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Enhanced photocatalytic water splitting with two-dimensional van der Waals heterostructures of BAs/WTeSe†

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The photocatalytic efficiency of monolayer materials can be significantly enhanced by constructing two-dimensional van der Waals heterostructures. This study presents first principles calculations based on density functional theory to investigate the electronic properties and photocatalytic mechanism of van der Waals heterostructures of boron arsenide (BAs) with the Janus MXY (M = W; X/Y = Se, Te) monolayers, with and without Se vacancies. Results from binding energies, phonon spectra, and *ab initio* molecular dynamics simulations indicate that the heterostructures are thermodynamically stable. The deformation potential theory was then used to investigate carrier mobility μ in 2D materials. All the heterostructures exhibit direct bandgaps with valence band maxima and conduction band minima suitable for water splitting. Additionally, these heterostructures possess high optical absorption coefficients in the visible and ultraviolet regions. In particular, our calculations predict BAs/WTeSe, with and without Se vacancies, as promising candidates for photocatalytic water splitting applications.

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Introduction

The finite nature of fossil fuels, coupled with their detrimental environmental impact and contribution to climate change, has exacerbated environmental contamination and energy crises, posing significant challenges to humanity.¹ Finding a sustainable energy source to replace finite fossil fuels is crucial. Hydrogen generated by photocatalytic water splitting is regarded as a sustainable, clean, and renewable energy source for the future.^{2–4} However, finding the right photocatalyst is essential to producing hydrogen energy from water splitting. The large band gap and improper alignment of band edge position are the main hurdles for a suitable photocatalyst for water splitting.^{5–7}

Graphene, a two-dimensional (2D) material, has been extensively studied for its applications in electronic and optoelectronic devices, even though it lacks a tunable band gap.⁸

Consequently, new types of 2D materials have been explored in recent years, including phosphorene,⁹ transition-metal dichalcogenides (TMDs),¹⁰ and MXenes.¹¹ Compared with the bulk counterparts, TMDs have enhanced charge transfer and separation capabilities due to their large specific surface area and are thus promising materials for application in optoelectronic devices, transistors and sensors.¹⁰ Among TMDs, WX₂ (X = S, Se, Te)¹² monolayers have been widely studied theoretically and experimentally. WS₂, WSe₂ and MoSe₂ monolayers are excellent high-temperature thermoelectric materials.

A recent breakthrough in the realm of 2D TMDs has led to the synthesis of Janus MXY (M = W, Mo; X/Y = Se, S, Te) monolayers through vulcanization or selenization experiments employing the CVD method.¹³ Unlike their mirror-symmetric counterparts MoS₂ and MoSe₂, the MoSSe monolayer is fabricated *via* selective selenization of the top atomic layer in a single-layer MoS₂ structure. The electronegativity difference between sulfur and selenium atoms induces interlayer polarization in Janus materials, enabling coupling with the inherent interlayer built-in polarization field. This unique property introduces an additional degree of freedom for modulating their physicochemical properties, garnering significant attention due to their novel features and promising applications. Moreover, Janus-WSeTe or Janus-WTeSe monolayers are direct band gap semiconductors with moderate band gaps.¹⁴ Unfortunately, not all 2D materials are suitable for water splitting

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because their band gap values may be too large and the band edge alignment may be located far from the redox potentials.

To address these limitations, the synthesis of van der Waals heterostructures (vdWHs) have been proposed as a promising strategy to engineer suitable band edge positions for water splitting.^{15–21} Arsenide semiconductors, such as boron arsenide (BAs),²² aluminum arsenide (AlAs),²³ gallium arsenide (GaAs),^{5,24} and indium arsenide (InAs),²⁵ have emerged as promising candidates for vdWHs due to their favorable band alignments and light absorption properties. These materials belong to groups III–V of compound semiconductor materials^{26–29} and play a vital role in making microwave devices,^{30,31} Hall devices,³² high-speed digital circuits,³³ lasers,^{34,35} magneto-resistive devices,^{7,35} and detectors.³⁶ Hexagonal BAs monolayer with a flat honeycomb structure like graphene has drawn attention as a potential photocatalyst for water splitting.³⁷ BAs has a direct band gap, high carrier mobility, and high in-plane mechanical stability and it is the only semiconductor material with an ultrahigh thermal conductivity of 2000 W m^{−1} K^{−1} at room temperature. Consequently, BAs monolayer has been recently combined with many other 2D materials such as the vdWHs BAs/MoTe₂,⁶ BAs/MoSe₂,³⁸ BAs/SnC,³⁹ and BAs/GaN.⁴⁰

This paper investigates using density functional theory (DFT) calculations the photocatalytic properties of BAs/Janus WSeTe and BAs/Janus WTeSe vdWHs as possible direct Z-scheme photocatalysts. Stacking-based approaches have been employed to investigate the electrical, optical and photocatalytic properties of all the suggested systems. We found that many configurations of these heterostructures are suitable for tuning optoelectronic characteristics for photocatalysis applications, either type-II or Z-scheme. Although the built-in electric field under various stacking configurations favours the development of Z-scheme charge transfer, not all of them belong to Z-scheme according to work function and charge density difference. In reality, the classification of the system as type-II or Z-scheme depends on the competition between charge separation and interlayer e[−]h⁺ recombination. Furthermore, by introducing Te vacancy defects, the charge transfer channel in Janus BAs/WTeSe heterostructures can be successfully flipped from type-II to Z-scheme, which may effectively regulate this competitive relationship. Remarkably, Te vacancy defects can increase the time interval between electron (hole) transmission. Our findings show that the BAs/Janus WTeSe vdWHs have strong photocatalytic activity and that the proposed class-II configuration is ideally suited for highly efficient water splitting.

Methods

All DFT calculations were conducted using the Vienna *ab initio* simulation package (VASP)⁴¹ and the projector augmented wave (PAW) potential method.⁴² Geometry optimization and *ab initio* molecular dynamics (AIMD) were performed using the Perdew–Burke–Ernzerhof (PBE) generalized gradient approximation (GGA)⁴³ in conjunction with Grimme's D3 (PBE-D3) method

to accurately describe the van der Waals interactions between BAs and Janus-WSeTe.⁴⁴ As generally PBE underestimates the band gap of semiconductors, single point calculations were conducted using the hybrid HSE06 (Heyd–Scuseria–Ernzerhof) functional. HSE06 was also used to compute the optical properties of the abovementioned heterostructures and monolayers.⁴⁵ Convergence tolerances were set to 10^{−6} eV for the energy and 0.01 eV Å^{−1} for the force. To ensure calculation accuracy, the *k*-point (Monkhorst–Pack) grid was set to (15 × 15 × 1) together with a cutoff energy of 520 eV for ensuring calculation accuracy. The vacuum space was set to 25 Å to minimise the interaction between adjacent layers. AIMD simulations were performed with a (4 × 4 × 1) supercell at the Γ point for 6 ps using a 1 fs timestep in the NVT ensemble (constant volume and temperature) at *T* = 300 K and *T* = 500 K.⁴⁶ The phonon dispersion curve calculations were also performed with a 4 × 4 × 1 supercell in the framework of frozen-phonon approximation.⁴⁷

