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Introduction

The electrochemical carbon dioxide reduction reaction (CO_2RR) has attracted enormous interest due to its mild reaction conditions and potential to be used for renewable electricity storage via the production of synthetic fuels. Particularly, the CO_2RR to ethylene (C_2H_4) aiming at high current densities and faradaic efficiencies (FE) is intensively studied because of the extremely high industrial value and the need to transition away from fossil fuel C_2H_4 production.¹ Copper-based $(Cu,^{2-4} CuO^{5-8}$ and $Cu_2O^{9-14})$ electrodes are proven to be the most used heterogeneous catalysts that tend to produce hydrocarbons and oxygenate compounds with competitive activity. However, as multi-step electron and proton transfer processes are involved in C_2H_4 formation, hydrogen (H₂) and other by-products such as

CeO₂-promoted Cu₂O-based catalyst sprayed on the gas diffusion layer for the electroreduction of carbon dioxide to ethylene[†]

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The development of efficient and selective catalysts for the carbon dioxide reduction reaction (CO₂RR) is crucial for sustainable energy and chemical synthesis. In this work, CeO_{2-y} (y = C (cubic) and R (rod)) was incorporated into Cu₂O nanocube electrocatalyst as a promoter for ethylene (C₂H₄) production. The results demonstrate that the catalyst with a loading of 5 wt% crystalline CeO_{2-C} exhibits competitive activity and stability for ethylene production compared to pristine Cu₂O. Under optimized reaction conditions of -250 mA cm⁻² current density and 1 M KOH electrolyte, the Cu₂O-5CeO_{2-C} catalyst achieved a faradaic efficiency (FE) of ~53% for C₂H₄ production, while maintaining stability over a period of 120 minutes. In contrast, non-promoted Cu₂O and CeO_{2-C} creates intrinsic sites (Cu^x-CeO_{2-x}; Cu^x = Cu²⁺, Cu⁺, and Cu⁰) for the binding of CO₂ and H₂O molecules. Moreover, the Cu₂O-5CeO_{2-C} catalyst outperforms other reported systems in terms of FE and partial current density for C₂H₄ production. It requires a lower potential (-0.98 V vs. RHE) to operate at the same electrolyte concentration. This finding highlights the promising nature of Cu₂O-5CeO_{2-C} as an efficient and cost-effective catalyst for C₂H₄ production.

methane (CH₄) will inevitably be produced during electrolysis. Therefore, the design of efficient catalysis systems specific for the CO_2RR to ethylene with both high selectivity and high FE as well as low overpotential is highly desirable.

Various catalyst design strategies have been reported to effectively regulate the selectivity of the CO_2RR .^{15–17} Notably, by altering the electronic structure of Cu-based catalysts by the addition of a second metal^{18–20} or metal oxide promoter²¹ phase provides a significant improvement for C_2H_4 production. Ceria (CeO₂) is widely used as a catalyst, support and promoter for a variety of heterogeneous catalytic reactions involving the hydrogenation of CO_2^{22} due to its acid–base and unique redox properties of oxygen storage and release.²³ The size and shape modification, surface/face reconstruction, and faceting of ceria at the nanoscale level can offer an important tool to govern activity and stability in these reactions.²⁴ Furthermore, strong interaction between noble metals and ceria leads to their higher dispersion, electronic modifications and enhanced catalytic activity.²⁵

Recently, the implementation of CeO₂ as a support material for the CO₂RR has been investigated by some researchers. For instance, a high methane (CH₄) faradaic efficiency (up to ~54% at -1.2 V vs. RHE) over Cu/CeO_{2-x} nanocrystalline heterodimers was reported by S. Varandili *et al.*²⁶ In another

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study, Cu/CeO₂ carbon nanofiber (Cu/CeO_x@CNF) catalysts synthesized by an electrospinning method were explored by X. Zong et al.²⁷ In this case, their optimized Cu/CeO_x@CNF_{S-2} catalyst exhibited a high CO faradaic efficiency (up to $\sim 59\%$ at -0.6 V vs. RHE) at a high current density of 100 mA cm⁻². Most recently, Zhao et al.28 have that the Cu electrode coating with CeO₂ nanoparticles largely enhanced the C₂₊ product selectivity during CO₂ electroreduction. The FE and partial current density of C₂₊ products on the CeO₂-Cu electrode achieved in a gas-tight H-type electrolytic cell with 0.1 M KHCO3 was 61% and 8.45 mA cm⁻² at -1.05 V vs. RHE, respectively. In that study, 1.5 mg cm⁻² was found to be the optimal CeO₂ coating. The improved CO2RR selectivity and activity of CeO2-Cu were attributed to the interface between Cu and CeO₂, which promotes C-C coupling towards C_{2+} production products. The CO₂RR to C_{2+} has also been evaluated over a CuO modified 20 wt% CeO2 catalyst using a flow cell at 1 M KOH. Under those conditions, the reported faradaic efficiency of the C_{2+} products was ~75.2% at a current density of 1.21 A cm⁻². S. Yan et al.²⁹ revealed that CeO2 and Cu and the subsurface Cu2O coexisted in CeO2/CuO during the CO₂RR and two competing pathways for C-C coupling were promoted separately, of which hydrogenation of *CO to *CHO is energetically favoured.

