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The multivariate interaction between Au and TiO₂ colloids: the role of surface potential, concentration, and defects†

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The established DLVO theory explains colloidal stability by the electrostatic repulsion between electrical double layers. While the routinely measured zeta potential can estimate the charges of double layers, it is only an average surface property which might deviate from the local environment. Moreover, other factors such as the ionic strength and the presence of defects should also be considered. To investigate this multivariate problem, here we model the interaction between a negatively charged Au particle and a negatively charged TiO₂ surface containing positive/neutral defects (e.g. surface hydroxyls) based on the finite element method, over 6000 conditions of these 6 parameters: V_{Part} (particle potential), V_{Surf} (surface potential), V_{Def} (defect potential), DD (defect density), $Conc$ (salt concentration), and R (particle radius). Using logistic regression, the relative importance of these factors is determined: $V_{\text{Surf}} > V_{\text{Part}} > DD > Conc > R > V_{\text{Def}}$, which agrees with the conventional wisdom that the surface (and zeta) potential is indeed the most decisive descriptor for colloidal interactions, and the salt concentration is also important for charge screening. However, when defects are present, it appears that their density is more influential than their potential. To predict the fate of interactions more confidently with all the factors, we train a support vector machine (SVM) with the simulation data, which achieves 97% accuracy in determining whether adsorption is favorable on the support. The trained SVM including a graphical user interface for querying the prediction is freely available online for comparing with other materials and models. We anticipate that our model can stimulate further colloidal studies examining the importance of the local environment, while simultaneously considering multiple factors.

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1. Introduction

The conventional Derjaguin–Landau–Verwey–Overbeek (DLVO) theory describes colloidal interactions by balancing the van der Waals attraction and the electrostatic repulsion between the double layers. If the resulting energy barrier is larger than the thermal energy (kT), then the colloidal suspension is predicted to be kinetically stable against aggregation,^{1,2} which is usually the desired behavior. However, there are occasions when the irreversible collision of particles is preferred, such as adsorbing metal nanoparticles onto oxide supports. To facilitate this adsorption, a common strategy is to mix metal and oxide colloids at a pH value

where the adsorbates and adsorbents are oppositely charged, such that they are attracted to each other (by both electrostatic and van der Waals contributions). Experimentally, this condition is fulfilled when the pH is between the isoelectric points (IEPs) of the two materials.³ The IEP is the pH when the particle has a net surface charge/zeta potential of zero. At a pH below the IEP, the material is positively charged, whereas at a pH above it, the material is negatively charged.⁴ Accordingly, by mixing the metal and oxide particles at a pH value between their IEPs, the two species are oppositely charged to encourage the electrostatically induced adsorption, while avoiding their self-agglomeration due to the repulsion between the same materials. As an example, Pt and TiO₂ have IEPs at around 2.8 and 6.4 respectively, and hence a pH value of 5 would be appropriate to induce the adsorption between them with an attractive Coulomb force.⁵ Importantly, the IEP is an experimentally measurable quantity reflecting the surface charges of particles, and hence it is usually a convenient metric for predicting whether the interaction between particles is favorable or not.^{6–10}

However, even when both particles are negatively charged and experience electrostatic repulsion between them, adsorption might still be possible. Nanoparticles with an absolute

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zeta potential larger than ± 30 mV generally possess an electrostatic energy barrier that is large enough to prevent the agglomeration and aggregation induced by the attractive van der Waals forces at room temperature.^{7–10} At 298 K, the thermal voltage ($\frac{kT}{e}$), which can be thought of as the potential equivalence of thermal energy (kT), is approximately 25.7 mV. This implies that if the absolute potential is smaller than 25.7 mV, particles following a Boltzmann distribution on average would possess sufficient kinetic energy to surpass the energy barrier, leading to coalescence.¹¹ Consequently, the common criterion of ± 30 mV is a conservative estimate, ensuring that the majority of particles in the Boltzmann distribution do not have adequate energy to overcome the barrier. On the contrary, with zeta potentials below this threshold, metal colloids can still be deposited onto a support because they are not repulsive enough. In this context, Wagener *et al.* studied the interaction between negatively charged Ag colloids and BaSO₄ microparticles. The zeta potential of the BaSO₄ support was kept constant at -2.5 mV, but that of the Ag nanoparticles was adjusted to more negative values by adding more citrate ligands to the mixture (-25 mV to -75 mV). It was observed that the adsorption between Ag and BaSO₄ only occurred when the Ag colloids obtained a zeta potential less negative than -30 mV.¹² However, it is also noted that the inclusion of surfactants in this case can complicate the interpretation because it also adds a small steric contribution to the question.

Although the zeta potential is usually a satisfactory descriptor for the electrostatic interaction between particles, it is only the weighted average potential at the shear/slipping plane of all the particles in the dispersion.¹³ It is known that the nanoenvironment around nanoparticles is different from the bulk, and the local surface potential might not be well represented by the global zeta potential.¹⁴ Following the DLVO theory, Au and P25 TiO₂ colloids would have an energy barrier larger than kT in the electrostatically repulsive regime of pH 8.5, hindering the adsorption between them.³ Nevertheless, quantitative adsorption of Au onto TiO₂ was observed up till 20 wt%,³ suggesting that the zeta potential and the traditional DLVO theory failed to capture the essence of this interaction. A possible reason for that is the deviation of the local environment from the bulk average. As an example, atomic force microscopy (AFM) experiments performed on single crystals of rutile TiO₂ demonstrated that different facets do not share the same IEP, *e.g.* (110): 5.5–4.8 and (100): 3.7–3.2.¹⁵ Unsurprisingly, the bulk IEP of rutile TiO₂ (5.2–6.8)⁴ matches well with the dominant facet of (110). However, it is notable that while (100) only occupies a small proportion of the exposed surfaces in rutile, the IEP of (100) is lower than the bulk value by as much as 2 units. This facet-dependent trend in IEP was also reported for SrTiO₃ (110) and (100) in an independent AFM study by Su *et al.*¹⁶ Despite being on the same particle, (110) and (100) are expected to interact very differently with metal particles, and these distinctive facets cannot be collectively described by a single metric of zeta potential. Moreover, even within the same facet, the charge distribution does not necessarily have to be uniform across the whole

surface due to the presence of point or volume defects. It was demonstrated by simulation studies that surface heterogeneity and roughness can significantly reduce the repulsive energy barrier compared to a perfect surface.^{17,18} Specifically for TiO₂, defects such as oxygen vacancies are often referred to as the favorable adsorption sites for Au nanoparticles in the DFT and catalysis literature due to the distorted electron density near the surface.^{19–24} In addition, surface hydroxyls can also serve as local charge deviations,^{25–27} as well as affecting short-range contributions such as hydration forces.^{28,29} The pH of the system plays a crucial role in the behavior of surface hydroxyls. As the pH increases, more surface hydroxyls are deprotonated, leading to a more negatively charged oxide surface.^{30–34} However, the surface hydroxyls are only completely deprotonated at highly alkaline pH (*e.g.* 11–12), since the pK_a of TiO₂ basic hydroxyls is ~ 9 .^{35,36} This means that at moderately alkaline pH (*e.g.* 8–9), a considerable proportion of hydroxyls are still protonated and could serve as point defects, since they are less negative than the rest of the surface (formally neutral for terminal Ti–OH and positive for bridging Ti–OH–Ti). Furthermore, it is also suggested that the enriched electron density near the surface due to Ti³⁺/oxygen vacancies can further alter the abundance of surface hydroxyls.³⁷ Clearly, the local environment exhibits a much more subtle surface potential than the zeta potential manages to describe, and this should be considered particularly in the repulsive regime when the interacting particles carry the same charge sign.

