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Quaternary ammonium fluorides and difluorosilicates as nucleophilic fluorination reagents[†]

Michal Trojan, ^a Kateřina Kučnirová, ^a Šárka Bouzková, ^a Josef Cvačka, ^b Jan Čejka, ^c Gašper Tavčar, ^d Markéta Rybáčková ^a and Jaroslav Kvíčala ^{a*}

TBAT (tetrabutylammonium difluorotriphenylsilicate) is an excellent homogeneous nucleophilic fluorination reagent, but a high excess of the reagent was reported to be essential. We hence optimized the reaction conditions and compared its nucleophilic fluorination reactivity with that of other common commercial nucleophilic fluorination reagents, such as anhydrous TBAF and TASF (tris(dimethylamino)sulfonium difluorotrimethylsilicate). As the substrates, we employed a standard set of primary and secondary octyl substrates under identical conditions. To eliminate the possibility of hydrogen fluoride elimination in the above reagents, we prepared four quaternary ammonium fluorides lacking β -elimination possibility in the hydrocarbon chain, transformed them to the corresponding difluorotriphenylsilicates, and compared their reactivity with that of the commercial reagents. Furthermore, attempts to isolate analogous tetrabutylammonium difluoromethyldiphenylsilicate or difluorodimethylphenylsilicate failed, as was confirmed by comparison of the published experimental data with computed ¹⁹F NMR spectra. Finally, we studied the transition states of decomposition of various tetramethylammonium methylphenyldifluorosilicates by DFT methods and found that their relative energies increase with an increasing number of phenyl groups. The formation of difluorosilicates is a nearly barrierless process.

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Introduction

Fluorinated compounds play an indispensable role in industrial chemistry,¹ medicinal chemistry^{2–4} and agrochemistry.⁵ Among plentiful variants of their preparation methods, nucleophilic fluorination is the most traditional and still extensively applied approach.^{6,7} Two basic versions of nucleophilic fluorination are available, with deoxyfluorination of alcohols or phenols being probably the most popular approach.⁸ However, haloalkanes are common substrates in organic chemistry, and so far, methods for converting them to fluorides are limited and insufficient. The main drawback of cheap and therefore frequently used reagents such as KF is their poor solubility and high basicity, which can be solved by the use of ionic liquids, crown ethers or cryptands^{7,9}

Combination of crown ethers with an appropriate sterically hindered alcohol improved the selectivity of nucleophilic fluorination of primary substrates and was studied both experimentally and by DFT methods, but was not tested on critical secondary substrates.^{10–13} Promising results were obtained with KF and highly sophisticated calixcrown ethers in combination with sterically hindered alcohols as solvents even for one selected secondary bromide.^{14,15} Unfortunately, no standard secondary substrates were tested. The most common way to cope with the solubility problem is the use of tetrabutylammonium fluoride (TBAF, 1).¹⁶ However, its basicity still remains a critical issue and the elimination side reaction is the main obstacle in nucleophilic substitution reactions. To cope with that, a sterically hindered hydrogen donor, *tert*-butylalcohol, was successfully employed to control TBAF basicity, which significantly improved substitution/elimination ratio for primary substrates.¹⁷ Unfortunately, this system was again not tested for critical secondary substrates. Fluorination selectivity and the structure of such complexes were further studied both experimentally¹⁸ and by DFT methods.¹⁹ Moreover, NMR spectroscopy was employed to understand better the coordination of naked F[–] anions with ureas.²⁰ Probably the best recent solution for fluorination of secondary substrates is the use of modified phenanthrene complexed CuHF₂.²¹ Another problem of TBAF is that it crystallizes with three molecules of water, and attempts to remove the water by heating result in its

^aDepartment of Organic Chemistry, University of Chemistry and Technology, Prague, Technická 5, 166 28 Prague 6, Czech Republic. E-mail: kvicalaj@vscht.cz

^bInstitute of Organic Chemistry and Biochemistry of the Czech Academy of Sciences, Flemingovo náměstí 542/2, 160 00 Prague 6, Czech Republic

^cDepartment of Solid State Chemistry, University of Chemistry and Technology, Prague, Technická 5, 166 28 Prague 6, Czech Republic

^dDepartment of Inorganic Chemistry and Technology, "Jožef Stefan" Institute, Jamova cesta 39, Ljubljana, Slovenia

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decomposition.²² Synthesis of anhydrous TBAF was developed by DiMagno *et al.*,²³ and this solution is commercially available. To avoid the decomposition of the tetrabutylammonium cation and the presence of water, hypervalent silicon compounds have been introduced as a stable source of anhydrous fluoride anions. However, a large excess of tetrabutylammonium difluorotriphenylsilicate (TBAT, 2) has been reported to be used to achieve satisfactory yield in fluorination of octan-2-yl mesylate (4b). Moreover, only low yield of the fluorinated product was obtained from 2-bromo-octane (5b).²⁴ Tris(dimethylamino)sulfonium difluorotrimethylsilicate (TASF, 3) has been developed and proved to be highly efficient in fluorination of sterically hindered reagents such as protected saccharides²⁵ or amino acids (Fig. 1).²⁶ With the original aim of synthesizing hypervalent inorganic anions using quaternary ammonium fluorides lacking β -hydrogen and hence not prone to β -elimination, trimethylneopentylammonium fluoride has been synthesized.²⁷

Herein, we wish to report a systematic study comparing the reactivity of quaternary ammonium fluorides lacking β -hydrogen and the corresponding difluorosilicates with that of the abovementioned fluorination reagents. Furthermore, a computational DFT study dealing with the stability of various tetramethylammonium methylphenylsilicates and the transition states of their decomposition to the corresponding fluorosilanes and TBAF was performed (Scheme 1).

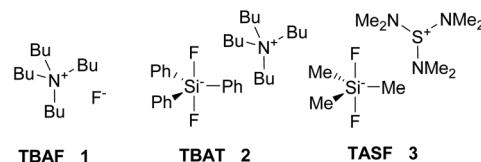
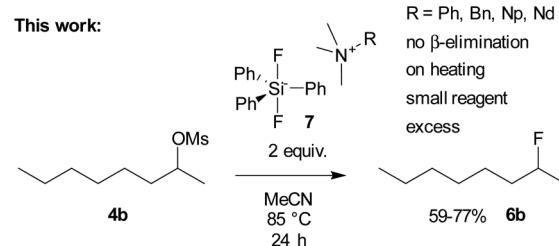
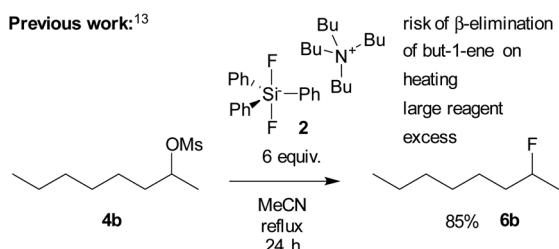


Fig. 1 Nucleophilic fluorination reagents used for comparison.



Scheme 1 Nucleophilic fluorination with difluorosilicates.

