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# Activated carbon from *Camellia oleifera* shells for adsorption of Y(III): experimental and DFT studies

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Yttrium is an important rare earth element and is widely used in fields such as special glass preparation, metallurgy, and materials science. However, it is difficult to recover yttrium ion waste from dilute solutions with traditional processes, resulting in a significant waste of rare earth resources. The simple, effective, and easy-to-operate adsorption method is the most promising method for recovering yttrium, which is of great significance for sustainable development of the rare earth industry. In this study, activated carbon was prepared from *Camellia oleifera* fruit shells (COS) using phosphoric acid activation, and efficient recovery of Y(III) from the *Camellia oleifera* fruit shell activated carbon was studied. Adsorption equilibrium data showed that this activated carbon had a Y(III) adsorption capacity of 35.41 mg g<sup>-1</sup>, indicating significant potential for recovery of yttrium ions. The adsorption of Y(III) by the activated carbon prepared from COS was consistent with the Langmuir model, and the adsorption data were consistent with the pseudo second-order kinetic model, indicating that the adsorption process was primarily chemical adsorption. After adsorption, the surface of the activated carbon contained large amounts of N, O, and Y, indicating that Y(III) was stably adsorbed. The mechanisms for adsorption of Y(III) on three types of activated carbon were studied through DFT calculations. The results showed that Y(III) interacted with the carbon atoms on the surfaces to form new chemical bonds. The yttrium ion adsorption capacities for the three different activated carbons decreased in the order C I > C II > C.

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## 1. Introduction

The rare earth elements include the 15 lanthanide elements (from La to Lu) plus scandium (Sc) and yttrium (Y).<sup>1,2</sup> They play an important role in manufacturing advanced materials and are widely used in sensing materials, luminescent materials, nanomaterials, battery components,<sup>3</sup> permanent magnets,<sup>4</sup> and superconducting materials. Yttrium is an important rare earth element<sup>5</sup> and is used in fields such as special glass preparation,<sup>6</sup> metallurgy,<sup>7</sup> and materials science.<sup>8</sup> As a nonrenewable rare earth, the scale of yttrium use has continuously been reduced by extensive development and utilization.<sup>9</sup> According to statistics, mining wastewater (including coal mine drainage) contains a large amount of low concentration rare earth ions, which have been explored as an attractive secondary source for rare earth recovery.<sup>10</sup> However, it is difficult to recover rare earth waste from dilute solutions *via* traditional processes, resulting in a significant waste of rare earth resources.<sup>11</sup> Therefore, the

recovery of valuable rare earth elements such as yttrium from dilute rare earth wastewater solutions is of great significance for sustainable development of the rare earth industry.

Researchers have developed several technologies for recovering the rare earths from wastewater, including biological methods,<sup>12</sup> chemical precipitation,<sup>13</sup> solvent extraction,<sup>14,15</sup> membrane filtration, electrodialysis, and adsorption.<sup>16</sup> Due to the cost-effectiveness and environmental friendliness of adsorption materials, recycling and utilization of the rare earths in wastewater has attracted much attention.<sup>17,18</sup> Currently, the research on rare earth adsorption materials includes zeolites, clays,<sup>19,20</sup> resins, activated carbon, *etc.*<sup>21</sup> Due to the low concentrations of rare earths in wastewater,<sup>22</sup> the ion adsorption materials used in recovering them must have high adsorption capacities, fast adsorption processes, high selectivities, and high enrichment efficiencies. It is difficult to meet these requirements with the currently used adsorption materials.

*Camellia oleifera* is widely planted and is one of the four major woody oil tree species in the world.<sup>23,24</sup> The yield of *Camellia oleifera* fruit husks is high, but the utilization rate is low.<sup>25</sup> According to statistics, over 3.0 × 10<sup>6</sup> tons of COS are discarded or burned per year,<sup>26</sup> which constitutes a waste of resources and environmental pollution.<sup>27,28</sup> The main components of COS are lignin, hemicellulose, and cellulose,<sup>29</sup> which

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are excellent precursor materials for carbon production.<sup>30</sup> Research has shown that the carbon prepared from COS adsorbs toxic chemicals<sup>31,32</sup> and harmful heavy metal ions such as lead, cadmium, and copper in aqueous solutions.<sup>33</sup> Therefore, the shells of *Camellia oleifera* have become a promising alternative raw material for producing activated carbon due to the high yield and low cost. Research on preparing high-performance activated carbon from COS to adsorb low concentrations of rare earth ions. Not only can a feasible treatment path be found for COS, but rare earth ions can also be recovered and environmentally friendly, achieving the goal of "treating waste with waste".

Therefore, this study used COS as the precursor raw material and optimized the phosphoric acid activation method to prepare high-performance COS activated carbon exhibiting a high recovery rate for the rare earth Y(III). This provides a theoretical reference for the application of *Camellia oleifera* shell activated carbon for adsorption of rare earths in wastewater.

## 2. Experimental

### 2.1. Materials and reagents

The shells of *Camellia oleifera* were used for the preparation of activated carbon and were collected from Ganzhou, Jiangxi, China. After washing with distilled water, the shells were dried at 105 °C for 24 h. Phosphoric acid was purchased from China National Pharmaceutical Group Chemical Reagent Co., Ltd. (Shanghai, China). All chemical reagents used in this study were of analytical reagent grade and were used without further purification. The water used in the experiment was deionized water.

### 2.2. Synthesis of activated carbon from COS

In the process for preparation of activated carbon for *Camellia oleifera* fruit husks, (1) the dried *Camellia oleifera* fruit husks were crushed with a stainless steel pulverizer to provide the required particle sizes; (2) five g of *Camellia oleifera* shell powder with the desired particle sizes was weighted and mixed with a certain proportion of phosphoric acid and soaked for a period; (3) the impregnated sample was transferred to a quartz boat and placed in a tube furnace. First, nitrogen gas was introduced to remove the air, and then the sample was heated to the desired temperature for 1 h. (4) After cooling to room temperature, the activated carbon was removed from the *Camellia oleifera* shells, washed with deionized water until the pH was constant, filtered and dried, and the dried sample was screened to 100 mesh.

### 2.3. Characterization of the activated carbon from COS

X-ray photoelectron spectroscopy (Thermo Scientific K-Alpha, USA) with an Al K $\alpha$  source was used to detect the surface binding state and elemental morphology of the activated carbon. The surface morphologies and energy spectra of the raw materials and activated carbon were examined with a scanning electron microscope (ZEISS Sigma 300 from Germany). An

inductively coupled plasma emission spectrometer (ICP-OES, France) was used to detect the Y<sup>3+</sup> concentrations with the manufacturer's recommended instrumental parameters.