While it might be experimentally challenging to examine the local environment when the particles come into contact, it is relatively easier to model particles with local surface charges deviating from the rest of the surface, and simulate how they interact. As discussed, when both the adsorbing nanoparticle and the support are negatively charged, the zeta potential alone might not be a sufficient predictor for the interaction outcome. Therefore, other parameters have to be taken into account, such as the density and potential of defects on the support. To evaluate the effect of each factor, it is desirable to perform a parametric sweep across a wide range of physically realistic conditions. To keep this computationally feasible while capturing the essence of the interaction, we constructed an empirical model based on the finite element method. We specifically consider the Au/TiO₂ system, which is a popular catalyst for reactions such as hydrogen production,³⁸ alcohol oxidation,^{39–41} and CO oxidation.^{42–44} Particularly for CO oxidation, the Au/TiO₂ interface has been identified as the active site for the reaction,^{45–48} and the strong metal–support interaction makes it a good catalyst.^{49,50} In our model (Fig. 1), we depict the interaction between a negatively charged sphere (Au) and a negative surface with positive/neutral defects (TiO₂), where the point defects (*e.g.* surface OH groups) are local sites with a surface potential different from the normal surface. To obtain a comprehensive picture, a total of 6000 unique conditions are simulated by systematically varying the following parameters: V_{Part} (particle potential), V_{Surf} (surface potential), V_{Def} (defect potential), DD (defect density), $Conc$ (salt concentration), and R (particle radius). The van der Waals inter-



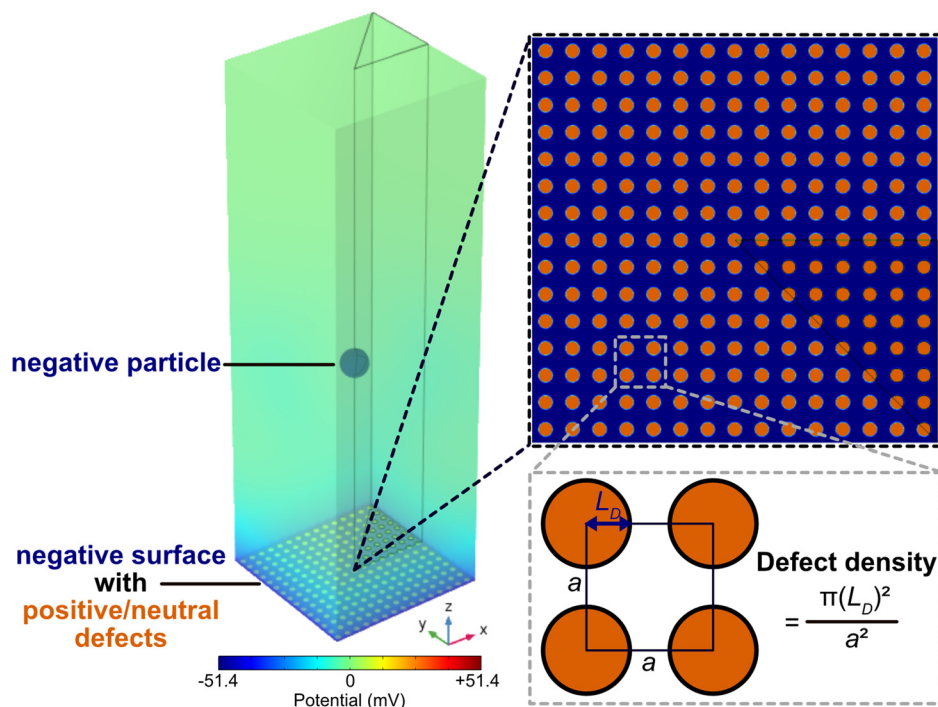


Fig. 1 Model configuration of a negative particle (e.g. Au) interacting with a negative surface (e.g. TiO₂) with positive/neutral defects. A defect density of 0.2 is illustrated in this example.

actions between typical Au and TiO₂ particles are also considered. Our results show that although the surface potential (V_{Surf} and V_{Part}) remains to be the most decisive factor, the defect density and the salt concentration in solution also seem to play a significant role in the adsorption between Au and TiO₂. Using logistic regression and a support vector machine (SVM), we also determined the transition boundary between the “favorable” (below kT) and “unfavorable” (above kT) cases. Hence, it is possible to predict the outcome of the interaction more confidently by considering multiple parameters at the same time, rather than only the zeta potential. Our study highlights the importance of local surface deviations in colloidal interactions, and we hope this work will encourage further research into the effects of local environments.

2. Methods

2.1. Model configuration and parameter space

We consider the interaction between Au and TiO₂ in an aqueous electrolyte solution, with a pH higher than the isoelectric points of both materials, but only at a mildly alkaline pH (e.g. 8). Under these conditions, both Au and TiO₂ are overall negatively charged, but not all surface hydroxyls on TiO₂ are deprotonated,^{30,31,35,51} and hence the remaining protonated species could serve as local charge deviations, which we refer to as defects. We model this particular Au/TiO₂ system by a negative sphere coming into contact with a negative surface containing positive/neutral defects, which is depicted schematically in Fig. 1. When Au nanoparticles are adsorbed

on TiO₂, the former is typically much smaller (at least 10 times) than the latter, and therefore the curvature of TiO₂ is also significantly smaller relative to the Au colloids. Accordingly, we model the TiO₂ surface as a flat plane. More importantly, TiO₂ particles have preferentially exposed facets, which result in a rod-like or elongated shape rather than being spherical. For instance, the most stable polymorph rutile TiO₂ is mostly terminated by (110) surfaces followed by (011) surfaces as predicted by Wulff construction, resulting in shapes ranging from a needle-like crystal to a short blocky habit,^{52,53} which can be sufficiently described as a flat plane for Au nanoparticles much smaller than TiO₂. We also only consider the case of diluted Au colloids so Au–Au interactions can be ignored. On the other hand, the defects are modeled by small cylinders with a radius (L_D) of 0.5 nm and a height of 0.1 nm. These defects are tiny protrusions in the z -direction because the oxygen atoms of surface hydroxyls are roughly 0.1 nm above the remaining (110) surface as measured experimentally by scanning tunneling microscopy⁵⁴ and atomic force microscopy.⁵⁵ For simplicity, the defects are uniformly arranged in a 2D primitive square array with a lattice parameter of a , where the defect density (DD) is defined as

$$DD = \frac{\pi(L_D)^2}{a^2}. \quad (1)$$

On the idealized rutile (110) surface, the surface oxygen atoms are ordered in a similar but rectangular fashion.⁵⁶ By tuning DD , it is possible to pack the defects more densely or sparsely to match the more realistic case.



Table 1 Parameter space of simulations, which covers a vast range of physically relevant conditions: 6000 unique combinations of these 6 parameters (V_{Part} , V_{Surf} , V_{Def} , DD , $Conc$, R) are considered. For each condition, the particle is moved towards the surface in 15 steps (D). Another 1260 conditions are simulated for perfect (uniformly charged) surfaces without defects (Table S2†)

	Parameter	Values	Unit
V_{Part}	Particle potential	−51.4, −38.5, −25.7, −12.8	mV
V_{Surf}	Surface potential	−51.4, −38.5, −25.7, −12.8	mV
V_{Def}	Defect potential	0.0, +12.8, +25.7, +38.5, +51.4	mV
DD	Defect density	0.1, 0.2, 0.25, 0.3, 0.35	—
$Conc$	Salt concentration	0.1, 0.5, 1.0	mM
R	Particle radius	0.5, 1.5, 2.5, 3.5, 4.5	nm
D	Particle–surface separation	0.1–4.1 (15 steps)	κ^{-1} (Debye length)

In order to study the interaction between Au and TiO_2 over an extensive range of realistic conditions, a parametric sweep was performed for these six factors: V_{Part} (particle potential), V_{Surf} (surface potential), V_{Def} (defect potential), DD (defect density), $Conc$ (salt concentration), and R (particle radius). A total of 6000 unique conditions are considered, and the parameter space is summarized in Table 1. Temperature is kept constant at 298 K throughout. Importantly, this parameter space is chosen to survey a large scope of physically relevant conditions. For example, both the particle and the flat surface have V_{Part} and V_{Surf} varied up to −51.4 mV, which is considerably repulsive given that an absolute zeta potential larger than ± 30 mV is generally considered to be sufficient to resist particle agglomeration and aggregation.^{7–10} Similarly, the maximum V_{Def} is set to be +51.4 mV for highly charged defects, while the minimum V_{Def} of 0 mV represents neutral defects.