Results and discussion

Quaternary ammonium fluorides 8

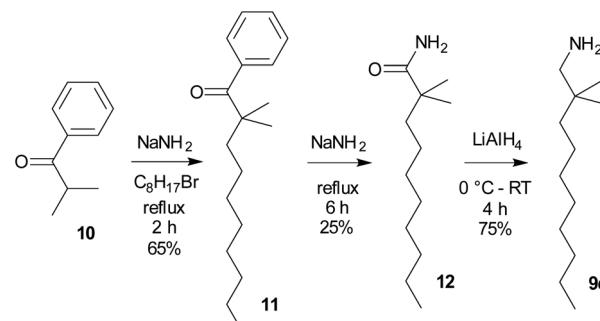
As the starting compounds, the corresponding primary amines 9 were employed. While amines 9a–9c are commercially available, 2,2-dimethyldecan-1-amine (9d) was prepared from 2-methyl-1-phenylpropan-1-one (isobutyrophenone, 10) by a three step procedure in analogy to ref. 28 (Scheme 2). This long-chained amine, 9d, was synthesized with the aim of improving the solubility of both fluoride 8d and difluorosilicate 7d in organic solvents.

Quaternary ammonium fluorides 8 were obtained from the starting amines 9 by quaternization with methyl iodide, followed by iodine-fluorine exchange with AgF in analogy to ref. 27 (Scheme 3).

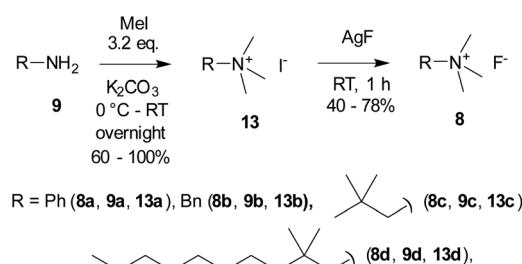
Quaternary ammonium difluorosilicates 7

The reaction of fluorides 8 with triphenylsilyl fluoride (14) at room temperature in acetonitrile afforded the target difluorosilicates 7 in poor to good yields as white solids (Scheme 4).

A single crystal of benzyltrimethylammonium difluorotriphenylsilicate (7b) was obtained and its X-ray structure is shown in Fig. 2. The ammonium cation is coordinated to one fluorine and two phenyl rings of the silicate anion. Although the X-ray structure of TBAT has been reported,²⁹ the position of the ammonium cation is not shown and hence the structures could not be compared. On the other hand, the X-ray structure of potassium difluorotriphenylsilicate complexed with [2.2.2]cryptand shows no such coordination due to the

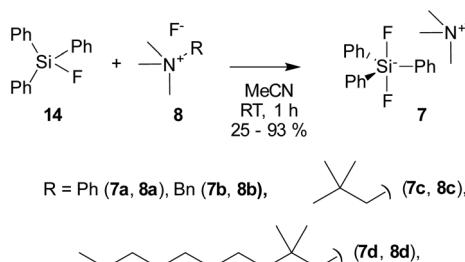


Scheme 2 Preparation of the starting long chain branched amine 9d.



Scheme 3 Preparation of quaternary ammonium fluorides 8.





Scheme 4 Preparation of quaternary ammonium difluorosilicates 7.

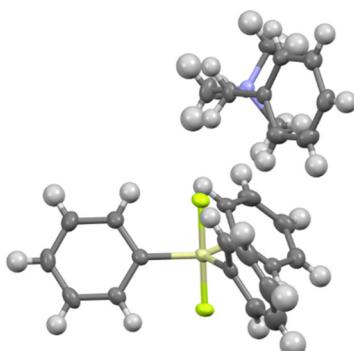


Fig. 2 Single crystal structure of benzylsilicate 7b. Crystallographic details are given in the ESI.†

bulky complexed cation.³⁰ The details of crystallographic analysis are given in the ESI.†

Nucleophilic fluorinations with fluorides 8 and difluorosilicates 7

To compare the activity of various fluorination reagents, two primary and two secondary substrates were chosen. In analogy to the paper devoted to TBAT,²⁴ we employed octyl mesylate (4a) and 1-bromoocetane (5a) as the primary substrates, while octan-2-yl mesylate (4b) and 2-bromoocetane (5b) were used as the secondary substrates. All reactions were performed in sealed Schlenk flasks in CD₃CN at 85 °C for 24 h using a two-fold excess of the reagent. In all cases, the average of two fluorination runs is given.

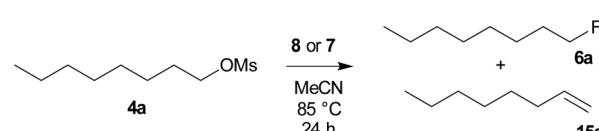
We first tested whether the published²⁴ high excess of 6 equivalents of reagent 2 is really needed. To our surprise, 2 equivalents were sufficient to obtain nearly quantitative conversion of octyl mesylate (4a) (Table 1) and we hence used it in all fluorinations. Lower temperatures gave significantly lower conversions.

Fluorination of octyl mesylate (4a). The results of fluorination of octyl mesylate (4a) (Scheme 5) are summarized in Table 2. Conversions and the ratios of products were determined from ¹H NMR spectra after careful baseline correction and phasing. Apart from long branched quaternary ammonium cation containing reagents 8d and 7d, for which substitution was accompanied by reasonable elimination, all other reagents gave nearly quantitative conversions and very

Table 1 Optimization of reaction conditions for fluorination of octyl mesylate (4a) with TBAT (2)

| TBAT equiv. | Time (h) | Temp. (°C) | Conversion (%) |
|----------------|----------|------------|----------------|
| 1.1 | 24 | 85 | 62 |
| 1.1 | 48 | 85 | 71 |
| 2 | 24 | 85 | 95 |
| 2 | 24 | 40 | ~3 |
| 2 | 24 | RT | 0 |
| 6 ^a | 24 | 85 | 96 |

^a High excess of reagent employed in ref. 24.



Scheme 5 Fluorination of octyl mesylate (4a) with reagent 8 or 7.

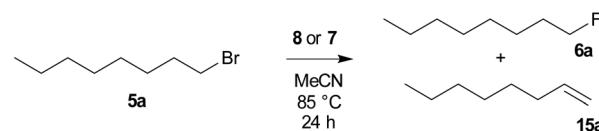
Table 2 Results of fluorination of octyl mesylate (4a) with quaternary ammonium fluorides 8 and difluorotriphenylsilicates 7

| Reagent | Conversion | Product ratio 6a : 15a |
|--|------------|------------------------|
| TBAF (1) | 99% | 60 : 40 |
| TBAT (2) | 95% | 97 : 3 |
| TASF (3) | 95% | 96 : 4 |
| PhN ⁺ Me ₃ F ⁻ (8a) | 73% | 98 : 2 |
| BnN ⁺ Me ₃ F ⁻ (8b) | 99% | 98 : 2 |
| NpN ⁺ Me ₃ F ⁻ (8c) ^a | 99% | 98 : 2 |
| NdN ⁺ Me ₃ F ⁻ (8d) ^b | 90% | 66 : 34 |
| PhN ⁺ Me ₃ Ph ₃ SiF ₂ ⁻ (7a) | 99% | 97 : 3 |
| BnN ⁺ Me ₃ Ph ₃ SiF ₂ ⁻ (7b) | 99% | 99 : 1 |
| NpN ⁺ Me ₃ Ph ₃ SiF ₂ ⁻ (7c) ^a | 77% | 98 : 2 |
| NdN ⁺ Me ₃ Ph ₃ SiF ₂ ⁻ (7d) ^b | 99% | 86 : 14 |

^a Np = 2,2-dimethylpropyl (neopentyl). ^b Nd = 2,2-dimethyldecyl (neododecyl).

good selectivities. TBAF (1) proved to be the worst reagent with nearly 40% elimination to oct-1-ene (15a).