### 2.4. Adsorption experiments

Y(III) ion solutions were prepared with different concentrations of Y(NO<sub>3</sub>)<sub>3</sub>·6H<sub>2</sub>O for adsorption experiments. Ten milligrams of sample was added to give 25 mL of the Y(III) solution for the batch adsorption experiments. The adsorption experiments were conducted in a constant temperature oscillator operated at 150 rpm for different durations. After completion of the experiment, the volume was measured following filtration through a 0.45  $\mu$ m filter membrane, and the concentration of yttrium ions was determined. Duplicate samples were prepared for each experiment. All results were the average of duplicate samples, with relative errors of less than 5%.

### 2.5. Kinetic and isothermal models

The pseudo first-order kinetic model and pseudo second-order kinetic model were used for data fitting and analyses with eqn (1) and (2) to determine the rate-controlling steps for Y<sup>3+</sup> adsorption on carbon samples:<sup>34,35</sup>

$$\ln(q_e - q_t) = \ln q_e - k_1 t \quad (1)$$

$$\frac{t}{q_t} = \frac{1}{k_2 q_e^2} + \frac{1}{q_e} t \quad (2)$$

where  $q_e$  and  $q_t$  are the adsorption equilibrium and adsorption capacity of Y<sup>3+</sup> at time  $t$  (mg g<sup>-1</sup>), respectively,  $t$  is the contact time (min), and  $k_1$  and  $k_2$  are the pseudo first-order and pseudo second-order rate constants, respectively.

The Langmuir model and Freundlich model were used for fitting and analyses to study the interactions between the activated carbon and the adsorbate. These two models were used to describe the electroadsorption process and the maximum adsorption capacity:<sup>36</sup>

$$q_e = \frac{q_m K_L C_e}{1 + K_L C_e} \quad (3)$$

$$\ln q_e = \ln K_F + \frac{1}{n} \ln C_e \quad (4)$$

where  $q_e$  and  $q_m$  are the adsorption capacity and the maximum adsorption capacity (mg g<sup>-1</sup>), respectively, at adsorption equilibrium;  $C_e$  is the concentration of adsorbate in solution at equilibrium (mg L<sup>-1</sup>);  $K_L$  is the Langmuir constant (L mg<sup>-1</sup>); and  $K_F$  is the Freundlich constant,  $\frac{1}{n}$  is the Freundlich exponent.

## 3. Results and discussion

### 3.1. Effects of preparation conditions for the activated carbon from COS on Y(III) adsorption

The granularity, soaking time, soaking ratio (mL H<sub>3</sub>PO<sub>4</sub> per g COS) and activation temperature were the key parameters used in preparing the camellia shell-derived activated carbon, and they affected the Y(III) adsorption performance. To prepare



highly efficient activated carbon from the COS, we studied the effects of the preparation conditions on COS adsorption.

The effect of the particle size on the performance of the activated carbon is shown in Fig. 1(a). When the particle size was  $-0.106$  mm, the COS activated carbon exhibited the highest efficiency for adsorption of  $Y(III)$ , with an adsorption capacity of  $19.55 \text{ mg g}^{-1}$ . The particle sizes had a significant impact on the activated carbon from the COS, and as the particle sizes decreased, the efficiency of  $Y(III)$  adsorption increased. The reason for this may be that as the particle sizes decreased, more micropores were formed in activated carbon, which facilitated the adsorption of  $Y(III)$ . Therefore, the optimal particle sizes of  $-0.106$  mm was used to prepare the activated carbon from the COS.

To investigate the effect of impregnation time on the adsorption capacities for the rare earth ions, activated carbons were prepared with different impregnation times (0, 2, 3, 6, 12 h) while other parameters remained unchanged. As shown in Fig. 1(b), the  $Y(III)$  adsorption capacity was rapidly increased when the soaking time was increased from 1 h to 2 h, reaching its maximum value of  $19.55 \text{ mg g}^{-1}$  at approximately 2 h. When the time was increased further to 12 h, the  $Y(III)$  adsorption capacity decreased slowly. This may have occurred because hydrolysis of the lignin in the COS with phosphoric acid was enhanced as the soaking time increased, which formed more micropores in the activated carbon. However, longer time may have led to excessive erosion of the carbon body and

reduced the performance. Therefore, the optimal soaking time of 2 h was used to study the effect of the soaking ratio on the adsorption capacity of the activated carbon.

To investigate the effect of the impregnation ratio on the adsorption capacities of rare earth ions, activated carbon preparation experiments were conducted with different impregnation ratios while the other parameters remained unchanged. As shown in Fig. 1(c), when the impregnation ratio (mL phosphoric acid per g COS) was increased from 7.5:5 to 17.5:5, the efficiency for  $Y(III)$  adsorption on the activated carbon first increased and then decreased. When the impregnation ratio was increased from 7.5:5 to 12.5:5, the  $Y(III)$  adsorption capacity rapidly increased and reached the maximum of  $19.55 \text{ mg g}^{-1}$  at approximately 12.5:5. As time was increased further to 17.5:5, the capacity for  $Y(III)$  adsorption on the activated carbon decreased. Therefore, the optimal impregnation ratio was 12.5:5.

To investigate the effect of the activation temperature on the adsorption capacities of rare earth ions, experiments were conducted with activated carbons prepared at different activation temperatures ( $300^\circ\text{C}$ ,  $400^\circ\text{C}$ ,  $500^\circ\text{C}$ ,  $600^\circ\text{C}$ ,  $700^\circ\text{C}$ ), while other parameters remained unchanged. Fig. 1(d) shows that with increasing activation temperature, the  $Y(III)$  adsorption capacity increased. When the activation temperature was increased from  $300^\circ\text{C}$  to  $600^\circ\text{C}$ , the capacity for  $Y(III)$  adsorption increased from  $8.35 \text{ mg g}^{-1}$  to  $35.41 \text{ mg g}^{-1}$ . As the temperature was increased to  $700^\circ\text{C}$ , the  $Y(III)$  adsorption

