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Strategic design of VO₂ encased in N-doped carbon as an efficient electrocatalyst for the nitrogen reduction reaction in neutral and acidic media†

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Electrocatalytic nitrogen fixation to ammonia (NH₃), a precursor for fertilizer production and a promising energy carrier, has garnered widespread interest as an environment-friendly and sustainable alternative to the energy-intensive fossil-feedstock-dependent Haber–Bosch process. The large-scale deployment of this process is contingent on the identification of inexpensive, Earth-abundant systems that can operate efficiently, irrespective of the electrolyte pH for the selective production of NH₃. In this regard, we discuss the scalable synthesis of VO₂ anchored on N-doped carbon (VO₂@CN), and its applicability as a robust electrocatalyst for the nitrogen reduction reaction (NRR). Benefitting from the presence of exposed VO₂, which presumably acts as the active site for nitrogen reduction, and its activity over a broad pH range (from acidic to neutral), VO₂@CN exhibits a high NH₃ yield of 0.31 and 0.52 μmol h⁻¹ mg_{cat}⁻¹ and a maximum Faradaic efficiency (FE) of 67.9% and 61.9% at -0.1 V vs. RHE, under neutral and acidic conditions, respectively. The obscured reaction intermediates of the NRR were identified from *in situ* ATR-IR studies under both electrolyte conditions. Additionally, the high selectivity of the catalyst was ascertained from the absence of hydrazine production and the competing hydrogen evolution reaction (HER). However, ammonia production underwent a reduction over 12 h of continuous operation presumably owing to the leaching of catalyst under these electrolysis conditions, which was more pronounced in electrolytes with acidic pH. Overall, the present report unveils the performance of an earth-abundant vanadium oxide-based system as an efficient electrocatalyst for the NRR under acidic and neutral pH conditions.

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Introduction

Ammonia (NH₃) obtained from the reduction of nitrogen is perceived as a clean energy carrier and a replacement for gaseous hydrogen owing to its compression to the liquid state at a low pressure of ~8 bar and room temperature and its high energy density in the liquid state.^{1,2} In addition, NH₃ is vital in agriculture, in the production of fertilizers to feed the ever-increasing global population. The industrial production of

NH₃ relies on the Haber–Bosch process, which requires harsh temperature (400–600 °C) and pressure (20–40 MPa) conditions, has an energy efficiency of ~65%, and accounts for about 2% of global energy consumption, releasing a large amount of CO₂ into the atmosphere causing an increase in the global temperature.^{3–5} This emphasizes the necessity of developing a clean, sustainable alternative process for nitrogen fixation, which precludes the necessity of elevated temperatures and reduces the carbon footprint.^{6,7} Various routes for nitrogen fixation, including photochemical, electrochemical and biological approaches have been explored by the research community to arrive at environmentally safe and energy-efficient methods to artificially fix nitrogen.^{8–11} Among these approaches, the electrochemical NRR is considered a promising alternative to NH₃ production due to its mild reaction conditions and environmentally friendly precursors of N₂ and H₂O, which can revolutionize green ammonia production when used in conjunction with renewable energy resources.^{5,12} Intensive efforts have been devoted worldwide to the design

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and investigation of catalysts for the selective production of NH_3 via the NRR with a high FE.^{13–19} It is extremely arduous to dissociate the inert dinitrogen molecule (bond energy of N_2 is 941 kJ mol^{-1}), and mechanistically, the NRR involves multiple proton-coupled electron-transfer steps.²⁰ In addition, the utility of the process is limited by the competing hydrogen evolution reaction (HER), which has faster kinetics and operates under a similar potential window.