In terms of the CO_2RR to C_2H_4 , the stabilization of Cu^+ within a CuO-CeO₂ interface has been reported by S. Chu et al.³⁰ They suggested that tuning the CuO/CeO₂ interfacial interaction permits dramatic suppression of proton reduction and enhancement of CO2 reduction. The C2H4 faradaic efficiency (up to $\sim 50\%$ at -1.1 V vs. RHE) was obtained over an optimised CuO/30 wt% CeO2 using a liquid H-type cell with continuous CO₂ bubbling in 0.1 M KHCO₃ electrolyte. Lately, the effect of exposed facets of CeO_2 (cubes (100), rods (110), and octahedral (111)) has also been investigated under similar reaction conditions by S. Chu et al.³¹ In the CeO₂-supported Cu nanoparticles, they found that CeO₂ changed the oxidation state of Cu atoms towards Cu⁺ at the CuO-CeO₂ interface. This fact was mainly identified on the Cu/CeO₂(110), flowed by Cu/ $CeO_2(100)$ and $Cu/CeO_2(111)$. The existence of Cu^+ species was supposed to be likely the adsorption and active sites for CO₂ activation followed by further C-C coupling to yield C₂H₄. The FE towards C_2H_4 was ~39% on the Cu/CeO₂(110) at a mild overpotential of 1.13 V. In another study performed on a flowcell at high current density and under 1 M KOH, Ce-doped Cu nanoparticles (Ce-Cu NPs) were reported to be highly selective to C₂H₄. A high faradaic efficiency (FE) of C₂H₄ close to 53% at a current density of 150 mA cm⁻² was achieved using a Ce/(Cu + Ce) precursor ratio = 10%. J. Shan *et al.*³² suggested that the high performance, which is 2.8 times higher than Cu nanoparticles under the same conditions, is due to the shrinking of the particle sizes due to Ce doping resulting in more catalytic active sites with oxygen defects. It seems that the proximity of the Ce atoms can boost the local electronic distribution on the Cu nanoparticles. Under similar reaction conditions, anchoring CeO₂ quantum dots on to the CuO surface has been evaluated by S. Wang et al.33 Their optimal catalyst consisting of 100 wt% CeO2 and 50 wt% CuO achieved a high faradaic efficiency of \sim 50% for ethylene with a partial current density of 197 mA cm⁻², attributed to the CuO/CeO₂ interfaces that simultaneously stabilize Cu⁺ and key intermediates.

Despite these advances in CuO supported CeO₂ electrocatalysts for CO₂RR to C₂H₄, the role of the CeO₂ as a promoter using Cu₂O instead of CuO has not been investigated yet. Recent studies have revealed that the Cu₂O phase provides a higher selectivity to ethylene compared to CuO.³⁴ Furthermore, the economic viability of most large-scale chemical processes relies on the use of simple and low-cost catalytic materials. Therefore, the implementation of CeO₂ as a promoter instead of bulk support is a practical way to design cost-effective electrocatalytic materials. In this line, the promoter content optimization into Cu₂O can provide a highly active-stable, and low-cost electrocatalytic formulation for C₂H₄ production.

Here, we developed a CeO2-promoted Cu2O-based catalyst sprayed on a carbon-based gas diffusion layer for the CO₂RR to C_2H_4 . A series of $Cu_2O-xCeO_2$ were synthesized by the simple liquid phase reduction method. The influence of CeO₂ crystallinity and promoter loading (x = 5-20 wt%) was studied under relevant reaction conditions (current density (i) = -50300] mA cm⁻², 1 M KOH, and flow of CO₂ (F_{CO_2}) = 200 mL min⁻¹). Furthermore, the influence of the KOH concentration (0.1-3 M) and electrolyte type (KOH, KCl, and KHCO₃,) over the optimized Cu₂O-5CeO₂ GDE was investigated to identify the best reaction components. Lastly, its stability was additionally tested under the selected reaction conditions ($j = -250 \text{ mA cm}^{-2}$, electrolyte = 1 M KOH, and $F_{CO_2} = 200 \text{ mL min}^{-1}$). The physico-chemical properties responsible for catalyst performance were studied using scanning electron microscopy-energy dispersive X-ray spectroscopy (SEM-EDX), transmission electron microscopy (TEM), X-ray diffraction (XRD), and Raman and photoelectron spectroscopy (XPS).

Experimental section

Chemicals and reagents

All chemical reagents used were of analytical grade without further treatment. Copper(II) sulfate [CuSO₄, 99%] and Nafion[®] perfluorinated resin solution (5 wt% in a mixture of lower aliphatic alcohols and water, 45% water) were purchased from Sigma Aldrich. Cerium(III) nitrate hexahydrate [Ce(NO₃)₃·6H₂O, 99.5%] was purchased from Thermo Scientific. L(+)-Ascorbic acid [C₆H₈O₆, 99%] was purchased from Acros Organics. Sodium hydroxide [NaOH, 97%], potassium hydroxide [KOH], and ethanol absolute [C₂H₆O, 99.8%] were purchased from Fisher Scientific. Ultrapure water was provided by Milli-Q Millipore source (18.2 M Ω cm, 20 °C).

Preparation of the catalysts

The promoter phase, CeO_2 , was first synthesized using a hydrothermal method. In a typical synthesis, 3.25 g of $Ce(NO_3)_3$ ·6H₂O was dissolved in 15 mL of deionized water to obtain cerium nitrate aqueous solution. Then, a NaOH aqueous solution was prepared in a Teflon-lined steel reactor using 9 g

of NaOH and 50 mL of deionized water. The cerium nitrate aqueous solution was later added dropwise into the NaOH aqueous solution. After magnetic stirring for 30 min, the reactor was sealed and heated at two different temperatures of 80 °C and 240 °C for 24 h to obtain rod (R) and cubic (C) ceria particles, respectively. After cooling down to room temperature, the resulting precipitate was separated by centrifugation (4000 rpm for 5 min) and washed three times with excess ethanol, followed by drying in an oven overnight at 80 °C.

The series of CeO₂-promoted Cu₂O (Cu₂O–*x*CeO_{2-y}; *x* = 5, 10, 15 and 20 wt% and y = rod (R) and cubic (C) shape) samples were synthesized using a simple liquid phase reduction method. CuSO₄ (0.2689 g) and the respective amount of the promoter (CeO₂) were dissolved into 50 mL of distilled water under sonication for 30 min and later by magnetic stirring for 30 min. Sequentially, 0.8 g NaOH was dissolved in 50 mL of distilled water step, 0.7045 g of C₆H₈O₆ as a reducing agent, was added to the mixture and kept constantly stirring for 30 min. Then, the CeO₂-promoted Cu₂O catalysts were centrifuged at 4000 rpm for 5 min and washed three times with excess ethanol. Finally, they were dried overnight in an oven at 80 °C. A non-promoted Cu₂O was also synthesized using the same procedure.

