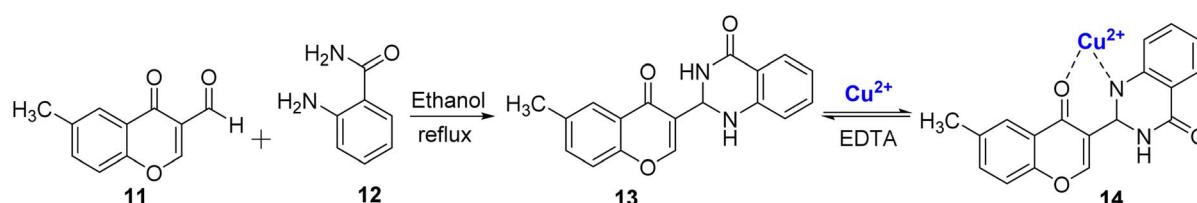
gradual addition of Cu^{2+} ions, a new absorption band emerged at 421 nm, accompanied by a reduction in the intensity of the bands at 315 nm and 278 nm. The binding stoichiometry of 1 : 2 between the complex and Cu^{2+} ions was confirmed through computational studies, Job's plot, and HRMS. The sensor exhibited reliable performance within a pH range of 5–9, and the sensing mechanism was found to be reversible, enhancing its practical utility. The binding constant was determined to be $4.5 \times 10^5 \text{ M}^{-1}$ for complex 10, with a low detection limit of $11.4 \times 10^{-7} \text{ M}$. For practical applications, it was demonstrated that the most noticeable color change can occur at a Cu^{2+} concentration of 10^{-4} M , using a paper strip test which underscores the efficacy of the probe 9 in real-world scenarios.

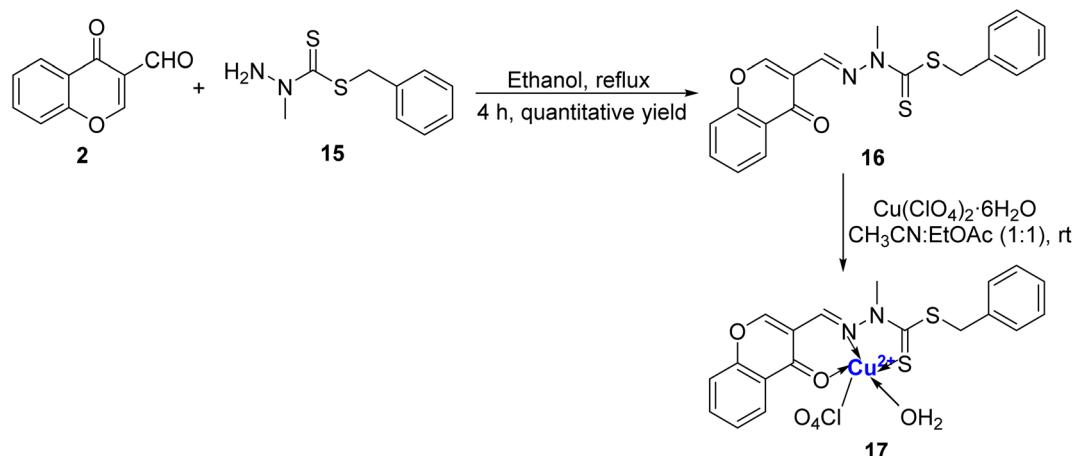
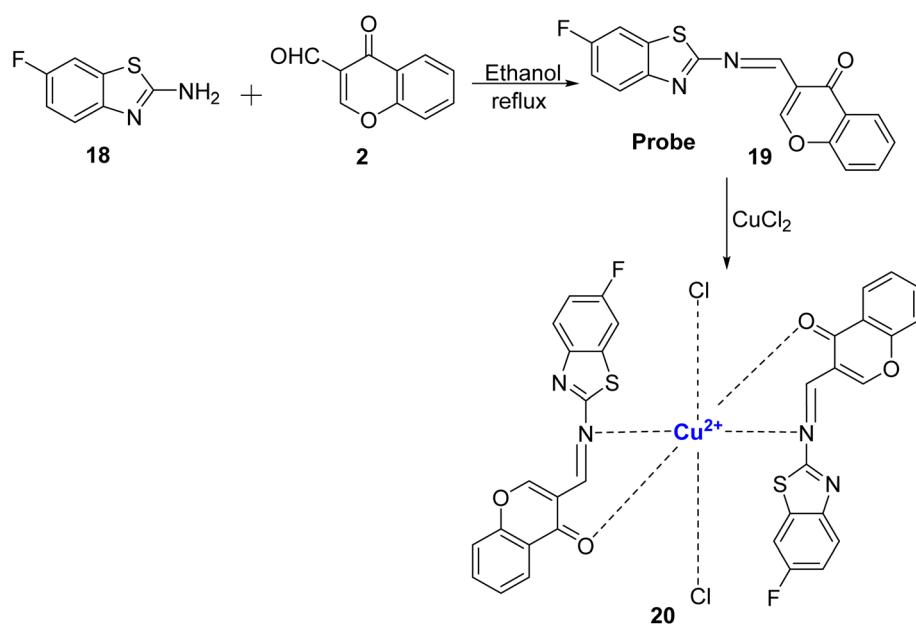
In another study, Mohammadi *et al.*⁷⁸ reported a novel optical sensor 13 for selective sensing of Cu^{2+} ions in acetonitrile–water solution. The sensor was synthesized by condensation reaction between 11 and 2-amino benzamide 12 followed by cyclization resulting in quinazolinone ring formation (Scheme 4). The chromone-based sensor 13 exhibited exceptional selectivity by undergoing a visible color change, transitioning from colorless to yellow when just $1 \mu\text{M}$ of Cu^{2+} ions were added. Notably, it retained its colorless appearance when tested individually with all twelve other metal ions at the same concentration. Upon the subsequent introduction of Cu^{2+} ions to the synthesized complex, a reduction in the intensity of the strong absorption band of the free ligand at 379 nm was observed, accompanied by the appearance of new absorption bands at approximately 306 nm and 456 nm. The stoichiometric ratio of 1 : 1 between the colorimetric sensor 13 and Cu^{2+} ions was corroborated through Job's plot and DFT studies. The binding constant for the 13– Cu^{2+} complex was determined to be $3.27 \times 10^4 \text{ M}^{-1}$, emphasizing the strength of the interaction of the sensor and Cu^{2+} ions. In anticipation of potential commercial applications, detection test strips were developed by coating filter

paper with this chemosensor. These test strips allowed for the immediate naked-eye detection of copper ions in a wide concentration range, spanning from 10^{-3} M to 10^{-7} M , which enhances the practical utility and versatility of this sensor. The near nano detection limit ($4.6 \times 10^{-7} \text{ M}$), wide concentration range, short response time, and reversible properties of 14 were some of the salient features of the developed sensor.

Rahman *et al.*⁷⁹ prepared a sensitive and selective chromone and benzylidethiocarbazate based ligand 16 for colorimetric detection of Cu^{2+} ion in HEPES buffer media in a mixture of $\text{H}_2\text{O}/\text{CH}_3\text{CN}$ (1 : 1). The ligand 16 was synthesized as a colourless cotton-like solid by taking *N*-methyl-*S*-benzylidethiocarbazate 15 with 3-formylchromone 2 in ethanol under reflux (Scheme 5). 100% selectivity was achieved in the case of Cu^{2+} with colourless to yellow colour change and there was no change in the presence of other metal ions and anions. Single crystal analysis of the green complex 17 formed with Cu^{2+} and 16 was found to be penta-coordinated with square pyramidal orientation. Complex 17 showed an absorption peak around 420 nm and also displayed a 1 : 1 stoichiometric ratio for Cu^{2+} : 16. The detection limit was 0.12 nM and the association constant for 16 with Cu^{2+} was $5.24 \times 10^6 \text{ M}^{-1}$. A colorimetric detection kit for Cu^{2+} ions was developed using the sensor and it was found that it could easily detect even 2 μM concentration of Cu^{2+} that is far beneath the WHO acceptable level.

Kouser *et al.*⁸⁰ reported the development of a chemosensor probe denoted as 19, based on the chromone molecule. This probe, 19, was synthesized by reacting 2 with 2-amino-6-fluorobenzothiazole 18 under reflux conditions, (Scheme 6). Notably, probe 19 exhibited exceptional selectivity and sensitivity towards Cu^{2+} ions as opposed to a range of other metal ions. Two distinct absorption bands at 371 nm and 260 nm were assigned to $\pi-\pi^*$ and $\pi-\pi^*$ transitions. Upon the introduction of Cu^{2+} ions

Scheme 4 Synthesis of the colorimetric sensor 13 and its complex with Cu^{2+} ions 14.

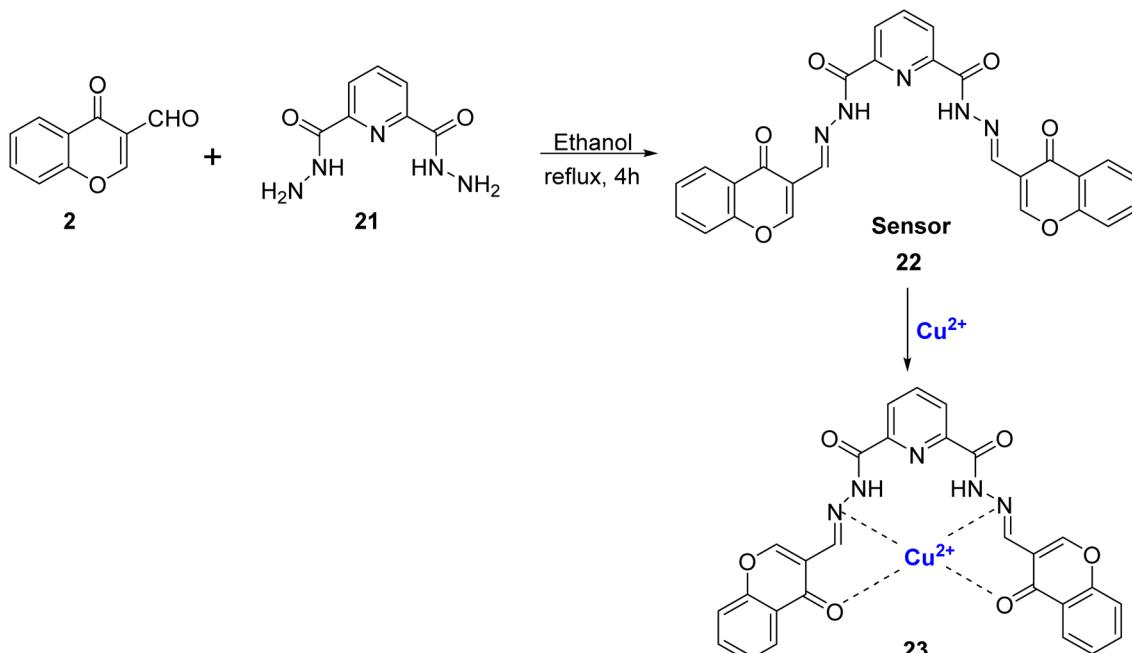
Scheme 5 Synthesis of ligand 16 and its complex 17 with Cu^{2+} ion.Scheme 6 Synthesis of chemosensor probe 19 and its complex with Cu^{2+} ion 20.

into the solution of **19**, there was a noticeable change from daffodil yellow to green, and new bands emerged at 355 and 452 nm. These changes were attributed to the formation of charge transfer complexes, d-d transitions, and $\pi-\pi^*$ transitions. The sensor demonstrated rapid recognition of Cu^{2+} ions and a low detection limit of 0.273×10^{-6} mol L⁻¹. Furthermore, it displayed an increase in fluorescence intensity and a high association constant of 8.48×10^8 M⁻² with the addition of Cu^{2+} ions. By utilizing Job's plot analysis, it was determined that the stoichiometric ratio between probe **19** and Cu^{2+} in the complex was 2 : 1. The sensor's effectiveness in detecting copper-containing biomolecules was also assessed in various biological samples, including healthy liver tissues, *F. gigantica*-infected liver tissues, and adult *F. gigantica* worms.

Rezaeian *et al.* reported a chemo-sensor **22** obtained by the condensation reaction of **2** and 2,6-pyridinedicarbohydrazide **21**

in high yields (Scheme 7).⁸¹ The synthesized ligand **22** was used to sense and spot Cu^{2+} and Zn^{2+} ions both visually and spectrophotometrically among the different cations tested. With the increase in Cu^{2+} ions absorbance at 310 nm decreased with the concomitant formation of complex **23** and a new band at 425 nm. Cu^{2+} ion and ligand **22** had 1 : 1 stoichiometry and a binding constant of 3.41×10^4 M⁻¹. From the experimental data, the detection limit for Cu^{2+} obtained was 5.5×10^{-7} mol L⁻¹ which was lower as compared to the limit set for safe drinking water by WHO. Further, multiple and complex Boolean operations such as OR, NOR, AND, INH, and half adders such as AND and XOR were also conducted by varying anions and cations as chemical inputs and the absorption intensity as the logic output.

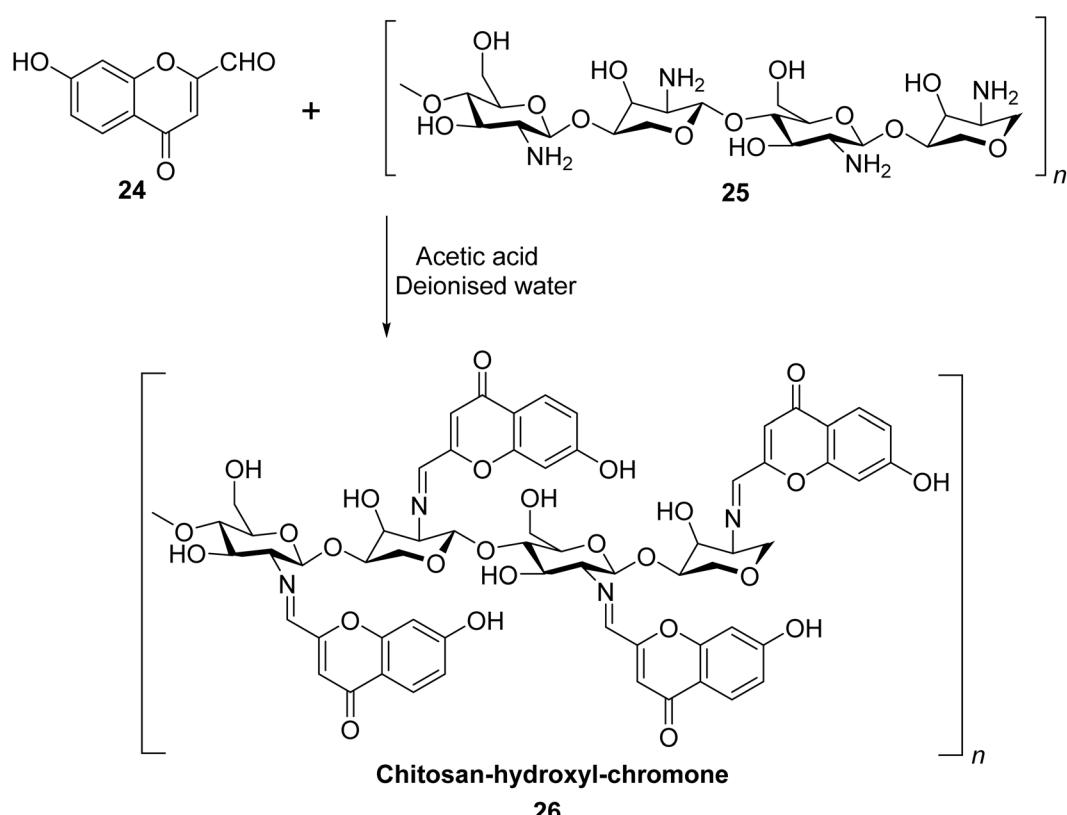
Gaidhane *et al.* used chitosan, which is a naturally occurring, environment-friendly bio-polysaccharide derived from chitin to synthesize a new chromone-based polymer **26**.⁸² Polymer **26** was



Scheme 7 Synthesis of chemo-sensor 22 and its complex with Cu^{2+} ions 23.

prepared by the treatment of 24 with chitosan 25 in deionized water (Scheme 8). It displayed specific, selective, and pH-influenced chelating efficiency with Cu^{2+} ions in comparison to Ni^{2+} , Co^{2+} , and Cd^{2+} heavy metal ions. At pH 4–6, polymer 26

showed selectivity in the order of $\text{Cu}^{2+} > \text{Ni}^{2+} > \text{Cd}^{2+} > \text{Co}^{2+}$. The adsorption capacity was also highest for Cu^{2+} with a concentration of 3 mmol g^{-1} . Chromone-based chitosan polymer 26 significantly inhibited lipid peroxidation emulsion system. It



Scheme 8 Synthesis of chitosan-hydroxyl-chromone 26.

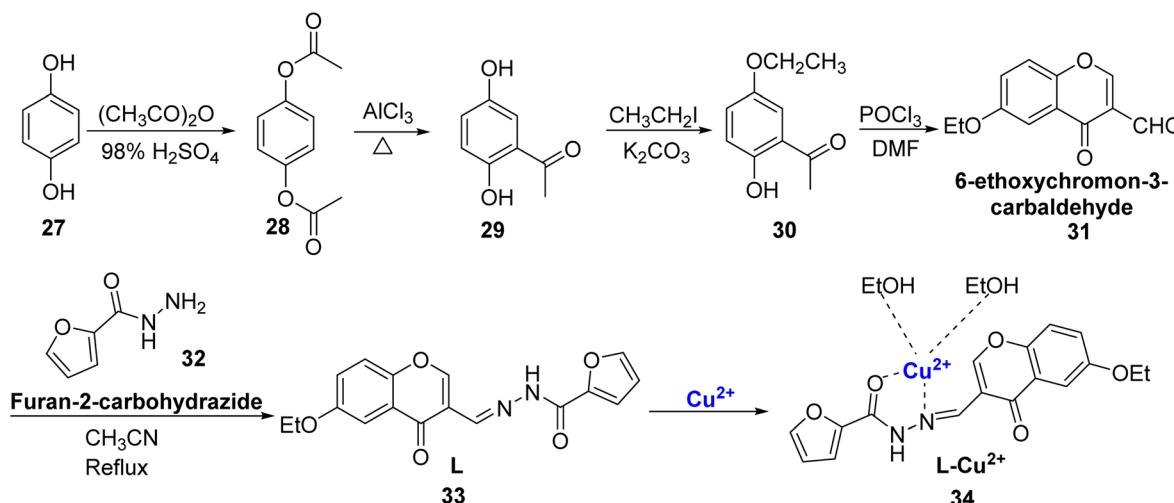


also demonstrated excellent results against Gram-negative bacteria *Pseudomonas aeruginosa* and decent activities with *Escherichia coli*, *Staphylococcus aureus*, and *Candida albicans*.