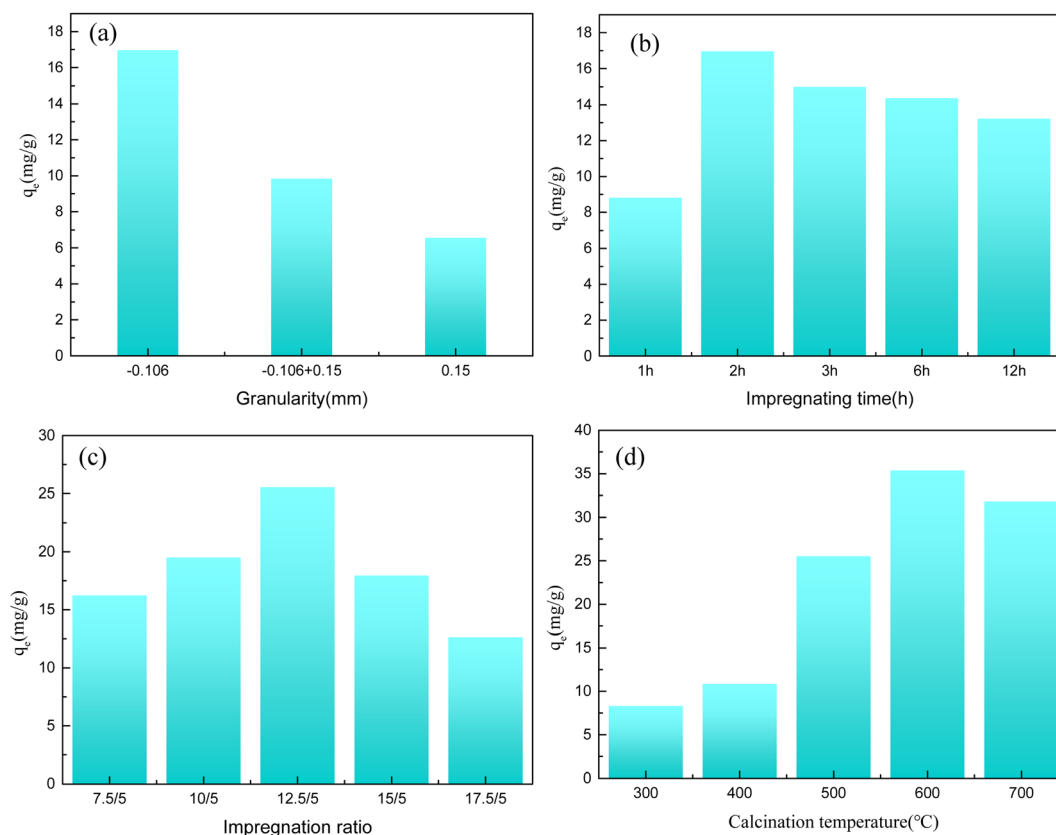


Fig. 1 Effects of (a) the particle size, (b) impregnation times, and (c) impregnation ratio, (d) activation temperature on  $Y(III)$  adsorption.



capacity decreased to  $31.87 \text{ mg g}^{-1}$ . Therefore, after considering the activation temperature carefully,  $600^\circ\text{C}$  was chosen.

### 3.2. Kinetics and isotherms

The activated carbon was prepared from the COS with a soaking time of 2 h, a soaking ratio of 12.5 : 5, and an activation temperature of  $600^\circ\text{C}$ . The amounts of  $\text{Y(III)}$  adsorbed on the activated carbon with different initial concentration conditions were studied to explore the adsorption performance and mechanism. The capacity for  $\text{Y(III)}$  adsorption on the activated carbons with different contact times was studied to determine the main steps controlling the adsorption process.

The effect of contact time on the adsorption performance of the activated carbon prepared from the COS is shown in Fig. 2(a). With increasing time, the adsorption capacity of activated carbon from the COS for  $\text{Y(III)}$  rapidly increased, reaching 90% of the maximum adsorption capacity within 30 minutes. As the time was increased further, the growth rate slowed and equilibrium was reached after 120 minutes, at which point the adsorption capacity for  $\text{Y(III)}$  was  $35.41 \text{ mg g}^{-1}$ . The effect of different rare earth ion initial concentrations on adsorption by the activated carbon from the COS was also studied. As shown in Fig. 2(b), as the initial concentration of  $\text{Y(III)}$  was increased from  $30 \text{ mg L}^{-1}$  to  $200 \text{ mg L}^{-1}$ , the  $\text{Y(III)}$  adsorption capacity increased from  $22.39 \text{ mg g}^{-1}$  to  $35.41 \text{ mg g}^{-1}$ . As the initial concentration was increased to  $200 \text{ mg L}^{-1}$ , the growth rate for the adsorption capacity slowed, and adsorption gradually approached equilibrium.

Nonlinear fitting was applied to the adsorption data with pseudo first-order and pseudo second-order kinetic models to explore the  $\text{Y(III)}$  adsorption on the activated carbon from the COS. The resulting fitting curves were drawn, as shown in Fig. 3(a) and (b). The adsorption data were fitted with the Langmuir model and the Freundlich model to evaluate the mechanism for adsorption of  $\text{Y(III)}$  on the activated carbon. The fits for the Langmuir model and Freundlich model are shown in Fig. 3(c) and (d).

Fig. 3(a) and (b) shows that, compared with the pseudo first-order kinetic model, the pseudo second-order kinetic model provided a better fit with the adsorption data, and the

correlation coefficient was greater than 0.99. Therefore, it can be concluded that the adsorption of  $\text{Y(III)}$  on the activated carbon from *Camellia oleifera* fruit shells primarily involved chemical adsorption. Fig. 3(c) and (d) shows that the correlation coefficient for the Langmuir model was 0.99, which was higher than the correlation coefficient of the Freundlich model, 0.93. This indicated that the adsorption of  $\text{Y(III)}$  on the activated carbon was more consistent with the Langmuir model. Therefore, the adsorption of  $\text{Y(III)}$  by the activated carbon from the COS primarily involved monolayer adsorption.

### 3.3. XPS analyses of the activated carbon from the COS

XPS analyses of the activated carbons prepared at different temperatures were used to understand the effect of temperature on the amounts of oxygen and nitrogen functional groups present on the surface of the material.

Fig. 4 show that the temperature significantly affected the surface oxygen, phosphorus, and nitrogen functional groups present on the activated carbon from the COS. As the temperature was increased with the range  $400^\circ\text{C}$  to  $600^\circ\text{C}$ , the N 1s peak for the activated carbon increased significantly, while the O 1s peak was significantly weakened. Table 1 show that when the temperature was increased from  $400^\circ\text{C}$  to  $500^\circ\text{C}$ , the N content increased from 1.06% to 1.23%, and the O content decreased from 10.93% to 7.75%, indicating the presence of P. When the temperature was further increased to  $500^\circ\text{C}$ , the N content increased to 1.69%, and the O content decreased to 6.52%. At this time, the proportion of P was 0.6%. This indicated that the increased temperatures increased the proportions of N-containing functional groups and P-containing functional groups on the surface of activated carbon but reduced the proportion of O-containing functional groups, which improved the rare earth ion adsorption capacities of the activated carbon from the COS.