Preparation of gas diffusion electrode (GDE)

The GDE usually consists of a gas diffusion layer (GDL) on which a catalytic layer is deposited. The carbon-based GDL (Freudenberg H23C6) was used as the support. A catalytic ink was formulated by mixing 20 mg of the as-prepared catalyst, 0.5 mL of ethanol absolute and 140 μ L of Nafion[®] perfluorate. The ink deposition on the GDL was performed using the spray coating technique. The catalyst loading mass was fixed to 1 ± 0.05 mg cm⁻² according to the previous control test, see Fig. SI1 (ESI†). After the catalyst deposition, GDEs were dried at room temperature. The GDEs were denoted as Cu₂O, Cu₂O–5CeO_{2-C}, Cu₂O–10CeO_{2-C}, Cu₂O–15CeO_{2-C}, Cu₂O–20CeO_{2-C}, CeO_{2-C}, Cu₂O–5CeO_{2-R}, and CeO_{2-R}. Fig. SI2 (ESI†) shows a systematic representation of the catalyst and GDE preparation.

CO₂ electroreduction set-up

The electrocatalytic tests were carried out using a compact H-type electrochemical cell.³⁵ The cell is shown in Fig. SI3 (ESI⁺), and included a current collector made of stainless steel with a built-in spiral flow field, which connected the cathode compartment and the GDE of the cell and guaranteed homogeneous distribution of the inlet reactants. In the first series of experiments, the influence of the promoter phase (CeO₂) was investigated using a CO₂ flow rate of 200 mL min⁻¹, which was adjusted by a manual flow meter (FR2000, Key Instruments). The working gas diffusion electrodes had a projected area of 1 cm² to the 1 M KOH aqueous electrolyte, while a platinum square mesh of 1×1 cm² was used as the counter electrode, located at the anode compartment and separated from the catholyte compartment by a proton exchange membrane (PEM, Nafion[®] 115) and immersed in 1 M KOH aqueous electrolyte. The total volume of the electrolyte was 150 mL. The reference

electrode was an Hg/HgO (1 M NaOH) electrode. The potential values were then translated into RHE (reversible hydrogen electrode) voltages by using eqn (1) and verified by a hydrogen reference electrode (Hydroflex^(R)(M-v01-0071)).

$$E(vs. \text{ RHE}) = E\left(vs. \frac{\text{Hg}}{\text{HgO}}\right) + 0.14 \text{ V} + 0.059 \text{ V} \times \text{pH} \quad (1)$$

For each electrode, the faradaic efficiency (FE) was evaluated by varying the current density from -50 to -300 mA cm⁻² ($\Delta j = -50$ mA cm⁻²) and using cronopotentiometry (PGSTAT204, Metrohm-Autolab). At each current density, the presence of gas products from the cathode outlet stream was examined for 5 minutes. During this time, the volume of the outlet products was measured and then analysed using an on-line gas chromatography (Shimadzu GC 2030) equipped with a Porapak Q 80/ 100 column. The gases (H₂, CO, CH₄ and C₂H₄) were detected using a thermal conductivity detector (TCD) and a flame ionization detector (FID).

After the identification of the most promising Ce-promoted copper-based GDE, the influence of the KOH concentration (0.1 M–3 M) and electrolyte type (KCl, KOH, KHCO₃,) over the optimized Cu₂O–5CeO_{2-C} GDE was additionally investigated to find the most effective electrolyte conditions. Finally, a stability test was performed at the selected reaction conditions (j = -250 mA cm⁻², electrolyte = 1 M KOH, $F_{CO_2} = 200$ mL min⁻¹).

The FE_X of the X obtained products, such as C_2H_4 , CH_4 , CO and H_2 , were estimated by using the following equation (eqn 2):

$$FE_{X} = \frac{Q_{X}}{Q_{Total}} = \frac{n_{X}N_{X}F}{Q_{Total}}$$
(2)

where Q_X and Q_{Total} are the charge passed to produce product X and the total passed charge (C) during CO₂RR, n_X represents the electron transfer number of product X, N_X is the product amount (mol) of X measured by the GC, and *F* is the Faraday constant (96 485 C mol⁻¹).

Catalyst characterization

The series of catalysts sprayed on the carbon-based gas diffusion layer were characterized by different physico-chemical techniques like scanning electron microscopy-energy dispersive X-ray spectroscopy (SEM-EDX), transmission electron microscopy (TEM), X-ray diffraction (XRD), and Raman and photoelectron spectroscopy (XPS).

The surface morphology and elemental composition were investigated using a JEOL JSM-7200F high-resolution SEM equipped with an EDX detector (Oxford instruments) and AZtecEnergy software. Prior to EDX analysis, catalysts were fixed over an aluminium holder. Elemental analysis was carried out at 20 kV and restricted to Cu, Ce, and O to avoid inconsistent results. TEM images were recorded on a FEI Tecnai T12 electron microscope operating at 120 kV. The powder samples were dispersed by ultrasound in ethanol. Suspension drops were deposited on a holey carbon-coated copper grid of 300 mesh and then dried in air.

The crystal structure of the catalysts was examined using a Bruker type XRD D2 Phaser diffractometer. X-ray powder

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diffraction patterns were acquired by applying a Cu K α radiation ($\lambda = 1.5406$ Å), a voltage of 40 kV, and a current of 40 mA. Continuous scans were collected over the 2 θ range of 10° to 80°. The crystalline phases present in the samples were identified by comparison with reference patterns available in the JCPDS database. The crystallite sizes of the main Cu^x species (Cu₂O, CuO and Cu⁰) and cerium oxide (CeO₂) were estimated using the Scherrer's equation at 2 $\theta = 36.46^{\circ}$ for Cu₂O (111), 38.76° for CuO (111), 43.32° for Cu (111), and 28.53° for CeO₂ (111); $D = (K\lambda/\beta \cos \theta)$, where λ is the X-ray wavelength, β is the full width of the diffraction line at half maximum (FWHM), and θ is the Bragg angle.

