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Sensitive Cr³⁺ sensor based on novel poly(luminol-co-1,8-diaminonaphthalene)/CeO₂/MWCNTs nanocomposites

 Salsabeel Al-Sodies,^{ab} Abdullah M. Asiri,^{ac} M. M. Alam,^{id} Khalid A. Alamry,^{id} Mahmoud A. Hussein^{id}*^{ae} and Mohammed M. Rahman^{id}*^{ac}

In this study, poly(luminol-co-1,8-diaminonaphthalene) (PLim-DAN) was synthesized and subsequently modified with MWCNTs and CeO₂ NPs. The synthesized nanocomposites were analyzed using IR, SEM, TEM, and XRD. Furthermore, a comprehensive set of thermal behavior measurements were taken using TGA/DTG analysis. Next, the electroactivity of the developed nanocomposites was tested as an electrochemical sensor to measure the concentration of Cr³⁺ ions in phosphate buffers. The GCE adapted with the PLim-DAN/CeO₂/CNTs-10% nanocomposite (NC) exhibited the highest current response among the other compositions and copolymers. The fabricated nanocomposite sensor showed high sensitivity, with a value of 19.78 μA μM⁻¹ cm⁻², and a low detection limit of 4.80 ± 0.24 pM. The analytical performance was evaluated by plotting a current calibration curve *versus* the concentration of Cr³⁺ ions. It was found to be linear ($R^2 = 0.9908$) over the range of 0.1 nM to 0.1 mM, identified as the linear dynamic range (LDR). This electrochemical sensor demonstrated that it could be a useful tool for environmental monitoring by accurately detecting and measuring carcinogenic Cr³⁺ ions in real-world samples.

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Introduction

Considering the widespread use of polymers in various applications, scientists have focused their attention on the polymer activity of organic compounds in multiple fields. Given the need for a comprehensive understanding of these polymers, luminol polymeric moiety systems have not been studied extensively.¹

Luminol and its derivatives have been employed in various analytical chemistry and biotechnology applications as efficient electrochemiluminescence conducting polymers.^{2,3} The primary application of electrochemiluminescence in crime scene investigations has resulted in the creation of discoveries of polymeric luminol, which enhances the efficiency, sensitivity, and quantum yield of the process.⁴ An amino group in luminol enhances chemiluminescence intensity by acting as an electron-donating group in the system. This positively impacts

the electrochemical properties of poly(luminol) and its potential use in electrochemical and biosensor materials.⁵⁻¹⁰

From the same family, a new type of multifunctional polymer material, poly(di-aminonaphthalene), can be synthesized from aromatic diamines, such as 1,5-, 1,8-, or 2,3-diaminonaphthalene, through electrochemical or chemical oxidative polymerization with properties similar to those of polyaniline and polypyrrole.¹¹ The substance 1,8-diaminonaphthalene displays properties such as electroconductivity, electrocatalysis, electroactivity, permselectivity, electrochromism, and other properties that stem from the chemical reactivity of the functional amino groups on its macromolecular structure.¹²⁻¹⁶

Furthermore, one of the most noteworthy and rapidly advancing fields in materials science and analytical chemistry is the design of novel electrochemical sensing materials that display exceptional electrocatalytic properties, enduring stability, consistent repetition, high reliability, and enhanced sensitivity and selectivity.¹⁷⁻²⁰ Carbon nanotubes (CNTs) are highly valued for developing sensors and biosensors because of their remarkable chemical resistance, large surface area, high tensile strength, excellent electrical conductivity, and distinct one-dimensional structure facilitating rapid electron transfer.²¹ CNTs possess a broad spectrum of potential applications, including electronics, polymer composites, energy storage, catalysis, gas storage, and sensors.²² Multi-walled carbon nanotubes (MWCNTs) exhibit good mechanical strength and enhanced surface activity with a high specific surface area,

^aChemistry Department, Faculty of Science, King Abdulaziz University, P. O. Box 80203, Jeddah, 21589, Saudi Arabia. E-mail: mahussein74@yahoo.com; maabdo@kau.edu.sa; mmrahman@kau.edu.sa; mmrahmanh@gmail.com

^bDepartment of Chemistry, Faculty of Science, Taibah University, Al-Madinah Al-Munawarah 30002, Saudi Arabia

^cCenter of Excellence for Advanced Materials Research (CEAMR), King Abdulaziz University, Jeddah 21589, Saudi Arabia

^dDepartment of Chemical Engineering, Z. H. Sikder University of Science and Technology (ZHSUST), Shariatpur 8024, Bangladesh

^eChemistry Department, Faculty of Science, Assiut University, Assiut, 71516 Egypt



making them ideal for thermally stable materials, biological applications, water filtration, structural materials, and sensors.^{23,24} Carbon nanotubes (SWCNTs and MWCNTs) have been extensively used as sensing materials to fabricate various nanocomposites that have been successfully employed to detect a wide range of bioanalytes, including uric acid, ascorbic acid, dopamine,²⁵ styrene, epinephrine,²⁶ glutathione, glucose,^{27,28} and toxic metal analytes, such as Ga³⁺, Fe³⁺, Hg²⁺, Pb²⁺, and Cu²⁺.^{29–31}

To further improve the sensing performance, nanomaterials have been fabricated and loaded on the surface of designed polymers. Cerium oxide (CeO₂) has garnered attention among the various nanoparticles because of its impressive properties, including proficient photocatalytic activity, high surface area, oxygen ion conductivity, high chemical stability, high specific capacitance, and non-toxicity. CeO₂ is a rare earth metal oxide with a broadband gap (3.4 eV) and a cubic fluorite structure. Each Ce⁴⁺ ion was surrounded by eight O²⁻ ions in a (fcc) arrangement. Each O²⁻ ion was tetrahedrally surrounded by four Ce⁴⁺ ions. Consequently, CeO₂ has been used in various applications, including solid state supercapacitors devices, solar cells, photocatalysis, and sensors.^{32–38}

Heavy metals have been targeted for sensing applications due to their widespread presence as one of the most important pollutants found in water bodies and are highly toxic to ecosystems.^{29–31} Trivalent chromium (Cr³⁺) is critical in metabolizing lipids, nucleic acids, and proteins in biochemical processes, making it an essential element in trace concentrations in humans and animals.^{39,40} However, exposure to higher concentrations of Cr³⁺ ions and a deficiency of Cr³⁺ can result in diseases associated with sugar metabolism disorders, such as diabetes, cataracts, cardiovascular disease, uremia, and blindness.^{41,42} Moreover, Cr³⁺ ions can impair the quality and quantity of edible agricultural products and animals.⁴³ Therefore, Cr³⁺ is considered a significant environmental pollutant and health hazard, and there is an urgent need for an easy and reliable method for detecting Cr³⁺ ions in aqueous media to ensure a sustainable environment. To this end, several efficient and sophisticated analytical methods, such as HPLC, ICP-AES, DPP, and X-ray fluorescence, have been employed to detect trace Cr³⁺ ions. However, these techniques are expensive, time consuming, and unsuitable for frequent analysis. Thus, there is a pressing need to develop a convenient probe that can rapidly and selectively detect Cr³⁺ ions using a sensor designed explicitly for cations.^{44,45}

The combination of CeO₂ NPs and MWCNTs in sensors has been reported in several studies.^{29,46–51} N. Dogra *et al.* developed a chemiresistive sensor for detecting ammonia vapor at room temperature using CeO₂/MWCNTs composites. The resistance of the sensor increased in the presence of ammonia vapor, and the fabricated sensor had a response time of 35 s, a relatively short recovery time of approximately 100 s, high sensitivity to ammonia, and stable and reproducible characteristics over a broad range of humidity levels.⁴⁶ Furthermore, rice-like CeO₂/MWCNTs nanocomposite were synthesized *via* a simple hydrothermal method and were investigated by S. Shanavas *et al.* for detecting ammonia and ethanol gases at

concentrations ranging from 0 to 500 ppm by employing a fiber-optic clad modification technique. The findings suggest that CeO₂/MWCNT displays a high sensitivity of 78 counts/ppm, rapid response time of 17 s, and recovery time of 9 s in response to ethanol gas.⁴⁷ The use of CeO₂/MWCNTs in the gas sensor was expanded to include stretchable sensors for the detection of NO₂. The fabricated sensor was formed by depositing the MWCNTs/CeO₂ composites onto silicon rubber and a jelly-based substrate to develop a flexible and biodegradable sensor. The jelly-based device exhibited faster response and recovery times (22.9/345.2 s) in compared to the stretchable device.⁴⁸ Furthermore, the two nano-structured CeO₂/MWCNTs were successfully explored and constructed as an electrochemical sensor for neonicotinoid insecticides (nitenpyram) with a low detection limit of 0.72 μM,⁴⁹ acetaldehyde (LOD = 7.4 nM),⁵⁰ dopamine (LOD = 0.03 μM),⁵¹ and Hg²⁺, Pb²⁺, and Cu²⁺ heavy metals with LOD of 1.98, 1.10 and 3.53 μg L⁻¹, respectively.²⁹

As the importance of the above individual components emerged, this study investigates a sensor probe to detect Cr³⁺ ions through electrochemical means using a novel copolymer of luminol and 1,8-diaminonaphthalene. To the best of our knowledge, this is the first reported study of constructed CeO₂/MWCNTs as a sensor for Cr³⁺ ions, as well as the polymeric luminol and 1,8-diaminonaphthalene. The copolymers modified with MWCNTs and CeO₂ NPs enhanced the electrochemical activity, producing nanocomposites of PLim-DAN/CeO₂/CNTs. The performance of the fabricated Cr³⁺-ion sensor with GCE was evaluated and was extremely sensitive and selective for Cr³⁺ ions. Finally, a newly developed Cr³⁺ ion sensor was applied to environmental samples to detect Cr³⁺ ions. Developing heavy metal ion sensors can provide a unique and reliable approach to environmental applications.

Experimental

Chemicals and reagents

Luminol, 1,8-diaminonaphthalene, ammonium persulfate (NH₄)₂S₂O₈, and dimethyl sulfoxide 98% purity (DMSO) were purchased from Aldrich. Multi-walled carbon nanotubes (MWCNTs) and cerium oxide nanoparticles (CeO₂ NPs: 25–50 nm) were purchased from Nano Tech Co. Ltd. Egypt. Analytically graded inorganic salts of As³⁺, Cd²⁺, Co²⁺, Cr³⁺, Ga³⁺, Hg²⁺, Pb²⁺, Sb³⁺, and Sn²⁺ were obtained from the supplier, as required for this study. All chemicals were used as received without further purification.

Instrumentation

Fourier transform infrared (FT-IR) spectra were recorded using a PerkinElmer Spectrum 100 FT-IR device in the 4000–500 cm⁻¹ range. The morphologies and elemental distributions of the polymers were examined using scanning electron microscopy (SEM, TESCAN VEGA 3, Czech Republic). Samples were mounted on aluminum microscopy stubs using carbon tape and then coated with gold (Au) for 120 s using a Quorum Techniques Ltd. sputter coater (Q150t, UK). Transmission electron microscopy



(TEM) Thermo Fisher Scientific, multi-purpose, Talos F200i S/TEM was used to investigate the interactions between the polymers and nanomaterials with high-resolution imaging and analysis applications operating at 200 kV. X-ray diffraction (XRD) for the designated materials data was collected using a Phillips X-ray unit (Phillips's Generator PW-1710) diffractometer with a Cu K α irradiation source. The $2\theta = 5\text{--}80^\circ$ range was scanned at a 1°min^{-1} rate. Thermogravimetric analysis (TGA) was performed on a Shimadzu TGA 50 instrument at a heating rate of $10^\circ \text{C min}^{-1}$ in air. A Keithley electrometer was used to construct the electrochemical cell for the electrochemical (I-V) analysis.

Synthesis of poly(luminol-co-1,8-diaminonaphthalene) copolymers (PLim-DAN)

The copolymer PLim-DAN was prepared in an equal ratio (1 : 1) using the original oxidative polymerization process for poly(luminol).⁵² Equipped with an N₂ atmosphere, a mixture of H₂O : DMSO (1 : 9) was added to luminol (1 mmol) and 1,8-diaminonaphthalene (1 mmol) in a two-neck round flask. The reaction mixture was stirred at room temperature for 30 min to ensure monomer solubility. A solution of ammonium persulfate (NH₄)₂S₂O₈ (3 mmol) was added to initiate the polymerization process. The reaction mixture was stirred for 24 h at room temperature, and the formed precipitate was collected by filtration, washed thoroughly with water multiple times, and dried at 70 °C for 48 h.