Results and discussion

Structural properties

To construct the initial models of 2D BAs–Janus WSeTe vdWHs, the structures of the isolated hexagonal BAs and WSeTe monolayers were optimized first. The optimized structural parameters of BAs and WSeTe are summarized in Table 1. The lattice constants of BAs are *a* = *b* = 3.39 Å, and the B–As bond length is 1.96 Å. In the literature, the lattice constants of monolayer BAs obtained by DFT calculations have been reported as *a* = *b* = 3.39 Å and the bond length between B

Table 1 Structural and energetic parameters of the isolated hexagonal BAs and WSeTe monolayers and of the 2D BAs–Janus WSeTe vdW heterostructures: lattice parameter *a* (Å), band gap (*E_g*), formation energy Δ*E* (eV), and distance *d* (Å) between BAs and Janus-WSeTe monolayers

| System | <i>d</i> (Å) | | | <i>E_g</i> (eV) | | |
|------------------------------|--------------|-------|--------|---------------------------|-------|------------|
| | <i>a</i> (Å) | PBE | PBE-D3 | HSE | PBE | Δ <i>E</i> |
| BAs | 3.390 | — | — | 1.37 | 0.75 | — |
| WSeTe | 3.451 | — | — | 1.80 | 1.35 | — |
| BAs/Janus WSeTe | | | | | | |
| Stacking-A | 3.392 | 4.208 | 3.71 | 1.29 | 0.712 | −0.198 |
| Stacking-B | 3.392 | 4.214 | 3.52 | 1.43 | 0.796 | −0.225 |
| Stacking-C | 3.391 | 4.214 | 3.61 | 1.30 | 0.745 | −0.208 |
| Stacking-D | 3.392 | 4.216 | 3.44 | 1.18 | 0.682 | −0.252 |
| Stacking-E | 3.390 | 4.208 | 3.78 | 1.28 | 0.694 | −0.189 |
| Stacking-F | 3.391 | 4.172 | 3.39 | 1.28 | 0.709 | −0.251 |
| BAs/Janus WTeSe | | | | | | |
| Stacking-A | 3.390 | 4.470 | 3.99 | 0.86 | 0.374 | −0.195 |
| Stacking-B | 3.390 | 4.259 | 3.78 | 0.89 | 0.382 | −0.223 |
| Stacking-C | 3.391 | 4.448 | 3.85 | 0.82 | 0.309 | −0.203 |
| Stacking-D | 3.391 | 4.128 | 3.61 | 0.66 | 0.192 | −0.254 |
| Stacking-E | 3.391 | 4.462 | 4.03 | 0.92 | 0.466 | −0.186 |
| Stacking-F | 3.389 | 4.117 | 3.56 | 0.93 | 0.530 | −0.264 |
| BAs/Janus WSeTe (Se vacancy) | 3.375 | — | 3.51 | 1.33 | 0.72 | — |
| BAs/Janus WTeSe (Te vacancy) | 3.381 | — | 3.61 | 0.59 | 0.31 | — |



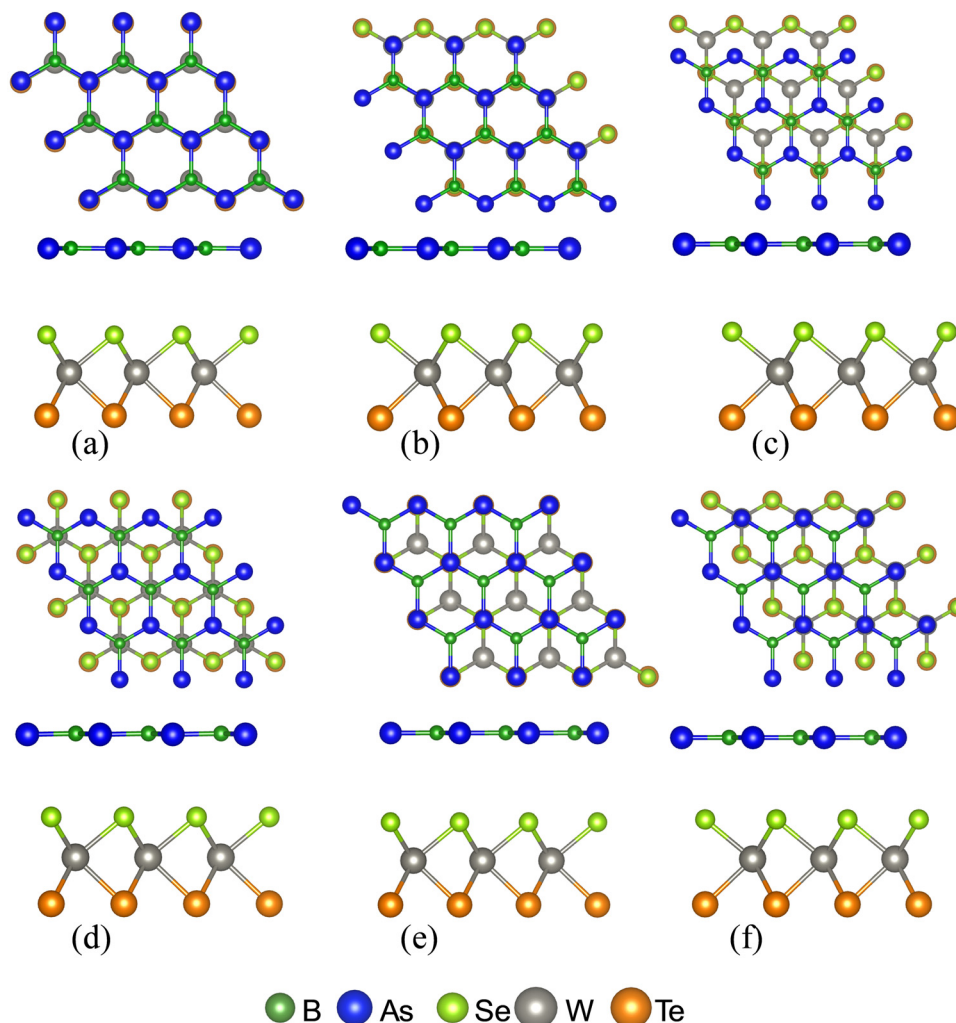


Fig. 1 Top and side views of different configurations of vdWHs obtained from the stacking of BAs and Janus-WSeTe: (a) stacking-A – the B and As atoms sit on the top of W and Se atoms, respectively; (b) stacking-B – the B and As atoms sit on the top of Se and W atoms, respectively; (c) stacking-C – the As atom is on the hollow site of the hexagonal site, and the B atom occupies the top of the Se atom; (d) stacking-D – the As atom is on the hollow site of the hexagonal site, and the B atom occupies the top of the W atom; (e) stacking-E – the As atom sits on the top of the Se atom, while the B atom is on the center of the hexagonal site; (f) stacking-F – the As atom sits on the top of the W atom, while the B atom is on the center of the hexagonal site, BAs/Janus WSeTe vdW heterostructure. The BAs/Janus WTeSe vdWHs are formed by swapping the positions of Se and Te atoms in the six configurations (a)–(f). The color scheme for each atom is given at the bottom.