Fluorination of 1-bromoocetane (5a). The results of fluorination of 1-bromoocetane (5a) (Scheme 6) are summarized in Table 3. With less reactive 1-bromoocetane 5a, the best selectivity was achieved with TBAT (2) and neopentyl-based difluorosilicate 7c; however, the conversion was somewhat lower. The most reactive reagents TBAF (1) and TASF (3) also gave much elimination, around 40%. Probably the best combination of conversion and selectivity was achieved with long branched difluorosilicate 7d.



Scheme 6 Fluorination of 1-bromoocetane (5a) with reagent 8 or 7.

Table 3 Results of fluorination of 1-bromoocetane (5a) with quaternary ammonium fluorides 8 and difluorotriphenylsilicates 7

| Reagent | Conversion | Product ratio 6a : 15a |
|--|------------|------------------------|
| TBAF (1) | 99% | 61 : 39 |
| TBAT (2) | 75% | 99 : 1 |
| TASF (3) | 99% | 69 : 41 |
| PhN ⁺ Me ₃ F ⁻ (8a) | 52% | 70 : 30 |
| BnN ⁺ Me ₃ F ⁻ (8b) | 93% | 67 : 33 |
| NpN ⁺ Me ₃ F ⁻ (8c) ^a | 99% | 70 : 30 |
| NdN ⁺ Me ₃ F ⁻ (8d) ^b | 64% | 61 : 39 |
| PhN ⁺ Me ₃ Ph ₃ SiF ₂ ⁻ (7a) | 99% | 84 : 16 |
| BnN ⁺ Me ₃ Ph ₃ SiF ₂ ⁻ (7b) | 82% | 85 : 15 |
| NpN ⁺ Me ₃ Ph ₃ SiF ₂ ⁻ (7c) ^a | 65% | 98 : 2 |
| NdN ⁺ Me ₃ Ph ₃ SiF ₂ ⁻ (7d) ^b | 99% | 87 : 13 |

^a Np = 2,2-dimethylpropyl (neopentyl). ^b Nd = 2,2-dimethyldecyl (neododecyl).

Fluorination of octan-2-yl mesylate (4b). The results of fluorination of octan-2-yl mesylate (4b) (Scheme 7) are summarized in Table 4. Secondary mesylate 4b is highly prone to elimination. Among the reagents tested, the best ratio of conversion and selectivity was given by TBAT (2), closely followed by benzyl containing difluorosilicate 7b. Quite surprisingly, the selectivity of the highly reactive commercial reagent TASF (3) was quite poor.

Fluorination of 2-bromoocetane (5b). The results of fluorination of 2-bromoocetane (5b) (Scheme 8) are summarized in Table 5. As expected, 2-bromoocetane (5b) proved to be the most demanding substrate for nucleophilic fluorination reactions. Again, TBAT (2) was the most successful reagent, closely followed by benzyl-based difluorosilicate 7b. TASF reagent (3) again showed the worst selectivity with complete elimination. Among the prepared reagents, difluorosilicates 7 gave better selectivities than fluorides 8.

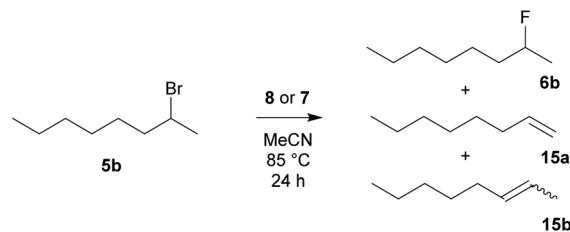
The main advantage of the newly prepared reagents 8 and 7 is that their synthesis avoids highly corrosive conditions such as HF, as well as the use of expensive anhydrous TBAF. Moreover, while the use of TBAT (2) at higher temperatures cannot fully exclude that the actual fluorinating agent is tetrabutylammonium hydrogen difluoride formed by decomposition of TBAT, this is not possible for reagents 8 and 7. Among the new reagents, benzyltrimethylammonium difluorotriphenylsilicate (7b) was probably the most efficient for difficult secondary substrates with results close to those of the commercial TBAT reagent.

Attempted synthesis of difluoromethylphenylsilicates 16 and 17. Synthesis of tetrabutylammonium difluoromethyl-

Table 4 Results of fluorination of octan-2-yl mesylate (4b) with quaternary ammonium fluorides 8 and difluorotriphenylsilicates 7

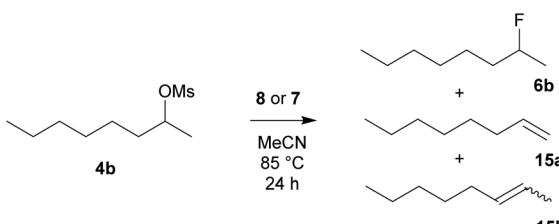
| Reagent | Conversion | Product ratio 6b : 15b : 15a |
|--|------------|------------------------------|
| TBAF (1) | 99% | 59 : 30 : 11 |
| TBAT (2) | 99% | 79 : 17 : 4 |
| TASF (3) | 99% | 29 : 37 : 34 |
| PhN ⁺ Me ₃ F ⁻ (8a) | 99% | 62 : 29 : 9 |
| BnN ⁺ Me ₃ F ⁻ (8b) | 99% | 57 : 27 : 16 |
| NpN ⁺ Me ₃ F ⁻ (8c) ^a | 99% | 57 : 27 : 16 |
| NdN ⁺ Me ₃ F ⁻ (8d) ^b | 96% | 39 : 52 : 9 |
| PhN ⁺ Me ₃ Ph ₃ SiF ₂ ⁻ (7a) | 94% | 69 : 26 : 5 |
| BnN ⁺ Me ₃ Ph ₃ SiF ₂ ⁻ (7b) | 99% | 77 : 19 : 4 |
| NpN ⁺ Me ₃ Ph ₃ SiF ₂ ⁻ (7c) ^a | 81% | 80 : 16 : 4 |
| NdN ⁺ Me ₃ Ph ₃ SiF ₂ ⁻ (7d) ^b | 94% | 59 : 34 : 7 |

^a Np = 2,2-dimethylpropyl (neopentyl). ^b Nd = 2,2-dimethyldecyl (neododecyl).

