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# Lattice capacity-dependent activity for CO<sub>2</sub> methanation: crafting Ni/CeO<sub>2</sub> catalysts with outstanding performance at low temperatures†

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In the pursuit of understanding lattice capacity threshold effects of oxide solid solutions for their supported Ni catalysts, a series of Ca<sup>2+</sup>-doped CeO<sub>2</sub> solid solutions with 10 wt% Ni loading (named Ni/Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub>) was prepared using a sol-gel method and used for CO<sub>2</sub> methanation. The lattice capacity of Ca<sup>2+</sup> in the lattice of CeO<sub>2</sub> was firstly determined by the XRD extrapolation method, corresponding to a Ca/(Ca + Ce) molar ratio of 11%. When the amount of Ca<sup>2+</sup> in the Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> supports was close to the CeO<sub>2</sub> lattice capacity for Ca<sup>2+</sup> incorporation, the obtained Ni/Ca<sub>0.1</sub>Ce<sub>0.9</sub>O<sub>y</sub> catalyst possessed the optimal intrinsic activity for CO<sub>2</sub> methanation. XPS, Raman spectroscopy, EPR and CO<sub>2</sub>-TPD analyses revealed the largest amount of highly active moderate-strength alkaline centers generated by oxygen vacancies. The catalytic reaction mechanisms were revealed using *in situ* IR analysis. The results clearly demonstrated that the structure and reactivity of the Ni/Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> catalyst exhibited the lattice capacity threshold effect. The findings offer a new venue for developing highly efficient oxide-supported Ni catalysts for low-temperature CO<sub>2</sub> methanation reaction and enabling efficient catalyst screening.

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## 1. Introduction

With the significant consumption of carbon-based energy sources, the amount of CO<sub>2</sub> emitted into the atmosphere is continually increasing and has caused great damage to the environment.<sup>1-4</sup> Currently, to achieve the goal of carbon neutrality, the relevant research around the capture, conversion and utilization of CO<sub>2</sub> has received extensive attention.<sup>5-7</sup> In this context, CO<sub>2</sub> methanation reaction has attracted much attention due to its potential to enable the reaction of industrial waste CO<sub>2</sub> with sustainable hydrogen gas to obtain high value-added methane and alleviate environmental problems.<sup>8-12</sup> CO<sub>2</sub> methanation reaction is an extremely exothermic reaction, presenting significant kinetic chal-

lenges.<sup>8</sup> Thus, there is a strong desire to develop cost-effective catalysts capable of achieving high activity at low temperatures for practical applications (Fig. 1a). Compared to extensively studied supported precious metals, Ni-based catalysts offer several distinct advantages including a readily available source, cost-effectiveness, and high performance.<sup>13-18</sup> Thus, oxide-supported Ni-based catalysts have garnered significant attention for CO<sub>2</sub> methanation.<sup>11,16-18</sup>

Among the various Ni-based catalysts,<sup>19-25</sup> CeO<sub>2</sub> support has been widely employed due to its abundant oxygen vacancies and robust oxygen storage capacity,<sup>18,26-30</sup> which holds the potential to enhance the low-temperature reaction activity.<sup>11,18,21,29</sup> In the pursuit of preparing high and low temperature-active catalysts, strategies such as hydrogen-rich stream,<sup>23</sup> modifying with Ce,<sup>24</sup> and constructing interfaces<sup>25</sup> have shown promising results.

In addition, oxide solid solutions offer a promising avenue for enhancing the low-temperature activity.<sup>11,31</sup> It has been reported that metal oxide solid solutions exhibit distinct properties. When compared to their individual components, they typically demonstrate improved and stimulating physico-chemical characteristics, thereby exerting a positive impact on the catalytic performance in various reactions.<sup>32-36</sup> The creation of a solid solution structure can lead to substantial enhancements in the thermal endurance, porous structure and abundance of surface-active oxygen sites. These improve-

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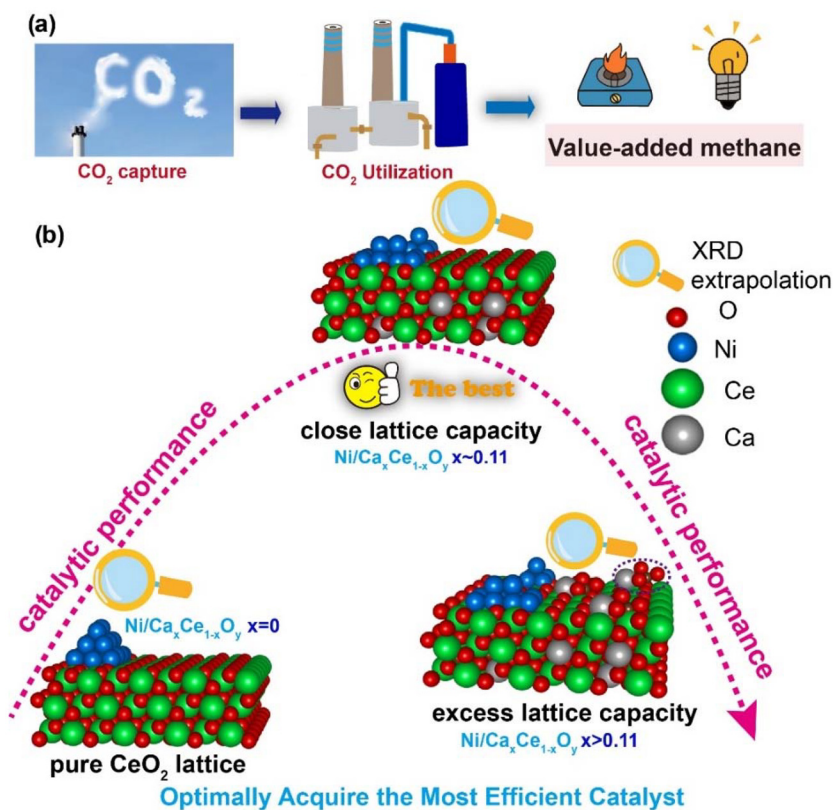
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**Fig. 1** Overview of this study. (a) Hydrogenating industrial CO<sub>2</sub> emissions with renewable hydrogen to produce high-value methane to achieve carbon neutrality. (b) In a groundbreaking approach, the lattice capacity of Ca<sup>2+</sup> in CeO<sub>2</sub> was precisely quantified utilizing the XRD extrapolation method.

ments are usually attributed to lattice distortion and the generation of lattice defects, both of which typically contribute to enhanced catalytic performance.<sup>37–39</sup> In our previous work, Ce–M–O solid solutions were investigated, where M represents cations such as Mg, Ca, Sr, and Ba.<sup>11</sup> The creation of stable solid solution arrangement structures for M, such as Mg<sup>2+</sup> and Ca<sup>2+</sup>, boasts a notably higher amount of surface oxygen vacancies and more moderate alkaline sites when compared to pure CeO<sub>2</sub>, which not only substantially augments the oxygen storage capacity but also significantly boosts the turnover frequency and enhances the low-temperature reaction activity for CO<sub>2</sub> methanation. Conversely, the alkaline earth metals Sr and Ba, with their large ionic radii, cannot fit into the ceria lattice to form solid solutions, which leads to fewer active oxygen vacancies. Instead, they predominantly accumulated on the support surface as highly alkaline oxides and carbonates. This excessive CO<sub>2</sub> adsorption capacity resulted in poor catalytic performance.<sup>40</sup> Solid solution catalysts have been extensively investigated and have frequently exhibited enhanced activity and stability when compared to individual metal oxides.<sup>10,26,34,37,38</sup>

In our previous publications, we devised an XRD extrapolation technique to quantify the lattice capacity of a solute cation within the lattice of a solvent metal oxide in oxide solid solutions<sup>39,40</sup> and discovered that pure oxide solid solutions act as catalysts with a lattice capacity threshold effect on

reactions.<sup>10,26,36,37</sup> This study marks a significant departure from our prior research on solid solution catalysts as it is based on our preliminary discovery that many metal oxide solid solutions can themselves act as catalysts, exhibiting a lattice capacity threshold effect in specific reactions. In this work, we delve into the lattice capacity threshold effect in supported catalysts, with a particular focus on Ni catalysts supported on Ca<sup>2+</sup>-doped CeO<sub>2</sub> solid solutions (Fig. 1b). A series of Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> supports with diverse contents of calcium oxide was prepared by a sol-gel method ( $x = 0.05–0.55$ ), and the supported Ni materials were obtained by an incipient impregnation method with 10 wt% Ni loading. The XRD extrapolation method was adopted to measure the lattice capacity; the activity of the Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub>-supported Ni catalysts exhibited the lattice capacity threshold effect for CO<sub>2</sub> methanation. The nature of the threshold effect was revealed by various characterization methods, such as XRD, H<sub>2</sub>-TPR, Raman, XPS, and CO<sub>2</sub>-TPD.

## 2. Experimental

### 2.1 Catalyst preparation

A series of Ca<sup>2+</sup>-doped CeO<sub>2</sub> supports with varying amounts of Ca cations was synthesized using a sol-gel method. Initially, Ce(NO<sub>3</sub>)<sub>3</sub>·6H<sub>2</sub>O (Shanghai Macklin, A.R) and Ca(NO<sub>3</sub>)<sub>2</sub>·4H<sub>2</sub>O

(A.R.) were dissolved in 50 mL solution, with different Ca/Ce molar ratios named  $x$ . Subsequently, citric acid (Shanghai Macklin, A.R.), equivalent to 0.6 times the molar ratio of  $\text{Ce}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$  and  $\text{Ca}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O}$ , was added. Next, the citric acid was burned off at 220 °C. Subsequently, the material was calcined at a heating rate of 2 °C  $\text{min}^{-1}$ , reaching 700 °C, and was held at this temperature for 4 h to produce  $\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$ . It should be noted that the synthesis protocol for  $\text{CeO}_2$  was identical to that of  $\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$ , with the sole distinction being the absence of Ca metal.

The 10 wt%  $\text{NiO}/\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$  materials were synthesized using an impregnation method. Typically, a specific quantity of  $\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$  (*vide supra*) was dropped into  $\text{Ni}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$  solution and stirred for 2 h. Subsequently, it was dried at 110 °C. Finally, the material was calcined at 450 °C to obtain  $\text{NiO}/\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$ . The final catalysts were defined as  $\text{Ni}/\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$ . The samples underwent reduction with a 10%  $\text{H}_2/\text{Ar}$  gas mixture, maintained under 450 °C. The chemical compositions of the samples were validated *via* ICP-AES analysis.