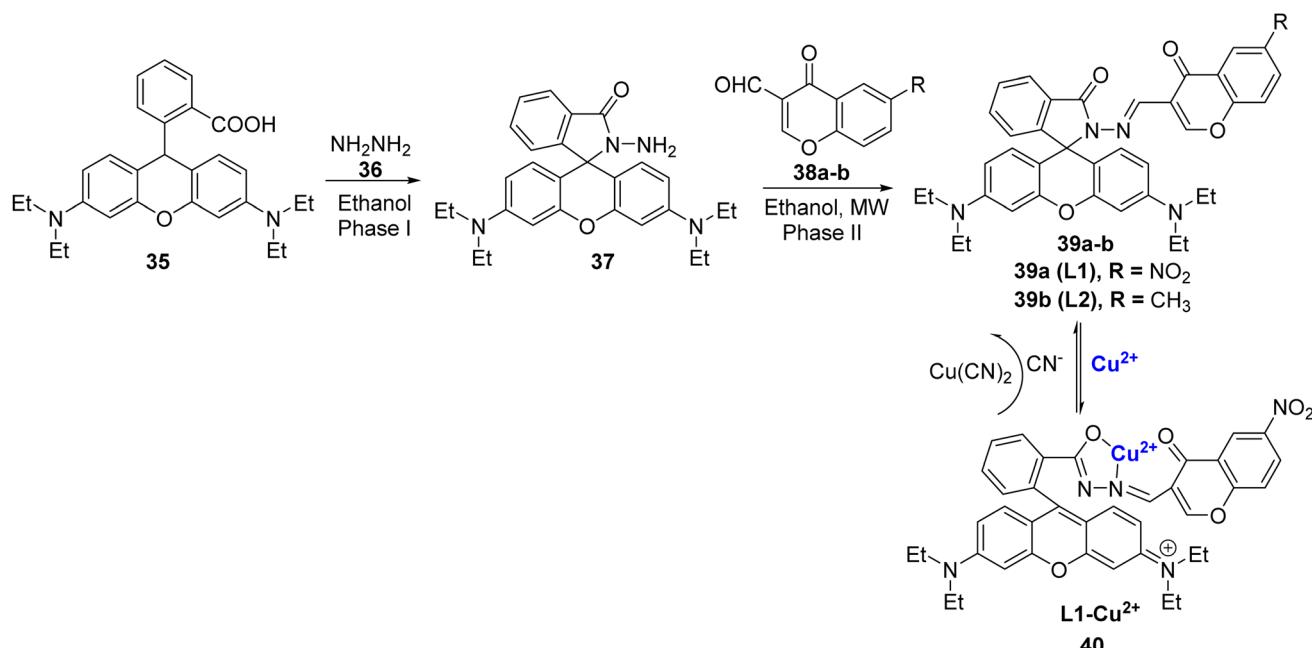
Yang *et al.* reported a new probe *i.e.*, 6-ethoxychromone-3-carbaldehyde-(furanyl)hydrazone 33 with turn-on fluorescence and colorimetric properties.⁸³ The dual functional probe 33 was designed by refluxing a mixture of furan-2-carbohydrazide 32 and 6-ethoxychromon-3-carbaldehyde 31 in acetonitrile for 24 hours (Scheme 9). It showed colorimetric changes and high selectivity for Al^{3+} and Cu^{2+} as compared to other metal ions. When exposed to Cu^{2+} ions, it underwent a visible color change from colorless to yellow. When Cu^{2+} was introduced, the absorption peak at 319 nm steadily diminished, and a novel absorption band appeared at 422 nm. Analyzing the data

through Job's plot and HRMS confirmed that the 33– Cu^{2+} complex involved a 1:1 coordination ratio. The probe 33 demonstrated an impressive detection limit of 2.857×10^{-7} M for Cu^{2+} ions, which met the standards set by the WHO. Furthermore, the practical applicability of 33 as a solid-state probe was assessed by utilizing silica-coated slides and test papers immersed in the chromone-based probe 33. Based on the findings, it is evident that 33 could serve as a prototype for a wide range of practical applications in environmental and biological systems.

Abebe *et al.* synthesized two rhodamine B-based chemosensors, 39a and 39b using microwave irradiation protocol (Scheme 10).⁸⁴ They were prepared in a two-step process, firstly by reacting rhodamine B 35 and hydrazine hydrate 36 in ethanol



Scheme 9 Synthesis of probe 33 and Cu^{2+} -probe complex 34.



Scheme 10 Synthesis of ligands (39a) and (39b) and their complexation with Cu^{2+} ions 40.



as green solvent to afford compound **37**, which further underwent condensation reaction with chromone aldehydes **38a–b** to furnish the final products **39a–b** (Scheme 10). The ligands **39a–b** were found to be selective for Cu^{2+} ions in the aqueous media. The ability of both the ligands **39a–b** was measured by using fluorescence titration, Job's plot, ^1H NMR, and DFT studies. Out of the two synthesized sensors, **39a** with electron-withdrawing substituent *i.e.*, the nitro group exhibited better response, good selectivity, and high sensitivity to Cu^{2+} ions in the aqueous solution. The fluorescence emission wavelengths for **39a** with Cu^{2+} ions appeared at 580 nm. The fluorescence intensity was observed mainly due to the formation of ring-open spirolactam **40** with Cu^{2+} ions whereas the other metal ions produced no substantial effect. Job's plot confirmed that both sensors bind to copper ions in 1:1 stoichiometry. The association constant of **39a** with Cu^{2+} was estimated to be $3.12 \times 10^4 \text{ M}^{-1}$ and the limit of detection was found to be $2.11 \mu\text{M}$. The *in situ* synthesis of the L1– Cu^{2+} complex **40** displayed excellent selectivity for cyanide ions through the metal-displacement method. It is noteworthy that the addition of cyanide ions to complex **40** led to a change in colour from pink to colourless. However, anions, such as Cl^- , I^- , F^- , ClO_4^- , CH_3COO^- , HSO_4^{2-} , $\text{H}_2\text{SO}_4^{2-}$, PO_4^{3-} , SCN^{2-} , HS^- , and OH^- , did not give any results.

Recently, Alorabi *et al.* designed a multidentate Schiff base ligand, **42** through condensation of **2** and 2-aminophenol **41** (Scheme 11).⁸⁵ They examined its role as a colorimetric chemosensor. There was no significant colour change for metal ions other than Cu^{2+} , Fe^{3+} , and V^{5+} . It was proposed that the synthesized Schiff base ligand coordinates with metal ions through nitrogen atom from the azomethine group and with the oxygen atom from the phenolic group, and the ketonic group. Utilizing Job's plot, it was determined that the binding stoichiometry between the ligand and Cu^{2+} ions was 2:1. The detection limit for copper ions was $7.03 \mu\text{M}$, and the binding constant was calculated to be $1.37 \times 10^4 \text{ M}^{-1}$. To evaluate the real-world practicality of the chemosensor, various water samples were gathered, encompassing distilled water, household water sources, and a sample from the Al-Aqiq water reservoir dam.

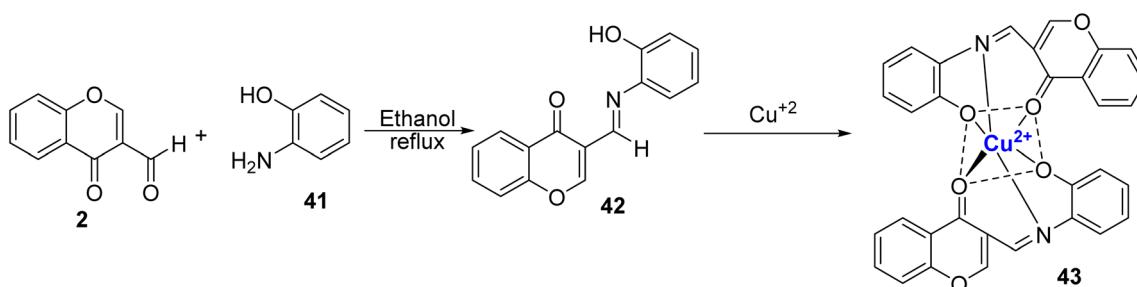
3. Pharmacological activities of chromone-based Schiff bases

Transition metal ion such as copper that is biocompatible combined with a ligand framework of Schiff base induces

a unique multi-modal mechanism of biological action. Incorporation of copper ions into the chromone moiety enhances their chemical and biological diversity. Thus, Schiff base-based metallo-drugs has great potential to combat disease like cancer, and microbial infections and circumvent multi-drug resistance problem. Copper complexes of chromone-based Schiff bases can cause oxidative DNA damage, base modification, and strand breaks.⁸⁶ The inhibition potential of metal complexes is far better than their ligands because of chelation which causes the delocalization of electrons over the chelate ring and the sharing of positive charge between metals and ligands enhances the lipophilic nature of the complex.⁸⁷ These factors boost the penetration of the metal complexes across the lipid membrane of the microbial cell wall thereby disrupting the membranes, blocking the enzyme's metal binding sites, and stopping the microbial growth.⁸⁸ Moreover, metal complexes can generate the bioactive compound *in situ* which could be another mode of action for metallodrugs. In this regard, chromone-based Cu^{2+} complexes Schiff bases with invigorating structural and electronic properties have received substantial attention.

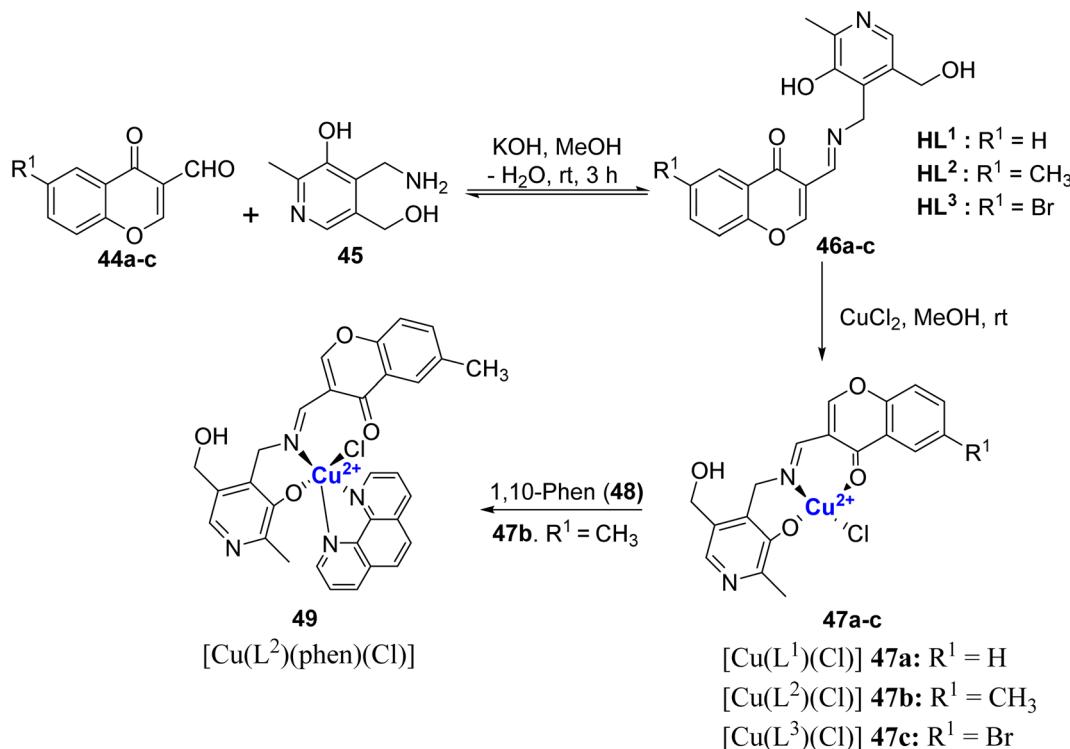
In 2022, Nunes *et al.*⁸⁹ reported chromone Schiff bases **46a–c** synthesized *via* the condensation of 6-substituted-3-formyl-chromones **44a–c** with pyridoxamine **45** (Scheme 12). The Schiff bases **46a–c** were then treated with copper(II) chloride in MeOH at room temperature to furnish $\text{Cu}(\text{II})$ -Schiff base complexes $[\text{Cu}(\text{L}^2)\text{Cl}]$ **47a–c**. Another ternary complex **49** $[\text{Cu}(\text{L}^2)(\text{phen})\text{Cl}]$ was also synthesized by the addition of bidentate 1,10-phenanthroline ligand **48** to complex **47b**. The synthesized complexes were subjected to elemental analysis such as FT-IR, ^1H - and ^{13}C -NMR spectroscopy, UV-Vis, EPR, and mass spectrometry. The investigations indicated that the chloride ion is bonded to the metal ions, which are coordinated to Schiff bases through donor atoms such as *O*-phenolate, *N*-imine, and *O*-carbonyl in compounds **47a–c** and **49**.

Amongst all the complexes, $[\text{Cu}(\text{L}^2)(\text{phen})\text{Cl}]$ **49** was the most stable and displayed higher stability in aqueous media. The presence of BSA protein further increased its stability. For other metal complexes it was predicted that ligand substitution, metal hydroxide formation, precipitation and partial oxidation in aqueous solution could be the reason for their instability. Compounds **46b**, **46c** as well as **47c** and **49** were tested for their cytotoxic characteristics on human cancer cells (MC7, HeLa, A2780, LN-229, and MCF7) and normal cells (RPE-1) with concentration 1.6–100 μM . The incubation periods ranged from 24–72 h with an interval of 24 h each. The metal complexes



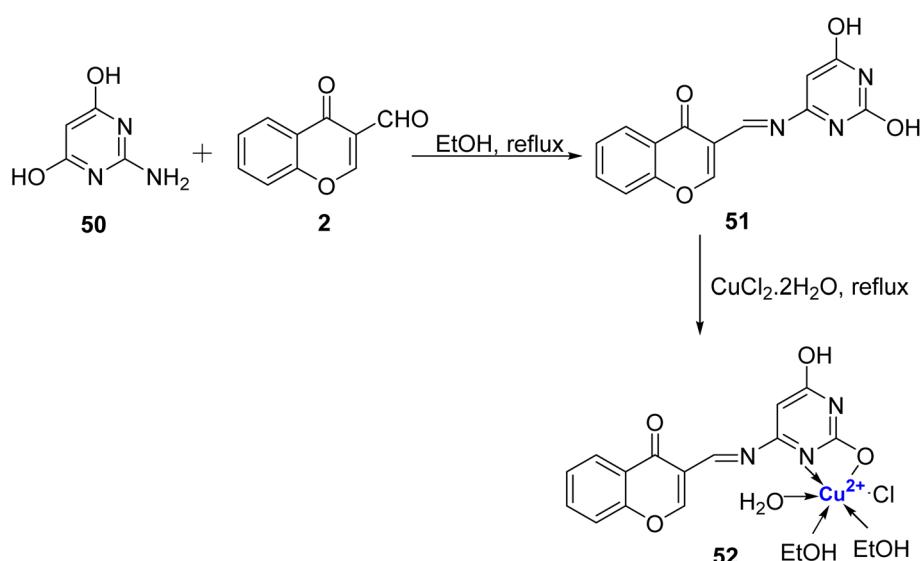
Scheme 11 Synthesis of ligand **42** and their complexation with Cu^{2+} ions **43**.



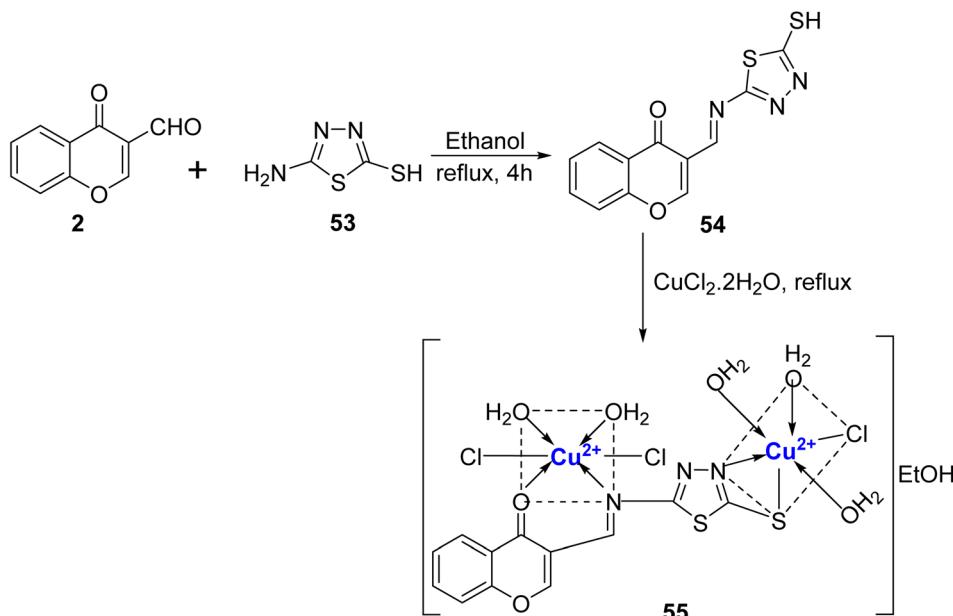
Scheme 12 Synthesis of ligands 46a–c and their complexation with Cu²⁺ ions 47a–c and 49.

showed cytotoxicity in comparison to their ligand precursors but were not selective toward cancer cells. Complex 49 induced DNA cleavage, nicking supercoiled DNA to completion, activated higher levels of ROS and genotoxic damage. It also triggered cell death through apoptosis, as determined by apoptotic body formation, condensation, and fragmentation of DNA and TUNEL assay. The major challenge in this work was low aqueous solubility and less selectivity for cancer *vs.* normal cells.

In 2019, Gaber *et al.*⁹⁰ developed a Schiff base ligand, (*E*)-3-((2,6-dihydroxypyrimidin-4-ylimino)-4*H*-chromen-4-one 51 that was prepared by condensation of 2 with 4-amino-2,6-dihydroxypyrimidine 50 in absolute ethanol in good yields. Further, Cu²⁺ metal complexes were made using CuCl₂·2H₂O. The spectral and magnetic studies revealed that the binding affinity of Cu(II) complex 52 was greater than the ligand 51 and the complexe have octahedral structure (Scheme 13). The ligand and Cu(II), complexes were assessed against *E. coli* and *S. aureus*



Scheme 13 Synthesis of ligand 51 and its Cu(II)-complex 52.