The literature reveals a series of transition metal-based systems that have been reported to facilitate the NRR under ambient conditions.^{21–28} Of these, Mo, Fe, Rh and Ru are at the top of the volcano diagram, and hence, they show the best activity.²³ However, the competing HER lowers their efficiency for ammonia production. Early transition metals including Zr, Ti, Sc, V *etc.*, have been reported to bind N-atoms more strongly compared to H-atoms, thereby enhancing the production of NH_3 compared to H_2 .²³ Among these, vanadium is noteworthy, as it plays a vital role in the nitrogenase enzyme, responsible for replenishing bioavailable ammonia from atmospheric dinitrogen.²⁹ Vanadium carbide, oxynitrides and phosphides have been extensively reported as electrocatalysts for the NRR mostly under neutral pH conditions,^{7,24,30,31} but vanadium oxides have been relatively less explored in this regard, despite the surmise that the NRR actually operates at O-supported vanadium Lewis acidic sites.^{27,28} DFT calculations indicate that N_2 activation is energetically favoured on O-supported V centres by $\sim 18 \text{ kcal mol}^{-1}$ compared to that on N-supported V centres.³² A survey of the literature disclosed that the coordination environment and oxophilicity of the metal are important in modulating the performance of electrocatalysts, as these properties affect the binding of N atoms and N_2 molecules on the substrate and the stability of the intermediates culminating in the release of NH_3 .³³ Catalysts with electron-deficient centres were reported to be beneficial for NRR performance enhancement.²⁷ On the other hand, carbon-based nanomaterials have garnered widespread attention for electrocatalytic applications owing to their mechanical stability, superior electrical conductivity and enhanced surface areas. Doping of heteroatoms on the carbon surface creates defect sites and a redistribution of electrons to regulate the electron transfer and alters the sorption of reactants, intermediates and products on the catalyst surface. N-doped carbon substrates have been reported to facilitate N_2 adsorption and $\text{N}\equiv\text{N}$ cleavage, though they show very mild electrocatalytic NRR activity.^{13,34} Though there are a couple of reports on the utility of vanadium oxide and N-doped carbon individually for the electrocatalytic NRR, their combined efficacy is yet to be explored for the NRR, to the best of our understanding.

Motivated by the above discussions, we present vanadium dioxide encased in an N-doped carbon support ($\text{VO}_2@\text{CN}$) as an electrocatalyst for the NRR. A multidimensional structure due to a combination of vanadium dioxide and N-doped carbon is expected to improve the exposure of active sites and increase electrical conductivity, which might facilitate electrochemical nitrogen fixation. Considering the advantage of reactive Lewis acidic oxo-vanadium centres for efficient N_2 adsorp-

tion and the stability of vanadium oxides over broad potential and pH ranges,³⁵ we have explored $\text{VO}_2@\text{CN}$ as an electrocatalyst for the NRR under acidic and neutral electrolyte conditions. Facilitated by the effectual N_2 adsorption, and stabilization of vanadium in multiple oxidation states to modulate multiple electron-transfer steps, $\text{VO}_2@\text{CN}$ exhibits a high NH_3 yield of 0.31 and $0.52 \mu\text{mol h}^{-1} \text{mg}_{\text{cat}}^{-1}$ and a maximum FE of 67.9% and 61.9% at -0.1 V vs. RHE , under neutral and acidic conditions, respectively. The complexities of the NRR could be minimized by the detection of the intermediates formed during the course of the reaction under both electrolyte conditions using *in situ* ATR-IR studies. This in turn paves the way for a better understanding of the reaction kinetics on the $\text{VO}_2@\text{CN}$ surface.

Results and discussion

Synthesis and structural characterization

A simple solution-processing route followed by carbonization (Fig. 1a) was employed to synthesize $\text{VO}_2@\text{CN}$. First, vanadium pentoxide was dissolved in 3 M HNO_3 to form $\text{VO}_2(\text{NO}_3)$ species. The acidic solution of oxo-vanadium species was thoroughly mixed with a solution of melamine in methanol. The resultant precipitate was evaporated to dryness, where the oxo-vanadium ions were trapped by the melamine moieties, resulting in the formation of $\text{VO}_2(\text{NO}_3)@\text{melamine}$ (referred to as $\text{VO}_2@\text{mel}$ hereafter). The reaction proceeded by carbonization under an inert atmosphere at $600 \text{ }^\circ\text{C}$ to yield the $\text{VO}_2@\text{CN}$ material. NH_3 liberated during the thermolysis of melamine at elevated temperatures acted as a reducing agent to convert V^{+5} in $\text{VO}_2@\text{mel}$ to V^{+4} in $\text{VO}_2@\text{CN}$.³⁶ The coordination of dinitrogen to transition metal centres is usually facilitated by a lower oxidation state of the metal.³⁷ In the case