## 2.2 Activity evaluation

$\text{CO}_2$  methanation evaluation using 10 wt%  $\text{Ni}/\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$  catalysts was carried out in the following manner. Initially, the catalyst powder was mixed with an appropriate amount of quartz sand and subjected to compression molding (catalyst : quartz sand = 1 : 1). 100 mg of fresh catalyst underwent *in situ* reduction at 450 °C for 2 h with a gas mixture comprising 10%  $\text{H}_2$  and 90% Ar. Subsequently, the temperature was reduced to room temperature, followed by the introduction of high-purity  $\text{H}_2$  (99.99%) and high-purity  $\text{CO}_2$  (99.99%) ( $\text{H}_2/\text{CO}_2 = 4$ ).  $\text{CO}_2$  and  $\text{H}_2$  reactants were mixed in 1 : 4 stoichiometric ratio, resulting in a WHSV of 36 000 mL per h per gram of the catalyst. The outlet gas was condensed, then analyzed online using a GC9310 gas chromatograph. High-purity argon gas (99.99%) was employed as the carrier gas in chromatography. A TDX-01 molecular sieve column with an inner diameter of 2 mm and a length of 3 m was utilized for the separation of  $\text{H}_2$ ,  $\text{CO}$ ,  $\text{CH}_4$ , and  $\text{CO}_2$ . The tail gas was subsequently passed through a gas chromatograph equipped with a TDX-01 column, and carbon conservation was applied using the peak area normalization method to calculate the conversion rate of the reaction. The detection of the separated substances (reactants and products) post-column separation was conducted using a thermal conductivity detector (TCD) in the GC9310 chromatograph. Before the test, steady-state kinetic data measurements were conducted. The  $\text{CO}_2$  conversion ( $X_{\text{CO}_2}$ ),  $\text{CH}_4$  selectivity ( $S_{\text{CH}_4}$ ) and  $\text{CH}_4$  yield ( $Y_{\text{CH}_4}$ ) were calculated as follows.

$$X(\text{CO}_2) = \frac{n(\text{CH}_4)_{\text{out}} + n(\text{CO})_{\text{out}}}{n(\text{CH}_4)_{\text{out}} + n(\text{CO}_2)_{\text{out}} + n(\text{CO})_{\text{out}}} \times 100\%$$

$$S(\text{CH}_4) = \frac{n(\text{CH}_4)_{\text{out}}}{n(\text{CH}_4)_{\text{out}} + n(\text{CO})_{\text{out}}} \times 100\%$$

$$Y(\text{CH}_4) = X(\text{CO}_2) \times S(\text{CH}_4)$$

where  $n(\text{CO}_2)$ ,  $n(\text{CH}_4)$ , and  $n(\text{CO})$  represent the respective molar amounts of each gas.

## 2.3 Catalyst characterization

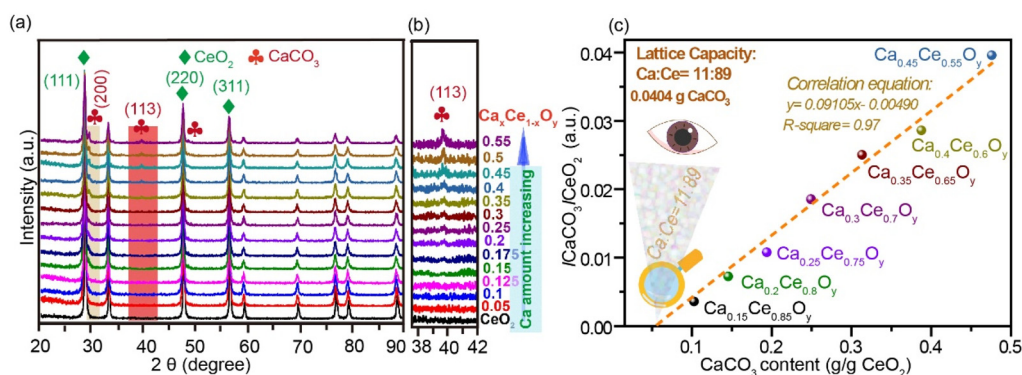
The XRD patterns of the samples were acquired using a Bruker AXS D8Focus X-ray diffractometer instrument with  $\text{Cu K}\alpha$  radiation (operating at 40 kV and 30 mA). Scans spanned from 20° to 90° ( $2\theta$ ) at a 2° per minute ramp rate. Mean crystallite sizes were calculated using the Scherrer equation ( $d = K\lambda/(B1/2 \cos \theta)$ ,  $K = 0.89$ ,  $\lambda = 1.5406 \text{ \AA}$ ). The Raman spectra were obtained with an argon laser excitation source at 532 nm.  $\text{CO}_2$ -TPD was performed using a Micromeritics Auto Chem 2920 apparatus. High-purity helium was employed as the carrier and pretreatment gas, while high-purity  $\text{CO}_2$  served as the adsorption gas. The reduced  $\text{Ni}/\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$  catalysts were pre-heated to 400 °C, then cooled to room temperature, and purged with ultrahigh purity He for 0.5 h. Afterwards, the instrument program was ramped up to 700 °C. XPS was performed at room temperature under ultra-high vacuum conditions (300 W and 15 kV). EPR test was done at 77 K using a JEOL FA-200 EPR with 100 kHz field modulation and microwave frequency of 9067.558 MHz.  $\text{H}_2$  temperature-programmed desorption ( $\text{H}_2$ -TPD) was conducted using a Micromeritics Auto Chem 2920 apparatus. Initially, 50 mg of the sample underwent heating at 400 °C for 30 min to eliminate any potential impurities. Subsequently, the sample was cooled to 50 °C and exposed to a flow of  $\text{H}_2$  at 30  $\text{mL min}^{-1}$  for 1 h to saturate the surface. This was followed by purging with ultrahigh purity Ar at the same flow rate for 30 min to remove any physically adsorbed  $\text{H}_2$ . Following these pretreatments, the catalyst was heated from 50 to 700 °C at a rate of 10 °C  $\text{min}^{-1}$ .

$\text{H}_2$ -TPR experiments were conducted on a FINESORB 3010C instrument using 50 mg of the sample. Catalysts were pre-treated in Ar flow at 120 °C for 30 min to remove impurities. Temperature was ramped from room temperature to 900 °C at 10 °C  $\text{min}^{-1}$  in a 10%  $\text{H}_2/\text{Ar}$  flow of 30  $\text{mL min}^{-1}$ .  $\text{H}_2$  consumption was monitored using a TCD with  $\text{CuO}$  (99.99%) as the calibration standard. Utilizing an *in situ* FTIR equipped with an MCT detector, capturing the intermediate species in the catalysts is a pivotal method for probing the reaction mechanisms. Before commencing experiments, the samples were pre-treated at 300 °C for 1 h in a high-purity Ar atmosphere. Subsequently, they were cooled to 50 °C, and the background spectra were obtained under Ar atmosphere. Following this, the gas was switched to a mixture of 4%  $\text{H}_2$ , 1%  $\text{CO}_2$ , and 95% Ar for  $\text{CO}_2$  hydrogenation testing, with a controlled gas flow rate of 20  $\text{mL min}^{-1}$ . The temperature was ramped from 50 °C to 400 °C to monitor the dynamic changes of the intermediate species within the samples. Infrared spectra were collected once all species reached a steady state at the set temperatures.

## 3. Results and discussion

### 3.1 Quantifying the lattice capacity of $\text{Ca}^{2+}$ in $\text{CeO}_2$ through the XRD extrapolation method

XRD experiments were used to analyze  $\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$  phases and  $\text{Ca}^{2+}$  lattice capacity in  $\text{CeO}_2$ . Fig. 2 and Fig. S1† display the



**Fig. 2** (a and b) Complete and enlarged XRD patterns of the  $\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$  support; (c) the lattice capacity quantification of  $\text{CeO}_2$  for  $\text{Ca}^{2+}$  cations. The lattice capacity of  $\text{Ca}^{2+}$  in the lattice of the  $\text{CeO}_2$  was firstly determined by the XRD extrapolation method, corresponding to a  $\text{Ca}/(\text{Ca} + \text{Ce})$  molar ratio of 11% (see Fig. 2, S1† and Table 1).

results, and Table 1 summarized the phase compositions. For pure  $\text{CeO}_2$ , three prominent peaks at  $2\theta$  values of  $28.90^\circ$ ,  $47.77^\circ$ , and  $56.76^\circ$  confirmed the cubic fluorite crystalline phase (Fig. 2).<sup>11</sup> The peaks at  $31.13^\circ$ ,  $42.91^\circ$ , and  $50.24^\circ$  indicated the presence of  $\text{CaCO}_3$ . Low Ca content solid solution samples (e.g.,  $\text{Ca}_{0.05}\text{Ce}_{0.95}\text{O}_y$  and  $\text{Ca}_{0.1}\text{Ce}_{0.9}\text{O}_y$ ) showed only  $\text{CeO}_2$  peaks, suggesting that  $\text{Ca}^{2+}$  incorporation into the  $\text{CeO}_2$  lattice formed a solid solution. Increasing Ca content ( $x = 0.15$  to  $0.55$ ) led to  $\text{CaCO}_3$  phase emergence and weakened  $\text{CeO}_2$  peaks (Fig. 2b), signaling excess Ca exceeding the lattice capacity and reacting with  $\text{CO}_2$ .<sup>11</sup> In Ce-rich Ce–Ca binary oxide supports with  $\text{Ca}/\text{Ce} < 15/85$  (e.g.,  $\text{Ca}_{0.05}\text{Ce}_{0.95}\text{O}_y$  and  $\text{Ca}_{0.1}\text{Ce}_{0.9}\text{O}_y$ ), only broadened  $\text{CeO}_2$  peaks were observed, indicating the  $\text{Ca}^{2+}$ -incorporated  $\text{CeO}_2$  lattice. As  $\text{Ca}/\text{Ce}$  approached  $15/85$ , subtle  $\text{CaCO}_3$  peaks emerged at  $42.91^\circ$ . For clarity, Table 1 summarizes the phase compositions of supports with varying  $\text{Ca}/\text{Ce}$  ratios. The  $\text{Ce}^{4+}$  in  $\text{CeO}_2$  possesses a coordination number (CN) of 8 and a radius measuring  $0.97 \text{ \AA}$ . The radius of  $\text{Ca}^{2+}$  cations under a coordination number (CN) of 8 is  $1.12 \text{ \AA}$ . In principle, the  $\text{Ca}^{2+}$  cations meet the required radius criteria for the formation of a solid solution structure.<sup>41,42</sup> Thus, there is high likelihood that  $\text{Ca}^{2+}$  cations were successfully integrated into the  $\text{CeO}_2$  lattice matrix to form a solid solution. To prove this hypothesis, we meticulously determined and compared the  $2\theta$  and  $d$ -values associ-

ated with the most intense (111) peak of the  $\text{CeO}_2$  phase as well as the mean crystallite sizes of  $\text{CeO}_2$  (Fig. S1† and Table 1). Evidently, the  $2\theta$  and  $d$  values in all the  $\text{Ca}^{2+}$ -doped samples deviate from those of the parent  $\text{CeO}_2$ , confirming the successful incorporation of  $\text{Ca}^{2+}$  cations into the  $\text{CeO}_2$  matrix, replacing a portion of the  $\text{Ce}^{4+}$  cations in the formed solid solution structure.

