




Cite this: *RSC Adv.*, 2024, 14, 3782

Adsorption of molecular hydrogen on $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl: theoretical insights for catalysis, hydrogen storage, gas separation, sensing, and environmental applications

Waqas Amber Gill, ^a Norah Alhokbany^b
and Muhammad Ramzan Saeed Ashraf Janjua ^{*c}

Employing a combination of Density Functional Theory (DFT) calculations and Molecular Dynamics (MD) simulations, the adsorption of molecular hydrogen (H_2) on $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl, a prominent silicate mineral, has been studied. The crystal structure of beryl, which consists of interconnected tetrahedral and octahedral sites, provides a fascinating framework for comprehending H_2 adsorption behavior. Initial investigation of the interaction between H_2 molecules and the beryl surface employed DFT calculations. We identified favorable adsorption sites and gained insight into the binding mechanism through extensive structural optimizations and energy calculations. H_2 molecules preferentially adsorb on the exposed oxygen atoms surrounding the octahedral sites, producing weak van der Waals interactions with the beryl surface, according to our findings. To further investigate the dynamic aspects of H_2 adsorption, MD simulations employing a suitable force field were conducted. To precisely represent interatomic interactions within the $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl- H_2 system, the force field parameters were meticulously parameterized. By subjecting the system to a variety of temperatures, we were able to obtain valuable information about the stability, diffusion, and desorption kinetics of H_2 molecules within the beryl structure. The comprehensive understanding of the H_2 adsorption phenomenon on $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl is provided by the combined DFT and MD investigations. The results elucidate the mechanisms underlying H_2 binding, highlighting the role of surface oxygen atoms and the effect of temperature on H_2 dynamics. This research contributes to a fundamental understanding of hydrogen storage and release in beryllium-based silicates and provides valuable guidance for the design and optimization of materials for hydrogen storage, catalysis, gas separation, sensing and environmental applications.

Received 2nd November 2023
Accepted 12th January 2024

DOI: 10.1039/d3ra07480c

rsc.li/rsc-advances

1. Introduction

The efficient storage and release of molecular hydrogen (H_2) is of utmost importance in various fields, including renewable energy and transportation sectors.^{1,2} Silicate minerals, with their abundant and diverse structures, have emerged as promising candidates for hydrogen storage materials due to their potential for high volumetric storage capacity.^{3,4} In particular, $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl, a prominent silicate mineral, exhibits a unique crystal structure comprising interconnected tetrahedral and octahedral sites, offering an intriguing framework for exploring hydrogen adsorption behavior.⁵

In this research paper, we investigate the adsorption behavior of H_2 over $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl using a combination of density functional theory (DFT) calculations, molecular dynamics (MD) simulations, and force field analysis.⁶ DFT, a quantum mechanical method, provides a powerful tool to understand the electronic structure, energetics, and adsorption properties of materials at the atomic scale.⁷ By employing DFT calculations, we aim to investigate the adsorption sites and energetics of H_2 molecules within the silicate's framework. Furthermore, DFT allows us to analyze the electronic properties of $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl, shedding light on the fundamental mechanisms governing H_2 adsorption.⁷ To complement the DFT results and capture the dynamical behavior of H_2 molecules, molecular dynamics (MD) simulations are conducted.⁸ MD simulations provide insights into the thermodynamic properties, diffusion kinetics, and overall stability of H_2 adsorption within the zeolite structure. By studying the time evolution of the system, we can observe the interactions between H_2 molecules and silicates, contributing to

^aDepartamento de Química Física, Universidad de Valencia, Avda Dr Moliner, 50, E-46100 Burjassot, Valencia, Spain

^bDepartment of Chemistry, King Saud University, Riyadh 11451, Saudi Arabia

^cDepartment of Chemistry, Government College University Faisalabad, Faisalabad 38000, Pakistan. E-mail: Janjua@gcu.edu.pk; Dr_Janjua2010@yahoo.com



a comprehensive understanding of the adsorption process.⁸ In addition to DFT and MD, we incorporate force field analysis to validate the accuracy and reliability of our simulation results.^{9,10} Force fields, empirical mathematical models, are commonly used to describe interatomic interactions within a system.¹¹ By comparing the results obtained from DFT calculations and MD simulations with those derived from the force field, we can assess the level of agreement and provide insights into the validity of our computational approach.^{9,10}

Previous studies have explored the adsorption of H₂ on silicate materials, providing valuable insights into the interaction mechanisms and storage capacities. For instance, a comprehensive review on hydrogen storage in beryllium-based silicates, highlighting the potential of these materials and the challenges associated with their practical implementation.¹² H₂ adsorption on different silicate surfaces, revealing the influence of surface chemistry and crystallographic orientation on the adsorption behavior investigated by DFT calculation.¹³ Hydrogen diffusion in silicate minerals, demonstrating the importance of temperature and structural features in governing the mobility of H₂ molecules by the utilization of MD simulation.¹⁴

We present a set of systematic potentials that characterize the interactions between molecular hydrogen and Be₃Al₂(SiO₃)₆-beryl using the Improved Lennard-Jones (ILJ) potential. These potentials strike a balance between accuracy and simplicity, making them suitable for molecular dynamics simulations. Molecular dynamics simulations will be used to compute second virial coefficients in order to compare the effectiveness of the proposed force fields. We can evaluate the effectiveness of the various force fields by comparing these results to experimental values reported in scientific literature. In addition, we will use precise techniques, including density functional theory (DFT), to predict the preferred interaction sites of gas molecules on Be₃Al₂(SiO₃)₆-beryl and determine the optimal orientation of molecular hydrogen.