Raman measurements were recorded on a Raman spectrometer (Horiba LabRAM HR800) equipped with a CCD detector, using a laser with a wavelength of 532 nm as the excitation source. The Raman spectra were obtained using a x50LWD objective with an incident power of 0.1 mW (5 acquisitions, 60 s of exposure time).

X-ray photoelectron spectroscopy (XPS) experiments were performed in a PHI 5500 Multitechnique System (from Physical Electronics) with a monochromatic X-ray source (Al K α 1486.6 eV).

Results and discussion

Characterization

SEM images displayed in Fig. SI4 (ESI[†]) evidenced that the formation of particles with cubic-like morphology was achieved for both the active phase (Cu₂O) and promoter phase (CeO_{2-C}) prepared at 240 °C. In the case of CeO_{2-R} prepared at 80 °C, its morphology was not possible to identify by SEM because the particle sizes were below the detection limits. When the electrocatalysts were synthetised in the presence of CeO_{2-C}, a catalytic system composed of particles with a cubic shape was also achieved for the series of CeO_{2-C}-promoted Cu₂O-based catalyst. Furthermore, EDX analysis performed over all catalysts proved the presence of their main components (Cu, Ce, and O). As a representative example, the EDX elemental mapping image exhibited in Fig. 1 demonstrated the experimental composition (82 wt% Cu, 5 wt% Ce, and 13 wt% O) and interaction between the cubic particles of the Cu₂O-5CeO_{2-C} catalyst.

A summary of the elemental composition of all the asprepared catalysts is shown in Table 1. The experimental composition of the CeO₂ promoter phase falls within the range of 6–21 wt% \pm 1, which closely matches the theoretical values. On the other hand, a slight difference in the elemental composition was mainly identified between the two synthesized promoters (CeO_{2-R} (81 wt% Ce and 19 wt% O) and CeO_{2-C} (83 wt% Ce and 17 wt% O)), suggesting the possible presence of different oxidation state of cerium (Ce³⁺ and Ce⁴⁺) over these samples. Therefore, the used temperature for the synthesis of CeO₂ nanoparticles can be a key parameter for the generation of Ce³⁺ and Ce⁴⁺.

The cubic-like morphology of the CeO_{2-C} promoter (Fig. 2(a)) and pristine Cu_2O (Fig. 2(b)) was additionally confirmed by



Fig. 1 Surface morphology and elemental composition analysis for the $Cu_2O-5CeO_{2-C}$ GDE. (a) Image of the as-prepared GDE. (b) SEM image showing the catalyst layer and the carbon-based GDL. (c) SEM image of the electron micrograph region, (d) EDX spectrum, (e) and distribution of Cu, Ce and O in elemental mapping.

Table 1 Elemental composition of the catalysts

	Elemental composition		
	Cu	Ce	0
Sample		[wt% ±2]	
Cu ₂ O	89	_	11
Cu ₂ O-5CeO _{2-C}	82	5	13
Cu ₂ O-10CeO _{2-C}	75	9	16
Cu ₂ O-15CeO _{2-C}	70	12	18
Cu ₂ O-20CeO _{2-C}	62	17	21
CeO _{2-C}	_	83	17
CeO _{2-R}	_	81	19
Cu ₂ O-5CeO _{2-R}	81	4	15

TEM. Regarding the CeO_{2-R}, rod-like morphologies with particle sizes of ~ 8 nm and irregular length can be identified, see Fig. SI5 (ESI[†]). Between cubic particles, the main difference observed was the size. The cubic CeO_{2-C} particles (~ 29 nm)



Fig. 2 TEM images of the (a) CeO $_{\rm 2-C}$ (b) Cu_2O, and (c) Cu_2O–5CeO $_{\rm 2-C}$ catalysts.

were smaller than cubic Cu₂O particles (~165 nm). The interaction between the CeO_{2-C} and Cu₂O particles was clearly detected over a series of Cu₂O–CeO_{2-C} catalysts. Fig. 2(c) is an example of the strong Cu₂O–CeO₂ interaction achieved over the Cu₂O catalyst promoted by 5 wt% CeO_{2-C}. For this CeO_{2-C}promoted Cu₂O catalyst, a decrease in the Cu₂O particle sizes was found, suggesting that the addition of CeO₂ during the synthesis inhibits the growth of Cu₂O particles. The average particle sizes estimated for CeO_{2-C} and Cu₂O were 43 nm and 160 nm, respectively.

The XRD patterns of the CeO_{2-C} promoter, pristine Cu₂O and the series of CeO_{2-C}-promoted Cu₂O are shown in Fig. 3. The XRD profiles confirmed the phase purity of the polycrystalline CeO2 (JCPDS: 00-034-0394) and Cu2O (JCPDS: 00-005-0667). For the CeO_{2-C} sample, the high-intensity reflections are observed at $2\theta = 28.53$, 33.09, 47.48, 56.33, 59.08, 69.41, 76.69, and 79.04° corresponding to the CeO₂ (111), (200), (220), (311), (222), (400), (331), (420) lattice planes. On the other hand, the diffraction peak at around 2θ = 29.61, 36.46, 42.34, 61.42, and 73.56° represent the Cu₂O (110), (111), (200), (220), and (311) lattice planes. Over the series of Cu₂O-CeO_{2-C} catalysts, CeO₂ and Cu_2O reflections were identified at a similar 2θ position, but with differences in their relative intensity, in agreement with the composition. With higher CeO₂ content, the intensity of the reflections of CeO2 increased, while the reflections of Cu2O were decreased. Compared to CeO_{2-C} promoter, the XRD profiles of CeO_{2-R} suggested that the sample appeared to be composed of small particles. Only four reflections were identified for the CeO_{2-R}. As can be seen in Fig. SI6 (ESI[†]), the reflections related to CeO2 phase in the Cu2O-CeO2-R were under the detection limit. These reflections were identified at 2θ = 28.53, 33.09, 47.48, 56.33° and associated to the CeO₂ (111), (200), (220), (311) lattice planes.