Our previous research has demonstrated that in the case of non-solid solutions, solute cations typically integrated into the matrix structure of a host metal oxide with a specific lattice capacity,<sup>41,42</sup> a quantifiable parameter that can be determined using the XRD extrapolation method developed by our team (Fig. 2c, fitting diagram). As outlined in our study, for some results of this partial substitution process, XRD analysis may not detect the presence of  $\text{CaCO}_3$  phases due to their low concentration at this stage. Once the calcium content nears its saturation point within the lattice, the catalyst surface starts to facilitate the formation of polymeric and amorphous calcium carbonate species. Nevertheless, unless the  $\text{CaCO}_3$  crystallites attained a specific crystalline dimension, typically about  $5 \text{ nm}$ , the  $\text{CaCO}_3$  phase remains undetectable by XRD analysis.<sup>11,41,42</sup> Therefore, due to this limitation, for a more precise and dependable determination of the capacity threshold of calcium in the  $\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$  solid solution, our research group has developed the XRD extrapolation method, which is widely recognized and employed.<sup>11,41,42</sup> The lattice capacity can only be extrapolated through the correlated line (Fig. 2c). Theoretically, the presence of crystalline  $\text{CaCO}_3$  was intricately linked to the overall calcium content within the Ca–Ce binary oxide catalysts. The strength or prominence of its diffraction peaks serves as an indicator of the quantity of the  $\text{CaCO}_3$  crystalline phase. The intensity of  $\text{CaCO}_3$  at the (113) crystal plane that represents the peak of  $\text{CaCO}_3$  was employed to quantify the extent of crystalline  $\text{Ca}^{2+}$  cations that entered the  $\text{CeO}_2$  lattice matrix. In detail, the intensity ( $I_0$ ) of the strongest peak of  $\text{CeO}_2$  (111) was employed to normalize the  $I$  to get a series of  $I/I_0$  ratios, which correlate to the quantity of  $\text{CaCO}_3$  crystalline phase in the samples.  $I/I_0$  ratios were plotted against the

**Table 1** XRD results of  $\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$  supports

Supports	$\text{CeO}_2$ (111) $2\theta$ ( $^\circ$ )	$d^a$ ( $\text{\AA}$ )	Phase composition	Crystalline size of $\text{CeO}_2$ (nm)
$\text{CeO}_2$	28.919	3.085	$\text{CeO}_2$	23.1
$\text{Ca}_{0.05}\text{Ce}_{0.95}\text{O}_y$	28.579	3.121	$\text{CeO}_2$ SS <sup>b</sup>	21.8
$\text{Ca}_{0.1}\text{Ce}_{0.9}\text{O}_y$	28.360	3.144	$\text{CeO}_2$ SS	18.2
$\text{Ca}_{0.125}\text{Ce}_{0.875}\text{O}_y$	28.400	3.140	$\text{CeO}_2$ SS + $\text{CaCO}_3$	19.0
$\text{Ca}_{0.25}\text{Ce}_{0.75}\text{O}_y$	28.420	3.138	$\text{CeO}_2$ SS + $\text{CaCO}_3$	19.8
$\text{Ca}_{0.4}\text{Ce}_{0.6}\text{O}_y$	28.421	3.136	$\text{CeO}_2$ SS + $\text{CaCO}_3$	20.7

<sup>a</sup> Calculated using Scherrer's equation for the diffraction pattern of (111) facets of  $\text{CeO}_2$ . <sup>b</sup> SS: solid solution.

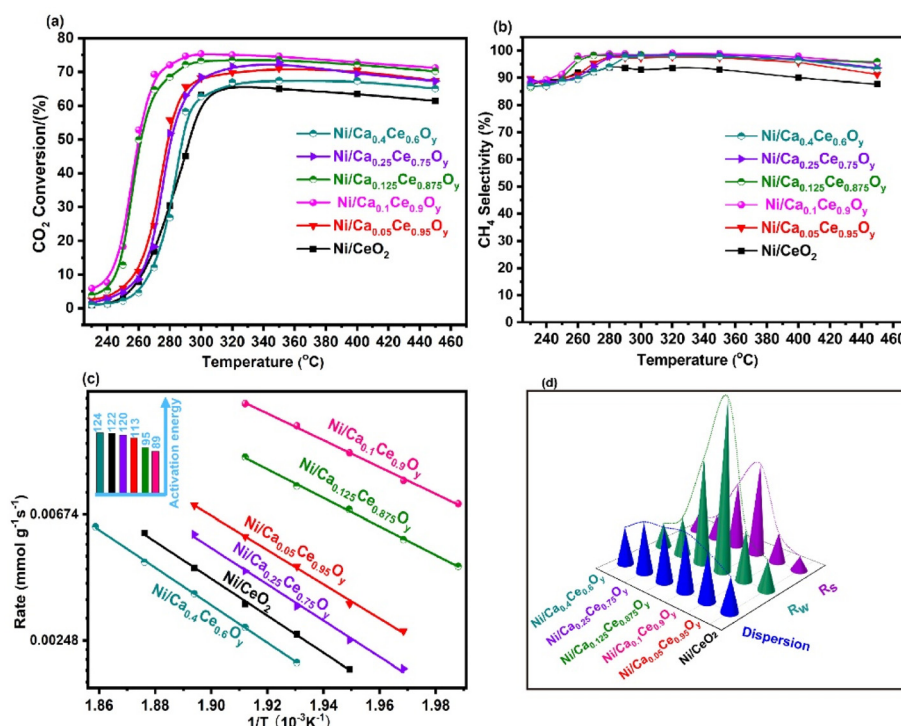
CaCO<sub>3</sub> quantity, which was normalized by the CeO<sub>2</sub> content. A correlated line was obtained, which intersects the x axis at a point having an x-axis value of 0.0404 g CaCO<sub>3</sub> per gram CeO<sub>2</sub>, which was equivalent to a Ca/Ce molar ratio of 11/89, corresponding to 11% of the Ce<sup>4+</sup> cations that can be maximally substituted by Ca<sup>2+</sup>, forming a stable solid solution (Fig. 2c). Once the Ca<sup>2+</sup> content exceeds the lattice capacity, the surplus Ca<sup>2+</sup> were could form free-state CaCO<sub>3</sub> species, aggregating and enriching on the surface. Typically, the CaCO<sub>3</sub> micro-crystallites can be clearly detected by XRD when their size reached or approached about 5 nm. Furthermore, their diffraction intensity increases with higher Ca content, as observed in the case of Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> (x = 0.15–0.55). Notably, for Ca<sub>0.05</sub>Ce<sub>0.95</sub>O<sub>y</sub> and Ca<sub>0.1</sub>Ce<sub>0.9</sub>O<sub>y</sub>, since the Ca content remained within the lattice capacity of CeO<sub>2</sub> (Fig. 2a and b), all Ca<sup>2+</sup> cations were incorporated into the lattice structure of CeO<sub>2</sub>, thereby eluding detection through XRD analysis.<sup>43,44</sup>

### 3.2 The lattice capacity threshold effect of Ca<sup>2+</sup> in CeO<sub>2</sub> on CO<sub>2</sub> methanation by Ni/Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> catalysts

To explore the lattice capacity threshold effect of these solid solutions, all the Ni/Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> catalysts were evaluated for CO<sub>2</sub> methanation (Fig. 3). For the parent Ni/CeO<sub>2</sub> catalyst, the overall activity increased as the temperature rose until reaching 350 °C, achieving a CO<sub>2</sub> conversion of 65% and CH<sub>4</sub> selectivity of 95%. Ni/Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> catalysts displayed the overall activity sequence of Ni/Ca<sub>0.1</sub>Ce<sub>0.9</sub>O<sub>y</sub> > Ni/Ca<sub>0.05</sub>Ce<sub>0.95</sub>O<sub>y</sub> > Ni/Ca<sub>0.25</sub>Ce<sub>0.75</sub>O<sub>y</sub> > Ni/

Ca<sub>0.4</sub>Ce<sub>0.6</sub>O<sub>y</sub> > Ni/CeO<sub>2</sub>; however, they have a similar CH<sub>4</sub> selectivity (Fig. 3). The Ni/Ca<sub>0.1</sub>Ce<sub>0.9</sub>O<sub>y</sub> catalyst exhibited the highest overall activity, achieving a remarkable 75% CO<sub>2</sub> conversion and complete CH<sub>4</sub> selectivity at 290 °C. The response data trends revealed that when the amount of Ca incorporated into the solid solution was below its lattice capacity, there was a gradual increase in CO<sub>2</sub> conversion. As the Ca content rose, the catalytic activity was further augmented. Conversely, when the Ca content surpassed the lattice capacity, the activity diminished with increasing Ca content. The optimal activity of the loaded Ni catalyst was achieved when the Ca doping content was aligned precisely with the lattice capacity threshold, enabling the creation of the highest possible quantity of a pure-phase solid solution on the carrier material (*i.e.*, 10 wt% Ni/Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub>). This indicated the presence of a lattice capacity threshold effect for dopant elements in oxide solid solution catalysts supporting Ni, highlighting its pivotal role in catalyst performance.

For better understanding the modification effects of varying calcium oxide content on the intrinsic activity, Arrhenius plots (Fig. 3c) were generated for all the catalysts. These plots were constructed using CO<sub>2</sub> conversion data below 10% to eliminate potential influences from mass and heat transfer effects. To provide a clearer understanding, we calculated *R<sub>w</sub>* (mass normalized rate) and *R<sub>s</sub>* (surface area normalized rate), representing CO<sub>2</sub> conversion rates at 250 °C, which are normalized by catalyst weights and surface areas and deter-



**Fig. 3** Performance assessment of CO<sub>2</sub> methanation on Ni/Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> catalysts involves (a) CO<sub>2</sub> conversion, (b) CH<sub>4</sub> selectivity, and (c) Arrhenius plots for CO<sub>2</sub> methanation. (d) The intrinsic activity and the dispersion of the catalyst. The activity tests reveal a lattice capacity threshold effect, with the maximum solid-solution pure-phase sample being at the CeO<sub>2</sub> lattice capacity threshold. This results in the generation of more high-activity oxygen vacancies, leading to its optimal catalytic performance (see Fig. 3, 5, 6 and Table 2).