In Section 2, we will describe the computational specifics used for this investigation. The focus of Section 3 will be the analysis of the derived potential energy surface, while Section 4 will describe the molecular dynamics simulation. In the concluding section, Section 5, the study's findings will be summarized.

2. Computational details

The widely used Lennard-Jones (LJ) potential is often employed to describe the adsorption of gases on silicates, graphene, metal organic frame works and zeolites.^{15,16} However, the LJ potential is known to have deficiencies at both short and long ranges.^{17–19} To overcome these limitations, Pirani *et al.* proposed an Improved Lennard-Jones (ILJ)^{17,18} potential that addresses the shortcomings of the LJ model. The ILJ potential introduces an additional parameter, enhancing its flexibility and accuracy. It has demonstrated excellent performance in capturing dispersion interactions in noble gas systems and has shown efficacy in high angular resolution and energy conditions, accurately reproducing vibrational springs. Through careful parameter

selection, the ILJ potential proves particularly valuable for molecular dynamics simulations, especially for non-covalent interactions.^{18,19} Its simplicity and accuracy make it a useful tool for accurately representing the adsorption behavior of gases on Be₃Al₂(SiO₃)₆-beryl. In many cases, electrostatic contributions to the interaction are incorporated using a coulombic expression, where partial charges are assigned to the system.^{20,21} However, the inclusion of these partial charges is often done without thorough consideration of the actual improvement in accuracy achieved by this term. It is important to note that the assigned partial charges lack true physical significance, as they are not quantum observables. Furthermore, there is often limited evidence to suggest that their inclusion significantly impacts the obtained results. The Improved Lennard-Jones (ILJ) interaction potentials utilized in this study were derived from first principles calculations performed at the dispersion-corrected density functional theory level. Specifically, the B97D functional, which incorporates an empirical dispersion term, was employed. The B97D functional is recognized as a robust approach among semi-empirical generalized gradient approximation (GGA) functionals, offering efficiency and precision in studying large systems that interact *via* dispersion forces.²² While the B97D functional has demonstrated its value in accurately describing non-covalent interactions,^{22–25} it is important to note that its performance may not be universally reliable for all systems.

In this study, we utilize the Improved Lennard-Jones (ILJ) potential to construct a series of force fields suitable for assessing van der Waals interactions, particularly within the Be₃Al₂(SiO₃)₆-beryl and molecular hydrogen system. Additionally, we explore the impact of the electrostatic component on the force field parameters by incorporating various charge schemes such as Hirshfeld,²⁶ CHelpG,²⁷ MBS (Minimum Basis Set),²⁸ MK (Merz-Kollman),²⁹ NBO (Natural Bonding Orbitals),³⁰ evaluating their effectiveness and feasibility.

To account for the fact that Be₃Al₂(SiO₃)₆-beryl is an infinite molecule, we employed truncated system for the single point energy calculations. This model was designed to be large enough to capture the essential interactions. Previous studies have demonstrated that Be₃Al₂(SiO₃)₆-beryl can serve as a suitable model for silicates in B97D calculations. All monomers were initially optimized at the B3LYP/6-31G** level.^{31,32} In Table 1, the distances of the cell matrix are 9.21 Å, −4.605 Å, 7.97609 Å and 9.17 Å, respectively, for Be₃Al₂(SiO₃)₆-beryl. The cell parameters are *A* = 9.21 Å with $\alpha = 90^\circ$, *B* = 9.21 Å with $\beta = 90^\circ$ and *C* = 9.17 Å with $\gamma = 120^\circ$ and with the cell volume 673.62660 Å³. Subsequently, these optimized structures were

Table 1 Crystallographic parameters and angles of the Be₃Al₂(SiO₃)₆-beryl unit cell, insight into silicate structure

	<i>x</i>	<i>y</i>	<i>z</i>	Parameters	Angles
<i>A</i>	9.21000	0.00000	0.00000	9.21000	$\alpha = 90^\circ$
<i>B</i>	−4.60500	7.97609	0.00000	9.21000	$\beta = 120^\circ$
<i>C</i>	0.00000	0.00000	9.17000	9.17000	$\gamma = 120^\circ$



treated as rigid in all subsequent interaction energy calculations.

Fig. 1 illustrates the representation of a crystal lattice in crystallography, where a unit cell is commonly utilized to describe the crystal structure. The unit cell serves as the smallest repeating unit that encompasses the entire crystal lattice. It is defined by three parameters: the lengths of its edges (A , B , and C) and the angles between them (α , β , γ). The parameters A , B , and C correspond to the lengths of the edges of the unit cell, determining the size of the unit cell in each direction. Typically, these values are expressed in angstroms (Å) or picometers (pm). On the other hand, α , β , and γ denote the angles between the edges of the unit cell, influencing the shape and symmetry of the unit cell. Specifically, α represents the angle between edges B and C , β denotes the angle between edges A and C , and γ indicates the angle between edges A and B . These angles are measured in degrees. $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl, a specific type of silicate, exhibits a distinctive framework structure composed of interconnected channels and cages. The unit cell parameters (A , B , and C) and angles (α , β , γ) describe the arrangement of atoms within the crystal lattice of $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl. By knowing the specific values of A , B , C , α , β , and γ , crucial information regarding the size, shape, and symmetry of the $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl crystal structure can be deduced. These parameters play a vital role in understanding the crystallographic properties and can be employed to unravel various physical and chemical characteristics associated with $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl.