**Scheme 8** Fluorination of 2-bromoocetane (5b) with reagent 8 or 7.**Table 5** Results of fluorination of 2-bromoocetane (4b) with quaternary ammonium fluorides 8 and difluorotriphenylsilicates 7

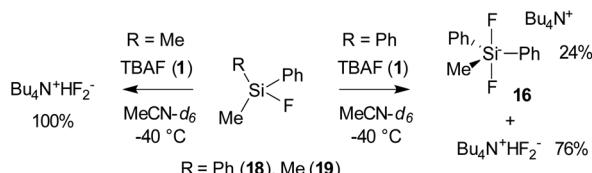
| Reagent | Conversion | Product ratio 6b : 15b : 15a |
|--|------------|------------------------------|
| TBAF (1) | 99% | 28 : 62 : 10 |
| TBAT (2) | 92% | 39 : 53 : 8 |
| TASF (3) | 99% | 0 : 81 : 19 |
| PhN ⁺ Me ₃ F ⁻ (8a) | 69% | 12 : 77 : 11 |
| BnN ⁺ Me ₃ F ⁻ (8b) | 99% | 14 : 72 : 14 |
| NpN ⁺ Me ₃ F ⁻ (8c) ^a | 90% | 19 : 69 : 12 |
| NdN ⁺ Me ₃ F ⁻ (8d) ^b | 99% | 19 : 71 : 10 |
| PhN ⁺ Me ₃ Ph ₃ SiF ₂ ⁻ (7a) | 65% | 33 : 59 : 8 |
| BnN ⁺ Me ₃ Ph ₃ SiF ₂ ⁻ (7b) | 92% | 36 : 56 : 8 |
| NpN ⁺ Me ₃ Ph ₃ SiF ₂ ⁻ (7c) ^a | 82% | 34 : 57 : 9 |
| NdN ⁺ Me ₃ Ph ₃ SiF ₂ ⁻ (7d) ^b | 41% | 20 : 72 : 8 |

^a Np = 2,2-dimethylpropyl (neopentyl). ^b Nd = 2,2-dimethyldecyl (neododecyl).

**Scheme 7** Fluorination of octan-2-yl mesylate (4b) with reagent 8 or 7.

diphenylsilicate (16) and difluorodimethylphenylsilicate (17) has been reported about twenty years before.^{31,32} However, the reported ¹⁹F NMR shifts (−140 to −150 ppm) do not correspond to the ¹⁹F NMR shift of TBAT (−96.0 ppm²⁴) and potassium difluoromethyldiphenylsilicate complexed with [2.2.2] cryptand (−85.6 ppm).³⁰ The previously reported data are thus erroneous and correspond to the HF₂⁻ anion. When we mixed a solution of fluoromethyldiphenylsilane (18) in CD₃CN with “anhydrous” TBAF in an NMR tube and immediately recorded the ¹⁹F NMR spectrum at −40 °C, two major signals in a 1 : 3 ratio were observed (Scheme 9). The minor signal (−86.4 ppm, bs, 24%) corresponds to the target difluorosilicate in good





Scheme 9 Attempted synthesis of difluoromethylphenylsilicates **16** and **17**.

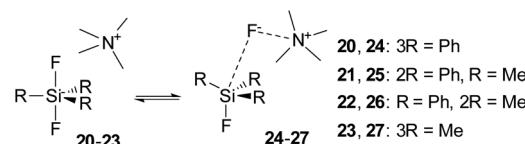
agreement with the computed value (see Table 6), while the major signal (-151.0 ppm, bd, $^1J_{H-F} = 128$ Hz, 76%) is the signal of the HF_2^- anion (see Fig. S37 in the ESI†). After stirring the mixture for 1 h at RT, only the signal of the HF_2^- anion was observed.

In an analogous experiment with fluorodimethylphenylsilane (**19**) and “anhydrous” TBAF (Scheme 9), only the signal of the HF_2^- anion was observed in the ^{19}F NMR spectrum (see Fig. S38 in the ESI†).

Computations

Both previous and our failures to synthesize difluorosilicates modified with methyl groups resulted in our aim to study the stability of these compounds and the connected fluorosilane–quaternary ammonium fluoride complexes by DFT methods. To simplify our calculations, we substituted the tetrabutylammonium cation in TBAT (2), and in **16** and **17**, the tetramethylammonium cation. The computational results imply that all difluorosilicates **20–23** should be stable compounds in contrast to the experimental data. For the obtained minima of difluorosilicates **20–23** and the corresponding fluorosilane–ammonium fluoride complexes **24–27** (Scheme 10), chemical shieldings of fluorine atoms and the respective computed chemical shifts obtained by comparison with CCl_3F standard shielding are listed in Table 6.

While the computed results agree well with the experimental chemical shifts for TBAT (2), they strongly differ from the erroneously reported shifts for methylated difluorosilicates **16** and **17**. Furthermore, we were interested in how high is the energy barrier of decomposition of difluorosilicates **20–23** to fluorosilane–fluoride complexes **24–27**. We hence started a DFT study of the corresponding potential energy surfaces. The results show that the transition state Gibbs free energy gradually increases with an increasing number of phenyls in the



Scheme 10 Equilibrium among difluorosilicates **20–23** and fluorosilane–TMAF complexes **24–27**.

corresponding difluorosilicates (see Fig. 3 for the individual barriers and Fig. 4 for critical structures). Interestingly, the computational results show that the formation of difluorosilicates from fluorosilane and fluoride anions is a nearly barrierless process.

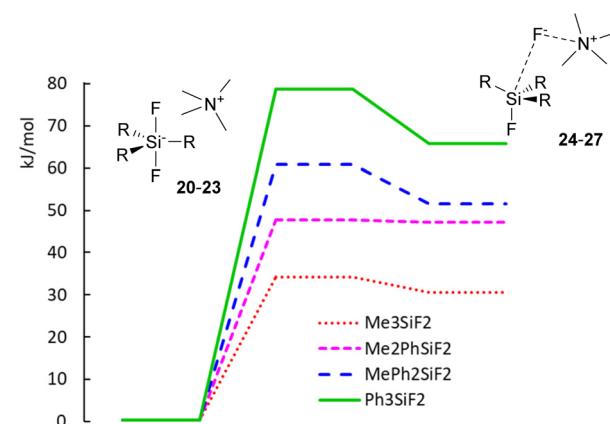


Fig. 3 Potential energy surfaces of decomposition of difluorosilicates **20–23** to fluorosilane–ammonium fluoride complexes **24–27**.



Fig. 4 Examples of computed critical structures – decomposition of tetramethylammonium difluorotriphenylsilicate (**20**) to fluorosilane–fluoride complex **24**.