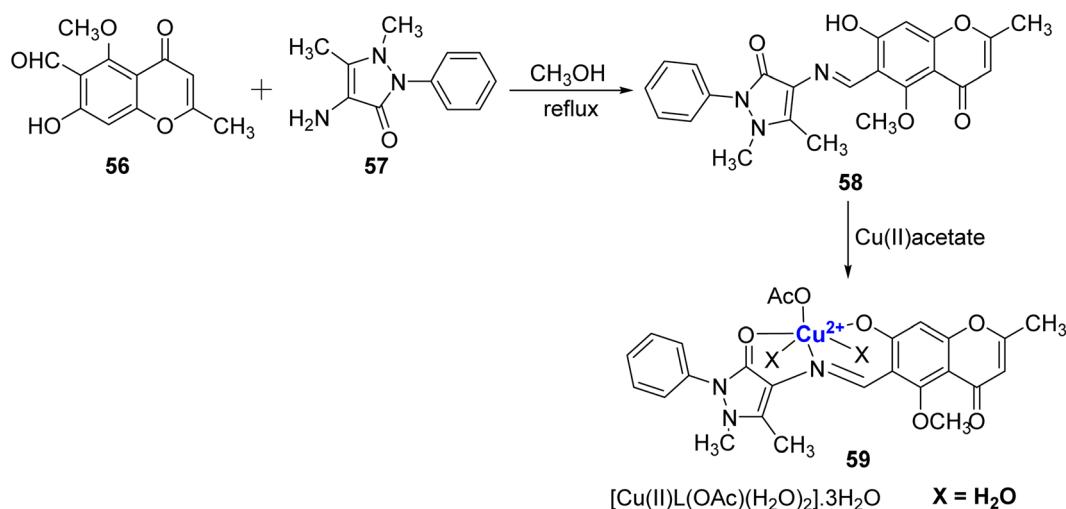
Scheme 14 Synthesis of ligand 54 and its copper(II)-complex 55.

for their antibacterial activity and *C. albicans* and *Aspergillus niger* for antifungal activity. However, complex 52 as well as ligand 51 did not exhibit promising antibacterial and antifungal activity, and inhibition results against human cancer cell line HepG2 was also not very encouraging. The IC₅₀ values for ligand 51 and its corresponding complex 52 against HepG2 were found to be 39.56 and 62.73 $\mu\text{g mL}^{-1}$. The intrinsic DNA-binding constant for Cu(II) complex 52 was found to be $4.3 \times 10^5 \text{ M}^{-1}$ indicating better intercalative-binding interaction with DNA.

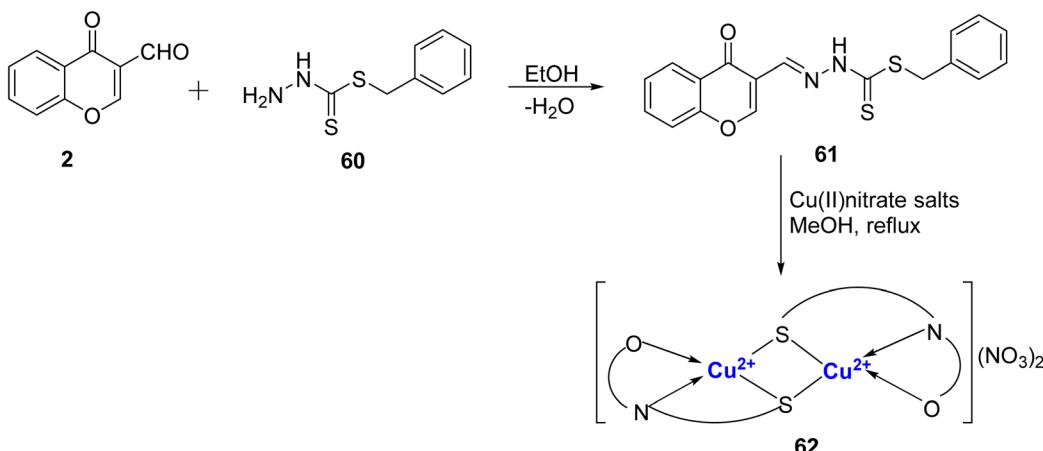
The same research group⁹¹ also designed novel Schiff base, 54 and its Cu(II)-complexes (Scheme 14). The Schiff base 54 was synthesized from chromon-3-carbaldehyde 2 and 53 in ethanol and its Cu(II) complex 55 was prepared by refluxing with hydrated metal chloride, CuCl₂·2H₂O. Characterization of the

Cu(II) complex 55 was done by elemental analysis, inductive coupled plasma, X-ray, IR, EI mass, UV-Vis, ESR, and ¹H- and ¹³C-NMR which revealed octahedral shape for the complex. The complexes showed weak antimicrobial and anticancer activity. The IC₅₀ values for the ligand 54 as well as its complex with Cu²⁺, 55, were quite high *i.e.* 23.00 $\mu\text{g mL}^{-1}$ and 18.40 $\mu\text{g mL}^{-1}$, respectively against HepG2 in comparison to that of doxorubicin, the standard drug (IC₅₀ = 4.73 $\mu\text{g mL}^{-1}$).

A Cu²⁺ complexes of a new chromone Schiff base was prepared by Shakdofa *et al.*⁹² The ligand 58 was synthesized by refluxing 56 and 57 in a 1:1 M ratio (Scheme 15). Copper(II)-based chromone Schiff base complex 59 displayed a tetragonal distorted octahedral geometry. It was synthesized by refluxing a methanolic solution of copper(II)-acetate and the ligand 58 for



Scheme 15 Synthesis of ligand 58 and its copper(II)-complex 59.



Scheme 16 Synthesis of tridentate monobasic ligand 61 and its complex 62 with Cu^{2+} ions.

three hours. The designed molecules were analyzed for inhibition of tumor suppressor protein p53 ubiquitination. Ligand 58 exhibited an *in vivo* IC_{50} value of $12.13 \mu\text{M}$. It is worth noting that inhibiting the interaction between p53 and the oncoprotein, MDM2, resulted in MDM2 blocking the ability of p53 to activate transcription. Copper complex 59 was found to disrupt the p53-MDM2 binding with IC_{50} values of $0.21 \mu\text{M}$ *in vitro*. Moreover, complex 59 with $\text{IC}_{50} = 1.79 \mu\text{M}$ also activated the tumor suppressor p53 present in cancer cells and showed promising results for p53 ubiquitination *in vivo* when compared to the reference drug *i.e.*, diphenylimidazole ($\text{IC}_{50} = 0.26 \mu\text{M}$).

A new monobasic tridentate ligand 61 with sulfur, nitrogen and oxygen of γ -pyrone chelating centers was prepared by the condensation reaction of chromone-3-carboxaldehyde 2 and 5-benzylidithiocarbazate 60 by Adly *et al.* (Scheme 16).⁹³ Further, dimeric complex with Cu^{2+} , was made. The energy gap of the prepared complexes was lower than the free ligand and because of that these complexes were more reactive. The HOMO \rightarrow LUMO electron transition and energy gap (ΔE) of the $\text{Cu}(\text{II})$ complex 62 was found to be 0.434 eV . The smaller energy gap of validates its high reactivity and low kinetic stability. The dimeric complex 62 was found to exist in square planar geometry. The ligand 61 (IC_{50} value = 14.90 mg mL^{-1}) and $\text{Cu}(\text{II})$ -complex 62 (IC_{50} value = 7.37 mg mL^{-1}) were active against HepG2 cancer cell lines when compared to the standard cis-platin (IC_{50} value = 3.67 mg mL^{-1}) and it was the electrophilicity and electronegativity that played a crucial role in their antitumor activities. Complex 62 also showed antimicrobial activity towards the Gram-positive/-negative bacteria as well as with fungus strains.

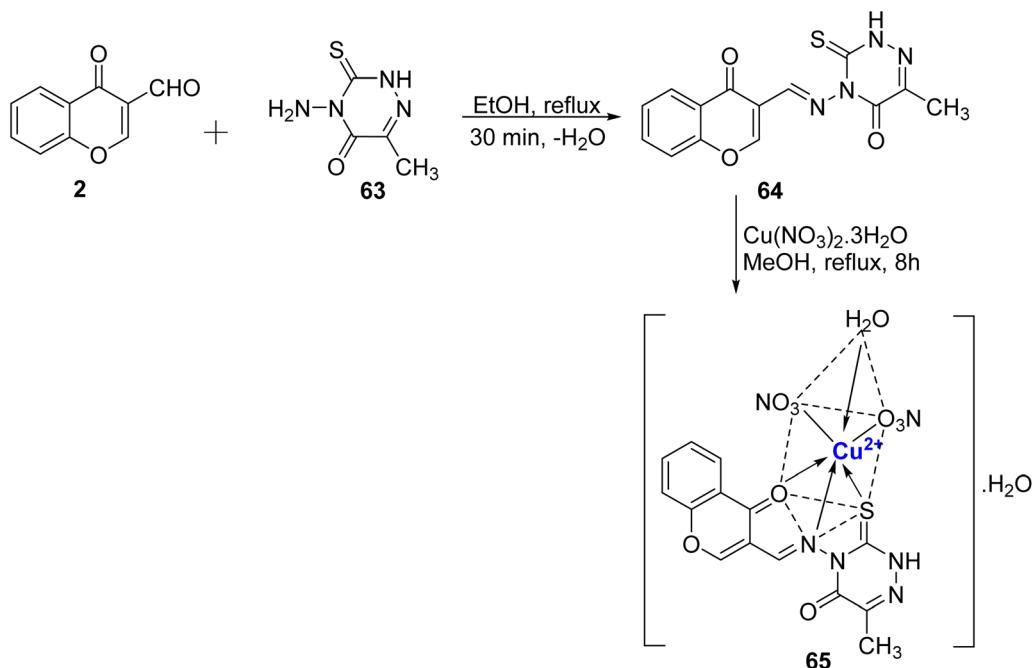
Recently, novel Cu^{2+} nano complex 65 based on hydrazone ligand 64 containing chromone and triazine moieties were synthesized.⁹⁴ The ligand 64 was synthesized by refluxing a mixture of 63 and 2 in ethanol (Scheme 17). The Cu^{2+} nano complex 65 was prepared by refluxing the methanolic solution of the neutral tridentate ligand 64 and $\text{Cu}(\text{NO}_3)_2 \cdot 3\text{H}_2\text{O}$ and was found to possess nano-rode morphology. The nanocomplex 65 exhibited excellent inhibition towards the cancer cell line HepG2 with IC_{50} value of $0.027 \mu\text{M}$. The effectiveness of this

sensor can be attributed to two key mechanisms. Firstly, it operates through reactive oxygen species generated by copper ions. These reactive species are known to inflict damage on the DNA within tumor cells, making it a promising tool in cancer treatment. Secondly, the chelation theory comes into play, where the positive charge on the metal ion boosts the ligand's acidity, enabling it to accept protons. This, in turn, strengthens the hydrogen bonds that pivotally enhance the sensor's biological activity. Docking results revealed Cu^{2+} nano complex 65 as a probable inhibitor of the CDK2 enzyme. Cu^{2+} nano-complex 65 showed better orientation with amino acids Phe82, Leu83, Asp86, and Lys89. It was also noticed that nano-formulations result in better cytotoxicity and proliferation.

A new Schiff base ligand 67 *via* the condensation between 56 and cephradine drug 66 was developed (Scheme 18).⁹⁵ The ligand was monobasic bidentate, and its $\text{Cu}(\text{II})$, complex was also prepared. From experimental data, it was found that $\text{Cu}(\text{II})$ -ligand complex 68 was in the stoichiometric ratio of $1 : 1$ (M : L). The spectrum of $\text{Cu}(\text{II})$ -L complex showed emission band at 443 and 726 nm. The intense fluorescence band emission was assigned to the intra-ligand fluorescence and intramolecular charge transfer between ligand and ion. As corroborated by the photostability studies, the complex 68 was less photostable than the Schiff base ligand 67 as absorption of photonic energy caused irreversible photochemical decomposition-bleaching. The photobiological larvicidal activity was conducted using mosquito larvae to examine the efficiency of complexes in controlling insects using direct sunlight. However, $\text{Cu}(\text{II})$ complexes did not show appreciable results. Cytotoxicity of the ligand 67 and complex 68 was also studied on the HepG2 cell line. The Schiff base ligand 67 displayed a better inhibitory effect of 55% towards the HepG2 cancer cell lines when compared to its $\text{Cu}(\text{II})$ complex 68 with 35% inhibition.

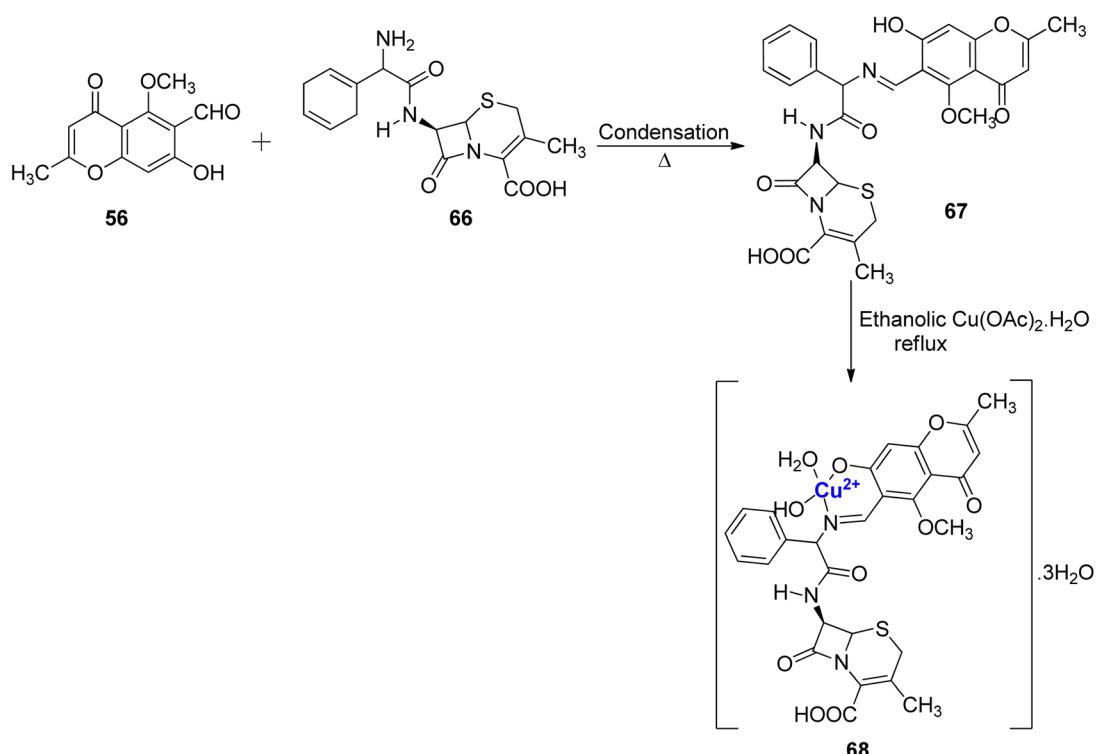
In 2018, Kalaiarasi *et al.*⁹⁶ prepared new Schiff base ligands 70a-b by the reaction of 24 with 69a and phenyl semicarbazone 69b (Scheme 19). The $\text{Cu}(\text{II})$ -complexes 71a-b were synthesized by the reaction of ligand 70a-b with copper(II)-nitrate salts in methanol under reflux. Single crystal XRD studies established the structure of cationic complex 71a as distorted square

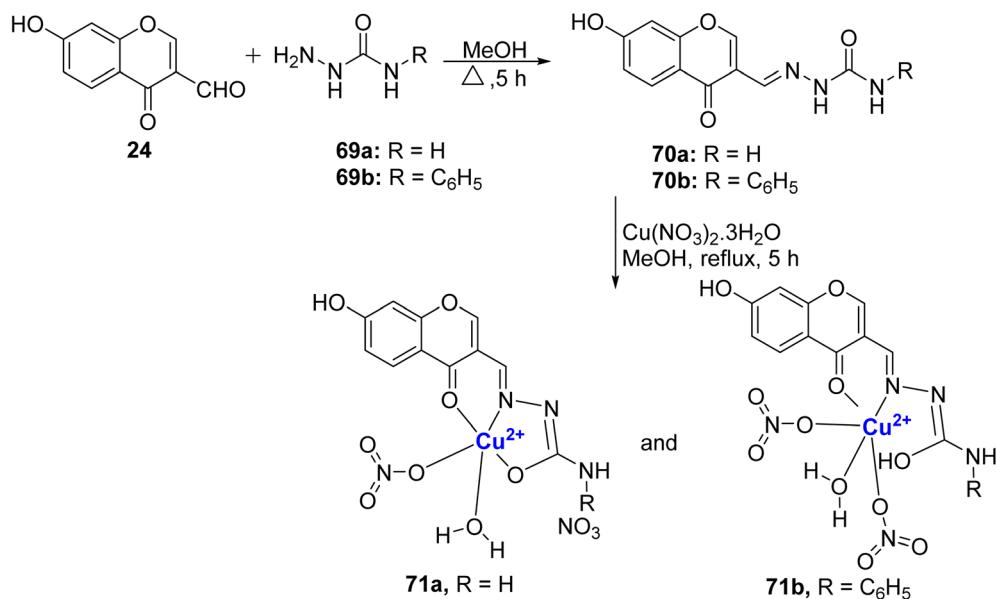


Scheme 17 Synthesis of hydrazone ligand 64 and its complex 65 with Cu^{2+} .

pyramidal geometry and the neutral complex **71b** as octahedral geometry. The intercalative mode of binding compounds with calf thymus DNA was confirmed by ethidium bromide displacement and viscosity measurement studies. Most pronounced EB-DNA fluorescence emission suppression was demonstrated by complex **71a**. The same complex was observed

to induce an increase in the separation of base pairs at the intercalation site within DNA, subsequently leading to an extension in the molecular length of the DNA. Furthermore, in a protein binding study, it was demonstrated that the ligands and complexes exhibited a binding capacity with both BSA and HSA through a static quenching mechanism. To assess their

Scheme 18 Synthesis of ligand 67 and its complexation with Cu^{2+} ions to give 68.