and As atoms is 1.95 \AA .⁴⁸ The Janus WSeTe monolayer has lattice constants of $a = b = 3.45 \text{ \AA}$, with bond lengths of 2.55 \AA and 2.71 \AA between W–Se and W–Te atoms, respectively. Sun *et al.* had investigated this structure and found the lattice parameters of the hexagonal WSeTe to be $a = b = 3.405 \text{ \AA}$.^{49,50} The BAs with Janus-WSeTe and BAs with Janus-WTeSe monolayers were used to construct 12 vdW heterostructure configurations. They were categorized into two classes: (i) BAs/Janus WSeTe vdWHs, where the BAs monolayer is close to the Se atom of the Janus-WSeTe monolayer; (ii) BAs/Janus WTeSe vdWHs, where the BAs layer is close to the Te atoms of the Janus WTeSe monolayer. For these two classes of heterostructures, six different stacking configurations are possible, as shown in Fig. 1.

The computed values of lattice parameters and interlayer distances of the optimized BAs/WSeTe vdWHs are listed in

Table 1. Their formation energies (ΔE) were calculated according to the following equation:

$$\Delta E = E_{\text{heter}} - E_{\text{layer1}} - E_{\text{layer2}} \quad (1)$$

where E_{heter} , E_{layer1} , and E_{layer2} are the total energies of the heterostructure, BAs, and Janus WSeTe monolayer, respectively. The two most stable heterostructures according to the values of formation energy are the BAs/Janus WSeTe (stacking-D) vdWH in Fig. 1(d) and BAs/Janus WTeSe (stacking-F) vdWH in Fig. 1(f). The formation energies of these structures are smaller than those of the other structures by at least 10 meV. These structures (BAs/Janus WSeTe (stacking-D) and BAs/Janus WTeSe (stacking-F)) were considered to investigate the vibrational, electronic, photocatalytic (water splitting), and optical properties: (i) BAs/Janus WSeTe (H-I); (ii) BAs/Janus WTeSe (H-II) and



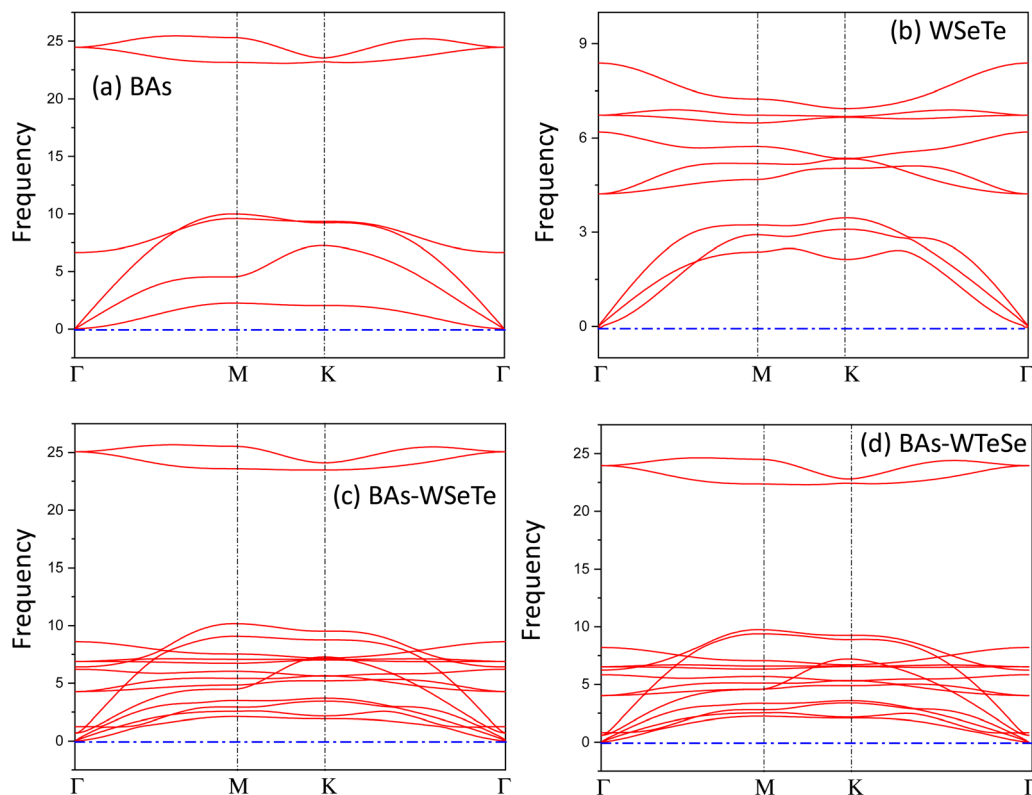


Fig. 2 The phonon spectra of the (a) BAs, (b) WSeTe, (c) BAs-WSeTe and (d) BAs-WTeSe heterostructures.

(iii) BAs/Janus WTeSe with Te vacancies (H-III) and (iv) BAs/Janus WTeSe with Se vacancies (H-IV). The phonon dispersions of BAs, WSeTe, and the H-I, H-II heterostructures in Fig. 2 confirm their structural stability: no imaginary frequencies are found in these phonon spectra. Furthermore, the H-III heterostructure exhibits dynamic stability, as corroborated by its phonon spectra presented in Fig. S1 of the ESI.[†]

The stability of the heterostructures was investigated using AIMD simulations. For each system, simulations were conducted at $T = 300$ K and 500 K. Fig. S2 (ESI[†]) confirms that the temperatures remained stable around the target values for BAs, WSeTe, BAs-WSeTe, and BAs-WTeSe. Representative configurations from the AIMD trajectories (Fig. S3–S6, ESI[†]) further illustrate the thermodynamic stability of these heterostructures at both temperatures. Notably, simulations without dispersion correction showed increased interlayer spacing between BAs and WSeTe (or BAs and WTeSe) monolayers, ultimately leading to separation. This highlights the crucial role of dispersion interactions for accurately modeling such systems when using DFT methods.