**Table 2** The reactivity of the catalysts

Catalysts	$S_{\text{BET}}^a$	Dispersion <sup>b</sup> (%)	$R_w^c$ ( $10^{-3}$ mmol s <sup>-1</sup> g <sup>-1</sup> )	$R_s^c$ ( $10^{-3}$ mmol s <sup>-1</sup> m <sup>-2</sup> )	TOF <sup>d</sup> (s <sup>-1</sup> )	$E_a$ [kJ mol <sup>-1</sup> ]
Ni/CeO <sub>2</sub>	21	2.22	1.97	0.94	0.035	122.4
Ni/Ca <sub>0.05</sub> Ce <sub>0.95</sub> O <sub>y</sub>	17	3.12	3.36	1.97	0.054	113.1
Ni/Ca <sub>0.1</sub> Ce <sub>0.9</sub> O <sub>y</sub>	18	3.77	10.91	6.06	0.100	88.7
Ni/Ca <sub>0.125</sub> Ce <sub>0.875</sub> O <sub>y</sub>	16	3.54	6.92	4.32	0.088	94.5
Ni/Ca <sub>0.25</sub> Ce <sub>0.75</sub> O <sub>y</sub>	16	3.36	2.54	1.59	0.042	119.7
Ni/Ca <sub>0.4</sub> Ce <sub>0.6</sub> O <sub>y</sub>	15	2.46	1.62	1.08	0.040	124.1

<sup>a</sup> Measured by BET results. <sup>b</sup> The absolute dispersion measured by H<sub>2</sub>-TPD results. <sup>c</sup>  $R_w$  and  $R_s$  were differential rates at 250 °C, standardized based on the catalyst's weight and surface area, respectively. <sup>d</sup> Calculated based on the steady state CO<sub>2</sub> conversion at 250 °C.

mined the apparent activation energies (Table 2) for the catalysts using the formulas derived from the Arrhenius plots. In the case of the two samples containing a pure CeO<sub>2</sub> solid solution phase, Ni/Ca<sub>0.05</sub>Ce<sub>0.95</sub>O<sub>y</sub> and Ni/Ca<sub>0.1</sub>Ce<sub>0.9</sub>O<sub>y</sub>, both the  $R_w$  and  $R_s$  values show an increasing trend as the calcium oxide doping content rises. However, in contrast, among the Ni/Ca<sub>0.25</sub>Ce<sub>0.75</sub>O<sub>y</sub> and Ni/Ca<sub>0.4</sub>Ce<sub>0.6</sub>O<sub>y</sub> catalysts, these three heterogeneous solid solution samples exhibited the presence of CaCO<sub>3</sub> on the surface of the Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> support. In this case, both the  $R_w$  and  $R_s$  values showed a significant downward trend (Fig. 3d and Table 2). The observed activity sequence, which reflected the catalytic performance of Ni/Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub>, appeared to be closely linked to the location of calcium within/on the CeO<sub>2</sub> lattice. It was noteworthy that both the  $R_w$  and  $R_s$  values followed the sequence Ni/Ca<sub>0.1</sub>Ce<sub>0.9</sub>O<sub>y</sub> > Ni/Ca<sub>0.125</sub>Ce<sub>0.875</sub>O<sub>y</sub> > Ni/Ca<sub>0.05</sub>Ce<sub>0.95</sub>O<sub>y</sub> > Ni/Ca<sub>0.25</sub>Ce<sub>0.75</sub>O<sub>y</sub> > Ni/Ca<sub>0.4</sub>Ce<sub>0.6</sub>O<sub>y</sub> ≫ Ni/CeO<sub>2</sub>. These analysis results, illustrating the incorporation of Ca<sup>2+</sup> into the CeO<sub>2</sub> lattice and forming pure solid solution structure, have potential to enhance the activity compared to heterogeneous solid solution samples accompanied by CaCO<sub>3</sub> on the Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> support surface. Additionally, the activation energy results also provided additional evidence that the incorporation of CaO into the CeO<sub>2</sub> lattice can reduce the activation energy in the case of pure solid solution samples. Next, the turnover frequencies (TOF) for CO<sub>2</sub> conversion on the exposed surface Ni active sites were calculated at 250 °C (Table 2). It is noteworthy that all the solid solution Ni/Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> catalysts exhibited significantly higher TOF values (ranging from 0.040 to 0.100 s<sup>-1</sup>) compared to the unmodified Ni/CeO<sub>2</sub> catalyst (0.035 s<sup>-1</sup>).

In a briefly general summary, solid solution samples Ni/Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> exhibited a notable enhancement in their intrinsic activity. The enhanced activity indicated the presence of a lattice capacity threshold effect in the reactivity of solid solution oxides.

The pattern of presentation was closely linked to the location of Ca (*i.e.*, in/on the support) and pure solid solutions generated the highest oxygen vacancy concentration. Hence, a pure solid solution crystal phase (*i.e.*, calcium ions incorporated into the crystalline lattice structure of CeO<sub>2</sub>) and proper condense of moderate alkalinity are key factors to promote low-temperature activity.

The optimal sample's catalytic performance was subjected to a 20-hour stability test at 290 °C, during which no decline in

the catalytic stability was observed. The stability plot is depicted below, with the corresponding text and descriptions highlighted in red font within the main text. The alkaline earth metal Ca modified the CeO<sub>2</sub> support, generating active oxygen vacancies. This facilitated CO<sub>2</sub> activation, leading to improved catalyst stability (Fig. S5†).

### 3.3 Studying interactions on metal and supports through H<sub>2</sub>-TPR

To gain insight into the redox behaviors of Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> solid solution supports and the metal-support interaction, H<sub>2</sub>-TPR experiments were performed on Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> and NiO/Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> (Fig. 4 and Table S3†). The parent CeO<sub>2</sub> exhibited two distinct reduction peaks, one at approximately 480 °C and another at about 780 °C. These peaks are associated with the conversion of surface Ce<sup>4+</sup> into Ce<sup>3+</sup> and the transformation of a fraction of the bulk CeO<sub>2</sub> into Ce<sub>2</sub>O<sub>3</sub>.<sup>45,46</sup>

Compared to the unaltered CeO<sub>2</sub>, all Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> supports demonstrated similar reduction peaks. However, there was a noticeable shift of the reduction peak at ~480 °C towards a lower temperature range, suggesting the increased reducibility of the surface CeO<sub>2</sub> species. For Ni/CeO<sub>2</sub>, a distinctive peak emerges at approximately 336 °C when comparing with the CeO<sub>2</sub> support profile. This peak can be confidently ascribed to the conversion of NiO species into metallic Ni through reduction. On the contrary, the reduction peak of NiO in the profiles of the Ni/Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> catalysts shifted to higher temperatures and became broader. This phenomenon directly indicated that the incorporation of Ca cations strengthened the interplay between NiO and the Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> supports. Hence, when compared to the unmodified Ni/CeO<sub>2</sub> material, the reduced Ni/Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> catalysts exhibited a smaller metallic Ni and a higher Ni dispersion (Tables S2 and S3†). Additionally, the capacity threshold effect seemed to be evident in the interaction between the metal and the support material. However, the disparity in the strength of interaction between the metal and the support was not as pronounced; thus, the impact on the activity was relatively minor. The determination of H<sub>2</sub> consumption and the H/Ni atomic ratios of the samples were carried out after subtracting the consumption associated with the Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> supports (Table S3†). The catalysts exhibited H/Ni atomic ratios that were close to the stoichiometric value of 2. This observation substantiated that NiO

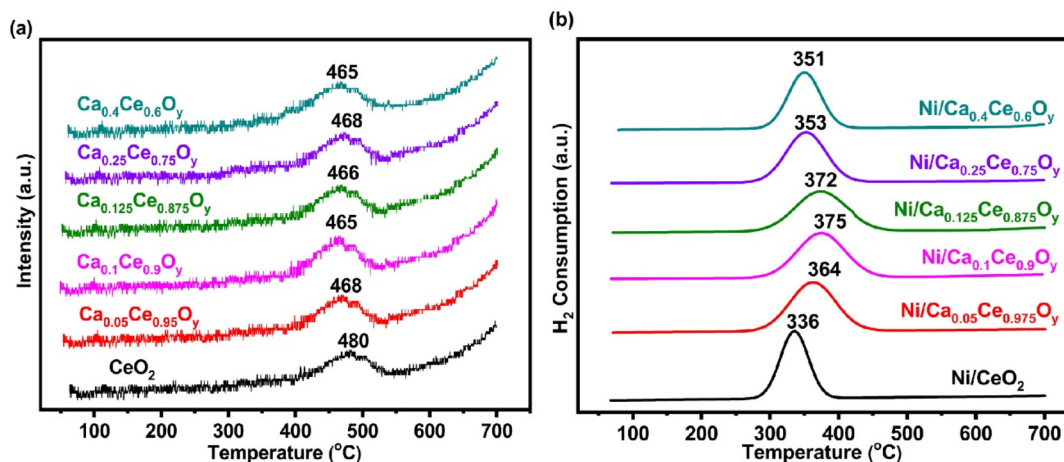


Fig. 4 H<sub>2</sub>-TPR profiles of (a) the Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> support and (b) the fresh Ni/Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> catalysts. The capacity threshold effect appeared in the interaction between the metal and the support material. Nevertheless, the difference in the strength of this interaction was not very pronounced, resulting in a relatively minor impact on the activity (see Fig. 4 and Table S3<sup>†</sup>).

was the prevailing component in all of the samples and underwent complete reduction after hydrogen reduction pretreatment.