The energy calculations for the molecular hydrogen over $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl system, were performed using Gaussian 09 (ref. 33) and the EMSL Basis Set Exchange. The calculations were based on the B97D/TZV2P^{22,31} level. This specific combination of the B97D functional and the split-valence triple-zeta basis set with two polarizations (TZV2P) has been previously

determined as adequate for accurately representing polarization effects in comparable systems.^{34,35}

A total of 50 random orientations³³ for the H_2 molecule were generated by fixing $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl. Each orientation was then meticulously scanned for intermolecular distances ranging from 2.5 to 12.0, with a focus on sampling near the equilibrium distance. There was a total of 25 distances considered, which encompassed all 50 relative orientations. Using the B97D/TZV2P method, the interaction energy for each of the 1250 conformations was then calculated.

3. Potential energy surface

It is common to divide a potential energy surface into distinct components, such as electrostatic and non-electrostatic forces, when constructing one. For the potential energy surface of H_2 over $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl, electrostatic forces result from coulombic interactions between charged particles within molecules, while non-electrostatic forces result from van der Waals interactions between neutral particles. The total potential energy can be expressed as the sum of electrostatic and non-electrostatic contributions if these components are treated separately and their separability is assumed. This assumption simplifies the construction procedure and facilitates the independent management of each contribution.

To facilitate efficient calculation of potential energy contributions, it is assumed that both parts of the potential energy surface can be represented as sums over A and B molecule centres. The interactions between two molecules can be described in a straightforward and computationally efficient manner by totaling the contributions from each atom or centre in both molecules. It is essential to observe, however, that these assumptions may not apply to all molecular systems, and that certain circumstances may necessitate the use of advanced methods.

$$\Delta E = \sum_{\substack{i \in A \\ j \in B}} \left(V_{ij}^{\text{ele}}(r_{ij}) + V_{ij}^{\text{non-ele}}(r_{ij}) \right) \quad (1)$$

There are different ways to describe the resulting interactions between molecules, which can include coulombic sums over assigned atomic charges. The coulombic sum is given by

$$E_{\text{elec}} = E_{\text{Coul}} = \sum_{A,B} \frac{q_A q_B}{R_{AB}} \quad (2)$$

In the potential energy surface, distinct contributions, such as electrostatic and non-electrostatic forces, can be identified. Coulombic interactions between charged particles within the molecules generate electrostatic forces, whereas van der Waals forces between inert particles are responsible for the non-electrostatic forces.

$$V_{\text{tot}}(R) = V_{\text{nelec}}(R) + V_{\text{elec}}(R) = V_{\text{ILJ}}(R) + V_{\text{Coul}}(R) \quad (3)$$

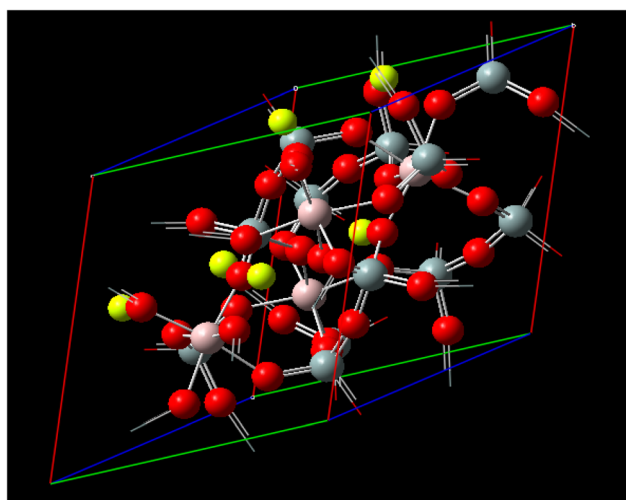


Fig. 1 Unit cell matrix of $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl are 9.21 Å, -4.605 Å, 7.97609 Å and 9.17 Å, respectively. The cell parameters are $A = 9.21$ Å with $\alpha = 90^\circ$, $B = 9.21$ Å with $\beta = 90^\circ$ and $C = 9.17$ Å with $\gamma = 120^\circ$ and with the cell volume 673.62660 Å³.



In molecular simulations, the ILJ potential is used as a model to characterize the non-electrostatic portion of the potential energy surface. In molecular simulations, non-electrostatic interactions are typically attributed to van der Waals forces caused by fluctuations in the electron cloud encircling atoms and molecules.

$$V(R) = \epsilon \left[\frac{m}{n(R) - m} \left(\frac{r_0}{R} \right)^{n(R)} - \frac{n(R)}{n(R) - m} \left(\frac{r_0}{R} \right)^m \right] \quad (4)$$

where

$$n(R) = \beta + 4 \left(\frac{R}{r_0} \right)^2 \quad (5)$$

ϵ is the well depth of the dissociation curve depicted by the ILJ potential, while r_0 is its position. On the other hand, β is a dimensionless parameter that adds extra flexibility.

In Table 2, the higher value for β_{AlH} is 7.893, indicating a stronger attractive interaction between aluminum (Al) and hydrogen (H). On the other hand, the lower value for β_{AlH} is 7.356. The higher value for β_{BeH} is 7.982, indicating a stronger attractive interaction between beryllium (Be) and hydrogen (H). The lower value for β_{BeH} is 7.521. The higher value for β_{SiH} is 8.853, indicating a stronger attractive interaction between silicon (Si) and hydrogen (H). The lower value for β_{SiH} is 8.264. The higher value for β_{OH} is 7.813, indicating a stronger attractive interaction between oxygen (O) and hydroxyl (OH) groups. It is worth noting that there is a discrepancy in the fifth value (7.637), which may or may not be considered an anomaly. The higher values of β_{AlH} , β_{BeH} , β_{SiH} , and β_{OH} represent stronger attractive interactions between the respective elements (Al, Be, Si, or O) and hydrogen (H) or hydroxyl (OH) groups. The lower values indicate relatively weaker attractive interactions.