Preparation of nanocomposite poly(luminol-co-1,8-diaminonaphthalene) copolymers with CeO₂ NPs/MWCNTs

A new series of poly(luminol-co-1,8-diaminonaphthalene)/CeO₂ NPs/MWCNTs was prepared *via* an *in situ* oxidative polymerization method using a fixed percentage (5%) of CeO₂ NPs and different loadings (1, 3, 5, and 10%) of MWCNTs concerning unmodified copolymers. For each formulation, a mixture of CeO₂ NPs (5%) and different ratios of MWCNTs in H₂O : DMSO (1 : 9) was sonicated for 1 h, followed by the addition of equimolar amounts of the two monomers (luminol and 1,8-diaminonaphthalene), and sonication continued for 30 min. The reaction mixture was stirred for 30 min. A solution of ammonium persulfate (NH₄)₂S₂O₈ (3 mmol) was added to initiate the polymerization process for each formulation. The workup of the formed nanocomposites was similar to that of the unmodified copolymers, and the obtained composites were labeled PLim-DAN/CeO₂/CNTs-1%, PLim-DAN/CeO₂/CNTs-3%, PLim-DAN/CeO₂/CNTs-5%, and PLim-DAN/CeO₂/CNTs-10%.

Modification of GCE with active poly(luminol-co-1,8-diaminonaphthalene)/CeO₂ NPs/MWCNTs

Hare, GCE, and PLim-DAN/CeO₂/CNTs nanocomposites were used to modify the desired electrochemical sensor using the I-V method. A slurry of PLim-DAN/CeO₂/CNTs NCs was prepared in ethanol as a thin and uniform layer, which was subsequently deposited onto a GCE with a surface area of 0.0316 cm². After drying the slurry, a drop of Nafion (5% Nafion suspension in ethanol) was added to the modified GCE to obtain the desired

stability. The GCE was then thoroughly dried in an oven at 35 °C for an adequate amount of time. An electrochemical cell was assembled using a Keithley electrometer, with PLim-DAN/CeO₂/CNTs NCs/binder/GCE as the working electrode and a simple Pt wire as the counter electrode. A chromium(III) ion solution was prepared and used as the target analyte. A calibration curve was plotted from the linear relationship between the current and concentration of Cr³⁺ ions, and the analytical performance of the sensor, such as the sensitivity and detection limit (DL), was estimated from the slope of the calibration curve. The linear dynamic range (LDR) was determined by considering the maximum linearity (R^2) of the calibration curve. During the electrochemical investigation, the phosphate buffer solution in the detection beaker was kept constant at 10.0 mL throughout the experiment. An electrochemical sensor using a Keithley electrometer is a simple two-electrode (working and counter) system.

Results and discussion

Chemistry

In the present work, the polymerization process consisted of two main steps: the oxidative polymerization of the genuine copolymers and the preparation of loaded nanomaterials using a typical procedure with minor modifications to distribute the nanoparticles (Fig. 1). The electrochemical performances of the designed copolymer and its composites as sensitive sensors for Cr³⁺ ions were examined. The polymerization occurred in a water and DMSO media mixture with ammonium persulfate (NH₄)₂S₂O₈ as an oxidizing agent.^{12,52} In the first approach, luminol and 1,8-diaminonaphthalene monomers were converted to polymeric form to design a novel copolymer, PLim-DAN, in an equimolar ratio under an inert atmosphere. The following process involved the modification of the pure copolymers with several ratios using a fixed percentage of cerium oxide nanoparticles (5%) (CeO₂ NPs) and different loadings of

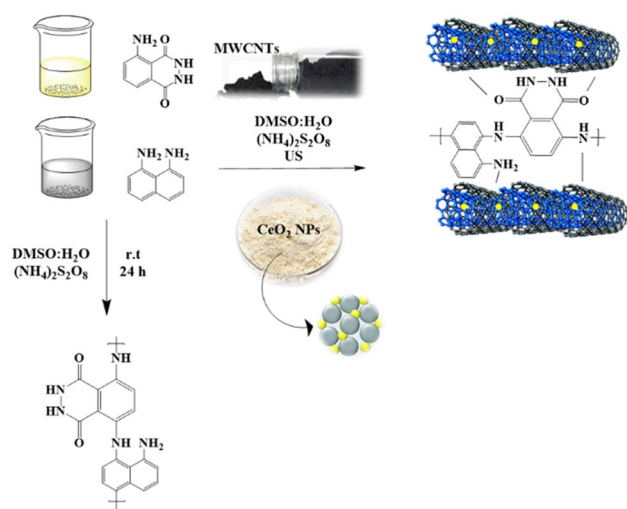


Fig. 1 Systematic illustration of PLim-DAN and PLim-DAN nanocomposite synthesis and fabrication.



MWCNTs (1%, 3%, 5%, and 10%). The new series of composites PLim-DAN/CeO₂/CNTs-1%, PLim-DAN/CeO₂/CNTs-3%, PLim-DAN/CeO₂/CNTs-5%, and PLim-DAN/CeO₂/CNTs-10% were produced through *in situ* oxidative polymerization by applying the same principle of pure copolymers with the introduction of the ultrasound technique to ensure complete distribution of the loaded nanoparticles.⁵³ The designed copolymers were characterized using FT-IR spectroscopy, and their morphologies were studied using SEM, XRD, and TEM. At the same time, thermal behavior was investigated using TGA and derivative thermogravimetry (DTG).

FT-IR analysis

The innovatively synthesized copolymer PLim-DAN and its nanocomposites were investigated using FT-IR analysis. Fig. 2 shows the FT-IR spectra of the two monomers, luminol and 1,8-diaminonaphthalene, compared with the newly designed copolymer PLim-DAN. In the luminol spectra (Fig. 2a), NH₂ appeared as two bands at 3471 and 3418 cm⁻¹, while the NH of the amide group was observed at 3322 cm⁻¹. The spectrum of 1,8-diaminonaphthalene (Fig. 2b) shows bands at 3453 and 3325 cm⁻¹ associated with the amino group. For the pure copolymer (Fig. 2c), the spectrum demonstrates the absence of NH₂ bands overlapping with the NH peak, confirming the formation of a polymeric bond.^{12,52} The C-H symmetric aromatic stretching band appeared in monomers and the copolymer in the 3004–3083 cm⁻¹ range. The amide group C=O band was present in the luminol spectrum (Fig. 2a) at 1662 cm⁻¹ and in the copolymer at 1656 cm⁻¹ (Fig. 2c). The bands at 1593 and 1586 cm⁻¹ in monomers and at 1597 cm⁻¹ in the copolymer represent symmetric and asymmetric aromatic ring stretching of (C=C) overlapped with the vibrations bending of (N-H) amide group, while the peaks located between

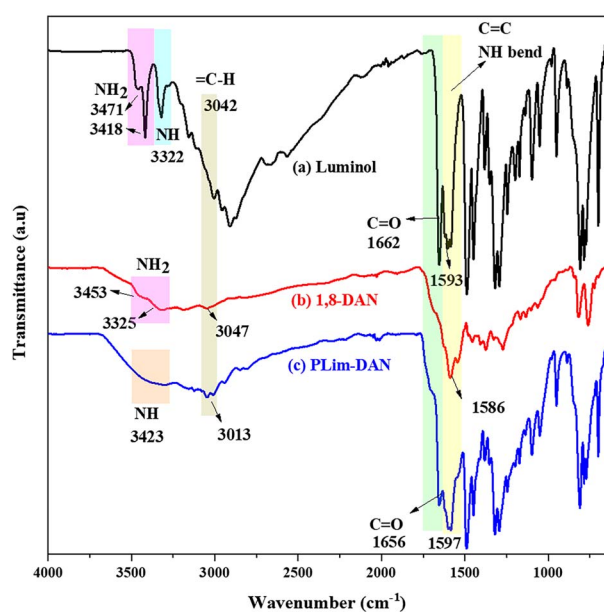


Fig. 2 FT-IR spectra for (a) luminol, (b) 1,8-diaminonaphthalene, and (c) copolymer PLim-DAN.

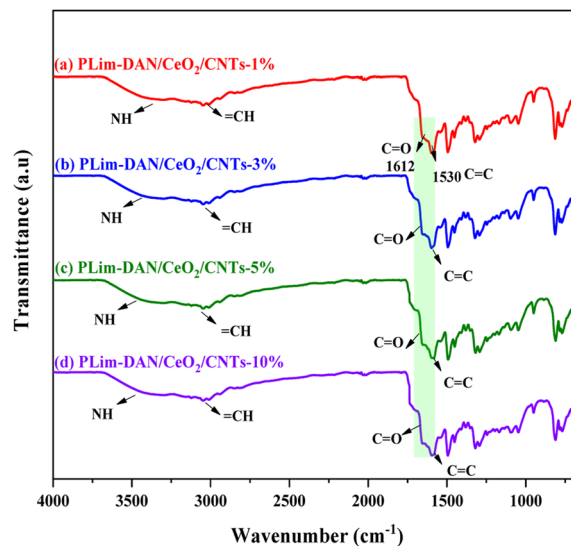


Fig. 3 FT-IR spectra for (a) PLim-DAN/CeO₂/CNTs-1%, (b) PLim-DAN/CeO₂/CNTs-3%, (c) PLim-DAN/CeO₂/CNTs-5% and (d) PLim-DAN/CeO₂/CNTs-10% nanocomposites.

809–818 cm⁻¹ in (Fig. 2a–c) may be related to aromatic C–H bending. The fingerprint area bands of the copolymer spectrum combined both bands in luminol and 1,8-diaminonaphthalene with a broad shape.^{54,55}

Fig. 3 presents the FT-IR spectra of the nanocomposites PLim-DAN/CeO₂/CNTs-1%, PLim-DAN/CeO₂/CNTs-3%, PLim-DAN/CeO₂/CNTs-5%, and PLim-DAN/CeO₂/CNTs-10%. In all spectra (Fig. 3a–d), the band at 1612 cm⁻¹ was assigned to the C=C stretching vibration of the MWCNTs. In contrast, the broad peak at 1530 cm⁻¹ was attributed to the C–C plane vibrations of the graphitic walls of the MWCNTs.⁵³ The peak at 500 cm⁻¹ might be correlated with the Ce–O stretching vibration. The major bands of the copolymer (Fig. 2c) and those related to the MWCNTs appeared in the spectra of the other nanocomposites (Fig. 3a–d). The FT-IR characterization confirmed the successful formation of the copolymer and its nanocomposites. Nevertheless, compared to the unfabricated copolymer, the nanocomposite spectra show similar peaks regions only with deformation in presentation due to the coating effect of MWCNTs and CeO₂ NPs, impacting the vibrational mode.⁵⁶

Morphology analysis

Systematic analysis using SEM was conducted at different magnifications between 5k \times to 25k \times and scale bar from 5 μ m to 1 μ m, to elucidate further the surface morphologies of the newly synthesized copolymer and its nanocomposites. Fig. 4 presents the SEM images of the bare copolymer PLim-DAN (Fig. 4a–c) and PLim-DAN/CeO₂/CNTs-10% nanocomposite (Fig. 4d–f). The pure copolymer matrix (Fig. 4a–c) displays an irregular spherical morphology encompassing large and small particles with no discernible dots or spots on the surface. Some spherical grains aggregated in certain areas, which became clearer at high magnification (Fig. 4c). On the other hand, Fig. 4d–f of PLim-DAN/CeO₂/CNTs-10% displays scattered bright spots as small