Table 6 Comparison of computed and reported experimental ^{19}F NMR shifts for difluoromethylphenylsilicates **2**, **16**, **17**, **20–23** and the fluorosilane–tetramethylammonium fluoride complexes **24–27**

| R | Silicate | Exp. δ | Silicate | Comp. δ | Silane–F $^-$ complex | Comp. δ |
|--------------------|-----------|---------------------------------|-----------|----------------|-----------------------|----------------|
| Ph ₃ | 2 | -96.0 (ref. 24) | 20 | -100.2 | 24 | -203.9 |
| Ph ₂ Me | 16 | -157.3 (ref. 31) ^{a,b} | 21 | -88.6 | 25 | -204.9 |
| PhMe ₂ | 17 | -150.9 (ref. 31) ^a | 22 | -86.6 | 26 | -196.3 |
| Me ₃ | — | -144.2 (ref. 31) ^a | 23 | -76.7 | 27 | -189.3 |

^a Erroneous values. ^b -86.4 ppm was observed by us for **16** in good agreement with computations.



Experimental

Materials and methods

All reactions were performed under an argon atmosphere in oven dried flasks using a standard inert technique, unless otherwise noted. Fluorinations were performed in sealed vials. ¹H NMR spectra were recorded with an Agilent 400-MR DDR2 spectrometer at working frequencies of 399.94 MHz for ¹H NMR, 376.29 MHz for ¹⁹F NMR and 100.58 MHz for ¹³C NMR or with a JEOL-ECZL400G spectrometer at working frequencies of 399.78 MHz for ¹H NMR, 376.17 MHz for ¹⁹F NMR and 100.53 MHz for ¹³C NMR, in deuterated solvents. Chemical shifts (δ) are reported in parts per million (ppm) with reference to the residual solvent peak. Signals are described as s = singlet, d = doublet, t = triplet, m = multiplet, bs = broad singlet. Coupling constants (J) are reported in Hz. Mass spectra were recorded on an LTW Orbitrap XL (Thermo Fisher Scientific) instrument. Octyl mesylate (**4a**),³³ octan-2-yl mesylate (**4b**),³⁴ fluoromethyl diphenylsilane (**18**) and fluorodimethylphenylsilane (**19**)³² were prepared according to the published procedures.

2,2-Dimethyl-1-phenyldecan-1-one (11). To a flask containing NaNH₂ (0.92 g; 24 mmol; 1.4 equiv.) equipped with a mechanical stirrer, dry toluene (10 mL) and isobutyrophenone (**10**, 3.0 g; 20 mmol; 1.2 equiv.) were added and the mixture was heated to reflux for 2 h. Then a solution of octyl bromide (3.2 g; 17 mmol; 1.0 equiv.) in anhydrous toluene (10 mL) was added and heated to reflux for 6 h. The mixture was cooled to r.t., quenched with water and the organic layer was washed with water (3 \times 20 mL), brine (3 \times 20 mL) and dried with anh. MgSO₄. The solvent was evaporated to give ketone **11** (2.8 g, 65%). ¹H NMR (399.94 MHz, CDCl₃): δ 0.86 (t, $^3J_{H-H}$ = 6.9 Hz, 3H, CH₃); 1.18–1.33 (m, 14H, CH₂); 1.29 (s, 6H, C(CH₃)₂); 7.35–7.42 (m, 1H, Ar–H); 7.42–7.48 (m, 2H, Ar–H); 7.63–7.68 (m, 2H, Ar–H) ppm. ¹³C NMR (100.58 MHz, CDCl₃): δ 14.2 (1C, CH₃); 22.8 (1C, CH₂); 24.9 (1C, CH₂); 26.3 (2C, C(CH₃)₂); 29.3 (1C, CH₂); 29.5 (1C, CH₂); 30.3 (1C, CH₂); 32.0 (1C, CH₂); 41.2 (1C, CH₂); 48.0 (1C, C(CH₃)₂); 127.7 (2C, Ar–CH); 128.2 (2C, Ar–CH); 130.9 (1C, Ar–CH); 139.4 (1C, Ar–C); 209.5 (1C, C=O) ppm. HRMS (ESI): calcd for C₁₈H₂₉O [M + H]⁺ 261.2213, found 261.2214.

2,2-Dimethyldecanamide (12)

To a flask containing NaNH₂ (0.84 g; 22 mmol; 2.0 equiv.), a solution of ketone **11** (2.8 g, 11 mmol; 1.0 equiv.) in anhydrous toluene (15 mL) was added and the mixture was heated to reflux. After 6 h, the reaction mixture was cooled to r.t., quenched with water and the organic layer was washed with water (3 \times 20 mL), brine (3 \times 20 mL), and dried with anh. MgSO₄. Then the solvent was evaporated to give amide **12** (0.53 g, 25%). ¹H NMR (399.94 MHz, CDCl₃): δ 0.87 (t, $^3J_{H-H}$ = 7.0 Hz, 3H, CH₃); 1.18 (s, 6H, C(CH₃)₂); 1.26 (m, 10H, CH₂); 1.49 (m, 2H, CH₂C); 5.21 (s, 1H, H₂N); 5.56 (s, 1H, H₂N) ppm. ¹³C NMR (100.58 MHz, CDCl₃): δ 14.3 (1C, CH₃); 22.8 (1C, CH₂); 25.0 (1C, CH₂); 25.7 (2C, C(CH₃)₂); 29.4 (1C, CH₂); 29.7 (1C, CH₂); 30.3 (1C, CH₂); 32.0 (1C, CH₂); 41.6 (1C, CH₂); 42.2

(1C, C(CH₃)₂); 180.7 (1C, C=O) ppm. HRMS (ESI): calcd for C₁₂H₂₅NNaO [M + Na]⁺ 222.1828, found 222.1829.

2,2-Dimethyldecan-1-amine (9d)

Amide **12** (0.92 g; 4.6 mmol; 1.0 equiv.) was dissolved in anhydrous Et₂O (25 mL) and the mixture was cooled to 0 °C in an ice bath. To the reaction mixture, LiAlH₄ (4.6 mL; 19 mmol; 4.0 equiv., 4 M sol. in Et₂O) was slowly added. After 4 h, the reaction mixture was cooled down in an ice bath and was slowly quenched with water. To the mixture, NaOH was added (10 mL, 10% aqueous solution) and the mixture was stirred for 30 min. The organic layer was separated and washed with water (3 \times 20 mL), brine (2 \times 20 mL), and dried with anh. MgSO₄. After evaporation, amine **9d** (0.65 g, 75%) was obtained. ¹H NMR (399.94 MHz, CDCl₃): δ 0.79 (s, 6H, C(CH₃)₂); 0.85 (t, $^3J_{H-H}$ = 7.0 Hz, 3H, CH₃); 1.07–1.32 (m, 14H, CH₂); 2.39 (s, 2H, H₂NCH₂) ppm. ¹³C NMR (100.58 MHz, CDCl₃): δ 14.3 (1C, CH₃); 22.8 (1C, CH₂); 24.0 (2C, C(CH₃)₂); 24.8 (1C, CH₂); 29.5 (1C, CH₂); 29.8 (1C, CH₂); 30.7 (1C, CH₂); 32.0 (1C, CH₂); 34.5 (1C, C(CH₃)₂); 39.6 (1C, CH₂); 52.9 (1C, H₂NCH₂) ppm. HRMS (ESI): calcd for C₁₂H₂₈N [M + H]⁺ 186.2216, found 186.2217.