The computed lattice parameters of the monolayers WSeTe and BAs are 3.45 Å and 3.39 Å, and of the heterostructures H-I, H-II, H-III and H-IV are 3.39 Å, 3.38 Å, 3.38 Å and 3.37 Å, respectively. Due to the higher value of lattice mismatch, many defects are invariably created at the interface when two single-crystalline materials are formed or fused together. For BAs and Janus-WSeTe, the lattice mismatch is only 3.3%, which is small when compared to other heterostructures.^{51,52} Finally, in

Table 1 the inclusion of dispersion correction in the PBE functional has a significant influence on the stability of heterostructures.⁵³ The most stable heterostructures H-I (stacking-D) and H-II (stacking-F) have lower interlayer spacing, $d = 3.44$ Å and $d = 3.56$ Å, respectively.

Electronic properties

The band structure provides insights into the electronic properties of a material. The electronic properties of the BAs and Janus WSeTe monolayers are first discussed before considering those of the BAs/Janus WSeTe heterostructure. The band structures of Janus WSeTe and BAs, computed at the PBE and HSE06 levels of theory, are presented in Fig. 3. The band gap values for the Janus WSeTe and BAs monolayers calculated using the HSE06 hybrid functional are 1.37 and 1.8 eV, respectively. Both Janus-WSeTe and BAs are direct band gap semiconductors and their conduction band minimum (CBM) and valence band maximum (VBM) reside at the same high-symmetry k -point in the Brillouin zone. Importantly, these monolayers retain their direct band gap character even after forming heterostructures.

To investigate the electronic properties of H-I, H-II and H-III heterostructures in more detail, the total density of states (TDOS) and partial density of states (PDOS) were calculated. The CBM and VBM of the WTeSe monolayer are formed by the Te 5p, Se 4p and W 4d orbitals, respectively.⁵⁰ Furthermore, the VBM of the BAs monolayer is primarily composed of As 4p orbitals, while the CBM is dominated by B 2p orbitals. These results are consistent with previous studies.²² This distinct



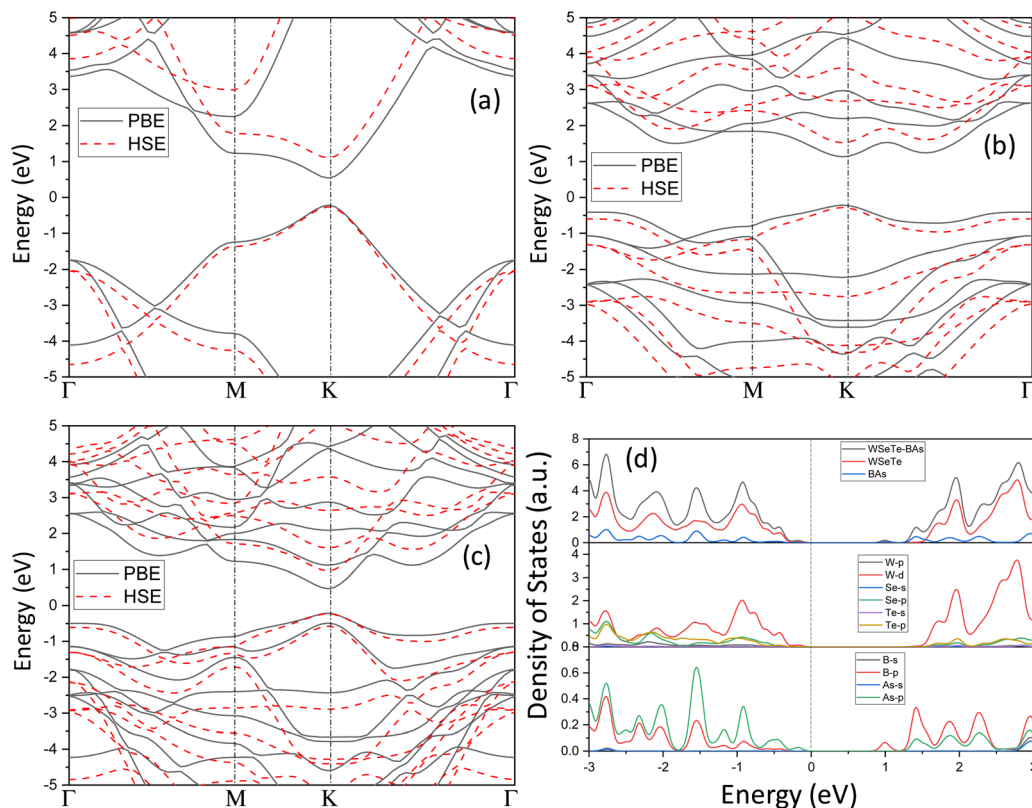


Fig. 3 The band structures of (a) BAS, (b) Janus WTeSe, and (c) H-I heterostructure computed using the PBE (black solid lines) and HSE06 (red dashed lines) functionals. (d) Density of states of the H-I heterostructure.

electronic structure of BAS allows for the formation of unique band alignments at the interface when combined with other materials. Notably, both type-II and Z-scheme configurations can be achieved, depending on the specific partner material. These band alignments create favorable conditions for efficient separation of photogenerated electron-hole pairs, making BAS-based heterostructures promising candidates for photocatalytic water splitting applications. Fig. 4 shows the projected band structures of H-I and H-II. In both heterostructures, the VBM is localized on WSeTe or WTeSe (represented by blue color), while

the conduction band minimum (CBM) is localized on BAS (represented by red color). H-II exhibits a smaller band gap (0.93 eV) compared to H-I (1.19 eV). The presence of Se vacancies in H-II further reduces the band gap to 0.59 eV, as shown in Fig. S7 of the ESI.†

Carrier mobilities

Within the framework of effective mass approximation (EMA), deformation potential (DP) theory has been a widely used approach to investigate carrier mobility (μ) in 2D materials.

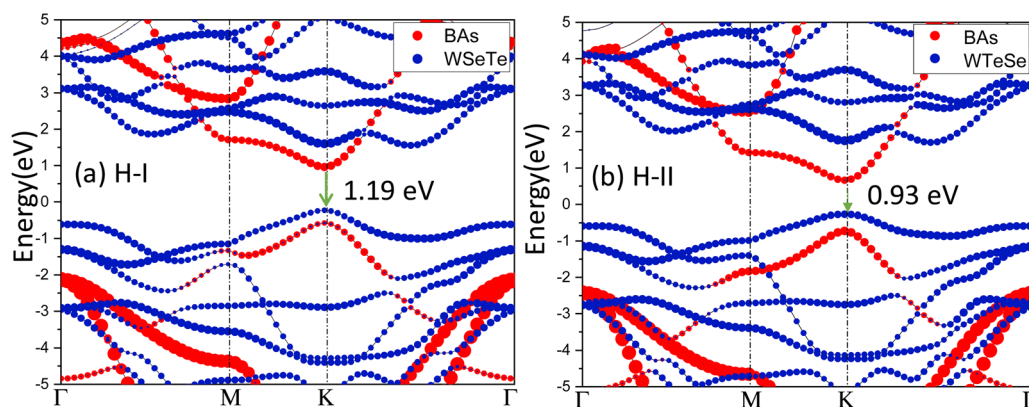


Fig. 4 Projected band structures of the (a) H-I and (b) H-II heterostructures.