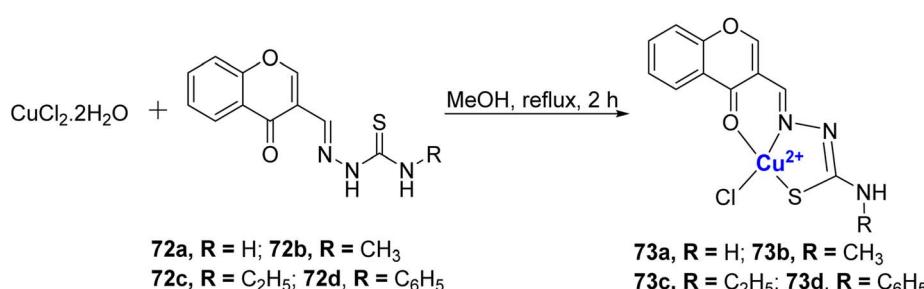
Scheme 19 Synthesis of the ligands 70a–b and their complexation with Cu^{2+} metal ion 71a–b.Table 1 IC_{50} values for ligands 70a and 70b as well as for complexes 71a and 71b for cell lines MCF-7 and A549

Compounds	IC_{50} (μM)	
	MCF-7	A549
$\text{Cu}(\text{NO}_3)_2 \cdot 3\text{H}_2\text{O}$	>50	>50
70b	18.44 ± 0.16	19.15 ± 0.14
70a	15.96 ± 0.14	17.19 ± 0.29
71b	3.52 ± 0.09	3.69 ± 0.06
71a	2.49 ± 0.10	3.33 ± 0.09
Cisplatin	15.10 ± 0.05	16.79 ± 0.08

potential medical relevance, the compounds were tested against MCF-7 and A549 cancer cell lines (Table 1). Notably, complexes 71a and 71b displayed superior cytotoxicity results when compared to the drug cis-platin. The enhanced activity of 71a may be attributed to its cationic nature, which sets it apart from the neutral octahedral complex 71b. These results were further checked by lactate dehydrogenase release assay and nitric oxide assay. The complexes 71a–b also displayed antimicrobial properties. The antibacterial activity of the compounds had the

order: 71b > 71a > 70b > 70a for *S. aureus*, *A. baumannii* as well as *S. Pneumoniae*, while for *P. aeruginosa*, the compounds 71a and 71b had similar MIC values. For *Aspergillus niger*, *Candida tropicalis*, and *Aspergillus fumigatus*, compound 71b had more potency as an antifungal agent over 71a.

The same research group also reported four water-soluble $\text{Cu}(\text{II})$ -chromone complexes 73a–d.⁹⁷ The complexes 73a–d were synthesized from $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ and 3-formyl chromone-4(*N*)-substituted thiosemicarbazones 72a–d ligands and it coordinated with the metal in a tridentate monobasic ONS donor fashion (Scheme 20). The binding ability of synthesized copper thiosemicarbazone to calf thymus DNA was analyzed and the results were in the order 73c > 73b > 73a > 73d. Notably, among the four complexes, the one with an ethyl group, 73c, demonstrated a stronger affinity for DNA as evidenced by its performance in the ethidium bromide (EB) displacement assay and viscosity measurements. The interaction of complexes 73a–d with plasmid pBR322 DNA showed that they cleave the supercoiled DNA. A static quenching mechanism was noted for BSA and HSA serum albumins by the complexes. These $\text{Cu}(\text{II})$ complexes 73a–d displayed considerable antibacterial as well as antifungal activity, where 73c was the most potent. Anti-



Scheme 20 Synthesis of complexes 73a–d.

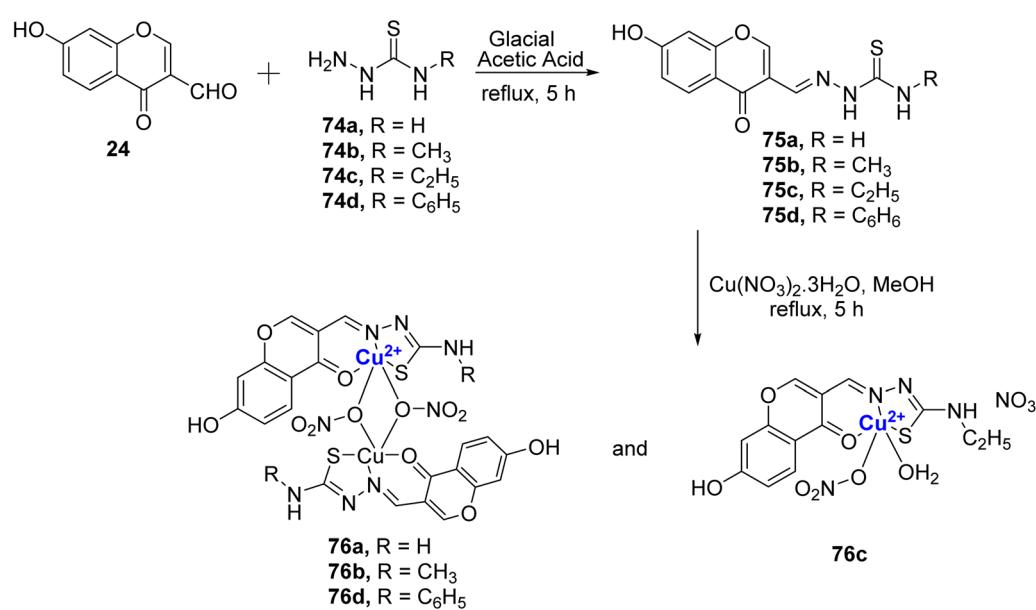
proliferation studies on MCF-7, HeLa, and HaCaT revealed that complexes **73a–d** could overcome cis-platin resistance in both cell lines. Significant cytotoxicity was observed with IC_{50} values of 5.05 ± 0.09 to $7.64 \pm 0.13 \mu\text{M}$ for the MCF-7 cell line and for A549, it was 7.57 ± 0.10 to $8.67 \pm 0.13 \mu\text{M}$ whereas for cisplatin the values are $16.79 \pm 0.08 \mu\text{M}$ for MCF-7 and $15.10 \pm 0.05 \mu\text{M}$ for A549. The synthesized compounds **72a–d** and **73a–d** were non-toxic toward the HaCaT. Complex **73c** was the most effective of all.

The same group further reported four Schiff bases **75a–d** and their resultant water-soluble Cu(II)-complexes **76a–d**.⁹⁸ The Cu(II)-complexes **76a–d** were prepared from different thiosemicarbazones **75a–d**, which were initially synthesized by refluxing the **24** with thiosemicarbazide **74a–d** in glacial acetic acid for 5 hours (Scheme 21). The compounds showed good antibacterial and antifungal properties towards the different tested microbial species. For fungi *T. rubrum* and *C. albicans*, the compounds displayed the activity in the following order: **76c** > **76b** > **76a** > **76d**; for *C. tropicalis*: **76c** > **76d** > **76b** > **76a**; for *A. fumigatus*: **76c** > **76a** > **76b** > **76d**; for *A. niger*: **76c** > **76b** > **76d** > **76a**. Complex **76c** was the most effective against *S. aureus* and *P. aeruginosa* and activity order was as follows, **76c** > **76d** > **76b** > **76a**. In the case of *S. pneumoniae*, the activities of the complexes followed the order **76c** > **76b** > **76a** > **76d**, and against *A. baumannii*, complex **76d** stood out with **76d** > **76a** > **76c** > **76b**. The cytotoxic assessment was also conducted against MCF-7 and A549 cells. Cu(II)-complexes followed the order with IC_{50} values for MCF-7 as **76c** ($2.94 \pm 0.09 \mu\text{M}$) > **76b** ($3.71 \pm 0.06 \mu\text{M}$) > **76d** ($3.87 \pm 0.08 \mu\text{M}$) > **76a** ($4.21 \pm 0.09 \mu\text{M}$); for A549 the inhibitory activity was as follows: **76c** ($2.31 \pm 0.07 \mu\text{M}$) > **76b** ($3.20 \pm 0.05 \mu\text{M}$) > **76d** ($4.00 \pm 0.09 \mu\text{M}$) > **76a** ($4.30 \pm 0.09 \mu\text{M}$). The results were better than the standard drug cis-platin for which IC_{50} for MCF-7 is $15.10 \pm 0.05 \mu\text{M}$ and for A549 it is $16.79 \pm 0.08 \mu\text{M}$. It is noteworthy that complex **76c** which contains an electron-rich

ethyl group displayed the highest activity that further authenticated by LDH and NO release assays. The binding affinity of the complexes with calf thymus DNA confirmed intercalative binding mode, which was also supported by EB displacement and viscosity measurements. Also, the copper complexes quenched the fluorescence of serum albumins through a static mechanism. It is motivating to note that *N*-terminal ethyl substituted thiosemicarbazone exhibited higher cytotoxicity and thus it will be interesting to assess compounds with longer alkyl chain lengths.

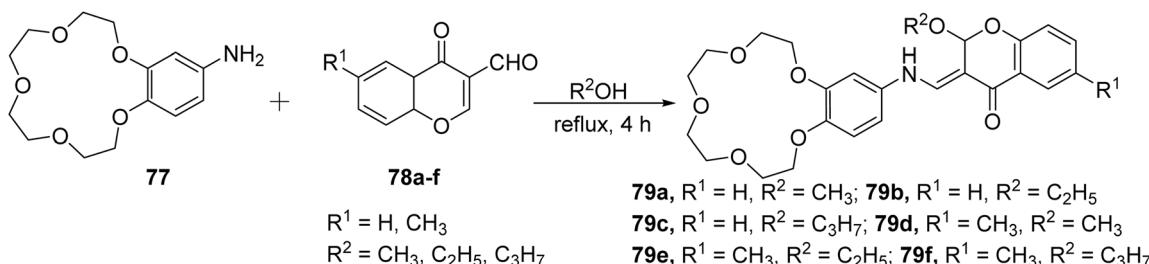
Sahin Gul *et al.* synthesized a series of chromone-crown ether-based Schiff base ligands **79a–f**.⁹⁹ The chromone-crown ethers **79a–f** were prepared by the reaction of **77** with 6-substituted-3-formylchromones **78a–f** (Scheme 22). The synthesized Schiff base ligands **79a–f** were tested as chemosensors using UV-visible and fluorescence spectroscopy and were found to be selective for Cu²⁺ and Fe³⁺ ions in the presence of various other competing ions. A sharp blue shift in the absorption spectrum and fluorescence quenching was detected for chromone compounds **79a–f**, on increasing the metal concentration. The synthesized ligands were also found active against both Gram-positive, Gram-negative bacteria and displayed good antifungal activity.

Synthesis of a Cu(II)-complex **82** from Schiff base ligand **81** has been attempted.¹⁰⁰ The ligand **81** was prepared by the reaction amid **80** and **56** (Scheme 23). Characterization of both the ligand **81** and the complex **82** was done by elemental analysis, Mass, FT-IR, thermal analysis, electronic spectra, magnetic susceptibility measurements, and conductivity. The analytical data confirmed 2 : 1 stoichiometry for the Cu₂-**81** complex with square planar geometry. However, the copper complex showed minimal biological activity towards *S. aureus*, *B. Subtilis*, *P. aeruginosa* and *E. coli* bacteria. This may be due to their lower permeability across the bacterial cell membrane. The studies

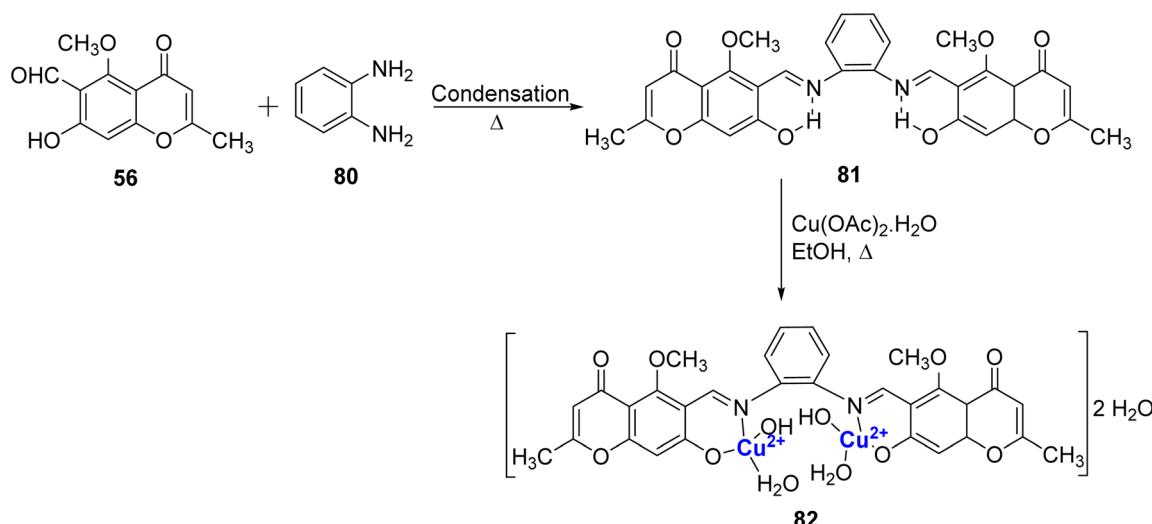


Scheme 21 Schiff base ligands **75a–d**, were synthesized, and subsequently, Cu(II)-complexes **76a–d** were formed.





Scheme 22 Synthesis of chromone crown-ether based ligands 79a-f.



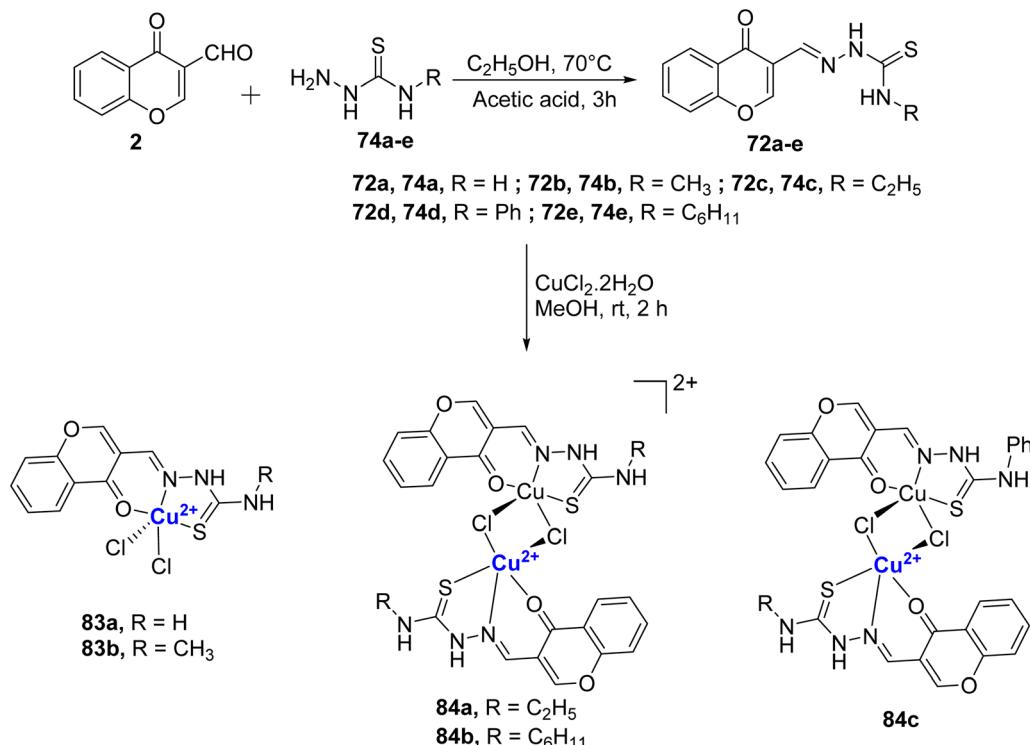
Scheme 23 Synthesis of ligand 81 and its Cu(II)-complex 82.

show that penetration of the complexes through the lipid membrane is crucial for the inhibition of microbial growth.

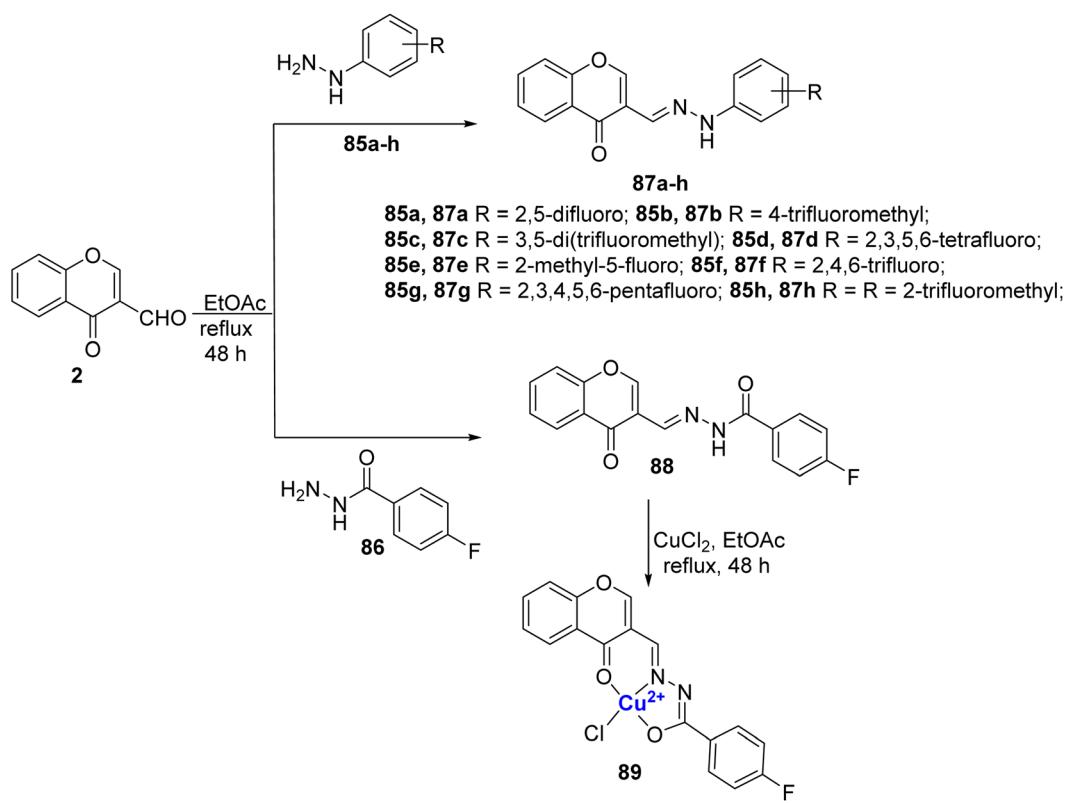
Balakrishnan *et al.* synthesized different chromone appended thiosemicarbazone ligands 72a-e by refluxing 2 with thiosemicarbazides 74a-e and treated them with copper salt to procure copper complexes 83a-b and 84a-c (Scheme 24).¹⁰¹ It was inferred that the variation in the terminal *N*-substitution in the ligands influenced the complexes' stoichiometry and due to the bulkiness of -C₂H₅, -C₆H₁₁ and -C₆H₅ in the ligands, dicationic bimetallic complexes 84a-b and neutral bimetallic 84c were formed. Monometallic complexes of Cu(II), 83a and 84b were obtained from 72a and 72b, respectively. The complexes were found to be stable under physiological conditions. The designed complexes 83a-b and 84a-c displayed catecholase-mimicking activity, and the result showed that except 84c with the bulky phenyl group, all other complexes 83a-b, 84a-b could oxidize 3,5-di-*tert*-butylcatechol into 3,5-di-*tert*-butylquinone molecule in the presence of air or aerobic conditions. It was seen that the bimetallic complexes 84a and 84b underwent dissociation into monomers as a necessary step to participate in the catalytic cycle. The catalytic activity followed the order 84a > 84b > 83b > 83a. Phosphatase like activity of the complexes were also studied with the help of 4-nitrophenylphosphate (4-NPP). The catalytic ability of the Cu(II) complexes to hydrolyse the phosphomonoester followed the

order 84a > 84b > 83a > 83b > 83d. Complexes 84a-b showed superior radical scavenging activity due to their cationic nature and electron-releasing group. On the other hand, mononuclear 83b with electron-donating methyl group showed better activity in comparison to the binuclear complex 84c with an electron-withdrawing group. Each complex exhibited the capacity to prevent hemolysis and, importantly, did not display any toxicity towards red blood cells. Three complexes 84a-c also showed better cytotoxicity towards the HeLa-cancer cells exhibiting IC₅₀ values of 2.24 μM (84a), 2.25 μM (84b), and 3.77 μM (84c) that is two times higher activity when compared to the standard drug cis-platin. The complex 84a-b displayed full inhibition of the colony formation at 10 μM. The results are promising and further development could lead to a potential anticancer metallodrug.