### 3.4 Oxygen properties of supports and catalysts analyzed by Raman and EPR spectroscopy

Raman characterization was employed to study the effects of Ca modification on the structure and surface oxygen vacancies. For the Raman spectrum of Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> supports (Fig. 5a–c), three distinct peaks were observed at 460, 570, and 1068 cm<sup>-1</sup>, corresponding to octahedrally symmetrical vibration modes (F<sub>2g</sub>), vacancy-induced modes, and superoxide (O<sub>2</sub><sup>-</sup>) modes,<sup>47</sup> respectively. Although all samples exhibited the same signal, there was a significant variation in intensity. The ratio of the peak intensities at 570 nm and 460 nm, referred to as *I*<sub>570</sub>/*I*<sub>460</sub>, is commonly used as an indicative measure of the concen-

tration of oxygen vacancies in CeO<sub>2</sub>-based materials. The *I*<sub>570</sub>/*I*<sub>460</sub> values obey the order Ca<sub>0.1</sub>Ce<sub>0.9</sub>O<sub>y</sub> > Ca<sub>0.125</sub>Ce<sub>0.875</sub>O<sub>y</sub> > Ca<sub>0.05</sub>Ce<sub>0.95</sub>O<sub>y</sub> > Ca<sub>0.25</sub>Ce<sub>0.75</sub>O<sub>y</sub> > Ca<sub>0.4</sub>Ce<sub>0.6</sub>O<sub>y</sub> > CeO<sub>2</sub>, where Ca<sub>0.1</sub>Ce<sub>0.9</sub>O<sub>y</sub> displayed the most pronounced concentration of surface oxygen vacancies. This indicated that the density of oxygen vacancies in the solid solution also exhibited a lattice capacity threshold effect. Additionally, to semi-quantify the amount of O<sub>2</sub><sup>-</sup>, we calculated the *I*<sub>1068</sub>/*I*<sub>460</sub> values for Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> (Table S4<sup>†</sup>). Interestingly, the order of the *I*<sub>1068</sub>/*I*<sub>460</sub> values mirrors that of the *I*<sub>570</sub>/*I*<sub>460</sub> values, illustrating that O<sub>2</sub><sup>-</sup> species were primarily generated through the adsorption of oxygen at these oxygen vacancies.<sup>48</sup> Furthermore, the modification involving Ca<sup>2+</sup> cations induced the formation of a greater number of O<sub>2</sub><sup>-</sup> species compared to unmodified CeO<sub>2</sub>. Thus, when comparing the Ca-modified supports to the parent CeO<sub>2</sub>, the former exhibited a broader F<sub>2g</sub> Raman signal, indicating that Ca<sup>2+</sup> ions were likely incorporated into the CeO<sub>2</sub>

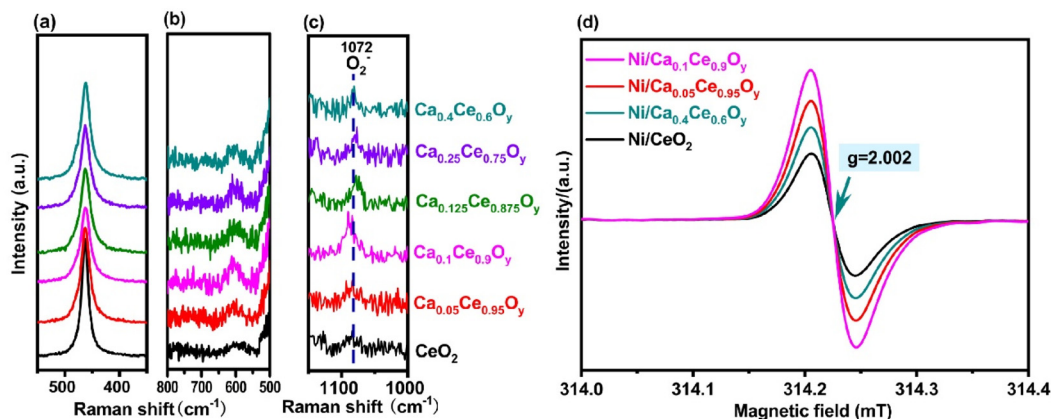


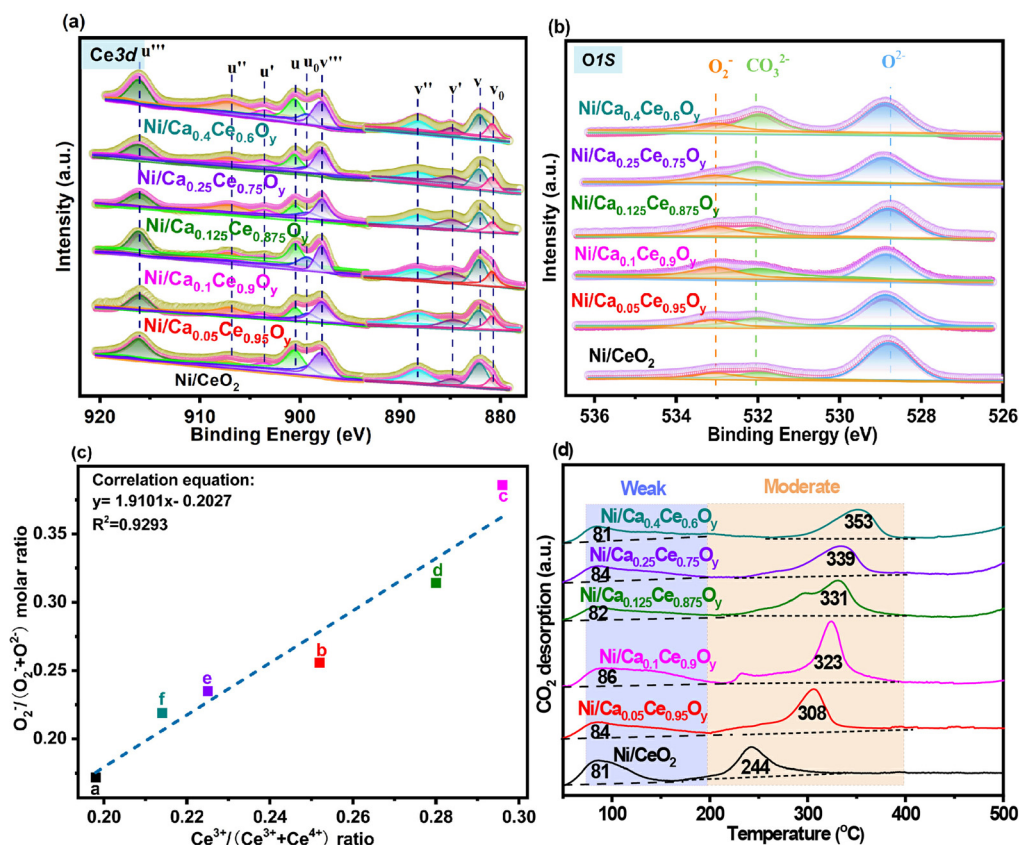
Fig. 5 Raman spectra of the Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> supports. (a) Octahedrally symmetrical vibration modes (F<sub>2g</sub>); (b) vacancy-induced modes; (c) superoxide (O<sub>2</sub><sup>-</sup>) modes and (d) EPR spectra of some typical Ni/Ca<sub>x</sub>Ce<sub>1-x</sub>O<sub>y</sub> catalysts. The results reveal a lattice capacity threshold effect, with the maximum solid-solution pure-phase sample being at the CeO<sub>2</sub> lattice capacity threshold (see Fig. 5, 6, 8 and Table S4<sup>†</sup>).

lattice matrix, which in turn affects the crystallinity (Fig. 5a). Interestingly, it is noteworthy that  $\text{Ca}_{0.1}\text{Ce}_{0.9}\text{O}_y$  exhibited the highest concentration of  $\text{O}_2^-$  species among all the  $\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$  supports. Raman spectroscopy results confirmed that introducing different levels of  $\text{Ca}^{2+}$  cations elevated both the density of oxygen vacancies and the quantity of  $\text{O}_2^-$  species, exhibiting a lattice capacity threshold effect. The modification of the solid solution supports has a notable impact on the surface characteristics of  $\text{Ni}/\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$ , which will be investigated further in conjunction with XPS and  $\text{CO}_2$ -TPD analysis.

EPR spectroscopy was employed to further characterize the  $\text{O}_2^-$  species with electromagnetism on the catalyst surface for some typical reduced catalysts (Fig. 5d). Notably, distinct EPR signals corresponding to  $\text{O}_2^-$  emerged, specifically with a  $g$  value of 2.002. The intensity of the  $\text{O}_2^-$  signal peaks followed the sequence  $\text{Ni}/\text{Ca}_{0.1}\text{Ce}_{0.9}\text{O}_y > \text{Ni}/\text{Ca}_{0.05}\text{Ce}_{0.95}\text{O}_y > \text{Ni}/\text{Ca}_{0.4}\text{Ce}_{0.6}\text{O}_y > \text{Ni}/\text{CeO}_2$ , where  $\text{Ni}/\text{Ca}_{0.1}\text{Ce}_{0.9}\text{O}_y$  showcased the most pronounced accumulation of surface oxygen vacancies, aligning with the Raman spectroscopy results. In summary, both EPR and Raman results consistently demonstrated a lattice capacity threshold effect for oxygen vacancies in both the support and supported Ni-based catalysts.

### 3.5 Exploring the lattice capacity threshold effect of oxygen vacancies and moderate alkali center contents *via* XPS and $\text{CO}_2$ -TPD

The surface elements of the reduced catalysts were assessed through XPS. The Ce 3d spectra were deconvoluted and analyzed (Fig. 6a and Table S5<sup>†</sup>). Each spectrum was deconvoluted into 10 groups corresponding to the surface species of  $\text{Ce}^{3+}$  and  $\text{Ce}^{4+}$ , with the full width at half maximum (FWHM) of a single Ce peak fixed in the range of 2.9–3.3 eV.<sup>49</sup> The six characteristic peaks at  $u'''$  (916.1 eV),  $u''$  (907.2 eV),  $u'$  (900.7 eV),  $v'''$  (898.0 eV),  $v''$  (888.6 eV) and  $v$  (882.2 eV) were assigned to the  $\text{Ce}^{4+}$  species. In contrast, the remaining four peaks at  $u'$  (903.9 eV),  $u_0$  (899.3 eV),  $v'$  (884.4 eV), and  $v_0$  (880.6 eV) were attributed to  $\text{Ce}^{3+}$  species. To conduct an investigation into the promotional effect of calcium on the surface  $\text{Ce}^{3+}$  concentration, the surface  $\text{Ce}^{3+}/(\text{Ce}^{3+} + \text{Ce}^{4+})$  ratios were determined by dividing the ratio of the integrated areas of the  $\text{Ce}^{3+}$  peaks with the sum of the integrated areas of both the  $\text{Ce}^{3+}$  and  $\text{Ce}^{4+}$  peaks.<sup>50</sup> Previous works have indicated a strong correlation between the quantity of oxygen vacancies on the surface of  $\text{CeO}_2$  and the concentration of  $\text{Ce}^{3+}$  ions on the surface.<sup>51,52</sup> The observed surface ratios of  $\text{Ce}^{3+}/(\text{Ce}^{3+} + \text{Ce}^{4+})$  followed the