### 3.4. SEM-EDS analyses of $\text{Y(III)}$ adsorption on activated carbon from the COS

SEM-EDS was used to analyse the surface morphology and elemental distribution of the activated carbon after adsorption of  $\text{Y(III)}$  to determine whether the activated carbon prepared

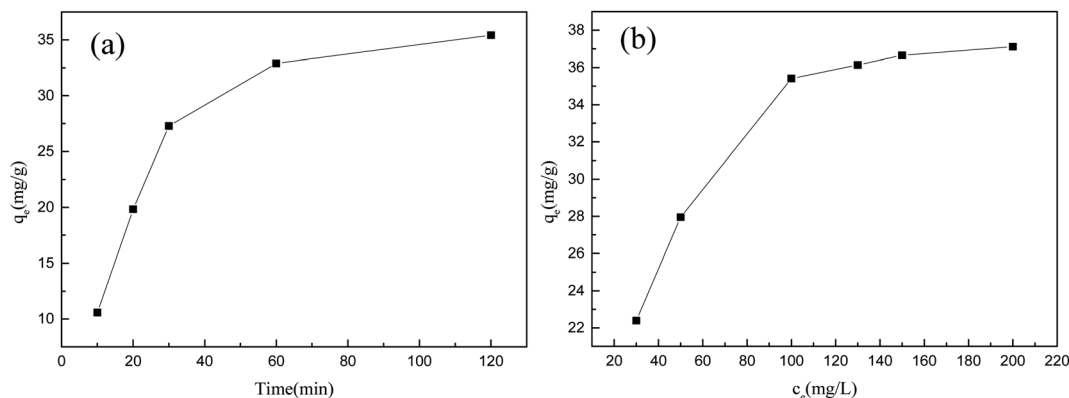


Fig. 2 (a) Kinetics and (b) isotherms for  $\text{La}^{3+}$  adsorption on activated carbon from the COS.

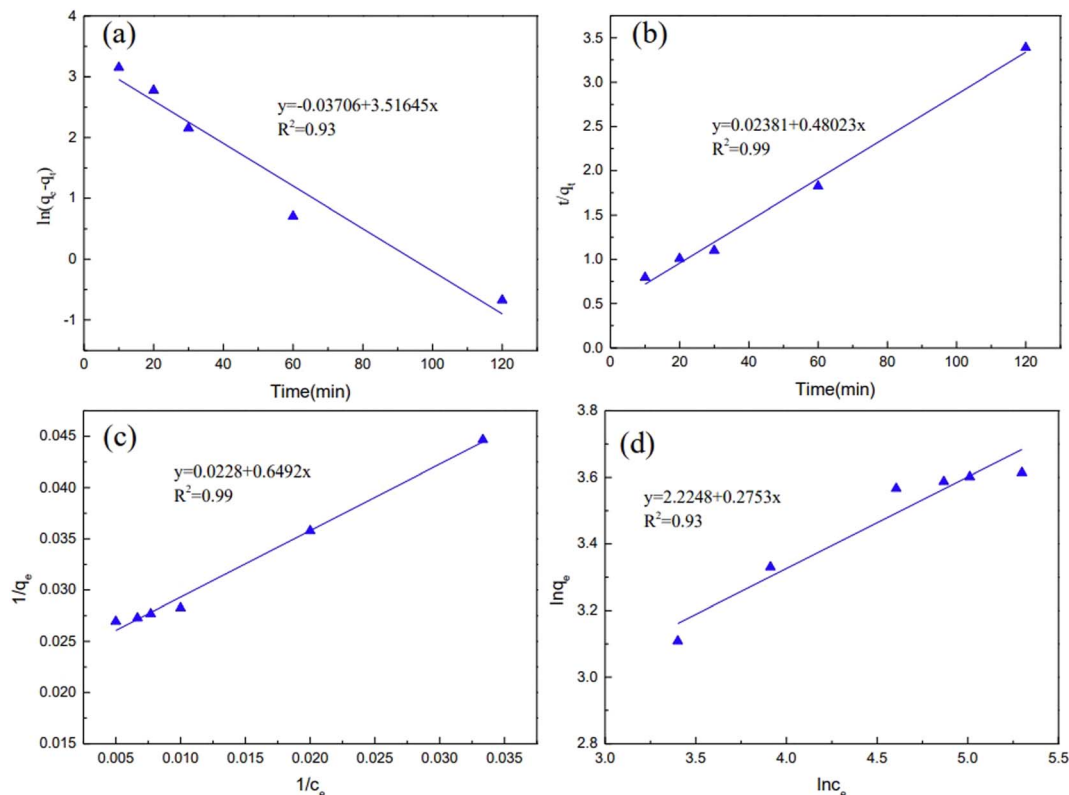


Fig. 3 Kinetics and isotherm fitting curves ((a): pseudo first-order dynamic model; (b): pseudo second-order dynamic model; (c): Langmuir model; (d): Freundlich model).

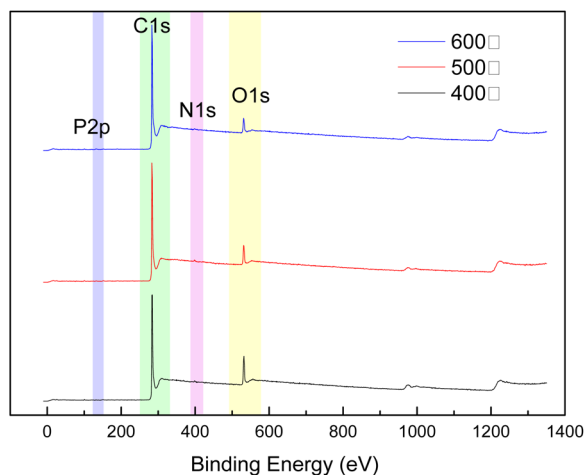


Fig. 4 XPS analyses of activated carbons prepared at different temperatures.

Table 1 Proportions of C, N, P, and O seen at different temperatures

Temperature	C (%)	N (%)	P (%)	O (%)
400 °C	88.01	1.06	0	10.93
500 °C	91.02	1.23	0	7.75
600 °C	91.19	1.69	0.60	6.52

from the COS captured Y(III) efficiently. Fig. 5 shows that the activated carbon had a porous structure; the surface contained large amounts of C, N, O, and P, and yttrium was also detected.

### 3.5. Adsorption capacity

To evaluate the practical application value of COS for Y(III) recovery, Table 2 summarizes the maximum adsorption capacity (reported in  $\text{mg g}^{-1}$ ) of other adsorbents for Y(III). Compared with other adsorbents reported in the literature, COS have higher adsorption performance and great potential in Y(III) recovery.

### 3.6. DFT study on the adsorption of Y(III) by the activated carbon from the COS

Activated carbon comprises an accumulation of irregular and small graphene crystals. Magnetic resonance imaging (MRI) showed that the small graphene crystals were composed of several aromatic ring clustered together. Therefore, a single-layer graphene structure is frequently used to model carbon materials in quantum chemistry,<sup>42</sup> and good results have been obtained in those calculations.<sup>43</sup> Chen and Yang<sup>44</sup> found that graphene models composed of six or more benzene rings effectively simulated carbon-based materials. Because the activated carbon was irregularly formed,<sup>45</sup> two main surface defect structures were used herein to approximate the true structure of the activated carbon from the COS, and the optimized structure is shown in Fig. 6.