Slomiak *et al.*¹⁰² prepared a library of hydrazine 87a-h and hydrazide derivatives 88 of 3-formylchromone 2. The ligands 87a-h and 88 were synthesized by reacting 2 with different hydrazines 85a-h or hydrazide 86 (Scheme 25). The neutral mononuclear copper(II)-complex 89 was synthesized by the treatment of the ligand 88 with copper(II) chloride. On studying their antimicrobial and antiproliferative properties, it was observed that the compounds 87a-h, 88, and 89 were capable of inhibiting the growth of microorganisms. Complex 89 was found to have improved antiproliferative properties than ligand



Scheme 24 Synthesis of ligands 72a–e and their mono- 83a–b/bi-metallic 84a–c Cu(II) complexes.



Scheme 25 Hydrazine derivatives 87a–h and hydrazide derivatives 88, originating from 3-formylchromone, were synthesized, and Cu(II)-complex denoted as 89 was also prepared.



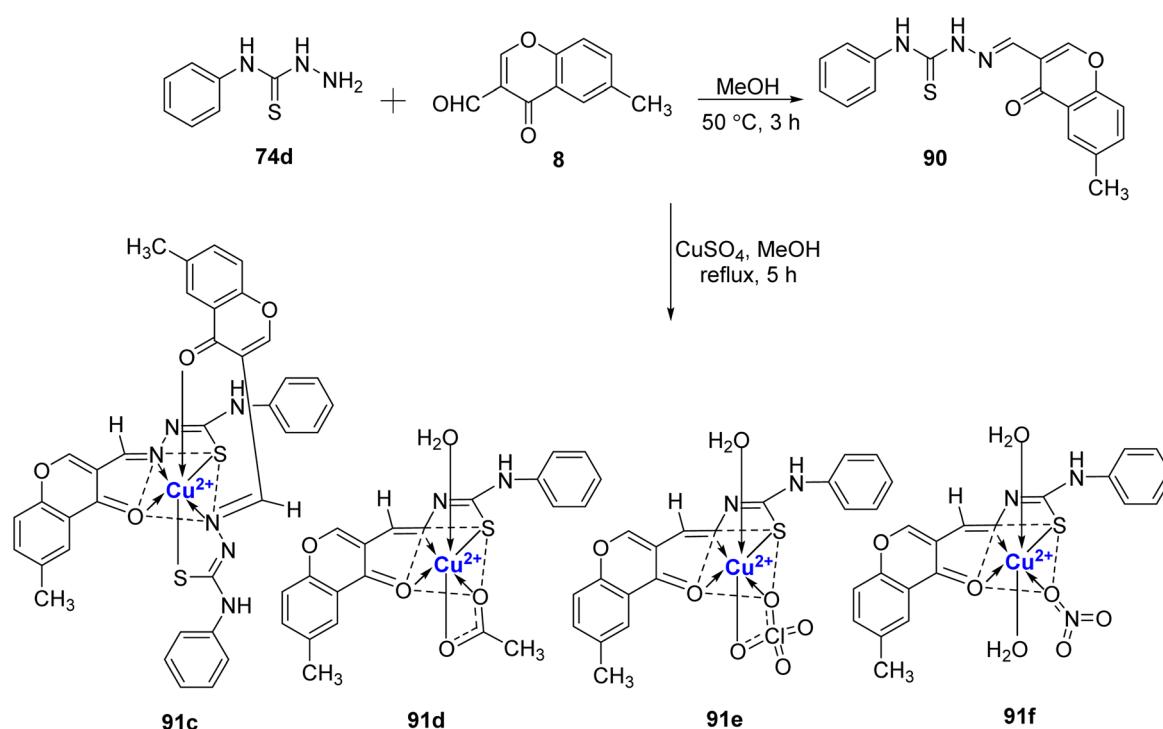
88, with its IC_{50} value $35.01 \mu\text{mol L}^{-1}$ for the L929 and $0.04 \mu\text{mol L}^{-1}$ for EA.hy926.

Copper complexes **91a–f** were synthesized from the ligand, 3-formyl-6-methylchromone-4-phenylthiosemicarbazone **8** using copper(II) salt solutions in molar ratio of 1:2 (M:L) by Ilies *et al.*¹⁰³ The ligand was synthesized by treating 4-phenylthiosemicarbazide **74d** with 3-formyl-6-methylchromone **8** in a methanolic solution (Scheme 26).¹⁰⁴ Complex **91a** was found to have a distorted square-planar shape and complex **91b** was found to be square-pyramidal. All of the Cu(II)-complexes **91a–f** exhibited antimicrobial activity against *S. aureus*, *E. faecalis*, *E. coli*, *S. enteritidis* and *C. albicans*. Complex **91e** displayed the most promising activity towards all the strains with MIC values between $16 \mu\text{g mL}^{-1}$ to $64 \mu\text{g mL}^{-1}$ and it was attributed to the bulky ClO_4^- anion. The antifungal and antibacterial data showed that the metal complexes **91a–f** have higher activity in comparison to the free ligand **90**.

Furochromone-based Cu(II)-complex of Schiff base, resulting from the reaction of **57** and **92**, has been reported.¹⁰⁵ Initially, the furochromone based ligand **93** was synthesized from **57** and **92** (Scheme 27). Copper complex **94** was prepared by the treatment of ligand **93** with the metal salt $\text{Cu}(\text{CH}_3\text{COO})_2 \cdot \text{H}_2\text{O}$. The structure of the complex **94** was reported to be of distorted octahedral geometry that was confirmed *via* spectral techniques and elemental analysis. All substances were tested for their *in vitro* antimicrobial activity. The findings demonstrated that copper complex **94** to be effective against *C. albicans* and *A. niger*. It showed moderate activity against *E. coli* and *S. aureus* whereas with *A. faecalis* and *B. subtilis*, very little activity was reported.

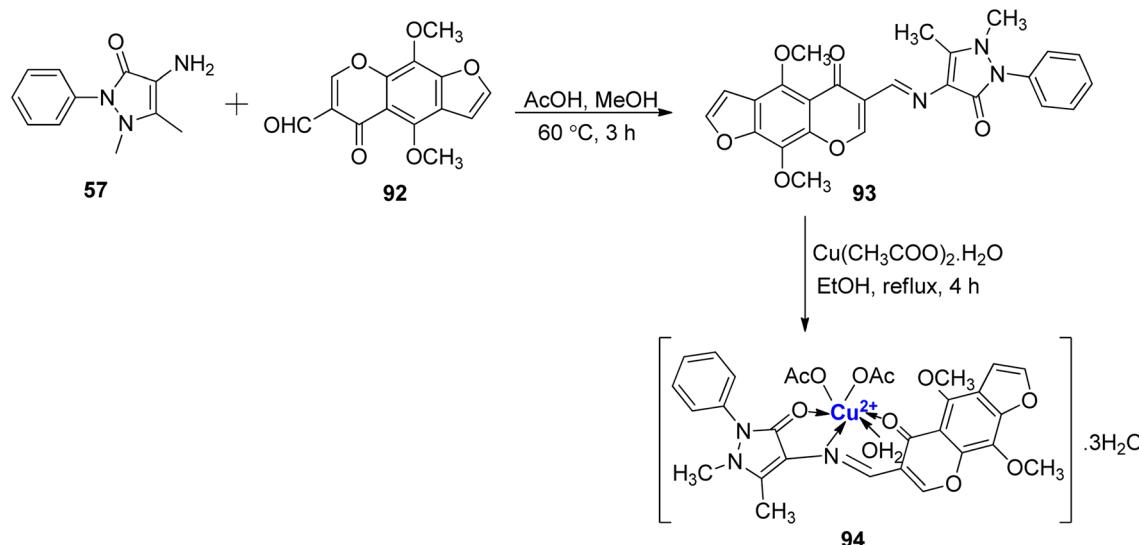
In the same year, Cu(II)-chromanone complexes **98a–b** was reported by Jose *et al.*¹⁰⁶ Firstly, the methoxy-substituted ligands **97a–b** were synthesized by the reaction of **95** with 2-amino-pyridine **96a**/2-amino-5-nitropyridine **98b** in the presence of methanol (Scheme 28). The complexes **98a–b** were then prepared by treating the ligands **97a–b** with the metal salt at 45°C . The Cu(II)-complexes **98a–b** thus synthesized were found to show exceptional stability due to tetragonal distortion and Jahn-Teller effect. Among the prepared compounds, The complex CuL **98b** exhibited the most potent α -amylase inhibitory activity, yielding an IC_{50} value of $0.251 \pm 0.2 \text{ mM}$. Conversely, the complex CuL **98a** demonstrated the highest α -glucosidase activity, with an IC_{50} value of $0.060 \pm 0.3 \text{ mM}$. Furthermore, these compounds were screened for their antimicrobial effectiveness. Moreover, the complexes **98a–b** also displayed excellent antibacterial activity with MIC of $15.3 \mu\text{g mL}^{-1}$ against *S. aureus* which was found equivalent to that of the standard drug.

A Cu(II), complex was designed and prepared using Schiff base **100** was synthesized by Pahontu *et al.*¹⁰⁷ The ligand **100**, was prepared from **57** and **99** in ethanol (Scheme 29), and it was further treated with methanolic CuBr_2 solution to yield the desired complex **101** in 73% yield. The structure of complex **101** was confirmed as tetrahedral with the aid of analytical techniques. Copper complex **101** showed both bacteriostatic and bactericidal properties against Gram-positive (concentration: $0.007\text{--}0.25 \mu\text{g mL}^{-1}$) as well as Gram-negative bacteria (concentration: $0.0312\text{--}0.5 \mu\text{g mL}^{-1}$). Additionally, ten cancer cell lines, namely MSC, A375, B16 4A5, HT-29, MCF-7, HEp-2, BxPC-3, RD, MDCK, and L20B, were employed to evaluate the

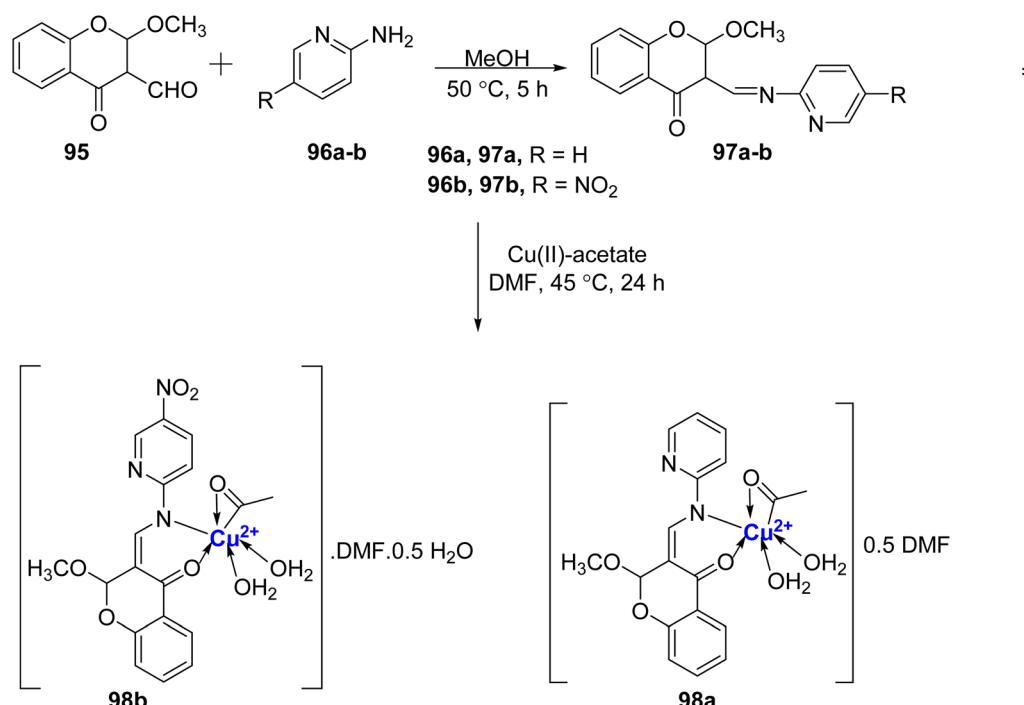


Scheme 26 Synthesis of ligand **90** and their Cu(II)-complexes **91c–f**.





Scheme 27 Synthesis of ligand 93 and its copper(II)-complex 94.



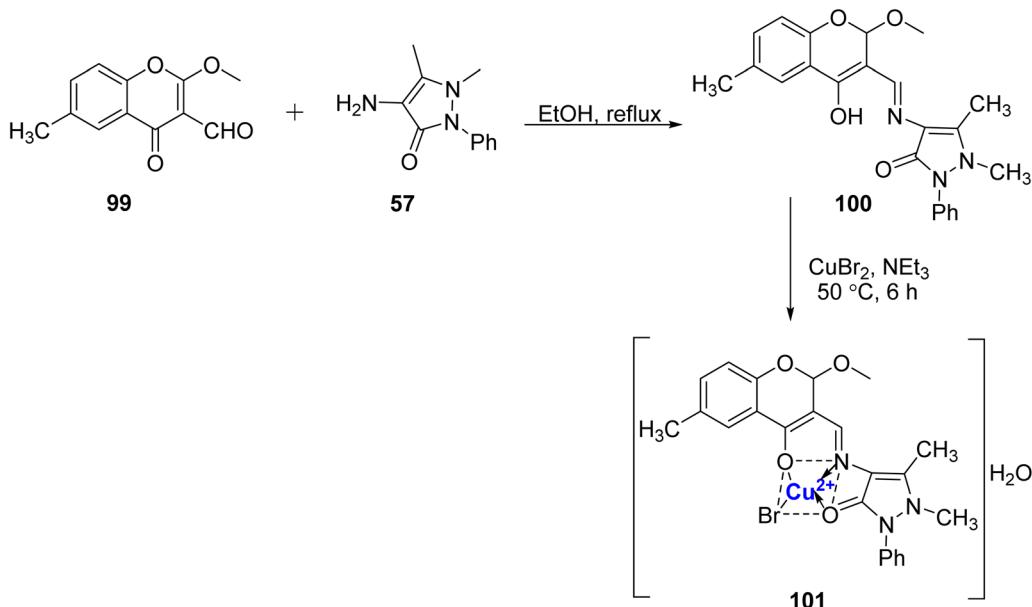
Scheme 28 Synthesis of 2-methoxy-4-chromanone based ligands 97a-b and its copper(II)-complex 98a-b.

in vitro antiproliferative properties of the ligand and complex. It was found that ligand **100** also has promising antimicrobial properties. The ligand **100** showed promising antifungal results, but the copper complex **101** was found to be inactive in comparison to the standard drugs nystatin and miconazole.

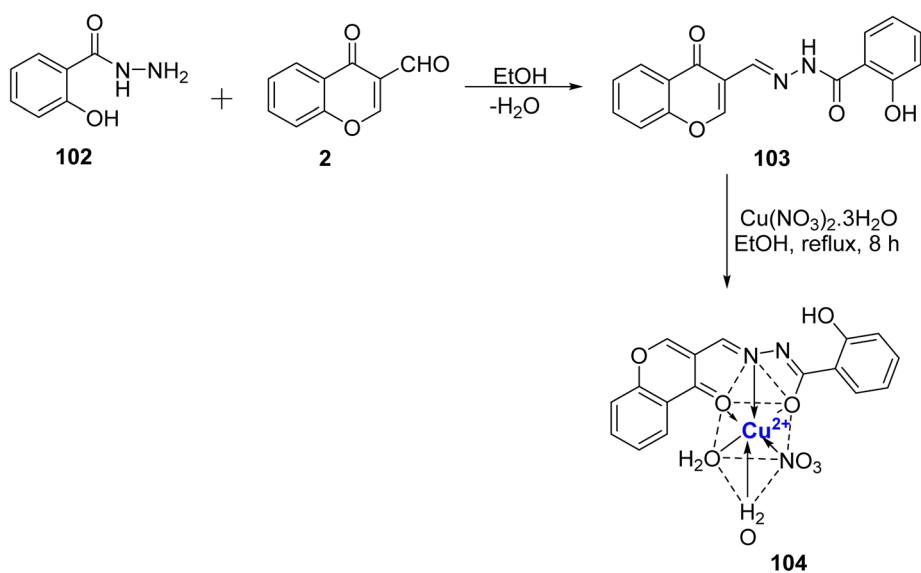
A new class of octahedral nano-complex of Cu^{2+} **104** with chromone Schiff base **103** was prepared by Saif *et al.*¹⁰⁸ (Scheme 30). The complex **104** was synthesized by refluxing **103** with copper-nitrate salt in ethanol as a reaction medium. The synthesized complex **104** exhibited excellent antioxidant activity

($\text{IC}_{50} = 0.93 \mu\text{M}$) as compared to the standard used (ascorbic acid). The $\text{Cu}(\text{II})$ nano-complex **104** demonstrated significant effectiveness in inhibiting the growth of EAC cells, with an IC_{50} value of $47 \mu\text{M}$. This efficacy surpassed that of its parent compound and the other complexes that were synthesized. Moreover, $\text{Cu}(\text{II})$ **104** nano-complex was also found to show cytotoxic effects and was less toxic than *cis*-platin. The chemical structure of complex **104** was confirmed *via* elemental analysis.

Kavitha *et al.*¹⁰⁹ used 3-formylchromone **2** and 2-amino-pyridine **96a** as the reactants to synthesize the ligand **106**, and



Scheme 29 Synthesis of ligand 100 and its Cu(II)-complex 101.