Preparation of ammonium iodides 13

General procedure. To a suspension of K₂CO₃ (1.2 equiv.) in anhydrous ethanol, amine **9** (1.0 equiv.) was added and the mixture was cooled to 0 °C in an ice bath. Then MeI (3.2 equiv.) was slowly added to it, and the mixture was allowed to warm to r.t. and stirred overnight. Solids were filtered off and the solvent was evaporated. The crude mixture was dissolved in a small amount of methanol, then Et₂O was added to it and white precipitate was collected, washed with Et₂O and dried under vacuum.

Trimethylphenylammonium iodide (13a). From aniline (**9a**, 0.51 g; 5.0 mmol), K₂CO₃ (0.92 g; 6.4 mmol), and MeI (2.4 g; 17 mmol), ammonium iodide **13a** (1.4 g, 60%) was obtained according to the general procedure. ¹H NMR (399.94 MHz, DMSO-*d*₆): δ 3.61 (s, 9H, N(CH₃)₃); 7.22–8.26 (m, 5H, Ar–H) ppm.³⁵

Benzyltrimethylammonium iodide (13b). From benzylamine (**9b**, 0.52 g; 4.6 mmol), K₂CO₃ (0.80 g; 5.6 mmol), and MeI (2.1 g; 15 mmol), ammonium iodide **13b** (1.3 g, 100%) was obtained according to the general procedure. ¹H NMR (399.94 MHz, DMSO-*d*₆): δ 3.04 (s, 9H, N(CH₃)₃); 4.58 (s, 2H, CH₂); 7.53 (5H; Ar–H) ppm.³⁶

Trimethylneopentylammonium iodide (13c). From neopentylamine (**9c**, 5.0 g; 57 mmol), K₂CO₃ (9.5 g; 68 mmol) and MeI (26 g; 0.18 mol), ammonium iodide **13c** (10.7 g, 73%) was obtained according to the general procedure. ¹H NMR (399.94 MHz, DMSO-*d*₆): δ 1.12 (s, 9H, C(CH₃)₃); 3.14 (s, 2H, CH₂); 3.28 (s, 9H, N(CH₃)₃) ppm.³⁷

(2,2-Dimethyldecyl)trimethylammonium iodide (13d). From neododecylamine **9d** (3.3 g; 18 mmol), K₂CO₃ (3.3 g; 22 mmol) and MeI (8.1 g; 57 mmol), ammonium iodide **13d** (4.0 g, 66%) was obtained according to the general procedure. ¹H NMR (399.94 MHz, DMSO-*d*₆): δ 0.85 (t, $^3J_{H-H}$ = 7.0 Hz, 3H, CH₃);



1.10 (s, 6H, C(CH₃)₂); 1.26 (m, 14H, CH₂); 3.15 (s, 9H, N(CH₃)₃); 3.28 (s, 2H, NCH₂) ppm. ¹³C NMR (100.58 MHz, DMSO-*d*₆): δ 14.0 (1C, CH₃); 22.1 (1C, CH₂); 23.0 (1C, CH₂); 26.6 (2C, C(CH₃)₂); 28.7 (1C, CH₂); 28.9 (1C, CH₂); 29.7 (1C, CH₂); 31.3 (1C, CH₂); 35.5 (1C, C(CH₃)₂); 42.3 (1C, CH₂); 54.6 (t, $^1J_{C-N}$ = 4.0 Hz, 3C, (CH₃)₃N⁺); 74.1 (1C, NCH₂) ppm. HRMS (ESI): calcd for C₁₅H₃₄N [M - I]⁺ 228.2686, found 228.2687.

Preparation of ammonium fluorides 8

General procedure. Ammonium iodide **13** (1.0 equiv.) was dissolved in anhydrous methanol (10 mL) and slowly added to a suspension of AgF (1.1 equiv.) in anhydrous methanol (10 mL). The suspension was stirred for 1 h, and then the solution was filtered first through Celite and then through filter paper to yield a clear solution. After evaporation of the solvent, solid ammonium fluoride **8** was dried under vacuum (0.1 kPa) for 14 days while heated to 50 °C. Dry fluorides **8** were stored in a glove box.

Trimethylphenylammonium fluoride (8a). From ammonium iodide **13a** (1.0 g; 3.8 mmol) and AgF (0.53 g; 4.2 mmol), ammonium fluoride **8a** (0.41 g, 61%) was obtained according to the general procedure. ¹H NMR (399.94 MHz, DMSO-*d*₆): δ 3.64 (s, 9H, N(CH₃)₃); 7.22–8.27 (m, 5H, Ar-H) ppm. ¹⁹F NMR (376.29 MHz, DMSO-*d*₆): δ -101.6 (bs, 1F, F⁻) ppm. ¹³C NMR (100.58 MHz, CD₃OD): δ 57.7 (3C, (CH₃)₃N); 121.0 (1C, Ar-CH); 131.7 (2C, Ar-CH); 131.7 (2C, Ar-CH); 148.6 (1C, Ar-C) ppm. HRMS (ESI⁺): calcd for C₉H₁₄N [M - F]⁺ 136.1121, found 136.1121. HRMS (ESI⁻): calcd for C₁₈H₁₅F₂Si [M]⁻ 297.0917, found 297.0918.

Benzyltrimethylammonium fluoride (8b). From ammonium iodide **13b** (1.0 g; 3.6 mmol) and AgF (0.50 g; 3.9 mmol), ammonium fluoride **8b** was obtained (0.25 g, 40%) according to the general procedure. ¹H NMR (399.94 MHz, DMSO-*d*₆): δ 3.02 (s, 9H, N(CH₃)₃); 4.54 (s, 2H, CH₂); 7.53 (5H, Ar-H) ppm. ¹⁹F NMR (376.29 MHz, DMSO-*d*₆): δ -108.24 (s, 1F, F⁻) ppm. ¹³C NMR (100.58 MHz, CD₃OD): δ 53.1 (t, $^1J_{C-N}$ = 4.2 Hz, 3C, N(CH₃)₃); 70.5 (t, $^1J_{C-N}$ = 2.8 Hz, CH₂, 1C); 129.1 (1C, Ar-C); 130.4 (2C, Ar-CH); 132.0 (1C, Ar-CH); 134.0 (2C, Ar-CH) ppm.