The higher value for ϵ_{AlH} is 0.346 kJ mol⁻¹, indicating a deeper potential well for the interaction between aluminum (Al) and hydrogen (H). The lower value for ϵ_{AlH} is 0.127 kJ mol⁻¹. The higher value for ϵ_{BeH} is 0.625 kJ mol⁻¹, indicating a deeper potential well for the interaction between beryllium (Be) and

hydrogen (H). The lower value for ϵ_{BeH} is 0.269 kJ mol⁻¹. The higher value for ϵ_{SiH} is 0.853 kJ mol⁻¹, indicating a deeper potential well for the interaction between silicon (Si) and hydrogen (H). The lower value for ϵ_{SiH} is 0.162 kJ mol⁻¹. The higher value for ϵ_{OH} is 0.096 kJ mol⁻¹, indicating a deeper potential well for the interaction between oxygen (O) and hydroxyl (OH) groups. The lower value for ϵ_{OH} is 0.025 kJ mol⁻¹. Higher values of ϵ_{AlH} , ϵ_{BeH} , ϵ_{SiH} , and ϵ_{OH} represent deeper potential wells, indicating stronger attractive interactions between the respective elements (Al, Be, Si, or O) and hydrogen (H) or hydroxyl (OH) groups. Conversely, lower values indicate shallower potential wells and relatively weaker attractive interactions. These values help characterize the depth of the interaction potentials, which are important for understanding molecular interactions and behaviors.

The higher value for r_{AlH} is 3.462 Å, indicating a larger distance at which the potential between aluminum (Al) and hydrogen (H) reaches zero. The lower value for r_{AlH} is 3.364 Å. The higher value for r_{BeH} is 3.964 Å, indicating a larger distance at which the potential between beryllium (Be) and hydrogen (H) reaches zero. The lower value for r_{BeH} is 3.673 Å. The higher value for r_{SiH} is 3.918 Å, indicating a larger distance at which the potential between silicon (Si) and hydrogen (H) reaches zero. The lower value for r_{SiH} is 3.621 Å. The higher value for r_{OH} is 3.782 Å, indicating a larger distance at which the potential between oxygen (O) and hydroxyl (OH) groups reaches zero. The lower value for r_{OH} is 3.529 Å. Higher values of r_{AlH} , r_{BeH} , r_{SiH} , and r_{OH} represent larger distances at which the potential between the respective elements (Al, Be, Si, or O) and hydrogen (H) or hydroxyl (OH) groups reaches zero. Conversely, lower values indicate smaller distances. These values provide information about the range of the interactions and help understand the spatial extent over which the attractive forces are effective.

The Fig. 2, represent interaction energies and charges related to the ILJ potential and its fitting to a specific system. Interaction energies (pink) represent the calculated interaction energies between the H₂/Be₃Al₂(SiO₃)₆-beryl system and some other entities, likely obtained through the B97D/TZV2P method. The pink color indicates that these energies are derived from calculations. Red points represent averaged interaction energies obtained from random conformations of the H₂/Be₃Al₂(SiO₃)₆-beryl system. Each point represents an averaged energy value derived from multiple conformations of the system. Orange solid line represents the optimized ILJ potential. It is obtained by fitting the ILJ potential to the interaction energies calculated or averaged from the H₂/Be₃Al₂(SiO₃)₆-beryl system. The optimized ILJ potential aims to best capture and reproduce the observed interaction energies. Black points represent the Hirshfeld charges of the H₂/Be₃Al₂(SiO₃)₆-beryl system. Hirshfeld charges are a type of atomic charge calculation method. The ILJ potential is also fitted to these charges, meaning that it tries to reproduce the observed charges using its parameters. The Fig. 2, shows the fitting of an ILJ potential to interaction energies and atomic charges in the H₂/Be₃Al₂(SiO₃)₆-beryl system. The pink interaction energies are calculated, the red points represent averaged energies from random conformations, the orange line represents the optimized ILJ potential,

Table 2 To obtain Improved Lennard-Jones (ILJ) potentials for the H₂/Be₃Al₂(SiO₃)₆-beryl, certain parameters are fitted using B97D/TZV2P calculations

Parameters	Fitted	Hirshfeld	CHelpG	MBS	MK	NBO
q_{H} (e ⁻)	0.419	0.291	0.443	0.416	0.426	0.568
β_{AlH}	7.356	7.482	7.367	7.528	7.651	7.893
ϵ_{AlH} (kJ mol ⁻¹)	0.127	0.265	0.346	0.156	0.398	0.219
r_{AlH} (Å)	3.364	3.389	3.462	3.375	3.521	3.502
β_{BeH}	7.851	7.982	7.634	7.911	7.521	7.964
ϵ_{BeH} (kJ mol ⁻¹)	0.269	0.625	0.347	0.512	0.275	0.488
r_{BeH} (Å)	3.761	3.964	3.931	3.826	3.725	3.673
β_{SiH}	8.264	8.759	8.853	8.561	8.672	8.379
ϵ_{SiH} (kJ mol ⁻¹)	0.162	0.199	0.178	0.166	0.137	0.182
r_{SiH} (Å)	3.658	3.782	3.621	3.873	3.918	3.653
β_{OH}	7.571	7.639	7.452	7.527	7.637	7.813
ϵ_{OH} (kJ mol ⁻¹)	0.025	0.096	0.035	0.067	0.085	0.053
r_{OH} (Å)	3.529	3.628	3.693	3.575	3.782	3.556
R^2	0.994	0.995	0.993	0.998	0.994	0.992
MAES (kJ mol ⁻¹)	0.412	0.727	0.448	0.467	0.481	0.458