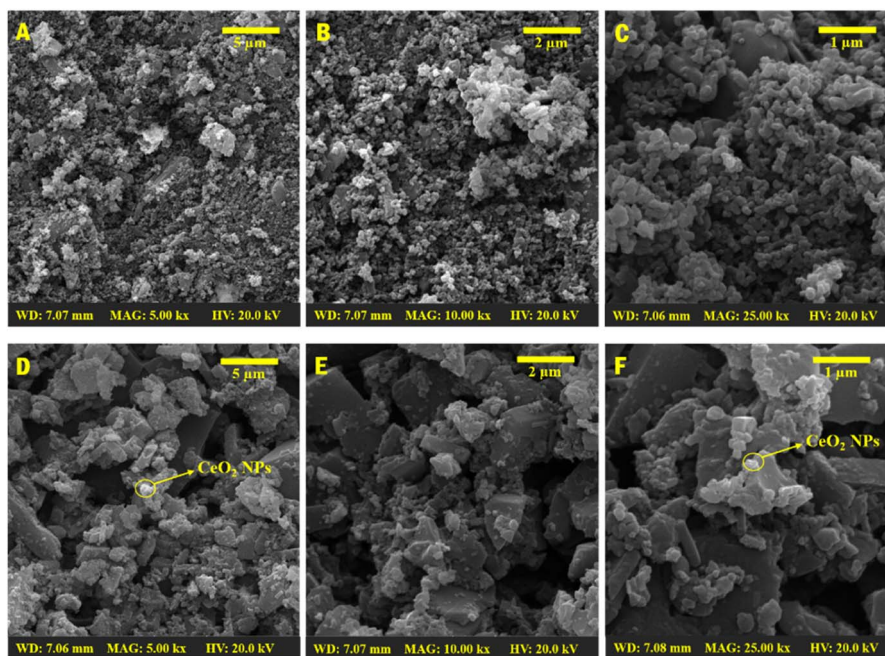


Fig. 4 SEM images of pure copolymer PLim-DAN (A–C) and PLim-DAN/CeO₂/CNTs-10% nanocomposites (D–F) ($\times = 5000, 10\,000, \text{ and } 25\,000$).

aggregated globules appearing on the surface, signifying the existence of cerium oxide nanoparticles (CeO₂ NPs).^{57,58} Furthermore, the images revealed the dispersion and coating of MWCNT nanofillers, with the polymer surface irregularly shaped growth on the MWCNTs being much denser and thicker for PLim-DAN/CeO₂/CNTs-10% with clustered MWCNTs on the surface.^{59,60} The nanofiller (MWCNT) had a large surface-to-volume ratio, which provided sorption sites for the luminol and 1,8-diaminonaphthalene monomers that polymerized, resulting in significant coverage of the MWCNTs.⁶¹

The elemental compositions of the synthesized copolymer PLim-DAN and the PLim-DAN/CeO₂/CNTs-10% nanocomposites were analyzed using electron diffraction X-ray analysis (EDX). The EDX spectrum showed the presence of C, N, and O in the unmodified copolymers (Fig. 5a).

Meanwhile, PLim-DAN/CeO₂/CNTs-10% nanocomposite confirmed the loading of CeO₂ NPs by demonstrating an additional peak attributed to Ce (Fig. 5b). Notably, the presence of

the Au peak in the EDX spectra can be attributed to the gold coating applied to the samples before analysis. Moreover, Fig. 6a and b illustrate the elemental mapping of the pure copolymer PLim-DAN and PLim-DAN/CeO₂/CNTs-10%, respectively. As seen in Fig. 6a, the C, O, and N elements were homogeneously distributed on the copolymer matrix. In contrast, the mapping in Fig. 6b exhibits an additional Ce element uniformly dispersed on the surface of the nanocomposite. The SEM/EDX outcomes established the successful synthesis of the copolymer and its nanocomposite through *in situ* oxidative polymerization.

The TEM technique was studied to provide a more detailed description of CeO₂ NPs and MWCNT dispersion within the synthesized copolymer matrix in the form of nanocomposites. TEM images of the PLim-DAN and PLim-DAN/CeO₂/CNTs-10% nanocomposites are presented in Fig. 7. The pure copolymer displayed conglomeration with an irregular spherical morphology (Fig. 7a and b).

The copolymer network in the MWCNT host showed strong interactions in the form of thin, dark threads that spread in the nanocomposite backbone of PLim-DAN/CeO₂/CNTs-10% (Fig. 7c–e), confirming the intercalation of MWCNTs into the PLim-DAN matrix and the small dark spots assigned to the presence of CeO₂ NPs.^{58–60} The stable intersection of the nanofiller enhanced the mass transport and electron transfer between the MWCNT and PLim-DAN copolymers through donor–acceptor interplay, which improved the electrochemical performance of the prepared nanocomposites.^{62,63}

The proposed structures of the pure copolymer and its nanocomposites were inspected using XRD, as illustrated in Fig. 8, to understand the crystalline nature and loading of nanoparticles on the polymeric matrix. The crystallographic

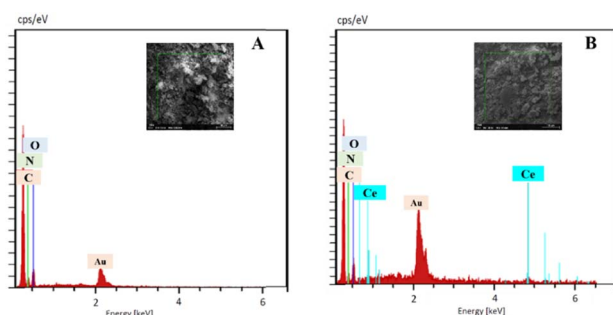


Fig. 5 EDX spectroscopy of pure copolymer PLim-DAN (A) and PLim-DAN/CeO₂/CNTs-10% nanocomposites (B).

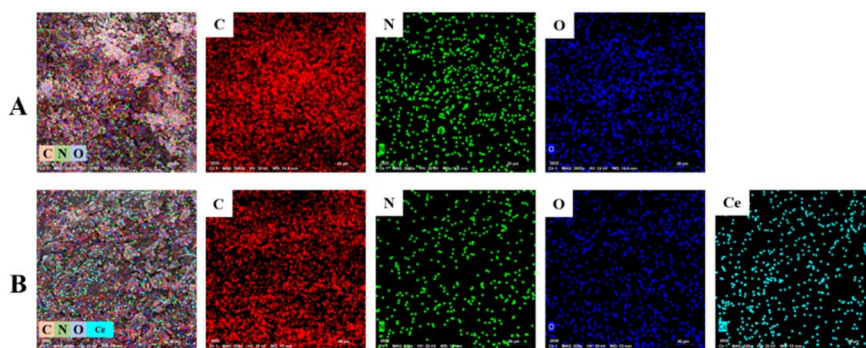


Fig. 6 EDX elemental maps of (A) pure copolymer PLim-DAN (C, N and O), and (B) PLim-DAN/CeO₂/CNTs-10% nanocomposites (C, N, O and Ce).

patterns of the pure copolymer and nanocomposites exhibited diffraction lines of a primarily crystalline nature. All spectra showed peaks at the low angle region between 6° and 25° which may indicate to the formation of ordered lamellar in the crystalline phase for polymeric (luminol-*co*-1,8-diaminonaphthalene) and the nanocomposites as (001) diffraction peaks at $2\theta = 8.6^\circ$ and the (110) plane at $2\theta = 14.5^\circ$ and 26° as main peaks for the π - π stacking distance of polymer network.^{64,65} Furthermore, the literature indicates that the primary diffraction peaks at between $2\theta = 14^\circ$ and 25° are associated with periodicities parallel and perpendicular to the polymer chains.⁶⁶ The XRD pattern of the pure copolymer PLim-DAN (Fig. 8a) revealed that the copolymer featured a semi-amorphous form with broad peaks of (001), (110) at 8° , 14° and 26° .^{54,55} The spectra of PLim-DAN/CeO₂/CNTs-3%, PLim-DAN/CeO₂/CNTs-5%, and PLim-DAN/CeO₂/CNTs-10% nanocomposites (Fig. 8b–d) show the same bands in the low angle region and extra bands at 28° , 37° , 54° , 73° are ascribed to the (111), (200), (220), and (400) lattice planes of the cubic structure of CeO₂ nanoparticles, demonstrating successful loading of NPs and in accordance with the

data registered in the literature (JCPDS 43-1002).^{67,68} Nevertheless, the studied nanocomposites (Fig. 8b–d) display the diffraction pattern of MWCNTs appearing at 2θ values of 26° (002), corresponding to reflections of the pseudo-graphite structure.⁶⁹ The peak intensity increased gradually with increasing MWCNT loading from 1% to 10%, confirming the insertion of MWCNTs into the copolymer molecules and the successful dispersion of the filler into the polymer matrix. Above that, the diffraction spectra revealed the effect of MWCNTs on the crystallinity of the nanocomposites. The characteristic diffraction peaks became sharper than before with the increase in the MWCNT loading, clearly confirming the crystalline phase of the nanocomposite.⁷⁰ The results were consistent with the outcomes obtained from the thermal analyses, where thermal stability increased with increasing MWCNT loading.

Thermal analysis

The thermal performance of the synthesized copolymers and nanocomposites was studied using TGA and derivative

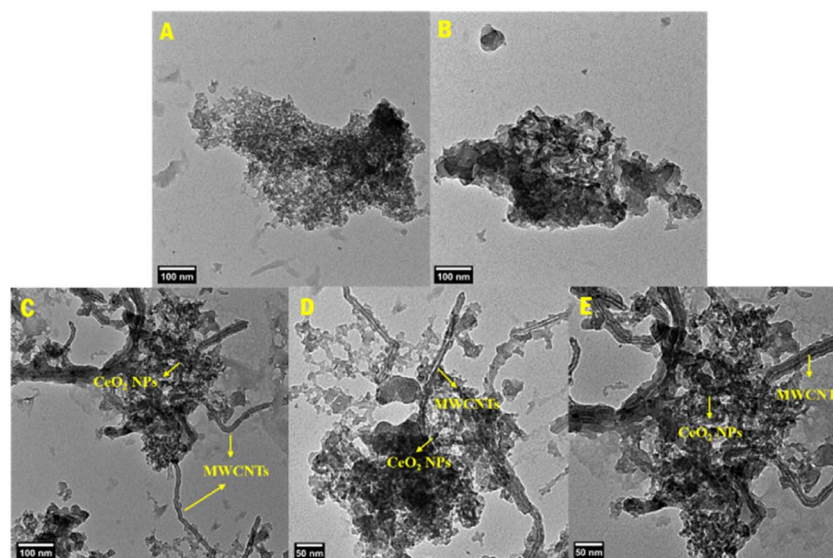


Fig. 7 TEM images of pure copolymer PLim-DAN (A and B) and PLim-DAN/CeO₂/CNTs-10% nanocomposites (C–E).



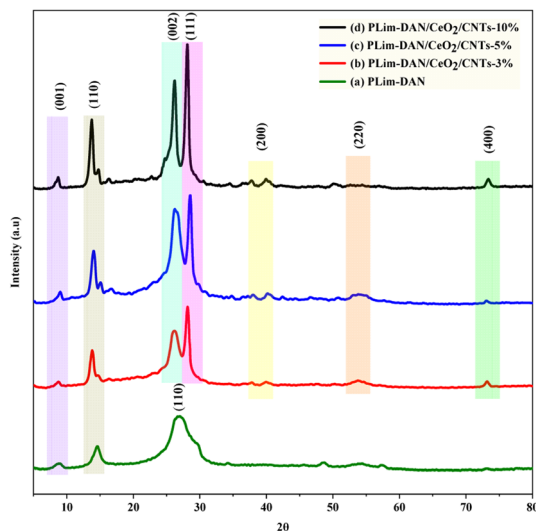


Fig. 8 XRD diffraction patterns for (a) PLim-DAN, (b) PLim-DAN/CeO₂/CNTs-3%, (c) PLim-DAN/CeO₂/CNTs-5%, and (d) PLim-DAN/CeO₂/CNTs-10% nanocomposites.