**Fig. 6** XPS spectra of  $\text{Ni}/\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$  catalysts (a) Ce 3d; (b) O 1s; (c) the correlation of surface  $\text{O}_2^-/(\text{O}_2^- + \text{O}^{2-})$  ratios; (d)  $\text{CO}_2$  desorption profiles of the  $\text{Ni}/\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$  catalysts. (a)  $\text{Ni}/\text{CeO}_2$ ; b.  $\text{Ni}/\text{Ca}_{0.05}\text{Ce}_{0.95}\text{O}_y$ ; c.  $\text{Ni}/\text{Ca}_{0.1}\text{Ce}_{0.9}\text{O}_y$ ; d.  $\text{Ni}/\text{Ca}_{0.125}\text{Ce}_{0.875}\text{O}_y$ ; e.  $\text{Ni}/\text{Ca}_{0.25}\text{Ce}_{0.75}\text{O}_y$ ; f.  $\text{Ni}/\text{Ca}_{0.4}\text{Ce}_{0.6}\text{O}_y$ ). The lattice capacity threshold effect in the presence of oxygen vacancies and moderate-strength alkali centers was primarily attributed to the lattice capacity threshold effect (see Fig. 5, 6 and Table S5<sup>†</sup>).

particular sequence of  $\text{Ni}/\text{Ca}_{0.1}\text{Ce}_{0.9}\text{O}_y > \text{Ni}/\text{Ca}_{0.125}\text{Ce}_{0.875}\text{O}_y > \text{Ni}/\text{Ca}_{0.05}\text{Ce}_{0.95}\text{O}_y > \text{Ni}/\text{Ca}_{0.25}\text{Ce}_{0.75}\text{O}_y > \text{Ni}/\text{Ca}_{0.4}\text{Ce}_{0.6}\text{O}_y > \text{Ni}/\text{CeO}_2$ . These results aligned well with the density of oxygen vacancies on the surface of  $\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$ , as determined through Raman spectroscopy (Tables S4 and S5<sup>†</sup>). This further validated the presence of an oxygen vacancy lattice capacity threshold effect in the sample with the best catalytic performance. The surface  $\text{Ca}/(\text{Ce} + \text{Ca})$  ratios of  $\text{Ni}/\text{Ca}_{0.05}\text{Ce}_{0.95}\text{O}_y$  and  $\text{Ni}/\text{Ca}_{0.1}\text{Ce}_{0.9}\text{O}_y$  catalysts were found to be slightly elevated compared to their stoichiometric ratios in the bulk material. This observation suggested that a substantial portion of the  $\text{Ca}^{2+}$  cations was effectively incorporated into the  $\text{CeO}_2$  lattice, resulting in the formation of a solid solution structure. Conversely, the surface  $\text{Ca}/(\text{Ce} + \text{Ca})$  ratios in the  $\text{Ni}/\text{Ca}_{0.125}\text{Ce}_{0.875}\text{O}_y$ ,  $\text{Ni}/\text{Ca}_{0.25}\text{Ce}_{0.75}\text{O}_y$ , and  $\text{Ni}/\text{Ca}_{0.4}\text{Ce}_{0.6}\text{O}_y$  catalysts were significantly elevated compared to their stoichiometric ratios in the  $\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$  bulk materials, indicating that  $\text{Ca}^{2+}$  cations were mainly found as scattered carbonate deposits on the support surfaces. Higher surface oxygen vacancies were expected to lead to more active oxygen species on the catalysts' surface due to the increased adsorption of gas-phase  $\text{O}_2$  molecules. We deconvoluted and analyzed the O 1s spectra for the reduced catalysts (Fig. 6b). Three distinct peaks representing  $\text{O}_2^-$  (533.0–533.1 eV),  $\text{CO}_3^{2-}$  (531.9–532.2 eV), and the lattice oxygen  $\text{O}^{2-}$  (528.6–529.2 eV) were identified based on a single oxygen peak with an FWHM of 1.8–2.0 eV.<sup>53</sup>  $\text{O}_2^-/(\text{O}_2^- + \text{O}^{2-})$ ; the ratio between the surface  $\text{O}_2^-$  and lattice oxygen

amount was quantified (Table S5<sup>†</sup>). The intimate correlation of surface  $\text{O}_2^-/(\text{O}_2^- + \text{O}^{2-})$  ratios with the  $\text{Ce}^{3+}/(\text{Ce}^{3+} + \text{Ce}^{4+})$  ratios (Fig. 6c) showcased that surface-bound  $\text{O}_2^-$  originates from  $\text{O}_2$  adsorption at the oxygen vacancies.

$\text{CO}_2$  adsorption/desorption technique was employed to assess the surface alkalinity of the reduced catalysts (Fig. 6d and Table S5<sup>†</sup>). Several distinct  $\text{CO}_2$  desorption peaks were observed at temperatures below 450 °C. Peaks occurring at temperatures below 200 °C and between 200 and 450 °C were identified as weak and moderate alkaline sites, respectively. Here, the weak alkaline sites were reported as surface OH groups, followed by oxygen vacancies as moderate alkaline sites.<sup>51</sup> The combined quantity of alkaline sites and the quantity of moderate alkaline sites followed the sequence  $\text{Ni}/\text{Ca}_{0.1}\text{Ce}_{0.9}\text{O}_y > \text{Ni}/\text{Ca}_{0.125}\text{Ce}_{0.875}\text{O}_y > \text{Ni}/\text{Ca}_{0.05}\text{Ce}_{0.95}\text{O}_y > \text{Ni}/\text{Ca}_{0.25}\text{Ce}_{0.75}\text{O}_y > \text{Ni}/\text{Ca}_{0.4}\text{Ce}_{0.6}\text{O}_y > \text{Ni}/\text{CeO}_2$  (Table S5<sup>†</sup>). This trend aligned with the intrinsic activity ( $R_s$ ) sequence (Tables 2 and S5<sup>†</sup>), which reaffirmed the lattice capacity threshold effect on the concentration of medium-strength alkali centers generated by oxygen vacancies in solid solutions.  $\text{Ni}/\text{Ca}_{0.1}\text{Ce}_{0.9}\text{O}_y$  and  $\text{Ni}/\text{Ca}_{0.125}\text{Ce}_{0.875}\text{O}_y$ , the two highly active catalysts, not only exhibited a significantly higher quantity of moderate alkaline sites but also demonstrated elevated  $\text{CO}_2$  desorption temperatures compared to the other samples, indicating that both weak and moderate alkaline sites facilitated the  $\text{CO}_2$  methanation reaction, with the moderate alkaline sites playing a primary role in determining the activity.

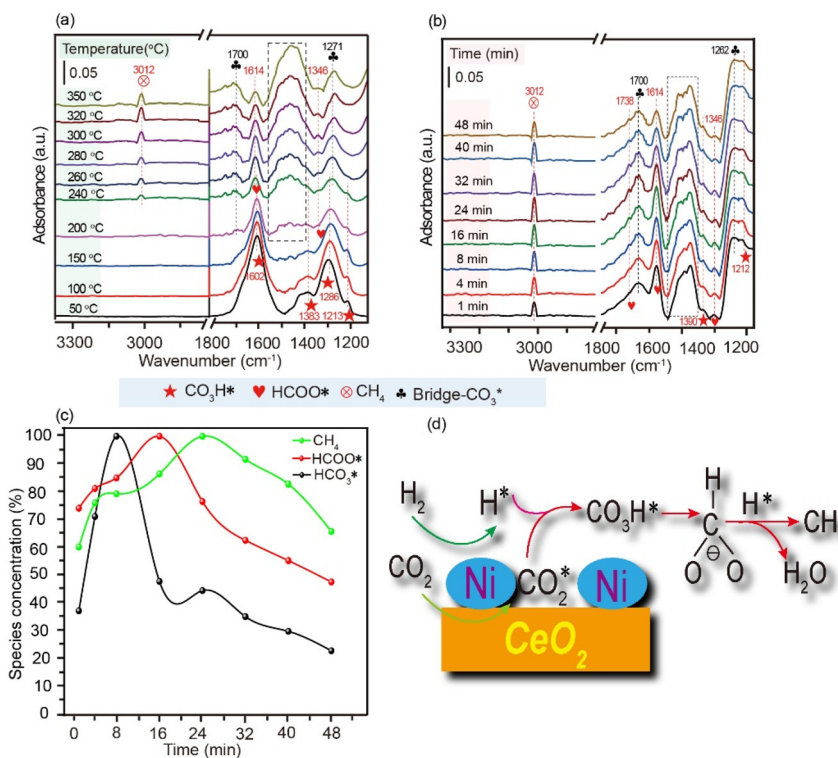


Fig. 7 *In situ* FTIR spectra on the  $\text{Ni}/\text{CeO}_2$  catalyst, (a) tendency over temperature; (b) tendency over time; (c) dynamic evolution of reaction intermediates chart; (d) proposed possible reaction mechanism ( $\text{CO}_3\text{H}^* \rightarrow \text{HCOO}^* \rightarrow \text{CH}_4$ ).

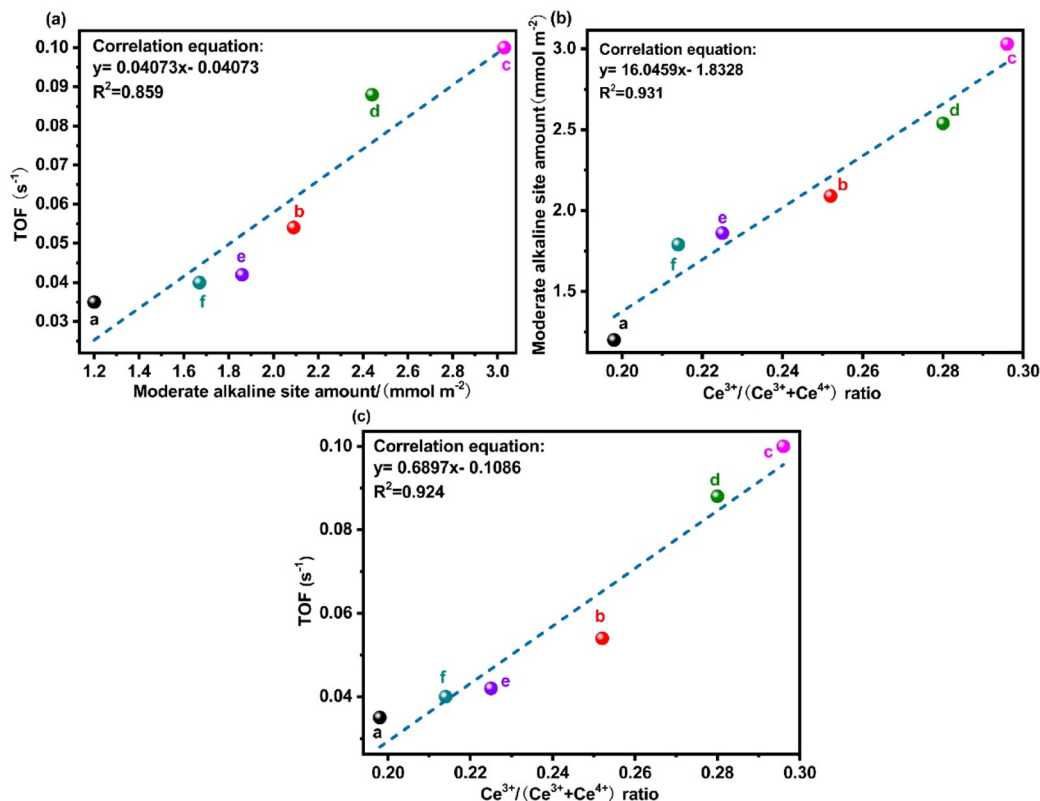
### 3.6 *In situ* FTIR investigation of the reaction path of Ni/CeO<sub>2</sub> for CO<sub>2</sub> methanation