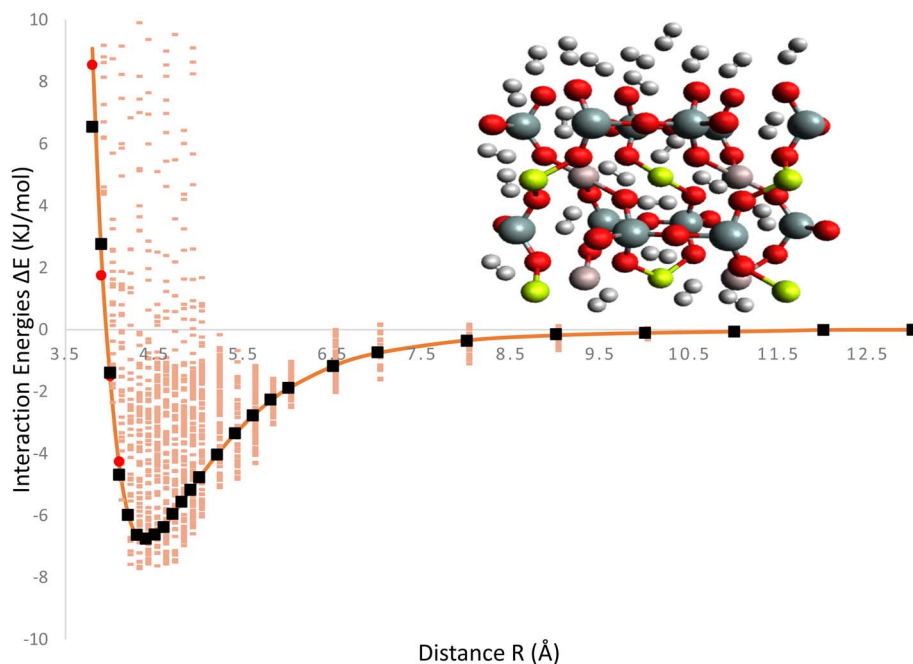


Fig. 2 Interaction energies (pink) from B97D/TZV2P calculations, red points generated from averaged $\text{H}_2/\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl random conformations, orange solid line optimized at ILJ potential and black point Hirshfeld charges also fitted by ILJ potential.

and the black point represents the Hirshfeld charges. This analysis helps in understanding the system's behavior and developing a potential energy model that accurately represents its interactions.

Mean Absolute Error (MAE) is a metric used to evaluate the accuracy of the ILJ potential's fit to the interaction energies or

charges of the $\text{H}_2/\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl system as shown in Fig. 3. The MAE measures the average magnitude of the deviations between the predicted values (obtained from the fitted ILJ potential) and the observed values (the calculated or average interaction energies and the Hirshfeld charges). For each data point, the absolute differences between predicted and observed

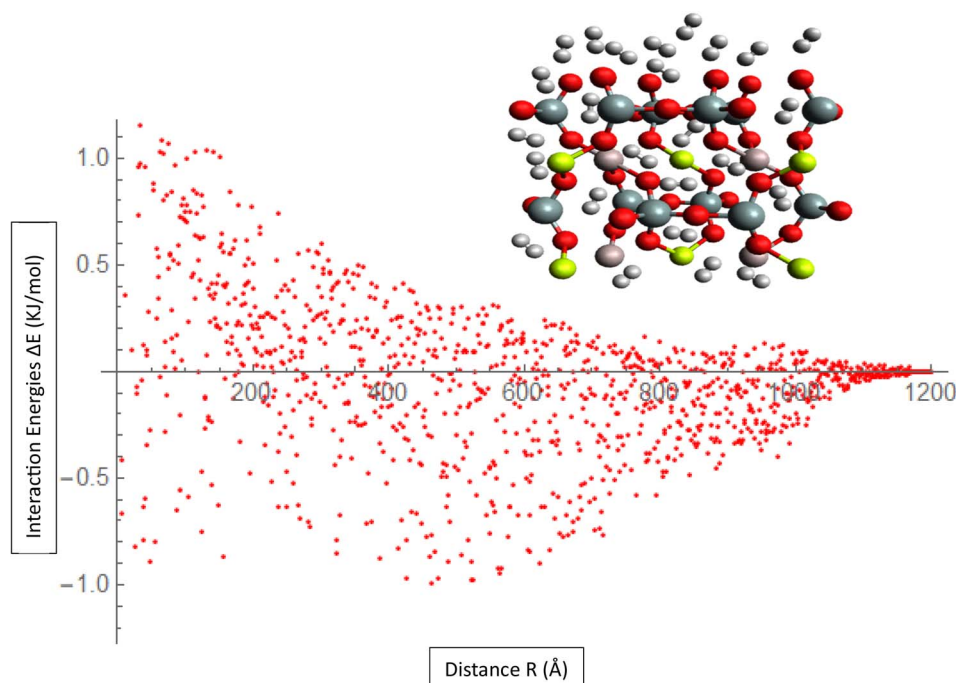


Fig. 3 Fitting ILJ potential to the $\text{H}_2/\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl system, including evaluation of interaction energies, charges, and MAE.



values are computed to determine the MAE. This value represents the mean absolute error. The MAE measures how well the ILJ potential fits the observed data, with smaller values indicating a superior fit. The MAE can be utilized to assess the efficacy of the ILJ potential and compare various potential models. It enables researchers to evaluate the accuracy with which the potential reproduces the interaction energies or charges and offers insight into the quality of the fitting procedure. By minimizing the MAE, one seeks to identify the set of ILJ potential parameters that best represents the observed data, resulting in a more precise description of the system's behavior.