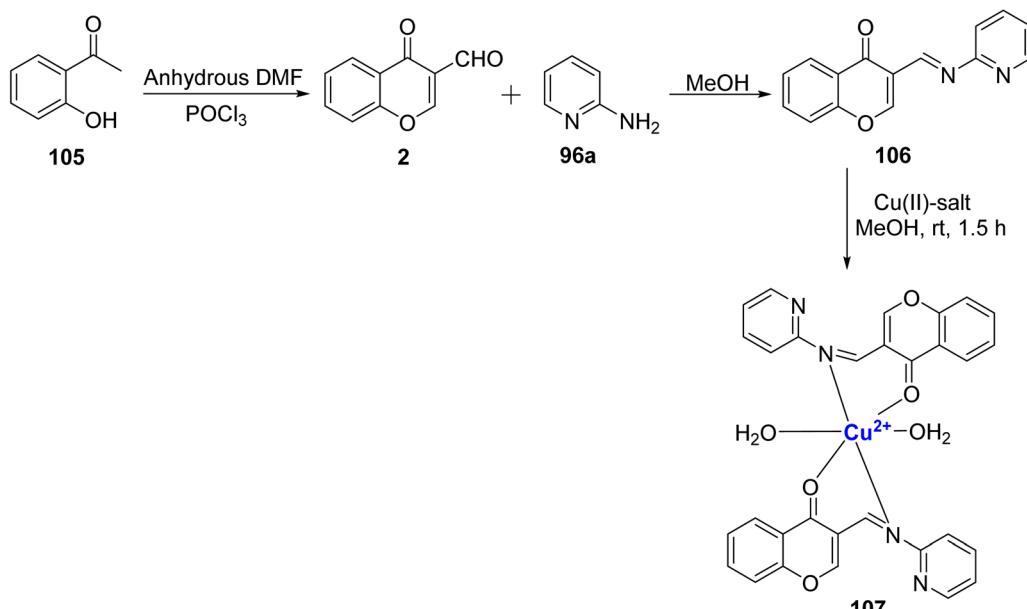


Scheme 30 Synthesis of ligand 103 and its octahedral nano-copper(II)-complex 104.

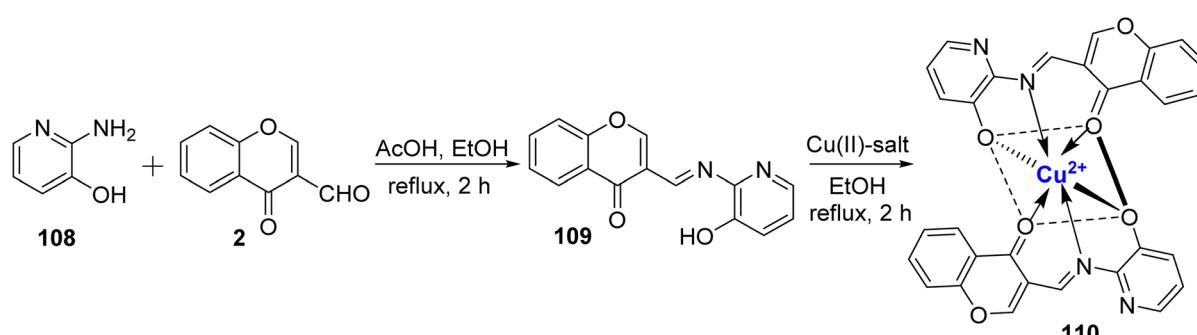
its complex **107** with Cu²⁺ ion (Scheme 31). Characterization of the ligand **106** and complex **107** was done by analytical methods including IR, ESR, XRD, and SEM. Based on magnetic and electronic spectrum data, the complexes had octahedral geometry. The nematicidal and antibacterial effects of the metal complex were stronger than those of the parent ligand. In the presence of H₂O₂, the ligand and its metal complex DNA cleaving activity was detected.

Ammar *et al.*¹¹⁰ conducted a study in which they synthesized a metal Cu(II) complex **110**. These complexe were derived from a tridentate ligand **109**, which was prepared using readily available starting materials, **108** and **2**, with the assistance of

a catalytic amount of acetic acid (Scheme 32). Both the ligand **109** and its corresponding complex **110** were thoroughly characterized using spectral data and elemental analysis. The octahedral geometry of complex **110** was verified through various methods, including DFT calculations, UV-Vis spectroscopy, and ligand field parameters. Furthermore, the researchers assessed the antibacterial properties of both the ligand and the metal complex *in vitro* against a range of bacterial and fungal strains. The collected findings support the investigated compounds potential as bactericides and fungicides. The most effective cytotoxic compound against malignant cells is the Cu(II) complex. Against *B. subtilis*, every substance exhibited



Scheme 31 Synthesis of ligand 106 and its copper(II)-complex 107.



Scheme 32 Synthesis of ligand 109 and its copper(II)-complex 110.

antibacterial action. The data demonstrates that the Cu(II) complex had substantial activity against *S. aureus* and *E. coli*, respectively. Cu(II) complex also displayed the antioxidant activity.

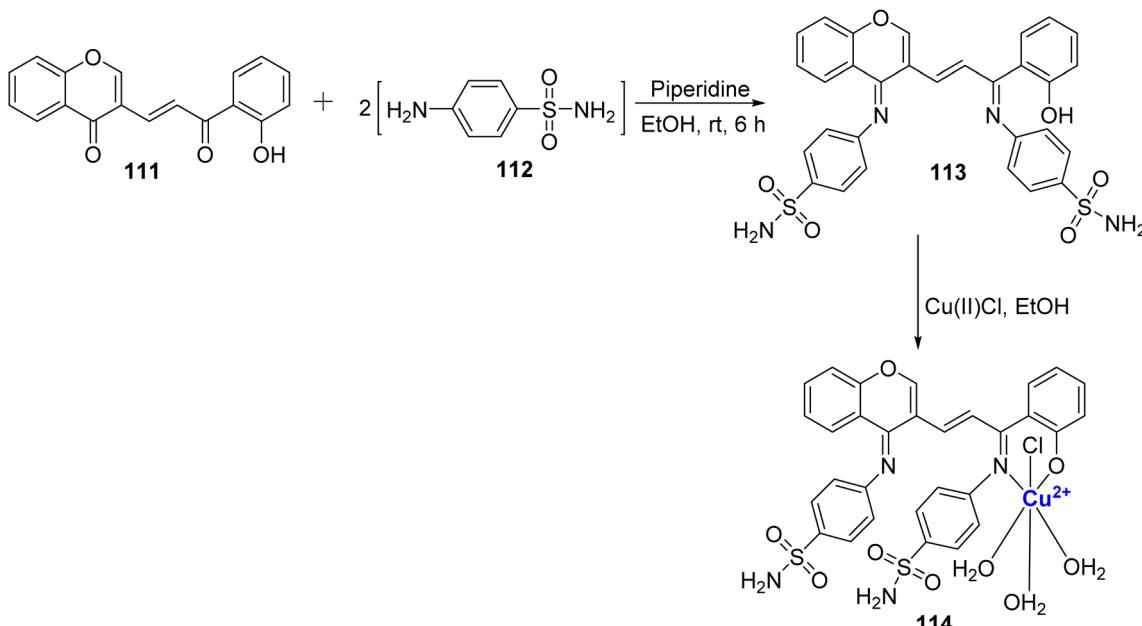
A fluorescent, octahedral, and non-electrolytic Cu(II)-complex 114 was reported by Sumathi *et al.*¹¹¹ (Scheme 33). The ligand 113 was synthesized using 111 and sulphanilamide 112 in the presence of piperidine. The structure of ligand 113 and complex 114 was further evaluated *via* various analytical techniques such as IR, NMR, *etc.* Generally, mostly metal chelates show higher biological activity because of the chelation theory.¹¹² In this context as well, it's worth noting that most metal chelates exhibit superior antimicrobial activity when compared to ligand 113. The developed complex 114 may also assist as a photoactive compound as shown by its fluorescence studies.

Padmaja *et al.*¹¹³ synthesized a Cu(II) complexes with ligand 116 and further characterized both the ligand 116 as well as the metal complex 117 *via* elemental/thermal analysis, ESR studies,

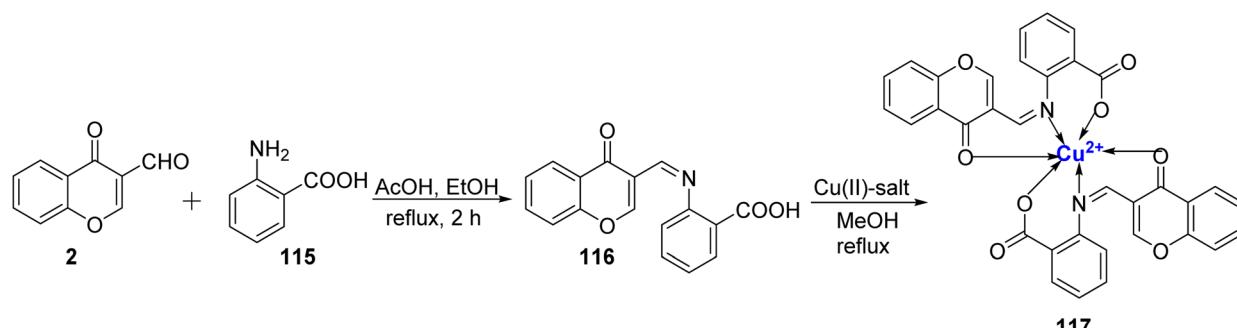
magnetic susceptibility, and spectroscopic techniques. All six complexes share a common octahedral coordination geometry surrounding the metal ion. Ligand 116 was synthesized through a condensation reaction between 2 and 115 (Scheme 34). In these complexes, the ligand 116 interacts with the metal ion in a 1:2 stoichiometric ratio. It's worth noting that the metal complexes 117 displayed superior antimicrobial properties compared to the unbound ligand 116.

In yet another research by Kavitha *et al.*,¹¹⁴ fluorescent Cu(II)-complexes 120a-d were synthesized with four Schiff bases, 119a, 119b, 119c, and 119d (Scheme 35). The ligands, 119a-d, along with the corresponding complexes 120a-d, were subjected to comprehensive characterization using techniques such as mass spectrometry, as well as ¹H and ¹³C nuclear magnetic resonance (NMR) spectroscopy. The complexes 120a-d adopted a tetragonally distorted octahedral geometry, while the ligands 119a-d coordinated with the Cu(II) metal ion in a tridentate manner. Significantly, the complexes 120a-d displayed superior antimicrobial properties when compared to the





Scheme 33 Synthesis of ligand 113 and its copper(II)-complex 114.



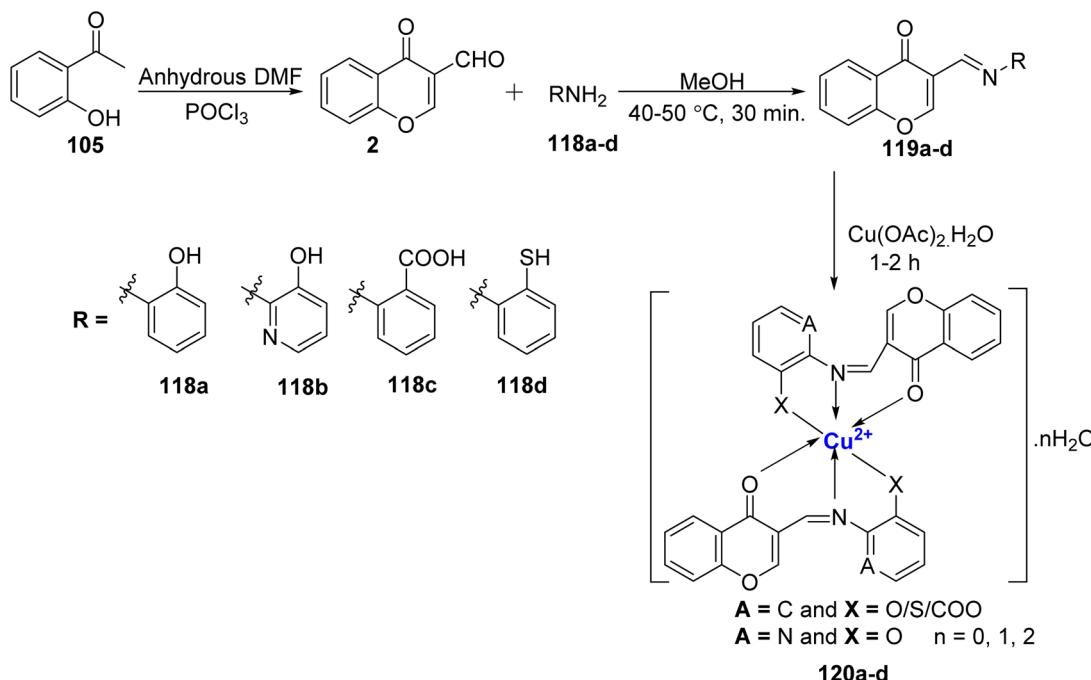
Scheme 34 Synthesis of ligand 116 and its complex 117 with Cu(II).

ligands **119a–d**. Furthermore, the radical scavenging activities of the synthesized compounds were assessed based on their IC_{50} values. Notably, complex **120a** exhibited a particularly low IC_{50} value of $0.16 \mu\text{g mL}^{-1}$, indicating strong radical scavenging activity. Further, the scavenging properties based on the IC_{50} values was in the following order: **120a** > **120b** > **120c** > **120d**.

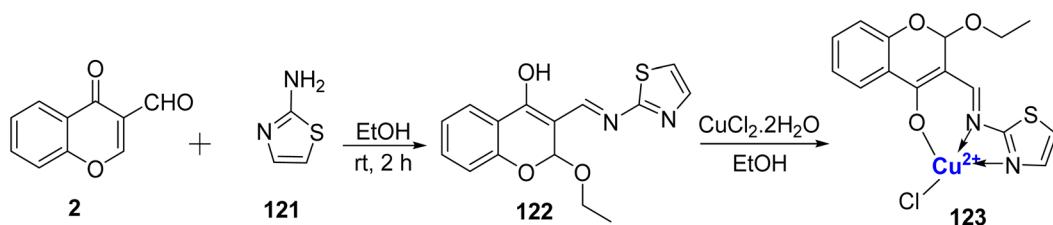
In 2011, Kalanithi *et al.*¹¹⁵ synthesized coordination compound involving Cu(II), with Schiff base **122**. This Schiff base was obtained through the condensation of 3-formyl chromone **2** and 2-aminothiazole **121**, as outlined in Scheme 36. The structural confirmation of these complex was established through various spectroscopic techniques, including EPR, NMR, mass spectrometry, and magnetic susceptibility measurements. The ligand **122** coordinated to the metal ion from three sites namely, enolic oxygen, the nitrogen of the thiazole ring as well as from nitrogen of the azomethine group, thereby behaving as a tridentate entity. Due to chelation effects, the copper complex **123** was potent against pathogens. Therefore, chelation affected the biological outcome of the

synthesized copper complex. The inhibition ability of the complex **123** against bacteria *C. albicans* also showed promising outcomes.

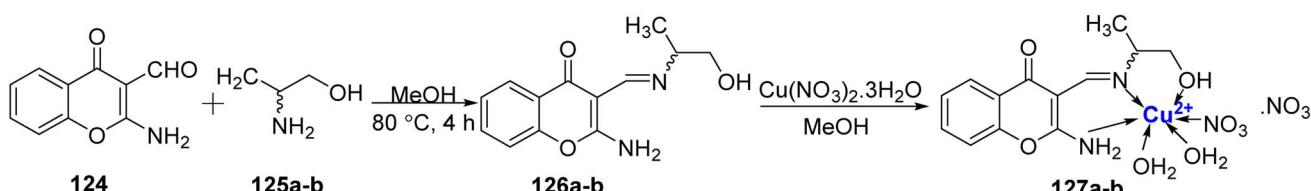
DNA cleavage ability of Cu(II)-metal complex **127a–b** was studied by Arjmand *et al.*¹¹⁶ in 2012. (*R*)- and (*S*)-2-amino-3-((1-hydroxypropan-2-yl)imino)methyl)-4*H*-chromen-4-one **126a–b** were prepared from **124** and **125a–b**, which on complexation with Cu(II) ion afford complexes **127a–b** (Scheme 37). They were then characterized by NMR, mass, IR, elemental analysis, and molar conductance calculations. It was observed that the complexes **127a–b** preferred to attach themselves to the guanine-cytosine region of the DNA molecule and the (*R*) enantiomer **127a** was found to be more active as compared to the (*S*) enantiomer **127b**. Topoisomerase II inhibition property as well as the cytotoxic effects of the complexes **127a–b** against human carcinoma lines were also examined. Complex **127a** was found to be selective for two cancer cell lines: A2780 (GI_{50} value = $17.6 \mu\text{g mL}^{-1}$) as well as MCF-7 (GI_{50} value = $18.4 \mu\text{g mL}^{-1}$), however complex **127b** displayed moderate activity with GI_{50}



Scheme 35 Synthesis of ligands 119a-d and their Cu(II)-complexes 120a-d.



Scheme 36 Synthesis of ligand 122 and its Cu(II)-complex 123.



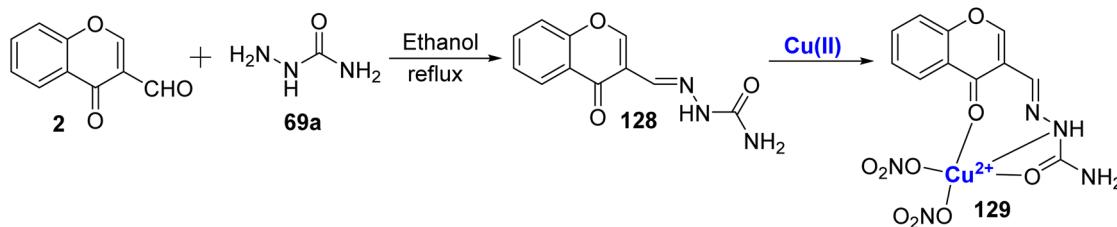
Scheme 37 Synthesis of ligand 126 and their Cu(II)-complexes 127a-b.

value of $20 \mu\text{g mL}^{-1}$ for A2780 and $26.6 \mu\text{g mL}^{-1}$ for MCF-7 cell lines. For cell lines Zr-75-1, SiHa as well as A549, the activity was observed to be very marginal.

In 2010, Li *et al.*¹¹⁷ also utilized Cu(II)-complex 129 in studying DNA binding properties. The characterization of 3-carbaldehyde-chromone semicarbazone 128 and its complex with Cu(II) was carried out using a variety of methods, including crystallography. Complexes showed better binding interactions as compared to the free ligand 128 as studied by spectroscopic measurements (Scheme 38). The developed complex 129 interacted with DNA through intercalation binding mode. The

antioxidant activity of these compounds is quantified by their IC₅₀ in μM against hydroxyl radicals (HO^\cdot). The IC₅₀ values for ligand 128 and its Cu(II) against HO^\cdot are 10.170 , and $1.195 \mu\text{M}$, respectively. Notably, the metal complex exhibit significantly higher scavenging activity against hydroxyl radicals when compared to standard antioxidants like mannitol (IC₅₀: $10.19 \mu\text{M}$). For scavenging superoxide anions, the IC₅₀ values for the ligand, and Cu(II) complex are 32.810 , and $0.943 \mu\text{M}$, respectively. The metal complex 129 exhibit superior antioxidant properties compared to the ligand. Thus, the study showed that the metal ions can act as selective scavenging agents in





Scheme 38 Synthesis of ligand 128 and its complex 129 with Cu(II) ions.

biological systems and paved a pathway for further advancements in this field.