*In situ* FTIR experiments were conducted to detect the reaction intermediates during CO<sub>2</sub> hydrogenation on the Ni/CeO<sub>2</sub> catalyst. The signals (1213, 1286, 1383 cm<sup>-1</sup>, and 1602 cm<sup>-1</sup>) were attributed to bicarbonate (CO<sub>3</sub>H\*) species<sup>54–56</sup> at 50 °C (Fig. 7a). As the temperature increased, the intensity of these CO<sub>3</sub>H\* species gradually decreased and nearly disappeared at 260 °C. The peaks at 1271 and 1700 cm<sup>-1</sup> were associated with bridged carbonate,<sup>57</sup> while peaks in the range of 1550–1350 cm<sup>-1</sup> correspond to carbonate species.<sup>58</sup> As the CO<sub>3</sub>H\* species were consumed, formates (HCOO\*) began to emerge (1346 and 1614 cm<sup>-1</sup>).<sup>54,56,59</sup> The intensity of HCOO\* significantly increased between 200 °C and 280 °C and then declined. Notably, a new peak at 3012 cm<sup>-1</sup> emerged, which was attributed to the CH<sub>4</sub>\* species under 240 °C. These results suggested that CH<sub>4</sub> formation may originate from HCOO\* species during CO<sub>2</sub> methanation on the Ni/CeO<sub>2</sub> catalyst.

To better illustrate the mutual conversion among intermediate species and their relationships (CO<sub>3</sub>H\*, HCOO\*, CH<sub>4</sub>\*), *in situ* infrared testing was conducted at 240 °C (Fig. 7b). Prior to conducting the experiment on the temporal

dynamics of the intermediate species, the catalyst underwent a pre-treatment at 300 °C in a high-purity Ar atmosphere for 1 h to eliminate surface impurities. Following this, the temperature was reduced to room temperature. After background subtraction, a mixed gas containing 1% CO<sub>2</sub>, 4% H<sub>2</sub>, and 95% Ar was introduced. Subsequently, rapid heating from ambient temperature to 240 °C was carried out to investigate the dynamic variations among the intermediate species at this temperature. The formation of formate (HCOO\*) initially increased, reaching its maximum at 16 min, while methane (CH<sub>4</sub>\*) also exhibited an initial increase tendency, reaching its maximum at 24 min (Fig. 7b).

To illustrate the dynamic changes in the intermediate species and the methane formation mechanism, we plotted the integrated peak areas of all the carbonate, formate, and methane species, setting the peak area with the highest value to 100%. This allowed us to observe the percentage content trends of these three species (Fig. 7c). It was evident that CO<sub>3</sub>H\* initially increased, reached a maximum at 8 min, and then gradually decreased, stabilizing after 40 min. Concurrently, as CO<sub>3</sub>H\* was rapidly consumed, HCOO\* exhibited a similar trend, increasing until 16 min and then stabilizing after 40 min. As HCOO\* was consumed, the CH<sub>4</sub> content



**Fig. 8** (a) Correlation between the TOF values at 250 °C and the amount of surface moderate alkaline sites, (b) the correlation between the surface moderate alkaline sites and Ce<sup>3+</sup>/(Ce<sup>3+</sup> + Ce<sup>4+</sup>) ratio, and (c) the correlation between the TOF values at 250 °C and Ce<sup>3+</sup>/(Ce<sup>3+</sup> + Ce<sup>4+</sup>) ratios. (a) Ni/CeO<sub>2</sub>; b. Ni/Ca<sub>0.05</sub>Ce<sub>0.95</sub>O<sub>y</sub>; c. Ni/Ca<sub>0.1</sub>Ce<sub>0.9</sub>O<sub>y</sub>; d. Ni/Ca<sub>0.125</sub>Ce<sub>0.875</sub>O<sub>y</sub>; e. Ni/Ca<sub>0.25</sub>Ce<sub>0.75</sub>O<sub>y</sub>; f. Ni/Ca<sub>0.4</sub>Ce<sub>0.6</sub>O<sub>y</sub>). The lattice capacity threshold effect in the presence of oxygen vacancies and moderate-strength alkali centers was primarily attributed to the lattice capacity threshold effect.

followed a pattern of initial increase, reaching its maximum at 24 min. Throughout this process,  $\text{CO}_3\text{H}^*$ ,  $\text{HCOO}^*$ , and  $\text{CH}_4$  co-existed, indicating a dynamic equilibrium in their conversion.

Based on the aforementioned findings, a  $\text{CO}_2$  methanation mechanism was proposed.  $\text{CO}_2$  interacted with the surface OH sites, forming  $\text{CO}_3\text{H}^*$  at low temperatures. As the temperature increased,  $\text{CO}_3\text{H}^*$  was transformed into  $\text{HCOO}^*$ , which subsequently evolved into  $\text{CH}_4^*$  (Fig. 7d).

### 3.7 The network of structure–activity relationship

It should be noted that surface oxygen vacancies play a pivotal role as highly active alkaline sites for  $\text{CO}_2$  adsorption and conversion.<sup>11,60</sup> These sites, stemming from surface oxygen vacancies, serve as affordably reactive moderate alkaline sites on  $\text{CeO}_2$ -based catalysts for various  $\text{CO}_2$  conversion reactions.<sup>11,51,61,62</sup> Hence, to better understand the relationship between oxygen vacancies, intermediate alkali centers, and turnover frequency (TOF), the relevant correlation lines were plotted. The linear correlation between surface oxygen vacancies and moderate alkaline sites further proved that the moderate alkaline sites were primarily assigned to the surface oxygen vacancies of  $\text{Ni}/\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$  catalysts (Fig. 8a). A clear and direct linear relationship was evident between the TOF values and the abundance of moderate alkaline sites (Fig. 8b). Catalysts with a higher concentration of moderate alkaline sites exhibited superior intrinsic activity for  $\text{CO}_2$  methanation, underscoring the pivotal role of these sites in the catalytic performance. Notably, surface oxygen vacancies, serving as alkaline sites, are expected to enhance  $\text{CO}_2$  activation. Evidently, there was a linear correlation between intrinsic catalytic performance and the amount of surface oxygen vacancies (Fig. 8c). Catalysts with a higher concentration of surface oxygen vacancies exhibited favorable intrinsic activity for  $\text{CO}_2$  methanation.

In summary, oxygen vacancies in solid solutions can serve as highly active catalytic sites at the reaction interface during the  $\text{CO}_2$  methanation reaction, which enhances  $\text{CO}_2$  adsorption activation and significantly boosts low-temperature reactivity. The presence of moderately strong alkaline centers in solid solution materials originated from oxygen defect sites, and the content of these alkaline centers determined the reaction activity. Here,  $\text{Ni}/\text{Ca}_{0.1}\text{Ce}_{0.9}\text{O}_y$ , due to the presence of the maximum amount of the pure solid solution phase, generates the highest number of oxygen vacancies. The lattice capacity effect resulting from the oxygen vacancies in this sample was the primary reason for the threshold effect observed in the methane conversion reactivity.

## 4 Conclusion

A series of  $\text{Ni}/\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$  catalysts were synthesized through the sol–gel method and employed for  $\text{CO}_2$  methanation to investigate the impact of  $\text{Ca}^{2+}$  incorporation into the  $\text{CeO}_2$  lattice on the structure–activity relationship of supported Ni catalysts, specifically focusing on the lattice capacity threshold

effect. XRD analysis was initially employed to study the solid solution structure and cell parameters, leading to the first-ever determination of the lattice capacity for Ca in  $\text{CeO}_2$  through the XRD extrapolation method. The lattice capacity of  $\text{Ca}^{2+}$  in  $\text{CeO}_2$  was detected for the first time, corresponding to a Ca : Ce molar ratio of 11 : 89. The catalytic materials in the  $\text{Ni}/\text{Ca}_x\text{Ce}_{1-x}\text{O}_y$  series exhibited a lattice capacity threshold effect due to the presence of alkaline centers generated by oxygen vacancies. This lattice capacity threshold effect was the primary reason for the enhanced activity observed in these catalysts during the  $\text{CO}_2$  methanation reaction. While the interaction between the metal and the support as well as the metal Ni exhibiting lattice capacity threshold effects promote the enhancement of the activity, they were not the primary factors contributing to the lattice capacity threshold effect in the active sites. Furthermore, the potential reaction pathway of  $\text{Ni}/\text{CeO}_2$  catalyst in the  $\text{CO}_2$  methanation was revealed ( $\text{CO}_3\text{H}^* \rightarrow \text{HCOO}^* \rightarrow \text{CH}_4$ ).

This study established a distinct correlation between oxygen vacancies and intrinsic activity through XRD extrapolation. The initial exploration unveiled the presence of a lattice capacity threshold effect in composite catalysts composed of metal oxide-supported Ni, which offers crucial insights for optimizing low-temperature  $\text{CO}_2$  methanation performance.

## Conflicts of interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

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## References

- 1 F. Tian, X. Wu, J. Chen, X. Sun, X. Yan and G. Liao, *Dalton Trans.*, 2023, **52**, 11934–11940.
- 2 G. Ding, C. Li, Y. Ni, L. Chen, L. Shuai and G. Liao, *EES Catal.*, 2023, **1**, 369–391.
- 3 Q. Zhang, G. Liao, B. Yang, Y. Zhang, G. Ge, A. Lipovka, J. Liu, R. D. Rodriguez, X. Yang and X. Jia, *Appl. Surf. Sci.*, 2023, **638**, 157989.
- 4 G. Liao, C. Li, S.-Y. Liu, B. Fang and H. Yang, *Phys. Rep.*, 2022, **983**, 1–41.
- 5 G. Liao, Y. He, H. Wang, B. Fang, N. Tsubaki and C. Li, *Device*, 2023, **1**, 100173.