4. Molecular dynamics simulation

The adsorption of H₂ on Be₃Al₂(SiO₃)₆-beryl,³⁶ which is a type of silicate mineral composed of beryllium, aluminum, and silicon, can be studied using molecular dynamics simulations. We create a virtual model of the Be₃Al₂(SiO₃)₆-beryl surface and the H₂ molecules. By employing mathematical equations and force fields that describe the interactions between atoms, the simulation calculates the movements and forces acting on each atom during the simulation timeframe. Through molecular dynamics simulations, we can observe how individual H₂ molecules approach and interact with the Be₃Al₂(SiO₃)₆-beryl surface. The simulations³⁷ can provide insights into the adsorption process, including the binding energies between the H₂ molecules and the Be₃Al₂(SiO₃)₆-beryl surface, the distribution of adsorbed H₂ molecules, and the dynamics of the system. By analyzing the simulation results, we can derive important information such as the adsorption energy, which indicates the strength of the interaction between H₂ and Be₃-Al₂(SiO₃)₆-beryl. They can also investigate the diffusion of H₂ molecules on the Be₃Al₂(SiO₃)₆-beryl surface, which provides insights into the mobility and stability of the adsorbed H₂. These simulations can help to understand the fundamental mechanisms³⁸ of H₂ adsorption on Be₃Al₂(SiO₃)₆-beryl and evaluate its potential as an adsorbent material for hydrogen storage or other applications. They can provide valuable insights into the thermodynamics and kinetics of the adsorption process and guide experimental studies.

DL_POLY³⁹ was utilized for molecular dynamics calculations. The microcanonical (NVE) ensemble used periodic boundary conditions in the *x*, *y*, and *z* directions under standard conditions (273 K and 1 atm). A timestep of 1 fs was utilized for a run time of 5 000 000 steps, including an equilibration of 3 000 000 steps, resulting in a total run time of 5 ns, while monitoring of temperature and energy demonstrated convergence. To assure the standard density of methane, the van der Waals and coulombic cutoffs were set to 18 Å, and 100H₂ molecules were randomly distributed in a cube with 160 Å sides.

We have calculated the diffusion coefficient using the following Einstein equation:

$$D = \frac{1}{2d} \lim_{\tau \rightarrow \infty} \frac{d}{d\tau} \langle [\vec{r}(\tau) - \vec{r}(0)]^2 \rangle \quad (6)$$

where *d* is the dimension of the system and $\vec{r}(\tau)$ is the position of the particle of interest at time τ . The term $\langle [\vec{r}(\tau) - \vec{r}(0)]^2 \rangle$ is

called the mean squared displacement (MSD) and varies linearly over time. To assure accurate statistics, we calculated the diffusion coefficient from five distinct starting configurations and sampled two million MSD values *versus* time for each starting configuration. We evaluated the diffusion coefficient over one million different time origins by dividing timestep 1 to one million, then timestep 2 to one million, and so on. The adsorption behavior of hydrogen molecules on Be₃Al₂(SiO₃)₆-beryl was investigated using two different calculation approaches by density functional theory (DFT) and molecular dynamics (MD) simulations. The obtained results are summarized in Table 2, which includes the distances between the adsorbed hydrogen molecules and the silicate surface, as well as the corresponding configurational energies (*E*_{cfg}) in kilojoules per mole (kJ mol^{−1}).

In Table 3, the distance obtained from the DFT calculation is 2.352 Å. This distance represents the separation or bonding arrangement between the relevant atoms or molecules in the system. The configurational energy obtained from the DFT calculation is −27.58 kJ mol^{−1}. Configurational energy refers to the potential energy of the system associated with a specific arrangement or configuration of the atoms or molecules. In this case, the negative value indicates that the system is in a stable or energetically favorable state with respect to this configuration. The distance obtained from the MD simulation is 2.338 Å. This distance represents the average separation between the relevant atoms or molecules during the simulation at the specified temperature. The configurational energy obtained from the MD simulation at a temperature of 273.15 K is −17.29 kJ mol^{−1}. This energy value represents the potential energy associated with the system's configuration during the simulation under the given conditions. The DFT approach yielded a distance of 2.352 Å, while the MD simulation resulted in a slightly shorter distance of 2.338 Å. The small difference in distance indicates a minor variation in the molecular arrangement or interatomic distances between the two methods. It's important to note that this discrepancy could arise from inherent differences in the underlying methodologies, approximations, or computational models used in DFT and MD. The configurational energy obtained from the DFT approach is −27.58 kJ mol^{−1}, whereas the MD simulation resulted in a lower value of −17.29 kJ mol^{−1}. This significant difference in energy can be attributed to several factors. DFT calculations typically provide more accurate electronic structure information but might not capture certain dynamic or thermodynamic effects present in MD simulations. Additionally, the different parameterizations and potential functions

Table 3 Insights from DFT calculations and MD simulations with modified ILJ parameters about the adsorption of hydrogen molecules (H₂) on Be₃Al₂(SiO₃)₆-beryl

Types of calculation	Distances (Å)	<i>E</i> _{cfg} (kJ mol ^{−1})
DFT approach	2.352	−27.58
MD (<i>T</i> = 273.15 K)	2.338	−17.29



used in the two methods can contribute to variations in the calculated energies.