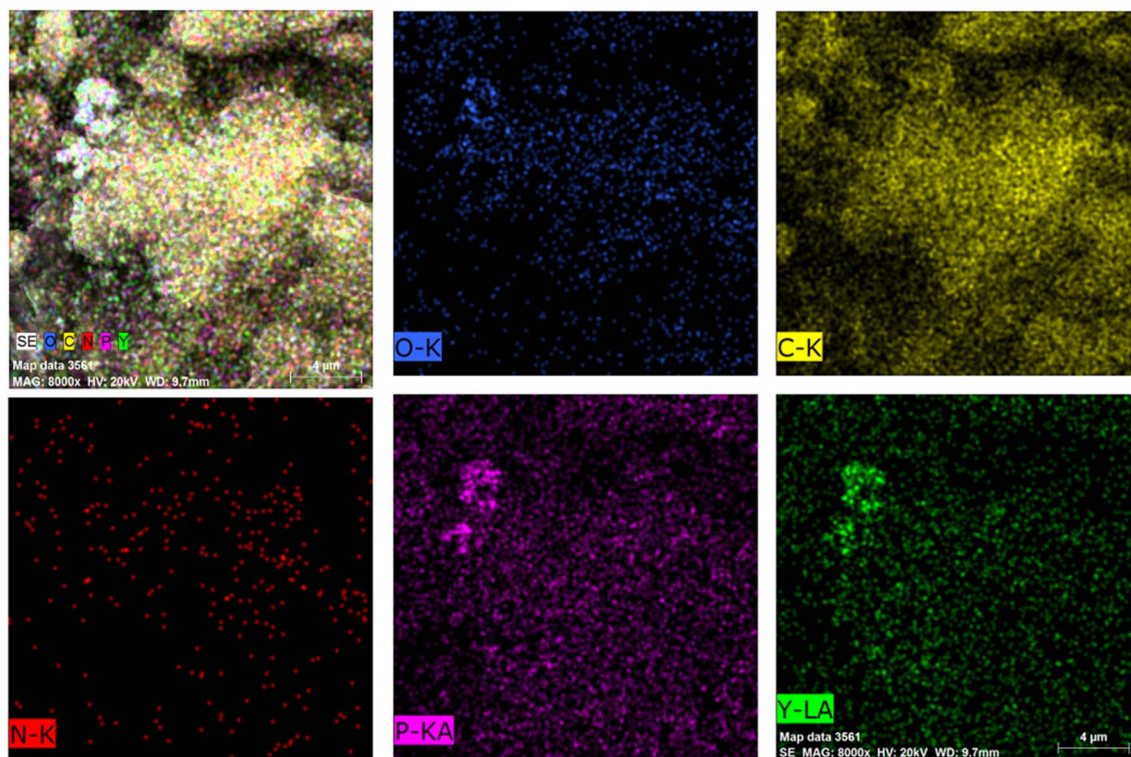


Fig. 5 Adsorption of Y(III) on the activated carbon from the COS.

Table 2 Adsorption capacity of different adsorbents for Y(III)

Sorbent	Max adsorption capacity (mg g <sup>-1</sup> )	Ref.
Cobalt iron oxide alginate nanocomposite	78.2	37
Fe <sub>3</sub> O <sub>4</sub> @SiO <sub>2</sub> @polyaniline-graphene oxide	8.10	38
Graphene oxide-tris(4-aminophenyl) amine composites nanoparticles	10.52	39
Multiwalled carbon nanotubes	7.75	40
Fe <sub>3</sub> O <sub>4</sub> @TiO <sub>2</sub> @P <sub>2</sub> O <sub>4</sub> nanoparticles	6.1	41
Nanometer-sized TiO <sub>2</sub>	35.41	This study
COS		

**3.6.1 Adsorption of Y(III) on the C-type activated carbon structure.** Y(III) was placed on the adsorption sites of the C-type activated carbon surface for the calculations, and the results are shown in Fig. 7. After relaxation, the Y(III) ion was significantly closer to the surface of the activated carbon. Before adsorption, the distance between the Y(III) and the C atoms was outside the bonding range. After adsorption and stabilization, the farthest distance between the Y(III) and C4 was 2.7614 Å; the shortest distance, from C2, was 2.70666 Å. The calculations indicated that new chemical bonds were formed between the Y(III) ions and C1, C2, C3, C4, C5, and C7 on the surface of activated carbon. Therefore, the Y(III) ions were stably adsorbed on the carbon surface with the C-type structure.

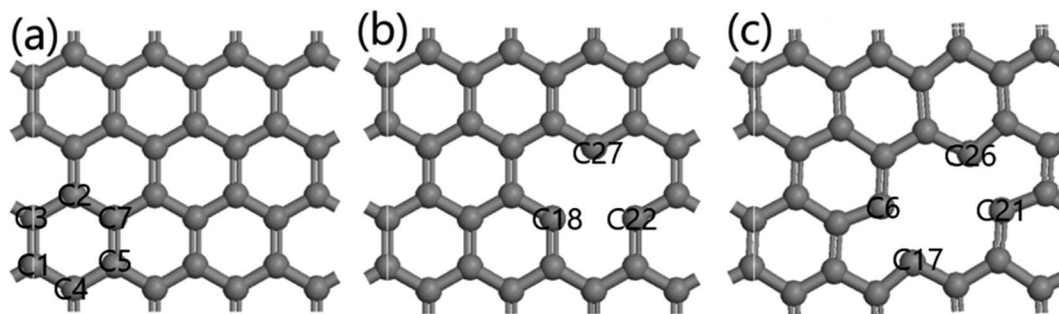


Fig. 6 Structure of the activated carbon prepared from the COS ((a) is the defect-free structure marked as C; (b and c) are the defect-activated carbon structures marked as C I and C II).