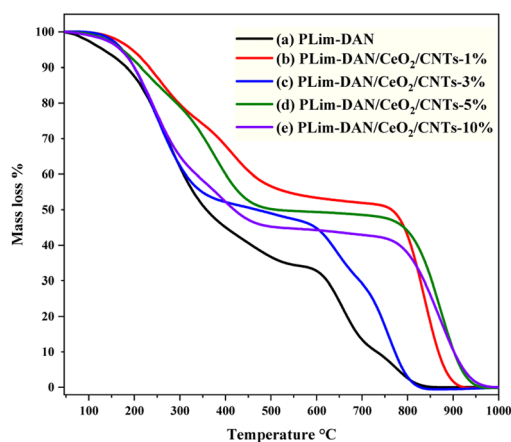


Fig. 9 TGA curves of (a) PLim-DAN (b–e) PLim-DAN/CeO₂/CNTs nanocomposites.

thermogravimetry (DTG) in the temperature range of 25–1000 °C at a heating rate of 10 °C min⁻¹ to provide more insights into the possible applications (Fig. 9a–e and 10a–e). The TGA/DTG curve exhibited a three-step weight loss system. The first stage occurred between 25 °C and 280 °C, attributed to moisture withdrawal. The second stage of weight loss followed between 280 °C and 760 °C owing to the degradation of the polymer backbone.^{71,72}

The third stage, observed between 760 °C and 940 °C, where the nanocomposites exhibited high thermal stability up to 800 °C for the 10%, 5% and 1% loading of MWCNTs with approximately 50% of the mass loss. Moreover, the nanocomposites (Fig. 9b–e) showed similar thermal performance when a fixed loading of CeO₂ NPs and different loadings of MWCNTs were introduced to the copolymer, implying that the nanocomposites displayed similar decomposition paths. The nanofiller addition of CeO₂ NPs and MWCNTs resulted in a shift in the second and

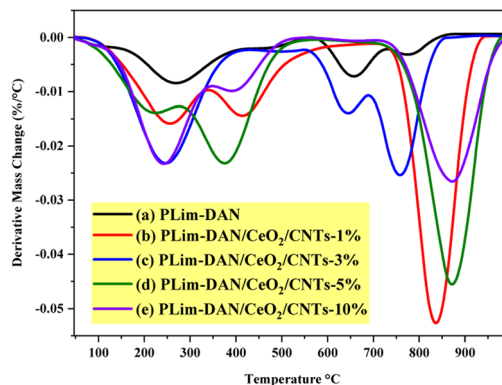


Fig. 10 DTG curves of (a) PLim-DAN (b) PLim-DAN/CeO₂/CNTs nanocomposites (b–e).

third stages of nanocomposite decomposition to higher temperature values. This result suggests that the nanofiller improved the thermal stability of the copolymer.^{57,73–75} Table 1 presents a comprehensive comparison of T_{10} , T_{25} , and T_{50} , illustrating thermal decomposition at 10%, 25%, and 50%, respectively. The T_{10} , T_{25} , and T_{50} values indicated a pattern performance controlled by the amount of MWCNTs loaded. As illustrated in Fig. 9 and Table 1, the increase in the percentage of the nanofiller was matched by an increase in thermal decomposition at 10%, 25%, and 50%.

Table 1 outlines the final polymer degradation temperature (PDT_{final}) and the maximum polymer decomposition temperature (PDT_{max}). As exemplified in Table 1, the PLim-DAN/CeO₂/CNTs-1% nanocomposite showed high optimum stability at T_{10} , T_{25} , and up to T_{50} from a thermal point of view. As the table and TGA clarified, the PDT_{final} values ranged from 815 °C to 950 °C, while the DTG revealed PDT_{max} in the 272–874 °C. The nanocomposite PLim-DAN/CeO₂/CNTs-3% demonstrated the lowest values of PDT_{final} and PDT_{max} in comparison with the other nanocomposites, whereas PLim-DAN/CeO₂/CNTs-10% displayed the highest values at both degradation temperatures.

Electrochemical study

Detection of chromium ions (Cr³⁺) employing PLim-DAN/CeO₂/CNTs NCs an electrochemical sensor that could specifically detect Cr³⁺ ions was manufactured using a glassy carbon electrode (GCE) and a composite of active PLim-DAN/CeO₂/CNTs-10% NCs. A Nafion conducting binding agent was employed to create a thin, uniform layer on the GCE. Nafion improves the adhesion of the PLim-DAN/CeO₂/CNTs NCs to the GCE and enhances the electron transfer rate of the desired electrochemical sensor through I–V analysis.^{76,77} The resulting electrochemical sensor demonstrated good sensitivity, a meager DL, a wide LDR, and long-term stability in a phosphate buffer medium with good reproducibility. During the initial stages of the I–V study, several heavy metal ions at a concentration of 0.1 μM and an applied potential of 0 to +1.5 V were analyzed in a phosphate buffer medium with a pH of 7.0. Cr³⁺ displays the highest I–V response among the electrochemical responses of As³⁺, Cd²⁺, Co²⁺, Cr³⁺, Ga³⁺, Hg²⁺, Pb²⁺, Sb³⁺, and Sn²⁺ ions (Fig. 11a).



Table 1 Thermal behavior of PLim-DAN and PLim-DAN/CeO₂/CNTs nanocomposites

| Sample | Temperature °C for various percentage decompositions ^a | | | PDT _{final} ^a °C | PDT _{max} ^b °C |
|-------------------------------------|---|-----------------|-----------------|--------------------------------------|------------------------------------|
| | T ₁₀ | T ₂₅ | T ₅₀ | | |
| PLim-DAN | 180 | 248 | 357 | 815 | 272 |
| PLim-DAN/CeO ₂ /CNTs-1% | 233 | 341 | 764 | 921 | 837 |
| PLim-DAN/CeO ₂ /CNTs-3% | 201 | 252 | 464 | 828 | 757 |
| PLim-DAN/CeO ₂ /CNTs-5% | 213 | 327 | 703 | 940 | 870 |
| PLim-DAN/CeO ₂ /CNTs-10% | 201 | 257 | 427 | 950 | 874 |

^a The values were determined by TGA at a heating rate of 10 °C min⁻¹. ^b The values were determined from the DTG curves.

The Cr³⁺ ion solution with concentrations ranging from 1.0 mM to 0.1 nM was investigated (Fig. 11b). The electrochemical responses were visibly distinguishable at different concentrations of Cr³⁺ ions from lower to higher. The analytical performance of the projected Cr³⁺ ions based on the PLim-DAN/CeO₂/CNTs-10% NCs/GCE was determined by plotting a calibration curve between the current and the concentration of Cr³⁺ ions, as shown in Fig. 11c. The measured slope of the calibration curve was used to determine the sensitivity and DL of the Cr³⁺ sensor. The obtained values were 19.78 μA μM⁻¹ cm⁻² and 4.80 ± 0.24 pM, respectively.

Furthermore, a contentious distribution of the current data at an applied potential of +1.5 V along the linear plot over the concentration range of 0.1 nM to 0.1 mM was obtained (Fig. 11c), signifying the LDR. LDR was determined to have a significantly wide range of concentrations. A curve of the current vs. log(concentration of Cr³⁺ ions) was plotted to assess the linearity of LDR (Fig. 11d). The current data were fitted with

a regression coefficient of $R^2 = 0.9908$, providing evidence of linear LDR.

The control experiment (Fig. 12) was conducted using 0.1 μM Cr³⁺ solutions in a buffer environment with modified GCE containing the pure copolymer and diverse percentages of MWCNT compositions (1–10%). The adapted GCE with the PLim-DAN/CeO₂/CNTs-10% NCs exhibited the highest current response among the other compositions and pure PLim-DAN. Hence, a 10% composition of CNT is the optimum composition for studying Cr³⁺ ions using an electrochemical method.

The reproducibility of the electrochemical sensor is a crucial reliability test. The test was performed using a 0.1 μM concentration of Cr³⁺ ion solution and an applied potential ranging from 0 to +1.5 V (Fig. 13a). The results displayed a perfectly indistinguishable (I–V) response, and the seven replicated runs showed no signs of alteration, even after washing the electrode following each run. These results confirm the reliability of the proposed Cr³⁺ ion sensor. The relative standard deviation (RSD) of the current data at an applied potential of +1.5 V was measured to evaluate the accuracy of this reproducibility test. As a result, a significant RSD value of 0.90 was obtained.

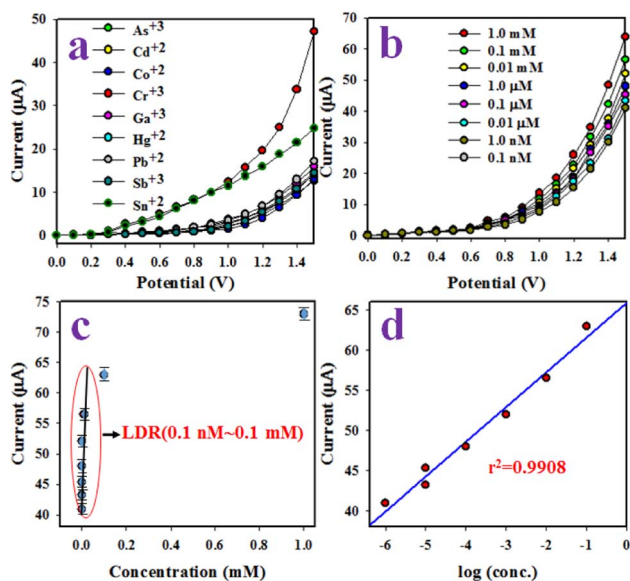


Fig. 11 Identification of sensor behavior using the electrochemical (I–V) approach. (a) Estimation of selectivity, (b) I–V responses based on the variation in the concentration of Cr³⁺ ions from lower to higher, (c) calibration curve, and (d) investigation of the linearity of LRD.

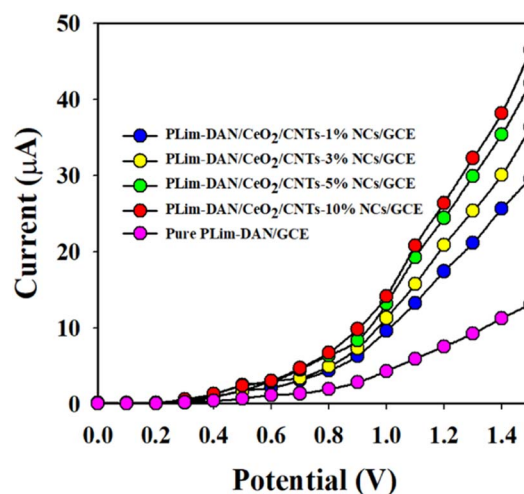


Fig. 12 A control experiment executed at 0.1 μM Cr³⁺ solutions in a buffer medium with modified GCE containing PLim-DAN and a 1–10% MWCNT composition of nanocomposites.



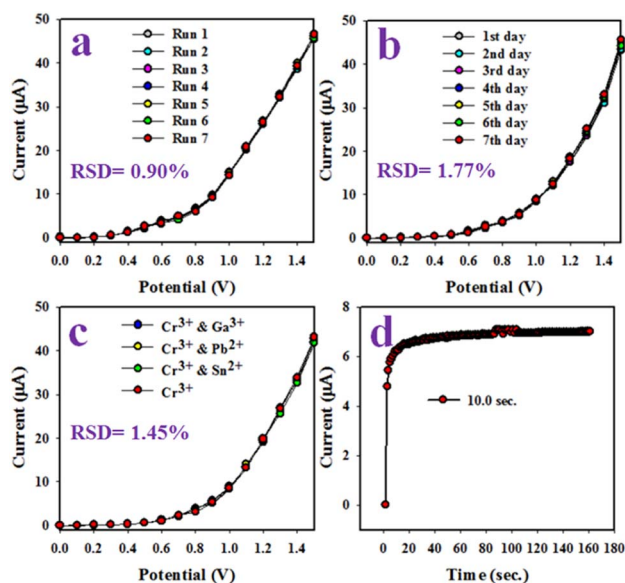


Fig. 13 (a) reproducibility test, (b) validity test of the Cr^{3+} ion sensor based on PLim-DAN/CeO₂/CNTs-10% NCs/GCE, (c) interference effect estimation, and (d) response time.

Similarly, a reproducibility test was conducted over approximately seven days to evaluate the stability and performance of the Cr^{3+} ion sensor.