The carrier mobility is defined as^{6,54}

$$\mu_{2D} = \frac{e\hbar^3 C_{2D}}{K_B T m_c^* m_d (E_1^i)^2} \quad (2)$$

The terms in the above equation are as follows: m_d , the average effective mass, which is defined as $m_d = (m_a^* m_b^*)^{1/2}$ (m_a and m_b are the mobility of electrons or holes in the armchair (x) and zigzag (y) directions); m_c^* , the carrier effective mass; T , the temperature; E_1^i , the DP constant, which describes the shift of the top of the valence band for holes and the bottom of the conduction band for electrons induced by the strain in the range of $\pm 1.5\%$ ($E_1^i = \Delta E_i / (\Delta l / l_0)$; see Fig. S8 in the ESI† for details); and C_{2D} , the elastic modulus of a and b , uniformly deformed crystals for emulating the lattice distortion activated by the strain. For the 2D structure, the in-plane stiffness can be determined by $C_{2D} = [\partial^2 E / \partial \delta^2] / S_0$, where E is the total energy of the supercell, δ is the applied uniaxial strain, and S_0 is the area of the equilibrium supercell. The details of the calculation of C_{2D} are given in Fig. S9 of the ESI†. The large difference between the electron and hole carrier mobilities will suppress the probability of recombination of photoinduced carriers and further improve the separation of electron-hole pairs.

Photocatalytic water splitting

The design of direct Z-scheme photocatalysts either experimentally or theoretically remains a challenge.^{55,56} DFT calculations can be used to identify the type of the heterostructure, particularly whether it follows a direct Z-scheme or type-II configuration.⁵⁷ The charge density difference ($\Delta\rho$), the built-in electric field, the electron-hole recombination, the effective mass, charge carrier mobilities, and the interlayer band gaps are important parameters to understand the mechanism. The $\Delta\rho$ at the interface was computed according to the following equation:

$$\Delta\rho = \rho_{\text{BAS/WSeTe}} - \rho_{\text{BAS}} - \rho_{\text{WSeTe}} \quad (3)$$

where $\rho_{\text{BAS/WSeTe}}$, ρ_{BAS} , and ρ_{WSeTe} represent the charge density values of the BAS/WSeTe (or BAS/WTeSe) vdWHs and of the BAS and WSeTe monolayers, respectively. The $\Delta\rho$ plots are given in the insets of Fig. 5. When two monolayers are in contact, electrons will flow through the interface due to the difference in their Fermi energies until these levels reach equilibrium.⁵⁸ The work function is given by the following expression:

$$\Phi = E_{\text{vac}} - E_{\text{F}} \quad (4)$$

where E_{vac} and E_{F} are the vacuum and Fermi energies, respectively. According to this equation, a higher work function (more energy required to remove an electron) corresponds to a lower Fermi energy.

The work functions for BAS and WSeTe are $\Phi = 4.97$ eV and $\Phi = 5.32$ eV, respectively. Due to its higher work function, BAS exhibits a lower Fermi energy compared to WSeTe. Consequently, charge transfer occurs at the interface, with electrons moving from BAS to WSeTe. This creates an electrostatic potential (built-in electric field) pointing from BAS to WSeTe,

which eventually halts further charge transfer once the Fermi energies of both materials equilibrate. This process leads to the formation of a stable heterostructure (H-I, H-II, H-III, and H-IV). Since our focus is on Z-scheme heterostructures for water splitting, WTeSe (H-II) and (H-III) were selected from the studied configurations. These configurations possess suitable band edge positions for efficient charge separation in a Z-scheme system (as shown in Table 2). While WTeSe (H-I) and (H-IV) exhibit high carrier mobilities, their band edge positions and the direction of the built-in electric field are not ideal for water splitting (Table 2). Consequently, these latter configurations may be more suitable for applications in electronic devices like diodes or transistors.^{59,60} However, Janus-WSeTe and Janus-WTeSe are polar materials, possessing an intrinsic electric field. This intrinsic field points in opposite directions within H-I compared to H-II and H-III. Consequently, the intrinsic field can either offset or reinforce the built-in electric field generated at the interface, leading to a net effective electric field (E_{eff}) across the heterostructure. In H-I, the opposing fields partially cancel each other, resulting in a weaker net E_{eff} . Conversely, for H-II and H-III, the fields align, leading to a stronger net E_{eff} . The value of E_{eff} depends on electrostatic potential ($D\phi$) and is given by the following expression:

$$E_{\text{eff}} = \frac{D\phi}{ed} \quad (5)$$

where d is the thickness. The values of $D\phi$ for the H-I, H-II and H-III heterostructures are 0.39, 0.57 and 0.42 eV, respectively.^{12,61} The effective electric field produced in the heterostructure could effectively inhibit the recombination of photo-generated electron-hole (e^-/h^+) pairs and enhance carrier separation in H-I, H-II, H-III and H-IV heterostructures, as shown in Fig. 5(g). The E_{eff} value of H-II is considerably higher than that of H-I. This indicates a faster e^-/h^+ recombination rate in the H-I heterostructure and a slower rate in the H-II heterostructure. Consequently, the H-II heterostructure exhibits superior catalytic performance and aligns with the type-II photocatalytic mechanism.^{14,26} In type-II or Z-scheme photocatalysts, the interlayer e^-/h^+ recombination should be much higher than the intralayer band gap to work without any mediators. This is only possible if the heterostructures have much smaller interlayer band gaps than intralayers and a larger ratio of the effective masses ($D = m_h^* / m_e^*$). Thus, this indicates that the e^-/h^+ recombination will be suppressed to enhance the catalytic performance (Table S1 in the ESI†). It is important to note that WTeSe (H-II) exhibits two highly stable configurations with a small difference in the formation energies: -0.254 eV and -0.264 eV. The most stable configuration possesses a band gap value of 0.93 eV, while the second most stable configuration has a band gap of 0.66 eV (see Fig. S12 in the ESI†). Notably, the smaller band gap of the second configuration may lead to a faster rate of e^-/h^+ recombination at the interface of a heterostructure. The above results show that H-I is type-II, while the H-II heterostructure is a direct Z-scheme photocatalyst.