On the other hand, Raman spectroscopy analysis was additionally performed over the most promising fresh samples (CeO_{2-C}, Cu₂O, and Cu₂O–5CeO_{2-C}) taking into consideration that the formation of defective sites can be interpreted by the alteration in M–O vibration frequency after CeO₂ introduction to the Cu₂O structure. Fig. 4 proved that the fluorite-type cubic



Fig. 3 XRD Patterns of the catalysts.





crystal structure of ceria in the CeO_{2-C} sample exhibits one firstorder Raman active fundamental mode located at 461 cm⁻¹.³⁶ This F_{2g} mode was associated with the symmetric vibrations of oxygen ions around Ce⁴⁺ ions in octahedral Ce–O₈.³⁷ The bands at 480 and 592 cm⁻¹ were attributed to the oxygen vacancies due to the presence of reduced Ce³⁺ cations (Ce³⁺O₈)³⁸ and the vacancy-interstitial Frenkel-type oxygen intrinsic defects in pure ceria,³⁹ respectively. Additionally, solid-state phonons assigned to second-order features $2\omega R(X)$ were observed at 1065 cm⁻¹.⁴⁰

In the non-promoted Cu₂O sample, the most intense Raman peak identified at 215 cm⁻¹ was related to the second-order overtone $2\Gamma_{12}^{-}$.⁴¹ The remaining Raman bands were assigned as follows: an inactive mode Γ_{12}^{-} at 104 cm⁻¹, an IR active mode $\Gamma_{15}^{(1)}$ (LO) at 145 cm⁻¹, a fourth order overtone $4\Gamma_{12}^{-}$ at 412 cm⁻¹, and an IR active mode $\Gamma_{15}^{(1)}$ (TO) at 629 cm⁻¹.⁴²⁻⁴⁴ In addition to the prominent peaks observed for Cu₂O, a weak band was also observed at 320 cm^{-1} , which corresponds to the A_g mode of CuO.⁴⁵ On the other hand, the introduction of CeO₂ into the Cu₂O structure induced the formation of a new extrinsic MO₈ site capable of delivering oxygen under reducing conditions, *i.e.* part of a Frenkel defect.⁴⁶ This new Raman peak identified over the Cu₂O-5CeO_{2-C} catalyst located at 648 cm⁻¹ can be probably an effect of the strong interaction between Cu⁺ and Ce³⁺ cations.⁴⁷ Furthermore, the presence of the CuO phase was not identified in comparison with the Cu₂O sample, implying that CeO₂ avoids overoxidation of Cu₂O and preserves the Cu(I) oxidation state.

The surface elemental composition and chemical state of the fresh CeO_{2-C}, Cu₂O and Cu₂O–5CeO_{2-C} samples were examined by X-ray photoelectron spectroscopy. The survey scan in Fig. SI7 (ESI[†]) shows the presence of C, Cu, Ce and O. The C 1s peak at 284.4 eV was used as a reference peak for calibration of all the spectra. The high-resolution XPS spectra of Cu 2p are presented in Fig. 5. The peaks at 932.1 eV and 951.9 eV were linked to Cu⁺ 2p_{3/2} and Cu⁺ 2p_{1/2}, respectively. Furthermore, characteristic shoulder peaks can also be resolved at 934.8 and 954.8 eV, which were associated with Cu²⁺ 2p_{1/2} and Cu²⁺ 2p_{3/2}, respectively. The presence of satellite peaks at the binding Paper



Fig. 5 Cu2p XPS spectra of CeO_{2-C}, Cu₂O and Cu₂O-5CeO_{2-C} catalysts.

energies of 943.4 indicated that $\mathrm{Cu}(\mathrm{I})$ was the major valence state for Cu species. 48

The peaks of the Cu₂O–5CeO_{2-C} nanocomposites shift in the direction of high binding energy compared with that of the Cu₂O nanocrystals. The difference between the two peaks is 0.5 eV from the binding energy, which indicates that there is an electronic exchange between Cu₂O and CeO₂. In particular, the Cu⁰ phase was identified in the Cu₂O–5CeO_{2-C} sample, suggesting that CeO₂ promotes the reduction of CuO. Regarding the CeO₂ phase (see Fig. SI8, ESI[†]), peaks at 881.9, 888.1, 897.9,

899.8, 901.8, and 916.1 eV were only detected in the Ce 3d spectrum of the CeO₂ sample, corresponding to the binding energies of Ce⁴⁺ $3d_{5/2}$, Ce³⁺ $3d_{5/2}$, SU₁, Ce⁴⁺ $3d_{3/2}$, SU₂, and Ce³⁺ $3d_{3/2}$, respectively.⁴⁹ Fig. SI9 (ESI[†]) depicts the high-resolution O 1s spectra of the different catalysts. The characteristic peaks resolved in the binding energy range of 534.6 to 526.3 eV were assigned to lattice oxygen (O_{lat}) and oxygen vacancies (O_{Vs}).^{50,51} Compared to Cu₂O, the relative proportion of oxygen vacancies was higher for the Cu₂O–5CeO₂. The Ov_s are renowned for their possession of weakly bound electrons, serving as exceptional Lewis base sites for CO₂ adsorption. These electrons contribute to the formation of the CO₂•⁻ intermediate by providing electron donation.^{52,53} Therefore, this indicates that the Cu₂O–5CeO₂ exhibits the strongest ability for CO₂ adsorption and subsequent electrochemical reduction.