The Cr^{3+} ion sensor displayed consistent results over a long period, with a % RSD of 1.77 (Fig. 13b). This result confirms the long-term stability of the Cr^{3+} ion sensor to a constant outcome of the I-V response. An interference test of the Cr^{3+} ion sensor based on PLim-DAN/CeO₂/CNTs-10% NCs/GCE was conducted (Fig. 13c). The findings exhibited high selectivity for Cr^{3+} ion sensors with no interference effects caused by other cations, such as Ga^{3+} , Pb^{2+} , and Sn^{2+} . Additionally, the response time of the sensor was measured to be approximately 10.0 seconds when a 0.1 μM concentration of Cr^{3+} ions was used in a phosphate buffer medium, as depicted in Fig. 13d. Overall, the sensor demonstrated adequate efficiency and performance.

The Cr^{3+} ion electrochemical sensor based on PLim-DAN/CeO₂/CNTs-10% NCs/GCE demonstrated impressive performance with high sensitivity ($19.7785 \mu\text{A} \mu\text{M}^{-1} \text{cm}^{-2}$), a wide LDR (0.1 nM to 0.1 mM), and a meager DL ($4.80 \pm 0.24 \text{ pM}$). Furthermore, the designed sensor exhibited reliable and consistent performance with a short response time (10.0 s) and excellent long-term stability in a phosphate buffer medium, making it a good candidate for assessment in actual

Table 2 Summarized sensing results of target PLim-DAN/CeO₂/CNTs-10% based fabricated sensor-probe towards different concentration of target analyte (Cr^{3+}) ions by electrochemical technique in terms of sensing response, sensing time, recovery time, and reversibility

| Trials | Analyte conc. | Sensing response (%) | Sensing time (s) | Recovery time (s) | Reversibility (%) |
|--------|--------------------|----------------------|------------------|-------------------|-------------------|
| 1 | 10.0 mM | 98.3 | 14 | 17 | 96.3 |
| 2 | 1.0 mM | 97.2 | 15 | 16 | 97.2 |
| 3 | 0.10 mM | 96.1 | 13 | 15 | 91.6 |
| 4 | 10.0 μM | 97.4 | 14 | 16 | 95.4 |
| 5 | 1.0 μM | 95.6 | 12 | 11 | 94.7 |
| 6 | 0.1 μM | 97.8 | 10 | 12 | 95.4 |
| 7 | 0.01 μM | 93.2 | 13 | 14 | 91.8 |
| 8 | 1.0 nM | 87.5 | 15 | 18 | 88.7 |
| 9 | 0.1 nM | 84.0 | 14 | 15 | 79.2 |
| 10 | 0.01 nM | 83.9 | 13 | 14 | 81.2 |

Table 3 Comparison of the performances of different electrochemical sensors for the determination of Cr^{3+}

| Modification material/electrode | Technique | Linear dynamic range | Limit of detection | Real sample composition | Ref. |
|-------------------------------------|-----------|---------------------------|---------------------|--|-----------|
| Chitosan-Au/SPCE | DPSV | 1–100 μM | 0.4 μM | Industrial wastewater | 78 |
| IIP (itaconic acid based)/CPE | DPV | 0.1–10 μM | 17.6 nM | River/sea water, urine | 79 |
| AgNP-LE/Pt | DPV | 10–90 μM | 0.804 μM | Lake/waste water | 80 |
| AuNPs/SPCE | LSV | 0.5–10 μM | 0.01 μM | Tea, coffee and mineral water | 81 |
| Bi film/GCE | FSV | 10^{-12} to 10^{-7} M | 0.3 pM | Tap water | 82 |
| 4-BBBSH/GCE | I–V | 100 pM to 100 mM | 95.5 pM | Coal/Red Sea/tap/well water, industrial effluent | 83 |
| AuNPs-mercapto succinic acid/ITO | ASV | 0.1–10 nM | 0.05 nM | Tap/pond/lake water | 84 |
| AuNP-3-mercaptopropionic acid/ITO | SWV | 200–500 ppb | 278 ppb | Pond water | 85 |
| PLim-DAN/CeO ₂ /CNTs/GCE | I–V | 0.1 nM to 0.1 mM | 4.80 pM | Industrial effluent/sea water/ground mineral water | This work |

^a SPCE: screen-printed carbon electrode; DPSV: differential pulse stripping voltammetry; CPE: carbon paste electrode; DPV: differential pulse voltammetry; LSV: linear sweep voltammetry; FSV: fast scan voltammetry; 4-BBBSH: (*E*)-*N'*-(4-bromobenzylidene)-4-benzenesulfonohydrazine; ITO: indium-doped tin oxide; ASV: anodic stripping voltammetry; SWV: square wave voltammetry.



environmental samples. The response and recovery times of the target PLim-DAN/CeO₂/CNTs-10% NCs/GCE sensor probe in the detection of target Cr³⁺ analyte by electrochemical method for 10 repetitive test and are included in Table 2.

Table 3 shows a comparison of the results obtained from the PLim-DAN/CeO₂/CNTs NCs/GCE modified sensor and several recently reported investigations in the literature for the detection of Cr³⁺ ions. The table comprises the modified material on the electrode, linear dynamic range, detection limits, and real sample test. As the table exemplified, the electrochemical

sensor, purposely designed for the detection of Cr³⁺ ions in this study, exhibits a low detection limit and a wide linear range. Additionally, the fabricated sensor reported here registered a lower detection limit than our previous research using the same (I-V) method.⁸³

Potential mechanism for the detection of Cr(III)

In this approach, the sensing mechanism can be explained as follows: the electrochemical detection of the PLim-DAN/CeO₂/CNTs is worked as a function of Cr³⁺ ions concentration interaction onto the fabricated surface at room conditions. The improved current response is observed with the fabricated electrode and probable mechanism is included in the Fig. 14. As obtained, the current response of the PLim-DAN/CeO₂/CNTs-10% film is significantly increased with the increasing concentration of target Cr³⁺ ions due to the large surface area of PLim-DAN/CeO₂/CNTs-10%, as well as electrochemical interaction and adsorption the target cations onto the sensor functional surface of PLim-DAN. The similar phenomena for toxic chromium ions detection with various nanocomposite materials have also been reported elsewhere.^{86,87} For a low concentration of Cr³⁺ ions in liquid medium, there is a smaller surface coverage of Cr³⁺ ions on PLim-DAN/CeO₂/CNTs-10%/GCE film and hence the surface reaction proceeds steadily. By increasing the Cr³⁺ ions concentration, the surface reaction is increased significantly (gradually increased the response) owing to surface

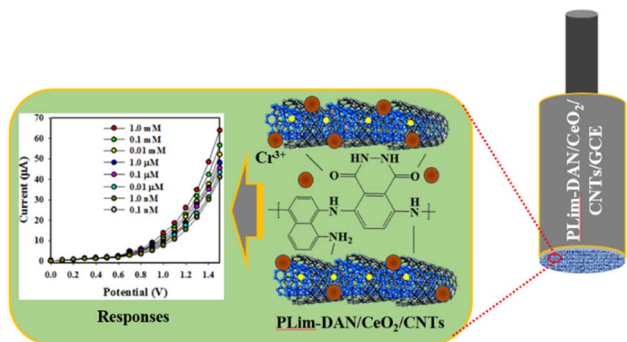


Fig. 14 Fabrication and mechanism of the probable interaction of Cr³⁺ with PLim-DAN/CeO₂/CNTs-10%/GCE with conducting 0.5% Nafion binders embedded onto GCE.

Table 4 Analyses of environmental samples with the PLim-DAN/CeO₂/CNTs-10% NCs/GCE sensor

| Samples | Added Cr ³⁺ ion concentration (µM) | Determined Cr ³⁺ conc. ^a by PLim-DAN/CeO ₂ /CNTs-10% NCs/GCE (µM) | | | Average recovery ^b (%) | RSD ^c (%) (n = 3) |
|----------------------|---|--|----------|-----------|-----------------------------------|------------------------------|
| | | R1 | R2 | R3 | | |
| Industrial effluent | 0.01000 | 0.009339 | 0.009689 | 0.0091616 | 93.97 | 2.86 |
| Sea water | 0.01000 | 0.009834 | 0.009627 | 0.009844 | 97.68 | 1.25 |
| Ground mineral water | 0.01000 | 0.009649 | 0.009908 | 0.009900 | 98.19 | 1.50 |

^a Mean of three repeated determinations (signal-to-noise ratio of 3) with PLim-DAN/CeO₂/CNTs-10% NCs/GCE. ^b Concentration of Cr³⁺ ions determined/concentration taken (unit: µM). ^c Relative standard deviation value indicates precision among three repeated measurements (R1, R2, and R3).

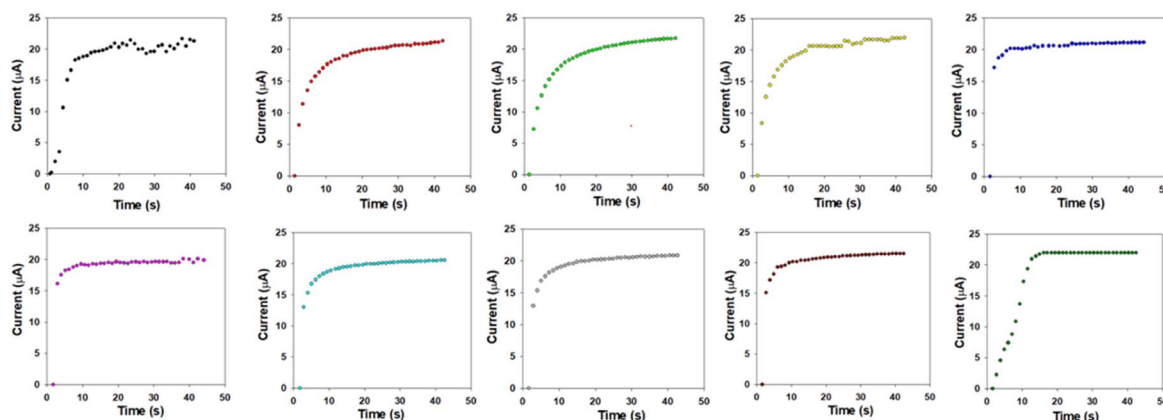


Fig. 15 Sensor responses towards target analyte (Cr³⁺) ions with prepared PLim-DAN/CeO₂/CNTs nanocomposite in identical conditions.



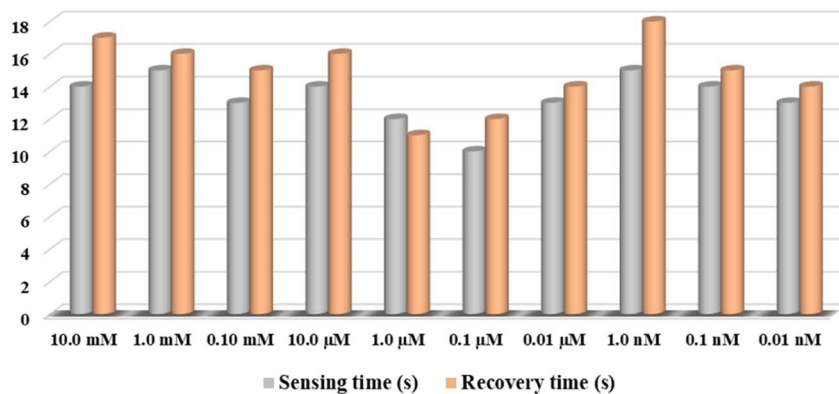


Fig. 16 The response time vs. recovery time for the modified PLim-DAN/CeO₂/CNTs-10% sensor during the range of 10.0 mM to 0.01 nM concentration.

area (assembly of PLim-DAN/CeO₂/CNTs-10%) contacted with Cr³⁺ ions (Fig. 14). Further increase of Cr³⁺ ions concentration onto PLim-DAN/CeO₂/CNTs-10%/GCE surface, it is exhibited a more rapid increasing the current responses, due to larger area covered by Cr³⁺ ions as well as the electrochemical interaction of the nitrogen and oxygen containing functional groups (Fig. 14) with the target cations. The interaction could be approaches as inter-molecular and intra-molecular interactions of the PLim-DAN layer with the target analyte.⁸⁸ Usually, the surface coverage of Cr³⁺ ions onto PLim-DAN/CeO₂/CNTs-10%/Naf/GCE surface is reached to the saturation level, based on the regular enhancement of current responses.