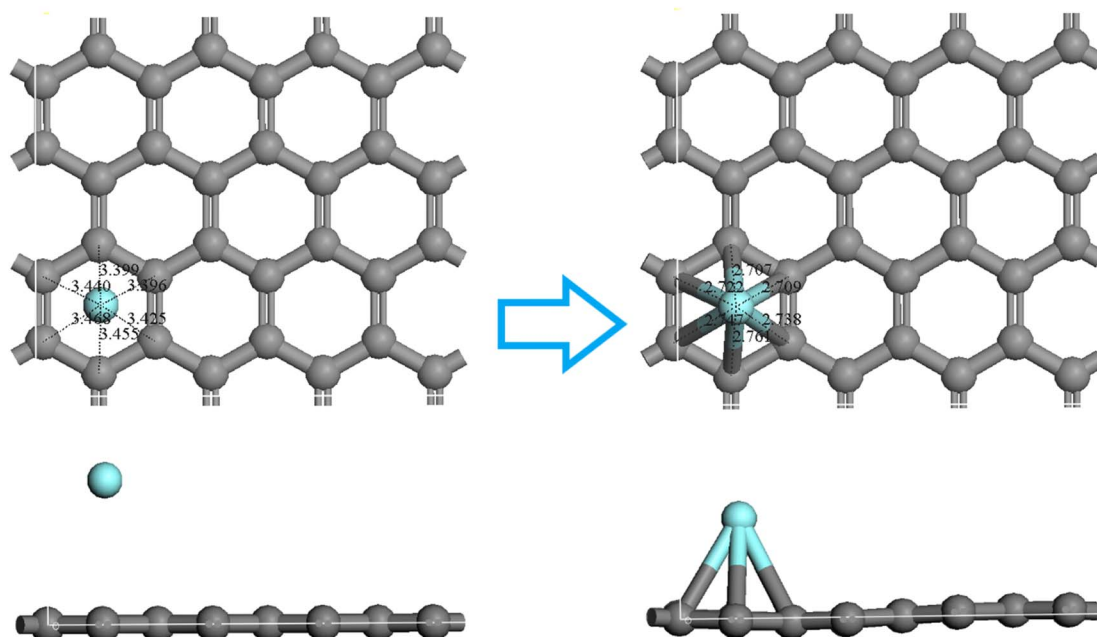


Fig. 7 Model for Y(III) before and after adsorption on the C-type surface.

To understand the mechanism for interaction between the Y(III) and the surface of the activated carbon further, we determined the density of states for the most stable adsorption configuration. Fig. 8 shows the partial wave density of states for C and Y after adsorption of the rare earth ion Y on the carbon

surface. The 5 s orbitals of the Y ions were distributed within the range  $-1$  to  $7.5$  eV at the bottom of the valence band and exhibited sharp peaks within the range  $-1$  to  $1.5$  eV. The 4p orbitals of the Y ions were distributed within the range  $1$  to  $10$  eV at the bottom of the valence band and exhibited sharp

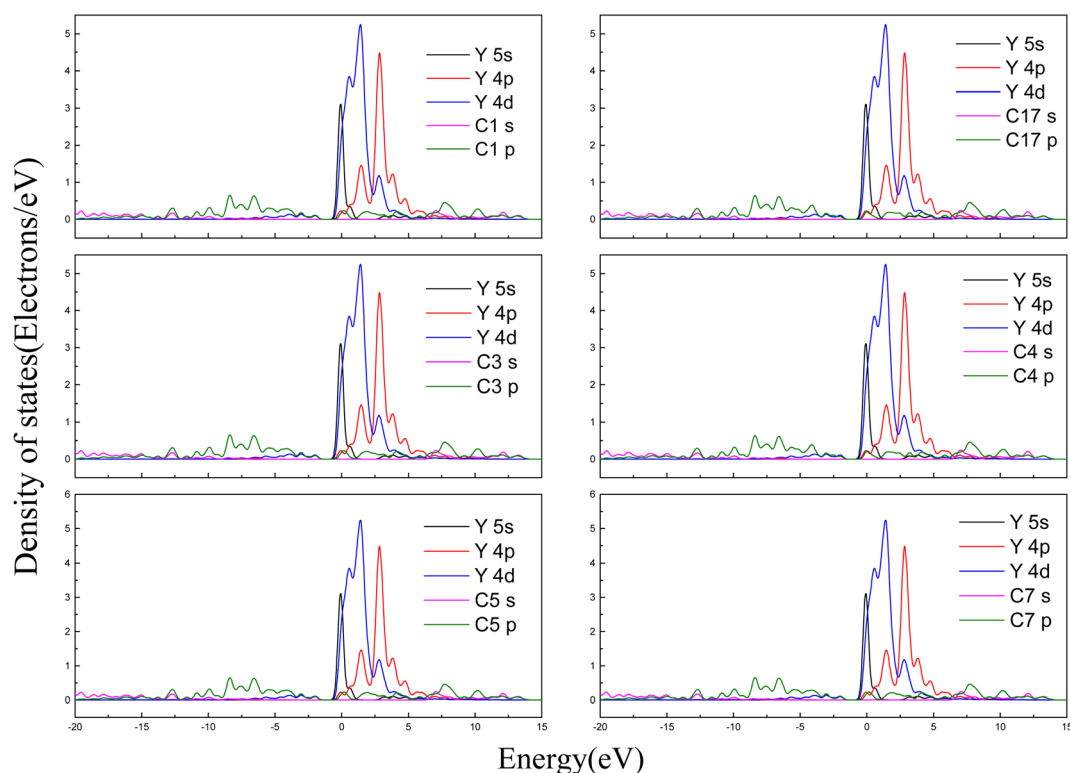


Fig. 8 Partial wave density of states for Y(III) after adsorption on a C-type surface.



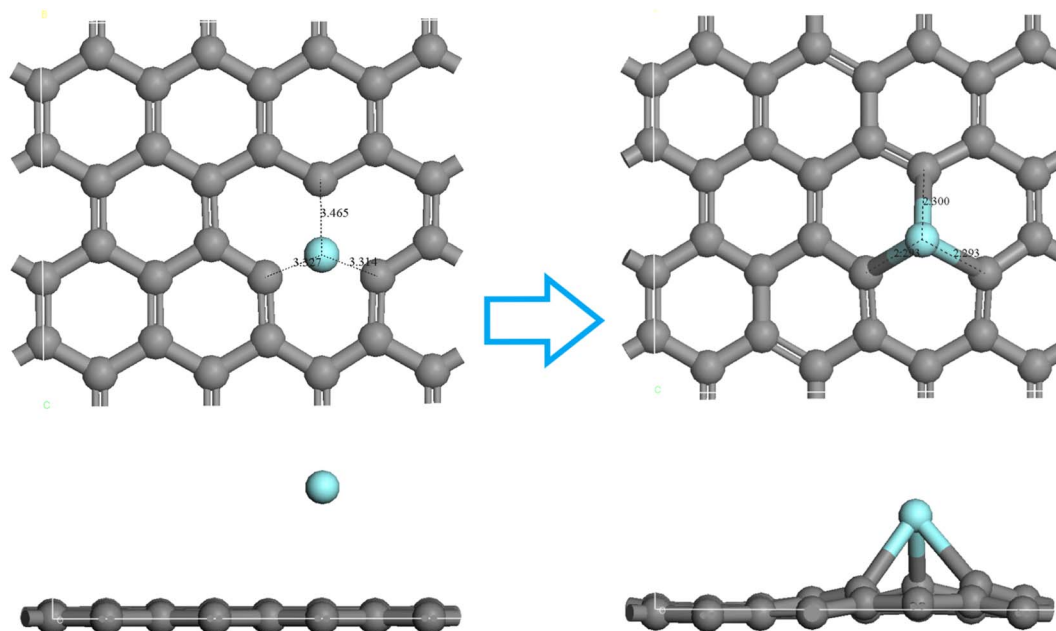


Fig. 9 Model for Y(III) before and after adsorption on the C I surface.

peaks within the ranges 2.5–3.5 eV and 1–2.5 eV. The 4d orbitals of the Y ions were distributed within the range –7.5 to 5 eV at the bottom of the valence band, and their highest peaks were within the range 1–2.5 eV and sharp peaks were found at –1 to 1 eV and 2.5–5 eV. The p orbitals of the C atoms were distributed within the range –15 to 17.5 eV, with multiple weak peaks present in the range –1 to –0.5 eV. The p orbitals of the C atom overlapped slightly with the 4d orbitals of the Y ions in the range –3 to –1.5 eV. Based on a comprehensive analysis of the bond populations results, the mechanism for adsorption of Y(III) on the C surface involved weak hybridization between the C atom p orbital and the Y ion 4d orbital on the surface.