The cases with minimal defects and the scenarios with full defect occupancy ($DD = 0.1$ – 0.35) are also considered. According to our definition of DD (eqn (1)), if all the surface oxygen atoms are protonated, this would give a DD of ~ 0.37 . (The surface density of unsaturated oxygen atoms is 5.2 nm^{-2} on rutile (110).⁵⁷ The diameter of an oxygen atom is $\sim 0.3 \text{ nm}$.⁵⁶ Therefore the maximum DD is $5.2 \times \pi \times (\frac{0.3}{2})^2 = 0.37$.) However, this extreme case should be highly unlikely at alkaline pH because the extent of deprotonation of surface OH increases with pH, so more moderate DD values (0.1, 0.2, 0.25, 0.3) are also investigated. In contrast to the defective (non-uniformly charged) surfaces, we also separately simulated a control model with only perfect (uniformly charged) surfaces ($DD = 0$). Details of the comparisons are provided in the ESI (Fig. S2–S6†). The objective is to identify the conditions where defects exert the largest influence at the nanoscale. To achieve this, we evaluated the interaction energy barriers when both perfect and defective surfaces have the same effective surface potential (V_{Eff} , the average potential of the entire surface). When defects play a more critical role, the deviations in the energy barriers are expected to be more pronounced between the two scenarios. To quantify this discrepancy, the coefficient of determination (R^2) is calculated, where a smaller R^2 value indicates a more significant impact of the defects. The results show that while there are some instances where the perfect surface offers a reasonable descrip-

tion of the defective surface with the same V_{Eff} , a considerable dispersion in the interaction energy barriers persists in most other cases, which can only be accounted for when defects are explicitly considered. Hence, defects are included as local charge deviations in our model.

$Conc$ spans salt concentrations commonly found in experiments (0.1, 0.5, 1.0 mM), and only symmetric monovalent electrolytes (e.g. NaCl) are considered, which means that the ionic strength (I) is equal to the salt concentration (c_0): $I = \frac{1}{2} \sum_i c_i z_i^2 = \frac{1}{2} [c_0(+1)^2 + c_0(-1)^2] = c_0$. We also note that it is the ionic strength which ultimately affects the Debye length, but not the salt concentration except for the case of symmetric monovalent salts. For R , the particle radius scales with the defect radius ($R = 1$ – $9 L_D = 0.5$ – 4.5 nm , where $L_D = 0.5 \text{ nm}$), so the actual physical size of defects is relatively less critical. To limit the simulation time, only one-eighth of the square prism is simulated (indicated by the solid black line in Fig. 1), and the rest of the model is generated by symmetry. Further simulation details and model snapshots can be found in the ESI (Fig. S1†).

For each particular set of conditions, the particle is moved progressively towards the defect-containing surface (similar to how real particles might diffuse) to examine how the system energy changes (0.1–4.1 Debye lengths, 15 steps). In this work, only two energy components are considered: the electrostatic free energy (F_{el}) stemming from the electrical double layer, and the van der Waals energy (E_{vdw}) resulting from the induced dipole interactions. Other contributions (e.g. hydrophobic forces and surface energies) are not included. While hydrophobic forces and surface energies might provide an additional thermodynamic driving force for the colloids to interact, they are only significant when the colloids are in close proximity. In other words, these energies are only crucial if the colloids have sufficient energy to first overcome the kinetic barrier to come close to each other. This kinetic barrier is computed by combining F_{el} and E_{vdw} in this study.

2.2. Electrostatic free energy (F_{el})

The electrostatic interaction between colloids in an aqueous medium is governed by the potential (ϕ or ψ) of the electrical double layer. We denote ϕ as the dimensional potential with a unit of V, and ψ as the dimensionless potential. The additional



subscript of φ_0 clarifies that this is the potential at the surface, but not that in the electrolyte. The potentials can be interconverted by multiplying/dividing by the thermal voltage (eqn (2)). At 298 K, 1 dimensionless potential translates to approximately 25.7 mV.

$$\psi = \frac{e\varphi}{kT} \quad (2)$$

For a more intuitive understanding, the dimensional potential φ is used predominantly throughout the text, which shares the same unit (V) with the experimentally measured zeta potential. However, it is often more convenient to express equations and run simulations with the dimensionless potential ψ , and therefore we present the equations here with a mixture of φ and ψ .

We calculate the potential and the corresponding electrostatic free energy following the expressions derived by Theodoor and Overbeek,⁵⁸ and adapted by Krishnan.⁵⁹ Since no assumptions were made for a particular geometry (e.g. perfect planes), these equations can be applied to our complicated surface with defects. The detailed derivation can be found in the above literature and also in the ESI.† Only the most important equations are presented here. Solving the Poisson–Boltzmann equation (eqn (3)) yields how the potential varies in the electrical double layer,

$$\nabla^2 \varphi = -\frac{\rho}{\epsilon_0 \epsilon_r} \quad (3)$$

where ρ is the charge density (C m⁻³), $\epsilon_0 \epsilon_r$ is the permittivity of the medium (F m⁻¹). For a symmetric monovalent electrolyte (e.g. NaCl), the expression simplifies to eqn (4).

$$\nabla^2 \psi = \kappa^2 \sinh(\psi) \quad (4)$$

κ^{-1} is the characteristic Debye length of an electrical double layer (eqn (5)),

$$\frac{1}{\kappa} = \sqrt{\frac{\epsilon_0 \epsilon_r kT}{2e^2 I}} \quad (5)$$

where e is the elementary charge and I is the ionic strength (number m⁻³).

In this work, the non-linear Poisson–Boltzmann equation is solved numerically using the COMSOL Multiphysics® software via the finite element method,⁶⁰ which yields a potential map in three dimensions. An example of that solution is visualized in Fig. 1, where the potential (φ) is indicated by different colors (e.g. navy blue when $\varphi = -51.4$ mV and light green when $\varphi = 0$ mV). A boundary condition of constant potential ($\varphi = \varphi_0$) is applied to the surfaces of the particle, the normal surface, and the defects, while the continuity condition ($n \cdot \nabla \varphi = 0$) is specified on the boundary of the model.

The computed spatial distribution of potential is subsequently used to calculate the electrostatic free energy (F_{el}), which is the sum of 3 energy components (eqn (6)).

$$F_{\text{el}} = U_{\text{el}} - T\Delta S + F_{\text{chem}} \quad (6)$$

The first term is the electrostatic potential energy (U_{el}) (eqn (7)), which describes the energy of surface and bulk charges in the local potential field.

$$U_{\text{el}} = \frac{1}{2} \int_A \sigma \varphi_0 dA - c_0 kT \int_V \psi \sinh(\psi) dV \quad (7)$$

The second component is the configurational entropy (ΔS) (eqn (8)), resulting from the uneven distribution of ions when mixing a dilute solution of charges.

$$\Delta S = 2c_0 k \int_V [-\psi \sinh(\psi) + \cosh(\psi) - 1] dV \quad (8)$$

The last contribution is the chemical free energy (F_{chem}) (eqn (9)), accounting for the change in free energy (chemical, non-electrical) due to the preferential adsorption of ions at the surface.

$$F_{\text{chem}} = - \int_A \sigma \varphi_0 dA \quad (9)$$

c_0 , σ , and φ_0 are the bulk salt concentration (number m⁻³), surface charge density (C m⁻²), and surface potential (V) respectively.