The adsorption behavior of hydrogen molecules on beryl's can be studied through the use of both density functional theory (DFT) calculations and molecular dynamics (MD) simulations. The trustworthiness of both approaches is highlighted by the fact that the predicted distances are compatible with one another when defining the spatial arrangement of the species that have been adsorbed. When looking at adsorption processes, it is essential to take into account both electronic structure calculations and dynamic impacts because the configurational energy values are so different from one another.

A snapshot refers to a particular moment or instant in time during a simulation where the positions and orientations of the particles or molecules are recorded. The "last configuration" indicates that this snapshot captures the arrangement of the system's particles at the end of the simulation run. $\text{H}_2/\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl system being simulated, composed of H_2 molecules interacting with the $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl surface. Beryl is a mineral composed of beryllium aluminum silicate, and in this context, it serves as the surface interacting with the H_2 molecules. The simulation is conducted at a temperature of 273.15 K. The random white-brown dumbbells are a visual representation of the H_2 molecules present in the system. They indicate that the H_2 molecules are distributed randomly and may exhibit different orientations and positions. The H_2 molecules are shown to be located above or interacting with the $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl surface. This suggests that the simulation is focused on studying the behavior of the H_2 molecules in relation to the surface and their potential interactions or adsorption onto the beryl surface (Fig. 4).

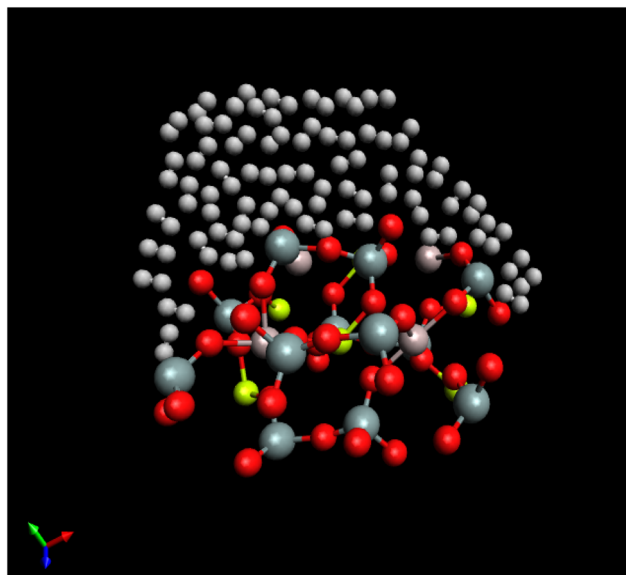


Fig. 4 Snapshot of the last configuration of the simulation for the $\text{H}_2/\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl system at $T = 273.15$ K. The random white-brown dumbbells indicate the H_2 molecules over $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl surface.

5. Conclusion

This study investigated the adsorption of molecular hydrogen (H_2) on $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl, an important silicate mineral, using a combination of Density Functional Theory (DFT) calculations and Molecular Dynamics (MD) simulations. Through extensive structural optimizations and energy calculations using DFT, the study identified favorable adsorption sites and elucidated the binding mechanism. The results revealed that H_2 molecules form mild van der Waals interactions with the exposed oxygen atoms surrounding the octahedral sites of the beryl surface. To investigate the dynamic aspects of H_2 adsorption, MD simulations were conducted with a force field that precisely represented the interatomic interactions within the system and was meticulously parameterized. By subjecting the $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl- H_2 system to a range of temperatures, valuable information regarding the stability, diffusion, and desorption kinetics of H_2 molecules within the beryl structure was obtained. The combined DFT and MD investigations provided a thorough comprehension of the H_2 adsorption phenomenon and shed light on the underlying mechanisms, highlighting the role of surface oxygen atoms and the impact of temperature on H_2 dynamics. In addition, the study demonstrated that the Improved Lennard-Jones (ILJ) potential utilized in the MD simulations was in excellent accord with the DFT calculations and observed system behavior. This agreement enhances the precision and dependability of the simulation method utilized in this study. This study contributes to the fundamental understanding of hydrogen storage and release in beryllium-based silicates, providing valuable insights for the design and optimization of materials for a variety of applications, including hydrogen storage, catalysis, gas separation, sensing, and environmental applications. The findings emphasize $\text{Be}_3\text{Al}_2(\text{SiO}_3)_6$ -beryl's potential as a candidate material for hydrogen-related applications and pave the way for future research in this area.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

The authors extend their sincere appreciation to Researchers Supporting Project number (RSP2024R253), King Saud University, Riyadh, Saudi Arabia.

References

- 1 A. M. Abdalla, S. Hossain, O. B. Nisfindy, A. T. Azad, M. Dawood and A. K. Azad, *Energy Convers. Manage.*, 2018, **165**, 602–627.
- 2 D. Teichmann, W. Arlt and P. Wasserscheid, *Int. J. Hydrogen Energy*, 2012, **37**(23), 18118–18132.
- 3 X. W. Sun, Y. X. Zhang and D. Losic, *J. Mater. Chem. A*, 2017, **5**(19), 8847–8859.
- 4 Y. Li, Z. Y. Fu and B. L. Su, *Adv. Funct. Mater.*, 2012, **22**(22), 4634–4667.