Fig. 1 (a) Schematic representation of the synthesis of $\text{VO}_2@\text{CN}$; (b) orbital interaction between N_2 and the VO_2 centre of the catalyst, responsible for N_2 reduction to NH_3 (NRR).

of vanadium, V^{+5} is the most stable oxidation state, while the lower oxidation states are often sensitive to air and moisture. The N-doped carbon matrix in $VO_2@CN$ not only aids in imparting stability to VO_2 , but it is expected to facilitate the NRR. It is envisaged that the VO_2 component of the catalyst would bring about σ -donation and counter π -back donation interactions between the highest occupied molecular orbitals (HOMOs) and lowest unoccupied molecular orbitals (LUMOs) of the catalyst active site and N_2 , as schematically illustrated in Fig. 1b. This interaction would help weaken the $N\equiv N$ bond strength, activating the molecule towards a successful reduction reaction. It is to be noted that V^{+5} in the precursor $VO_2@mel$ cannot favour π -back donation to weaken the $N\equiv N$ bond. The PXRD analysis of the as-synthesized material shows distinct peaks corresponding to vanadium dioxide, which can be assigned to the monoclinic phase of VO_2 (JCPDS file no. 01-082-0661, space group $P2_1/c$) (Fig. 2).³⁸ The presence of melamine in acidic solution prevents the fast aggregation of VO_6 octahedra to favour the exclusive formation of the stable monoclinic VO_2 (M) phase. The absence of any other phases in the as-prepared sample ascertains the purity of the VO_2 (M) phase. The presence of N-doped carbon was supported by the appearance of a broad signal at $\sim 2\theta = 23^\circ$ corresponding to the 002 plane. The FTIR spectrum of $VO_2@CN$ (Fig. S2†) shows the presence of a sharp band at $\sim 435\text{ cm}^{-1}$ characteristic of the long-range stretching vibrations of the V–O–V bridge and at $\sim 684\text{ cm}^{-1}$ indicative of the packing of VO_6 octahedra, respectively.³⁹ The breathing mode of triazine can be observed from a peak at $\sim 813\text{ cm}^{-1}$. In addition, the incorporation of nitrogen into the carbon framework can be identified from an FTIR signal at $\sim 1372\text{ cm}^{-1}$.⁴⁰ Broad bands at $\sim 3200\text{ cm}^{-1}$ are attributable to N–H stretching and hydrogen bonding in the $VO_2@CN$ sample. Raman spectroscopy was performed under the excitation wavelength of 532 nm, which revealed three peaks at 1373 cm^{-1} , 1598 cm^{-1} and 2760 cm^{-1} corresponding to the D, G and 2D bands of N-doped carbon, respectively (Fig. S3†). The calculated I_D/I_G value of 1.13 indicated the introduction of nitrogen atoms into the carbon framework



Fig. 2 PXRD pattern of the as-synthesized $VO_2@CN$ indicating the presence of VO_2 (M) and N-doped carbon.

resulting in significant lattice distortion.⁴¹ TGA measurements reveal multiple weight loss phases until $\sim 450^\circ\text{C}$ indicative of the release of adhered water molecules and interlayer water molecules, respectively, followed by the decomposition of N-doped carbon layers beyond 600°C (Fig. S4†).⁴²

The microenvironment of $VO_2@CN$ and its precursor $VO_2@mel$ was investigated from X-ray photoelectron spectroscopy (XPS) measurements (Fig. 3a–c, ES1†). As shown in Fig. 3a, the V $2p_{3/2}$ spectrum exhibited a V^{+4} feature at binding energy 515.9 eV, and a feature at 516.8 eV indicative of the surface oxidation of vanadium to the V^{+5} state, which has been reported for vanadium oxide samples.⁴³ A V $2p_{1/2}$ signal was observed at binding energy 523.8 eV. The O 1s spectrum revealed the presence of V–O bonds at 529.45 eV and $V^{+4}-O^{-2}$ bonds at 530.14 eV (Fig. 3b). Another peak at 531.77 eV indicated OH^- on the VO_2 surface.²⁸ The N 1s spectrum of $VO_2@CN$ depicted the presence of three peaks at binding energies of 398.7 eV, 399.6 eV and 401.0 eV, which can be assigned to pyridinic N, pyrrolic N and graphitic N, respectively (Fig. 3c).²⁷ In contrast, the analysis of the valence state of the precursor $VO_2@mel$ (Fig. S1†) revealed the prominent presence of the V^{+5} state in the V 2p XPS results (binding energy 516.9 eV for V $2p_{1/2}$). The N 1s XP spectrum of $VO_2@mel$ revealed the presence of three peaks ascribed to the ring C=N of melamine (398.5 eV), exocyclic $-NH_2$ of melamine (399.6 eV) and nitrate counter anion (406.0 eV), attributed to the presence of $VO_2(NO_3)$ in $VO_2@mel$.²⁸