Substituting eqn (7)–(9) back to F_{el} gives eqn (10).

$$F_{\text{el}} = -\frac{1}{2} \int_A \sigma \varphi_0 dA + c_0 kT \int_V \psi \sinh(\psi) dV - 2c_0 kT \int_V [\cosh(\psi) - 1] dV \quad (10)$$

Importantly, F_{el} consists of a surface integral and two volume integrals. The surface integral should be performed on all the surfaces. For our model, this would be the particle, the defects, and the normal surface. On the other hand, the volume integrals should be evaluated for the bulk electrolyte.

2.3. van der Waals energy (E_{vdw})

In addition to the electrostatic free energy F_{el} , the attractive van der Waals contributions are also considered. The system of Au and TiO₂ is geometrically modeled by a sphere interacting with a plane. For such geometry, the attractive van der Waals energy is given by

$$E_{\text{vdw}} = -\frac{A_{\text{H}} R}{6D} \quad (11)$$

where A_{H} , R , and D are the Hamaker constant, particle radius, and particle–surface separation respectively.^{1,61}

While the A_{H} of TiO₂ in water is fairly consistent in the literature (~50 zJ),^{62,63} that of Au varies considerably depending on the study (~100–300 zJ).^{64,65} For the sake of discussion, we initially limit ourselves to the mean value of 200 zJ as the A_{H} of Au, but it is allowed to vary later when fitting the data with a support vector machine (0–150 zJ for Au/TiO₂). Accordingly, the Hamaker constant describing the interaction between TiO₂ and Au can be estimated by taking the geometric mean of the individual A_{H} , i.e. $\sqrt{(A_{\text{H,TiO}_2})(A_{\text{H,Au}})} = \sqrt{(50)(200)} = 100\text{zJ}$. However, it is noted that changing the overall A_{H} slightly does



not affect the qualitative outcome, as indicated by the distributions of energy barriers with varying A_H in Fig. S7.†

2.4. Total interaction energy (E_{tot}) and energy barrier

The total interaction energy (E_{tot}) is given by the sum of the electrostatic free energy (F_{el}) and the van der Waals energy (E_{vdw}) (eqn (12)).

$$E_{\text{tot}} = F_{\text{el}} + E_{\text{vdw}} \quad (12)$$

For a particular set of parameters (V_{Part} , V_{Surf} , V_{Def} , DD , $Conc$, R), E_{tot} is evaluated at 15 distances ($0.1\text{--}4.1 \text{ } \kappa^{-1}$) between the particle and the surface to examine how the energy changes upon interaction. A representative example of how each individual energy component constitutes the total energy is illustrated in Fig. 2 (when $V_{\text{Part}} = -51.4 \text{ mV}$, $V_{\text{Surf}} = -51.4 \text{ mV}$, $V_{\text{Def}} = +12.8 \text{ mV}$, $DD = 0.2$, $Conc = 1 \text{ mM}$, $R = 2.5 \text{ nm}$). The whole dataset can be explored in our GitHub repository.

Since we are only interested in the relative change in energy when the particle moves towards the surface, the energy when the particle and surface are significantly apart ($>4 \text{ } \kappa^{-1}$) and are essentially non-interacting is defined as zero. The energy curves for defects and the normal surface in Fig. 2 are fitted to an exponential expression for better analysis: $E = a \exp(-bD)$, where a and b are the parameters fitted for each curve, with an average residual standard deviation of 0.14 and 0.16 for the fitted defective and surface curves respectively. Importantly, the resultant E_{tot} is in the magnitude of thermal energy (kT , $T = 298 \text{ K}$), and the maximum energy of E_{tot} can be extracted as the energy barrier for interaction. This operation is repeated for all the 6000 simulated conditions to acquire the corresponding energy barrier. For the particular example in Fig. 2,

the barrier is larger than kT . Further representative cases of how F_{el} and E_{vdw} affect the final E_{tot} are shown in Fig. S8.† It is also shown that the 15 steps in particle–surface separations are sufficient to capture the energy barriers (details in ESI, Fig. S9†).

3. Results and discussion

As depicted in Fig. 1, we consider the interaction between a negatively charged particle (Au) and a negatively charged plane with the presence of positive/neutral defects (TiO_2). This configuration represents the scenario at mildly alkaline pH (e.g. 8), which is above the IEP of both materials, and therefore both Au and TiO_2 are overall negatively charged. However, this moderately alkaline pH is only slightly above the IEP of TiO_2 , so not all surface hydroxyls on TiO_2 are deprotonated.^{30,31,35,51} Hence, the remaining protonated species could serve as local positive/neutral charges, and we refer to these local charge deviations as defects. In our defect-containing model, we systematically varied a total of six parameters (V_{Part} , V_{Surf} , V_{Def} , DD , $Conc$, R) in 6000 different combinations (Table 1). For each of these 6000 conditions, the total interaction energy is calculated by considering both the electrostatic free energy and van der Waals energy, in a fashion similar to the classical DLVO theory (Fig. 2, details in the Methods section). The resultant energy barriers are extracted from each case, and the distributions of these 6000 energy barriers are plotted by each parameter in Fig. 3(a).

By examining how the energy barriers change in response to varying parameters, it is possible to evaluate the importance of V_{Part} , V_{Surf} , V_{Def} , DD , $Conc$, R . For instance, if a factor is