Real environmental samples analysis

The PLim-DAN/CeO₂/CNTs-10% NCs/GCE for the Cr³⁺ ion electrochemical sensor was utilized to examine various real environmental samples, including industrial waste effluent, ground mineral water, and seawater. Considerable results were obtained, as summarized in Table 4.

Furthermore, the PLim-DAN/CeO₂/CNTs-10% based fabricated sensor-probe response time vs. current has been calculated and illustrated in Fig. 15 for each concentration.^{89,90} Fig. 16 presented a relation between response time and recovery time at different concentration between 10.0 mM to 0.01 nM. As shown in Table 2 and Fig. 16, the 0.1 μM of Cr³⁺ solution displayed the fastest response time of 10 s and 12 s of recovery time with 97.8% of the sensing response.

Conclusion

Copolymerization of luminol and 1,8-diaminonaphthalene was designed and quaternary PLim-DAN/CeO₂/CNTs nanocomposites with different loading values (1, 3, 5, and 10%) were successfully modified and characterized in this work. XRD diffraction spectra and TGA revealed the effect of MWCNTs on the crystallinity and thermal stability of the nanocomposites, whereby increasing the MWCNT loading sharpens the characteristic diffraction peaks. This result confirms the crystalline phase of the nanocomposite, along

with its high thermal stability. TEM images revealed strong interactions between PLim-DAN and MWCNTs, enhancing the efficiency of mass transport and electron transfer in electrochemical sensing applications. The electroactivity of the designed nanocomposites toward different heavy metal ions was studied, and the results showed a high sensitivity to Cr³⁺ ions. The GCE modified with the PLim-DAN/CeO₂/CNTs-10% NCs exhibited the highest current response among the other compositions. The chromium ion sensor was fabricated by attaching PLim-DAN/CeO₂/CNTs-10% NCs onto a GCE with a conducting Nafion-binding agent. The sensor was then used to detect Cr³⁺ ions in a phosphate buffer solution, showing excellent analytical performance regarding sensitivity, linear dynamic range, and DL. The sensor was also reliable, with a short response time and reproducible results.

Author contributions

S. Al-Sodies: methodology, investigation, writing – original draft & editing; M. M. Alam: methodology; K. A. Alamry: conceptualization, investigation; M. A. Hussein: conceptualization, investigation, writing – review & editing; A. M. Asiri: review & editing – final draft; M. M. Rahman: methodology, investigation, writing – review & editing.

Conflicts of interest

There are no conflicts to declare.

References

- C. M. Geiselhart, C. B. Kowolik and H. Mutlu, Untapped toolbox of luminol based polymers, *Polym. Chem.*, 2021, **12**, 1732–1748, DOI: [10.1039/d1py00034a](https://doi.org/10.1039/d1py00034a).
- C. A. Marquette and L. J. Blum, Applications of the luminol chemiluminescent reaction in analytical chemistry, *Anal. Bioanal. Chem.*, 2006, **385**, 546, DOI: [10.1007/s00216-006-0439-9](https://doi.org/10.1007/s00216-006-0439-9).
- C. Dodeigne, L. Thunus and R. Lejeune, Chemiluminescence as diagnostic tool. A review, *Talanta*, 2000, **51**, 415, DOI: [10.1016/s0039-9140\(99\)00294-5](https://doi.org/10.1016/s0039-9140(99)00294-5).



- 4 P. Khan, D. Idrees, M. A. Moxley, J. A. Corbett, F. Ahmad, G. von Figura, W. S. Sly, A. Waheed and M. I. Hassan, Luminol-based chemiluminescent signals: clinical and non-clinical application and future uses, *Appl. Biochem. Biotechnol.*, 2014, **173**, 333, DOI: [10.1007/s12010-014-0850-1](https://doi.org/10.1007/s12010-014-0850-1).
- 5 Y. Wang, S. Hamid, X. Zhang, N. Akhtar, X. Zhang and T. He, An electrochemiluminescent biosensor for dopamine detection using a poly(luminol-benzidine sulfate) electrode modified by tyramine oxidase, *New J. Chem.*, 2017, **41**, 1591, DOI: [10.1039/C6NJ03338E](https://doi.org/10.1039/C6NJ03338E).
- 6 L. Wei, Y. Zhang, N. Eziz, Y. Yang, G. Li and M. Guan, An ultrasensitive electrochemiluminescence immunosensor for alpha-fetoprotein based on a poly(aniline-luminol)/graphene oxide nanocomposite, *Anal. Bioanal. Chem.*, 2019, **411**, 5175, DOI: [10.1007/s00216-019-01897-w](https://doi.org/10.1007/s00216-019-01897-w).
- 7 M. B. Gholivand and E. Ahmadi, Square Wave Anodic Stripping Voltammetric Determination of Paracetamol at Poly Luminol/Functionalized Multi-walled Carbon Nanotubes Modified Glassy Carbon Electrode, *Russ. J. Electrochem.*, 2019, **55**(12), 1151, DOI: [10.1134/S102319351912005X](https://doi.org/10.1134/S102319351912005X).
- 8 Y. Yang, Y. Zhang, L. Wei, G. Li, M. Guan and S. Tian, A Highly Sensitive Electrochemiluminescence Choline Biosensor Based on Poly(aniline-luminol-hemin) Nanocomposites, *Electroanalysis*, 2019, **31**, 624, DOI: [10.1002/elan.201800582](https://doi.org/10.1002/elan.201800582).
- 9 R. Dong, Y. Zhang, J. Huang, M. Habibul and G. Li, Electrochemiluminescence DNA Biosensor for HBV Based on Coraloid Poly(Aniline-Luminol)-MWCNTs and Catalysis of Ferrocene, *Electroanalysis*, 2022, **34**, 1555, DOI: [10.1002/elan.202200020](https://doi.org/10.1002/elan.202200020).
- 10 S. Al-Sodies, A. M. Asiri, K. A. Alamrya and M. A. Hussein, Exploiting poly(safranin) and poly(luminol) for sensing applications. A mini review, *RSC Adv.*, 2023, **13**, 9697, DOI: [10.1039/D3RA00532A](https://doi.org/10.1039/D3RA00532A).
- 11 X.-G. Li, M.-R. Huang, W. Duan and Y.-L. Yang, Novel Multifunctional Polymers from Aromatic Diamines by Oxidative Polymerizations, *Chem. Rev.*, 2002, **102**(9), 2925, DOI: [10.1021/cr010423z](https://doi.org/10.1021/cr010423z).
- 12 X.-G. Li, M.-R. Huang and S.-X. Li, Facile synthesis of poly(1,8-diaminonaphthalene) microparticles with a very high silver-ion adsorbability by a chemical oxidative polymerization, *Acta Mater.*, 2004, **52**, 5363, DOI: [10.1016/j.actamat.2004.07.042](https://doi.org/10.1016/j.actamat.2004.07.042).
- 13 J. -W. Lee, D. -S. Park, Y. -B. Shim and S. -M. Park, Electrochemical Characterization of Poly(1,8-diaminonaphthalene): A Functionalized Polymer, *J. Electrochem. Soc.*, 1992, **139**, 3507, DOI: [10.1149/1.2069107](https://doi.org/10.1149/1.2069107).
- 14 B. J. Pałys, M. Skompska and K. Jackowska, Sensitivity of poly 1,8-diaminonaphthalene to heavy metal ions — electrochemical and vibrational spectra studies, *J. Electroanal. Chem.*, 1997, **433**, 41, DOI: [10.1016/S0022-0728\(97\)00144-7](https://doi.org/10.1016/S0022-0728(97)00144-7).
- 15 T. Akkaya, M. Gülfen and U. Olgun, Adsorption of rhodium(III) ions onto poly(1,8-diaminonaphthalene) chelating polymer: Equilibrium, kinetic and thermodynamic study, *React. Funct. Polym.*, 2013, **73**, 1589, DOI: [10.1016/j.reactfunctpolym.2013.09.001](https://doi.org/10.1016/j.reactfunctpolym.2013.09.001).
- 16 H. Park, T. Kwon, D.-S. Park and Y.-B. Shim, Electrocatalytic Reduction of Molecular Oxygen at Poly(1,8-diaminonaphthalene) and Poly(Co(II)-(1,8-diaminonaphthalene)) Coated Electrodes, *Bull. Korean Chem. Soc.*, 2006, **27**, 1763, DOI: [10.5012/BKCS.2006.27.11.1763](https://doi.org/10.5012/BKCS.2006.27.11.1763).
- 17 A. Rkoçı, M. Pumera, X. Llopis, B. Pérez, M. Del Valle and S. Alegret, New materials for electrochemical sensing VI: Carbon nanotubes, *TrAC, Trends Anal. Chem.*, 2005, **24**, 826, DOI: [10.1016/j.trac.2005.03.019](https://doi.org/10.1016/j.trac.2005.03.019).
- 18 Y. Jia, N. Zhang, Y. Du, X. Ren, H. Ma, D. Wu, D. Fan, Q. Wei and H. Ju, Nanoarrays-propped in situ photoelectrochemical system for microRNA detection, *Biosens. Bioelectron.*, 2022, **210**, 114291, DOI: [10.1016/j.bios.2022.114291](https://doi.org/10.1016/j.bios.2022.114291).
- 19 Y. Jia, R. Xu, K. Xu, S. Wang, X. Ren, N. Zhang, D. Wu, H. Ma, Y. Li and Q. Wei, Ternary Z-Scheme Ag-Embedded TiO₂-Ag₂S Nanojunction as a Novel Photoelectrochemical Converter for CD44 Detection, *Anal. Chem.*, 2022, **94**(33), 11713, DOI: [10.1021/acs.analchem.2c03046](https://doi.org/10.1021/acs.analchem.2c03046).
- 20 Y. Jia, X. Ren, X. Zhang, D. Wu, H. Ma, Y. Li and Q. Wei, Encapsulation of Tetraphenylethylene Derivative in Liposome Vesicles as Promising Aggregation Induced Electrochemiluminescence Emitter for Detection of Human Epidermal Growth Factor Receptor 2, *Anal. Chem.*, 2023, **95**(24), 9139, DOI: [10.1021/acs.analchem.3c01957L](https://doi.org/10.1021/acs.analchem.3c01957L).
- 21 L. Ma, X. Dong, M. Chen, L. Zhu, C. Wang, F. Yang and Y. Dong, Fabrication and water treatment application of carbon nanotubes (CNTs)-based composite membranes: A review, *Membranes*, 2017, **7**, 16, DOI: [10.3390/membranes7010016](https://doi.org/10.3390/membranes7010016).
- 22 X. Jia and F. Wei, Advances in Production and Applications of Carbon Nanotubes, in *Single-Walled Carbon Nanotubes. Topics in Current Chemistry Collections*, ed. Li Y. and Maruyama S., Springer, Cham, Switzerland, 2019, p. 299, DOI: [10.1007/978-3-030-12700-8_11](https://doi.org/10.1007/978-3-030-12700-8_11).
- 23 Y. Wang, L. Tang, Y. Li, Q. Li and B. Bai, Preparation of modified multi-walled carbon nanotubes as a reinforcement for epoxy shape-memory polymer composites, *Polym. Adv. Technol.*, 2021, **32**, 67, DOI: [10.1002/pat.5061](https://doi.org/10.1002/pat.5061).
- 24 M. Lai, L. Jiang, X. Wang, H. Zhou, Z. Huang and H. Zhou, Effects of multi-walled carbon nanotube/graphene oxide-based sizing on interfacial and tribological properties of continuous carbon fiber/poly(ether ether ketone) composites, *Mater. Chem. Phys.*, 2022, **276**, 125344, DOI: [10.1016/j.matchemphys.2021.125344](https://doi.org/10.1016/j.matchemphys.2021.125344).
- 25 A. Savk, B. Özdil, B. Demirkan, M. S. Nas, M. H. Calimli, M. H. Alma, A. M. Asiri and F. Sen, Multi-walled carbon nanotube-based nanosensor for ultrasensitive detection of uric acid, dopamine, and ascorbic acid, *Mater. Sci. Eng., C*, 2019, **99**, 248, DOI: [10.1016/j.msec.2019.01.113](https://doi.org/10.1016/j.msec.2019.01.113).
- 26 K. K. Reddy, M. Satyanarayana, K. Y. Goud, K. V. Gobi and H. Kim, Carbon nanotube ensembled hybrid nanocomposite electrode for direct electrochemical detection of epinephrine in pharmaceutical tablets and urine, *Mater. Sci. Eng., C*, 2017, **79**, 93, DOI: [10.1016/j.msec.2017.05.012](https://doi.org/10.1016/j.msec.2017.05.012).