The scanning electron microscopy (SEM) images of $VO_2@CN$ revealed vanadium dioxide particles as rectangular block-shaped structures with a length of $\sim 500\text{ nm}$ amidst larger crumpled sheets of N-doped carbon, which did not show any significant alteration from the morphology of pristine N-doped carbon materials (Fig. 3d and S5†). The larger size of the VO_2 rods could be attributed to the favourable grain growth of nanorods at elevated temperatures.³⁸ The presence and uniform distribution of the constituent elements V, O, C and N was attested from the EDX and elemental mapping analysis (Fig. S6†). The transmission electron micrographs (TEM) of $VO_2@CN$ supported the presence of VO_2 blocks encased in crumpled N-doped carbon layers (Fig. 3e). HRTEM confirmed the presence of VO_2 encapsulated in N-doped carbon from the lattice fringes corresponding to a d -spacing value of 0.32 nm (Fig. 3f) indicative of the (011) plane of VO_2 (M). The SAED pattern depicts the polycrystalline nature of the $VO_2@CN$ sample supporting the observation of N-doped carbon-encased VO_2 blocks (Fig. S7†).

Electrochemical NRR performance of $VO_2@CN$

Vanadium in various forms has been a popular candidate having abilities to reduce dinitrogen electrochemically.^{7,43} Mostly, vanadium nitrides are investigated for the NRR, but the involvement of lattice nitrogen and the unstable metal centre have prompted researchers to explore the oxy-nitride class of materials, where the oxygen-supported vanadium centres were proved to be energetically more favourable towards N_2 scission as compared to the nitrogen supported



Fig. 3 XPS spectra of $\text{VO}_2@\text{CN}$: (a) V 2p, (b) O 1s and (c) N 1s. (d) FESEM image of $\text{VO}_2@\text{CN}$ indicating the presence of blocks of VO_2 encapsulated in crumpled N-doped C layers. TEM and HRTEM images of $\text{VO}_2@\text{CN}$ are presented as (e) and (f), respectively.

ones.⁴⁴ Investigation of V_2O_5 nanodots as an electrocatalyst for the NRR afforded a faradaic efficiency (FE) of 28.7%, attributable to the presence of oxygen vacancies that participated in the electron transfer to the V^{+5} state, to effectively generate V^{+4} , thus underlining the importance of the V^{+4} species in electrocatalytic NRR.²⁸ Moreover, vanadium-based systems were mostly found effective for the NRR at neutral pH, and have rarely been explored under acidic conditions in order to suppress the competing HER. Since the challenging part in the NRR is the bond polarization of N_2 leading to its first protonation and subsequent conversion to ammonia, it forms the motivation for the catalyst ($\text{VO}_2@\text{CN}$) synthesis in this work. All the potentials reported in this work were calibrated with respect to a reversible hydrogen electrode (RHE) following eqn (1) (ESI†). It was shown by Biswas *et al.* that the choice of the electrolyte is important as different catalysts respond differently in the presence of the ionic environment of the electrolyte (effect of the counter ions), and it crucially determines the nitrogen reduction performance of the catalyst.^{45,46} The ions from the electrolyte form a Stern layer across the electrode that mediates the adsorption of N_2 and repels H^+ from the electrode.⁴⁷ It should be noted that the rationale behind avoiding the basic electrolyte is attributed to sample instability in the medium. Thus, in this study both acidic (0.1 M HCl) and neutral (0.1 M Na_2SO_4) electrolytes were explored, and a comparative study was accomplished. Before any electrochemical tests, the electrolyte solutions were degassed with Ar followed

by N_2 purging after passing the N_2 feed gas through base and acid traps since contamination is unavoidable in the NRR and this calls for all precautionary measures to achieve reliable data. The electrochemical cell set-up used in this study is shown in Fig. S8.† As shown in Fig. S9,† the trap solutions were subjected to colorimetric tests for the detection of any dissolved NO_x or NH_4^+ that would have been present in the feed gas. The UV-visible spectra for the N_2 -trapped solutions were overlaid with those of the bare ones, which indicated that there were negligible gaseous nitrogenous impurities in the feed gas itself.