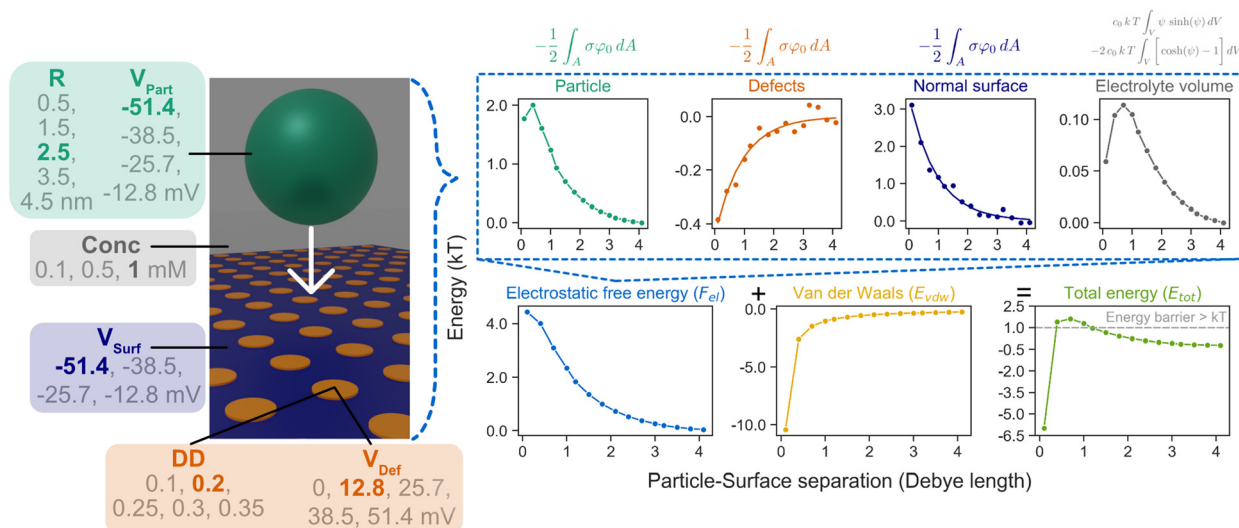


Fig. 2 The six parameters (V_{Part} , V_{Surf} , V_{Def} , DD , $Conc$, R) determine the electrostatic free energy (F_{el}) and the van der Waals energy (E_{vdw}), which are added up to give the total interaction energy (E_{tot}). A representative example is shown here when $V_{\text{Part}} = -51.4 \text{ mV}$, $V_{\text{Surf}} = -51.4 \text{ mV}$, $V_{\text{Def}} = +12.8 \text{ mV}$, $DD = 0.2$, $Conc = 1 \text{ mM}$, $R = 2.5 \text{ nm}$. The maximum E_{tot} is extracted as the energy barrier for each set of conditions. Further examples of how F_{el} and E_{vdw} affect the final E_{tot} are included in Fig. S8.† It is also shown that the 15 steps in particle–surface separations are sufficient to capture the energy barriers (details in section 5 of ESI†).



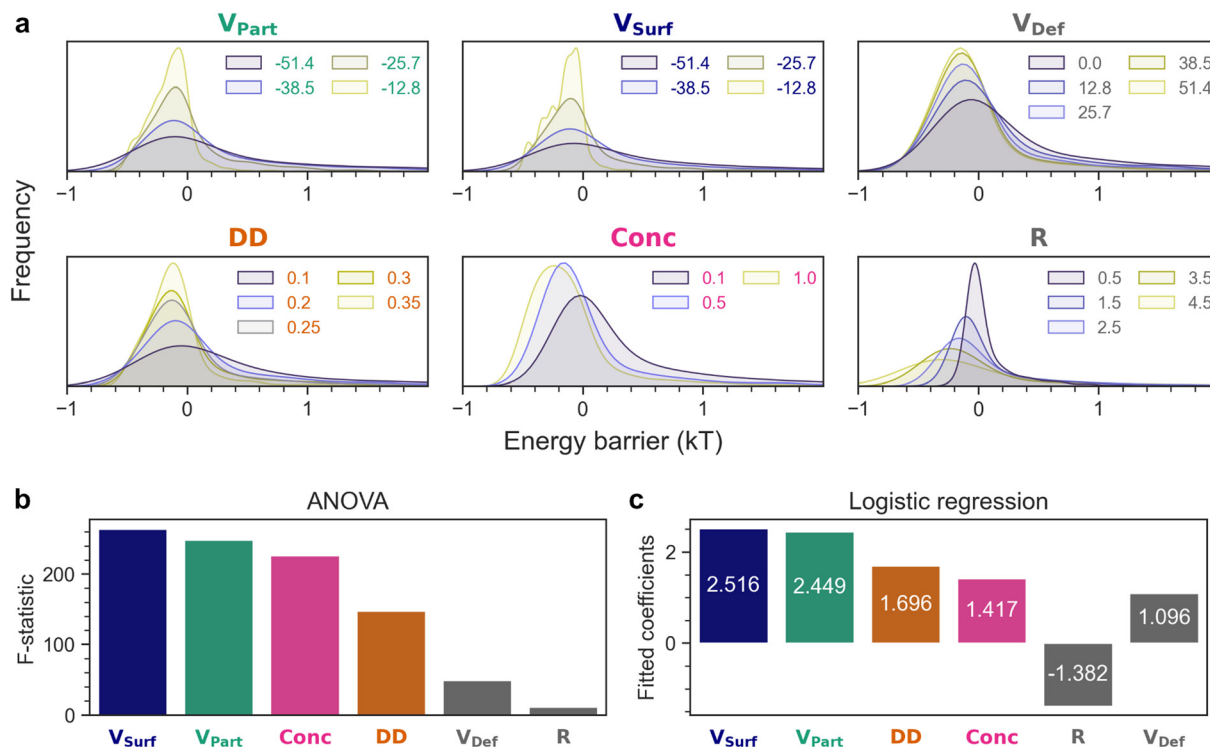


Fig. 3 (a) Grouping the same set of 6000 energy barriers by different parameters to examine the effect of each factor. Feature importance determined by (b) ANOVA and (c) logistic regression.

crucial for the interaction, changing its value should also shift the energy barrier significantly. By contrast, if a parameter has minimal effect on the outcome, the energy barrier should remain essentially the same. From Fig. 3(a), it is apparent that when modifying the value of each parameter, the distributions of energy barriers are all shifted to lower or higher kT values substantially, suggesting that all six factors contribute to the resultant energy barrier, though to a different extent. The effect of each parameter can be examined by varying it, and inspecting how the energy barriers change correspondingly:

- V_{Part} , V_{Surf} , V_{Def} : For V_{Part} and V_{Surf} , traversing the surface potential from -12.8 mV to -51.4 mV results in a broader distribution of energy barriers and more values above kT , since there is more electrostatic repulsion when the particle and the surface are more negatively charged. On the contrary, when the defects become more positively charged (V_{Def} from 0 mV to $+51.4$ mV), the energy barriers are reduced due to more electrostatic attraction.

- DD : Higher defect density (0.1–0.35) lowers the interaction energy barrier due to the neutral/positively charged defects.

- $Conc$: Higher salt concentrations (and ionic strengths) (0.1–1 mM) mean shorter Debye lengths and less repulsion between the surfaces due to the screening of charges.

- R : Larger particles (0.5–4.5 nm) lead to more electrostatic repulsion but also more van der Waals attraction (eqn (11)). The van der Waals contribution outweighs the electrostatic component in this case.

Although the behavior of each parameter follows what one might expect qualitatively, it would be desirable to quantify which factors are more decisive than others in determining the final interaction energy barrier, *i.e.* the feature importance. This is estimated by two separate methods here: ANOVA (analysis of variance) and logistic regression.

ANOVA compares the variances of different distributions and calculates the F -statistic to evaluate how likely the distributions are from the same population. In other words, a large F -statistic means the distributions do not resemble each other. Following this idea, if a parameter is more important, it should also shift the distributions of energy barriers more significantly, leading to more heterogeneous distributions and a larger value of F -statistic. By performing ANOVA on the distributions of energy barriers in Fig. 3(a), the F -statistics for each parameter can be obtained and are plotted in Fig. 3(b) in descending order: $V_{Surf} > V_{Part} > Conc > DD > V_{Def} > R$. We shall comment on the feature importance estimated by ANOVA in greater detail after the discussion of logistic regression, but it is immediately obvious that V_{Surf} (264) and V_{Part} (248) are the most influential factors as manifested by the largest F -statistics, which supports the experimental intuition to use zeta potential as a first hint to infer colloidal stability. In addition, $Conc$ (226) is also a decisive determinant for colloidal interactions followed by DD (147). On the contrary, V_{Def} (50) and R (11) do not seem to affect the outcome significantly. Instead of comparing several distributions at once with ANOVA, a pairwise comparison of the distributions was also



carried out by the Tukey's HSD (honestly significant difference) test (Fig. S10†), where the q -statistics suggest the effect of each parameter is even stronger when it is varied towards its extreme values (minimum or maximum).