**3.6.2 Adsorption of Y(III) on the C I-type structure of the activated carbon.** The adsorption calculation was performed by placing Y(III) at an adsorption site on the surface of the C I-type defect structure, and the results are shown in Fig. 9. After relaxation, the Y(III) was significantly closer to the surface of the activated carbon. Before adsorption, the distances between Y(III) and C18, C22, and C27 were 3.327 Å, 3.465 Å, and 3.334 Å, respectively, which were outside the bonding range. After adsorption stabilization, the distance between Y(III) and C18 changed from 3.327 Å to 2.495 Å, which reduced the distances from C22 and C27 to 2.293 Å and 2.29963 Å. The calculated adsorption structure indicated that the Y(III) on the surface of the activated carbon underwent bonding with C18, C22, and C27. Therefore, the rare earth Y(III) ion was stably adsorbed on the C surface with a type C I defect structure.

To further understand the interaction mechanism between Y(III) and the surface of the activated carbon, we calculated the density of states for the most stable adsorption configuration, as shown in Fig. 10. The 5s orbitals of the Y ions were distributed within the range –5 to 5 eV at the bottom of the valence band and exhibited sharp peaks at –3 to 2 eV, –0.5 to 0.5 eV,

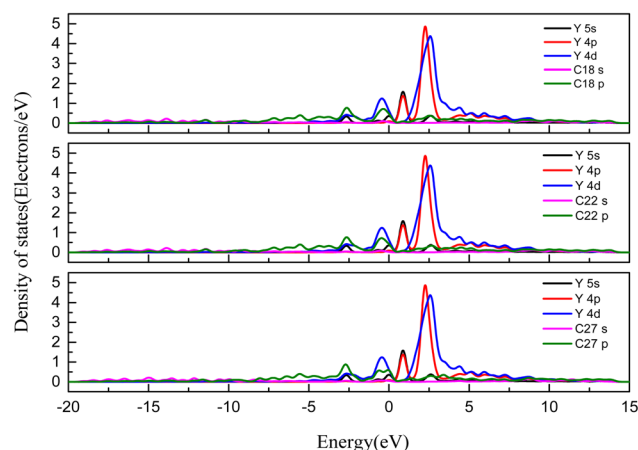


Fig. 10 Partial wave density of states for Y(III) adsorbed on the C I-type surface.

and 2–3 eV. The 4p orbitals of the Y ions were distributed in the range 0–10 eV at the bottom of the valence band and exhibited sharp peaks in the ranges 0–2 eV and 2–3 eV. The 4d orbitals of the Y ions were distributed within the range –5 to 10 eV at the bottom of the valence band, with sharp peaks at –3 to 2 eV, –2 to 1 eV, and –1.5 to 3 eV. The p orbitals of the C atoms exhibited sharp peaks within the ranges –2 to –1.5 eV, –1 to –0.5 eV, –7.5 to 10 eV, and 2–4 eV. The p orbitals of the C atom and the 4d orbitals of the Y ion overlapped significantly within the ranges –2 to 1.5 eV and –1 to 0.5 eV, and the 5s orbitals of the Y ions overlapped significantly within the ranges –3 to 2 eV, –0.5 to 0.5 eV, and 2–3 eV, resulting in significant hybridization. In summary, the mechanism for Y(III) adsorption on the surface of activated carbon involved hybridization of the C atom p orbitals





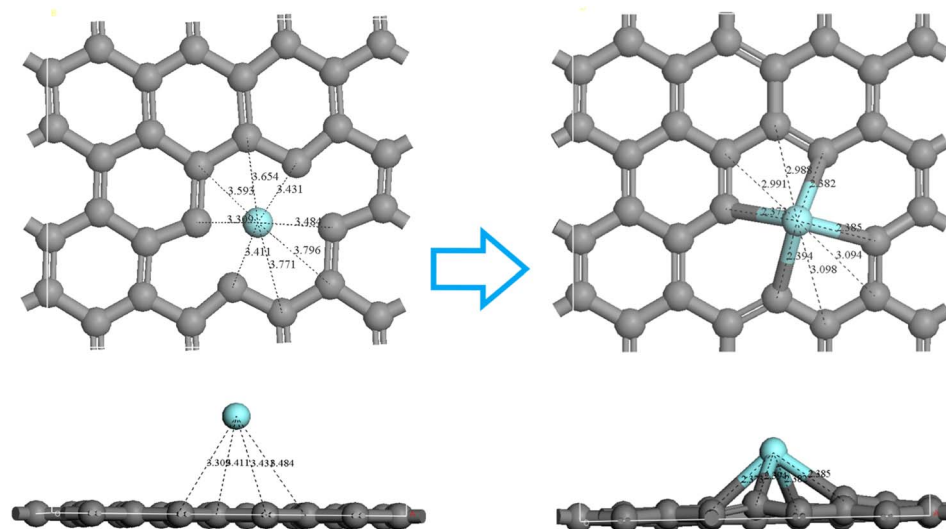


Fig. 11 Models for Y(III) before and after adsorption on the C II surface.

and the Y ion 5s and 4d orbitals on the surface, thereby forming stable chemical bonds.

**3.6.3 Adsorption of Y(III) on the C II-type activated carbon structure.** A Y(III) ion was placed at an adsorption site on the surface of the C II defect structure for to calculate the adsorption energy. Fig. 11 shows the changes of the Y(III) occurring after adsorption on the surface of the activated carbon: after relaxation, the Y(III) ion was significantly closer to the surface of the activated carbon. Before adsorption, the distances between Y(III) and C6, C17, C21, and C26 were 2.589 Å, 3.568 Å, 2.38478 Å, and 2.38231 Å, respectively, which were outside the bonding range. After adsorption and stabilization, the distances between

Y(III) and C6, C17, C21, and C26 decreased to 2.37282 Å, 2.39354 Å, 2.38478 Å, and 2.38231 Å, respectively. The calculation of the surface indicated that the Y(III) on the activated carbon underwent bonding with C6, C17, C21, and C26. Therefore, the rare earth Y(III) was stably adsorbed on the C surface with the type II defect structure.