Thereafter, an elaborate study with chronoamperometry (CA) at different potentials was carried out in both the electrolytes for 1 h each (Fig. S10a and c†). The post-electrochemistry UV-visible studies with the electrolytes stained with colouring agents (indophenol blue method), depicted a maximum absorbance corresponding to the indophenol dye formed out of the produced ammonia at -0.1 V vs. RHE (Fig. S10b and d†). The material $\text{VO}_2@\text{CN}$ displayed a better ammonia yield rate (calculated using eqn (2), ESI†) in 0.1 M HCl as shown in the bar plot in Fig. 4a, while the FE (calculated using eqn (2), ESI†) in Fig. 4b was better in the case of Na_2SO_4 . In fact, in the case of HCl, the FE showed a drastic reduction as we approached more closely towards a negative potential, owing to enhanced hydrogen evolution reaction (HER) kinetics. The acidic electrolyte with abundant proton sources helped to protonate the produced ammonia, which remained dissolved in the electrolyte



Fig. 4 NRR performance comparison of the VO₂@CN catalyst in both 0.1 M Na₂SO₄ and 0.1 M HCl electrolytes. (a and b) NH₃ yield rate and FE of the catalyst over the wide-ranging potential window from 0.0 V to -0.4 V vs. RHE in the presence of continuous N₂ purging; (c) quantitative analyses of NRR performance of VO₂@CN using UV-visible and NMR methods; CV curves of VO₂@CN and VO₂@mel samples in (d) 0.1 M HCl and (e) 0.1 M Na₂SO₄ showing an enhancement in the current density upon graphitization of melamine due to increased conductivity; (f) NH₃ yield rate and (g) FE comparison of VO₂@CN with control samples like VO₂@mel and g-CN in both acidic and neutral electrolyte media. The error bars represent the standard deviation values for three different data-sets.

in the form of NH₄⁺ and ensured no loss as NH₃ gas. However, at the same time, it also induced proton adsorption on the catalyst surface since the HER is a less energy-intensive process than the NRR, which caused the catalyst to incur a sacrificial FE over the entire potential range. With a maximum FE of 67.9% and 61.9% in Na₂SO₄ and HCl respectively at -0.1 V vs. RHE, the material was also proved to be selective towards ammonia production as there was no interference from side-product (hydrazine) formation (Fig. S10e and f†). This could be justified with respect to a very recent report by Huang *et al.* that experimentally and theoretically suggested that the NRR

on V active sites could energetically prefer a distal pathway negating the chances of the formation of hydrazine, rightly corroborating our experimental results.⁴⁸ Thus, the NRR on our active material of VO₂@CN was kinetically facilitated yielding 0.31 and 0.52 μmol h⁻¹ mg_{cat}⁻¹ of NH₃ in Na₂SO₄ and HCl, respectively. The concentrations of ammonia in all cases were calculated using the standard calibration curves and the corresponding linear regression plot in the relevant acidic or neutral electrolytes (Fig. S11 and S12†). The interference from hydrazine formation was interpreted from the standard calibration and corresponding linear-fitted curves in both the elec-

trolytes, as shown in Fig. S13.† It is noteworthy to mention that the bare Ti foil (pre-cleaned and used instantaneously in order to avoid any surface oxidation layer) showed an FE of 3.24% and 9.5% in 0.1 M Na₂SO₄ and 0.1 M HCl, respectively, as shown in the bar diagram in Fig. S14,† which indicated that there was only a slight interference of the substrate towards the NRR and it was the right choice to purposely use Ti foil as a substrate as it helped to elevate material performance.