- 27 F. Tahernejad-Javazmi, F. M. Shabani-Nooshabadi and H. Karimi-Maleh, Analysis of glutathione in the presence of acetaminophen and tyrosine via an amplified electrode with MgO/SWCNTs as a sensor in the hemolyzed erythrocyte, *Talanta*, 2018, **176**, 208, DOI: [10.1016/j.talanta.2017.08.027](https://doi.org/10.1016/j.talanta.2017.08.027).
- 28 W. Shao, J. Mai and Z. Wei, Nonenzymatic lactic acid detection using cobalt polyphthalocyanine/carboxylated multi-walled carbon nanotube nanocomposite-modified sensor, *Chemosensors*, 2022, **10**, 83, DOI: [10.3390/chemosensors10020083](https://doi.org/10.3390/chemosensors10020083).
- 29 A. Mourya, B. Mazumdar and S. K. Sinha, Application of CeO₂-MWCNTs nanocomposite for heavy metal ion detection in aqueous solutions by electrochemical technique, *Cleaner Mater.*, 2021, **2**, 100021, DOI: [10.1016/j.clema.2021.100021](https://doi.org/10.1016/j.clema.2021.100021).
- 30 D. F. Katowah, M. A. Hussein, M. M. Alam, S. H. Ismail, O. I. Osman, T. R. Sobahi, A. M. Asiri, J. Ahmed and M. M. Rahman, Designed network of ternary core-shell PPCOT/NiFe₂O₄/C-SWCNTs nanocomposites. A Selective Fe³⁺ ionic sensor, *J. Alloys Compd.*, 2020, **834**, 155020, DOI: [10.1016/j.jallcom.2020.155020](https://doi.org/10.1016/j.jallcom.2020.155020).
- 31 D. F. Katowah, M. A. Hussein, M. M. Alam, T. R. Sobahi, M. A. Gabal, A. M. Asiri and M. M. Rahman, Poly(pyrrole-co-o-toluidine) wrapped CoFe₂O₄/R(GO-OXSWCNTs) ternary composite material for Ga³⁺ sensing ability, *RSC Adv.*, 2019, **9**, 33052, DOI: [10.1039/c9ra03593a](https://doi.org/10.1039/c9ra03593a).
- 32 Y. Wei, M. Li, S. Jiao, Q. Huang, G. Wang and B. Fang, Fabrication of CeO₂ nanoparticles modified glassy carbon electrode and its application for electrochemical determination of UA and AA simultaneously, *Electrochim. Acta*, 2006, **52**, 766, DOI: [10.1016/j.electacta.2006.06.006](https://doi.org/10.1016/j.electacta.2006.06.006).
- 33 J. Roh, S. H. Hwang and J. Jang, Dual-functional CeO₂: Eu³⁺ nanocrystals for performance-enhanced dye-sensitized solar cells, *ACS Appl. Mater. Interfaces*, 2014, **6**, 19825, DOI: [10.1021/am505194k](https://doi.org/10.1021/am505194k).
- 34 A. Umar, R. Kumar, M. Akhtar, G. Kumar and S. Kim, Growth and properties of well-crystalline cerium oxide (CeO₂) nanoflakes for environmental and sensor applications, *J. Colloid Interface Sci.*, 2015, **454**, 61, DOI: [10.1016/j.jcis.2015.04.055](https://doi.org/10.1016/j.jcis.2015.04.055).
- 35 Y. Wei, Q.-A. Huang, M.-G. Li, X.-J. Huang, B. Fang and L. Wang, CeO₂ nanoparticles decorated multi-walled carbon nanotubes for electrochemical determination of guanine and adenine, *Electrochim. Acta*, 2011, **56**, 8571, DOI: [10.1016/j.electacta.2011.07.048](https://doi.org/10.1016/j.electacta.2011.07.048).
- 36 S. Yang, G. Li, G. Wang, L. Liu, D. Wang and L. Qu, Synthesis of highly dispersed CeO₂ nanoparticles on N-doped reduced oxide graphene and their electrocatalytic activity towards H₂O₂, *J. Alloys Compd.*, 2016, **688**, 910, DOI: [10.1016/j.jallcom.2016.07.113](https://doi.org/10.1016/j.jallcom.2016.07.113).
- 37 A. S. Razavian, S. M. Ghoreishi, A. S. Esmaeily, M. Behpour, L. M. Monzon and J. M. D. Coey, Simultaneous sensing of L-tyrosine and epinephrine using a glassy carbon electrode modified with Nafion and CeO₂ nanoparticles, *Mikrochim. Acta*, 2014, **181**, 1947, DOI: [10.1007/s00604-014-1284-8](https://doi.org/10.1007/s00604-014-1284-8).
- 38 B. Pandit, N. Kumar, P. M. Koinkar and B. R. Sankapal, Solution processed nanostructured cerium oxide electrode: Electrochemical engineering towards solid-state symmetric supercapacitor device, *J. Electroanal. Chem.*, 2019, **839**, 96, DOI: [10.1016/j.jelechem.2019.02.047](https://doi.org/10.1016/j.jelechem.2019.02.047).
- 39 A. Pechova and L. Pavlata, Chromium as an essential nutrient: A review, *Vet. Med.*, 2007, **52**, 1, DOI: [10.17221/2010-VETMED](https://doi.org/10.17221/2010-VETMED).
- 40 N. Zhang, J. S. Suleiman, M. He and B. Hu, Chromium (III)-imprinted silica gel for speciation analysis of chromium in environmental water samples with ICP-MS detection, *Talanta*, 2008, **75**, 536, DOI: [10.1016/j.talanta.2007.11.059](https://doi.org/10.1016/j.talanta.2007.11.059).
- 41 A. R. Walsh, J. Ohalloran and A. M. Gower, Some effects of elevated levels of chromium (III) in sediments to the mullet *Chelon labrosus* (R), *Ecotoxicol. Environ. Saf.*, 1994, **7**, 168, DOI: [10.1006/eesa.1994.1014](https://doi.org/10.1006/eesa.1994.1014).
- 42 Z. Zhou, M. Yu, H. Yang, K. Huang, F. Li, T. Yi and C. Huang, FRET-based sensor for imaging chromium (III) in living cells, *Chem. Commun.*, 2008, 3387, DOI: [10.1039/B801503A](https://doi.org/10.1039/B801503A).
- 43 H. Zhang, G. Zhang, J. Xu, Y. Wen, S. Ming, J. Zhang and W. Ding, Highly sensitive "turn-on" fluorescent chemical sensor for trace analysis of Cr³⁺ using electro-synthesized poly (N-(9-fluorenylmethoxycarbonyl)-L-histidine), *Spectrochim. Acta, Part A*, 2018, **191**, 79, DOI: [10.1016/j.saa.2017.09.008](https://doi.org/10.1016/j.saa.2017.09.008).
- 44 H. A. Zamani, G. Rajabzadeh and M. R. Ganjali, Highly selective and sensitive chromium (III) membrane sensors based on 4-amino-3-hydrazino-6-methyl-1, 2, 4-triazin-5-one as a new neutral ionophore, *Sens. Actuators, B*, 2006, **119**, 41, DOI: [10.1016/j.snb.2005.11.048](https://doi.org/10.1016/j.snb.2005.11.048).
- 45 A. K. Singh, V. K. Gupta and B. Gupta, Chromium (III) selective membrane sensors based on Schiff bases as chelating ionophores, *Anal. Chim. Acta*, 2007, **585**, 171, DOI: [10.1016/j.aca.2006.11.074](https://doi.org/10.1016/j.aca.2006.11.074).
- 46 N. Dogra, M. Singh, A. Kumar, *et al.*, Selective room-temperature ammonia sensing using CeO₂-multiwalled carbon nanotube composite, *Appl. Phys. A*, 2023, **129**, 24, DOI: [10.1007/s00339-022-06283-5](https://doi.org/10.1007/s00339-022-06283-5).
- 47 S. Shanavas, T. Ahamad, S. M. Alshehri, A. Sultan, R. Acevedo and P. M. Anbarasan, Development of high-performance fiber optic gas sensor-based rice-like CeO₂/MWCNT nanocomposite synthesized by facile hydrothermal route, *Opt Laser. Technol.*, 2020, **123**, 105902, DOI: [10.1016/j.optlastec.2019.105902](https://doi.org/10.1016/j.optlastec.2019.105902).
- 48 A. Molina, M. Al-Sardar, V. Rodriguez-Gonzalez, V. Escobar-Barrios, A. A. Zakhidov, A. I. Mtz-Enriquez, A. Encinas and J. Oliva, Efficient NO₂ detection and the sensing mechanism of stretchable/biodegradable MWCNT based sensors decorated with CeO₂ nanoparticles, *Synth. Met.*, 2022, **287**, 117091, DOI: [10.1016/j.synthmet.2022.117091](https://doi.org/10.1016/j.synthmet.2022.117091).
- 49 J. Ai, X. Wang, Y. Zhang, *et al.*, A sensitive electrochemical sensor for nitrophenol detection based on CeO₂/MWCNTs nanocomposite, *Appl. Phys. A*, 2022, **128**, 831, DOI: [10.1007/s00339-022-05952-9](https://doi.org/10.1007/s00339-022-05952-9).
- 50 R. M. Shereema, S. R. Nambiar, S. S. Shankar and T. P. Rao, CeO₂-MWCNT nanocomposite based electrochemical