- 6 G. Liao, G. Ding, B. Yang and C. Li, *Precis. Chem.*, 2024, **2**, 49–56.
- 7 K. Wei, H. Guan, Q. Luo, J. He and S. Sun, *Nanoscale*, 2022, **14**, 11869–11891.
- 8 F. M. Mota and D. H. Kim, *Chem. Soc. Rev.*, 2019, **48**, 205–259.
- 9 S. Rönsch, J. Schneider, S. Matthischke, M. Schlüter, M. Götz, J. Lefebvre, P. Prabhakaran and S. Bajohr, *Fuel*, 2016, **166**, 276–296.
- 10 M. Younas, L. L. Kong, M. J. K. Bashir, H. Nadeem, A. Shehzad and S. Sethupathi, *Energy Fuels*, 2016, **30**, 8815–8831.
- 11 K. Liu, X. Xu, J. Xu, X. Fang, L. Liu and X. Wang, *J. CO<sub>2</sub> Util.*, 2020, **38**, 113–124.
- 12 B. Li, B. Ma, S.-Y. Wang, M.-M. Yu, Z.-Q. Zhang, M.-J. Xiao, H. Zhang, J.-F. Wu, Y. Peng, Q. Wang and H.-L. Zhang, *Nanoscale*, 2022, **14**, 9736–9742.
- 13 L. P. L. Gonçalves, M. Meledina, A. Meledin, D. Y. Petrovykh, J. P. S. Sousa, O. S. G. P. Soares, Y. V. Kolen'ko and M. F. R. Pereira, *Carbon*, 2022, **195**, 35–43.
- 14 L. Zhong, H. Zheng, C. Du, W. Du, G. Liao, C. S. Cheung and H. Gao, *J. Catal.*, 2020, **384**, 208–217.
- 15 C. Italiano, J. Llorca, L. Pino, M. Ferraro, V. Antonucci and A. Vita, *Appl. Catal., B*, 2020, **264**, 118494.
- 16 X. Jia, X. Zhang, N. Rui, X. Hu and C.-J. Liu, *Appl. Catal., B*, 2019, **244**, 159–169.
- 17 J. K. Prabhakar, P. A. Apte and G. Deo, *Chem. Eng. J.*, 2023, **471**, 144252.
- 18 R.-P. Ye, Q. Li, W. Gong, T. Wang, J. J. Razink, L. Lin, Y.-Y. Qin, Z. Zhou, H. Adidharma, J. Tang, A. G. Russell, M. Fan and Y.-G. Yao, *Appl. Catal., B*, 2020, **268**, 118474.
- 19 C. Li, B. Cheng, H. Lu, G. Ding, Z. Jiang and G. Liao, *Inorg. Chem.*, 2023, **62**, 6843–6850.
- 20 G. Liao, Z. Xiao, X. Chen, C. Du, L. Zhong, C. S. Cheung and H. Gao, *Macromolecules*, 2020, **53**, 256–266.
- 21 S. Tada, T. Shimizu, H. Kameyama, T. Haneda and R. Kikuchi, *Int. J. Hydrogen Energy*, 2012, **37**, 5527–5531.
- 22 G. Liao, L. Zhang, C. Li, S.-Y. Liu, B. Fang and H. Yang, *Matter*, 2022, **5**, 3341–3374.
- 23 B. Nematollah, M. Rezaei and E. N. Lay, *J. Rare Earths*, 2015, **33**, 619–628.
- 24 R. Darouhegi, F. Mowshkani and M. Rezaei, *Chem. Eng. Sci.*, 2020, **2302**, 41–50.
- 25 Z. Zhang, Y. Tong, X. Fang, J. Xu, X. Xu and X. Wang, *Fuel*, 2022, **316**, 123191.
- 26 N. Rui, X. Zhang, F. Zhang, Z. Liu, X. Cao, Z. Xie, R. Zou, S. D. Senanayake, Y. Yang, J. A. Rodriguez and C.-J. Liu, *Appl. Catal., B*, 2021, **282**, 119581.
- 27 S. Lin, Z. Hao, J. Shen, X. Chang, S. Huang, M. Li and X. Ma, *J. Energy Chem.*, 2021, **59**, 334–342.
- 28 L. Lin, C. A. Gerlak, C. Liu, J. Llorca, S. Yao, N. Rui, F. Zhang, Z. Liu, S. Zhang, K. Deng, C. B. Murray, J. A. Rodriguez and S. D. Senanayake, *J. Energy Chem.*, 2021, **61**, 602–611.
- 29 Y. Xie, J. Chen, X. Wu, J. Wen, R. Zhao, Z. Li, G. Tian, Q. Zhang, P. Ning and J. Hao, *ACS Catal.*, 2022, **12**, 10587–10602.
- 30 G. Liao, J. Fang, Q. Li, S. Li, Z. Xu and B. Fang, *Nanoscale*, 2019, **11**, 7062–7096.
- 31 X. Xu, L. Liu, Y. Tong, X. Fang, J. Xu, D.-E. Jiang and X. Wang, *ACS Catal.*, 2021, **11**, 5762–5775.
- 32 W.-T. Chen, K.-B. Chen, M.-F. Wang, S.-F. Weng, C.-S. Lee and M. C. Lin, *Chem. Commun.*, 2010, **46**, 3286–3288.
- 33 J. Sun, Y. Lu, L. Zhang, C. Ge, C. Tang, H. Wan and L. Dong, *Ind. Eng. Chem. Res.*, 2017, **56**, 12101–12110.
- 34 J. Y. Wei, Y. X. Zhu, L. Y. Duan and Y. C. Xie, *Chin. J. Catal.*, 2003, **24**, 414–418.
- 35 X. Yao, C. Tang, Z. Ji, Y. Dai, Y. Cao, F. Gao, L. Dong and Y. Chen, *Catal. Sci. Technol.*, 2013, **3**, 688–698.
- 36 P. Zhang, H. Lu, Y. Zhou, L. Zhang, Z. Wu, S. Yang, H. Shi, Q. Zhu, Y. Chen and S. Dai, *Nat. Commun.*, 2015, **6**, 8446.
- 37 X. Yao, Y. Xiong, W. Zou, L. Zhang, S. Wu, X. Dong, F. Gao, Y. Deng, C. Tang, Z. Chen, L. Dong and Y. Chen, *Appl. Catal., B*, 2014, **144**, 152–165.
- 38 P. F. Zhang, H. F. Lu, Y. Zhou, L. Zhang, Z. L. Wu, S. Z. Yang, H. L. Shi, Q. L. Zhu, Y. F. Chen and S. Dai, *Nat. Commun.*, 2015, **6**, 10.
- 39 Z. Zhang, Y. Tong, X. Fang, J. Xu, X. Xu and X. Wang, *Fuel*, 2022, **316**, 123191.
- 40 M. Sugiura, *Catal. Surv. Asia*, 2003, **7**, 77–87.
- 41 Q. Sun, X. Xu, H. Peng, X. Fang, W. Liu, J. Ying, F. Yu and X. Wang, *Chin. J. Catal.*, 2016, **37**, 1293–1302.
- 42 X. Xu, F. Liu, X. Han, Y. Wu, W. Liu, R. Zhang, N. Zhang and X. Wang, *Catal. Sci. Technol.*, 2016, **6**, 5280–5291.
- 43 Y. Xie and Y. Tang, *Adv. Catal.*, 1990, **37**, 1–43.
- 44 X. Xu, F. Liu, J. Tian, H. Peng, W. Liu, X. Fang, N. Zhang and X. Wang, *ChemPhysChem*, 2017, **18**, 1533–1540.
- 45 X. Lin, S. Li, H. He, Z. Wu, J. Wu, L. Chen, D. Ye and M. Fu, *Appl. Catal., B*, 2018, **223**, 91–102.
- 46 J. Xu, L. Peng, X. Fang, Z. Fu, W. Liu, X. Xu, H. Peng, R. Zheng and X. Wang, *Appl. Catal., A*, 2018, **552**, 117–128.
- 47 J. Xu, Y. Zhang, X. Xu, X. Fang, R. Xi, Y. Liu, R. Zheng and X. Wang, *ACS Catal.*, 2019, **9**, 4030–4045.
- 48 Z.-Y. Pu, X.-S. Liu, A.-P. Jia, Y.-L. Xie, J.-Q. Lu and M.-F. Luo, *J. Phys. Chem. C*, 2008, **112**, 15045–15051.
- 49 F. Larachi, J. Pierre, A. Adnot and A. Bernis, *Appl. Surf. Sci.*, 2002, **195**, 236–250.
- 50 Z. Hu, X. Liu, D. Meng, Y. Guo, Y. Guo and G. Lu, *ACS Catal.*, 2016, **6**, 2265–2279.
- 51 F. Wang, M. Wei, D. G. Evans and X. Duan, *J. Mater. Chem. A*, 2016, **4**, 5773–5783.
- 52 Z. Bian, Y. M. Chan, Y. Yu and S. Kawi, *Catal. Today*, 2020, **347**, 31–38.
- 53 V. J. Ferreira, P. Tavares, J. L. Figueiredo and J. L. Faria, *Ind. Eng. Chem. Res.*, 2012, **51**, 10535–10541.
- 54 C. Binet, M. Daturi and J. C. Lavalley, *Catal. Today*, 1999, **50**, 207–225.
- 55 Q. Pan, J. Peng, T. Sun, S. Wang and S. Wang, *Catal. Commun.*, 2014, **45**, 74–78.

- 56 Q. Pan, J. Peng, S. Wang and S. Wang, *Catal. Sci. Technol.*, 2014, **4**, 502–509.
- 57 C. Li, Y. Sakata, T. Arai, K. Domen, K.-I. Maruya and T. Onishi, *J. Chem. Soc., Faraday Trans. 1*, 1989, **85**, 929–943.
- 58 G. Marban, I. Lopez and T. Valdes-Solis, *Appl. Catal., A*, 2009, **361**, 160–169.
- 59 S. Haq, J. G. Love, H. E. Sanders and D. A. King, *Surf. Sci.*, 1995, **325**, 230–242.
- 60 B. Yang, X. Li, Q. Zhang, X. Yang, J. Wan, G. Liao, J. Zhao, R. Wang, J. Liu, R. D. Rodriguez and X. Jia, *Appl. Catal., B*, 2022, **314**, 121521.
- 61 J. Graciani, K. Mudiyansele, F. Xu, A. E. Baber, J. Evans, S. D. Senanayake, D. J. Stacchiola, P. Liu, J. Hrbek, J. F. Sanz and J. A. Rodriguez, *Science*, 2014, **345**, 546–550.
- 62 F. Wang, S. He, H. Chen, B. Wang, L. Zheng, M. Wei, D. G. Evans and X. Duan, *J. Am. Chem. Soc.*, 2016, **138**, 6298–6305.