Since the nuclear magnetic resonance (NMR) method provides more accurate quantification, it was adopted to verify the NH₃ yield rate and FE in acidic and neutral electrolytes. The details are provided in the ESI.† It is worth noting from Fig. S15† that, in both media, while there was no peak under Ar-purged conditions, a triplet appeared in the presence of N₂ resembling ¹⁴NH₄⁺ (coupling constant value of 52 Hz calculated using eqn (3) in the ESI†). Thereafter, a quantitative calculation using maleic acid as the standard (eqn (4) from the ESI†) revealed similar NH₃ yield rates from both UV-visible and NMR methods, as shown in Fig. 4c, that justify the reliability of the experimental data. However, it is true that without control experiments it is not rational to claim the superiority of the catalyst. Thus, we took into consideration two control samples, *i.e.* the precursor VO₂@mel, and g-CN for the study as the former would show the importance of a conducting carbon framework to anchor the active centres while the latter would justify the role of VO₂ active sites of the catalyst for the NRR (detailed NRR studies and UV-visible data are shown in Fig. S16†). As expected, the CV curves at 10 mV s⁻¹ in Fig. 4d and e in 0.1 M HCl and 0.1 M Na₂SO₄ respectively show an enhancement in the areal current as well as cathodic current density for VO₂@CN as compared to VO₂@mel signifying the enhancement of the electron conductivity due to the graphitization of melamine to the CN framework. However, the NH₃ yield rate and FE for VO₂@mel, although sound enough, were lower when compared to that of VO₂@CN (Fig. 4f and g) because V was primarily in the V⁺⁵ oxidation state, which reduced the π-back donation to the LUMO of the N₂ molecule; thus the cleavage of N≡N bond was relatively less. So, a long-term CA experiment would result in a loss of active sites from the catalyst surface. In contrast to this, the NRR performance of the g-CN sample was very low, being devoid of VO₂ units (Fig. 4f and g). This clearly manifested the role of VO₂ as the NRR active site with the simultaneous importance of the conjugated carbon–nitride framework in the pronounced electrocatalytic performance of our final material VO₂@CN towards the NRR.

Mechanistic investigation and identification of NRR intermediates by *in situ* ATR-FTIR studies. The electrochemical NRR is a complex multi-step process, where the N₂ activation and reduction depend largely on the nature of the catalyst active site. The orbital interactions accompanying σ-donation and π-back donation between the active site and N₂ serve as the primary steps of the NRR initiating N₂ adsorption and activation. V is known to catalyze the NRR and, in this work, the V-center of the VO₂ unit in VO₂@CN is considered the active unit where the presence of adjacent O atoms could reorganize

the electron occupancy of the molecular orbital levels of metallic VO₂ further assisted by the N-doped carbon support.⁴⁹ Thus, the d_{||} and π* orbitals of VO₂ could likely participate in the respective bonding–anti-bonding orbital interactions with the σ-highest occupied molecular orbital (HOMO) and π*-lowest unoccupied molecular orbital (LUMO) of N₂, as schematically shown in Fig. 5a. This interaction is followed by the N≡N bond polarization and the protonation steps that produce several obscured reaction intermediates on the catalyst surface. Out of these, the stable intermediates are detected in the *in situ* ATR-FTIR spectra including those attributed to peaks at 1307 and 1460 cm⁻¹ observed in both electrolytes *i.e.*, 0.1 M Na₂SO₄ and 0.1 M HCl (Fig. 5b and c), resembling H–N–H rocking and N–H bending, respectively.⁵⁰ It is important to note that the peak corresponding to N–N stretching could be attributed to a stable H₂N–NH₂ intermediate on the catalyst surface and the corresponding peak appeared to be slightly red-shifted in HCl (1139 cm⁻¹) than in the neutral electrolyte (1143 cm⁻¹) suggesting the better extent of N–N bond polarization in the acidic medium and faster reaction kinetics. This also supported our claim for product (NH₃) selectivity of the material as we did not obtain any trace of N₂H₄ from the UV-visible spectra. Furthermore, the abundance of protons in the acidic electrolyte activated the polarized N₂ molecules giving a better yield rate for NH₃ production with a comparatively faster kinetics. However, it is also noteworthy that the FTIR spectra in the acidic electrolyte were much more noisy owing to the simultaneous evolution of H₂ (HER) on the catalyst surface along with the NRR, which although produced NRR intermediates and products, the current thus obtained contributed to the HER resulting in a sacrificial FE for the NRR in HCl rather than in Na₂SO₄. Thus, it is interesting to note that keeping all the electrochemical conditions the same, different electrolytes create different local electronic/ionic environments at the catalyst–electrolyte interface and variedly influence the NRR at the active site.