CO₂RR

CeO₂ loading. Fig. 6 illustrates the influence of CeO_{2-C} loading on a series of Cu₂O-based gas diffusion electrodes (GDEs). During the CO₂ electroreduction reaction in 1 M KOH, the main products detected across all GDEs were C₂H₄, CH₄, CO, and H₂. Within the evaluated range of current densities (-50 to -300 mA cm⁻²), the faradaic efficiencies for C₂H₄ production were consistently higher in the CeO_{2-C}-promoted Cu₂O-based GDEs compared to the non-promoted



Fig. 6 The faradaic efficiency [%] and potential [V vs. RHE] as a function of the current density [mA cm⁻²] over the series of Cu₂O-based GDEs.

 Cu_2O GDE. This indicates that the incorporation of the CeO_{2-C} promoter phase is beneficial for enhancing C_2H_4 formation during the CO_2RR .

The product formation on the CeO_{2-C} electrode starts at approximately -0.93 V vs. RHE, see Fig. 6. The achieved faradaic efficiencies (FEs) for C₂H₄ were lower compared to the FEs for other sub-products such as H₂, CO, and CH₄. The maximum FE_{C₂H₄} observed for this electrode was 0.6% at -1.34 V vs. RHE, and this was achieved at a high current density (j = -250 mA cm⁻²). In contrast, when the current densities exceeded -150 mA cm⁻², a noticeable increase in FE_{CO} (>75% at -1.33 V vs. RHE) and a decrease in FE_{CH4} (<3%at -1.33 V vs. RHE) were observed. Comparing with the CeO_{2-C} electrode, the Cu₂O electrode exhibited a maximum FE_{C2H4} of approximately 41% at -0.87 V vs. RHE, which was achieved at a current density of -200 mA cm⁻². The onset potential for C₂H₄ formation (approximately -0.52 V vs. RHE) was lower compared to the CeO_{2-C} electrode.

The observed behaviour of faradaic efficiencies (FEs) for the Cu_2O-CeO_{2-C} series indicates that a lower amount of the CeO_2 promoter is favoured to enhance C_2H_4 production. In this case, the potential at which C_2H_4 formation starts ranged between -0.52 and -0.67 V ν s. RHE. It is noteworthy that the maximum FEs for C_2H_4 were shifted towards lower current densities due to the increased CeO_2 loading. Furthermore, the FE_{CO} and FE_{CH4} increased, while FE_{H2} decreased with both increasing CeO₂ loading and current density. Among the Cu₂O-CeO_{2-C} electrodes, Cu₂O with 5 wt% of CeO_{2-C} was determined to be the optimum electrode for C_2H_4 production. It achieved a maximum FE_{C2H4} of approximately 53% at -0.98 V ν s. RHE, at a current density of -250 mA cm⁻².

To investigate the influence of the CeO₂ particle morphology, a CO₂RR test was conducted using rod-like ceria particles (CeO_{2-R}) with the optimized promoter content of 5 wt% (previously found for cubic-like ceria particles, CeO_{2-C}), under the same reaction conditions (1 M KOH, j = -50 to -300 mA cm⁻²). Fig. SI10 and SI11 (ESI[†]) revealed that the FEs for C₂H₄

formation are dependent on the shape of the promoter particles.

In the CeO_{2-R} promoter electrode, the main products observed were CO, H₂, CH₄, and C₂H₄, similar to the CeO_{2-C} promoter. However, compared to CeO_{2-C}, the CeO_{2-R} electrode required a more negative potential (-0.94 V vs. RHE) to initiate C₂H₄ formation. Furthermore, the maximum FE for C₂H₄ ($\sim 0.15\%$ at -1.26 V vs. RHE) over CeO_{2-R} was detected at a current density of -150 mA cm⁻². This suggests that the rodlike ceria morphology is less favourable for C₂H₄ production, as a higher potential was achieved at relatively low current densities. Regarding the FEs for other subproducts, CeO_{2-R} exhibited a lower FE for CO and higher FE for H₂ and CH₄ compared to CeO_{2-C}, as shown in Fig. SI10 (ESI†).

On the other hand, the combination of Cu₂O with CeO_{2-R} did not have a positive effect on enhancing the faradaic efficiency for C₂H₄, see Fig. SI11 (ESI[†]). The maximum FE_{C₂H₄} achieved over the Cu₂O–CeO_{2-R} electrode was only ~50% at -0.90 V *vs.* RHE. This low faradaic efficiency can be attributed to the insignificant promotion of CO₂ and H₂O activation by the presence of CeO_{2-R}, as shown in Fig. SI9 (ESI[†]). Therefore, the results suggest that CeO_{2-C} promotes more suitable activation of both CO₂ and H₂O molecules, which is crucial for ensuring the selective formation of C₂H₄. It should be noted that the optimum catalyst exhibits a low cathodic polarization compared to pristine Cu₂O (Fig. SI12, ESI[†]) attributed to the local CO formation at the CeO_{2-C} promoter and simultaneous CO₂ and CO electroreduction. Thus, the combination of Cu₂O with CeO_{2-C} demonstrates improved performance in terms of FE_{C₂H₄ compared to Cu₂O–CeO_{2-R}.}

Electrolyte. The influence of three different electrolytes, namely KOH, KCl, and KHCO₃, was investigated on the optimized Cu₂O-5CeO_{2-C} catalyst while maintaining a constant electrolyte concentration of 1 M and varying the current density $(j = -[50-300] \text{ mA cm}^{-2})$. In terms of C₂H₄ production, the KOH electrolyte exhibited higher C₂H₄ faradaic efficiencies (\geq 9% at -0.52 V vs. RHE) at all tested current densities, except for 50 mA cm⁻². The selectivity trend for C₂H₄ among the



Fig. 7 The faradaic efficiency [%] and potential [V vs. RHE] as a function of current density [mA cm⁻²] over the $Cu_2O-5CeO_{2-C}$ GDE using different electrolytes.

different electrolytes was observed as follows: KOH > KCl > KHCO₃, as shown in Fig. 7. As the current density increased, the preferential occurrence of the hydrogen evolution reaction was observed in the electrolytes based on KCl and KHCO₃. Overall, KHCO₃ exhibited the most negative potentials (> -1.10 V *vs.* RHE) compared to the other evaluated electrolytes. Therefore, the selection of the electrolyte is crucial for the CO₂ reduction reaction on Cu₂O-5CeO_{2-C} gas diffusion electrodes, with alkaline KOH conditions enabling preferential C₂H₄ production.