To understand the mechanism for the interaction between the Y(III) and the surface of the activated carbon, we calculated the density of states for the most stable adsorption configuration. Fig. 12 shows the partial wave density of states for C and Y after adsorption of the Y(III) on the carbon surface. The 5s orbitals of Y ions were distributed within the range  $-3$  to  $10$  eV

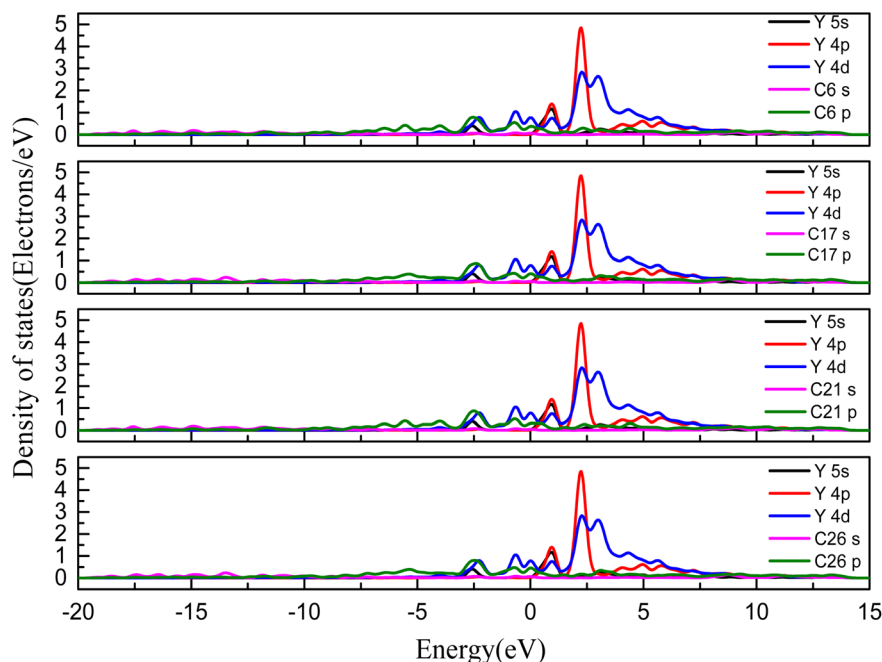


Fig. 12 Partial wave density of states for Y(III) adsorbed on the C II surface.



**Table 3** Energies for adsorption of Y(III) on different activated carbon surfaces (units: eV)

Surface	Adsorbate	$E_{\text{adsorbate/slab}}$	$E_{\text{adsorbate}}$	$E_{\text{slab}}$	$E_{\text{ads}}$
C	Y	−5152.9	−4963.27	−188.51	−1.12
C I	Y	−4994.22	−4800.04	−188.51	−5.67
C II	Y	−4838.42	−4644.72	−188.51	−5.19

at the bottom of the valence band, with sharp peaks at −3 to 2 eV and 0–2 eV and weak peaks at 3–4 eV and 8–9 eV. The 4p orbitals of the Y ion were distributed within the range 0–10 eV at the bottom of the valence band and showed the highest peak within the range of 2–3 eV, followed by a subpeak within the range 0.2–2 eV and a long and wide peak within the range 3–7.5 eV. The 4d orbitals of Y were distributed within the range 3 to 9 eV at the bottom of the valence band, and there were three strong peaks within the range 2 to 4 eV, appearing as peaks at −3 to 2 eV, −1 to −0.2 eV, and 0.5 to 1.8 eV. The p orbitals of the C atoms were distributed within the range −8 to 10 eV, exhibiting sharp peaks within the range −2.8 to −2 eV, multiple peaks within the range −1.8 to 0.5 eV, and weak, broad peaks within the range 2–5 eV and 2–4 eV. The p orbitals of the C atoms overlapped significantly with the 4d orbitals of the Y ion within the ranges −2.8 to −2 eV and −1.8 to −0.5 eV, and the 5s orbitals of the Y ion overlapped significantly within the ranges −2.8 to −2.3 eV and −0.3 to 1.8 eV. The 4p orbitals of the Y ion overlapped significantly within the range −0.3 to 1.8 eV, indicating significant hybridization between the p orbitals of the C atom and the 5s, 4p, and 4d orbitals of the Y ion. In summary, the mechanism for Y(III) adsorption on the surface of the activated carbon involved hybridization of the C p orbitals on the surface with the 5s and 4d orbitals of Y ions, thereby forming stable chemical bonds.

**3.6.4 Analysis of the Y(III) adsorption energy on different activated carbon surfaces.** Adsorption energy calculations were conducted to determine the strengths of Y(III) adsorption on the different activated carbon surfaces, and the results are shown in Table 3. The energies for adsorption of Y(III) on the C, C I, and C II surfaces were −1.12 eV, −5.67 eV, and −5.19 eV, respectively, indicating that Y(III) was stably adsorbed on the surfaces of the activated carbons, mainly through chemical adsorption. The Y(III) adsorption stabilities for the three activated carbon structures (C, C I, and C II) decreased in the order C I > C II > C.

## 4. Conclusions

This study used *Camellia oleifera* shells as the precursor and prepared activated carbon *via* phosphoric acid activation. The effects of the particle sizes, impregnation times, impregnation ratios, and activation temperatures on the adsorption of Y(III) were investigated. The activated carbon from the COS showed significant adsorption capacity and great potential for use in rare earth ion recovery. The kinetic studies indicated that adsorption occurred *via* monolayer adsorption and chemical adsorption. The SEM-EDS data indicated that Y(III) was stably adsorbed on the surface of the activated carbon from the COS.

The mechanism for adsorption of Y(III) on the surface of the activated carbon was studied through first principles calculations, which showed that Y(III) interacted with carbon atoms on the surfaces of C, C I, and C II to form new chemical bonds. The calculated adsorption energies shows that the stabilities for Y(III) adsorption on the three different activated carbons decreased in the order C I > C II > C. This work provides an inexpensive and efficient adsorbent for environmentally friendly recovery of rare earth Y(III) waste from dilute solutions.

## Data availability

Data will be made available on request.

## Author contributions

Bin Zeng: conceptualization, methodology, investigation, data curation, writing – original draft. Xiangrong Zeng: conceptualization, methodology, investigation, writing – original draft, writing—review and editing. Lianghui Hu: investigation, data curation. Lijinhong Huang: investigation, data curation, writing—review and editing, funding acquisition. Yuxiang Huang: investigation, data curation. Yi Zhou: investigation, data curation. Guoliang Liu: investigation, data curation. Wanfu Huang: data curation, conceptualization, supervision, funding acquisition.

## Conflicts of interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

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