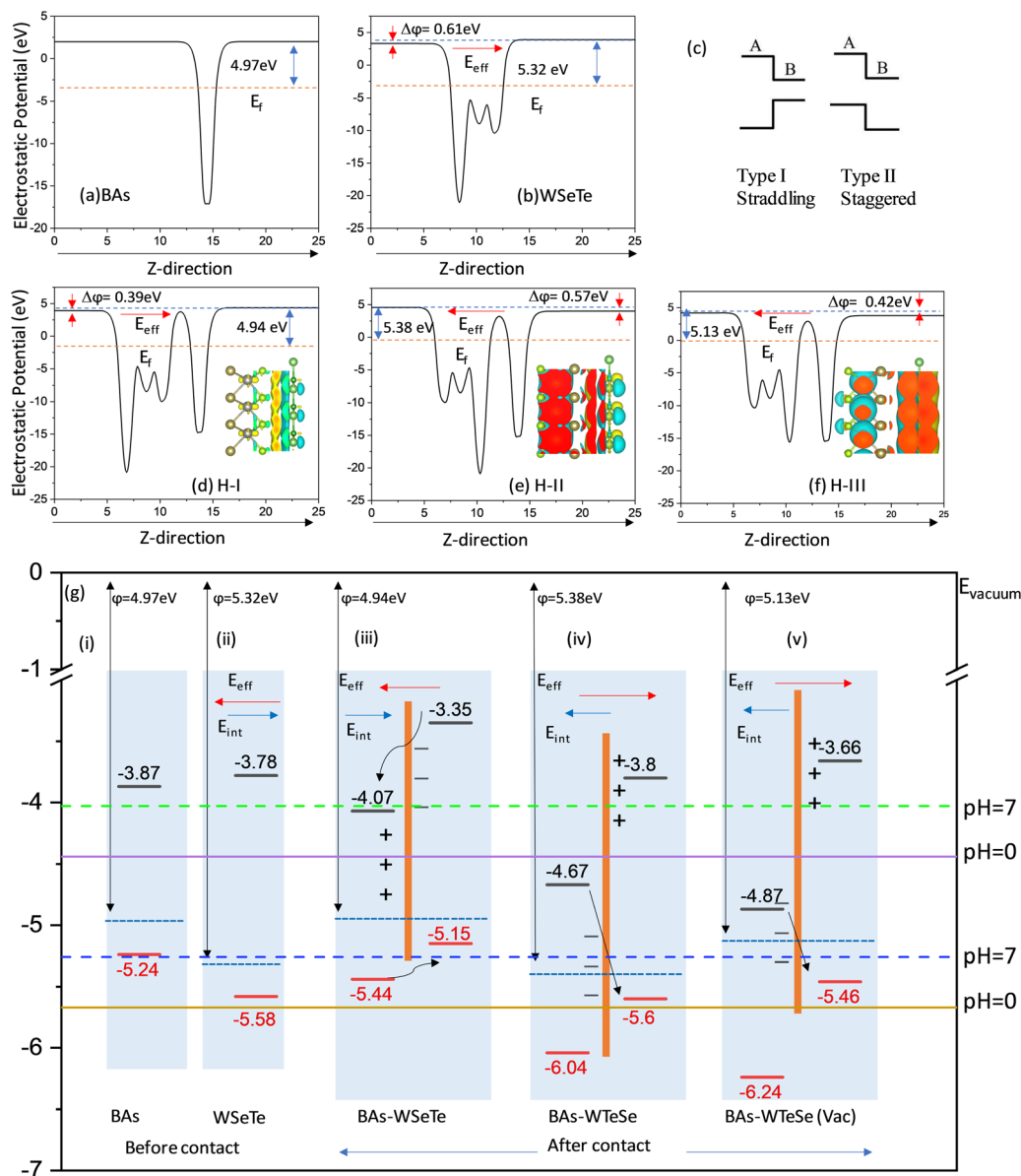


Fig. 5 Electrostatic potential of (a) BAS; (b) WSeTe; (d) H-I; (e) H-II; and (f) H-III. (c) Mechanisms of type-I and type-II heterostructures. (g) Schematic diagram of the Z-scheme mechanism for photocatalysis (VBM in red and CBM in black).

Table 2 The values of deformation potential constant E_1^i (eV), in-plane stiffness C_{2D} (Nm⁻¹), carrier effective mass m^* (m_e) and carrier mobility μ (cm² V⁻¹ S⁻¹) in both zig-zag and armchair directions

| Models | | m_a^* | m_b^* | E_{1a}^i | E_{1b}^i | $C_a/2D$ | $C_b/2D$ | $\mu_a/2D$ | $\mu_b/2D$ |
|-------------|---|---------|---------|------------|------------|----------|----------|------------|------------|
| BAS | e | 0.173 | 0.256 | 2.00 | 2.02 | 156.09 | 155.69 | 23422.60 | 15475.30 |
| | h | 0.172 | 0.260 | 3.01 | 3.02 | 156.09 | 155.69 | 10378.36 | 6808.04 |
| WTeSe | e | 0.390 | 0.433 | 2.01 | 3.17 | 134.91 | 132.36 | 4543.00 | 597.77 |
| | h | 0.535 | 0.675 | 5.22 | 4.91 | 134.91 | 132.36 | 337.21 | 296.29 |
| BAS-WSeTe | e | 0.184 | 0.255 | 4.64 | 4.51 | 292.67 | 293.06 | 7468.53 | 5706.25 |
| (H-I) | h | 0.372 | 0.448 | 4.64 | 4.52 | 292.67 | 293.06 | 1961.31 | 1715.22 |
| BAS-WTeSe | e | 0.192 | 0.277 | 4.52 | 4.51 | 292.53 | 292.67 | 7071.91 | 4927.57 |
| (H-II) | h | 0.360 | 0.498 | 4.64 | 4.64 | 292.53 | 292.67 | 1953.07 | 1411.68 |
| BAS-WTeSe-v | e | 0.265 | 0.211 | 4.69 | 4.87 | 285.31 | 279.94 | 4528.44 | 5177.09 |
| (H-III) | h | 0.582 | 0.522 | 5.52 | 3.65 | 281.31 | 279.94 | 630.19 | 1598.54 |
| BAS-WSeTe-v | e | 0.271 | 0.213 | 3.11 | 5.13 | 281.91 | 276.14 | 3586.91 | 218.05 |
| (H-IV) | h | 2.488 | 1.601 | 3.66 | 4.04 | 281.91 | 276.14 | 2971.92 | 204.25 |