Trimethylneopentylammonium fluoride (8c). From ammonium iodide **13c** (1.0 g; 3.9 mmol) and AgF (0.54 g; 4.3 mmol), ammonium fluoride **8c** was obtained (0.45 g, 78%) according to the general procedure. ¹H NMR (399.94 MHz, DMSO-*d*₆): δ 1.12 (s, 9H, C(CH₃)₃); 3.14 (s, 2H, CH₂); 3.28 (s, 9H, N(CH₃)₃) ppm. ¹⁹F NMR (376.29 MHz, DMSO-*d*₆): δ -102.11 (bs, 1F, F⁻) ppm.³⁸

(2,2-Dimethyldecyl)trimethylammonium fluoride (8d). From ammonium iodide **13d** (1.0 g; 2.8 mmol) and AgF (0.41 g; 3.3 mmol), ammonium fluoride **8d** was obtained (0.35 g, 48%) according to the general procedure. ¹H NMR (399.94 MHz, CD₃OD): δ 0.89 (t, $^3J_{H-H}$ = 7.0 Hz, 3H, CH₃); 1.22 (s, 6H, C(CH₃)₂); 1.37 (m, 14H, CH₂); 3.25 (s, 9H, N(CH₃)₃); 3.37 (s, 2H, NCH₂) ppm. ¹⁹F NMR (376.29 MHz, CD₃OD): δ -152.09 (s, 1F, F⁻) ppm. ¹³C NMR (100.58 MHz, CD₃OD): δ 14.5 (1C, CH₃); 23.7 (1C, CH₂); 24.7 (1C, CH₂); 27.4 (2C, C(CH₃)₂); 30.5 (1C, CH₂); 30.7 (1C, CH₂); 31.4 (1C, CH₂); 33.1 (1C, CH₂); 37.1 (1C, CH₂); 44.0 (1C, C(CH₃)₂); 56.1 (t, $^1J_{C-N}$ = 4.0 Hz, 3C, (CH₃)₃N);

77.0 (t, $^1J_{C-N}$ = 2.2 Hz, 1C, NCH₂) ppm. HRMS (ESI): calcd for C₁₅H₃₄N [M - F]⁺ 228.2686, found 228.2687.

Preparation of difluorosilicates 7

General procedure. To a solution of ammonium fluoride **8** (1.0 equiv.) in anhydrous MeCN (2 mL), a solution of fluorotriphenylsilane (**14**) (1.0 equiv.) in anhydrous MeCN (2 mL) was added and the mixture was stirred for 1 h. Then the solvent was evaporated to yield crude difluorosilicate **7**, which was purified by dissolving the white powder in CH₂Cl₂, precipitating it with hexane and drying *in vacuo*.

Trimethylphenylammonium difluorotriphenylsilicate (7a). From ammonium fluoride **8a** (100 mg; 0.36 mmol) and Ph₃SiF (**14**, 56 mg; 0.36 mmol), difluorosilicate **7a** was obtained according to the general procedure (145 mg, 93%). ¹H NMR (399.94 MHz, DMSO-*d*₆): δ 7.04–7.16 (m, 9H, Ar-H); 7.51–7.69 (m, 5H, Ar-H); 7.83–8.02 (m, 6H, Ar-H) ppm. ¹⁹F NMR (376.29 MHz, DMSO-*d*₆): δ -97.85 (s, 2F, Si-F₂) ppm. ¹³C NMR (100.58 MHz, DMSO-*d*₆): δ 56.4 (3C, N(CH₃)₃); 120.4 (2C, Ar-CH); 126.0 (6C, Ar-CH); 126.3 (6C, Ar-CH); 130.1 (3C, Ar-CH); 137.1 (3C, Ar-CH); 147.3 (1C, Ar-CH); 150.5 (3C, Ar-CH) ppm. HRMS (ESI⁺): calcd for C₉H₁₄N [M]⁺ 136.1121, found 136.1121. HRMS (ESI⁻): calcd for C₁₈H₁₅F₂Si [M]⁻ 297.0917, found 297.0918.

Benzyltrimethylammonium difluorotriphenylsilicate (7b). From ammonium fluoride **8b** (209 mg; 0.75 mmol) and Ph₃SiF (**14**, 127 mg; 0.75 mmol), difluorosilicate **7b** was obtained according to the general procedure (244 mg, 73%). ¹H NMR (399.94 MHz, DMSO-*d*₆): δ 2.99 (s, 9H, CH₃); 4.49 (s, 2H, CH₂); 7.05–7.16 (m, 10H, Ar-H); 7.49–7.57 (m, 5H, Ar-H); 7.85–7.89 (m, 5H, Ar-H) ppm. ³¹¹⁹F NMR (376.29 MHz, DMSO-*d*₆): δ -97.82 (s, F, Si-F₂) ppm. ³¹¹³C NMR (100.58 MHz, DMSO-*d*₆): δ 51.7 (t, $^1J_{C-N}$ = 3.6 Hz, 3C, N(CH₃)₃); 67.8 (t, $^1J_{C-N}$ = 2.6 Hz, 1C, NCH₂); 126.1 (6C, Ar-CH); 126.3 (6C, Ar-CH); 128.3 (1C, Ar-C); 129.0 (2C, Ar-CH); 130.3 (2C, Ar-CH); 132.8 (1C, Ar-CH); 137.0 (3C, Ar-CH); 150.5 (t, $^3J_{C-F}$ = 41.1 Hz, 3C, Ar-C) ppm. HRMS (ESI⁺): calcd for C₁₀H₁₆N [M]⁺ 150.1277, found 150.1278. HRMS (ESI⁻): calcd for C₁₈H₁₅F₂Si [M]⁻ 297.0917, found 297.0919.

Trimethylneopentylammonium difluorotriphenylsilicate (7c). From ammonium fluoride **8c** (107 mg; 0.63 mmol) and Ph₃SiF (**14**, 176 mg; 0.63 mmol), difluorosilicate **7c** was obtained according to the general procedure (216 mg, 80%). ¹H NMR (399.94 MHz, DMSO-*d*₆): δ 1.10 (s, 9H, C(CH₃)₃); 3.11 (s, 9H, N(CH₃)₃); 3.24 (s, 2H, CH₂); 7.09 (m, 9H, Ar-H); 7.80–7.89 (m, 6H, Ar-H) ppm. ²⁷¹⁹F NMR (376.29 MHz, DMSO-*d*₆): δ -97.31 (bs, 2F, Si-F₂) ppm. ²⁷¹³C NMR (100.58 MHz, DMSO-*d*₆): δ 29.3 (3C, C(CH₃)₃); 32.9 (1C, C(CH₃)₃); 54.5 (t, $^1J_{C-N}$ = 3.8 Hz, 3C, N(CH₃)₃); 75.5 (t, $^1J_{C-N}$ = 2.3 Hz, 1C, NCH₂); 126.1 (6C, Ar-CH); 126.3 (6C, Ar-CH); 137.0 (3C, Ar-CH); 150.8 (m, 3C, Ar-C) ppm. HRMS (ESI⁺): calcd for C₈H₂₀N [M]⁺ 130.1590, found 130.1591. HRMS (ESI⁻): calcd for C₁₈H₁₅F₂Si [M]⁻ 297.0917, found 297.0915.