Yang *et al.*¹¹⁸ developed the transition metal complexes of **83a** to study its fluorescence as well as DNA binding by spectral and viscosity studies (Scheme 39). The ligand **83a** was synthesized from ethanolic solutions of 3-carbaldehyde chromone **2** and thiourea **74a**. Ligand **83a** was reacted with copper(II)nitrate in ethanol at reflux to afford the complex **130**. Moreover, the antioxidant properties (superoxide dismutase activity) for ligand **83a** ($IC_{50} = 263.028 \mu\text{M}$) and its corresponding Cu^{2+} complex **130** ($0.799 \mu\text{M}$) were also tested and found to be significant and higher than that of the standards used (IC_{50} for vitamin C = $852 \mu\text{M}$). The ligand **83a** and complex **130** were characterized by various structural methodologies. The Cu(II) complex demonstrates superior antioxidant activity against superoxide and hydroxyl radicals and exhibits stronger scavenging effects.

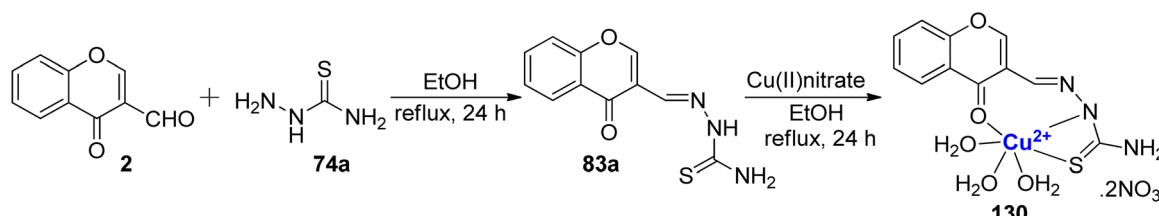
Rosu *et al.*¹¹⁹ conducted a study in which they synthesized coordination compounds **132a-f** with Cu(II) ions, using the Schiff base ligand **131**. The Schiff base, **131**, was obtained through the reaction between **57** and **8** as outlined in Scheme 40. To characterize these compounds, various techniques such as NMR, FT-IR, UV-Vis, ESR spectroscopy, X-ray diffraction, molar electric conductibility, and elemental analysis were employed. Furthermore, the antibacterial activity of the synthesized compounds was investigated *in vitro*. The results of the antibacterial study clearly indicated that the antibacterial properties of the Schiff base compounds were significantly enhanced when coordinated with metal ions.

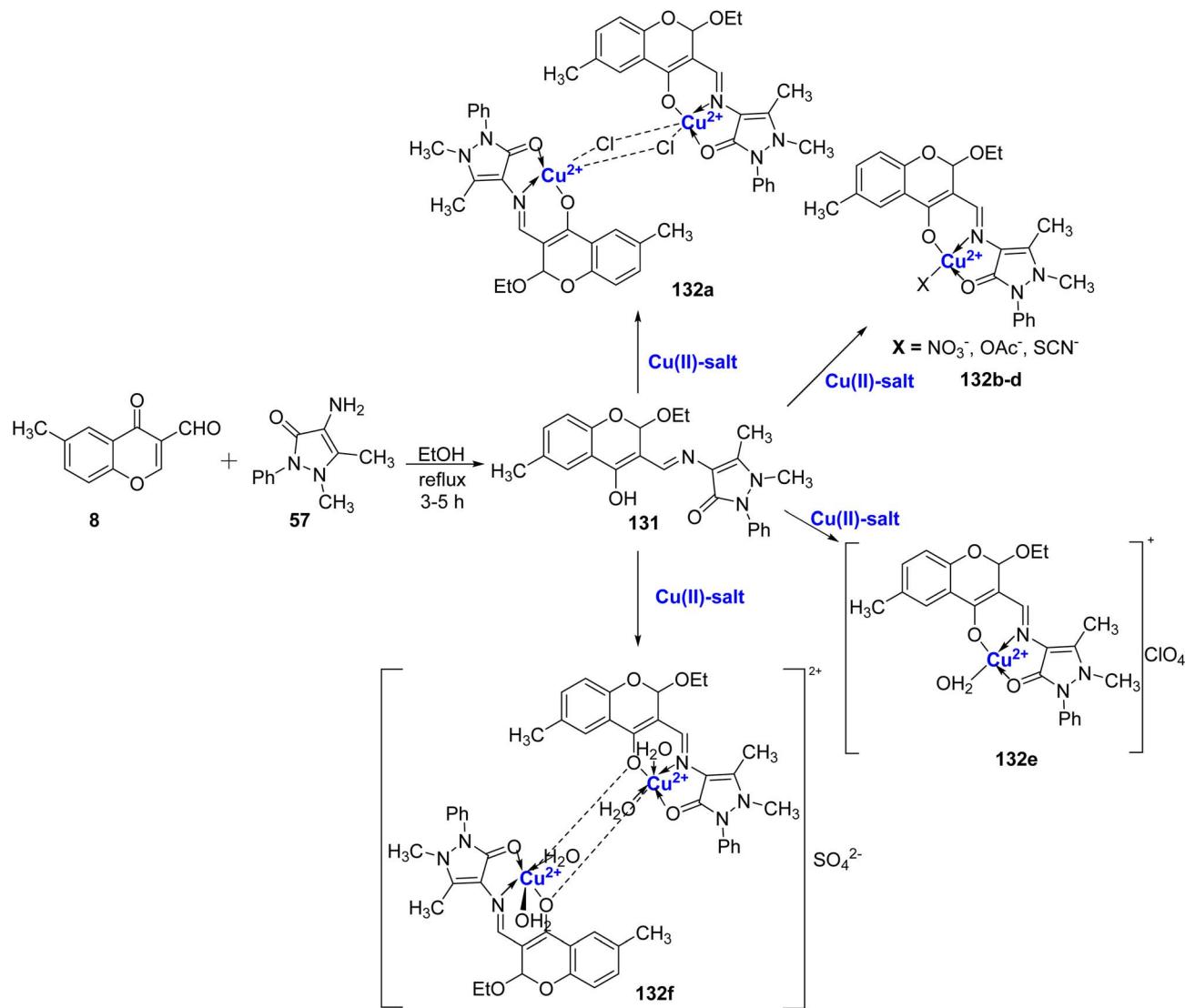
Anitha *et al.*¹²⁰ designed 1:2 complex of Cu(II) ions in conjunction with azo Schiff base **136**. Conductance data indicate that the complex are generally non-electrolytic in nature. The synthesis of Schiff base **136** involved the condensation of *p*-phenylenediamine **133**, **135**, and **2**. Subsequently, this Schiff base was treated with copper(II) chloride, resulting in the

desired metal complex **137** (Scheme 41). The Schiff bases and their corresponding metal complex were also subjected to antibacterial and antifungal studies. The Cu(II) complexes showed promising antibacterial activity towards bacteria, *S. aureus*, *E. coli*, *S. enterica typhi*, and *B. subtilis*.

Mendu *et al.*¹²¹ explored the synthesis of the Schiff base known as 4-chloro-2-((4-oxo-4H-chromen-3-yl)methyleneamino) benzoic acid **139**, along with a Cu(II) complex (Scheme 42). The Schiff base ligand **139** obtained by the reaction of **2** and 4-chloroanthranilic acid **138**. The synthesized ligand **139** and its respective complex were also subjected to antimicrobial assessments against different bacteria using well disc and fusion methods. The metal complex exhibited higher potency against microorganisms when compared to the free Schiff base ligand. The Cu(II)-complex function through redox chemistry in cleaving DNA. The structure (octahedral geometry) and characterization were done *via* analytical and spectroscopic measurements. The complexation of the ligand **139** was found to occur from three sites: the nitrogen atom of azomethine, the oxygen atom of ketonic functionality as well as the hydroxyl of the carbonyl functional group.

Bheemarasetti *et al.*¹²² developed a novel Schiff base ligand known as 3-(((1H-1,2,4-triazol-3-yl)imino)methyl)-4H-chromen-4-one (L) **141** through the reaction of **2** and **140** in methanol (Scheme 43). Additionally, they synthesized Cu(II), complex. Comprehensive characterization of both the ligand **141** and the complex **142** was carried out using techniques such as FT-IR, ESR, UV-Vis, SEM, NMR, mass spectrometry, TGA, and X-ray analysis. The shape of complex **142** was found to be square planar. The complex **142** also demonstrated good antiproliferative and anticancer results compared to the remaining compounds. They could be used as a promising antitumor agent. Additionally, DNA binding studies revealed that these compounds interact with CT-DNA *via* an intercalative mode.

Scheme 39 Synthesis of ligand **83a** and its Cu(II)-complex **130**.



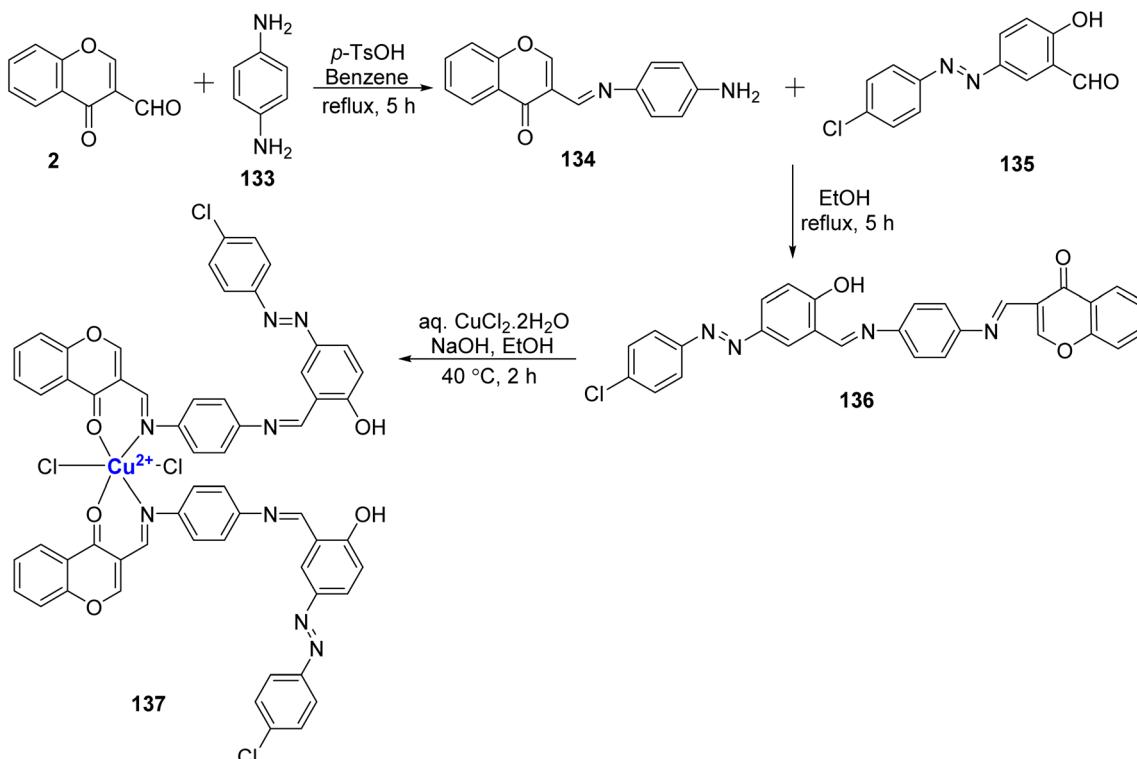
Scheme 40 Synthesis of ligand 131 and their Cu(II)-complexes 132a-f.

In 2011, Arjmand *et al.*¹²³ conducted a study in which they synthesized a Cu(II) complex based on chromone and investigated its interactions with DNA molecules, as well as its DNA cleavage properties. The ligand **144** was prepared by combining **124** and **143**, (Scheme 44). The Cu(II)-complex **145** was then synthesized with a metal-to-ligand ratio of 1:2. Significant absorption enhancements were observed, confirming the binding of complex **145** with the DNA backbone, particularly in the 5'-GMP region. Consistent findings supporting this interaction were obtained through other experimental measurements. Gel electrophoretic mobility assay was employed to assess the complex's artificial nuclease activity, revealing that complex **145** effectively cleaves plasmid pBR322 DNA, transforming it from form I to form II and ultimately generating linearized form III as the complex concentration increases.

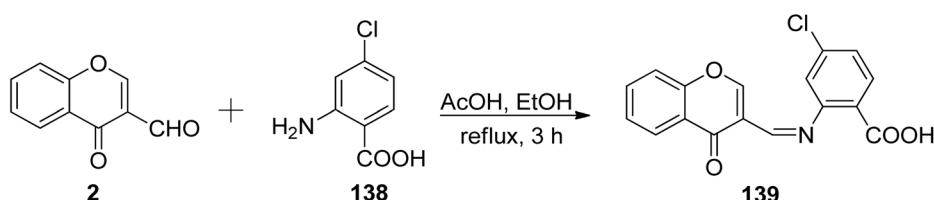
In a study conducted by Alturiqi *et al.*,¹²⁴ they synthesized new complex of Cu(II), **148** by combining **146** with **2**, as outlined

in Scheme 45. The complex was then evaluated for their potential as agents against microbial infections and tumors. The researchers assessed the antimicrobial properties of complex against a range of bacteria, as well as fungi. The hexacoordinated octahedral copper complex **148** demonstrated the most robust antibacterial activity, with a zone of inhibition ranging from 13 to 20 mm. This enhanced activity was attributed to the larger atomic radius and electronegativity of the Cu(II) ion, which reduced the effective positive charge on the complex and facilitated its interaction with cellular membranes. Additionally, the complex exhibited outstanding scavenging activity against the DPPH radical, with the copper complex showing an IC₅₀ value of 83.28. However, when it came to their impact on human tumor cell lines, including A427, LCLC-103H, and SISO, none of the tested compounds displayed toxicity. Remarkably, the Cu(II) complex displayed moderate activity, with IC₅₀ values of 18.365 μ M.

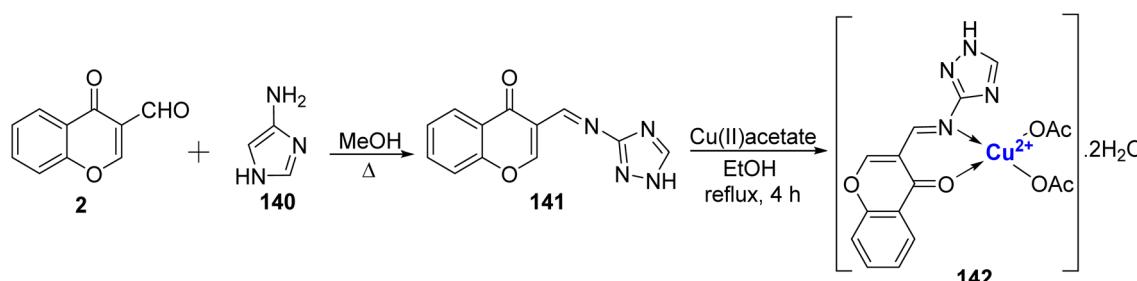




Scheme 41 Synthesis of ligand 136 and its complex with Cu(II) ion 137.



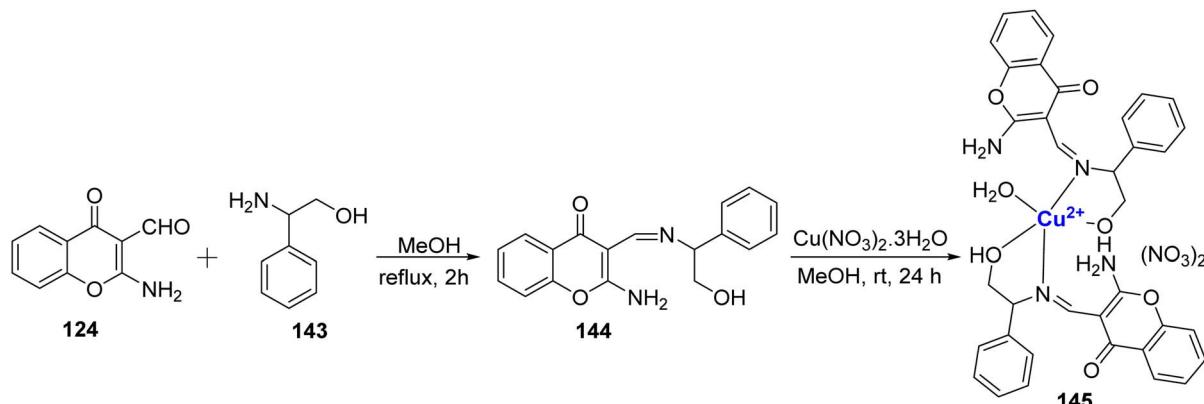
Scheme 42 Synthesis of Schiff base ligand 139.



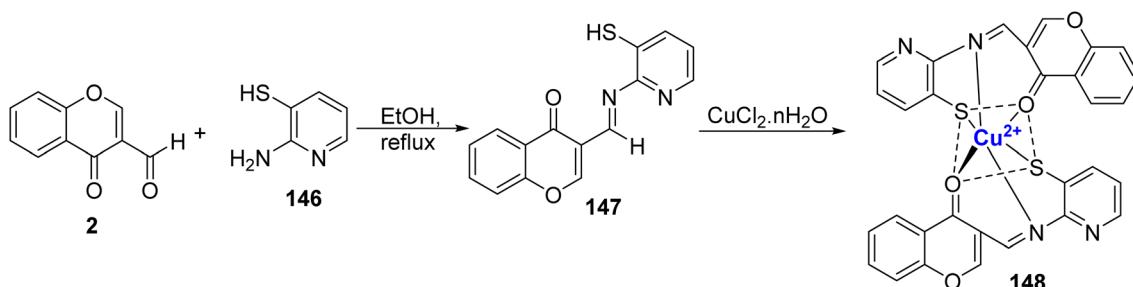
Scheme 43 Synthesis of ligand 141 and Cu(II)-ion complex 142.