Water splitting requires a photocatalyst with band edges strategically positioned relative to the water redox potentials. Moreover, water splitting can only occur on the surface of heterostructures if the CBM of semiconductor-II is more positive than the reduction potential of H^+/H_2 , and the VBM is more negative than the oxidation potential of $\text{O}_2/\text{H}_2\text{O}$. Therefore, the band alignments of H-I, H-II, and H-III were investigated, with the results presented in Fig. 5. The BAS and WTeSe monolayers are presented for comparison. The reduction potential of H^+/H_2 and the oxidation potential of $\text{O}_2/\text{H}_2\text{O}$ at different pH values are given by^{6,62}

$$E_{\text{H}^+/\text{H}_2}^{\text{red}} = -4.44 \text{ eV} + \text{pH} \times 0.059 \text{ eV} \quad (6)$$

$$E_{\text{O}_2/\text{H}_2\text{O}}^{\text{oxd}} = -5.67 \text{ eV} + \text{pH} \times 0.059 \text{ eV} \quad (7)$$

At a pH of 7, the reduction and oxidation potentials are -4.03 eV and -5.26 eV , respectively. From Fig. 5, the CBM energies of the BAS and WSeTe monolayers and of the H-I and H-II heterostructures are higher than the reduction potential of H^+/H_2 , while their VBM energies are lower than the oxidation potential of $\text{O}_2/\text{H}_2\text{O}$. This band alignment is favorable for photocatalytic water splitting. Fig. 6 also illustrates the photocatalytic reaction mechanism and the alteration of photoexcited carriers for H-I, H-II, and H-III heterostructures. Upon interaction with light, the heterostructure absorbs energy, causing electrons in the VBM to transition to the CBM, generating holes in the VBM and electrons in the CBM.

Subsequently, the photoinjected electrons in the CBM of BAS transfer to the VBM of Janus WSeTe, as shown in Fig. 5(g)(iii). However, in this configuration, the interlayer band gap of 1.19 eV leads to a rapid electron-hole recombination process due to the small value of $\Delta\phi$ (0.39 eV) and small difference between electron and hole carrier mobilities (Table 2).

Swapping the Se atoms with Te atoms in the H-I heterostructure significantly reduces the interlayer band gap to 0.93 eV . This reduction in interlayer band gap slows down the e^-/h^+ recombination rate, facilitating charge accumulation at the VBM and CBM due to the higher value of $\Delta\phi = 0.57 \text{ eV}$. This, in turn, enhances the catalytic efficiency of the heterostructure. Furthermore, introducing Te vacancies in the H-III heterostructure decreases the band gap to 0.59 eV as well as $\Delta\phi = 0.42 \text{ eV}$. These vacancies generate electron and hole trapping states at the VBM and CBM, which effectively increase the mass of electrons and holes and suppress e^-/h^+ recombination at the interlayer, potentially influencing the catalytic performance. In Fig. 6(a), the charge density with dipole correction at the interface is presented to elaborate the Te vacancy defects in contact with BAS monolayer. As we increase the vacancy defects, the charge density flips and H-III gains a high e^-/h^+ recombination rate. Fig. 6 illustrates the dipole correction across the interface between BAS and the Te atom of Janus WTeSe. Without dipole corrections, the electron density difference ($\Delta\rho$) appears delocalized in the middle of the interface (Fig. 6(b)). However, after considering the dipole correction term, $\Delta\rho$ becomes localized (Fig. 6(a)). Moreover, with Te vacancy

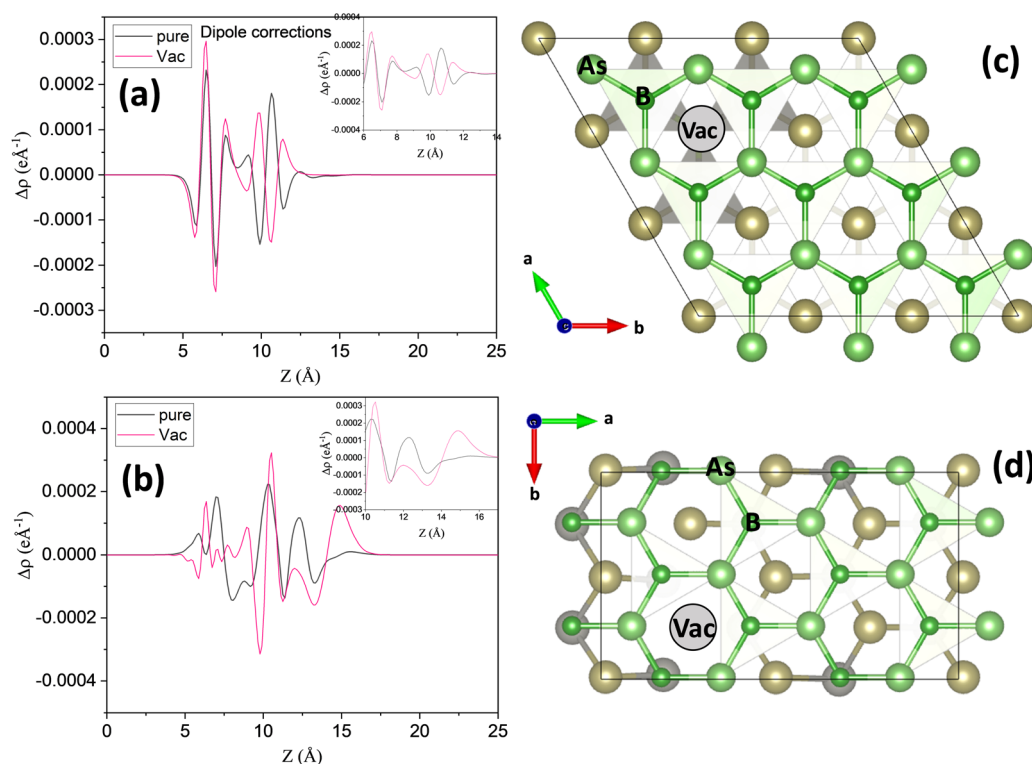


Fig. 6 Charge density at the interfaces of pure BAS–Janus WTeSe (black) with vacancies (red) computed (a) with dipole correction and (b) without dipole correction. Vacancy defect models: (c) hexagonal and (d) orthorhombic for strain energy calculations.