(2,2-Dimethyldecyl)trimethylammonium difluorotriphenylsilicate (7d). From ammonium fluoride **8d** (107 mg; 0.63 mmol) and Ph₃SiF (**14**, 176 mg; 0.63 mmol), difluorosilicate **7d** was



obtained according to the general procedure (216 mg, 80%). ^1H NMR (399.94 MHz, DMSO- d_6): δ 0.86 (t, $^3J_{\text{H-H}} = 6.8$ Hz, 3H, CH_3); 1.09 (s, 6H, $\text{C}(\text{CH}_3)_2$); 1.19–1.37 (m, 14H, CH_2); 3.13 (s, 9H, $\text{N}(\text{CH}_3)_3$); 3.26 (s, 2H, NCH_2); 7.01–7.16 (m, 10H, Ar–H); 7.84–7.89 (m, 5H, Ar–H) ppm. ^{19}F NMR (376.29 MHz, DMSO- d_6): δ −97.83 (bs, 2F, Si–F₂) ppm. ^{13}C NMR (100.58 MHz, DMSO- d_6): δ 14.0 (1C, CH_3CH_2); 22.1 (1C, CH_2); 23.0 (1C, CH_2); 26.7 (1C, CH_2); 28.7 (2C, $\text{C}(\text{CH}_3)_2$); 29.0 (1C, CH_2); 29.8 (1C, CH_2); 31.3 (1C, CH_2); 35.5 (1C, CH_2); 42.4 (1C, $\text{C}(\text{CH}_3)_2$); 54.7 (t, $^1J_{\text{C-N}} = 4.2$ Hz, 3C, $\text{N}(\text{CH}_3)_3$); 74.3 (1C, NCH_2); 126.0 (s, 6C, Ar–CH); 126.3 (s, 6C, Ar–CH); 137.0 (3C, Ar–CH) ppm. HRMS (ESI $^+$): calcd for $\text{C}_{15}\text{H}_{34}\text{N}$ [M] $^+$ 228.2686, found 228.2686. HRMS (ESI $^-$): calcd for $\text{C}_{18}\text{H}_{15}\text{F}_2\text{Si}$ [M] $^-$ 297.0917, found 297.0918.

Fluorinations

General procedure. A 5 mL Schlenk flask was charged with the fluorination reagent (2.0 equiv.), substrate (1.0 equiv., 20 mg) and CD_3CN (0.7 mL). The flask was sealed and heated on a metallic block to 85 °C for 24 h. After cooling, the samples were measured by NMR (16 scans, 20 s relaxation delay) and the conversions were determined from the characteristic peaks of the products **14** and **15** given below.

1-Fluorooctane (6a). ^1H NMR (399.94 MHz, CD_3CN): δ 4.44 (dt, $^2J_{\text{H-F}} = 47.5$ Hz, $^3J_{\text{H-H}} = 6.2$ Hz, 2H, CH_2F) ppm.

2-Fluoroctane (6b). ^1H NMR (399.94 MHz, CD_3CN): δ 4.64 (ddqd, $^2J_{\text{H-F}} = 49.3$ Hz, $^3J_{\text{H-H}} = 7.5$ Hz, $^3J_{\text{H-H}} = 6.2$ Hz, $^3J_{\text{H-H}} = 4.7$ Hz, 1H, CHF) ppm.

Oct-1-ene (15a). ^1H NMR (399.94 MHz, CD_3CN): δ 5.85 (ddt, $^3J_{\text{H-H}} = 17.0$ Hz, $^3J_{\text{H-H}} = 10.2$, $^3J_{\text{H-H}} = 6.7$ Hz, 1H, $\text{CH}=\text{CH}_2$) ppm.

Oct-2-ene (15b). ^1H NMR (399.94 MHz, CD_3CN): δ 5.33–5.53 (m, 2H, $\text{CH}=\text{CH}$) ppm.

Attempted preparation of tetrabutylammonium difluoromethylidiphenylsilicate (16). An NMR tube was charged with fluoromethylphenylsilane **18** (30 mg, 27 μL , 0.14 mmol) and CD_3CN , followed by the addition of a solution of “anhydrous” TBAF in anhydrous THF (1 M, 140 μL , 0.14 mmol). The mixture was stirred for 1 h at RT and the ^{19}F NMR spectrum was recorded at RT and at −40 °C. At −40 °C, the minor signal of silicate **16** was observed (−86.4 ppm, bs, 24%), while the major signal corresponded to the signal of the HF_2^- anion (−151.0, bd, $^1J_{\text{H-F}} = 128$ Hz, 76%).

Attempted preparation of tetrabutylammonium difluorodimethylphenylsilicate (17). An NMR tube was charged with fluoromethylphenylsilane **19** (20 mg, 27 μL , 0.14 mmol) and anhydrous CD_3CN , followed by the addition of a solution of “anhydrous” TBAF in anhydrous THF (1 M, 140 μL , 0.14 mmol). The mixture was stirred for 1 h at RT and the ^{19}F NMR spectrum was recorded at RT and at −40 °C.

Computational details

Preliminary computations were performed using the Gaussian16³⁹ program suite using the pure M-06L functional,⁴⁰ which enabled the use of the RI (resolution of identity) approach,⁴¹ together with the double- ζ def2-SVP basis

set,⁴² which greatly accelerated the initial calculations. To better describe the anionic structures, we also used the double- ζ def2-SVPD basis set with additional diffuse functions.⁴³ The solvent (MeCN) was simulated using the SMD variant of the IEF-PCM method.⁴⁴ The results of preliminary computations are given in the ESI.† Higher level productive calculations were accomplished using the ORCA computational program,⁴⁵ which uses the efficient RIJCOSX approximation⁴⁶ to accelerate hybrid functional computation. We employed the M06-2X hybrid functional⁴⁰ together with the minimally augmented ma-def2-TZVP basis set.⁴⁷ Weigend's universal auxiliary basis set was used for the RI approximation calculations.⁴⁸ The MeCN solvent was simulated with the CPCM method⁴⁹ and the description of the non-covalent interactions was improved by the dispersion correction.⁵⁰

Results of the preliminary computations, copies of NMR spectra and xyz files of all computed structures are given in the ESI.†

Conclusions

We synthesized a series of new quaternary ammonium fluorides **8** and difluorotriphenylsilicates **7** resistant against β -elimination and compared their reactivity with that of standard nucleophilic fluorination reagents, TBAF (**1**), TBAT (**2**) and TASF (**3**), in the reaction with standard C_8 primary and secondary substrates, mesylates and bromides. We found that the reported²⁴ high excess of the TBAT reagent is not essential and a two-fold excess of the reagent is fully sufficient for successful fluorination. TBAF and TASF displayed significantly worse selectivity than TBAT. Newly prepared difluorosilicates **7** were more selective reagents than fluorides **8**, with benzyltrimethylammonium difluorotriphenylsilicate **7b** showing reactivity and selectivity close to those of TBAT (**2**). As an advantage, the synthesis of the newly prepared difluorosilicates **7** avoids the use of corrosive HF or expensive anhydrous TBAF. Although we were not able to isolate alkylated difluorosilicates, we observed the signal of tetrabutylammonium difluoromethylidiphenylsilicate (**16**) in the low temperature ^{19}F NMR spectra together with that of the HF_2^- anion. Analogous formation of tetrabutylammonium difluorodimethylphenylsilicate (**17**) was not observed. In contrast to that, DFT calculations showed that methylated difluorosilicates should be stable compounds and the discrepancy between theoretical and experimental data will be further studied. Transition state energies of the decomposition of difluorosilicates **20–23** increase with an increasing amount of phenyl substituents.

Conflicts of interest

There are no conflicts to declare.

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