While it is useful to calculate the exact interaction energy barrier, it is only meaningful when compared with the thermal energy (kT) available, which ultimately determines whether the interaction is favorable or not at a particular temperature. In other words, if the barrier is larger than kT , the interaction is considered “unfavorable”, and *vice versa*. However, we also note that colloids should exhibit a Boltzmann distribution of energies, and therefore simply comparing the energy barrier with kT might not be the most comprehensive metric, especially when they are of similar magnitude. Yet, by splitting the 6000 computed barriers into either “below kT ” or “above kT ” cases, this interaction problem essentially becomes a simple binary classification task, which is well-modeled by logistic regression. Similar to multivariate linear regression, logistic regression adds up all the variables (x_i) with appropriate coefficients (β_i) together with an intercept term (β_0) to give y (eqn (13)). However, there is an additional step of feeding y into a sigmoid function to introduce non-linearity (eqn (14)). The resultant z takes a value between 0 and 1, and is therefore suitable for predicting the probability when there are only two distinct scenarios.

$$y = \beta_0 + \sum_{i=1}^6 \beta_i x_i \quad (13)$$

$$z = \frac{1}{1 + \exp(-y)} \quad (14)$$

Putting this into the context of colloidal interactions, the calculated z is essentially the probability of “below kT ” (and $(1 - z)$ is the probability of “above kT ”). The goal of logistic regression is to take the six parameters (V_{Part} , V_{Surf} , V_{Def} , DD , $Conc$, R) as inputs, and predict whether it is more likely to be “below kT ” or “above kT ”. This prediction can be optimized by tuning the coefficients (β_i) of the factors. The fitted coefficients are shown in Fig. 3(c) and eqn (15), where the magnitude represents the relative importance of each parameter. To prevent large absolute numbers from potentially dominating the results, the parameters are scaled by z -score normalization prior to fitting ($\|x_i\| = \frac{x_i - \mu_i}{\sigma_i}$, μ_i and σ_i are the mean and standard deviation of the distribution). The extra $\| \|$ symbol is used to denote a normalized quantity. Further details of logistic regression can be found in the ESI.†

$$y = 6.572 + 2.516\|V_{\text{Surf}}\| + 2.449\|V_{\text{Part}}\| + 1.696\|DD\| + 1.417\|Conc\| - 1.382\|R\| + 1.096\|V_{\text{Def}}\| \quad (15)$$

The feature importance informed by logistic regression is as follows: $V_{\text{Surf}} > V_{\text{Part}} > DD > Conc > R > V_{\text{Def}}$. Given that ANOVA analyzes the absolute energy barriers, while logistic regression collapses the whole interaction problem into a binary decision, it is not surprising that the feature impor-

tance estimated by both methods does not match exactly. Interestingly, although the exact order differs, ANOVA and logistic regression share the top four factors (V_{Surf} , V_{Part} , DD , $Conc$). V_{Surf} and V_{Part} are among the most influential descriptors, which agree well with the common experimental observation that the zeta potential of colloids is a reasonably good metric for predicting their stability.^{6–10} The salt concentration ($Conc$) is also an important determinant, which is in accordance with the classical double layer theory that ions screen the charges in the solution, and therefore a higher salt concentration (and ionic strength) leads to a shorter Debye length and lower stability of colloids. However, when defects are also taken into account, it appears that the exact potential of defects (V_{Def}) does not play a pivotal role, but the defect density (DD) is much more crucial to the fate of the interaction. Nevertheless, it is noted that the interaction in question is ultimately a high-dimensional problem, and therefore no single factor can predict the outcome confidently, but a combination of parameters should be considered.

Since the top four factors (V_{Surf} , V_{Part} , DD , $Conc$) should be responsible for the majority of the results, it is therefore of interest to examine whether it is already sufficient to only consider these parameters, *i.e.* leaving out the two remaining ones (R , V_{Def}). A common engineering approach is to handcraft certain dimensionless numbers out of the parameters in consideration, and visualize the data on a 2D plot to inspect whether the data points cluster. Several dimensionless numbers (*e.g.* $\frac{(V_{\text{Surf}} + V_{\text{Part}})}{V_{\text{Def}}}$ and $\frac{\kappa^{-1}}{R}$) were constructed in an attempt to separate the data points into regions of “below kT ” and “above kT ” (Fig. S11†). However, none of them yields satisfactory clustering. Since the formulation of dimensionless numbers is confined by the intrinsic units of V_{Surf} , V_{Part} , DD , and $Conc$, these factors can only be assembled in a limited number of ways. To solve this problem, we exploit z -score normalization as a general way to combine parameters even though they do not necessarily have complementary units. Z -score normalization scales the inputs with respect to their standard deviations which have the same units, essentially rendering them dimensionless on their own ($\|x_i\| = \frac{x_i - \mu_i}{\sigma_i}$, where μ_i and σ_i are the mean and standard deviation of the distribution respectively). The mean and standard deviation for each variable are summarized in Table S3.† Importantly, with these dimensionless, normalized quantities, they can be simply added up together to construct dimensionless numbers, even though they might not have compatible units intrinsically.

However, to form a meaningful dimensionless number, the normalized variables must be combined linearly with appropriate weights. A natural choice of such weights is the coefficients previously fitted by logistic regression (β_i , summarized in (eqn 15) and Fig. 3(c)), since the logistic regression coefficients were obtained by processing normalized parameters in the first place. More specifically, we can add up the surface potential V_{Surf} and the particle potential V_{Part} with the logistic regression coefficients to construct the “Electrostatics”



metric ($2.516\|V_{\text{Surf}}\| + 2.449\|V_{\text{Part}}\|$). Similarly, the defect density DD and the salt concentration $Conc$ can be considered together to formulate the “Environment” descriptor ($1.696\|DD\| + 1.417\|Conc\|$). When these two descriptors (“Electrostatics” and “Environment”) are used as the selection criteria and plotted as the x and y axes in Fig. 4, the “below kT ” and “above kT ” cases are successfully clustered into distinctive regions as manifested by the separation of colors (green for “below kT ” and red for “above kT ”), which cannot be achieved if only individual parameters (V_{Surf} , V_{Part} , DD , $Conc$) are considered. The color of the data points represents the likelihood of “below kT ”, which is evaluated by considering the overlapping points ($N = 25$ in Fig. 4) in this low-dimensional space, inevitably existing because only 4 out of 6 parameters are included here. Each point carries a label of either “below kT ” (1) or “above kT ” (0), and by taking the average of the labels at the overlapping points, the probability of “below kT ” can be estimated. Notably in Fig. 4, the decision boundary between “above kT ” and “below kT ” is vividly visualized by the diagonal transition from red to yellow and eventually to green, indicating an increasing probability of “below kT ” from the bottom left to the top right of the graph. The diagonal nature of this shift demonstrates that both the x and y axes are vital to the fate of colloidal interactions. If the results are mainly determined by the x -axis (y -axis), the boundary would have been a vertical (horizontal) one. Given that only $\sim 9\%$ of the data points (554 out of 6000) are “above kT ”, this graphical method is proven to be an effective way to separate the “below kT ” and “above kT ” cases by considering the top four factors together. Overall, a prediction for the interaction can be made by simply inputting the normalized numbers of $\|V_{\text{Surf}}\|$, $\|V_{\text{Part}}\|$, $\|DD\|$, $\|Conc\|$, and reading the color off the diagram: green for “below kT ”, red for “above kT ”, and yellow for the uncertain cases.