NRR stability studies and post-stability material characterization. Since the VO₂@CN catalyst shows a significant NRR performance in both acidic and neutral media, it was obvious to note a distinct deviation in the CV curves for the material under both Ar- and N₂-purged electrolyte conditions (Fig. S17†). For the sake of negating the possibilities of false positive results, several control CA experiments, like (a) VO₂@CN in the presence of Ar at –0.1 V and (b) VO₂@CN in N₂ at the open circuit potential (OCP), were conducted (Fig. S18a–d†). Although, at the OCP, there was negligible production of ammonia, under an Ar atmosphere, there was a nominal yield rate and FE for NH₃ wherein, the lattice N could partially partake in the electrocatalytic process following the Mars van Krevelen mechanism.³² However, this was trivial when compared to that obtained in the presence of N₂ in both electrolytes. Furthermore, the sustainability of the catalyst was investigated with stability studies, where in both electrolytes after 12 h of continuous CA experiment (Fig. 6a–d), the catalyst showed a drop in ammonia production and NRR performance (Fig. 6c). This could be ascribed to (a) material loss from the



Fig. 5 (a) Orbital interaction between the VO_2 part of $\text{VO}_2@\text{CN}$ and N_2 in terms of a molecular orbital diagram; unprocessed *in situ* ATR-FTIR spectra of the catalyst at different working potentials under N_2 - and Ar-purged conditions in (b) 0.1 M Na_2SO_4 and (c) 0.1 M HCl , respectively.

electrode surface or (b) leaching of V-active sites from the catalyst upon prolonged exposure to the electrolyte under a constant potential of -0.1 V with the conditions being more prominent in the case of HCl (FE reduces from 61.9% to 45.7% after 12 h) than Na_2SO_4 (FE decreases from 67.9% to 55.2% after 12 h). These data are in accordance with several reports on vanadium-based systems for the NRR.^{51,52} However, they did not suggest any effect on the morphology of the post-stability $\text{VO}_2@\text{CN}$ sample as is evident from the FESEM and TEM images shown in Fig. S19a and b† along with the uniform presence of all the constituent elements like C, N, O and V on the electrode surface, as shown from the elemental mapping in Fig. S19c and d.† In fact, the XRD spectrum in Fig. S20a† displayed combined features of both $\text{VO}_2@\text{CN}$ and the underlying carbon cloth (used as a substrate for stability studies). To distinguish the peaks, the spectrum was zoomed and compared with that of the bare carbon cloth (CC) and fresh $\text{VO}_2@\text{CN}$ samples, wherein all the peaks appeared to have been retained in the post-electrolysis sample (Fig. S20b†) suggesting the persistent lattice structure of the material. However, there was some change in the chemical bonding of the material upon prolonged electrolysis, as demonstrated from the XPS studies.

XPS studies were conducted for $\text{VO}_2@\text{CN}$ samples after the electrochemical NRR stability experiments under HCl and Na_2SO_4 , respectively, and compared with the pristine sample

(Fig. 3a–c and 6d–f). The XPS data reveal minor alterations in the surface structure of the $\text{VO}_2@\text{CN}$ catalyst after electrochemical NRR. Under both pH conditions, the V 2p spectrum shifts towards a slightly lower binding energy, indicative of the removal of the surface oxidized V^{+5} (BE 516.4 eV in post-NRR samples *vs.* 516.8 eV in pristine) species that was observed in the pristine sample (Fig. 3a). In all cases, the pronounced V 2p_{3/2} maxima around the binding energy of 516 eV indicated V^{+4} as the major species (Fig. 6d).⁵³ However, there is a reduction in the atomic percentage of vanadium in the post-NRR samples, which can be ascribed to the leaching of VO_2 species in solution during the NRR. There is a significant increment in the post-NRR intensity of the O 1s signals around the binding energies of ~ 531 eV and 532.4 eV ascribed to the prominence of oxide and adsorbed hydroxyl species, respectively (Fig. 6e). Unlike the case of the oxynitride species reported in the literature, the shift of the N 1s signal to lower binding energies (*i.e.* towards 396.5 eV)³² and any significant depletion in the N content of N-doped C were not observed in our case (Fig. 6f). There was no evidence of N-incorporation in $\text{VO}_2@\text{CN}$ from the N_2 -saturated electrolyte under NRR conditions, consistent with the literature reports.⁵⁴ It should be mentioned here that the changes in the core level spectra of V 2p, O 1s and N 1s are induced by cathodic polarizations in HCl and Na_2SO_4 electrolytes, and not due to exposure under ambient conditions.⁵⁵ The greater ease in protonation of the