KOH concentration. As the electrolyte pH also plays an important role in the CO_2RR to C_2H_4 , the influence of the KOH concentration was evaluated by varying the current density $(j = -[50-300] \text{ mA cm}^{-2})$. Fig. 8 demonstrates that the electrochemical activity is highly dependent on the concentration of KOH. Notably, the lowest faradaic efficiencies for C_2H_4 production (\geq 18% at -0.60 V vs. RHE) were observed at the selected KOH concentration of 0.1 M. In contrast, there was an expected decline in the preference for the hydrogen evolution reaction with increasing concentration (>1 M). Similarly, at higher current densities (>250 mA cm^{-2}), the hydrogen evolution reaction became even less preferred. Concerning the potentials, more negative values (-1.05 V vs. RHE) were achieved with the increase of the KOH concentration. Particularly, CO and CH₄ were the main products promoted at those high negative potential values.

Stability test. The stability of Cu₂O–5CeO_{2-C} and Cu₂O catalysts shown in Fig. 9 was used to evaluate the effect of the promoter phase. The tests were conducted under the optimized reaction conditions of j = -250 mA cm⁻² and 1 M KOH electrolyte, which were found to yield the most promising results for the CeO_{2-C}-promoted Cu₂O gas diffusion electrode (GDE). Under these selected conditions, the Cu₂O–5CeO_{2-C} electrode exhibited a higher FE_{C₂H₄ (~53% at -0.98 V *vs.* RHE) and maintained its stability over a period of 120 minutes. The faradaic efficiencies for other byproducts (H₂, CO, and CH₄) remained constant throughout the reaction (see Fig. SI13, ESI†).}

In contrast, when using non-promoted Cu_2O , the initial FE_{C,H_4} reached a maximum value of approximately 38% at



Fig. 9 C_2H_4 FE [%] as a function of the time [min] over the Cu₂O-5CeO₂. _C GDE. Reaction conditions: electrolyte = 1 M KOH and j = -250 mA cm⁻²

-1.05 V vs. RHE but experienced partial deactivation after 45 minutes. The FE for C₂H₄ on the Cu₂O electrode decreased by approximately 5%. Notably, the FEs for CO and H₂ increased as the FE for C₂H₄ and CH₄ decreased on Cu₂O (see Fig. SI14, ESI†).

Consistently with these experimental findings, the 5 wt% CeO_{2-C} -promoted Cu_2O demonstrated enhanced performance compared to the non-promoted Cu_2O electrode in terms of ethylene production, aligning with the results obtained from linear sweep voltammetry (see Fig. SI15, ESI†). The Cu_2O - $5CeO_{2-C}$ electrode exhibited a more positive onset potential than Cu_2O , indicating that it was the most active catalyst for the CO_2RR .

XRD measurements were also conducted on the Cu₂O and Cu₂O–5CeO_{2-C} gas diffusion electrodes at different reaction times (t = 0 (before reaction: fresh), 5 and 120 min (after reaction: used)) to investigate the surface copper species during the stability test. In both the fresh and used electrodes, reflections at 18.15, 25.67, and 52.74° were attributed to the carbon phase (JCPDS: 00-047-0787 and 00-026-1076) of the carbon-paper support (see Fig. SI16, ESI[†]). Prior to sample activation (t = 0 min), characteristic reflections associated with the Cu₂O



Fig. 8 The faradaic efficiency [%] and potential [V vs. RHE] as a function of the current density [mA cm⁻²] over the $Cu_2O-5CeO_{2-C}$ GDE using different KOH concentrations.



Fig. 10 XRD patterns of the Cu₂O and Cu₂O-5CeO₂ GDEs at the reaction time of 120 minutes.

phase were observed in both Cu₂O (see Fig. SI17, ESI†) and Cu₂O–5CeO_{2-C} (Fig. SI18, ESI†) electrodes. Additionally, CeO₂ was detected on the Cu₂O–5CeO_{2-C} electrode.

After a reaction time of 5 min, the intensity of reflections corresponding to Cu_2O decreased in both gas diffusion electrodes, and two new phases, CuO (JCPDS: 01-078-0428) and metallic Cu^0 (JCPDS: 00-004-0836), were identified. A similar behaviour was observed for the intensity of the reflections related to the CeO_2 phase in the Cu_2O -5 CeO_{2-C} electrode. However, two new phases, Cu_4O_3 (JCPDS: 00-033-0480) and K₃CuO₂ (JCPDS: 00-038-0971), were specifically identified in the Cu_2O electrode.

At a total reaction time of 120 min (see Fig. 10), the reflections corresponding to CuO phases disappeared, and the intensity of reflections related to Cu₂O and Cu⁰ decreased. Additionally, new reflections were observed in the Cu₂O electrode, which were associated with the KCuO (JCPDS: 01-076-2437) and KO₂ (JCPDS: 01-084-1972) phases. In contrast, the reflections of Cu₂O-5CeO_{2-C} remained like those identified at 5 min, indicating that CeO₂ promoted the stability of the main copper species (Cu₂O, CuO, and Cu⁰) involved in the CO₂RR. These identified copper species were consistent with those detected through cyclic voltammetry measurements (see Fig. SI19, ESI†).

According to XRD analysis, the crystallite size of the main $Cu^x = Cu^{2+}$, Cu^+ , and C^0 species were modified after the reaction, see Table SI1 (ESI[†]). The increase of the Cu₂O crystallite size for the non-promoted Cu₂O catalyst was 15%, while for the Cu₂O-5CeO_{2-C} it was 1%. For Cu₂O-5CeO_{2-C}, the CeO₂ crystallite sizes were also slightly increased by 8%. Furthermore, the crystallite size corresponding to the Cu phase (20 nm) was higher for the Cu₂O sample. Therefore, the poor stability of the Cu₂O catalyst can result in the increase of the crystallite sizes of the main copper species.