Although the graphical method in Fig. 4 manages to capture the transition from “above kT ” to “below kT ” by taking the four most important factors into account, the intermediate cases with higher uncertainty can only be decided by including the remaining parameters. Moreover, throughout the previous discussion, a constant Hamaker constant (A_H) of 100 zJ is used to represent the average van der Waals attraction between TiO_2 and Au, which might slightly deviate from it depending on the actual Au system. Consequently, we expanded the dataset by allowing the Hamaker constant to vary (0–150 zJ, steps of 10 zJ). Further simulations of perfect (uniformly charged) surfaces without defects ($DD = 0$ and $V_{\text{Def}} = 0$) are also added to the dataset for comparative purposes. As a result, the extended dataset now encompasses a total of 116 160 conditions of seven parameters including the Hamaker constant (details in ESI†). To interact with this extensive dataset effectively, we fitted it to a support vector machine (SVM). The trained SVM model takes the seven factors (V_{Part} , V_{Surf} , V_{Def} , DD , $Conc$, R , A_H) as inputs, and returns “below kT ” (spontaneous adsorption expected) or “above kT ” (no adsorption expected) as the output, together with a decision function indicating the confidence of the prediction (a larger absolute value means higher confidence). An accuracy of 97% is achieved with a standard train/validation split of 80%/20%. Importantly, since our parameter space covers a large landscape of realistic conditions, and the SVM is trained based on these simulation data, it is also possible to predict the interaction between colloids for unseen but physically feasible cases, for perfect and defective surfaces. To facilitate the usability of the trained SVM, we prepared a Google Colab notebook with a simple graphical user interface, where users can enter a specific set of the seven parameters, and obtain a prediction for whether the interaction is favorable or not. This is made available in our GitHub repository.

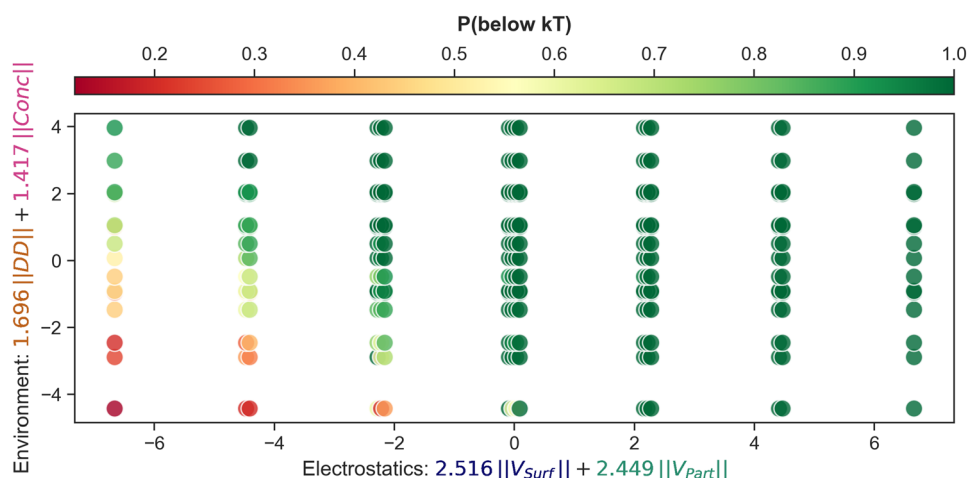


Fig. 4 Decision boundary between “below kT ” (green) and “above kT ” (red) as predicted by logistic regression, visualized on a 2D plot. The yellow/orange points indicate uncertain cases at the boundary. The “Electrostatics” ($2.516\|V_{\text{Surf}}\| + 2.449\|V_{\text{Part}}\|$) and “Environment” ($1.696\|DD\| + 1.417\|Conc\|$) descriptors are constructed by combining the top four factors (V_{Surf} , V_{Part} , DD , $Conc$) using the weights fitted by logistic regression. For each dot on this low-dimensional plot, there are 25 overlapping points ($N = 25$) which can only be differentiated if the remaining two parameters (R and V_{Def}) are also considered. P (below kT) is estimated by the fraction of “below kT ” at the overlapping points.



While it might be a neat approach to treat the colloidal interaction as a simple problem of binary classification (“below kT ” or “above kT ”), it is worth noting the limitations of our model. First, the simulations are based on the Poisson–Boltzmann equation (eqn (3)), which means that our model also suffers from the inadequacies inherited from the theory. For instance, it is only a mean-field method with no ion-ion correlations considered.⁶⁶ In addition, the Poisson–Boltzmann theory also tends to fail for highly charged regimes and multi-valent electrolytes.⁶⁷ Consequently, a symmetric monovalent salt (e.g. NaCl) has been deliberately assumed throughout this work, and hence the terms “salt concentration” and “ionic strength” are used interchangeably in the text. For simplicity, the Stern layer is also neglected, but it might be included in a similar way as described by Biesheuvel.⁶⁸ Apart from the shortcomings of the Poisson–Boltzmann theory, we have only considered the effect of seven parameters (V_{Part} , V_{Surf} , V_{Def} , DD , $Conc$, R , A_{H}), and other factors which might potentially influence the outcome are not captured in our model, such as the morphology beyond a spherical particle. Despite these limitations, we employed this parametrized model based on the Poisson–Boltzmann theory due to its efficiency, where a simulation only took roughly ten seconds to run on average (details in ESI†). This enables a parametric sweep across a wide range of physically interesting conditions, which cannot be easily accomplished by more sophisticated yet computationally expensive methods (e.g. molecular dynamics and Monte Carlo).

Furthermore, it has to be noted that surfactants are not considered in our model, which are often used to stabilize the metal colloids against aggregation. Although our model can possibly account for the electrostatic contribution from the surfactants, the additional steric component of the surfactants might still prevent adsorption from happening. Therefore, to encourage the favorable interaction between metal and oxide particles, surfactant-free colloids should be used. Common approaches to remove surfactants include thermal treatment and repeated centrifugation with a mixture of good and poor organic solvents.^{69,70} However, it is not always trivial to eliminate the surfactants without changing the structure and keeping the colloids stable. An alternative method is to synthesize particles without the use of surfactants from the beginning. This can be done by techniques such as laser ablation in liquid, where a bulk metal target (e.g. Au) is irradiated by a pulsed laser to create nanoparticles in water which are only stabilized electrostatically.⁷¹ When comparing our model to experiments, it is expected that these surfactant-free particles would follow the predicted adsorption behavior better than the surfactant-stabilized counterparts.

Until now, we have limited ourselves to Au/TiO₂ systems. However, since the factors (e.g. V_{Surf} , DD , A_{H}) are all empirical, it is expected that our model should largely generalize to similar metal/oxide materials, which could help understand the effect of the local environment on colloidal interactions. Experimentally, the zeta potential has always been the indi-

cator of colloidal stability due to its ease of measurement. Yet, the zeta potential is only a weighted average of the potentials at the shear/slipping plane of all the particles in the dispersion.¹³ This global metric might not fully represent the local surface potential, which ultimately determines the adsorption behavior. As an example, it has been shown by AFM that the major (110) and minor (100) facets of rutile TiO₂ have disparate IEPs separated by about two units, i.e. (110): 5.5–4.8 and (100): 3.7–3.2.¹⁵ This means that the surface (and zeta) potentials of these two facets are not the same at a particular pH, and are expected to interact differently with colloids. This facet-dependent interaction was demonstrated evidently by Su *et al.*¹⁶ when adsorbing silica nanoparticles on SrTiO₃ supports. With dual scale AFM, the authors established that the IEPs of major (110) and minor (100) on SrTiO₃ are ~4 and ~6 respectively. At a pH below their IEPs (3.5), negatively charged silica colloids were adsorbed on both facets which were positively charged. At a pH above their IEPs (10.8), no adsorption happened since all the silica, SrTiO₃ (110) and (100) were negatively charged. Interestingly, in the intermediate pH range (4–6) between the IEP of (110) and (100), negatively charged silica colloids selectively adsorbed on the positively charged (100), but not the negatively charged (110). The bulk IEP value of SrTiO₃ (~3.5) would not have predicted this facet anisotropy. In our model, different facets can be potentially described by two separate cases: (110) with a lower IEP and V_{Surf} , and (100) with a higher IEP and V_{Surf} . Instead of only considering the case when (110) is negatively charged and (100) is positively charged, it would also be possible to predict the adsorption when both facets are slightly negatively charged.