Fig. 6 (a and b) CA response of VO₂@CN during the NRR at -0.1 V vs. RHE in 0.1 M Na₂SO₄ and 0.1 M HCl, respectively, for a continuous 12 h period; (c) FE comparison of the catalyst after 1 and 12 h of NRR studies in both electrolytes; post-NRR XPS profiles for (d) V 2p, (e) O 1s and (f) N 1s, respectively, under HCl and Na₂SO₄ electrolyte conditions.

surface-bound N₂ on vanadium oxide suggests that the increased acidity of the vanadium active site bound to the electronegative oxygen atoms and encapsulated in N-doped carbon stabilized the intermediates of the NRR and facilitated dinitrogen cleavage. This trend has been mentioned for other metals with high oxophilicity, including NbO₂, which exhibited a high FE for the NRR.⁵⁶ Thus, it can be concluded that there are several factors that rule the choice of electrolyte in the NRR. It is not a good practice to go with the catalyst performance for a 1 h potential-dependent study as the electrolyte might have a poisoning effect on the catalyst active site by either blocking or causing leaching of the active centre. This fact has relevance with respect to this study as it reveals that although the yield of NH₃ was better in the acidic electrolyte, Na₂SO₄ stood out as a better contender in terms of FE and the stability study of the VO₂@CN catalyst.

Experimental section

Synthesis of VO₂@CN

4 g of melamine was added to a 250 mL round bottomed flask containing 50 mL of methanol. The solution was then stirred at room temperature for about 30 minutes. 300 mg of vanadium pentoxide dissolved in 20 mL of 3 M nitric acid was

added to the solution containing melamine under continuous stirring. The reaction mixture was stirred for another hour under ambient conditions. The precipitate was evaporated to dryness.

The obtained VO₂@melamine (VO₂@mel) was transferred into an alumina boat placed inside a tubular furnace and heated to 600 °C under continuous nitrogen flow for 3 h to obtain a black powder named VO₂@CN.

Conclusions

In our quest to identify an inexpensive, Earth-abundant material as a universal-pH electrocatalyst for the NRR, we explored vanadium oxide encased in N-doped carbon, an Earth-abundant transition metal oxide that is synthesizable in bulk quantities. The data reported here designate VO₂@CN as an efficient and selective electrocatalyst for nitrogen fixation with a high NH₃ yield of 0.31 and 0.52 μmol h⁻¹ mg_{cat}⁻¹ and a maximum FE of 67.9% and 61.9% at -0.1 V vs. RHE, under neutral and acidic conditions, respectively, without any observable incorporation of nitrogen into the vanadium oxide lattice. Furthermore, *in situ* ATR-FTIR studies at different potentials under N₂ and Ar revealed the obscured reaction intermediates of the NRR and justified the differential activity

trend of VO₂@CN under both electrolyte conditions, owing to the possible impact of the competitive HER. In addition, the selectivity of the material was ascertained from the absence of side-products like hydrazine and suppression of the competing hydrogen evolution reaction. Our results demonstrate the competence of vanadium oxide on N-doped carbon as an electrocatalyst for the NRR under variable pH conditions, and substantiates oxophilicity as an important factor in designing electrocatalysts.

Author contributions

AC: methodology, formal analysis, data curation (synthesis and characterization), and writing – original draft, AB: methodology, formal analysis, data curation (electrocatalytic NRR), and writing – original draft. SP: data curation (characterization) and writing – review and editing. RSD: funding acquisition, supervision (AB), and writing – review and editing. JM: funding acquisition, supervision (AC and SP), writing – original draft, and writing – review and editing.

Conflicts of interest

There are no conflicts to declare.

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