Additionally, SEM-EDS analysis was performed after the reaction to verify the catalyst composition, see Table SI2 (ESI†). Besides the main elements (Cu, Ce, and O), the potassium (K) phase was also detected on the surface of both used catalysts. However, a clear modification of the catalyst composition was identified for used Cu₂O. In this catalyst, the Cu phase decreased by 55% compared to its fresh estate. This behaviour



Fig. 11 CO_2RR to ethylene over $Cu_2O-5CeO_2$ GDE. (a) Plausible mechanism of CO_2RR to C_2H_4 . (b) Performance comparison with similar catalytic system reported in the literature. 0.1 M KHCO₃ was used for CuO-30 wt% CeO₂/CB³¹ and Cu/CeO₂(100),³¹ while 1 M KOH was used for CeO₂/50 wt% CuO,³³ Ce-Cu NPs,³² and Cu₂O-5 wt% CeO₂ (this work).

can be related to the presence of the new KOx and KCuOx phases (see Fig. SI20, ESI†), in agreement with post-catalysis XRD results. Regarding the used Cu₂O–5CeO₂, its composition was slightly modified, suggesting that CeO_{2-C} (see Fig. SI21, ESI†) inhibits the formation of potassium species, which are negative for ethylene production.

Based on the electrochemical and catalyst characterization conducted before and after the reaction, it can be concluded that the interaction between Cu₂O and CeO₂ creates intrinsic sites (Cu^{x} -CeO_{2-x}; $Cu^{x} = Cu^{2+}$, Cu^{+} , Cu^{0}) for the binding of CO₂ and H_2O . The proposed mechanism, illustrated in Fig. 11(a), suggests that CO_2 can be adsorbed around the Cu^x -CeO_{2-x} interface and reduced to *CO on the CeO_{2-x} site. This step is crucial as CO serves as the key intermediate for the production of C₂₊ products.^{30,33} Subsequently, *CO can undergo further reduction to C₂H₄ through enhanced *CO-*CO coupling, which occurs on $Cu^{x} = Cu^{2+}, Cu^{+}, and Cu^{0}$ sites, with the assistance of H⁺ species derived from adsorbed H_2O on Cu^+ -Ce O_{2-x} sites.^{30,54} Therefore, the improved and stable FE_{C.H.} achieved with the Cu₂O-5CeO_{2-C} catalyst supports the assertion that CeO_{2-C} plays a significant role in the CO₂RR. It promotes the formation of Cu^{x} -CeO_{2-x} sites, which govern the activity, selectivity, and stability of C₂H₄ production by synergistically activating CO₂ and H₂O molecules.

The performance comparison with catalysts shown in Fig. 11(b), in terms of faradaic efficiency (FE) and partial current density, indicates that Cu₂O-5CeO_{2-C} is a promising catalytic system for C₂H₄ production. Compared to other reported catalytic systems, Cu₂O-5CeO_{2-C} demonstrates a high FE for C₂H₄ (FE_{C₂H₄} ~ 53%) at a high partial current density ($j_{C_2H_4} = 132 \text{ mA cm}^{-2}$). Additionally, one of the most significant advantages of Cu₂O-5CeO_{2-C} is that it requires a lower potential (-0.98 V *vs.* RHE) than the reported CeO₂/50 wt% CuO system (-2.3 V *vs.* RHE) to operate at the same electrolyte concentration of 1 M KOH.³³

Conclusions

This study investigated the influence of CeO₂ promotion on the catalytic performance of Cu₂O-based catalysts for the electrochemical reduction of CO₂. The addition of the CeO_{2-C} promoter phase was found to enhance the formation of C₂H₄, a valuable commercial product, during the CO₂RR compared to the non-promoted Cu₂O catalyst. The characterization techniques, including SEM, TEM, XRD, Raman spectroscopy, and Xray photoelectron spectroscopy, provided valuable insights into the structural and compositional properties of the catalysts.

The results indicated that the synthesized catalysts exhibited well-defined morphologies, with cubic-like particles for Cu_2O and CeO_{2-C} , and rod-like morphologies for CeO_{2-R} . The XRD analysis confirmed the phase of Cu_2O and CeO_2 , while Raman spectroscopy revealed characteristic peaks for both materials and a new peak indicating the interaction between Cu^+ and Ce^{3+} cations. X-ray photoelectron spectroscopy further confirmed the presence of different copper oxidation states in the Cu_2O -based catalysts.

The electrochemical evaluation of the catalysts demonstrated that the CeO_{2-C}-promoted Cu₂O-based gas diffusion electrodes (GDEs) exhibited higher faradaic efficiencies for C₂H₄ production compared to the non-promoted Cu₂O GDE. The optimal loading of CeO_{2-C} for maximum C₂H₄ production was determined to be 5 wt%. The proposed Cu₂O–5CeO_{2-C} GDE demonstrates enhanced and stable C₂H₄ production through the interaction of Cu₂O and CeO₂, creating active sites (Cu^x– CeO_{2-x}; Cu^x = Cu²⁺, Cu⁺, Cu⁰) for the CO₂RR. It exhibits high faradaic efficiency (FE_{C₂H₄ ~ 53%) and partial current density ($j_{C_2H_4}$ = 132 mA cm⁻²) at a low potential (-0.98 V *vs.* RHE) with 1 M KOH electrolyte, making it a promising catalyst. As an important fact, the FEs for C₂H₄ in the Cu₂O–5CeO_{2-C} GDE were found to be influenced by the type of electrolyte and its concentration.}

These findings highlight the potential of CeO_2 promotion in improving the selectivity and activity of Cu_2O -based catalysts for the CO_2RR . The study contributes to the understanding of the structure–property relationship in CO_2 electrochemical conversion and provides a foundation for further development of more efficient catalysts.

Conflicts of interest

There are no conflicts to declare.

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