- sensor for acetaldehyde, *Anal. Methods*, 2015, 7, 4912, DOI: [10.1039/C5AY00747J](https://doi.org/10.1039/C5AY00747J).
- 51 A. Üge, D. K. Zeybek and B. Zeybek, An electrochemical sensor for sensitive detection of dopamine based on MWCNTs/CeO₂-PEDOT composite, *J. Electroanal. Chem.*, 2018, **813**, 134, DOI: [10.1016/j.jelechem.2018.02.028](https://doi.org/10.1016/j.jelechem.2018.02.028).
- 52 J. N'Diaye and K. Lian, Investigation of the chemical structure and electrochemical activity of a chemically polymerized luminol, *J. Electroanal. Chem.*, 2019, **839**, 90–95, DOI: [10.1016/j.jelechem.2019.03.024](https://doi.org/10.1016/j.jelechem.2019.03.024).
- 53 H. H. AL-Refai, A. A. Ganash and M. A. Hussein, Sensitive and selective voltammetric sensor based on polythiophene nanocomposite mixed MWCNT-G for the determination of Tartrazine, *Synth. Met.*, 2021, **280**, 116875, DOI: [10.1016/j.synthmet.2021.116875](https://doi.org/10.1016/j.synthmet.2021.116875).
- 54 S. Jadoun, S. M. Ashraf and U. Riaz, Microwave-assisted synthesis of copolymers of luminol with anisidine: Effect on spectral, thermal and fluorescence characteristics, *Polym. Adv. Technol.*, 2018, **29**, 1007, DOI: [10.1002/pat.4212](https://doi.org/10.1002/pat.4212).
- 55 U. Riaz, S. Jadoun, P. Kumar, R. Kumarc and N. Yadav, Microwave-assisted facile synthesis of poly(luminol-co-phenylenediamine) copolymers and their potential application in biomedical imaging, *RSC Adv.*, 2018, **8**, 37165, DOI: [10.1039/C8RA08373H](https://doi.org/10.1039/C8RA08373H).
- 56 S. Ansari, M. S. Ansari, N. Dev and S. P. Satsangee, CeO₂ nanoparticle-based electrochemical sensor for an anti-anginal drug, *Mater. Today: Proc.*, 2019, **18**, 1210, DOI: [10.1016/j.matpr.2019.06.583](https://doi.org/10.1016/j.matpr.2019.06.583).
- 57 A. L. Nikolaeva, I. V. Gofman, A. V. Yakimansky, E. M. Ivan'kova, I. V. Abalov, A. E. Baranchikov and V. K. Ivanov, Polyimide-Based Nanocomposites with Binary CeO₂/Nanocarbon Fillers: Conjointly Enhanced Thermal and Mechanical Properties, *Polymers*, 2020, **12**, 1952, DOI: [10.3390/polym12091952](https://doi.org/10.3390/polym12091952).
- 58 M. Mauro, M. Crosera, M. Monai, T. Montini, P. Fornasiero, M. Bovenzi, G. Adami, G. Turco and F. Laese Filon, Cerium Oxide Nanoparticles Absorption through Intact and Damaged Human Skin, *Molecules*, 2019, **24**, 3759, DOI: [10.3390/molecules24203759](https://doi.org/10.3390/molecules24203759).
- 59 Q. Wang, Y. Wang, B. Duan and M. Zhang, Modified Sol-Gel Synthesis of Carbon Nanotubes Supported Titania Composites with Enhanced Visible Light Induced Photocatalytic Activity, *J. Nanomater.*, 2016, 3967156, DOI: [10.1155/2016/3967156](https://doi.org/10.1155/2016/3967156).
- 60 P. M. Masipa, T. Magadzu and B. Mkhondo, Decoration of Multi-walled Carbon Nanotubes by Metal Nanoparticles and Metal Oxides using Chemical Evaporation Method, *S. Afr. J. Chem.*, 2013, **66**, 173. <https://org.za/pdf/sajc/v66/29.pdf>.
- 61 J. Zhuang, D. Gu, L. Xi, K. Lin, Y. Fang and R. Wang, Preparation method and underlying mechanism of MWCNTs/Ti6Al4V nanocomposite powder for selective laser melting additive manufacturing, *Powder Technol.*, 2020, **368**, 59, DOI: [10.1016/j.powtec.2020.04.041](https://doi.org/10.1016/j.powtec.2020.04.041).
- 62 R. J. Chen, Y. Zhang, D. Wang and H. Dai, Noncovalent Sidewall Functionalization of Single-Walled Carbon Nanotubes for Protein Immobilization, *J. Am. Chem. Soc.*, 2001, **123**(16), 3838–3839, DOI: [10.1021/ja010172b](https://doi.org/10.1021/ja010172b).
- 63 F. J. M. Hoeben, P. Jonkheijm, E. W. Meijer and A. P. H. J. Schenning, About Supramolecular Assemblies of π -Conjugated Systems, *Chem. Rev.*, 2005, **105**(4), 1491, DOI: [10.1021/cr030070z](https://doi.org/10.1021/cr030070z).
- 64 F. Neve and O. Francescangeli, Layered ω -Substituted Alkylpyridinium Salts with Inorganic Anions: Effects of H-Bonding Patterns on the Layer Thickness, *Cryst. Growth Des.*, 2005, **5**(1), 163, DOI: [10.1021/cg049894f](https://doi.org/10.1021/cg049894f).
- 65 S. Zhang, J. Liu, Y. Han and L. Wang, Polymer Electron Acceptors Based on Iso-Naphthalene Diimide Unit with High LUMO Levels, *Macromol. Chem. Phys.*, 2017, **218**, 1600606, DOI: [10.1002/macp.201600606](https://doi.org/10.1002/macp.201600606).
- 66 A. Mahudeswaran, J. Vivekanandan, A. Jeeva, J. Chandrasekaran and P. S. Vijayanand, Synthesis, characterization, optical and electrical properties of nanostructured poly(aniline-co-o-bromoaniline) prepared by in-situ polymerization method, *Optik*, 2016, **127**, 3984, DOI: [10.1016/j.ijleo.2016.01.095](https://doi.org/10.1016/j.ijleo.2016.01.095).
- 67 B. Pandit, B. R. Sankapal and P. M. Koinkar, Novel chemical route for CeO₂/MWCNTs composite towards highly bendable solid-state supercapacitor device, *Sci. Rep.*, 2019, **9**, 5892, DOI: [10.1038/s41598-019-42301-y](https://doi.org/10.1038/s41598-019-42301-y).
- 68 E. Kumar, P. Selvarajan and K. Balasubramanian, Preparation and Studies of Cerium Dioxide (CeO₂) Nanoparticles by Microwave-Assisted Solution Method, *Recent Res. Sci. Technol.*, 2010, **2**(4), 37.
- 69 Z. Xu, L. Liu, Y. Huang, Y. Sun, X. Wu and J. Li, Graphitization of polyacrylonitrile carbon fibers and graphite irradiated by γ rays, *Mater. Lett.*, 2009, **63**, 1814, DOI: [10.1016/j.matlet.2009.05.055](https://doi.org/10.1016/j.matlet.2009.05.055).
- 70 M. K. H. Bhuiyan, M. M. Rahman, M. F. Mina, M. R. Islam, M. Abdul Gafur and A. Begum, Crystalline morphology and properties of multi-walled carbon nanotube filled isotactic polypropylene nanocomposites: Influence of filler size and loading, *Composites, Part A*, 2013, **52**, 70, DOI: [10.1016/j.compositesa.2013.05.011](https://doi.org/10.1016/j.compositesa.2013.05.011).
- 71 H. Kang, L. Xu, Y. Cai, Y. Liu, F. Jiang, J. Xu and W. Zhou, Using boronic acid functionalization to simultaneously enhance electrical conductivity and thermoelectric performance of free-standing polythiophene film, *Eur. Polym. J.*, 2021, **144**, 110208, DOI: [10.1016/j.eurpolymj.2020.110208](https://doi.org/10.1016/j.eurpolymj.2020.110208).
- 72 R. Singh, A. K. Bajpai and A. K. Shrivastava, CdSe QDs reinforced poly(1, 8 diaminonaphthalene) (PDAN) offers improved thermal and AC conductivity properties, *SN Appl. Sci.*, 2019, **1**, 815, DOI: [10.1007/s42452-019-0835-3](https://doi.org/10.1007/s42452-019-0835-3).
- 73 D. F. Katowah, M. A. Hussein, M. M. Rahman, Q. A. Alsulami, M. M. Alam and A. M. Asiri, Fabrication of hybrid PVA-PVC/SnZnOx/SWCNTs nanocomposites as Sn²⁺ ionic probe for environmental safety, *Polym.-Plast. Technol. Eng.*, 2020, **59**, 642, DOI: [10.1080/25740881.2019.1673409](https://doi.org/10.1080/25740881.2019.1673409).
- 74 R. A. Qazi, R. Khattak, L. Ali Shah, R. Ullah, M. S. Khan, M. Sadiq, M. M. Hessien and Z. M. El-Bahy, Effect of MWCNTs Functionalization on Thermal, Electrical, and Ammonia-Sensing Properties of MWCNTs/PMMA and PHB/MWCNTs/PMMA Thin Films Nanocomposites, *Nanomaterials*, 2021, **11**, 2625, DOI: [10.3390/nano11102625](https://doi.org/10.3390/nano11102625).



- 75 D. Kumar and P. Jindal, Effect of multi-walled carbon nanotubes on thermal stability of polyurethane nanocomposites, *Mater. Res. Express*, 2019, **6**, 105336, DOI: [10.1088/2053-1591/ab3ad7](https://doi.org/10.1088/2053-1591/ab3ad7).
- 76 S. Ren, C. Li, X. Zhao, Z. Wu, S. Wang, G. Sun, Q. Xin and X. Yang, Surface modification of sulfonated poly(ether ether ketone) membranes using nafion solution for direct methanol fuel cells, *J. Membr. Sci.*, 2005, **247**, 59, DOI: [10.1016/j.memsci.2004.09.006](https://doi.org/10.1016/j.memsci.2004.09.006).
- 77 Z. Wang, G. Liu, L. Zhang and H. Wang, Electrochemical detection of trace cadmium in soil using a Nafion/stannum film-modified molecular wire carbon paste electrodes, *Ionics*, 2013, **19**, 1687, DOI: [10.1007/s11581-013-0891-4](https://doi.org/10.1007/s11581-013-0891-4).
- 78 S. Wu, N. C. Sekar, S. N. Tan, H. Xie and S. H. Ng, Determination of chromium (III) by differential pulse stripping voltammetry at a chitosan-gold nanocomposite modified screen printed electrode, *Anal. Methods*, 2016, **8**(5), 962–967.
- 79 T. Alizadeh, F. Raffei, N. Hamidi and M. R. Ganjali, A new electrochemical sensing platform for Cr (III) determination based on nano-structured Cr (III)-imprinted polymer-modified carbon composite electrode, *Electrochim. Acta*, 2017, **247**, 812–819.
- 80 A. Aravind, M. Sebastian and B. Mathew, Green synthesized unmodified silver nanoparticles as a multi-sensor for Cr (III) ions, *Environ. Sci.: Water Res. Technol.*, 2018, **4**(10), 1531–1542.
- 81 H. Filik and A. A. Avan, Dextran modified magnetic nanoparticles based solid phase extraction coupled with linear sweep voltammetry for the speciation of Cr(VI) and Cr(III) in tea, coffee, and mineral water samples, *Food Chem.*, 2019, **292**, 151–159.
- 82 G. Chen, J. Xie, Z. Zhang, W. Meng, C. Zhang, K. Kang, Y. Wu and Z. Guo, A portable digital-control electrochemical system with automatic ohmic drop compensation for fast scan voltammetry and its application to ultrasensitive detection of chromium (III), *Sens. Actuators, B*, 2019, **301**, 127135.
- 83 M. M. Hussain, A. M. Asiri, M. N. Arshad and M. M. Rahman, Synthesis, characterization, and crystal structure of (E)-N-(4-Bromobenzylidene)- benzenesulfonohydrazide and its application as a sensor of chromium ion detection from environmental samples, *J. Mol. Struct.*, 2020, **1207**, 127810.
- 84 Y. Gao, B. Feng, L. Miao, Y. Chen and J. Di, Determination of Cr (III) ions based on plasmonic sensing and anodic stripping voltammetry with amplification of Ag nanoparticles, *Microchem. J.*, 2020, **157**, 104995.
- 85 S. Y. Ejeta and T. Imae, Selective colorimetric and electrochemical detections of Cr(III) pollutant in water on 3-mercaptopropionic acid-functionalized gold plasmon nanoparticles, *Anal. Chim. Acta*, 2021, **1152**, 338272.
- 86 X. L. Cheng, H. Zhao, L. H. Huo, S. Gao and J. G. Zhao, ZnO nanoparticulate thin film: preparation, characterization and gas-sensing property, *Sens. Actuators, B*, 2004, **102**, 248, DOI: [10.1016/j.snb.2004.04.080](https://doi.org/10.1016/j.snb.2004.04.080).
- 87 J. K. Srivastava, P. Pandey, V. N. Mishra and R. Dwivedi, Structural and micro structural studies of PbO-doped SnO₂ sensor for detection of methanol, propanol and acetone, *J. Nat. Gas Chem.*, 2011, **20**, 179, DOI: [10.1016/S1003-9953\(10\)60168-5](https://doi.org/10.1016/S1003-9953(10)60168-5).
- 88 A. Masoumi, M. S. Gargari, G. Mahmoudi, B. Miroslaw, B. Therrien, M. Abedi and P. Hazendonk, Structural diversity in mercury(II) coordination complexes with asymmetrical hydrazone-based ligands derived from pyridine, *J. Mol. Struct.*, 2015, **1088**, 64, DOI: [10.1016/j.molstruc.2015.02.016](https://doi.org/10.1016/j.molstruc.2015.02.016).
- 89 N. B. Sonawane, K. V. Gurav, R. R. Ahire, J. H. Kim and B. R. Sankapal, CdS nanowires with PbS nanoparticles surface coating as room temperature liquefied petroleum gas sensor, *Sens. Actuators, A*, 2014, **216**, 78–83, DOI: [10.1016/j.sna.2014.05.012](https://doi.org/10.1016/j.sna.2014.05.012).
- 90 R. D. Ladhe, K. V. Gurav, S. M. Pawar, J. H. Kim and B. R. Sankapal, p-PEDOT:PSS as a heterojunction partner with n-ZnO for detection of LPG at room temperature, *J. Alloys Compd.*, 2012, **515**, 80–85, DOI: [10.1016/j.jallcom.2011.11.076](https://doi.org/10.1016/j.jallcom.2011.11.076).