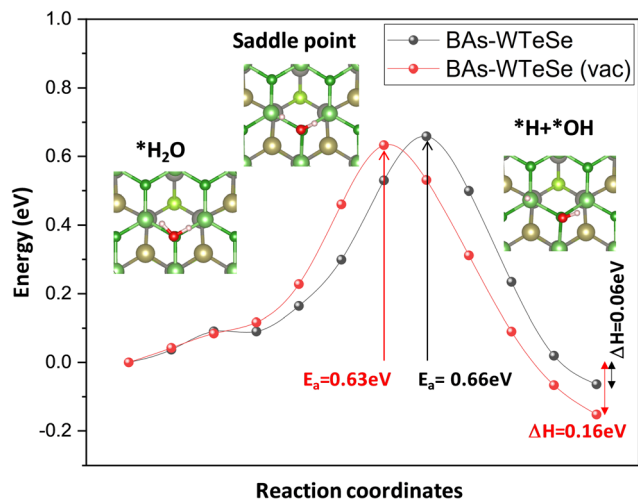


Fig. 7 Reaction profiles for the water splitting on BAS-WTeSe without Te vacancies (black line) and BAS-WTeSe with Te vacancies (vac) (red line).

defects, there is a flip in $\Delta\rho$ (Fig. 6(a)), shown in red. This flip in charge density significantly affects the built-in electric field, improving the Z-scheme heterostructure performance. Moreover, the charge carrier mobilities show that there is a large difference between electron and hole carrier mobilities that will improve the separation of e^-h^+ pairs (Table 2). This Z-scheme heterostructure works efficiently for water splitting due to suitable oxidation and reduction energy levels at -6.04 eV and -3.8 eV, respectively.

Further, climbing image nudged elastic band (CI-NEB) calculations were performed to investigate water splitting on BAS-WTeSe with and without Te vacancies. In both cases, the reaction from $*H_2O$ to $(*H + *OH)$ is exothermic ($\Delta H = 0.06$ eV and 0.16 eV), with activation energies of 0.63 eV and 0.66 eV, respectively. The presence of a Te vacancy slightly favors the water splitting reaction, as indicated by the lower activation energy (0.63 eV) compared to the pristine surface (0.66 eV). The saddle point was confirmed by frequency calculations for H-II and H-III, which showed only one imaginary frequency of 1237.6 cm^{-1} and 1238.0 cm^{-1} , respectively. These imaginary frequencies correspond to the vibrational mode of the breaking H-O bond in water, confirming the reaction pathway (Fig. 7).

Furthermore, to verify the stability of the BAS/WTeSe van der Waals heterostructures under reaction conditions, AIMD simulations of BAS-WTeSe, with and without Te vacancies, in contact with 90 water molecules were conducted. The simulations were performed in the NVT ensemble at 300 K for a total simulation time of 5 ps with a timestep of 1 fs . The structural integrity of the proposed heterostructures was maintained throughout the simulations, as evidenced in Fig. S10 (ESI[†]) by the well-preserved structure of the final configuration. Additionally, Fig. 11 (ESI[†]) demonstrates the stability of the total energy of the system over the simulation time, with no significant fluctuations observed. These results strongly suggest that the BAS/WTeSe heterostructures exhibit good stability in the presence of water molecules, making them promising candidates for water splitting applications.

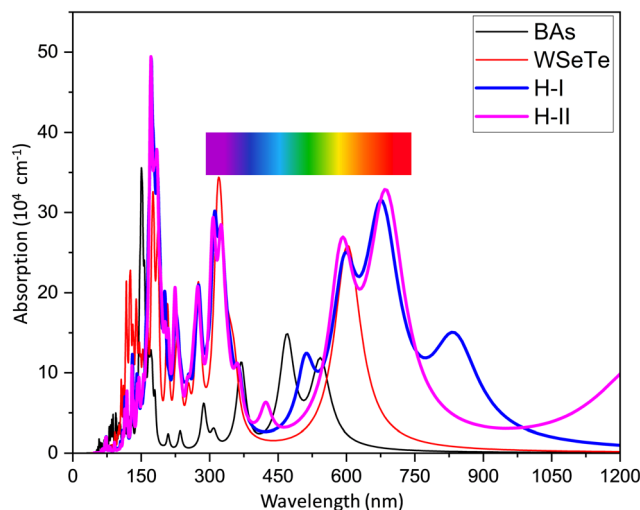


Fig. 8 The absorption coefficient of the monolayers BAS, WSeTe, and two different heterostructures.

Optical properties

The photocatalytic efficiency of a photocatalyst depends on its ability to capture photons in the visible region (380 to 760 nm). The light absorption coefficient is given as

$$\alpha = \frac{\sqrt{2}\omega}{c} \left\{ [\varepsilon_1^2(\omega) + \varepsilon_2^2(\omega)]^{1/2} - \varepsilon_1(\omega) \right\}^{1/2} \quad (8)$$

where c and ω show the speed of light and angular frequency in vacuum, and the dielectric constant has real $\varepsilon_1(\omega)$ and imaginary $\varepsilon_2(\omega)$ parts. Fig. 8 demonstrates the light absorption coefficients of Janus-WSeTe, BAS, H-I and H-II heterostructures. BAS exhibits maximum light absorption in the ultraviolet (150 , 210 , 280 nm) and visible regions (370 , 470 , 550 nm).¹² WSeTe monolayer has peaks at 322 nm and 610 nm . The absorption peaks for the WSeTe monolayer show a wide absorption range in the visible and ultraviolet regions which is better than that of BAS. Compared to monolayers, the heterostructures introduce two new peaks in the visible region: 424 nm for H-I (blueshift) and 514 nm for H-II (redshift). This not only significantly expands the light absorption range but also substantially increases the absorption intensity, demonstrating that the heterostructures exhibit superior light utilization in the visible and ultraviolet regions compared to monolayers. This enhancement in the absorption coefficient of light is attributed to charge transfer, vacancies, and interplanar coupling. These factors collectively improve the photocatalytic performance of heterostructures for water splitting.

Conclusions

In summary, DFT calculations were employed to investigate the electronic properties and photocatalytic mechanism of BAS/Janus-WSeTe and BAS/Janus-WTeSe heterostructures. The stability of these heterostructures was verified by phonon spectrum analysis and *ab initio* MD simulations. Our results revealed the formation of type-II and Z-scheme



heterostructures with direct band gap values of 1.19 eV and 0.93 eV, respectively. Charge carrier mobilities calculated using DP theory reveal that electron–hole recombination is significantly suppressed in the Z-scheme heterostructure with Te vacancies, leading to a reduced interlayer band gap of 0.59 eV. Charge redistribution at the interface induced a built-in electric field. Upon interaction with visible light, this built-in electric field accumulated photoinjected holes in the VBM of Janus-WSeTe or Janus-WTeSe and photoinjected electrons in the CBM of BAS. Overall, the Z-scheme (BAS/Janus-WTeSe) heterostructure exhibited superior catalytic efficiency compared to the type-II (BAS/Janus-WSeTe) heterostructure, demonstrating its potential as a promising catalyst for water splitting.

Author contributions

Conceptualisation of work: HBB, AJ, SZI, AGN and DDT; conducting of experiments: HBB and AGN; computations: HBB and AGN; data analyses: HBB, SZI, AGN and DDT; data dissemination and graphics: HBB, SZI, AGN, and DDT; writing of manuscript: HBB, AJ, AGN, and DDT; project support: DDT.

Conflicts of interest

There are no conflicts to declare.

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