Moreover, even within the same facet, the local surface potential does not necessarily have to be uniform across the entire surface due to the presence of defects, such as surface hydroxyls having a different charge than the rest of the surface. Using an OH-functionalized AFM tip, Wagner *et al.*⁷² directly probed the proton affinity of surface oxygen atoms on In₂O₃ (111), where the proton affinity of surface oxygen atoms is strongly correlated with the acidity ($\text{p}K_{\text{a}}$) of the resultant surface hydroxyls. It was found that the four types of surface hydroxyls on In₂O₃ (111) had varied degrees of acidity. This means that they are likely to be protonated/deprotonated to a different extent at the same pH. In our model, this local environment of surface hydroxyls at different pH values could be captured by modifying the defect density (DD) and defect potential (V_{Def}), and predictions can be made on whether a particular colloidal interaction is favorable. Overall, although the zeta potential is usually a reasonable initial estimate for colloidal interactions, it is the local environment which ultimately dictates the interaction outcome. We hope that our model will encourage further studies on examining the effect of facets and surface defects with well-defined methodology (e.g. quartz crystal microbalance and single crystals), as well as achieving the desired adsorption behavior by controlling the surface chemistry.



4. Conclusion

Colloidal interaction is ultimately a high-dimensional problem. Conventionally, the zeta potential is the most frequently used metric for predicting the stability of colloids in line with the established DLVO theory. The salt concentration and ionic strength are also known to affect the interaction following the Debye length relationship and the Schulze–Hardy rule. To gain more insights beyond the current understanding, it is important to consider all these well-studied factors together with other contributions which are relatively less explored, such as the effect of local defects.

Herein, we simulate the interaction between negatively charged Au and TiO₂ particles in the electrostatically repulsive regime based on the Poisson–Boltzmann equation and the finite element method, while varying a total of six parameters: V_{Part} (particle potential), V_{Surf} (surface potential), V_{Def} (defect potential), DD (defect density), $Conc$ (salt concentration), and R (particle radius). Defects were treated as local charge deviations compared to the regular surface. A possible identity of such defects is the surface hydroxyls on TiO₂, which are less negative than the rest of the surface when protonated. The van der Waals contribution was included by using a mean value of Hamaker constant (A_{H}) for Au/TiO₂. A wide range of experimentally relevant conditions is covered by simulating 6000 unique combinations of these six parameters, and the interaction energy barrier for each case is extracted. By analyzing the distribution of interaction energies with ANOVA and logistic regression, the relative importance of each factor is inferred: $V_{\text{Surf}} > V_{\text{Part}} > DD > Conc > R > V_{\text{Def}}$. The results demonstrate that the top four factors are V_{Surf} , V_{Part} , DD , and $Conc$, which agrees well with the conventional wisdom that the surface (and zeta) potential and the salt concentration (ionic strength) play a decisive role in the stability of colloids. Interestingly, when defects are included, it appears that the energy barrier can also be significantly lowered by a higher defect density (DD), while the defect potential (V_{Def}) seems to have a minimal impact on the interaction outcome.

To examine whether it is sufficient to consider only the top four factors, the interaction can be framed as a classification problem: for Au colloids to adsorb on the TiO₂ support, the energy barrier has to be smaller than the thermal energy (kT). Accordingly, the interaction energy barriers from the simulations can be divided into two groups relative to kT : “below kT ” for favorable interactions and “above kT ” for unfavorable interactions. This essentially simplifies the interaction problem to a binary classification task, which is well-modeled by logistic regression. Using the weights from logistic regression, the top four factors are combined to construct the “Electrostatics” ($2.516\|V_{\text{Surf}}\|+2.449\|V_{\text{Part}}\|$) and “Environment” ($1.696\|DD\|+1.417\|Conc\|$) descriptors, which successfully cluster the data points into regions of unfavorable (red), uncertain (yellow), and favorable (green) cases. However, even when combining the top four parameters, certain scenarios remain uncertain (yellow), and can only be confidently classified by taking the remaining parameters (R and V_{Def}) into account. Hence, it is

crucial to consider all the six parameters including the particle radius R and the defect potential V_{Def} for an accurate prediction.

In the preceding discussion, the van der Waals Hamaker constant (A_{H}) was fixed at the mean value for Au/TiO₂ (100 zJ). However, to improve the generalizability of this model towards other materials, A_{H} was also allowed to vary across the entire range of 0–150 zJ, yielding an expanded dataset with 116160 conditions encompassing seven parameters and their corresponding interaction energies. To predict the fate of the interaction while considering all parameters, a support vector machine (SVM) is trained with the extended dataset, which determines a decision boundary between “below kT ” and “above kT ” cases with 97% accuracy using a standard train/validation split of 80%/20%. The trained SVM model takes a total of seven variables (V_{Part} , V_{Surf} , V_{Def} , DD , $Conc$, R , A_{H}) as the input, and returns “below kT ” or “above kT ” as the output, together with a decision function indicating the confidence of the prediction. It is also noted that the model can be used for both seen and unseen cases, perfect and defective surfaces. To facilitate the usability of the trained SVM, we prepared a Google Colab notebook with a simple graphical user interface for querying the prediction given a set of parameters, where users can easily change different parameters and examine the effect.

Although the simulations are modeled after Au/TiO₂ systems, the factors (e.g. V_{Surf} and $Conc$) are all empirical, and therefore the model should largely generalize to similar metal/oxide materials. We anticipate that our results and tools can encourage further experimental and theoretical investigations into the effect of local surface chemistry (e.g. facets and defects) on colloidal interactions, rather than only assessing the established bulk metrics (e.g. zeta potential).

Abbreviations

β_0	Logistic regression intercept (—)
β_i	Logistic regression coefficient (—)
n	Normal (—)
ΔS	Configurational entropy (J K ^{−1})
ϵ_0	Vacuum permittivity (F m ^{−1})
ϵ_r	Relative permittivity (—)
κ^{-1}	Debye length (m)
μ_i	Mean of distribution (—)
ψ	Dimensionless potential (—)
ρ	Charge density (C m ^{−3})
σ	Surface charge density (C m ^{−2})
σ_i	Standard deviation of distribution (—)
φ	Dimensional potential (V)
φ_0	Dimensional surface potential (V)
a	Lattice parameter of defect array (nm)
A_{H}	Hamaker constant (J)
c_0	Number concentration of salt (number m ^{−3})
$Conc$	Salt concentration (mM)
D	Particle–Surface separation (κ ^{−1})
DD	Defect density (—)
e	Elementary charge (C)



E_{tot}	Total interaction energy (J)
E_{vdw}	van der Waals energy (J)
F_{chem}	Chemical free energy (J)
F_{el}	Electrostatic free energy (J)
I	Ionic strength (number m^{-3})
k	Boltzmann constant (J K^{-1})
L_{D}	Defect radius (0.5 nm)
R	Particle radius (nm)
T	Temperature (K)
U_{el}	Electrostatic potential energy (J)
V_{Def}	Defect potential (mV)
V_{Part}	Particle potential (mV)
V_{Surf}	Surface potential (mV)

Data availability

The whole dataset and the SVM model are available at https://github.com/kinranlau/COMSOL_colloid_interaction.

Author contributions

All authors conceptualized the idea. K. L. ran the simulations and analyzed the data under the supervision of B. G. and S. R. K. L. drafted the manuscript. All authors wrote the final version of the manuscript. S. B. and S. R. acquired the funding for this work.

Conflicts of interest

There are no conflicts to declare.

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