Singh *et al.*¹²⁵ conducted a study where they synthesized a metal complex involving Cu(II), using Schiff base ligand 150 (Scheme 46). Through UV absorption studies, they identified characteristic peaks for the Cu(II) complexes, designated as 151, in combination with the chromone ligand. These peaks appeared at 274 nm, 278 nm, and 279 nm, respectively,

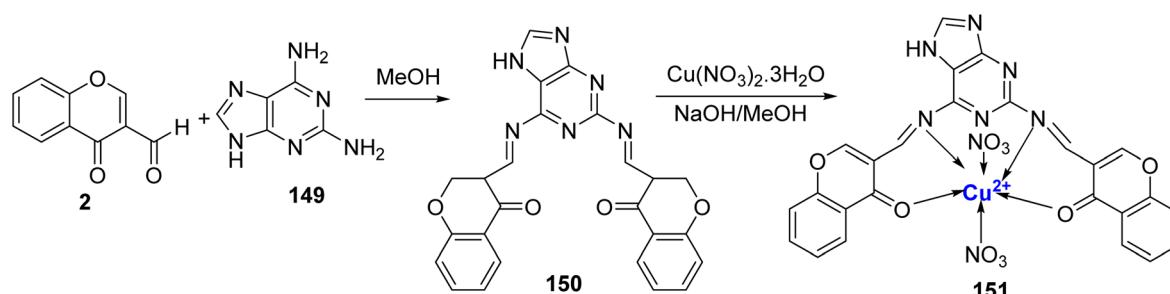
representing $\pi-\pi^*$ transitions, and also showed charge transfer transitions in the range of 465–481 nm. The geometric structure of the copper complexes was determined through DFT studies, revealing a distorted octahedral geometry. To investigate how these metal complexes interacted with DNA, the researchers employed various techniques, including UV-vis absorption



Scheme 44 Synthesis of ligand 144 and its Cu(II)-complex 145.



Scheme 45 Synthesis of ligand 147 and its Cu(II)-complex 148.



Scheme 46 Synthesis of ligand 150 and its Cu(II)-complex 151.

spectroscopy, fluorescence titration, and viscosity measurements. Changes in spectral position and intensity were observed to assess the interaction between the metal complexes and DNA. Notably, when the metal complexes bound to DNA, a hypochromic effect was observed, accompanied by a bathochromic shift. The copper complexes associated with chromone displayed absorption peaks at 299, 310, and 356 nm due to $\pi-\pi^*$ transitions. These complexes were found to bind to DNA through non-covalent interactions or by causing the uncoiling of the DNA double helix. As a result, the complexes adopted an intercalation mode of binding, effectively stacking among the aromatic chromophores and DNA base pairs. Fluorescence quenching studies revealed a decrease in emission intensity, signifying the interaction of the complexes with DNA. Further confirmation of the intercalation mode of interaction

was provided by replacing EB with the complex molecules in EB-bound DNA. Hydrodynamic viscosity measurements furnished additional evidence by showing that the complexes inserted themselves between the DNA base pairs. This action caused the separation of the DNA double helix, resulting in an increase in DNA length and viscosity.

4. Biomimetic catalytic activity of chromone-based Schiff bases

Copper has a vital role in different enzymatic activities.¹²⁶ The reason for the significant interest in Schiff base metal complexes is not only because of their high stability and biological activity^{127,128} but also because of their unique features of



functioning as catalysts.¹²⁹ Within the array of bio-inspired structures capable of emulating catecholase activity, copper complexes of Schiff bases emerge as a particularly promising candidate. While the active site of the catecholase enzyme typically features a hydroxo-bridged dicopper(II) center, it is widely acknowledged that numerous monometallic copper(II) complexes exhibit catecholase activity. In this context, we have presented a study on the synthesis, characterization, and catalytic attributes of certain chromone-based metal complexes of Schiff bases.

In 2016, Beyazit *et al.*¹³⁰ reported a novel tetradentate, unsymmetrical Schiff base ligand **153** and its Cu²⁺ ion complex **154**. The ligand **153** was obtained by reacting a mixture of **56** (prepared by oxidation of visnagin¹³¹) and 2-aminobenzylamine **152** in CHCl₃ for 2 hours under reflux condition (Scheme 47). The complex **154** was obtained as a dark green precipitate by refluxing the ligand **153** and Cu(CH₃COO)₂·H₂O in methanol for 5 hours. The ligand **153** and the complex **154** were confirmed by using characterization techniques such as NMR, mass, elemental analysis, and electronic spectra. The complex **154** complex exhibited catecholase-like biocatalytic activity against the oxidation of 3,5-di-*tert*-butylcatechol to quinone form. This revealed that complex **154** complex had moderate catalytic activity.

In continuation of their previous work (Scheme 47), they synthesized two chromone-based ligands **156a–b** and their corresponding transition metal complexes (Scheme 48).¹³² The complexes **157a–b** thus obtained were screened for their catechol oxidase activity. The study revealed that the presence of substituents on ligands **156a–b** significantly influenced the catalytic activity of the resulting metal complexes **157a–b**. Notably, electron-donating substituents led to an enhancement in catalytic performance. The oxidation of catechol was found to

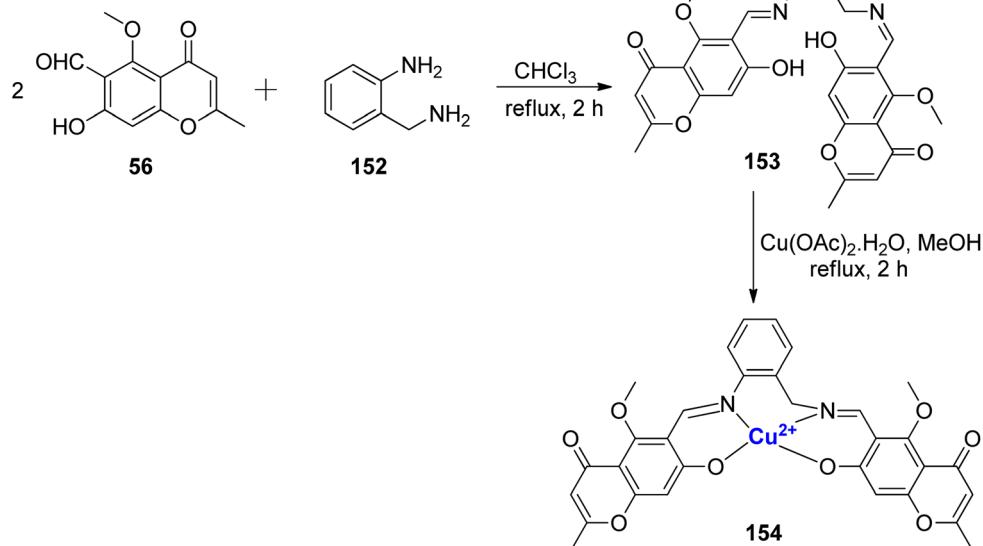
adhere to first-order kinetics, according to the results of the kinetic measurements. All the synthesized compounds **156a–b** and **157a–b** were well characterized *via* thermal/elemental analysis, FT-IR, UV-vis spectroscopy, and NMR.

Beyazit *et al.*¹³³ also utilized visnagin derivative **56** to synthesize Schiff base ligands **160** and **161** by reacting it with the 2,3-diaminonaphthalene **158** and 1,8-diaminonaphthalene **159** (Scheme 49). The synthesized ligands **160** and **161** were treated with copper salts in a mixture of EtOH/CHCl₃ to afford Cu²⁺–L complexes **162** and **163**. The structure of the ligands **160** and **161** as well as the complex **162** and **163** were confirmed *via* various spectroscopic techniques and elemental analysis. The complexes Cu(II) thus prepared were investigated for its catecholase potential and found to display moderate activity against 3,5-DTBC oxidation.

In 2021, Shebl *et al.*¹³⁴ synthesized novel mononuclear Cu(II)–hydrazone complexes **166a–c** with the aid of ligand **165** (Scheme 50). Characterization was done by elemental analysis, IR, TEM, powder XRD, thermal analysis, conductivity, mass, and ESR. Three different complexes of Cu(II) were reported: [Cu(L)(NO₃)]·EtOH **166a**, [Cu(L₂)]·5H₂O **166b**, and [Cu(L)(8-HQ)NO₃]·H₂O **166c**. These complexes **166a–c** also showed phenoxazinone synthase property by oxidizing 2-aminophenol into the 2-aminophenoxazine-3-one. The complexes **166a–c** were found to be active against *Candida albicans*. The ligand **165** also displayed activity against *Candida albicans* but was found to be less than that of the complexes **166a–c**.

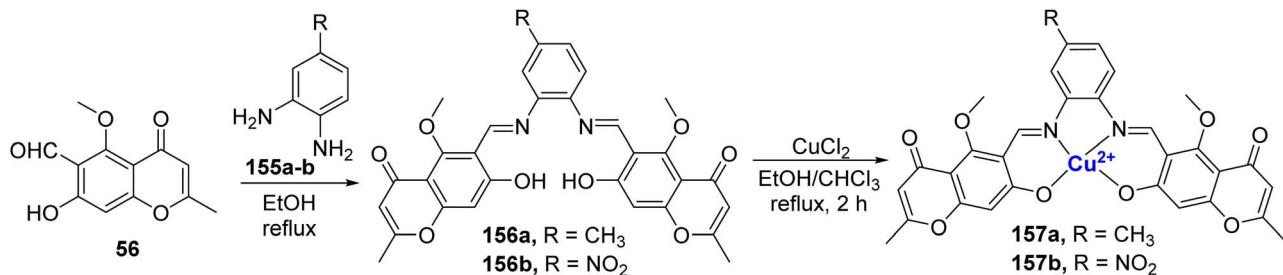
5. Miscellaneous

A tridentate chromone-based ligand **169** was developed by Alaghaz *et al.*¹³⁵ When an equimolar mixture of Girard T (2-ethyl-5-methoxy-6-formyl-7-hydroxy chromone) **167** and

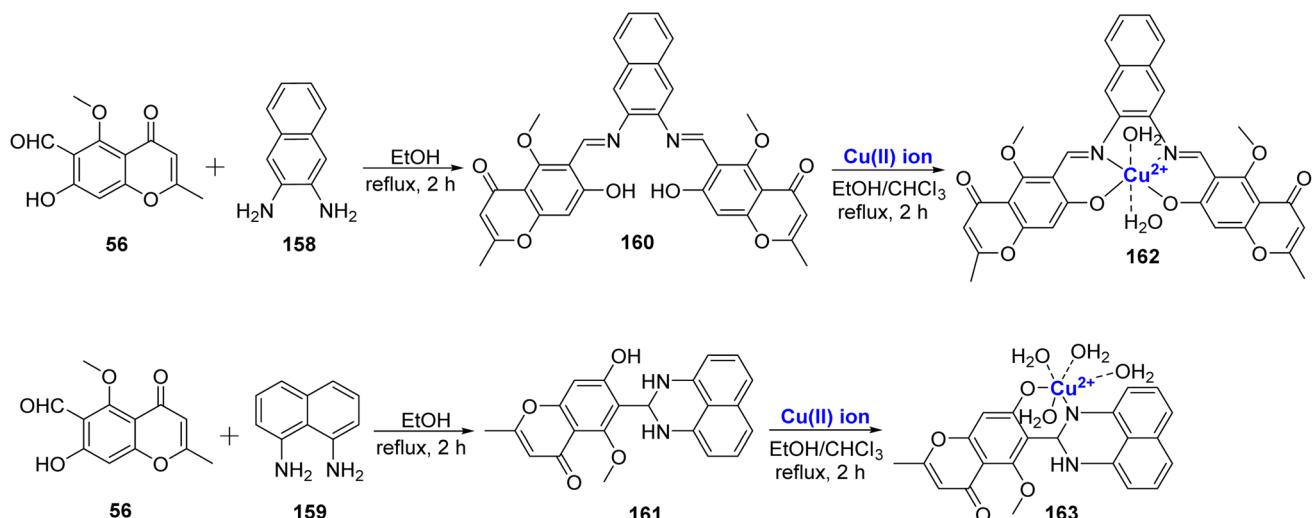


Scheme 47 Synthesis of ligand **153** and its Cu²⁺–L complex **154**.

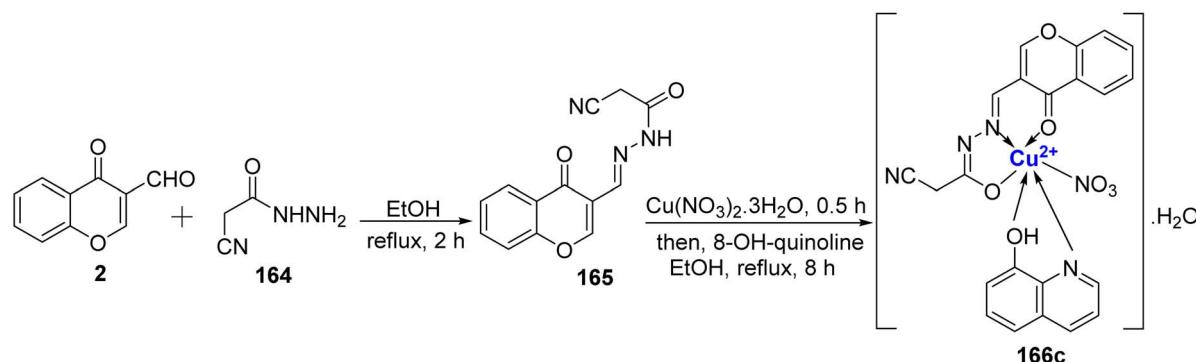




Scheme 48 Synthesis of ligands 156a–b and Cu(II)-metal complexes 157a–b.



Scheme 49 Synthesis of ligands 160–161 and Cu(II) ion complexes 162–163.

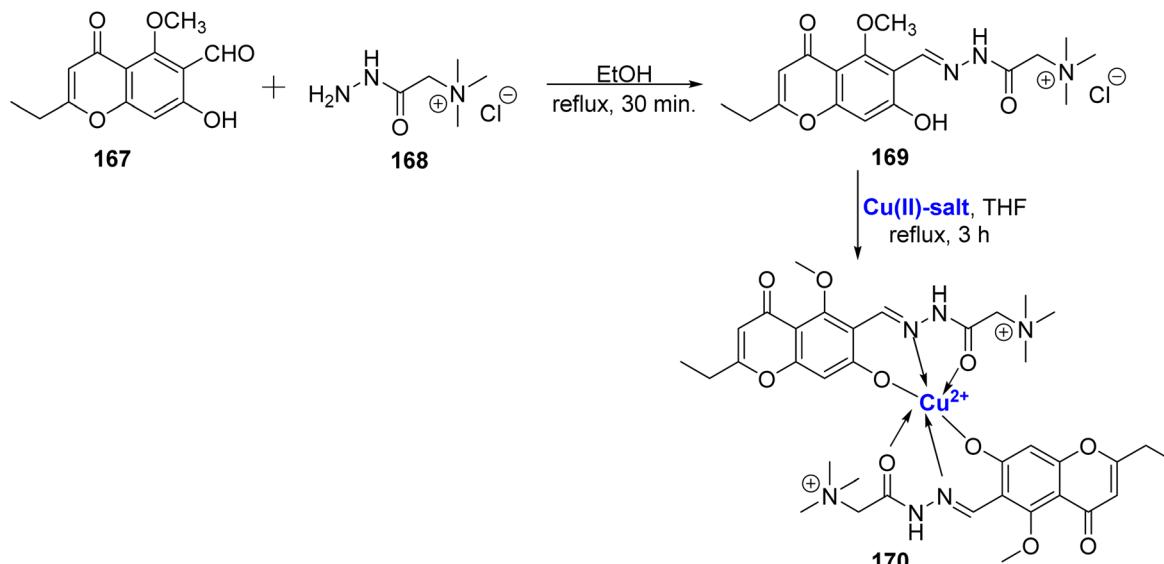


Scheme 50 Synthesis of ligand 165 and Cu(II)-ion complex 166c.

trimethylammoniumacetyl hydrazine chloride 168 was refluxed, it gave chromone-based Schiff base ligand 169 (Scheme 51). This tridentate ligand which was confirmed *via* an elemental analysis as well as spectral data, was further utilized by the authors to

synthesize its Cu(II)-complex 170 which was found to be non-electrolytic in nature. DFT calculations affirmed the octahedral geometry of complex 170.





Scheme 51 Synthesis of chromone Girard T ligand 169 and its Cu(II)-complex 170.

6. Conclusion and future perspectives

Schiff bases hold a significant position within the realm of organic compounds, owing to their pharmacological attributes and their capacity to establish stable bonds with transition metal ions, notably copper. Copper complexes featuring Schiff bases have garnered notable attention in recent times because of their diverse applications in biological processes and their role in the advancement of novel therapeutic drugs. Consequently, chromone-based Schiff base complexes with copper constitute a noteworthy category of compounds. It has been revealed that Cu(II) complexes formed with chromone Schiff bases exhibit potent *in vivo* antitumor properties by impeding DNA replication in tumor cells, thereby inhibiting tumor growth. These compounds have also gained prominence for their utility as not only anticancer agents but also as antimicrobial agents, antioxidants, and more.

Nevertheless, there remain certain challenges, particularly concerning water solubility and selectivity issues between cancer and normal cells. These challenges can potentially be addressed by incorporating them into nano-level formulations. Furthermore, Schiff base ligands have demonstrated their potential as excellent probes for detecting cupric ions. Chromone-based Schiff base ligands, in particular, exhibit high specificity and selectivity in detecting Cu²⁺ ions, with potential applications in environmental and biological monitoring, as well as various pharmaceutical applications.

Abbreviations

AAS Atomic absorption spectroscopy

HIV	Human immunodeficiency virus
NMR	Nuclear magnetic resonance
CS	Colorimetric sensor
U. S.	United States
FTIR	Fourier transform infra-red
UV-Vis	Ultraviolet-visible
EPA	Environmental Protection Agency
TGA	Thermal Gravimetric Analysis
EPR	Electron paramagnetic resonance
IR	Infra-red
EI Mass	Electron ionization mass
HRMS	High resolution mass spectrometry
TGA	Thermal gravimetric analysis
ICP	Inductively coupled plasma
5-BDTC	5-Benzylidithiocarbazate
DFT	Density functional theory
PNT	<i>Para</i> -Nitrotoluene
ESR	Electron spin resonance
FESEM	Field emission scanning electron microscopy
DTCB	<i>N</i> -methyl-5-benzylidithiocarbazate
TEM	Transmission electron microscopy
WHO	World Health Organization
SEM	Scanning electron microscopy
DTBC	Di- <i>tert</i> -butylcatechol
CTAB	Cetrimonium bromide
TG-	Thermogravimetry and differential scanning colorimetry
DSC	Differential scanning colorimetry
TD-	Time dependent density-functional theory
DFT	Density functional theory
EPDM	Ethylene propylenediene monomer
XRD	X-ray diffraction
DAN	Diaminonaphthalene
IC ₅₀	Half-maximal inhibitory concentration
DNA	Deoxyribonucleic acid
GMP	Guanosine monophosphate



GI ₅₀	Growth inhibition by 50%
CT-DNA	Circulating-tumor DNA

Conflicts of interest

There are no conflicts to declare.

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