- 5 M. J. Rosseinsky, *Microporous Mesoporous Mater.*, 2004, **73**(1–2), 15–30.
- 6 J. Rouquerol, F. Rouquerol, P. Llewellyn, G. Maurin and K. S. W. Sing, *Adsorption by Powders and Porous Solids: Principle, Methodology and Applications*, Academic Press, 2013.
- 7 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**(18), 3865–3868.
- 8 M. P. Allen and D. J. Tildesley, *Computer Simulation of Liquids*, Oxford University Press, 2nd edn, 2017.
- 9 B. Smit and T. L. M. Maesen, *Nature*, 2008, **451**(7179), 671–678.
- 10 D. Frenkel and B. Smit, *Understanding Molecular Simulations: from Algorithms to Applications*, Academic Press, San Diego, 3rd edn, 2023.
- 11 D. Dubbeldam, S. Calero, T. J. H. Vlugt, R. Krishna, T. L. Maesen, E. Beerdsen and B. Smit, *Phys. Rev. Lett.*, 2004, **93**(8), 088302.
- 12 J. A. Armstrong and M. T. Weller, *J. Am. Chem. Soc.*, 2010, **132**(44), 15679–15686.
- 13 S. A. Mian, L. C. Saha, J. Jang, L. Wang, X. Gao and S. Nagase, *J. Phys. Chem. C*, 2010, **114**(48), 20793–20800.
- 14 I. C. Bourg and C. I. Steefel, *J. Phys. Chem. C*, 2012, **116**(21), 11556–11564.
- 15 C. D. Wu, T. H. Fang, J. Y. Lo and Y. L. Feng, *J. Mol. Model.*, 2013, **19**(9), 3813–3819.
- 16 K. V. Kumar, K. Preuss, L. Lu, Z. X. Guo and M. M. Titirici, *J. Phys. Chem. C*, 2015, **119**, 22310–22321.
- 17 F. Pirani, M. Albertí, A. Castro, M. M. Teixidor and D. Cappelletti, *Chem. Phys. Lett.*, 2004, **394**(1–3), 37–44.
- 18 F. Pirani, S. Brizi, L. F. Roncaratti, P. Casavecchia, D. Cappelletti and F. Vecchiocattivi, *Phys. Chem. Chem. Phys.*, 2008, **10**(36), 5489–5503.
- 19 A. Lombardi and F. Palazzetti, *J. Mol. Struct.: THEOCHEM*, 2008, **852**(1–3), 22–29.
- 20 M. Albertí, A. Aguilar, J. M. Lucas and F. Pirani, *J. Phys. Chem. A*, 2012, **116**(22), 5480–5490.
- 21 M. K. Rana, H. S. Koh, H. Zuberi and D. J. Siegel, *J. Phys. Chem. C*, 2014, **118**(6), 2929–2942.
- 22 S. Grimme, *J. Comput. Chem.*, 2006, **27**(15), 1787–1799.
- 23 S. Grimme, *J. Comput. Chem.*, 2004, **25**(12), 1463–1473.
- 24 R. Peverati and K. K. Baldrige, *J. Chem. Theory Comput.*, 2008, **4**(12), 2030–2048.
- 25 L. M. Roch and K. K. Baldrige, *Phys. Chem. Chem. Phys.*, 2017, **19**(38), 26191–26200.
- 26 H. Liu, G. Qi, Q. Song and H. Wang, *J. Photochem. Photobiol., A*, 2018, **354**, 119–126.
- 27 A. V. Marenich, C. J. Cramer and D. G. Truhlar, *J. Phys. Chem. B*, 2009, **113**(18), 6378–6396.
- 28 J. Autschbach, *ChemPhysChem*, 2009, **10**(11), 1757–1760.
- 29 M. M. Ghahremanpour, P. J. Van Maaren and D. Van Der Spoel, *Sci. Data*, 2018, **5**(1), 1–10.
- 30 F. Weinhold and C. R. Landis, *Valency and Bonding: A Natural Bond Orbital Donor-Acceptor Perspective*, Cambridge University Press, 2005, p. 760.
- 31 A. Schäfer, C. Huber and R. Ahlrichs, *J. Chem. Phys.*, 1994, **100**(8), 5829–5835.
- 32 M. Metcalf, J. Reid and M. Cohen, *Modern Fortran explained*, Oxford University Press, New York, 2011.
- 33 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, V. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery, J. J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam, M. Klene, J. E. Knox, J. B. Cross, V. Bakken, C. Adamo, J. Jaramillo, R. Gomperts, R. E. Stratmann, O. Yazyev, A. J. Austin, R. Cammi, C. Pomelli, J. W. Ochterski, R. L. Martin, K. Morokuma, V. G. Zakrzewski, G. A. Voth, P. Salvador, J. J. Dannenberg, S. Dapprich, A. D. Daniels, Ö. Farkas, J. B. Foresman, J. V. Ortiz, J. Cioslowski, and D. J. Fox, *Gaussian-09 Revision D.01*, 2009.
- 34 B. S. Mallik, I. W. Kuo, L. E. Fried and J. I. Siepmann, *Phys. Chem. Chem. Phys.*, 2012, **14**(14), 4884–4890.
- 35 K. Srinivasu and S. K. Ghosh, *J. Phys. Chem. C*, 2012, **116**(48), 25184–25189.
- 36 H. Hao, Y. Cao, L. Li, G. Fan and J. Liu, *Appl. Surf. Sci.*, 2021, **537**, 147926.
- 37 S. Caratzoulas, D. G. Vlachos and M. Tsapatsis, *J. Am. Chem. Soc.*, 2006, **128**(2), 596–606.
- 38 Y. Zhou, D. Hou, H. Manzano, C. A. Orozco, G. Geng, P. J. Monteiro and J. Liu, *ACS Appl. Mater. Interfaces*, 2017, **9**(46), 41014–41025.
- 39 W. Smith and T. R. Forester, *J. Mol. Graphics*, 1996, **14**(3), 136–141.

